Macrocyclic shape-persistency of cyclo[6]aramide results in enhanced multipoint recognition for the highly efficient template-directed synthesis of rotaxanes†

Xiaowei Li,†a Xiangyang Yuan,†a Pengchi Deng,†a Lixi Chen,†a Yi Ren,†a Chengyuan Wang,†a Lixin Wu,b Wen Feng,a Bing Gong† and Lihua Yuan†a

Examples of using two-dimensional shape-persistent macrocycles, i.e. those having noncollapsible and geometrically well-defined skeletons, for constructing mechanically interlocked molecules are scarce, which contrasts the many applications of these macrocycles in molecular recognition and functional self-assembly. Herein, we report the crucial role played by macrocyclic shape-persistency in enhancing multipoint recognition for the highly efficient template-directed synthesis of rotaxanes. Cyclo[6]aramides, with a near-planar conformation, are found to act as powerful hosts that bind bipyridinium salts with high affinities. This unique recognition module, composed of two macrocyclic molecules with one bipyridinium ion thread through the cavity, is observed both in the solid state and in solution, with unusually high binding constants ranging from $\sim 10^{-13} \text{ M}^{-2}$ to $\sim 10^{-15} \text{ M}^{-2}$ in acetone. The high efficacy of this recognition motif is embodied by the formation of compact [3]rotaxanes in excellent yields based on either a ‘click-capping’ (91%) or ‘facile one-pot’ (85%) approach, underscoring the great advantage of using H-bonded aromatic amide macrocycles for the highly efficient template-directed synthesis of mechanically interlocked structures. Furthermore, three cyclo[6]aramides bearing different peripheral chains 1–3 demonstrate high specificity in the synthesis of a [3]rotaxane from 1 and 2, and a [2]rotaxane from 3 via a ‘facile one-pot’ approach, in each case as the only isolated product. Analysis of the crystal structure of the [3]rotaxane reveals a highly compact binding mode that would be difficult to access using other macrocycles with a flexible backbone. Leveraging this unique recognition motif, resulting from the shape-persistency of these oligoamide macrocycles, in the template-directed synthesis of compact rotaxanes may open up new opportunities for the development of higher order interlocked molecules and artificial molecular machines.

Introduction

The development of mechanically interlocked molecules (MIMs) has invariably been accompanied by emerging recognition motifs and template-directed synthetic protocols.1 Among various factors, the combination of a macrocycle and a thread component with sufficient binding affinity constitutes a crucial determinant to create such motifs. In this regard, many elegant examples have been reported. These include the use of macrocyclic hosts such as crown ethers,7 tetracationic cyclophanes,1 calixarenes,4 cucurbiturils,5 cyclodextrins,6 and pillarenes7 to interact with the chosen thread components. However, uncovering simple and efficient recognition motifs still represents a grand challenge for the synthesis of MIMs, especially those involving non-metal coordination for extremely compact interlocked molecules, with high atom economy.

Thus far, most known examples of highly efficient, template-directed synthesis of [3]rotaxanes have relied on metal–ligand coordination.9 In all these cases, the ‘wheels’ employed are based on macrocyclic molecules that are nonplanar and fairly flexible conformationally. In recent years, shape-persistent macrocycles, with noncollapsible and geometrically well-defined skeletons, such as phenylacetylene macrocycles,9 have attracted considerable attention due to their intriguing functions, including supramolecular gelation,10 channelized transportation,11 organic...
catalysis,\textsuperscript{12} molecular recognition\textsuperscript{13} and multifunctional self-assembly.\textsuperscript{14} Many shape-persistent macrocycles have well-defined surface topography and nanosized cavities with reorganized binding sites, and thus may exhibit enhanced complexation.\textsuperscript{15}

Inspired by the pioneering work on synthesizing MIMs using flexible aromatic amide macrocycles\textsuperscript{16} or H-bonded oligoamide foldamers,\textsuperscript{17} we reasoned that a shape-persistent macrocycle, when serving as a wheel component for a MIM, may overcome the limit of known systems by engaging effectively in cooperative recognition interactions. However, introducing two-dimensional (2D) shape-persistent macrocycles into interlocked structures still presents a significant challenge. To the best of our knowledge, few examples of MIMs based on these macrocycles are known,\textsuperscript{18} especially [3]rotaxanes. In fact, there is only one report, based on a pentagonal cyanostar macrocycle, on forming [3]rotaxanes.\textsuperscript{19}

H-Bonded aromatic amide macrocycles\textsuperscript{20} are a large class of recently emerged 2D shape-persistent cyclic compounds featuring full amide linkages with rigid backbones enforced through intramolecular hydrogen bonds. Among them, cyclo[6]aramides, the smallest member of the aromatic oligoamide macrocycles\textsuperscript{21} with the amide carbonyl oxygens pointing inwards, are particularly noteworthy. They resemble a crown ether in their oxygen-rich cavities, but differ considerably in their conformational rigidity. In particular, the highly favorable intramolecular H-bond assisted macrocyclization strategy makes numerous geometrially well-defined macrocycles readily available, which facilitates their further applications. With a well-defined cavity of ca. 8 Å in diameter, large π-surfaces and precisely positioned binding sites, these macrocycles have displayed rich host–guest (H–G) chemistry. Our recent studies revealed that cyclo[6]aramides could bind organic cations and hydrogen bond donors.\textsuperscript{22} Especially notable is the complex consisting of such macrocycles and diquat, an isomer of paraquat, with 2 : 1 stoichiometry, in which a diquat molecule resides between two macrocyclic molecules rather than threads into their cavities due to electronic repulsion and steric hindrance.\textsuperscript{23} Paraquat is one of the most widely used guests for studying H–G interactions.\textsuperscript{24} Despite numerous reports on paraquat recognition, threading two rings on a single paraquat molecule with high binding affinity both in solution and in the solid state is still very difficult to achieve. The cryptand reported by Gibson and co-workers could form a 2 : 1 H–G complex only observed in the crystal structure and the highest binding constant for 1 : 1 stoichiometry achieved so far is 5.0 × 10\textsuperscript{6} M\textsuperscript{−1} in acetonitrile.\textsuperscript{25} A crown ether, bis-p-xylyl[26]crown-6, was used by Chiu and co-workers to form a [3]pseudorotaxane-like complex both in solution and in the solid state. However, only a very low binding constant (K\textsubscript{1} = 700 M\textsuperscript{−1} and K\textsubscript{2} = 60 M\textsuperscript{−1}) was observed in acetonitrile.\textsuperscript{26} So far no successful examples have been reported on creating a tight 2 : 1 binding motif for constructing constrained [3]rotaxanes based on a single paraquat unit. Herein we report that cyclo[6]aramides, with their persistent shape and geometrically well-defined electron-rich cavity (Scheme 1a), could act as powerful hosts for the tight binding of bipyridinium salts in a 2 : 1 binding mode both in the solid state and in solution with exceedingly high binding constants. More importantly, this unique recognition motif has led to the template-directed synthesis of compact [3]rotaxanes in excellent yields based on either a “click-capping” or “facile one-pot” approach (Scheme 1b).

**Results and discussion**

**Evidence for 2 : 1 cooperative host–guest complexation**

The first sign indicating the H–G interaction came from a color change upon adding paraquat G1 to a solution of cyclo[6]aramide 1 in acetonitrile. An abrupt change from clear to light yellow was observed, indicating that a charge transfer (CT) interaction happened between 1 and G1. A CT band in the UV-vis absorption spectrum confirms the formation of the H–G inclusion complex (Fig. S69†). Then, the formation of the H–G complex was further explored using \textsuperscript{1}H NMR spectroscopy (Fig. 1). When 1.0 equiv. of 1 was added to a 2.0 mM solution of G1, two sets of signals from the bipyridinium ion were observed, which evolved into one set of signals with 2 equiv. of 1, indicating the slow exchange of the complex on the NMR time scale. Commensurate with this change is the appreciable shift for the aromatic and amide protons of the host upon guest binding. The binding event was supported with 2D nuclear overhauser effect spectroscopy (NOESY), which revealed through-space NOEs between bipyridinium protons and the internal aromatic protons Ha and Hb (Fig. S64†, S65–S68†). Such through-space NOE contacts can transpire only if these two macrocycles are mutually parallel but orthogonal to G1. Two-dimensional diffusion ordered spectroscopic (2D-DOSY) analysis provided additional evidence for the formation of very stable complexes between 1 and G1–G4 (Fig. S65–S68†). For each complex, the protons of 1 and the guest have the same diffusion constant in solution. Furthermore, Job plots supplied information on the binding stoichiometry of 1 and G1–G4 in solution (Fig. S71–S78†). For example, in the case of the complex 1\textsubscript{2} ⊃ G1, the maximum change in absorbency was observed at 0.67, indicating a macrocycle-cation ratio of 2 : 1. Examining the 2 : 1 mixture of 1 and G1 using matrix-assisted laser ionization time of flight mass spectrometry (MALDI-TOF-MS) uncovered the peak with the highest absorbance was observed at m/z = 5005.219, corresponding to the complex [1\textsubscript{2} + G1 – PF\textsubscript{6}]\textsuperscript{−}. The above results, taken in concert, clearly demonstrate the formation of a 2 : 1 complex 1\textsubscript{2} ⊃ G1. The same 2 : 1 stoichiometry was observed with guests G2 and G4, but G3 only shows 1 : 1 stoichiometry with 1 (Fig. S103–S109†). Another interesting observation made on complex 1\textsubscript{2} ⊃ G1 is its reversible redox-responsiveness, which was realized by the addition and removal of Zn powder (Fig. S143†).

The binding constants K\textsubscript{1} and K\textsubscript{2} for the complexation of cyclo[6]aramide 1 with bipyridinium salts G1–G4 were obtained using UV-vis titration methods (Table 1). The UV-vis titration experiments revealed surprisingly high binding constants, K\textsubscript{1} = 3.49 × 10\textsuperscript{7} M\textsuperscript{−1} and K\textsubscript{2} = 1.09 × 10\textsuperscript{8} M\textsuperscript{−1}, for complex 1\textsubscript{2} ⊃ G1. The binding of G1 with the second macrocycle is accompanied by a slightly negative cooperative effect (4K\textsubscript{2}/K\textsubscript{1} < 1),\textsuperscript{27} which is probably caused by the remaining pyridinium unit that has become less electron deficient after the threading of the first macrocycle. To provide insight into the role of the second pyridinium moiety when the first site is occupied, G2, with one
positive charge, was examined for binding to 1. Results from UV-vis titration show that, as compared to those of 1agle G1, the binding constants for G2 and 1 are drastically reduced to 
\[ K_1 = \frac{3.62 \times 10^{4}}{C} \text{M}^{-1} \] and 
\[ K_2 = \frac{2.28 \times 10^{4}}{C} \text{M}^{-1} \], which indicates the pivotal role played by the second positive charge of G1 for the high stability of complex 1agle G1. In addition, the positive cooperativity (\(4K_2/K_1 = 2.51\)) observed for 1agle G2, along with the similar Ka values of complex 1agle G3 to that of K1 for 1agle G1, suggests that additional inter-macrocycle \(\pi-\pi\) stacking may also assist the binding of the second macrocycle in the formation of 1agle G2. It is worth noting that G4 gives the highest binding constants for binding 1 (\(K_1 = 1.68 \times 10^{7} \text{M}^{-1}\) and \(K_2 = 2.47 \times 10^{8} \text{M}^{-2}\)), which should greatly facilitate the synthesis of rotaxanes (Fig. S83–S95 and Table S1†). A similar trend in binding constants was observed in a competitive solvent (acetone-d6/DMSO-d6, 9/1, v/v) (Table 1, Fig. S94 and S95†). Additional infrared experiments on the complexes.

### Table 1  Binding constants of bipyridinium salts G1–G4 with cyclo[6]aramide 1a

| Complex | Solvent | Molar ratio | \( K_1 (\text{M}^{-1}) \) | \( K_2 (\text{M}^{-1}) \) | \( K_{12} (\text{M}^{-2}) \) | Cooperativity \( \alpha (4K_2/K_1) \n\)
<table>
<thead>
<tr>
<th></th>
<th></th>
<th></th>
<th></th>
<th></th>
<th></th>
<th></th>
</tr>
</thead>
<tbody>
<tr>
<td>1agle G1</td>
<td>Acetone</td>
<td>2 : 1</td>
<td>3.49 \times 10^{7}</td>
<td>1.09 \times 10^{6}</td>
<td>3.80 \times 10^{13}</td>
<td>0.12</td>
</tr>
<tr>
<td>1agle G1</td>
<td>Acetone/DMSO (9/1, v/v)</td>
<td>2 : 1</td>
<td>3.29 \times 10^{10}</td>
<td>3.72 \times 10^{6}</td>
<td>1.22 \times 10^{12}</td>
<td>0.45</td>
</tr>
<tr>
<td>1agle G2</td>
<td>Acetone</td>
<td>2 : 1</td>
<td>3.62 \times 10^{4}</td>
<td>2.28 \times 10^{4}</td>
<td>8.25 \times 10^{8}</td>
<td>2.51</td>
</tr>
<tr>
<td>1agle G3</td>
<td>Acetone</td>
<td>1 : 1</td>
<td>1.20 \times 10^{4}</td>
<td>——</td>
<td>——</td>
<td>——</td>
</tr>
<tr>
<td>1agle G4</td>
<td>Acetone</td>
<td>2 : 1</td>
<td>1.68 \times 10^{4}</td>
<td>2.47 \times 10^{7}</td>
<td>4.15 \times 10^{15}</td>
<td>0.58</td>
</tr>
<tr>
<td>1agle G4</td>
<td>Acetone/DMSO (9/1, v/v)</td>
<td>2 : 1</td>
<td>8.98 \times 10^{6}</td>
<td>2.14 \times 10^{14}</td>
<td>1.92 \times 10^{13}</td>
<td>0.95</td>
</tr>
<tr>
<td>5agle G4</td>
<td>Acetone-d6</td>
<td>2 : 1</td>
<td>[7.29 \times 10^{6}]d</td>
<td>[2.50 \times 10^{10}]d</td>
<td>[1.82 \times 10^{15}]d</td>
<td>0.13</td>
</tr>
</tbody>
</table>

a The binding constant values were obtained through UV-vis titration experiments using a nonlinear fit curve calculated using Matlab 801. For each titration, at least 30 data points were collected. Typically, wavelength was monitored around the charge transfer absorption maxima for the complex formed, estimated errors for \(K_1\) and \(K_2\) = ±50%. b The stoichiometry was determined using Job plots based on NMR experiments. c The cooperativity with the interaction parameter \(\alpha\) according to the equation \(\alpha = 4K_2/K_1\). d The binding constants were obtained using NMR titration experiments. For full details, see the ESL.
prepared from pyridinium salts and macrocycle 1 show that the C=O stretching frequency shift induced by complex formation is in the order of G4 > G1 > G2 > G3 (Fig. S110–S113 and Table S17), which agrees well with the difference in the binding affinities of these guests with macrocycle 1. In order to demonstrate the crucial role by shape persistency, cyclo[6]aramide 5, which is partially H-bonded and bears two rotatable amide groups, was synthesized. Since the depletion of partial intramolecular H-bonds results in free rotation of the two amide groups, the shape-persistency as observed in 1 should be attenuated to a significant extent. Indeed, the results from the binding experiments show that it binds G4 in a 2:1 binding mode (Fig. S81–S82†) with binding constants of \( K_1 = 7.29 \times 10^4 \text{ M}^{-1} \) and \( K_2 = 2.50 \times 10^3 \text{ M}^{-1} \) (Table 1 and Fig. S95 and S96†), which are four orders of magnitude lower than that of complex 12 \( \supseteq \) G4. This significantly reduced binding affinity as compared to 1 strongly suggests the importance of shape-persistency in the binding event. Despite the great advances made in the past decades, few recognition modules have shown such unusually strong binding in organic media.

**X-ray crystal structure of [3]pseudorotaxane 32 \( \supseteq \) G1**

The slow evaporation of a solution containing 3 and G1 in a mixed solvent of acetone/CHCl3/methanol (10/1/0.3) afforded red crystals in 7 weeks. The analysis of the resulting solid state structure using single-crystal X-ray diffraction (XRD) experiments reveals that G1 is inserted into two macrocyclic molecules. In this complex, each of the pyridinium units is orientated orthogonally with respect to the two macrocycles and engages in multiple C–H···O H-bonding (Fig. 2a) and N···O ion–dipole interactions that are reinforced by face to face π–π stacking interactions (3.71 Å) (Fig. 2b). These hydrogen bonds all have very short donor–acceptor distances ranging from 2.30 to 2.74 Å due to the rigid macrocyclic backbone constrained by intramolecular three-centre H-bonds. The length of the N···O ion–dipole interactions varying from 3.32 to 4.73 Å is indicative of strong Coulomb interaction between the amide oxygen atoms and pyridinium cations. Therefore, the strong binding affinity between cyclo[6]aramide 1 and the bipyridinium guests is attributed to the result of the cooperative interplay of multipoint non-covalent forces. The geometrically well-defined and tightly packed solid-state structure of 32 \( \supseteq \) G1 is also observed which is stabilized by multiple van der Waals forces between the side chains of macrocycle 3 and G1 (Fig. 2c).

“Click-capping” approach for the synthesis of rotaxanes

Click reactions, with their high efficiency and simplicity, are widely used for constructing MIMs. By using the observed tight binding, we anticipated that compact rotaxanes [3]CR-C16 (n = 16, 12 and 6) might be attained by tethering Stopper-N4 to the [3]pseudorotaxane formed from cyclo[6]aramides 1–3 and guest G4 via a “click-capping” approach. Indeed, the templation reaction proceeded very well, leading to the specific formation of [3]rotaxanes (Table 2, entries 1–3). In the case of macrocycle 1, which is decorated peripherally with sterically crowded alkyl groups, the reaction offers [3]rotaxane [3]CR-C16 as the sole product in a yield of 86%. Using macrocycle 2, which tends to severely aggregate in solution due to the presence of linear alkyl chains,28 the yield of [3]rotaxane [3]CR-C16 was increased up to 91%. To the best of our knowledge, this is one of the rare examples of the synthesis of [3]rotaxanes in excellent yields reported hitherto using a non-metal coordination strategy.29 Due to its low solubility, macrocycle 3 forms [3]CR-C6 in a yield of 64%. Varying the solvent polarity and the ratios of reactants resulted in the controlled generation of [3]rotaxanes as the favourable product (Table 2, entries 4 and 5). To probe which binding mode (1 : 1 or 2 : 1) is more favourable for forming rotaxanes, a mixture of 1 equiv. of 1 and 1 equiv. of guest G4 was stirred in the presence of a catalyst. We speculated that the formation of [2]rotaxane should have dominated under this condition; however, besides [2]rotaxane in 34% yield, the yield of [3]rotaxane [3]CR-C16 still reached 36%, indicating that the
recognition event favours 2 : 1 stoichiometry, rather than 1 : 1, during the binding process. The highly selective formation of [3]rotaxanes with cyclo[6]aramides originates from the combination of exceedingly high binding affinities and the compact self-assembly mode. These features are rarely observed in other similar H–G systems. In addition, face to face π–π stacking interactions between the two near-planar macrocycles as observed in the crystal structure of 3_2 with G1 could provide an additional driving force that contributes to the high efficiency. It is worth noting that the preparation of [3]rotaxanes with shape-persistent macrocycles as “wheels” in such a high yield (>90%) under non-metal coordination conditions is preceded by few examples and remains a formidable task.16a,30

“Facile one-pot” approach for the synthesis of rotaxanes

The “facile one-pot” reaction, characterized by simply mixing and heating, and the absence of metallo-catalysts, is another useful approach in the synthesis of rotaxanes with high yields.11 The synthesis of [3]rotaxanes [3]R-C_n (n = 16 and 12) or [2]rotaxane [2]R-C_6 was based on this method achieved through mixing 2.0 equiv. of cyclo[6]aramide, 2.5 equiv. of Stopper-Br and 1.0 equiv. of 4,4′-bipyridine in CHCl_3/CH_3CN (1/1, v/v) (Table 3, entries 1–3). Highly efficient template-directed synthesis with macrocycle 1 or 2 was achieved with an excellent yield of 85%, which is rare in the known synthesis of [3]rotaxanes.12 Particularly noticeable was the formation of a [2]rotaxane as the only product when macrocycle 3, which bears short side chains, was used. This specificity in [2]rotaxane formation is rationalized according to the sparse dissolution of the macrocycle in the solution. Compound 3 alone was insoluble in CHCl_3/CH_3CN (1 : 1, v/v). However, gradual dissolution was observed to occur as the reaction progressed. This suggests a scanty concentration of macrocycle 3 in the reaction system as compared to that of macrocycle 1 or 2. Since the efficiency of forming [3]rotaxanes depends predominantly on the effective molar ratio (2 : 1) of the macrocycle and cationic guest in solution, the limited concentration of macrocyclic molecules with respect to that of the cationic axle tends to facilitate a binding process that favours a 1 : 1 binding mode, thereby leading to the specific formation of a [2]rotaxane.

<table>
<thead>
<tr>
<th>Entry</th>
<th>Macrocycle</th>
<th>n</th>
<th>Yield (%)</th>
</tr>
</thead>
<tbody>
<tr>
<td>1a</td>
<td>1</td>
<td>16</td>
<td>86</td>
</tr>
<tr>
<td>2a</td>
<td>2</td>
<td>12</td>
<td>91</td>
</tr>
<tr>
<td>3a</td>
<td>3</td>
<td>6</td>
<td>64</td>
</tr>
<tr>
<td>4a,b</td>
<td>1</td>
<td>16</td>
<td>36</td>
</tr>
<tr>
<td>5a</td>
<td>1</td>
<td>16</td>
<td>60</td>
</tr>
</tbody>
</table>

* Solvent is acetonitrile/[CHCl_3]/PF_6, and the macrocycle prior to the reaction did not lead to any difference in the rotaxane formation as compared to the common synthetic procedure in the Experimental section. Only 1.0 equiv. of macrocycle 1 was used. ** Solvent is acetonitrile/[CHCl_3]/PF_6, 40 °C, 48 h. n is the number of side chain carbons of the macrocycle. isolated yield after chromatography. The yield of [2][3]rotaxane was calculated based on the macrocycle. * Isolated yield after chromatography. The yield of [2][3]rotaxane was calculated based on the macrocycle. n.d. = not detected. For full details, see the ESI.

**Table 3** “Facile one-pot” approach for the synthesis of rotaxanes [3]R-C_n

<table>
<thead>
<tr>
<th>Entry</th>
<th>Macrocycle</th>
<th>n</th>
<th>Yield (%)</th>
</tr>
</thead>
<tbody>
<tr>
<td>1a</td>
<td>1</td>
<td>16</td>
<td>85</td>
</tr>
<tr>
<td>2a</td>
<td>2</td>
<td>12</td>
<td>85</td>
</tr>
<tr>
<td>3a</td>
<td>3</td>
<td>6</td>
<td>71</td>
</tr>
</tbody>
</table>

* Solvent is acetonitrile/[CHCl_3]/PF_6(1/1, v/v), 40 °C, 24 h. ** n is the number of side chain carbons of the macrocycle. † Isolated yield after chromatography. The yield of [2][3]rotaxane was calculated based on the macrocycle. n.d. = not detected. For full details, see the ESI.
The mass spectra of [3]RC16, [3]RC12 and [2]RC6 at m/z = 5236.883, m/z = 4563.225 and m/z = 2058.365, pointing to the formation of a [3]rotaxane or [2]rotaxane (Fig. 4d-f). In all cases, all the peaks were isotopically resolved and fully agreed with their calculated theoretical distributions. It should be noted that all the rotaxanes are in red due to charge-transfer interactions between the macrocycles and axles, as indicated by the CT band observed in the UV-vis absorption spectra (Fig. S142†).

X-ray crystal structure of [3]CR-C6

Single crystals of [3]rotaxane [3]CR-C6 were obtained by slowly diffusing methanol into an acetone solution containing [3]CR-C6 in about 8 weeks. The X-ray structure of [3]CR-C6 clearly shows that the thread Axle-1 penetrates through the cavities of the two neighboring macrocyclic molecules in a zigzag conformation (Fig. 5a and b). The mechanically interlocked structure is stabilized by twenty C–H⋯O H-bonds and twelve N+/O ion dipole interactions (Tables S4 and S5†). There are also a number of very weak face-to-face π⋯π stacking interactions (4.658 Å) between the macrocyclic molecules, resulting in a well-ordered and tightly packed solid-state structure (Fig. 5c). The numerous noncovalent forces revealed from the X-ray structure work cooperatively, leading to the surprisingly high stability of the [3]rotaxanes. In fact, when a mixture of Axle-1 and macrocycle 1 was heated under reflux for 3 hours in acetone-d6/DMSO-d6 (9:1, v/v), the 1H NMR spectra did not show any signs of the presence of rotaxanes, indicating the fact that threading does not occur (Fig. S117†), and thus it can be inferred that [3]rotaxane [3]CR-C6 is unlikely to experience a dethreading process under the conditions specified. The high stability is also demonstrated by the observation that complex [2]IC6 (or [3]IC4) and free macrocycle 1 were clearly seen for each of them on a TLC plate. Upon the addition of 10.0 equiv. of diethylamine (DEA) to [3]R-C16 or [2]R-C6 in acetone, only [2]R-C6 caused a color change from light yellow to blue, indicating the insensitivity of [3]R-C16 to redox responsiveness. Interestingly, trifluoroacetic acid (TFA) can reverse the redox process of [2]R-C6 (Fig. S144†).

Computational simulation of [3]R-C1

Since the growth of single crystals of the [3]rotaxane [3]R-Cn (n = 14, 12 and 6) synthesized according to the “facile one-pot” approach has proved to be extremely challenging, we resorted to molecular mechanics simulations to gain a better understanding of the noncovalent bonding interactions that direct rotaxane formation and stability. Further structural insights on [3]rotaxane [3]R-C1 were obtained through computational simulations based on the DFT method. Our computational study indicated that [3]rotaxane [3]R-C1 was built by assembling...
the two macrocyclic molecules of 4 in a near-planar conformation and the central motif Axle-1 in an interlocked orthogonal binding arrangement, which is in good agreement with the structure obtained from the single crystal of [3]CR-C6. Furthermore, multiple C–H⋯O H-bonds and N⋯O ion–dipole interactions are also observed in the modelling structure which direct the [3]rotaxane formation (Fig. S153†). It is worth noting that there are no face to face π–π stacking interactions formed between the two macrocycles, which is the same as the observation from the crystal structure of [3]CR-C6.

Conclusions

In conclusion, we have demonstrated the unusually high binding affinity ($K_a$ from $\sim 10^{13}$ M$^{-2}$ to $\sim 10^{15}$ M$^{-2}$ in acetone) of a novel threaded recognition motif comprising cyclo[6]aramides and bipyridinium salts in 2 : 1 (H : G) stoichiometry. The crystal structure of the [3]pseudorotaxane shows clear evidence of the high binding affinity, which is attributed as the result of the cooperative interplay of multipoint C–H⋯O H-bonds, N⋯O ion–dipole interactions and π–π stacking interactions between the two neighboring macrocycles. Furthermore, the highly efficient synthesis of compact [3]rotaxanes achieved using either a “facile one-pot” or “click-capping” approach presents a rare example of constructing MIMs based on 2D shape-persistent macrocycles. The high efficiency of the formation of these rotaxanes highlights the unique advantage of macrocyclic shape-persistence, which results in enhanced multipoint recognition for the highly efficient synthesis of compact mechanically interlocked molecules. The concept of utilizing macrocyclic shape-persistence for boosting multipoint binding affinities for the template-directed synthesis of rotaxanes might be useful for the design of novel MIMs and the development of artificial molecular machines.

Experimental section

General experimental procedure for the “click-capping” approach for the synthesis of [3]CR-C$_n$ ($n = 16, 12$ and $6$)

A mixture of cyclo[6]aramide (2.0 equiv.), guest G4 (1.0 equiv.) and Cu(CH$_3$CN)$_4$PF$_6$ (0.3 equiv.) was stirred in dry acetone at room temperature for 20 minutes under N$_2$. Then a solution of Stopper-N$_3$ (2.5 equiv.) and N,N-diisopropylethylamine (DIPEA) (1.2 equiv.) was injected. The mixture was further stirred for 40°C for 24 h. The resulting solution was washed with 16% aqueous EDTA tetra-sodium saturated ammonia solution (2 x 50 mL). The organic layer was retained and the aqueous layer extracted twice with CH$_2$Cl$_2$ (2 x 50 mL). The organic extracts were combined and washed with water, dried over Na$_2$SO$_4$ and dried in vacuo. Removal of the solvent afforded a red solid and the crude material was purified using flash column chromatography with silica gel (CHCl$_3$/CH$_3$OH, 20 : 1) to yield the corresponding red solid. The detailed synthetic procedures and full characterization of new compounds are provided in the ESI†.


A mixture of cyclo[6]aramide (2.0 equiv.), 3,5-di-tert-butylbenzyl bromide Stopper-Br (2.5 equiv.) and 4,4′-bipyridine (1 equiv.) was stirred in 6 mL of CHCl$_3$/CH$_3$CN (1/1, v/v) under N$_2$ at 40°C for 48 h. Removal of solvents afforded a pale red solid, which was dissolved in acetone/H$_2$O, followed by the addition of saturated aqueous NaPF$_6$. After stirring for 30 min, the organic solvent was then evaporated under reduced pressure. The precipitate was collected and washed with H$_2$O. Then the crude material was purified using flash column chromatography with silica gel (CHCl$_3$/CH$_3$OH, 30 : 1, then CHCl$_3$/CH$_3$OH, 10 : 1) to yield the corresponding red solid. The detailed synthetic procedures and full characterization of new compounds are provided in the ESI†.

Single X-ray crystal data

Crystallographic data (excluding structure factors) for all the structures reported in this article have been deposited with the Cambridge Crystallographic Data Centre in CIF (see the ESI†). Crystal data for 3$_2 \supset$ G1 (CCDC-1475246): C$_90$H$_{121}$N$_7$O$_{18}$PF$_6$, M
0.107 mm

1475247): C216H292N20O36P2F12,

Acknowledgements

This work is supported by the National Natural Science Foundation of China (21572143), Doctoral Program of the Ministry of Education of China (2013011110023), Open Project of State Key Laboratory of Supramolecular Structure and Materials (SKLSSM201629), and US National Science Foundation (CHE-1306326). The Comprehensive Training Platform of the Specialized Laboratory, College of Chemistry, Sichuan University, is acknowledged for NMR analysis.

Notes and references


