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A practically simple, catalyst free and scalable synthesis of *N*-substituted ureas in water

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A practically simple, mild and efficient method is developed for the synthesis of *N*-substituted ureas by nucleophilic addition of amines to potassium isocyanate in water without organic co-solvent. Using this methodology, a variety of *N*-substituted ureas (mono-, di- and cyclic-) were synthesized in good to excellent yields with high chemical purity by applying simple filtration or routine extraction procedures avoiding silica gel purification. The developed methodology was also found to be suitable for gram scale synthesis of molecules having commercial application in large volumes. The identified reaction conditions were found to promote a unique substrate selectivity from a mixture of two amines.

Introduction

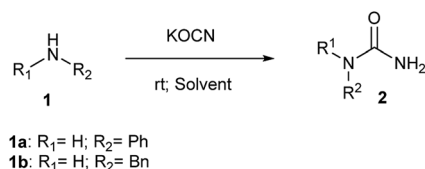
Development of resource-efficient as well as environment-friendly synthetic processes for manufacturing of chemicals and intermediates being used in large quantities is a must for sustainable industrial development.¹ Urea and its derivatives are an important class of molecules having diverse chemical and biological properties and hence being extensively employed in chemical,² pharmaceutical³ and agrochemical⁴ industries. Because of their diverse applications, a large number of urea derivatives are high production volume chemicals (HPV).⁵ Being molecules of commercial interest, not only their use but also their manufacturing methods are a potential source of environment pollution. There are strict regulatory guidelines for use and application of HPV chemicals. Unfortunately, that is not the case for their manufacturing processes.⁶ In general, industrial manufacturing processes for high volume chemicals are influenced by economy and ease of execution providing little consideration for environmental impacts.⁷ Synthetic chemists have obligations and challenges to develop resource-efficient as well as environment-friendly production methods.^{1d} *N*-Substituted ureas are an important subclass of urea derivatives, where one of the nitrogen atoms is substituted with an alkyl or aryl group. The *N*-substituted urea derivatives were not only found to have direct applications as agrochemicals⁴ or pharmaceutical³ agents but also serve as building blocks for various other important chemicals² of high commercial interest.

The most common method for the synthesis of *N*-substituted ureas involve the reaction of isocyanates or carbamoyl chlorides with ammonia.^{3f-i,8} The desired isocyanate and carbamoyl

chloride can be generated by the reaction of corresponding amine with phosgene.^{8f,g} This is neither environment friendly nor safe,⁹ but still one of the few methods which is assertively being used. Lately, phosgene was replaced with safer reagents such as triphosgene, 1,1-dicarbonyl imidazole and carbonates.¹⁰ Hoffmann, Curtius and Lossen rearrangement are complementary approaches which exploit *in situ* isocyanate generation.¹¹ However compared to phosgene, these methods are less attractive for commercial production. Importantly, all these methods involving isocyanate need ammonia gas as one of the reactant, which led to safety issues and operational complexities. Direct alkylation or arylation of urea appears to be straightforward approach but need harsh conditions.¹² Also, involvement of *O*-alkylation and/or dialkylation as competing reactions is a point of concern. Reaction of amines with 4-nitrophenyl-*N*-benzyl carbamate followed by hydrogenation,¹³ oxidation of thiourea derivatives,¹⁴ metal or base catalysed transamidation of urea,¹⁵ conversion of cyanamides *via* Tie-mann rearrangement¹⁶ and transition metal catalysed carbonylation¹⁷ of amines are some recently developed less explored methods for urea synthesis. These methods are mostly employed for synthesis of substituted ureas and though added huge intellectual addition on academic understandings, but unfortunately found to have no or very limited applicability for industrial production of *N*-substituted ureas. Transamidation of urea is an attractive approach, but often requires catalyst or/and heating.¹⁵ Interestingly, nucleophilic addition of amine to metal isocyanate looks like a most facile and safe method for commercial production of *N*-substituted ureas.¹⁸ However, this is least explored method compared to all with very limited substrate scope.^{18d-f} Moreover, most of the reports are limited to the use of either excess of acid or specialized acid as promoter and often requires high temperature or microwave irradiation.^{18a,b} Requirement of harsh conditions and fancy promoters not only restrict the scope but also make the

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Scheme 1

approach a resource inefficient and environmentally hazardous. Herein, we want to report a synthetic protocol for the synthesis of *N*-substituted ureas in water, which is practically simple, having broader substrate scope, found to be scalable and environment friendly as well. The newly developed method potentially has a possibility to be a method of choice for manufacturing of various *N*-substituted ureas, which are being used in industry as large volume commercial products and intermediates. Additionally, the developed methodology holds potential to provide a chemo-selectivity based on pK_aH of amines.

Results and discussion

To facilitate internal drug discovery efforts, we were keen to develop a mild and efficient method for synthesis of *N*-substituted ureas exploiting nucleophilic addition of amines to potassium isocyanate. Accordingly, reaction of aniline (**1a**) and benzylamine (**1b**) was attempted individually in presence of hydrochloric acid using different solvent systems at 25 °C (Scheme 1). The hydrochloric acid was used primarily to activate potassium cyanate as an electrophile. The outcome of this study is tabulated in Table 1. We planned to use 4 M solution of HCl in dioxane for this reaction with and without any additional co-solvent such as tetrahydrofuran (THF) and acetonitrile (ACN)

(entry 1, 2 and 3; Table 1). Reactions performed in either neat dioxane or using freshly dried THF as a co-solvent, gave no conversion for aniline and low conversion for benzylamine (entry 1 and 2) to corresponding ureas **2a** and **2b** respectively. Intriguingly, ACN as a co-solvent found to have a profound effect on reactivity of aniline and no effect on benzylamine unlike THF (entry 3 vs. entry 2). Using ACN as co-solvent not only gave better conversion for urea **2a** but also demonstrated a reversal in reactivity order of two amines. Aniline, in spite of being less nucleophilic gave better conversion compared to benzylamine using ACN as solvent. We started looking for an explanation to describe this unusual effect of ACN. It was noticed that, the ACN used from shelf bottle has a small amount of residual water as mentioned by manufacturer label. This instigated us to do a new experiment with fresh lot of ACN that was dried over molecular sieves before employing in reaction (entry 4).

The use of dry ACN, was found to be detrimental for conversion of aniline to its corresponding urea (**2a**; 18%) but did not affect the overall conversion of benzylamine (**2b**; 30%) as mentioned in entry 4. This made us analyse and explore the role of water on reactivity of amine under given reaction conditions (entry 5, 6 and 7). For reactions of aniline and benzylamine, corresponding *N*-substituted ureas were observed in high yields when reaction was performed using water as co-solvent (entry 6) or using water only as the only solvent (entry 7). Actually, rate of reaction, product conversion and isolated yields were found to be better when reaction was performed exclusively in water without any co-solvent (entry 7 vs. entry 6). Water a universal solvent, found to be a good promoter for higher yield of *N*-substituted urea. It was also observed that 2 equivalents of KOCN afford better conversion and yields in shorter time compared to 1.1 equivalents (entry 7 vs. entry 8).

Table 1 Optimization study of reaction conditions for Scheme 1

Entry	Amine 1	4 M HCl (dioxane)	Co-solvent	Conversion ^a (%) 2	Time (hrs)
1	1a	3 mL	No solvent	0	12
	1b			17	
2	1a	1 mL	THF (dry; 2 mL)	0	12
	1b			29	
3	1a	1 mL	ACN (2 mL)	88	12
	1b			30	
4	1a	1 mL	ACN (dry; 2 mL)	18	12
	1b			ND	
5	1a	1 mL	H ₂ O (0.5 mL)	<5	12
	1b			ND	
6	1a	1 mL	H ₂ O (2 mL)	81	12
	1b			98	
7	1a	NA	Aq. HCl (1 N, 3 mL)	>95	10
	1b			>95	
8	1a	NA	Aq. HCl (1 N, 3 mL)	>95	6 ^b
	1b			>95	
9	1a	NA	Aq. HCl (0.5 N, 3 mL)	66	12
10	1a	NA	Aq. HCl (4 N, 3 mL)	31	12
11	1a	NA	Water 3 mL	14	12

^a Based on ¹H NMR. ^b 2.2 equivalent of KOCN used.



Table 2 Synthesis of *N*-substituted urea to evaluate substrate scope

$ \begin{array}{c} \text{R}_2\text{NH} \\ \\ \text{R}_1 \end{array} + \text{KOCN} \xrightarrow[3\text{mL (Vol.)}; \text{rt}]{1\text{N aq HCl}} \begin{array}{c} \text{O} \\ \\ \text{R}_2\text{N}-\text{C}-\text{NH}_2 \\ \\ \text{R}_2 \end{array} $ <p>1: 2mmol 2.2eq 2</p>			
Entry	Product	Time (h)	Isolated yield (%)
1		6	94
2		8	90
3		12	71 ^a
4		12	93
5		12	82
6		12	80
7		12	80 ^a
8		12	63 ^a
9		4	97
10		4	88
11		12	90
12		12	67 ^a
13		12	89
14		12	83



Table 2 (Contd.)

$ \begin{array}{c} \text{R}_2\text{-NH} \\ \\ \text{R}_1 \\ \text{1: 2mmol} \end{array} + \text{KOCN} \xrightarrow[3\text{mL (Vol.)}; \text{rt}]{\text{2.2eq}} \xrightarrow{1\text{N aq HCl}} \begin{array}{c} \text{O} \\ \\ \text{R}_2\text{-N}-\text{C}-\text{NH}_2 \\ \\ \text{R}_2 \\ \text{2} \end{array} $			
Entry	Product	Time (h)	Isolated yield (%)
15		12	47 ^a
16		12	70 ^a
17		24	NR ^a
18		6	89
19		12	81 ^a
20		12	73 ^a
21		5	74
22		12	52 ^a
23		12	72 ^c
24		12	68 ^c
25		12	68 ^a
26		12	44 ^a
27		12	71 ^a
28		12	91 ^b
29		12	73 ^a



Table 2 (Contd.)

$$\text{R}_2\text{NH}-\text{R}_1 + \text{KOCN} \xrightarrow[3\text{ mL (Vol.)}; \text{rt}]{1\text{ N aq HCl}} \text{R}_2\text{N}-\text{C}(=\text{O})\text{NH}_2$$

1: 2 mmol

Entry	Product	Time (h)	Isolated yield (%)
30		12	70 ^a
31		12	60 ^a
32		12	75 ^a
33		12	68 ^a
34		12	53 ^a
35		12	18 ^a
36		12	63 ^a
37		12	80
38		12	30 ^a
39		12	70 ^a
40		12	70 ^a

^a Un-reacted amine recovered. ^b Amine was used in excess; NR = no reaction. ^c Yields based on ¹H NMR as product could not be isolated from starting amine.

The critical role of HCl (entry 11 vs. entry 6) and its optimum amount (1 N, 3 mL aqueous solution of HCl) was quickly established by varying the concentration of HCl (entry 9 and 10). We attempted to build on this and planned for a substrate scope evaluation using conditions of entry 8 of Table 1 using different amines. The outcome of this study is presented in Table 2.

We tried to generalize reaction with different aromatic, aliphatic and hetero aromatic amines using 1 N aqueous HCl as solvent. All the reactions of Table 2 were done using 2 mmol of amine, 2.2 equivalent of KOCN and 3 mL of 1 N aqueous HCl.

Most of the tested amines provided good to excellent isolated yields of the corresponding *N*-substituted ureas. The reaction profile was very clean for all the evaluated substrates irrespective of the amine type used. It was either desired *N*-substituted urea as a single product or a mixture of desired product with unreacted starting amine at the end of the reaction time. Reactions involving aniline derivatives were very facile and it was practically very easy to isolate the final products. The corresponding urea product of all tested anilines tends to precipitate out with progress of the reaction, which can be



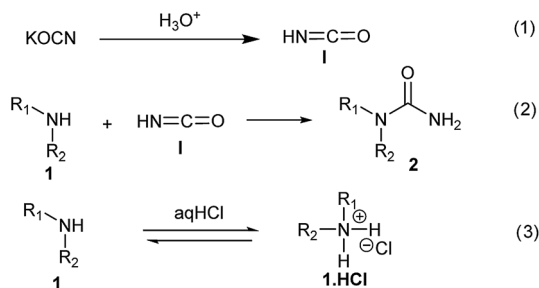
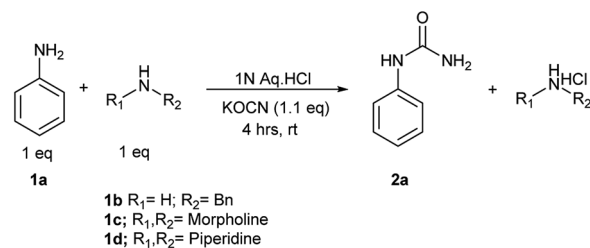


Fig. 1 Plausible reaction mechanism.

filtered out easily (entry 1 to 17). Compared to this, aliphatic ureas did not precipitate out as efficiently from the reaction mixture requiring an extraction with ethyl acetate in all reported cases. The reactions involving most of the aromatic amines proceeded smoothly and afforded good to excellent yields of the product within short reaction time (6–12 hours) except entries 3, 8, 15, 16, and 17. The higher reaction rate and better isolated yields were obtained for aniline substituted with electron donating groups *e.g.* entry 9 and 10. The lower conversion for entries 3, 8, 15 to 17 in Table 2 was attributed to the presence of electron withdrawing groups (*i.e.* Cl, CF₃, CN, OMe and NO₂) which seems to hamper the reactivity of the aniline. In contrast to aromatic amines, most of the aliphatic amines evaluated in Table 2 demonstrated slower reaction rates, which was a surprise, considering the higher nucleophilicity of aliphatic amines (entry 21, 22 and 28 to 32) compared to aniline. It is worth to mention that all of the reported reactions were very clean and one can recover unreacted aliphatic amines along with desired *N*-substituted ureas when there was an incomplete conversion. No chemical interference from other reactive groups such as –SH (entry 24), aliphatic –OH (entry 25) and phenolic –OH (entries 11 to 13) was found to be another pleasant outcome of the developed methodology. In addition, the developed methodology found to be a suitable approach to prepare *N*-carbamoylation of amino acids to generate corresponding ureas.^{18c} It is important to mention that, we only tested racemic counterpart of few selected amino acids for this study (entries 33 to 37). This methodology was found to work well for heterocyclic amines affording good yield of corresponding *N*-substituted ureas in high purity (entries 38 to 40), however yields were found to be low due to incomplete conversion of the starting amines.

The plausible mechanism for this reaction involves formation of a reactive isocyanic acid (I) from KOCN in presence of aqueous acidic conditions (eqn (1); Fig. 1). This is followed by a nucleophilic attack of amine resulting in the formation of corresponding *N*-substituted urea (eqn (2)). We tried to reason out the slow rate of reaction for aliphatic amines (entry 21, 22 and 28–32) compared to anilines (entry 1–11). We argued that, it is the higher basicity of the aliphatic amines or higher *pK_a* of their conjugate acids which is playing critical role in determining the rate of reaction. The average *pK_a*H (*pK_a* of conjugate acids) of aniline and its derivatives is low (range 2 to 5) compared to aliphatic amines (range of 8–12).¹⁹ Aliphatic



Scheme 2

amines such as diethyl amine (entry 29; Table 2) cyclohexyl amine (entry 30; Table 2), piperidine (entry 31; Table 2) and morpholine (entry 32; Table 2) tend to exist more as ammonium salt (eqn (3)) in aqueous HCl reaction conditions (pH between 2 to 4) because of their higher *pK_a*H values. Hence, higher population of the basic aliphatic amine exists as ammonium salt (1 HCl) and not as free nucleophilic amine (1) and this led to slow rate of reaction and poor conversion of desired product. Plausible reaction mechanism

We planned to test this argument. A reaction of KOCN (1.1 equiv.) was performed with an equimolar mixture of aniline (1a) and morpholine (1c) as mentioned in Scheme 2. After a reaction time of four hours, we observed exclusive formation of *N*-substituted urea (2a) of aniline in reaction mixture. There was no reaction of morpholine and hence no trace of corresponding urea 2c after reaction time of 4 hours. The un-reacted morpholine was recovered in aqueous layer after filtration of solid precipitates of 2a. Although we did observe traces of 2c *i.e.* urea of morpholine after a long reaction time of 6 hours as studied by ¹H NMR, but this was not significant. Exclusive reaction of aniline with KOCN even in presence of more nucleophilic morpholine supported our proposed hypothesis and role of *pK_a*H. We attempted two additional reactions of aniline, one each in presence of benzylamine (1b) and piperidine (1d). The results were same in these two cases as well. There was exclusive formation of aniline urea (2b) and we could recover unreacted benzylamine and piperidine respectively in aqueous layer after filtration of pure 2b. This demonstrated that, the developed reaction conditions offered unique chemo-selectivity influenced by *pK_a*H of the amines.

After successful demonstration of given protocol for synthesis of wide range of *N*-substituted ureas, we wanted to test the synthetic utility of the new protocol for industrial application. Accordingly, synthesis of few selected ureas was performed at incremental scale (1, 10 and 20 mmol) and yields of higher scale were compared to low mmol scales as mentioned in Table 3. For this study, amines were selected based on the commercial utility of their corresponding *N*-substituted ureas. For example, the *N*-substituted urea 2e of 4-*p*-aminocumene (1e, entries 1–3; Table 3) is a key intermediate for Isoproturon whereas, *N*-substituted urea of 3,4-dichloro aniline (entry 9 & 10) is a key structural motif for linuron and diuron.^{4a} Isoproturon, linuron and diuron are three herbicides which are extensively being used for crop protection. The *N*-substituted urea of methylamine (entry 7 & 8) and aniline (entries 4–6) are extensively

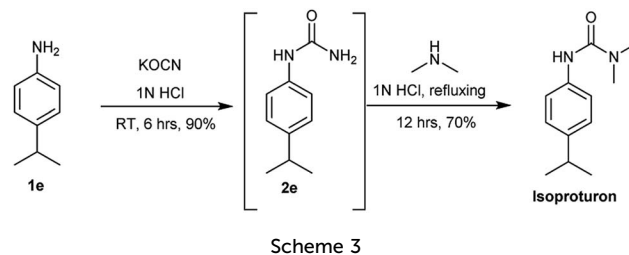


Table 3 Large scale preparation of *N*-substituted urea

Entry	Product	Scale (mmol)	Time (h)	Isolated yield (%)
1		1	4	88
2		10	3	95
3		20	2	99
4		1	6	90
5		10	5	89
6		20	5	90
7		1	5	74
8		20	5	78
9		1	12	71
10		20	12	72

being used as key raw material for various target molecules.²⁰ In all the scale up reactions of four different amines (Table 3), we found no detrimental effect on rate of reaction or overall isolated yields. In fact, in case of 4-isopropyl aniline, reaction gave considerably better isolated yield at 20 mmol scale compared to 1 or 10 mmol batch size. It is important to mention that a high level of chemical purity of the final product was obtained in 20 mmol reaction just by simple filtration without any additional workup or column chromatographic purification.

The industrial applicability of the developed protocol was further established by executing a one-pot synthesis of Isoproturon in water as mentioned in Scheme 3. Isoproturon is an agrochemical which is produced and used in large quantities. The *N*-substituted urea **2e** produced by developed methodology was used as a key intermediate and reacted *in situ* with dimethyl amine (2 molar THF solution) in same pot to afford isoproturon as major product. This is the first reported one-pot synthesis of Isoproturon using water as a solvent avoiding catalyst or isocyanate chemistry.²¹



Experimental

General information

Reagents and starting materials were obtained from commercial sources and used as received. All reactions were performed under an inert atmosphere. Flash chromatography was carried out using silica gel (100–200 mesh). Thin layer chromatography was performed on silica gel and was visualized by staining with KMnO_4 or iodine vapours. The ^1H and ^{13}C NMR spectra were recorded using a Bruker Avance 300 (300 MHz for ^1H and 75 MHz for ^{13}C) spectrometer in CDCl_3 or $\text{DMSO}-d_6$ or $\text{acetone}-d_6$ using TMS as an internal standard. The chemical shifts (δ) for ^1H and ^{13}C are given in ppm relative to residual signals of the solvent. Coupling constants are given in Hz. The following abbreviations are used to indicate the multiplicity: s, singlet; d, doublet; t, triplet; q, quartet; m, multiplet; brs, broad signal. The reactions were monitored by TLC (Merck). Evaporation of solvents was performed under reduced pressure using a rotary evaporator. Mass spectra were recorded on Advion Mass Express CMS instrument. Most of the compounds reported in this manuscript are known, and accordingly referred for their analytical data for characterization. For these known compounds, only ^1H NMR data is provided along with physical characteristics. However, for new molecules, ^1H , ^{13}C NMR and mass spectrometry data is provided along with their physical characterization.

General procedures for Scheme 1 and optimization studies of Table 1

To a stirring solution of 1 mmol amine (**1a** or **1b**) separately in 3 mL solvent (4 M dioxane HCl and co-solvent if any as mentioned in Table 1), 2.2 equivalent of potassium cyanate was added. The reaction was allowed to stir for specific time at room temperature. The progress of reaction was monitored by TLC. The reaction aliquot was extracted and dried under vacuum. The ^1H -NMR analysis of this vacuum dried aliquot was done to determine the quantification of the product.

Procedures for synthesis of 1-phenylurea (**2a**) from aniline (**1a**)

To a stirring solution of aniline (**1a**, 0.18 mL; 2 mmol), and 1 N aqueous HCl (3 mL), potassium cyanate (4.4 mmol; 357 mg; 2.2 equiv.) was added. The reaction mixture was allowed to stir at room temperature for 6 hours. The progress of the reaction was monitored by TLC. The precipitates thus formed at the end of reaction time was filtered off and washed with aqueous 1 N HCl (5 mL). The solid residue was air dried and collected as white



solid of 1-phenylurea^{18a} (250 mg, 94%); ¹H NMR (300 MHz, DMSO-*d*₆): δ 5.81 (s, 2H, NH), 6.88 (t, *J* = 15 Hz, 1H), 7.21 (t, *J* = 15 Hz, 2H), 7.38 (d, *J* = 6 Hz, 2H), 8.49 (s, 2H).

The compounds from Table 2 (entries 2–16, 19, 20, 36, 39, 40) were synthesized based on this method. The analytical data of these compounds is mentioned below:

Entry 2; 1-(4-chlorophenylphenyl)urea.^{18a} The product was isolated as white solid (308 mg, 90%); ¹H NMR (300 MHz, DMSO-*d*₆): δ 8.66 (s, 1H), 7.43–7.39 (m, 2H), 7.27–7.23 (m, 2H), 5.89 (s, 2H, NH).

Entry 3; 1-(3,4 dichlorophenyl)urea. The product was isolated as white solid (290 mg, 71%); ¹H NMR (300 MHz, DMSO-*d*₆): δ 8.75 (s, 1H, NH), 7.80–7.77 (m, 1H), 7.21–7.24 (m, 2H), 5.96 (s, 2H, NH). ¹³C-NMR (300 MHz, DMSO-*d*₆): δ 156.19, 141.28, 131.31, 130.28, 122.66, 119.12, 118.19. APCI MS: (M, M + 2), 204.9, 207.0, 163.9, 161.9, 129, 127.

Entry 4; 1-(2-bromophenyl)urea. The product was isolated as white solid (400 mg, 93%); ¹H NMR (300 MHz, DMSO-*d*₆): δ 8.02–8.06 (m, 1H), 7.85 (s, 1H, NH), 7.52 (d, *J* = 6 Hz, 1H), 7.27 (t, *J* = 18 Hz, 1H), 6.90 (m, 1H), 6.42 (s, 2H, NH). ¹³C-NMR (300 MHz, DMSO-*d*₆): δ 156.04, 138.32, 132.75, 128.34, 123.75, 122.41, 112.87. APCI MS: 214.9 (M+), 216.9 (M + 2), 173.9, 171.9, 170.9, 169.9.

Entry 5; 1-(3-bromophenyl)urea.^{18a} The product was isolated as white solid (352 mg, 82%); ¹H NMR (300 MHz, DMSO-*d*₆): δ 8.76 (s, 1H, NH), 7.85 (t, *J* = 3 Hz, 1H), 7.22–7.26 (m, 1H), 7.18 (d, *J* = 6 Hz, 1H), 7.07–7.11 (m, 1H), 5.98 (s, 2H, NH).

Entry 6; 1-(4-bromophenyl)urea.²² The product was isolated as light brown solid (344 mg, 80%); ¹H NMR (300 MHz, CD₃OD): δ 7.38 (d, *J* = 9 Hz, 2H), 7.34 (d, *J* = 9 Hz, 2H).

Entry 7; 1-(3-trifluoromethoxyphenyl)urea. The product was isolated as white solid (350 mg, 80%); ¹H NMR (300 MHz, DMSO-*d*₆): δ 8.87 (s, 1H, NH), 7.94 (s, 1H), 7.50–7.39 (m, 2H), 7.19 (d, *J* = 9 Hz, 1H), 5.98 (s, 2H, NH). ¹³C-NMR (300 MHz, DMSO-*d*₆): δ 156.23, 149.16, 142.84, 130.61, 116.68, 113.29, 110.11. APCI MS: (M + 1) 221.0, 204, 178, 135, 122, 107.

Entry 8; 1-(3-trifluoromethylphenyl)urea. The product was isolated as white solid (256 mg, 63%); ¹H NMR (300 MHz, DMSO-*d*₆): δ 8.83 (s, 1H, NH), 7.63 (s, 1H), 7.31 (t, *J* = 15 Hz, 1H), 7.20 (d, *J* = 6 Hz, 1H), 6.44 (d, *J* = 6 Hz, 1H), 5.96 (s, 2H, NH). ¹³C-NMR (300 MHz, DMSO-*d*₆): δ 156.32, 141.87, 122.94, 121.64, 117.72, 114.05. APCI MS: (M + 1) 205.0, 162, 142, 93.

Entry 9; 1-(3, 5-dimethylphenyl)urea. The product was isolated as white solid (320 mg, 97%); ¹H NMR (300 MHz, DMSO-*d*₆): δ 8.32 (s, 1H, NH), 7.00 (s, 2H), 6.53 (s, 2H), 5.77 (s, 2H, NH), 2.19 (s, 6H). ¹³C-NMR (300 MHz, DMSO-*d*₆): δ 156.43, 140.84, 137.91, 123.15, 115.92, 21.61. APCI MS: (M + 1) 165.1, 148, 133.1, 122.1.

Entry 10; 1-(4-isopropylphenyl)urea. The product was isolated as white solid (316 mg, 88%); ¹H NMR (300 MHz, DMSO-*d*₆): δ 8.40 (s, 1H, NH), 7.30 (d, *J* = 9 Hz, 2H), 7.09 (d, *J* = 9 Hz, 2H), 5.77 (s, 2H, NH), 2.84–2.79 (m, 1H), 1.18 (d, *J* = 6 Hz, 6H). ¹³C-NMR (300 MHz, DMSO-*d*₆): δ 156.53, 141.51, 138.71, 126.73, 118.40. APCI MS: (M + 1) 179.2, 162, 137.1, 120.1.

Entry 11; 1-(2-hydroxyphenyl)urea.² The product was isolated as white solid (275 mg, 90%); ¹H NMR (300 MHz, DMSO-*d*₆):

δ 9.91 (s, 1H, NH), 7.99 (s, 1H), 7.84–7.81 (m, 1H), 6.8–6.7 (m, 3H), 6.19 (s, 2H, NH).

Entry 12; 1-(3-hydroxyphenyl)urea. The product was isolated as pale yellow solid (203 mg, 67%); ¹H NMR (300 MHz, DMSO-*d*₆): δ 9.22 (s, 1H, NH), 8.4 (s, 1H), 7.0–6.9 (m, 2H), 6.71–6.7 (m, 1H), 6.34–6.30 (m, 1H), 5.79 (s, 2H, NH). ¹³C-NMR (300 MHz, DMSO-*d*₆): δ 158.09, 156.36, 142.05, 129.67, 108.70, 105.27. APCI MS: (M + 1) 153, 148.9, 134, 121, 110.

Entry 13; 1-(4-hydroxyphenyl)urea.²⁴ The product was isolated as white solid (270 mg, 89%); ¹H NMR (300 MHz, DMSO-*d*₆): δ 8.94 (s, 1H), 8.51 (s, 1H, NH), 7.15 (d, *J* = 9 Hz, 2H), 6.63 (d, *J* = 9 Hz, 2H), 5.64 (s, 2H, NH).

Entry 14; 4-ureidobenzoic acid.²⁵ The product was isolated as white solid (yield: 300 mg, 83%); ¹H NMR (300 MHz, DMSO-*d*₆): δ 9.14 (s, 1H, NH), 7.72 (d, *J* = 6 Hz, 2H), 7.29 (d, *J* = 6 Hz, 2H), 6.02 (m, 3H), 6.19 (s, 2H, NH), 5.50 (s, 1H).

Entry 15; 1-(4-cyanophenyl)urea. The product was isolated as white solid (150 mg, 47%); ¹H NMR (300 MHz, DMSO-*d*₆): δ 9.02 (s, 1H, NH), 7.73–7.61 (m, 2H), 7.58–7.55 (m, 2H), 6.11 (s, 2H, NH). ¹³C-NMR (300 MHz, DMSO-*d*₆): δ 155, 145, 133, 119, 117, 102. APCI MS: (M + 1) 162.0, 142, 119.1.

Entry 16; 4-methoxyphenyl urea.²⁶ The product was isolated as brown solid (234 mg, 70%); ¹H NMR (300 MHz, DMSO-*d*₆): δ 8.29 (s, 1H, NH), 7.26–7.29 (m, 2H), 6.78–6.81 (m, 2H), 5.70 (s, 2H, NH), 3.68 (s, 3H).

Entry 19; 2-(4-ureidophenyl)acetic acid. The product was isolated as off white solid (315 mg, 81%); ¹H NMR (300 MHz, DMSO-*d*₆): δ 8.99 (s, 1H, NH), 7.33 (d, *J* = 9 Hz, 2H), 7.08 (d, *J* = 6 Hz, 2H), 5.97 (s, 2H, NH), 5.50 (s, 1H, NH), 3.38 (s, 2H, NH). ¹³C-NMR (300 MHz, DMSO-*d*₆): δ 173.46, 156.50, 139.55, 130.30, 127.97, 118.14. APCI MS: (M + 1) 195.1, 178.0.

Entry 20; 4-(ureidomethyl)benzoic acid. The product was isolated as off white solid (285 mg, 73%); ¹H NMR (300 MHz, DMSO-*d*₆): δ 7.91 (d, *J* = 6 Hz, 2H), 7.34 (d, *J* = 9 Hz, 2H), 6.57 (t, *J* = 12 Hz, NH), 5.67 (s, 2H, NH), 4.25 (d, *J* = 6 Hz, 2H). ¹³C-NMR (300 MHz, D₂O): δ 159.40, 129.79, 126.99, 42.97. APCI MS: (M + 1) 194.9, 153, 135, 107.

Entry 36; N-carbamoyl-phenylalanine.³² The product was isolated as off white solid (260 mg, 63%); ¹H NMR (300 MHz, DMSO-*d*₆): δ 7.31–7.17 (m, 5H), 6.13 (d, *J* = 6 Hz, 1H), 5.60 (s, 2H, NH), 4.43–4.28 (m, 1H), 3.02–2.85 (m, 1H), 2.83–2.49 (m, 1H).

Entry 39; 1-(1H benzo[d]imidazol-2-yl)urea. The product was isolated as yellow solid (248 mg, 70%); ¹H NMR (300 MHz, DMSO-*d*₆): δ 7.11–7.14 (m, 2H), 6.87–6.90 (m, 2H), 6.26 (s, 2H, NH), 5.43 (s, 1H, NH). ¹³C-NMR (300 MHz, DMSO-*d*₆): δ 155, 138.27, 119.84, 111.94, 21.86. APCI MS: 159.9, 134.0, 92.1.

Entry 40; 1-(H-pyrazol-3-yl)-urea. The product was isolated as off white solid (177 mg, 70%). ¹H NMR (300 MHz, DMSO-*d*₆ + D₂O), δ 8.08 (d, *J* = 2.7 Hz, 1H), 6.27 (s, 1H), 6.26 (s, 2H, NH), 5.43 (s, 1H, NH). ¹³C-NMR (300 MHz, DMSO-*d*₆): δ 155, 138.27, 119.84, 111.94, 21.86. APCI MS: 159.9, 134.0, 92.1, 127 (M + 1), 110.1, 82.

Procedures for synthesis of N-benzyl urea (2b) from benzylamine (2a)

To a stirring solution of benzylamine (**1b**, 2 mmol), and 1 N aqueous HCl (3 mL), potassium cyanate (4.4 mmol; 2.2 equiv.) was added. The reaction mixture was allowed to stir at room



temperature for 6 hours. The progress of the reaction was monitored by TLC. The reaction mixture was diluted with 1 N aqueous HCl (3 mL) and extracted with dichloromethane (20 mL). Organic layer was washed with water (5 mL), dried over anhydrous sodium sulphate and concentrated under reduced pressure to afford pure as white solid (270 mg, 89%). ^1H NMR (300 MHz, $\text{DMSO}-d_6$): δ 7.34–7.29 (m, 2H), 7.29–7.28 (m, 3H), 6.40 (t, J = 12 Hz, 1H), 5.51 (s, 2H, NH), 4.17 (d, J = 6 Hz, 2H).²²

Compound from Table 2 (entries 17, 21–25, 27–29, 31–35, 37, 38) were prepared by this method.

Entry 21; *N*-methyl urea.²⁷ The product was isolated as brown viscous oil (110 mg, 74%); ^1H NMR (300 MHz, $\text{DMSO}-d_6$): δ 5.86 (s, 1H, NH), 5.45 (s, 2H, NH), 2.52 (s, 3H).

Entry 22; *N*-butyl urea.^{18a} The product was isolated as white solid (120 mg, 52%); ^1H NMR (300 MHz, $\text{DMSO}-d_6$): δ 5.94 (s, 1H), 5.39 (s, 2H, NH), 3.01–2.94 (m, 2H), 1.37–1.28 (m, 2H), 0.93–0.90 (d, J = 6 Hz, 2H).

Entry 23; chloroethylamine urea. The product was isolated as off white solid and found to be a mixture with starting amine. Attempts to purify this mixture by column chromatography were not successful. Conversion based on ^1H NMR 72%: ^1H NMR (300 MHz, D_2O): δ 7.91 (d, J = 6 Hz, 2H), 7.34 (d, J = 9 Hz, 2H), 6.57 (t, J = 12 Hz, NH), 5.63 (s, 2H, NH), 4.25 (d, J = 6 Hz, 2H).

The above compound was not purified by column chromatography, here we have reported conversion through NMR.

Entry 24; 1-(2-mercaptoethyl)urea. The product was isolated as viscous oil and found to be a mixture with starting amine. Attempts to purify this mixture by column chromatography were not successful. Conversion based on ^1H NMR: (68%); ^1H NMR (300 MHz, D_2O): 3.28 (t, J = 12 Hz, 2H), 3.17 (t, J = 15 Hz, 1H), 2.69 (t, J = 15 Hz, 2H). APCI MS: 121.0 (M + 1), 104, 91, 88.1, 70.1.

Entry 25; ethanolamine urea.²³ The product was isolated as viscous oil (140 mg, 68%); ^1H NMR (300 MHz, $\text{DMSO}-d_6$): δ 6.04 (s, 1H, NH), 5.48 (s, 2H, NH), 6.57 (t, J = 12 Hz, NH), 5.63 (s, 2H, NH), 4.25 (d, J = 6 Hz, 2H).

Entry 26; isopropylurea.^{15b} The product was isolated as off white solid (90 mg, 44%); ^1H NMR (300 MHz, $\text{DMSO}-d_6$): δ 5.79 (d, J = 6 Hz, 1H), 5.31 (s, 2H, NH), 3.65–3.67 (m, 1H), 1.04 (d, J = 6 Hz, 6H).

Entry 27; 1,1 dimethyl urea.^{18a} The product was isolated as off white solid (125 mg, 71%); ^1H NMR (300 MHz, $\text{DMSO}-d_6$): δ 5.75 (s, 2H, NH), 2.7 (s, 6H).

Entry 28; 1,1 dimethyl urea.^{18a} The product was isolated as off white solid (320 mg, 91%); ^1H NMR (300 MHz, $\text{DMSO}-d_6$): δ 5.75 (s, 2H, NH), 2.7 (s, 6H).

Entry 29; 1,1 diethyl urea. The product was isolated as light brown solid (212 mg, 73%); ^1H NMR (300 MHz, $\text{DMSO}-d_6$): δ 5.66 (s, 2H, NH), 2.49–2.48 (m, H), 0.97 (t, J = 14 Hz, 6H). ^{13}C -NMR (300 MHz, $\text{DMSO}-d_6$): δ 158, 40.24, 14.28. APCI MS: (M + 1) 117.1.

Entry 30; cyclohexyl urea.²⁸ The product was isolated as white solid (200 mg, 70%); ^1H NMR (300 MHz, $\text{DMSO}-d_6$): δ 5.80 (d, J = 6 Hz, 1H, NH), 5.27 (s, 2H, NH), 3.26 (m, 1H), 1.74–1.60 (m, 4H), 1.59–1.48 (m, 1H), 1.26–1.16 (m, 5H).

Entry 31; piperidine-1-carboxylic acidamide.²⁹ The product was isolated as white solid (153 mg, 60%); ^1H NMR (300 MHz,

$\text{DMSO}-d_6$): δ 5.82 (s, 2H, NH), 7.34 (t, J = 9 Hz, 4H), 1.52–1.46 (m, 2H), 1.41–1.30 (m, 4H).

Entry 32; morpholine-4-carboxamide.³⁰ The product was isolated as white solid (195 mg, 75%); ^1H NMR (300 MHz, $\text{DMSO}-d_6$): δ 6.0 (s, 2H, NH), 3.55 (t, J = 9 Hz, 4H), 3.26 (t, J = 9 Hz, 4H).

Entry 33; *N*-carbamoyl-glycine.³¹ The product was isolated as brown viscous oil (160 mg, 68%); ^1H NMR (300 MHz, $\text{DMSO}-d_6$): δ 8.20 (s, 2H, NH), 5.50 (t, J = 12 Hz, 2H), 3.7 (d, J = 6 Hz, 2H).

Entry 34; *N*-carbamoyl-valine.³² The product was isolated as viscous oil (170 mg, 53%); ^1H NMR (300 MHz, D_2O): δ 3.77 (s, 1H), 2.01–1.94 (m, 1H), 0.82 (d, J = 6 Hz, 3H), 0.75 (d, J = 9 Hz, 3H).

Entry 35; *N*-carbamoyl-glutamic acid.³² The product was isolated as off white solid (yield: 70 mg, 18%); ^1H NMR (300 MHz, D_2O): δ 4.24–4.29 (m, 2H), 2.40–2.47 (m, 1H), 2.27–2.32 (m, 2H), 2.03–2.08 (m, 1H).

Entry 37; *N*-carbamoyl-tyrosine.³² The product was isolated as off white solid (360 mg, 80%); ^1H NMR (300 MHz, D_2O): δ 6.95 (d, J = 6 Hz, 2H), 6.58 (d, J = 6 Hz, 2H), 4.08 (t, J = 12 Hz, NH), 2.93–2.87 (m, 1H), 2.74–2.67 (m, 1H).

Entry 38; methyl-2-ureidothiophene-2-carboxylate. The product was isolated as white solid (120 mg, 30%). ^1H NMR (300 MHz, $\text{DMSO}-d_6$): δ 9.23 (s, 1H, NH), 7.96 (d, J = 6 Hz, 2H), 7.79 (d, J = 3 Hz, 2H), 6.79 (s, 2H, NH), 3.84 (s, 3H). ^{13}C -NMR (300 MHz, $\text{DMSO}-d_6$): δ 164.07, 155.04, 147.01, 132.82, 122.34, 106.41, 52.16. APCI MS: 151.9, 126.0, 98.

Specific procedure for Scheme 2

To a stirring solution of aniline (**1a**; 0.09 mL; 1 mmol) and morpholine (**1c**; 0.086 mL, 1 mmol), was added 1 N aqueous HCl (3 mL). To this stirring reaction mixture, 1.1 mmol of potassium cyanate (89 mg) was added. The reaction mixture was stirred at room temperature for 6 hours. An aliquot were extracted after 4 hours and analysed for conversion of amines to corresponding *N*-substituted ureas using HPLC analysis and ^1H NMR spectroscopy. The reaction mixture was filtered off after 6 hours to collect off white precipitates of phenyl urea (**2b**). The collected white solid was washed with 1 N HCl (5 mL). The collected solid as well as filtrate was analysed using HPLC and ^1H NMR spectroscopy for quantification of *N*-substituted urea and left over amine. Other reactions of Scheme 2 were done using similar protocol.

Specific procedures for Table 3 for 20 mmol scale of phenyl urea (**1a**, entry 6, Table 3)

To a stirring solution of aniline (**1a**, 20 mmol, 1.83 mL), and 1 N aqueous HCl (30 mL), potassium cyanate (44.0 mmol; 2.2 eq., 3.56 g) was added. The reaction mixture was allowed to stir at room temperature for 6 hours. The progress of the reaction was monitored by TLC. The precipitates thus formed at the end of reaction time was filtered off and washed with aqueous 1 N HCl (50 mL). The solid residue was air dried and collected as white solid of 1-phenylurea (yield: 2.4 g, 90%).



One pot synthesis of isoproturon²¹ (Scheme 3)

4-Isopropylaniline (*p*-aminocumene; **1e**; 1 mmol, 178 mg) was mixed in 1 N aqueous HCl (3 mL) and potassium cyanate (2.2 mmol, 178 mg) was added. The reaction mixture was stirred at room temperature and reaction was monitored using TLC for maximum consumption of aniline (6 hours). Further, dimethylamine (3 mmol) was added to reaction mixture and the reaction was refluxed for additional 12 hours. The progress of the reaction was monitored by TLC. Reaction mixture was cooled to room temperature, diluted with water (5 mL) and extracting with dichloromethane (25 mL). The organic layer was collected, dried over anhydrous sodium sulphate and concentrated under vacuum to afford crude product. The crude was purified by column chromatography using ethyl acetate and hexane as eluent to afford pure product as light yellow solid (175 mg; 70%). ¹H NMR²¹ (300 MHz, CDCl₃), δ 7.23 (d, *J* = 9 Hz, 2H), 7.09 (d, *J* = 9 Hz, 2H), 6.24 (s, 2H, NH), 2.98 (s, 6H), 2.75–2.86 (m, 1H), 1.18 (d, *J* = 6 Hz, 6H).

Conclusion

We demonstrated a green method for the synthesis of *N*-substituted ureas, which are molecules of commercial interests and being produced and used in large quantities. The developed approach is mild, does not require any heating or pressure and reactions can be performed using water as a solvent without any organic co-solvent. In fact, water as a solvent found to provide a unique reactivity and selectivity under identified conditions. The developed methodology demonstrated a large substrate scope and practical ease for product isolation. The reactions can be done at gram scale using this methodology without any effect on yields. All these characters make this approach as an industrial amenable approach for manufacturing of *N*-substituted ureas.

Conflicts of interest

There are no conflicts to declare.

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