To catalyze or not to catalyze: elucidation of the subtle differences between the hexameric capsules of pyrogallolarene and resorcinarene†

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The functional mimicking of natural enzymes has been a very fascinating but also challenging research topic for decades. Several supramolecular structures have been identified that are able to catalyze reactions inside their enzyme-like pockets.† Nevertheless, neither the catalytic efficiency nor the selectivity of such systems, can usually compete with natural enzymes. Therefore, the refinement of known structures, as well as the development of new systems, is necessary. For the design of new artificial enzyme-like catalysts, it is of fundamental importance to understand the prerequisites for catalytic activity. We† and others† have shown that the hexameric resorcinarene capsule I, which self-assembles from six units of resorcinarene and eight water molecules (Fig. 1), is an efficient catalyst for a variety of cationic reactions. Nevertheless, the reasons for the high catalytic efficiency remain unclear. To learn more about the pivotal requirements for the catalytic activity of hexamer I, we became interested in the structurally closely related pyrogallolarene hexamer II.† It self-assemblys from six units of pyrogallolarene. Surprisingly, hexamer II displays a different encapsulation behavior compared to I in chloroform solution. It was reported that hexamer I encapsulates both tertiary amines as well as alkylammonium species,† while hexamer II only binds tertiary amines.† Very recently, the Cohen group disclosed that binding of ammonium salts in II can be observed to some extent in benzene solution. The high affinity of ammonium salts for II can be explained by strong cation-π interactions. However, the surprising exclusion of alkylammonium species from II in chloroform solution remained puzzling, especially since encapsulated tertiary amines were completely expelled after protonation by the addition of acid.† This seemingly contradictory encapsulation behavior of II has remained a mystery for the last decade. We herein elucidate the reasons for these differences in chloroform solution. Importantly, we report that capsule II is catalytically completely incompetent in cationic reactions which are efficiently accelerated inside I and probe the molecular mechanisms responsible for these surprising observations.

Our interest in capsule II started with the observation that, while I is an efficient catalyst for cationic reactions, hexamer II is catalytically incompetent in such reactions. The tail-to-head

Fig. 1 Structures of hexameric resorcinarene I and pyrogallolarene II capsules, optimized at the density functional theory (DFT: PBE-D3/def2-SVP) level and their respective building blocks 1 and 2.

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The closely related, self-assembled resorcinarene and pyrogallolarene capsules display contrasting and puzzling encapsulation behaviors. Herein, we elucidate the reasons for these differences by combining experimental studies and DFT calculations. Furthermore, we report that, in contrast to the resorcinarene capsule, the pyrogallolarene derivative is not capable of catalyzing reactions with cationic transition states. The molecular mechanisms responsible for these observations are probed in detail.
terpene cyclization of nerol (3) (Fig. 2) was investigated more closely with II. 1H-NMR analysis confirmed the formation of the host–guest complex 3@II (ESI-Fig. 14†). Therefore, the guest uptake cannot explain its catalytic inertness.

We first probed whether the absence of catalytic activity of II could arise from the reported resistance to bind alkyl ammonium salts, which may indicate a lack of cation–π stabilization inside the cavity. A cation–π stabilization of cationic transition states has been proposed as a possible reason for rate acceleration inside I, and may thus be a potential explanation for the observed catalytic differences. Interestingly, however, it was reported that tertiary amines bind well inside II. Recently, we demonstrated that a proton transfer from capsule I to the tertiary amines is responsible for their high affinity for I. This raised the question of whether the same is true inside II. Therefore, we investigated the encapsulation of differently sized tertiary amines inside II. The encapsulation of triethylamine (5a) in II is clearly evident in the 1H-NMR spectrum (ESI-Fig. 1†).

The integral of the phenolic groups of II diminishes upon treatment with 5a, while a new broad peak emerges between 3 and 7 ppm. Careful integration (ESI-Schemes 1 and 2, Table 2†) reveals that it accounts for the diminished phenolic protons, as well as for the water signal, which is no longer visible as a separate peak. These observations are consistent with our previous findings for capsule I, and thus indicate the concomitant protonation of tertiary amines upon encapsulation in II. The integrity of the hexameric encapsulation complex was confirmed by DOSY spectroscopy (ESI-Fig. 2†). Therefore, the formation of smaller (dimeric, or monomeric) pyrogallolarene–cation complexes observed in methanol solution and in the solid state can be excluded.

The size of the tertiary amine has a pronounced effect on protonation and encapsulation, as shown in Table 1a. The degree of protonation and encapsulation drops with the increasing size of the alkyl groups, a behavior that is consistent with a decreased cation–π interaction due to steric shielding. In the case of triethylamine (5a), the deprotonation degree of II is considerably lower than the encapsulation degree. Therefore, a [5H5]@II-cation complex, also observed in capsule I, is likely formed.

These results indicate that capsule II encapsulates amines in their protonated form and therefore is able to stabilize cations inside its cavity. Indeed, our quantum chemical density functional theory (DFT) calculations also indicate that capsule II stabilizes cations as well as or even more strongly than capsule I (within ca. 4–12 kcal mol⁻¹ for NEt₃⁺; see ESI-Table 4†). This seems to contradict the previous observation that the ammonium salt Hex₄NBr (6dBr⁻) was rejected by capsule II. We therefore repeated the encapsulation studies not only with 6dBr⁻ but with differently sized ammonium salts (Table 1b). The small ammonium salt Et₄NBr (6aBr⁻) is indeed encapsulated to a considerable extent (see ESI-Fig. 5†). However, quantification was hampered by precipitation of a dimeric complex (see ESI-Table 3†). With the increasing size of the alkyl residues, encapsulation drops dramatically: only 4% encapsulation is observed in the case of Pr₄NBr (6bBr⁻), while the longer ammonium salts Bu₄NBr (6cBr⁻) and Hex₄NBr (6dBr⁻) do not show any degree of uptake inside II. This confirms the earlier observation that larger ammonium salts are rejected by II, but also demonstrated that the uptake of the smaller salts is possible. Nevertheless, there is a striking difference between the capsules, as I encapsulates salts (6a–6cBr⁻) quantitatively and 6dBr⁻ to a large extent (see, ESI-Fig. 4†). This discrepancy is even more surprising considering that our DFT calculations suggest an even stronger stabilization of cations inside II relative to I. To elucidate the different behavior of I and II, we calculated the electrostatic potential (ESP) map of the capsules’

### Table 1: Encapsulation studies with 1 eq. of (a) amine 5 with capsule II and (b) ammonium compounds 6 with capsules I and II. The encapsulation of the ammonium guest by capsule II was increased by the addition of the large amine base 7

<table>
<thead>
<tr>
<th>Amino Acid</th>
<th>Capsule</th>
<th>Deprotonation degree of II</th>
<th>Encapsulation degree of II after addition of 7</th>
</tr>
</thead>
<tbody>
<tr>
<td>5a</td>
<td>C₆H₁₃</td>
<td>21%</td>
<td>7%</td>
</tr>
<tr>
<td>5b</td>
<td>C₇H₁₅</td>
<td>19%</td>
<td>12%</td>
</tr>
<tr>
<td>5c</td>
<td>C₆H₁₃</td>
<td>12%</td>
<td>6%</td>
</tr>
<tr>
<td>5d</td>
<td>C₆H₁₃</td>
<td>10%</td>
<td>7%</td>
</tr>
</tbody>
</table>

[a] Precipitation occurred (2 : 1 complex, ESI-Table 3).
inner surfaces. As shown in Fig. 3, the main difference between the two systems are high potential areas on the inner surface of I, which represent positively charged hydrogen atoms of the bound water molecules. These could potentially stabilize the anions of encapsulated ammonium salts via hydrogen bonds. Such a stabilization of anions inside II is lacking, and could therefore explain the dramatically weaker binding of ammonium salts. Our DFT optimizations of the capsules in presence of the ammonium salt 6a·Br− confirm this hypothesis: although capsule II stabilizes cations better than capsule I, capsule II binds the ammonium salt weaker (ca. 3–9 kcal mol−1, see ESI-Table 4†) than I.

The energetically unfavorable encapsulation of anions inside II raises the question of whether the anion is encapsulated at all. Its presence would weaken the cation–π interactions, as has been shown in cyclophane hosts in organic solvents of low dielectric constants. Therefore, the ammonium salt EtPr3N+MeSO3− (6a·MeSO3−) containing an organic anion, detectable by 1H and 13C-NMR spectroscopy, was investigated. 1H- and 13C-NMR investigations (ESI-Fig. 8–10†) confirm that the counterion is located outside of II (Fig. 4a). In capsule I, however, the ion pair is encapsulated (ESI-Fig. 11–13†). The minute uptake of 6+ by neutral II is likely achieved by an energetically unfavorable charge separation in CDCl3. This could explain the low encapsulation ratio of 6+ by II. If so, the addition of the large base 7 (Table 1), which cannot be encapsulated due to its size, should increase the uptake of ammonium species: it would deprotonate the capsule and form an ion pair with the bromide outside the capsule, as depicted in Fig. 4b. Indeed, upon the addition of 7, the encapsulation of ammonium species increases as anticipated (see also Table 1b). These findings solve the puzzling encapsulation behavior of capsule II and also explain the expulsion of encapsulated trihexylamine (5d) after HCl-addition28 (Fig. 4c). After the addition of HCl, the ion pair 5dH+Cl− is formed, which is rejected by capsule II.

The evidence presented clearly indicates that capsule II is able to stabilize cations inside its cavity due to cation–π interactions. Therefore, its catalytic incompetence originates from a different source. After successful substrate uptake, protonation is required for substrate activation (Fig. 2). Although II is able to protonate amines (Table 1a), the acidity of the system may be too low for activation of the alcohol substrate. Therefore, the acidity of hexamer II was determined analogously to I29 by a series of protonation experiments with amines of varying basicity. The pKₐ value of capsule II is approx. 9.5–10 (see ESI-Fig. 15 and 16;† ca. four pKₐ units higher than resorcinarene capsule I). These results are in excellent agreement with our DFT calculations at the PBE-D3/def2-SVP/elib = 4.81 level of theory, suggesting that the proton affinity of I is ca. 5 kcal mol−1 lower than that of II (ESI-Fig. 18†). The surprisingly low acidity of II may be a result of mesomeric destabilization of the phenolate (ESI-Fig. 17†). Our DFT calculations further show that the anionic defect can delocalize across several hydrogen bonds in I, while we observe a more localized defect in II, leading to a lower relative pKₐ in capsule I (ESI-Fig. 19†). Therefore, the low acidity of II is the likely cause of its catalytic incompetence which prevents activation of the substrate by proton transfer (see also Fig. 2). We tried to overcome this limitation by the addition of stronger external acids (see ESI 4.3†), but could not observe a difference to the background reaction, caused by the acid added. This result, however, is not too surprising, since an external acid forms an ion pair with the substrate, which will resist encapsulation (see also Fig. 4c).
The elucidation of the differences in these two systems allowed us to learn important lessons concerning catalytic activity in hydrogen-bond based molecular capsules. The identification of acidity as the crucial element of the catalytically active derivative I is essential for the design and construction of novel hydrogen bond-based supramolecular catalysts. These studies provided us with a first estimate on the required acidity for catalytic activity in such systems. Additionally, we demonstrated that externally added acid cannot function as a co-catalyst for capsule II since the substrate acid ion pairs formed are not encapsulated.

Conclusions

In conclusion, the differences between the closely related hexameric capsules I and II were elucidated for the first time. To the best of our knowledge this is the first study of two very closely related supramolecular host systems which completely differ in their catalytic activity. Evidence was presented showing that capsule II does not stabilize anions inside its cavity, as opposed to I. Therefore, alkyl ammonium salts are encapsulated only to a small extent with concomitant charge separation. Cations, nevertheless, are stabilized inside II even more strongly than inside I via cation–π interactions. The much lower acidity of capsule II was determined to be the cause of its catalytic incompetence. These findings are of great significance for future developments in the field of enzyme-like catalysis and have a profound impact on the design of new hydrogen bond-based catalytically active host systems.

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Notes and references


