Regulating the topology of 2D covalent organic frameworks by the rational introduction of substituents†

Zhong-Fu Pang, Tian-You Zhou, Rong-Ran Liang, Qiao-Yan Qi and Xin Zhao

The topology of a covalent organic framework (COF) is generally believed to be dictated by the symmetries of the monomers used for the condensation reaction. In this context, the use of monomers with different symmetries is usually required to afford COFs with different topologies. Herein, we report a conceptual strategy to regulate the topology of 2D COFs by introducing alkyl substituents into the skeleton of a parent monomer. The resulting monomers, sharing the same C2 symmetry, were assembled with a D2h symmetric tetraamine to generate a dual-pore COF or single-pore COFs, depending on the sizes of the substituents, which were evidenced using PXRD studies and pore size distribution analyses. These results demonstrate that the substituent is able to exert a significant influence on the topology of COFs, which is crucial for their application.
ethyl (2,5-diethoxyterephthalaldehyde, DETA) and n-butyl (2,5-dibutoxyterephthalaldehyde, DBTA). CPK modeling suggested that the size of a triangular pore in the DP-COF was big enough to accommodate three hydroxy groups but it became highly congested when the ethoxy groups were introduced (Fig. S1, ESI†). In the case of the butoxy groups, these were too big to be accommodated by the triangular pores. The experimental results indicate that, while the condensation of DHTA and ET TA gave rise to a COF with a dual-pore structure, the reactions of DETA or DBTA with ET TA led to the formation of COFs with a single-pore topology (Scheme 1). These results reveal that a substituent is able to exert significant influence on the topology of COFs, which has never been observed before.

**Results and discussion**

The condensation reactions were carried out by heating the mixtures of ET TA and the corresponding dialdehyde (1 : 2) in 1,4-dioxane-acetic acid (aq., 6 M) (1/0.1, v/v) in sealed glass ampoules at 120 °C for 72 h and the products were obtained as powders (for details, see the ESI†), which were named COF-DHTA (from DHTA and ET TA), COF-DETA (from DETA and ET TA) and COF-DBTA (from DBTA and ET TA). The as-obtained powders were insoluble in water and common organic solvents, and their colors varied from red to yellow with the increasing chain length of the alkyl substituents. The stretching vibrational band of C≡N was observed at 1619 cm\(^{-1}\) in Fourier transform infrared (FT-IR) spectra of the powders, which confirmed the existence of C≡N linkages in these materials (Fig. S2, ESI†). Furthermore, the vibration of the NH\(_2\) group in ET TA (around 3300 cm\(^{-1}\)) was largely attenuated after the condensation reactions, indicating a high degree of polymerization for the polymers. Several peaks in the range of 2850–3000 cm\(^{-1}\) were also observed for COF-DETA and COF-DBTA but not for COF-DHTA, which is consistent with the fact that in the first two materials the alkyl groups were incorporated into their skeletons. The powders were further examined with solid-state \(^{13}\)C cross polarization/magic angle spinning nuclear magnetic resonance (CP/MAS NMR) spectroscopy, which displayed the characteristic \(^{13}\)C resonance signal of C≡N that appeared at 163 ppm for COF-DHTA and at 152 ppm for the other two polymers, which again corroborated the formation of the imine-based polymers (Fig. S3–S5, ESI†). In addition to the signals of C≡N and the aromatic carbon atoms, the peaks of the alkyl groups were also identified in the range of 10–70 ppm for COF-DETA and COF-DBTA, clearly indicating the presence of the alkyl substituents in the polymers. Elemental analyses were also performed for these polymers, which revealed that their C, H, and N content was close to that of the corresponding calculated values (see the ESI† for details). As indicated by the thermogravimetric analyses, the as-prepared powders exhibited excellent thermal stabilities. Less than 8% weight losses were observed upon heating them to 400 °C (Fig. S6, ESI†). Scanning electron microscopy (SEM) revealed that they all exhibited irregular morphologies (Fig. S7, ESI†).

In the next step, theoretical simulations and powder X-ray diffraction (PXRD) experiments were carried out to determine the exact structures of the as-obtained powders (Fig. 1–3 and S8–S10, ESI†). In the simulations, two possible crystalline structures, that is, DP-COF and SP-COF, were constructed. Each crystalline structure was further divided into two stacking models, namely eclipsed packing (AA) and staggered packing (AB). The DP-AA and DP-AB structures were constructed with unit cell parameters of \(a = b = 37.9\ \text{Å}, c = 4.5\ \text{Å}\) (for DP-AA) or \(9.0\ \text{Å}\) (for DP-AB), \(a = \beta = 90°\), and \(\gamma = 120°\). Similarly, the SP-AA and SP-AB structures were constructed with unit cell parameters of \(a = 29.2\ \text{Å}, b = 25.4\ \text{Å}, c = 4.5\ \text{Å}\) (for SP-AA) or \(9.0\ \text{Å}\) (for SP-AB), and \(a = \beta = \gamma = 90°\) (see Fig. S9–S12† for details). The powder prepared from DHTA and ET TA was subjected to PXRD analysis and it gave a PXRD pattern similar to that of the previously

![Scheme 1](Image)
reported dual-pore COF (Fig. 1a). A close comparison of the experimental PXRD pattern with the simulated ones strongly suggests that the powder had a dual-pore structure and eclipsed packing (Fig. 1). Firstly, it exhibits a strong diffraction intensity, indicating the good crystallinity of the powder. Secondly, the experimental PXRD pattern was in good agreement with the simulated PXRD pattern generated from the proposed DHTA-based DP-AA structure. In the experimental PXRD pattern of COF-DHTA, diffraction peaks at 2.72°, 4.83°, 5.40°, 8.10°, and ca. 19.7° are observed, which can be assigned to (100), (110), (200), (300) and (001) diffractions, respectively (Fig. 1a and c). The peak positions and relative intensities could be well reproduced by the simulated PXRD pattern of a DP-COF with AA stacking. In contrast, the simulated PXRD patterns of the structure with single-pore topology exhibited significant differences from the experimentally observed PXRD pattern, suggesting again that the condensation of DHTA and ETTA gave rise to a dual-pore COF, not a single-pore COF. A Pawley refinement reproduced the experimentally observed PXRD peaks quite well and yielded unit cell parameters of $a = b = 37.48$ Å, $c = 4.52$ Å, $\alpha = \beta = 90.0^\circ$, and $\gamma = 120.0^\circ$, with $R_p = 3.50\%$ and $R_{wp} = 4.70\%$.

COF-DETA and COF-DBTA exhibited experimental PXRD patterns that were similar to each other. Two broad peaks at 5.0° and 9.9° were observed in the experimental PXRD pattern of COF-DETA (Fig. 2a). Obviously it did not match the simulated PXRD patterns for a dual-pore COF with AA or AB packing (comparing Fig. 2a with 2e and 2f). This result strongly suggested that the condensation of DETA and ETWA did not generate a COF bearing a similar topology as the above dual-pore COF-DHTA. The experimentally observed PXRD pattern was thereupon compared with the ones simulated for a single-pore COF. The simulated PXRD pattern of the DETA-based structure with single-pore topology and AA packing has two major diffraction peaks at $4.62^\circ$ and $9.24^\circ$ (Fig. 2b), which are assignable to (110) and (220) diffractions, respectively. At first glance, it seems that the experimental and simulated patterns do not agree very well with each other since the 2-theta values of the diffraction peaks in the experimental PXRD pattern are a little bit larger than those of the simulated one. Further analysis showed that this phenomenon is consistent with previous studies on alkyl modified COFs, in which the addition of alkyl substituents resulted in a broadening of the diffraction peaks, and a continuous shift of the peaks to a large 2-theta value was also observed with the increase of the alkyl chain length, in comparison with the unsubstituted parent COF. This result was attributed to the presence of alkyl groups.
which freely rotated within the pores. Therefore, taking this influence of the alkyl substituents on the PXRD of the COF into account, we could conclude that a single-pore COF with eclipsed packing was produced from the condensation of DETA and DHTA. A Pawley refinement was performed, which gave unit cell parameters of \( a = 29.19 \) Å, \( b = 25.40 \) Å, \( c = 4.50 \) Å, \( \alpha = \beta = 90.0^\circ \), and \( \gamma = 90.0^\circ \), with \( R_p = 2.01\% \) and \( R_{wp} = 2.48\% \). The difference plot between the experimental and the refined XRD patterns indicates that they match each other quite well.

Similarly, the simulated PXRD pattern of the DBTA-based structure with a single-pore topology and AA packing has two major peaks at 4.62° and 9.24°, which are assignable to (110) and (220) diffractions, respectively. The experimental PXRD pattern of COF-DBTA displays two broad diffraction peaks at 5.4° and 10.9° (Fig. 3). A comparison of the experimental PXRD pattern of COF-DBTA with the simulated ones also indicated that it did not match with a dual-pore topology but could match with a single-pore structure when the influence of the butyls on the PXRD pattern was taken into consideration. Therefore, a single-pore COF structure with AA packing was assigned to COF-DBTA. A Pawley refinement gave optimized unit cell parameters of \( a = 29.26 \) Å, \( b = 25.29 \) Å, \( c = 4.50 \) Å, \( \alpha = 90.16^\circ \), \( \beta = 90.16^\circ \), and \( \gamma = 91.44^\circ \), with \( R_p = 2.72\% \) and \( R_{wp} = 3.74\% \).

The assigned structures for COF-DHTA, COF-DETA, and COF-DBTA were further corroborated using nitrogen adsorption–desorption measurements. The \( N_2 \) isotherm of COF-DHTA exhibited good reversibility. It did not fit a typical type I model but a hybrid of type I and type IV sorption isotherms (Fig. 4a),11 suggesting that micropores and mesopores coexist in COF-DHTA. The Brunauer–Emmett–Teller (BET) model was applied to the isotherm with \( P/P_0 \) in the range of 0.05–0.3, which generated a BET surface area of 1869.32 m² g⁻¹ for COF-DHTA (Fig. S11, ESI†). Its theoretical maximum BET surface area was calculated using the Monte Carlo Metropolis method in Materials Studio,14 which yielded a theoretical surface area of 2103.89 m² g⁻¹ (Fig. S12, ESI†). This value is quite close to its Connolly surface area (2274.33 m² g⁻¹). The total pore volume of COF-DHTA was calculated to be 0.90 cm³ g⁻¹ (\( P/P_0 = 0.99 \)). Its pore size distribution curve was generated using nonlocal density functional theory (NLDFT). The two main distributions were observed around 6.3 Å and 26.5 Å, respectively, indicating that two kinds of pores were present in the material (Fig. 4b). These values are very close to the theoretical pore sizes of the dual-pore COF (6.3 and 25.2 Å as estimated by PM3 calculations), confirming again that COF-DHTA holds a dual-pore topology. COF-DETA and COF-DBTA displayed similar \( N_2 \) isotherms (Fig. 4c and e). On the basis of their \( N_2 \) isotherm data, their BET surface areas were calculated to be 458.11 and 379.04 m² g⁻¹, respectively (Fig. S13 and S14, ESI†). We failed to calculate their theoretical maximum BET surface areas due to the fact that no suitable model was available for them. The experimental pore size distribution of COF-DHTA shows a narrow peak around 8.4 Å, which is close to the theoretical value of the corresponding SP-COF structure (9.1 Å) (Fig. 4d). Different to COF-DETA, in COF-DBTA the butyl chains are long enough to divide a single rhombus-shaped pore into four parts with very similar pore sizes, which were theoretically estimated to be 6.7 Å (for two of the pores) and 7.0 Å (for the other two pores) (Fig. 4f). Fitting the isotherm data of COF-DBTA using NLDFT revealed a main pore distribution around 6.3 Å, which closely matched with the theoretically predicted pore size. These pore size distribution analyses further confirm that both COF-DETA and COF-DBTA hold a single-pore topology.

Conclusions

In summary, switching the topology of COFs has been successfully achieved through the rational introduction of alkyl substituents into the skeletons of the monomers. Although these monomers share the same symmetry, the presence of different alkyl groups exerts different influences on the formation of the polymeric structures. While a dual-pore COF was obtained from an unsubstituted dialdehyde, the introduction of ethyl or \( n \)-butyl groups resulted in COFs with a single-pore topology. The formation of single-pore COFs can be explained by the alleviation of steric repulsion between the substituents. This proof of concept opens up a new route to regulate the topologies of COFs, which in turn will modify their properties and give them new functions.

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Notes and references


8 (a) T. Ben and S. Qiu, CrystEngComm, 2013, 15, 17.


