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Flexible thermally treated 3D PIM-CD molecular sieve membranes exceeding the upper bound line for propylene/propane separation

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Polymers of intrinsic microporosity (PIM) incorporated with beta-cyclodextrin (β -CD) (referred to as PIM-CD) are prepared *via* nucleophilic substitution copolymerization and then thermally treated at elevated temperatures from 300 to 600 °C. After the decomposition of thermally labile CD, the spaces originally occupied by CD convert to micro-pores and cross-linking points in the polymer matrix. The thermally induced three-dimensional (3D) cross-linking network not only possesses ultra-fine micro-pores and interconnected microvoids but also shows a superior molecular sieve ability for propylene/propane separation. The C_3H_6/C_3H_8 selectivity of the thermally treated membranes increases remarkably, about 3 times compared to that of the original untreated membranes. In particular, the PIM-CD membranes thermally treated at 300 or 400 °C are flexible, and they have C_3H_6/C_3H_8 separation performance exceeding the upper bound line in both pure and mixed gas tests. As observed from the gas sorption isotherms, membranes with a higher content of CD units have a higher gas sorption capacity. Additionally, the diffusivity selectivity of PIM-CD membranes after thermal treatment increases more quickly than that of the PIM membrane and contributes more to the permeability selectivity.

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1. Introduction

Since propylene is the second largest feedstock for the production of commonly used plastics (*e.g.* polypropylene) and chemicals (*e.g.* isopropanol, acrylonitrile, and propylene oxide), the separation of propylene from propylene/propane mixtures is one of the most important processes in the petrochemical industry.¹ The current dominant process for C_3H_6 and C_3H_8 separation, cryogenic distillation, is extremely energy intensive due to their close boiling points, high volatility and condensability characteristics.^{2–5} Membrane technology,^{3–7} which possesses advantages of low energy consumption, easy operation and maintenance, small footprint and operational cost, and environmental friendliness, is one of the potential strategy to debottleneck the traditional separation process.

Until now, many kinds of membrane materials have been tried for propylene/propane separation, such as (1) pure polymeric membranes, (2) metal ion doped membranes, (3) mixed matrix membranes, and (4) carbon molecular sieve (CMS) membranes. Merits coexist with drawbacks in each kind of membrane material. Polymeric membranes are relatively inexpensive and easy to scale up, but their C_3H_6/C_3H_8 separation performance often suffers from the trade-off relationship between permeability and selectivity.^{1,3–5,7–15} For example, polyimide membranes prepared from Matrimid® 5218 have a high

C_3H_6/C_3H_8 selectivity of 10 but an extremely low C_3H_6 permeability of 0.1 barrer.³ As a result, it's a challenge to design polymeric membranes with both high permeability and good selectivity simultaneously.

Metal ion doped membranes often possess good C_3H_6/C_3H_8 separation performance due to specific interactions between the olefin's hybrid molecular orbitals and the metal's atomic orbitals.^{5,16–21} However, metal ion doped membranes also face the inherit problems of short life time and carrier poisoning induced by trace amounts of methyl acetylene and sulfur impurities in the C_3H_6/C_3H_8 feed mixture.²² Mixed matrix membranes consisting of MOF or ZIF nanoparticles in polymer matrices are promising materials because they combine the merits of both polymer and molecular sieving characteristics.^{9,23–27} However, poor compatibility between nanoparticles and polymer matrices often results in particle agglomeration and defective membranes.

Carbon molecular sieve (CMS) membranes have received considerable attention because of their good thermal and chemical stability and superior gas separation performance.^{3,28–33} However, CMS membranes suffer from the drawbacks of brittleness and high energy consumption and costs. Until now, no CMS membranes derived from the PIM precursor are reported for C_3H_6/C_3H_8 separation even though some data were available for C_2H_4/C_2H_6 separation.^{34,35} Therefore, we aim to explore if we can develop flexible thermally treated PIM membranes for C_3 separation with performance exceeding the upper bound of the trade-off relationship.

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Polymers of intrinsic microporosity (PIM) were chosen in this study because they possess extremely high free volume and surface areas.^{36–40} The ladder-type structure and rigidly contorted spiro-carbon sites can prevent polymer chains from becoming densely packed. However, the linear PIM polymer chains interact with each other by relatively weak van der Waals' forces which may easily slide over one another. To enhance inter-chain interactions and improve molecular sieve characteristics, Liu *et al.* incorporated β -cyclodextrin (β -CD) with a unique 3D hollow bowl structure into PIM chains (referred to as PIM-CD).⁴¹ The incorporation of a very small amount of β -CD into PIM significantly enhanced the gas permeability that surpassed the Robeson upper bound without much compromising the selectivity. The PIM-CD membrane also exhibited a greater resistance to physical aging than the PIM membrane.

Since β -CD is a thermally labile molecule containing multiple reactive hydroxyl groups, one may convert the highly permeable PIM-CD precursor to covalently cross-linked polymer networks with significantly improved molecular-sieving properties by thermal annealing. The thermally labile CD units may decompose and create microvoids as well as generate free radicals, which can effectively accelerate the self-cross-linking reaction within the rigidly contorted PIM polymer framework. As a result, the annealed PIM-CD membrane has not only a large free volume for gas transport but also a tighter structure to separate C_3H_6/C_3H_8 mixtures. Therefore, various annealing temperatures from 300 to 600 °C would be used to investigate their effects on the chemical structure and gas separation performance of PIM-CD membranes under both pure gas and mixed gas testing environments. These results may open up new avenues for molecular designing of highly permeable membranes for propylene/propane separation.

2. Experimental

2.1 Materials

β -Cyclodextrin (β -CD) was purchased from Cyclolab Ltd and dried overnight at 100 °C under vacuum prior to use. The molecular weight of the dehydrated β -CD (formula: $(C_6H_{10}O_5)_7$) is 1134. The outer diameter of the larger opening of the β -CD toroid is 15.3 Å, the side rim depth is 7.8 Å and the inner cavity diameter of β -CD is 6.0–7.8 Å; therefore, CD was usually used in separation science and technology because of its unique hollow bowl structure.^{41,42} 2,3,5,6-tetrafluoroterephthalonitrile (TFTPN, 99%) was acquired from Matrix Scientific and purified by sublimation under vacuum prior to syntheses. The monomer 5,5',6,6'-tetrahydroxy-3,3,3',3'-tetramethyl-1,1'-spirobisindane (TTSBI, 97%) was ordered from Alfa Aesar and purified *via* recrystallization from methanol. The spirocentre within TTSBI is a kinked and shape-persistent ladder structure with a dihedral angle of 90°. *N,N'*-Dimethylformamide (DMF, >99.5%) from Merck (Germany) was further purified *via* vacuum distillation at 65 °C before being utilized as the solvent for the polymer synthesis. Anhydrous potassium carbonate (K_2CO_3 , >99.5%) from Sigma Aldrich, methanol (MeOH, $\geq 99.9\%$) from Merck and dichloromethane (DCM, 99.99%) and hydrochloric acid (HCl, 37.5%) from Fisher Scientific (UK) were used as received.

All purified gases (at least 99.95%), such as C_3H_6 , C_3H_8 , and mixed C_3H_6/C_3H_8 (50%/50%) were supplied by SOXAL Pte. Ltd. (Singapore).

2.2 Syntheses of polymers and fabrication of membranes

The typical synthesis procedure of PIM-CD polymers and the fabrication of PIM-CD membranes were carried out according to a previously reported procedure.⁴¹ PIM-CD-1% means the molar ratio content of CD is 1 percent in copolymerization. The average membrane thickness was measured from 10 different points using a Digimatic indicator (IDC-112B-5) with an accuracy of 1 μ m. The cast dense membranes with a thickness of about $50 \pm 5 \mu$ m were used for further studies.

2.3 Thermal treatment of PIM-CD membranes

The thermal treatment of PIM and PIM-CD membranes was carried out using a VBF-1200X vacuum furnace from MTI Corporation. The oxygen concentration inside the furnace was measured with an EQ-W3000 Trace Oxygen Analyzer and controlled at about 10 ppm at all times. The vacuum furnace temperature was increased to 100 °C at a rate of 2 °C min^{-1} and held for a period of 1 h to dry the membranes. Subsequently the temperature was increased to a set value at a slow rate of 1 °C min^{-1} in order to prevent the membranes from cracking. Finally, the temperature was held isothermally at the set temperature for 2 h. After the thermal treatment, the membranes were cooled down to room temperature naturally in the vacuum furnace and stored in a dry box for further studies. The thermally treated membranes were labeled as "PIM-CD-content (temperature)", for example, PIM-CD-1% (300 °C). A clear coloration from yellow to dark brown and black colors could be observed for the thermally treated membranes depending on the annealing temperature. However, the membranes thermally treated at 700 °C were extremely fragile, hence leading to enormous difficulties in characterization and sample preparation for permeation tests.

2.4 Characterization

The chemical structures of pristine membranes and the thermally treated membranes were analyzed using a Shimadzu Fourier transform infrared spectroscopy (FTIR) 8400 spectrometer in the range of 400–4000 cm^{-1} . The thermogravimetric analyses (TGAs) of pristine membranes were performed by using a Shimadzu Thermal Analyzer (DTG-60AH/TA-60WS/FC-60A) with a heating rate of 10 °C min^{-1} from 50 °C to 800 °C under a nitrogen atmosphere with a flow rate of 50 mL min^{-1} . The TGA outlet was coupled on-line with a Perkin-Elmer FTIR Spectra 2000 spectrometer. Thus, the exhausted products from the TGA were analyzed in a gas cell and the IR spectra were obtained by using a KBr beam splitter. The decomposition process of PIM-CD membranes was recorded over the range of 50–800 °C.

The insoluble content of the thermally treated membranes was determined by soaking the respective membranes in a DCM solvent for 24 h. The remaining insoluble portions of the membranes were dried under vacuum at 120 °C for 24 h to



remove the residual solvent before weighing. The insoluble content was evaluated using eqn (1):

$$\% \text{ insoluble content} = M_1/M_0 \times 100 \quad (1)$$

where M_0 and M_1 are the weights of the thermally treated membrane before and after soaking in DCM, respectively.

The inter-chain spacing of the membranes was investigated by using a Bruker D8 Advance X-ray diffractometer (XRD) utilizing Cu K- α as the X-ray radiation source with a wavelength of 1.54 Å. The X-ray diffraction angles between 5 and 30 were studied. The average d -spacing was evaluated based on Bragg's law ($n\lambda = 2d \sin \theta$), where n is an integer (1, 2, 3,...), l represents the X-ray wavelength, d represents the intersegmental spacing between two polymer chains and θ denotes the X-ray diffraction angle.

2.5 Measurements of gas permeation

The gas permeation properties for both pure gases and mixed gases were determined on a variable-pressure constant-volume gas permeation cell. The detailed experimental design and procedures have been reported by our group.^{41,46} The dense membrane was mounted onto the permeation cell and vacuumed at 35 °C for 12 h before tests. Pure gas permeability was tested following the order of C₃H₆ and C₃H₈. The cell temperature was kept constant at 35 °C. The permeability tests were repeated at least three times for three different samples prepared under the same conditions and the average deviation obtained was less than 5%.

The ideal selectivity between two different gases across a polymeric membrane is the ratio of their single gas permeability as described in the following eqn (2):

$$\alpha_{A/B} = \frac{P_A}{P_B} \quad (2)$$

where P_A and P_B refer to the permeability coefficients of gases A and B, respectively. According to the solution-diffusion model, permeability is the product of diffusivity (D) and solubility (S). Therefore, the ideal permeability selectivity (α_P) can be defined as the product of diffusivity selectivity (α_D) and solubility selectivity (α_S) as described in eqn (3) and (4):

$$P = D \times S \quad (3)$$

$$\alpha_{A/B} = \alpha_P = \alpha_D \times \alpha_S = \frac{D_A}{D_B} \times \frac{S_A}{S_B} \quad (4)$$

where D_A and D_B are the diffusion coefficients of gases A and B, respectively ($\text{cm}^2 \text{s}^{-1}$); S_A and S_B are the solubility coefficients of gases A and B, respectively ($\text{cm}^3 \text{ (STP) cm}^{-3} \text{ polymer atm}$). In the case of mixed gas experiments, a binary mixture of 50% C₃H₆ and 50% C₃H₈ was used. The membranes were tested at 35 °C with a feed pressure of 3, 5 and 7 atm. The mixed gas composition was analyzed by using a Hewlett Packard (HP) 6890 Series gas chromatograph (GC) integrated with a HP 5973 mass selective detector.

2.6 Measurements of gas sorption

Pure gas sorption of the thermally treated membranes was measured *via* the dual-volume pressure decay method using a XEMIS microbalance system.⁴¹ The membranes weighing about 50–60 mg were placed into the sample chamber and a vacuum was maintained for 3 h before sorption tests were conducted. Both the reference and sample chambers were immersed in the same water bath ((Techne FTE-10DE) from Techne Cambridge LTD (UK) with an accuracy of ± 0.01 °C) to maintain the desired temperature. The sorption isotherm of each gas was obtained at 35 °C up to a maximum pressure of 5 atm. Sorption tests were conducted for C₃H₆ and C₃H₈ pure gases.

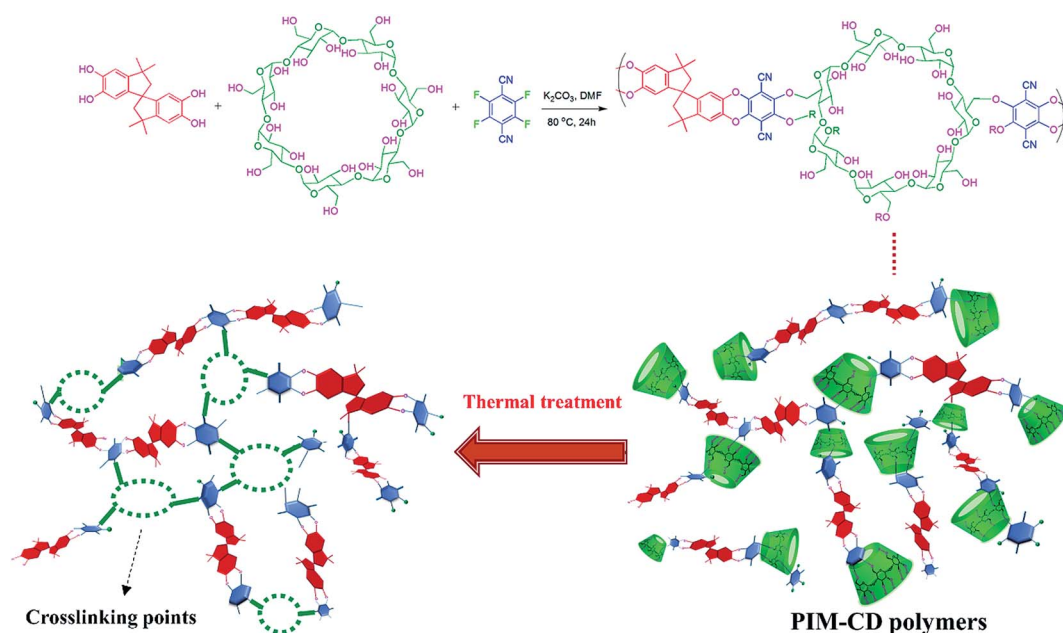


Fig. 1 Synthesis and thermal crosslinking schemes of PIM-CD polymers.



3. Results and discussion

3.1 Characterization of the polymer membranes

As presented in Fig. 1, the 3D PIM-CD polymer precursors were synthesized *via* nucleophilic substitution copolymerization of β -CD, TFTP, and TTSBI. Since β -CD comprises multiple reactive hydroxyl groups, it can be copolymerized into the main chains of PIM easily. The copolymerization was conducted in a polar aprotic anhydrous solvent of DMF at 80 °C for 24 h, using K_2CO_3 as the catalyst. The chemical structures of the original PIM and PIM-CD polymers were characterized and confirmed in our previous report.⁴¹ The thermal stability of the original PIM and PIM-CD membranes was evaluated by TGA. Fig. 2 shows their

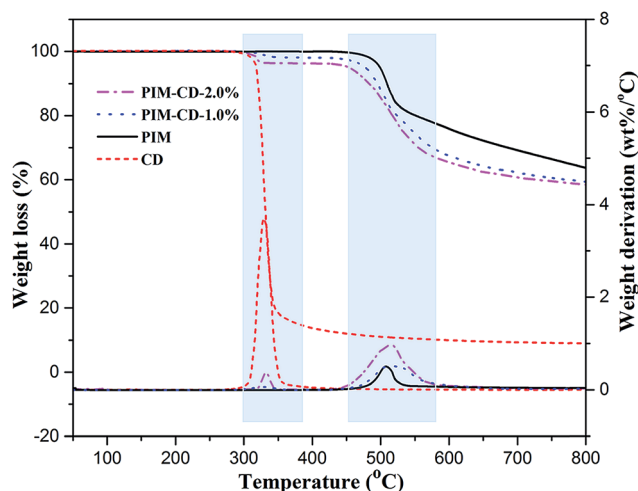


Fig. 2 TGA spectra of β -CD, pristine PIM and PIM-CD membranes.

thermal decomposition and derivatives of weight losses (*e.g.* decomposition rates) as a function of temperature in a nitrogen atmosphere. Pure β -CD is also included in TGA tests for comparison; it decomposes from about 300 to 400 °C. Compared to the pristine PIM membrane, the PIM-CD membranes exhibit a little lower thermal stability due to the weight loss of β -CD in their polymer chains. There are two stages of thermal degradation for PIM-CD membranes. The first stage of weight loss starts from 300 °C to 400 °C and the amount of weight loss increases with increasing CD loading, while the second stage of weight loss from 450 °C is attributed to the decomposition of the PIM polymer chains.

Fig. 3 presents the simultaneous 2D FTIR spectra of pyrolysis gases evolved from the membranes during TGA tests. Three major fragments are found in the decomposition of pure β -CD due to the rupture initiated from the C-O and C-C bonds in glucose units of β -CD. The release of compounds at about 3600, 2350 and 1750 cm^{-1} may be characteristic bands of OH, CO_2 and C_xH_yO (aldehydes or ketones), respectively. In addition, there is no other IR peak during the decomposition of pure β -CD. Only the CD structure decomposition contributes to the weight loss of the PIM-CD membranes during the first stage of weight loss from 300 to 400 °C. Compared with pure β -CD, the PIM-CD-1% and PIM-CD-2% have new characteristic peaks at about 3050 (C-H) and 1500 cm^{-1} ($C\equiv N$). These C-H fragments arise mainly from the decomposition of methyl groups, while the $C\equiv N$ fragments come from nitrile groups of PIM in the range of 400–600 °C. Clearly, the β -CD will decompose to form microvoids and cross-linking points when the membranes are treated at 300–400 °C, while the PIM polymer chains will decompose or partially carbonize when the membranes are treated at 500–600 °C. As shown in Fig. 4, after thermal

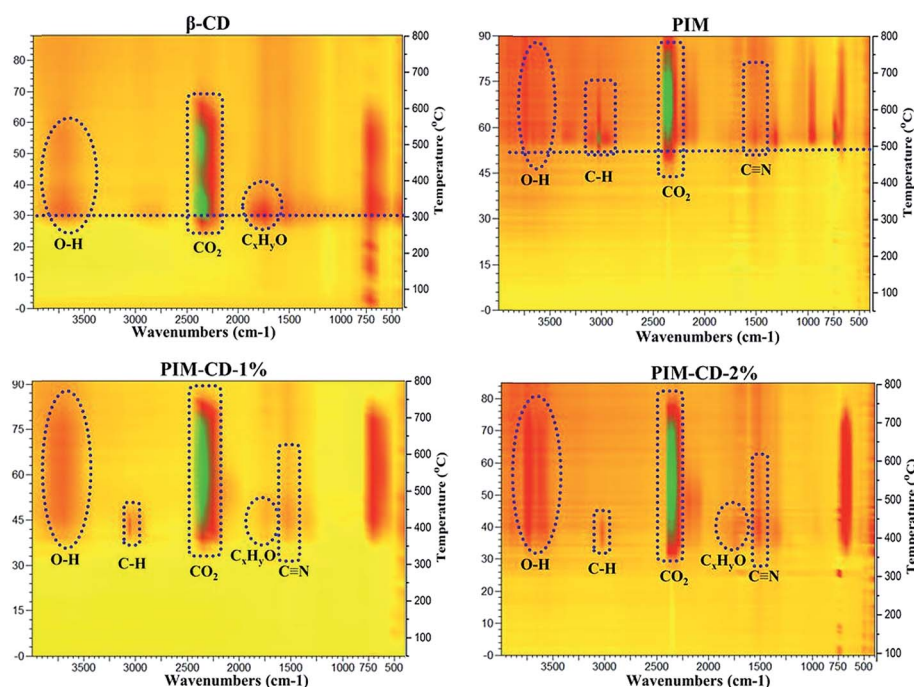


Fig. 3 FTIR-TGA spectra of β -CD, pristine PIM and PIM-CD membranes.



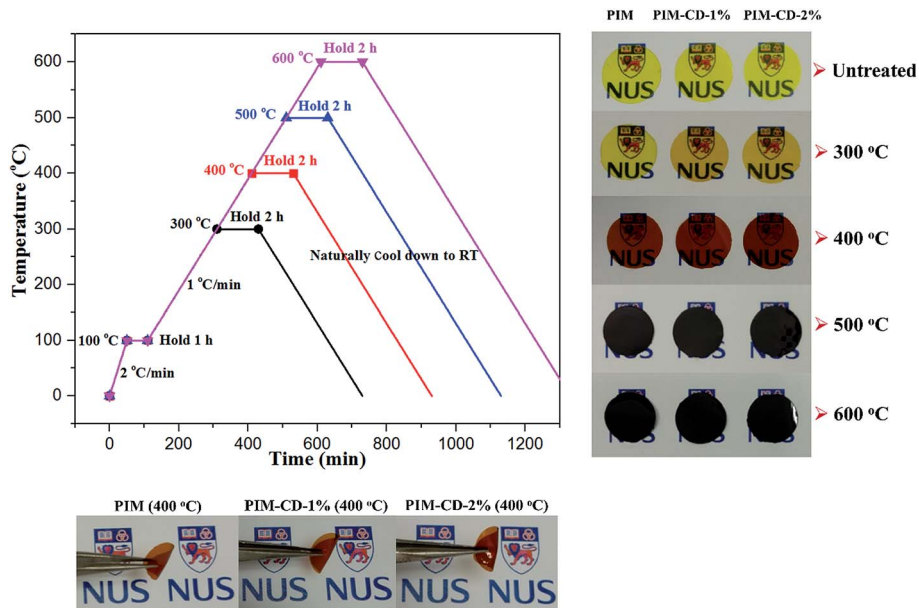


Fig. 4 Photographs and thermal treatment procedures of PIM-CD membranes.

treatments, the transparent membranes change from the original fluorescent yellow to dark brown at 400 °C and then to deep black at 500 °C. The PIM-CD membranes treated at 300 °C are slightly darker than the PIM membrane because of little more cross-linking points induced by β -CD and more formation of inter-chain charge transfer complexes (CTCs). All membranes thermally treated at and below 400 °C are flexible and can be bended tightly, while the membranes thermally treated at and above 500 °C are fragile.

The chemical structure changes and the formation of triazine after thermal treatment were confirmed by FTIR-ATR and XPS analyses. As illustrated in Fig. 5, the characteristic absorption at 2239 cm^{-1} is assigned to the stretching of $\text{C}\equiv\text{N}$

groups. The broad band at 3400 cm^{-1} is attributed to the O–H groups in PIM-CD polymers. The intensity of both $\text{C}\equiv\text{N}$ and O–H peaks decrease dramatically with increasing the annealing temperature. In contrast, the intensity at 1495 cm^{-1} attributed to the triazine ring slightly increases due to the thermal self-cross-linking of $\text{C}\equiv\text{N}$ groups.³⁸ Additionally, the $\text{C}=\text{C}$ bonds persist when the thermal treatment temperature is lower than the decomposition temperature of PIM polymer chains, while the $\text{C}=\text{N}$ bonds are gradually generated from the thermal self-cross-linking of $\text{C}\equiv\text{N}$ groups. Since the characteristic peaks of both the $\text{C}=\text{C}$ stretching of aromatic rings and the $\text{C}=\text{N}$ stretching of the triazine ring fall in approximately the same position 1610 cm^{-1} , the intensity of the absorption at 1610 cm^{-1} increases a little with increasing the thermal treatment temperature. There are no obvious peaks of the membranes thermally treated at 500 or 600 °C, probably because of the weak IR reflection of dark black membranes due to partial carbonization and formation of carbon molecular sieves.

Besides, XPS analyses were also applied to confirm the thermal cross-linking process and the generation of the triazine ring. Fig. 6 depicts the XPS spectra of the original and thermally treated membranes. The original PIM-CD-1% membrane presents a symmetrical Gaussian-type N_{1s} core signal with the maximum peak at 398.3 eV which is ascribed to the $\text{C}\equiv\text{N}$ (nitrile) groups of the PIM polymer chains. After thermal treatment, the symmetrical N_{1s} signal becomes broader and asymmetrical with a newly formed shoulder peak at about 399.5 eV, which is assigned to the nitrogen of $\text{C}=\text{N}$ groups in the thermally self-cross-linked triazine ring. In order to quantitatively detect the degree of the thermally self-cross-linking reaction, the N_{1s} XPS spectrum was deconvoluted into two peaks at 398.3 and 399.5 eV. In addition, the highest peak at 402.5 is probably due to the π -excitation which appears in all carbon compounds with double bonds. The ratio of integrated

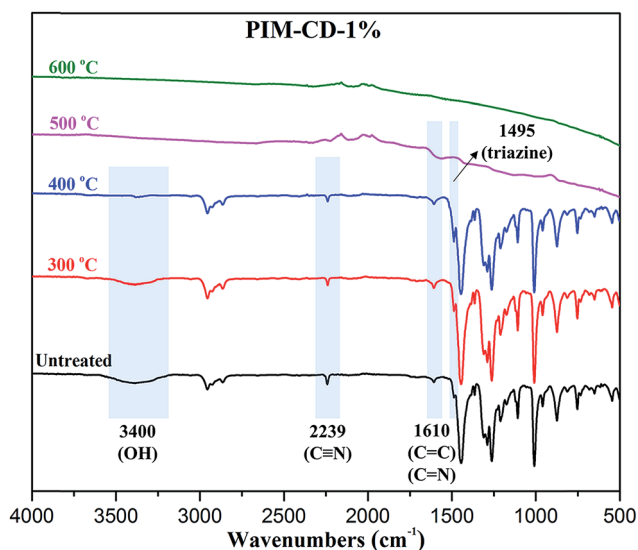


Fig. 5 FTIR-ATR spectra of the membranes after thermal treatment.



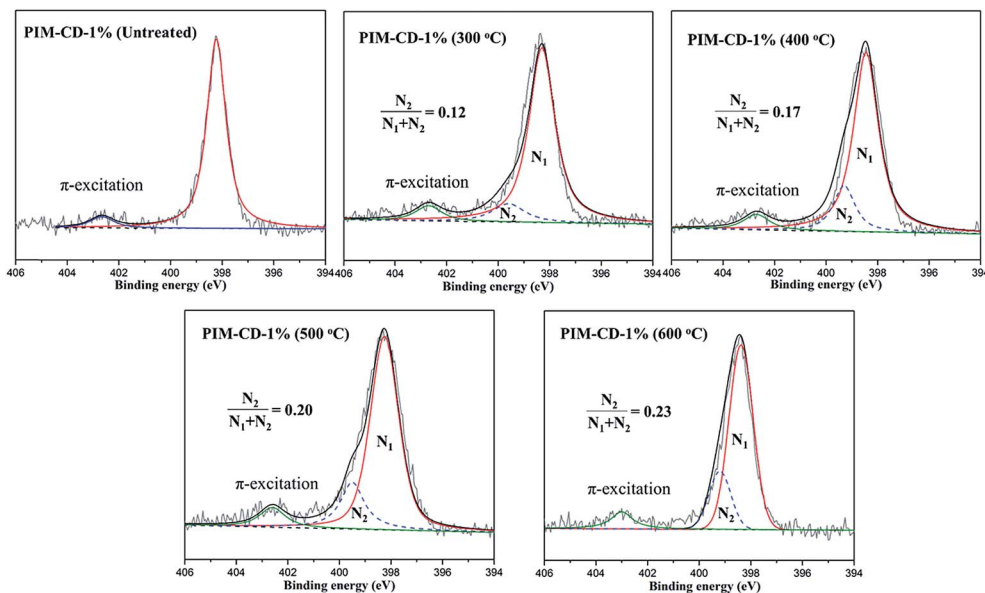


Fig. 6 N_{1s} XPS analyses of the untreated and thermally treated PIM-CD-1% membranes.

area $C=N/C\equiv N + C=N$ (*i.e.*, $N_2/(N_1 + N_2)$) shows an increasing trend from 0.12 to 0.23 with an increase in the thermal treatment temperature. This suggests that the content of the thermal cross-linking triazine ring is more pronounced at a higher temperature. The direct evidence for the inter-chain cross-linking reaction is the color change of the membrane from the original fluorescent yellow to dark brown and then to deep black. During the thermal treatment process, β -CD units first decompose to generate free radicals, which can effectively accelerate the self-cross-linking reaction of nitrile. The

formation of triazine rings tends to pull the PIM polymer chains closely together and allows the transfer of π -electrons from the electron donors (*i.e.*, ether groups) to the electron acceptor (*i.e.*, triazine ring). Thus, charge-transfer complexes between the polymer chains are formed that generates the dark colors.

A direct indication of the occurrence and the extension of the self-cross-linking reaction at different treatment temperatures is provided by the insoluble content of the membranes. As observed from Fig. 7, the degree of the self-cross-linking reaction increases with increasing the thermal treatment

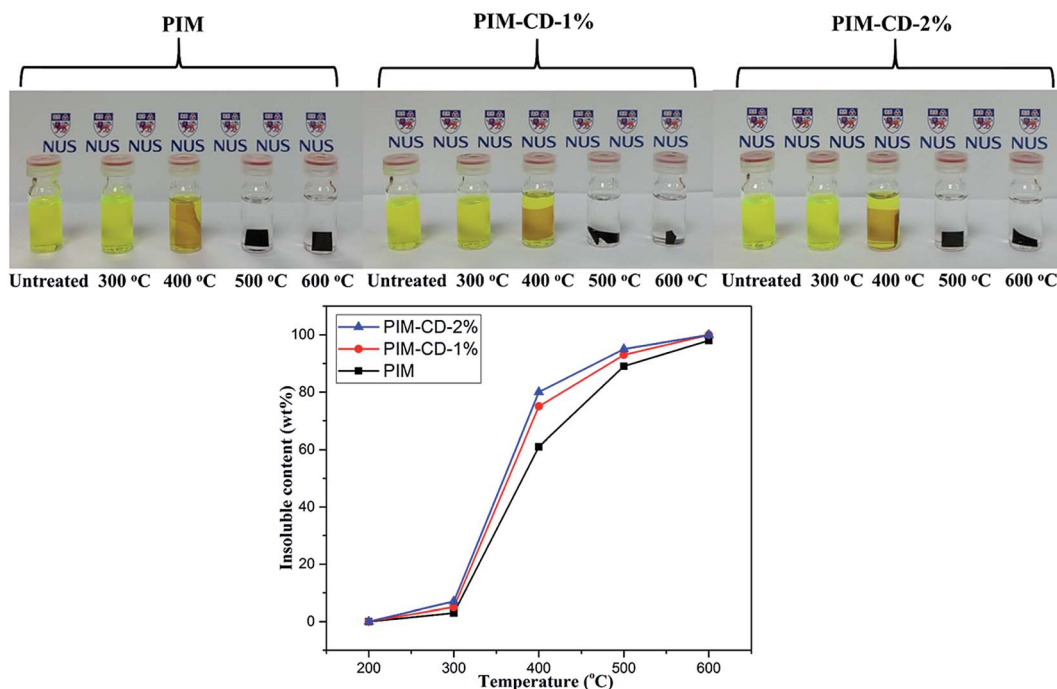


Fig. 7 The insoluble content of membranes treated at different temperatures and then soaked in DCM for 24 hours.



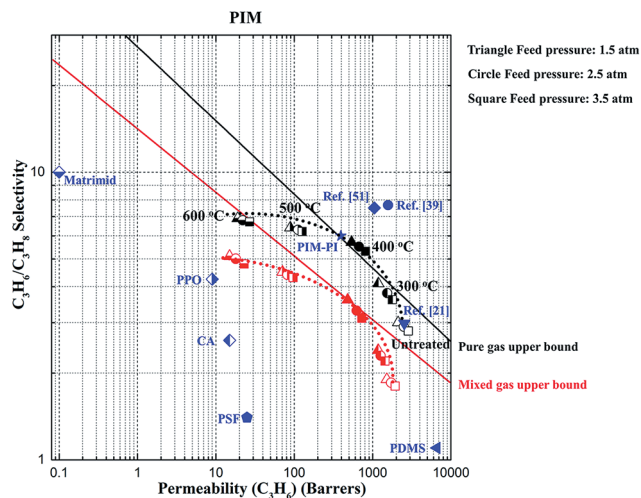


Fig. 8 Comparison of the pure and mixed gas performance of the thermally treated PIM membranes at 35 °C (lines are drawn to guide the eyes).

temperature. Clearly, there is more than 89% insoluble content for the membranes treated above 500 °C, the thermally cross-linked membranes are largely insoluble in DCM even after 24 hour immersion. Additionally, the insoluble content of PIM-CD membranes is higher than those of pristine PIM membranes, and the free radicals generated from β -CD decomposition can effectively facilitate the self-cross-linking reaction of nitrile.

3.2 Pure and mixed gas separation performance

Fig. 8 plots the pure and mixed gas separation performance against their corresponding upper bound lines.³ The previously

reported data of PIM-1, PIM-PI and commercial polymers for propylene/propane separation are also included in Table 2 and the Robeson plot. The gas separation performance of the thermally treated PIM-CD membranes is better than those of commercial polymeric membranes.^{43–45} In this study, for both pure and mixed gas tests, the thermally treated membranes have a higher selectivity than the original untreated ones, while the C_3H_6 permeability measured in mixed gases is always lower than that in pure gas because of competitive sorption of C_3H_6 and C_3H_8 in the membranes. Except the PIM membrane annealed at 400 °C, no other annealed PIM membranes can surpass the upper bound lines in either pure or mixed gas tests. In contrast, the thermally treated PIM-CD-1% membranes at 300 and 400 °C can all surpass the upper bound lines, as illustrated in Fig. 9.

Fig. 10 also shows the gas separation performance of the thermally treated PIM-CD-2% membranes against the upper bound lines. The C_3H_6/C_3H_8 selectivity increases with increasing the thermal treatment temperature and generally follows the order of PIM-CD-2% (600 °C) > PIM-CD-2% (500 °C) > PIM-CD-2% (400 °C) > PIM-CD-2% (300 °C), while the gas permeability decreases and follows the opposite order. Interestingly, the PIM-CD-2% membranes thermally treated at 300, 400 and 500 °C can all surpass the upper bound line for mixed gas separation. Since the main released gases derived from CD decomposition are C_2/C_3 compounds (Fig. 3) which have similar sizes with C_3H_6 and C_3H_8 , the transport channels resulted from CD's decomposition may possess the appropriate characteristics to discriminate C_3H_6 and C_3H_8 , thus improving the gas pair selectivity.^{47–49} In other words, the decomposition of thermal liable β -CD at high temperatures would generate free radicals and effectively accelerate the self-cross-linking reaction of nitrile to form triazine rings and pull the polymer chains closely

Table 1 Permeability, solubility and diffusion coefficients of the thermally treated membranes at 3.5 atm and 35 °C

Membranes	Pure gas permeability (barrer)		Permeability selectivity	Solubility coefficient ($\text{cm}^3 \text{cm}^{-3} \text{cm Hg}^{-1}$)		Solubility selectivity	Diffusion coefficient ($10^{-10} \text{cm}^2 \text{s}^{-1}$)		Diffusivity selectivity
	C_3H_6	C_3H_8		C_3H_6	C_3H_8		C_3H_6	C_3H_8	
PIM (untreated)	2861	1022	2.80	0.45	0.41	1.10	6358	2493	2.55
PIM (300 °C)	1792	498	3.60	0.43	0.38	1.13	4167	1311	3.18
PIM (400 °C)	817	154	5.31	0.37	0.33	1.12	2208	467	4.73
PIM (500 °C)	127	20	6.35	0.33	0.27	1.22	385	74	5.20
PIM (600 °C)	27	4	6.75	0.45	0.37	1.22	59	11	5.36
PIM-CD-1% (untreated)	4315	1725	2.50	0.46	0.43	1.07	9380	4012	2.34
PIM-CD-1% (300 °C)	2942	817	3.60	0.44	0.39	1.13	6686	2095	3.19
PIM-CD-1% (400 °C)	1217	214	5.69	0.40	0.35	1.14	3042	611	4.98
PIM-CD-1% (500 °C)	122	18	6.78	0.29	0.25	1.16	421	72	5.85
PIM-CD-1% (600 °C)	22	3	7.33	0.35	0.29	1.21	63	10	6.30
PIM-CD-2% (untreated)	5032	2097	2.40	0.47	0.45	1.04	10 706	4660	2.30
PIM-CD-2% (300 °C)	3642	958	3.80	0.44	0.38	1.16	8277	2521	3.28
PIM-CD-2% (400 °C)	2093	403	5.19	0.39	0.33	1.18	5367	1221	4.40
PIM-CD-2% (500 °C)	192	27	7.11	0.37	0.31	1.19	519	87	5.97
PIM-CD-2% (600 °C)	42	5	8.40	0.41	0.32	1.28	102	16	6.38



Table 2 A comparison of the C₃H₆/C₃H₈ transport properties with those of previously reported PIM-1, PIM-PI and commercial polymers^a

Membranes	Pure gas permeability (barrer)		Selectivity	Feed pressure (atm)	Temperature (°C)	Reference
	C ₃ H ₆	C ₃ H ₈				
Matrimid®	0.10	0.01	10.0	2.0	35	3
PIM-1	1575	205	7.68	1.0	25	39
PIM-1	1051	140	7.51	1.0	30	51
PIM-1	2534	851	2.98	3.5	35	21
PIM-1	2861	1022	2.80	3.5	35	This work
PIM-PI	393	65	6.00	2.0	35	52
PDMS	6600	6000	1.10	2.0	50	43
PPO	9	2	4.25	3.0	30	44
CA	15	6	2.60	3.0	30	45
PSF	25	18	1.40	3.0	30	45

^a Note: polydimethylsiloxane (PDMS), cellulose acetate (CA), polysulfone (PSF), poly(2,6-dimethyl-1,4-phenylene oxide) (PPO), the date of CA and PSF in ref. 45 were tested using a mixed gas containing 55% propylene and 45% propane.

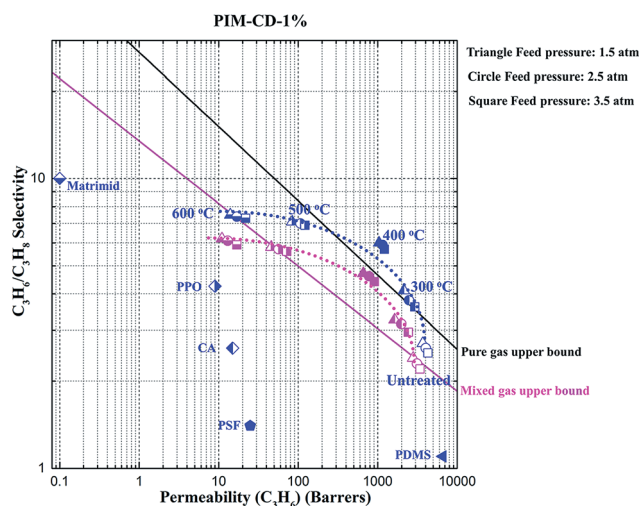


Fig. 9 Comparison of the pure and mixed gas performance of the thermally treated PIM-CD-1% membranes at 35 °C (lines are drawn to guide the eyes).

together. This results in a three-dimensional cross-linked network with narrow gates that not only offer a better size and shape selectivity but also keep the overall permeability sufficiently high for rapid diffusion of gas molecules. Therefore, the gas selectivity of the membranes treated at and below 400 °C increases rapidly while their gas permeability just decreases slightly compared to that of the original untreated one.

There are two competing factors influencing the gas permeability. The positive one is the provision of super highways by the CD units for gas transport, while the negative one is the shrinkage of the microspores when annealed at high temperatures. As observed from the TGA spectra, only CD decomposes during the first stage of weight loss from 300 to 400 °C. Clearly, after the CD decomposition, the spaces originally occupied by CD convert to micro-pores, which provide highways for gas transport. On the other hand, the heat treatment densifies the polymer chains, which is a negative factor for

gas transport. When the annealing temperature increases to 500 °C or higher (above the glass transition temperature of PIM), the polymer chains tend to pack more densely. Cross-linking reactions and partial carbonization may also occur that shrink the micropores to ultrafine micropores with pore dimensions able to discriminate the C₃H₆/C₃H₈ gas pair. Since gas selectivity is mainly determined by ultra-fine micro-pores, while gas permeability is largely related to the micro-pores, the gas permeability of PIM-CD membranes decreases while their selectivity increases with an increase in the thermal treatment temperature.

Fig. 11 shows the evolution of the *d*-spacing of the polymer chains as a function of thermal treatment temperature. The untreated PIM-CD-1% membrane has a peak at about 6.6 Å which is probably attributed to the loosely packed polymer chains derived from the kinked ladder-type backbone of PIM because of the formation of micro-pores among the polymer chains.⁵⁰ The next peak at 4.9 Å is the *d*-spacing of the efficiently packed

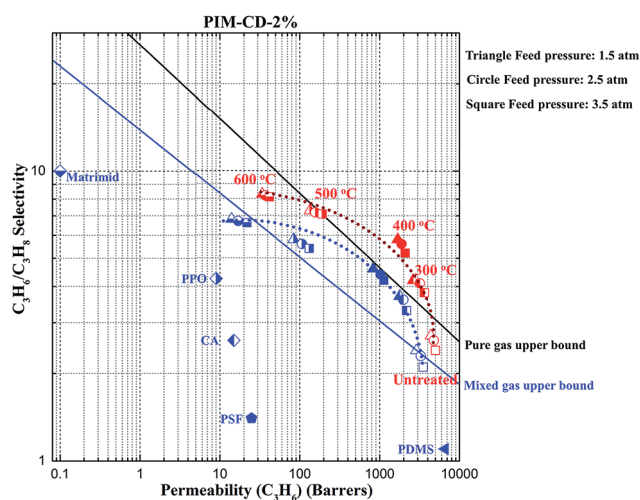


Fig. 10 Comparison of the pure and mixed gas performance of the thermally treated PIM-CD-2% membranes at 35 °C (lines are drawn to guide the eyes).



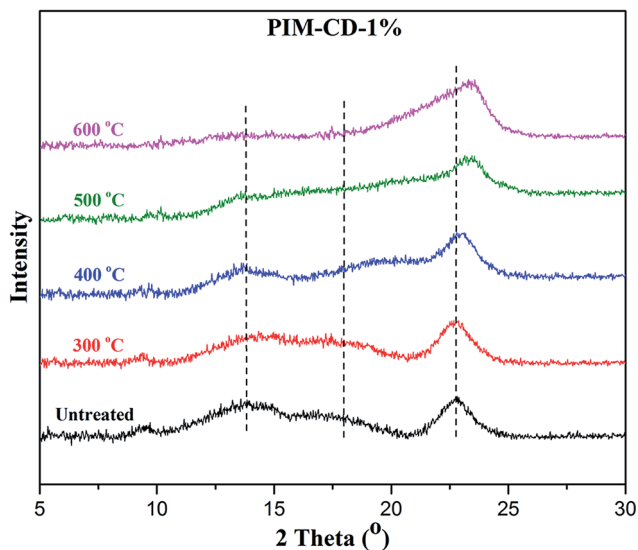


Fig. 11 XRD spectra of the untreated and thermally treated PIM-CD-1% membranes.

polymer chain-to-chain distance that contributes directly to the conformation of small micro-pores. After thermal treatment, the two peaks at 6.6 and 4.9 Å gradually shrink with increasing annealing temperature. They almost disappear when the annealing temperature increases to 500 or 600 °C (*i.e.*, above the glass transition temperature of PIM). These suggest the densification of the polymer chains and the occurrence of the self-cross-linking reaction or partial carbonization in the membrane. As a result, the micropores shrink to ultra-fine micropores. However, the last peak at 3.8 Å is always there regardless of the annealing temperature because it is a typical peak for aromatic systems and amorphous carbon configurations. The kinetic diameters of propylene and propane are about 4.0 and 4.2 Å, respectively. The ultra-fine micropores with a size about 4.1–3.5 Å (a broad distribution near 3.8 Å) play an important role as molecular sieves for propylene/propane separation. The thermally liable β -CD decomposes to generate free radicals after thermal treatment, which can effectively accelerate the self-cross-linking reaction of nitrile. The resultant 3D cross-linking networks have narrow gates as molecular sieves with better size and shape exclusion for propylene/propane separation. Clearly, the XRD results support our previous observation that the permeability always decreases while the C_3H_6/C_3H_8 selectivity increases with an increase in the thermal treatment temperature.

3.3 Gas sorption isotherms

The pure gas sorption isotherms of the untreated and thermally treated membranes at 35 °C up to 5 atm are illustrated in Fig. 12. For the same gas either C_3H_6 or C_3H_8 , the amount of gas sorption obeys the order of PIM-CD-2% > PIM-CD-1% > PIM because the addition of CD into membranes creates more micro-pores and free volume to capture gas molecules regardless of thermal treatment, while for the same membrane, the sorption amount is in the order of C_3H_6 > C_3H_8 , following the exact order of their condensability (*i.e.*, critical temperatures).

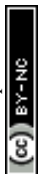
The PIM-CD membranes annealed at 300 °C have slightly lower C_3H_6 and C_3H_8 sorption concentrations than the untreated ones because of partial CD decomposition and slight chain densification at this temperature. When the annealing temperature is 400 °C, almost all CD decomposes but the C_3H_6 and C_3H_8 sorption concentrations are further reduced due to a greater degree of chain densification.

When the thermal treatment temperature is increased to 500–600 °C, the gas sorption isotherms present an interesting behavior. The C_3H_6 and C_3H_8 sorption concentrations increase quickly at low pressures, and then reach plateaus rapidly. Obviously, the total sorption amounts of C_3H_6 and C_3H_8 in the untreated membranes at 5 atm are higher than those in the 600 °C treated membranes. This is probably due to the partial carbonization of the PIM polymer chains that shrinks the micropores to ultrafine micropores during the pyrolysis. The sorption behavior of the carbonized membranes (600 °C) is in agreement with the Langmuir mode because it behaves like a hole-filling process (*i.e.*, the ultrafine micropores are filled with C_3H_6 or C_3H_8 rapidly at low pressures) and there is a sorption capacity limitation for each gas. The change of free volume could approximately be evaluated from the gas sorption isotherms. The free volume gradually shrinks according to the reduced C_3H_6/C_3H_8 gas sorption capability after the thermal treatment.

Since the amorphous carbon sheets may have strong π - π interactions with C_3H_6 and C_3H_8 , the carbonized membranes thermally treated at 600 °C have higher sorption concentrations of C_3H_6 and C_3H_8 than the ones treated at 500 °C. In contrast, the gas sorption behavior of the thermally treated membranes at lower temperatures (<500 °C) follows the dual-mode sorption model (*i.e.*, Henry and Langmuir modes).³⁹ Their sorption concentrations exhibit convex curves against the pressure axis. This non-linear behavior is a typical sorption pattern of glassy polymeric membranes.

3.4 Gas diffusivity and solubility

Since the solubility coefficient can be calculated from $S = c/p$, Table 1 summarizes the corresponding solubility coefficients and diffusion coefficients at 3.5 atm and 35 °C. For PIM, PIM-CD-1% and PIM-CD-2%, the sharp drop of their permeability with an increase in the annealing temperature is mainly attributed to the decrease of the diffusion coefficient. The diffusion coefficient of C_3H_6 decreases about 100-fold from its original value of $6358 \times 10^{-10} \text{ cm}^2 \text{ s}^{-1}$ in the pristine PIM to its minimum value of $59 \times 10^{-10} \text{ cm}^2 \text{ s}^{-1}$ in the 600 °C treated membrane; meanwhile, the diffusion coefficient of C_3H_8 decreases about 200-fold from its original value of $2493 \times 10^{-10} \text{ cm}^2 \text{ s}^{-1}$ to its minimum value of $11 \times 10^{-10} \text{ cm}^2 \text{ s}^{-1}$. Fig. 13 plots the diffusion coefficient of all membranes as a function of annealing temperature. The diffusion coefficient has a slower decreasing speed with the annealing temperature for the membranes thermally treated up to 400 °C than those treated from 500 to 600 °C, especially for PIM-CD membranes. This interesting characteristic is probably due to the fact that the bowl-shaped cavity of 3D β -CD can support the architecture and make the micro-pores unable to collapse when the thermal



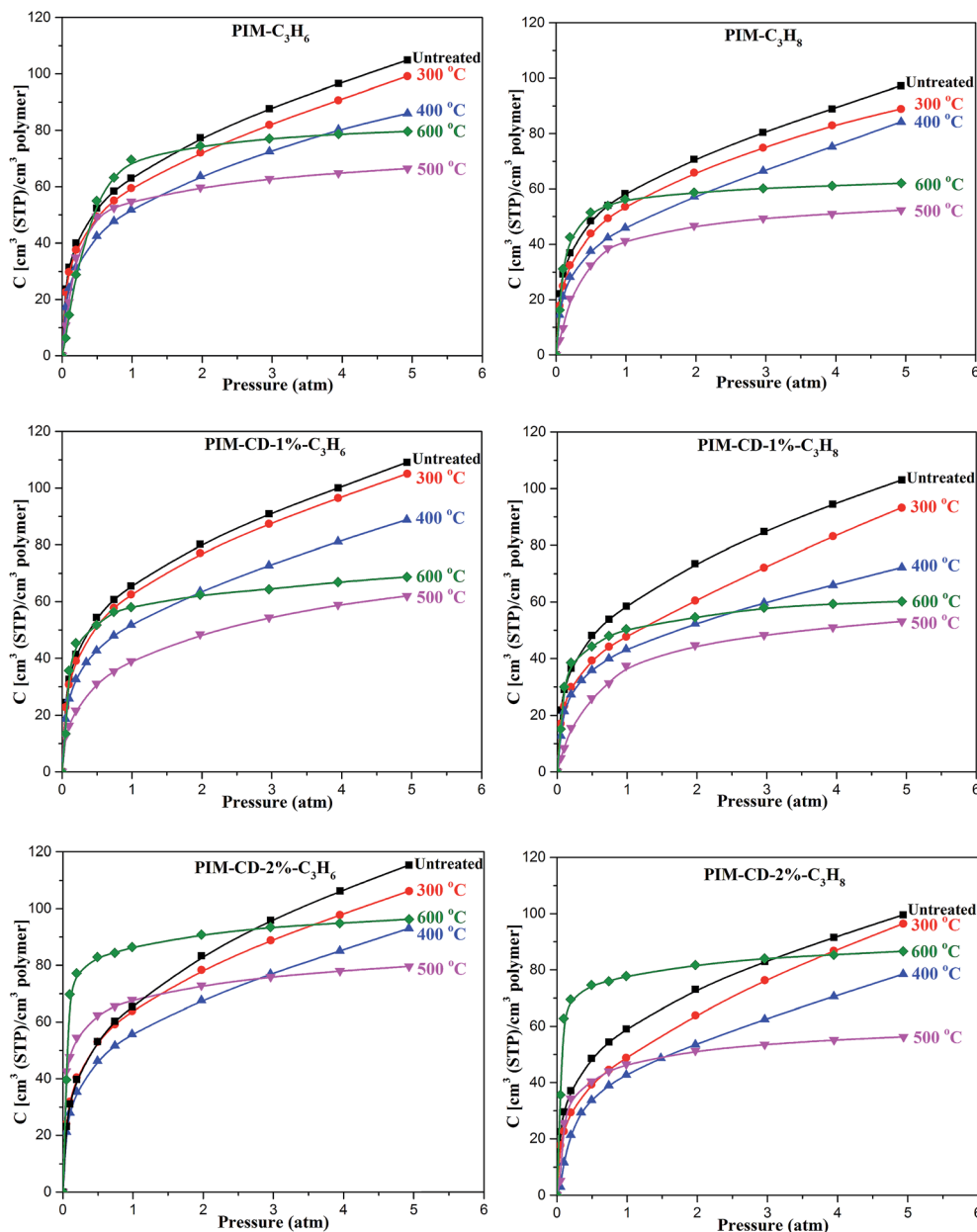
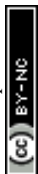


Fig. 12 Sorption isotherms of the untreated and thermally treated membranes at 35 °C.

treatment temperature is low. However, when the annealing temperature is higher than the glass transition temperature of PIM and the decomposition temperature of CD, the bowl-shaped cavity of 3D β -CD collapses. As a consequence, it densifies the polymer chains, shrinks the free volume and reduces the diffusion coefficients of the membranes treated at 500 or 600 °C significantly.

The diffusivity selectivity and solubility selectivity of the untreated and thermally treated membranes are shown in Fig. 14. Both the diffusivity selectivity and the solubility selectivity increase with increasing the thermal treatment temperature. The diffusivity selectivity increases faster than the solubility selectivity for all membranes. Since the ideal permeability selectivity can be defined as the product of diffusivity selectivity and solubility selectivity, the sharp increase of the gas

permeability selectivity is mainly attributed to the increase of the diffusivity selectivity. The PIM-CD membranes after thermal treatment have a greater increase in the diffusivity selectivity than the PIM membrane. Comparing the diffusivity selectivity between the untreated and 600 °C treated PIM membranes, it increases from 2.5 to 5.4 with an increase of 2.2 times, while it increases from 2.3 to 6.1 with an increase of 2.7 times for the untreated and 600 °C treated PIM-CD-1% membranes. Similarly, it increases from 2.3 to 6.6 with an increase of 2.9 times for the untreated and 600 °C treated PIM-CD-2% membranes. Since the diffusivity selectivity essentially matches the actual pure-gas permselectivity for the whole temperature range, the molecular sieving mechanism is dominant in these thermally self-cross-linked membranes.



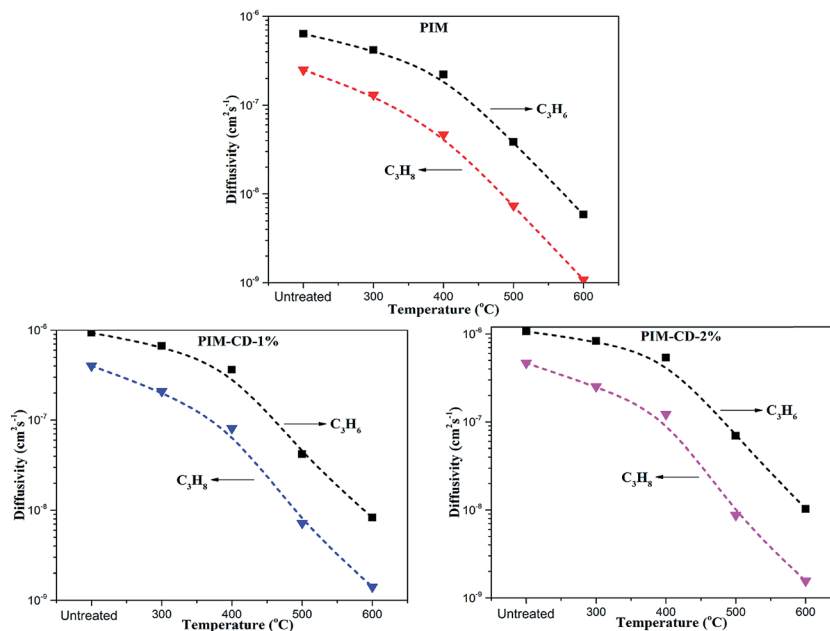


Fig. 13 The C_3H_6 and C_3H_8 diffusivity at 3.5 atm and 35 °C (lines are drawn to guide the eyes).

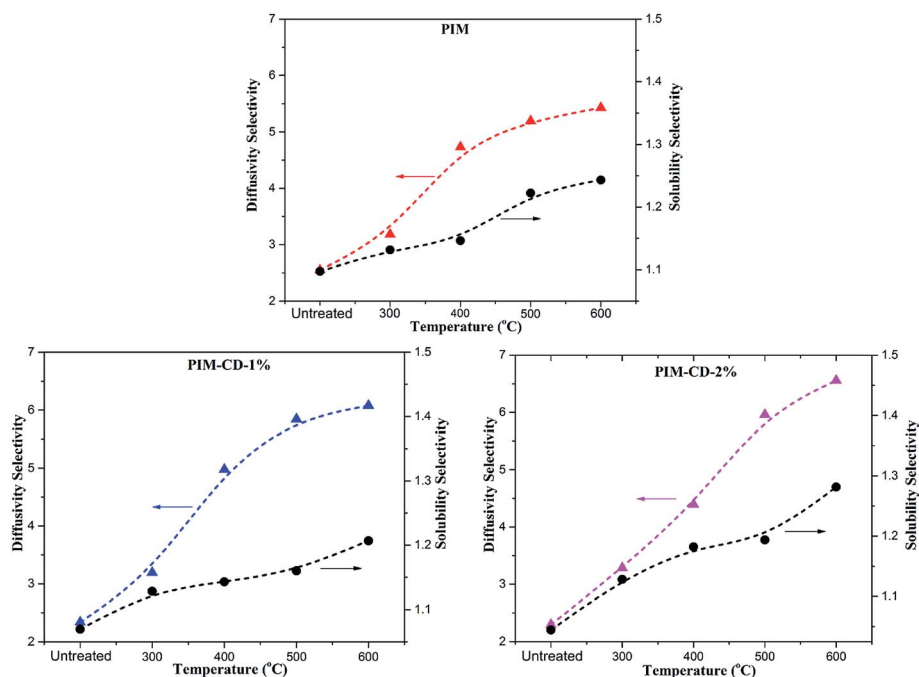


Fig. 14 The C_3H_6/C_3H_8 diffusivity and solubility selectivity at 3.5 atm and 35 °C (lines are drawn to guide the eyes).

4. Conclusions

A series of PIM-CD polymers were prepared by incorporating thermally labile 3D β -CD into PIM polymer chains *via* chemical bonds from nucleophilic substitution copolymerization of β -CD, TFTP, and TTSBI. Once the PIM-CD membranes are thermally treated at high temperatures, the thermal labile β -CD decomposes to generate free radicals, which can effectively accelerate the self-cross-linking reaction of nitrile. The formation of triazine rings

tends to pull the PIM polymer chains approaching one another closely. Since the spaces originally occupied by CD convert to micro-pores while the rigidly contorted PIM polymer framework persists, the resultant 3D cross-linking networks have narrow gates for remarkably better size and shape selectivity while the overall permeability is still maintained sufficiently high for rapid gas diffusion. As a result, the selectivity for C_3H_6/C_3H_8 rises rapidly while the permeability just drops slightly. In addition to showing high mechanical flexibility, the PIM-CD membranes



treated at 300 or 400 °C have separation performance surpassing the upper bound line in both pure and mixed gas tests.

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References

- 1 D. F. Sanders, Z. P. Smith, R. Guo, L. M. Robeson, J. E. McGrath, D. R. Paul and B. D. Freeman, Energy-efficient polymeric gas separation membranes for a sustainable future: a review, *Polymer*, 2013, **54**, 4729–4761.
- 2 R. B. Eldridge, Olefin/paraffin separation technology: a review, *Ind. Eng. Chem. Res.*, 1993, **32**, 2208–2212.
- 3 R. L. Burns and W. J. Koros, Defining the challenges for C₃H₆/C₃H₈ separation using polymeric membranes, *J. Membr. Sci.*, 2003, **211**, 299–309.
- 4 L. Liu, A. Chakma and X. Feng, Propylene separation from nitrogen by poly(ether block amide) composite membranes, *J. Membr. Sci.*, 2006, **279**, 645–654.
- 5 S. W. Kang, K. Char and Y. S. Kang, Novel application of partially positively charged silver nanoparticles for facilitated transport in olefin/paraffin separation membranes, *Chem. Mater.*, 2008, **20**, 1308–1311.
- 6 M. D. Guiver and Y. M. Lee, Polymer rigidity improves microporous membranes, *Science*, 2013, **339**, 284–285.
- 7 N. Peng, N. Widjojo, P. Sukitpeneit, M. M. Teoh, G. G. Lipscomb, T. S. Chung and J. Y. Lai, Evolution of polymeric hollow fibers as sustainable technologies: past, present, and future, *Prog. Polym. Sci.*, 2012, **37**, 1401–1424.
- 8 C. Zhang, Y. Dai, J. R. Johnson, O. Karvan and W. J. Koros, High performance ZIF-8/6FDA-DAM mixed matrix membrane for propylene/propane separations, *J. Membr. Sci.*, 2012, **389**, 34–42.
- 9 S. S. Chan, R. Wang, T. S. Chung and Y. Liu, C₂ and C₃ hydrocarbon separations in poly(1,5-naphthalene-2,2'-bis(3,4-phthalic)hexafluoropropane)diimide (6FDA-1,5-NDA) dense membranes, *J. Membr. Sci.*, 2002, **210**, 55–64.
- 10 R. W. Baker, Future directions of membrane gas separation technology, *Ind. Eng. Chem. Res.*, 2002, **41**, 1393–1411.
- 11 R. J. Swaidan, B. Ghanem, R. Swaidan, E. Litwiller and I. Pinnau, Pure- and mixed-gas propylene/propane permeation properties of spiro- and triptycene-based microporous polyimides, *J. Membr. Sci.*, 2015, **492**, 116–122.
- 12 Z. P. Smith, G. Hernandez, K. L. Gleason, A. Anand, C. M. Doherty, K. Konstas, C. Alvarez, A. J. Hill, A. E. Lozano, D. R. Paul and B. D. Freeman, Effect of polymer structure on gas transport properties of selected aromatic polyimides, polyamides and TR polymers, *J. Membr. Sci.*, 2015, **493**, 766–781.
- 13 D. Liu, X. Ma, H. Xi and Y. S. Lin, Gas transport properties and propylene/propane separation characteristics of ZIF-8 membranes, *J. Membr. Sci.*, 2014, **451**, 85–93.
- 14 R. Faiz and K. Li, Polymeric membranes for light olefin/paraffin separation, *Desalination*, 2012, **287**, 82–97.
- 15 R. Faiz and K. Li, Olefin/paraffin separation using membrane based facilitated transport/chemical absorption techniques, *Chem. Eng. Sci.*, 2012, **73**, 261–284.
- 16 T. C. Merkel, R. Blanc, I. Ciobanu, B. Firat, A. Suwarlim and J. Zeid, Silver salt facilitated transport membranes for olefin/paraffin separations: carrier instability and a novel regeneration method, *J. Membr. Sci.*, 2013, **447**, 177–189.
- 17 R. S. Murali, K. Y. Rani, T. Sankarshana, A. F. Ismail and S. Sridhar, Separation of binary mixtures of propylene and propane by facilitated transport through silver incorporated poly(ether-block-amide) membranes, *Oil Gas Sci. Technol.*, 2015, **70**, 381–390.
- 18 L. Liu, X. Feng and A. Chakma, Unusual behavior of poly(ethylene oxide)/AgBF₄ polymer electrolyte membranes for olefin–paraffin separation, *Sep. Purif. Technol.*, 2004, **38**, 255–263.
- 19 S. Hess, C. Staudt-Bickel and R. N. Lichtenthaler, Propene/propane separation with copolyimide membranes containing silver ions, *J. Membr. Sci.*, 2006, **275**, 52–60.
- 20 S. W. Kang, J. H. Kim, J. Won and Y. S. Kang, Suppression of silver ion reduction by Al(NO₃)₃ complex and its application to highly stabilized olefin transport membranes, *J. Membr. Sci.*, 2013, **445**, 156–159.
- 21 K. S. Liao, J. Y. Lai and T. S. Chung, Metal ion modified PIM-1 and its application for propylene/propane separation, *J. Membr. Sci.*, 2016, **515**, 36–44.
- 22 I. Pinnau and L. G. Toy, Solid polymer electrolyte composite membranes for olefin/paraffin separation, *J. Membr. Sci.*, 2001, **184**, 39–48.
- 23 C. Zhang, Y. Dai, J. R. Johnson, O. Karvan and W. J. Koros, High performance ZIF-8/6FDA-DAM mixed matrix membrane for propylene/propane separations, *J. Membr. Sci.*, 2012, **389**, 34–42.
- 24 J. Ploegmakers, S. Japip and K. Nijmeijer, Mixed matrix membranes containing MOFs for ethylene/ethane separation part A: membrane preparation and characterization, *J. Membr. Sci.*, 2013, **428**, 445–453.
- 25 M. Askari and T. S. Chung, Natural gas purification and olefin/paraffin separation using thermal cross-linkable copolyimide/ZIF-8 mixed matrix membranes, *J. Membr. Sci.*, 2013, **444**, 173–183.
- 26 S. Japip, H. Wang, Y. C. Xiao and T. S. Chung, Highly permeable zeolitic imidazolate framework (ZIF-71) nanoparticles enhanced polyimide membranes for gas separation, *J. Membr. Sci.*, 2014, **467**, 162–174.
- 27 N. Hara, M. Yoshimune, H. Negishi, K. Haraya, S. Hara and T. Yamaguchi, Diffusive separation of propylene/propane with ZIF-8 membranes, *J. Membr. Sci.*, 2014, **450**, 215–223.
- 28 M. L. Chng, Y. C. Xiao, T. S. Chung, M. Toriida and S. Tamai, Enhanced C₃H₆/C₃H₈ separation by an ingenious carbonaceous membrane derived from PAEK/azide interpenetrating network, *Carbon*, 2009, **47**, 1857–1866.



- 29 X. Ma, Y. S. Lin, X. Wei and J. Kniep, Ultrathin carbon molecular sieve membrane for propylene/propane separation, *AIChE J.*, 2016, **62**, 491–499.
- 30 K. M. Steel and W. J. Koros, An investigation of the effects of pyrolysis parameters on gas separation properties of carbon materials, *Carbon*, 2005, **43**, 1843–1856.
- 31 K. Okamoto, S. Kawamura, M. Yoshino, H. Kita, Y. Hirayama, N. Tanihara and Y. Kusuki, Olefin/paraffin separation through carbonized membranes derived from an asymmetric polyimide hollow fiber membrane, *Ind. Eng. Chem. Res.*, 1999, **38**, 4424–4432.
- 32 J. Hayashi, H. Mizuta, M. Yamamoto, K. Kusakabe, S. Morooka and S. H. Suh, Separation of ethane/ethylene and propane/propylene systems with a carbonized BPDA-pp'ODA polyimide membrane, *Ind. Eng. Chem. Res.*, 1996, **35**, 4176–4181.
- 33 X. L. Ma, B. K. Lin, X. T. Wei, J. Kniep and Y. S. Lin, Gamma-alumina supported carbon molecular sieve membrane for propylene/propane separation, *Ind. Eng. Chem. Res.*, 2013, **52**, 4297–4305.
- 34 O. Salinas, X. H. Ma, E. Litwiller and I. Pinnau, Ethylene/ethane permeation, diffusion and gas sorption properties of carbon molecular sieve membranes derived from the prototype ladder polymer of intrinsic microporosity (PIM-1), *J. Membr. Sci.*, 2016, **504**, 133–140.
- 35 O. Salinas, X. H. Ma, E. Litwiller and I. Pinnau, High-performance carbon molecular sieve membranes for ethylene/ethane separation derived from an intrinsically microporous polyimide, *J. Membr. Sci.*, 2016, **500**, 115–123.
- 36 P. M. Budd, B. S. Ghanem, S. Makhseed, N. B. McKeown, K. J. Msayib and C. E. Tattershall, Polymers of intrinsic microporosity (PIMs): robust, solution processable, organic nanoporous materials, *Chem. Commun.*, 2004, 230–231.
- 37 P. M. Budd, E. S. Elabas, B. S. Ghanem, S. Makhseed, N. B. McKeown, K. J. Msayib, C. E. Tattershall and D. Wang, Solution-processed, organophilic membrane derived from a polymer of intrinsic microporosity, *Adv. Mater.*, 2004, **16**, 456–459.
- 38 F. Y. Li, Y. C. Xiao, T. S. Chung and S. Kawi, High-performance thermally self-cross-linked polymer of intrinsic microporosity (PIM-1) membranes for energy development, *Macromolecules*, 2012, **45**, 1427–1437.
- 39 P. Li, T. S. Chung and D. R. Paul, Gas sorption and permeation in PIM-1, *J. Membr. Sci.*, 2013, **432**, 50–57.
- 40 N. Y. Du, M. M. Dal-Cin, I. Pinnau, A. Nicalek, G. P. Robertson and M. D. Guiver, Azide-based cross-linking of polymers of intrinsic microporosity (PIMs) for condensable gas separation, *Macromol. Rapid Commun.*, 2011, **32**, 631–636.
- 41 J. T. Liu, Y. C. Xiao, K. S. Liao and T. S. Chung, Highly permeable and aging resistant 3D architecture from polymers of intrinsic microporosity incorporated with beta-cyclodextrin, *J. Membr. Sci.*, 2017, **523**, 92–102.
- 42 A. Alsbaiee, B. J. Smith, L. Xiao, Y. Ling, D. E. Helbling and W. R. Dichtel, Rapid removal of organic micropollutants from water by a porous β -cyclodextrin polymer, *Nature*, 2016, **529**, 190–194.
- 43 K. Tanaka, A. Taguchi, J. Q. Hao, H. Kita and K. Okamoto, Permeation and separation properties of polyimide membranes to olefins and paraffins, *J. Membr. Sci.*, 1996, **121**, 197–207.
- 44 S. Bai, S. Sridhar and A. A. Khan, Metal-ion mediated separation of propylene from propane using PPO membranes, *J. Membr. Sci.*, 1998, **147**, 131–139.
- 45 S. Sridhar and A. A. Khan, Simulation studies for the separation of propylene and propane by ethylcellulose membrane, *J. Membr. Sci.*, 1999, **159**, 209–219.
- 46 P. S. Tin, T. S. Chung, Y. Liu, R. Wang, S. L. Liu and K. P. Pramoda, Effects of cross-linking modification on gas separation performance of Matrimid membranes, *J. Membr. Sci.*, 2003, **225**, 77–90.
- 47 Y. C. Xiao and T. S. Chung, Grafting thermally labile molecules on cross-linkable polyimide to design membrane materials for natural gas purification and CO₂ capture, *Energy Environ. Sci.*, 2011, **4**, 201–208.
- 48 M. Askari, M. L. Chua and T. S. Chung, Natural gas purification and olefin/paraffin separation using cross-linkable, 6FDA-Durene/DABA co-polyimides grafted with α , β , and γ -cyclodextrin, *J. Membr. Sci.*, 2012, **390–391**, 141–151.
- 49 M. Askari, T. X. Yang and T. S. Chung, Natural gas purification and olefin/paraffin separation using cross-linkable dual-layer hollow fiber membranes comprising β -cyclodextrin, *J. Membr. Sci.*, 2012, **423–424**, 392–403.
- 50 F. Y. Li, Y. C. Xiao, Y. K. Ong and T. S. Chung, UV-rearranged PIM-1 polymeric membranes for advanced hydrogen purification and production, *Adv. Energy Mater.*, 2012, **2**, 1456–1466.
- 51 M. M. Khan, G. Bengtson, S. Shishatskiy, B. N. Gacal, M. M. Rahman, S. Neumann, V. Filiz and V. Abetz, Cross-linking of Polymer of Intrinsic Microporosity (PIM-1) via nitrene reaction and its effect on gas transport property, *Eur. Polym. J.*, 2013, **49**, 4157–4166.
- 52 R. J. Swaidan, B. Ghanem, R. Swaidan, E. Litwiller and I. Pinnau, Pure- and mixed-gas propylene/propane permeation properties of spiro- and triptycene-based microporous polyimides, *J. Membr. Sci.*, 2015, **492**, 116–122.

