

Cite this: *RSC Adv.*, 2017, 7, 55450Received 30th September 2017
Accepted 28th November 2017

DOI: 10.1039/c7ra10808g

rsc.li/rsc-advances

Process simulation of coal-direct chemical looping gasification for syngas production

Lang Liu,^a Yan Cao,^{*b} Dongran Ma,^c Qingcai Liu^c and Jian Yang^c

The coal gasification process is used in the commercial production of syngas as a means toward the clean use of coal. In this paper, the process of syngas production *via* the coal-direct chemical looping gasification process (CDCLG) is modeled under thermochemical equilibrium with the Gibbs free energy approach, and the model is developed using a complete and comprehensive simulation model. Due to the kinetic limits of char gasification, the carbon conversion efficiency is not suggested to be larger than 0.9 in the fuel reactor during the CDCLG process. Besides, the simulation results suggest the residual carbon should be burnt in the air reactor. Moreover, the effects of various conditions, which include carbon conversion, temperature, pressure and the steam feed ratio, on the input and energy output parameters, including the air feed ratio, efficiency of electricity and syngas quality, are studied and revealed in this paper.

1. Introduction

Syngas, one of the most important energy carriers in sustainable energy systems, can be produced from carbonaceous materials, such as natural gas, oil, coal and biomass.^{1–3} From these mentioned energy sources, coal has been regarded as the most attractive option, due to its relative rich abundance and low cost. Hydrogen has already been generated from coal *via* centralizing gasification processes, in which the coal is first partially oxidized to syngas.^{4,5} Fuel gas with a higher H₂ concentration can be produced *via* steam-char gasification, which attracts much research interest. Additionally, more highlights can be found in the steam gasification process, such as maximizing gas productivity with a high heating rate, advantageous residence time and high char reduction efficient. However, steam gasification reactions are highly endothermic. To provide the necessary heat input, air is introduced to burn parts of the available fuels (char and gas). This favors the dilution with nitrogen of the product gas and reduces its caloric value. A number of researchers have studied the use of metal oxides for providing the heat, among which the exothermic reaction between CuO and carbon results in copper oxide being one of the most popular replacements. Also, CuO has been proven to be an oxygen releasing material,^{6,7} which can

contribute to kinetic improvement of the reaction between CuO and carbon.⁸

Chemical looping combustion and gasification (CLC&G⁹) has been suggested to be the most promising technology, which combines the inherent separation of CO₂ with a reduced energy penalty.^{10–12} Oxygen carriers (OC) and replacing air are used as the oxygen source to prevent mixing nitrogen with the CO₂ stream.

Different chemical looping configurations have been studied by Andersen *et al.*¹³ and Ertesvåg *et al.*¹⁴ Consonni and Viganò¹⁵ analyzed a pre-combustion setup to co-generate power and hydrogen. Besides, simulation software has been widely employed to investigate the utilization of chemical looping strategies in coal gasification processes, such as optimizing syngas productivity, energy balances, the circulation rates of oxygen carriers, the oxygen carrier mass loadings, the carbon burnout, the oxygen partial pressure, *etc.*^{16–18} The coal-direct chemical looping gasification (CDCLG) system is a promising technology for capturing post-combustion CO₂ from power plants, in which the fuel reactor is regarded as the gasifier, producing syngas, and the air reactor is the combustor in the gasification process. In this study, the kinetics of char-CO₂ gasification are studied in a pressurized thermo-gravimetric analyzer (TGA-HP150S) before simulation. Then the integration of a chemical looping concept and partial-gasification-combustion is evaluated using simulation software. Notably, we focused on a new configuration of chemical looping gasification, which featured high but not full carbon conversion in the fuel reactor, with residual carbon then being combusted in the air reactor. We proposed that the char residue should be removed from the gasification stage and considered in the combustion stage for the continuation of carbon conversion.

^aGuizhou Institute of Technology, Chemical Engineering Institute, Guiyang, Guizhou, 550003, China. E-mail: l.liu.git@qq.com

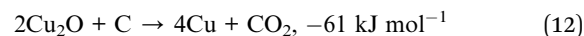
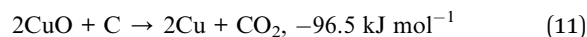
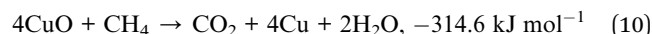
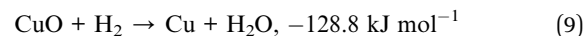
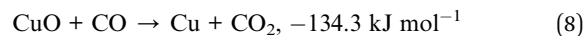
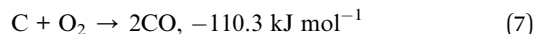
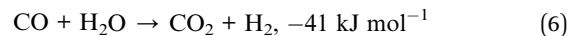
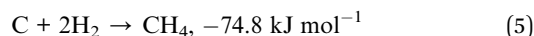
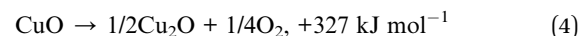
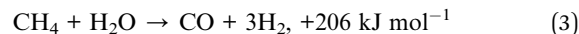
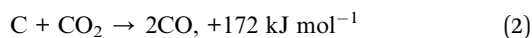
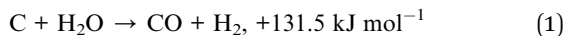
^bInstitute for Combustion Science and Environmental Technology, Chemistry Department, Western Kentucky University, Bowling Green, KY 42101, USA. E-mail: yan.cao@wku.edu

^cCollege of Material Science & Engineering, Chongqing University, Shapingba, Chongqing 400044, China



2. Process descriptions

A schematic diagram of the chemical looping system is shown in Fig. 1. There are two main reactors in this system, *i.e.*, a fuel reactor and an air reactor. The major reactions (with CuO as the oxygen carrier) in the fuel reactor are listed as follows:



Aiming to simplify the explanation, just the parts of the reactions involved into the thermodynamic calculations are listed here. Furthermore, coal is considered as pure carbon in these reactions. Meanwhile, hydrogen and other combustible or gaseous constituents have also been considered in the thermodynamic calculations.¹⁹

Heat balance requires the highly endothermic reactions (1)–(4) to occur in the fuel reactor; the air reactor has a higher temperature than that of the fuel reactor. Therefore, the oxygen carrier solids bring extra heat to maintain the necessary heat for the gasification stage. Moreover, reactions (8)–(12) are applied to supply the reaction heat; these reactions have exothermic properties and occur when feeding oxygen carriers into the fuel reactor.

In the air reactor, the reduced oxygen carrier is fully reoxidized to CuO according reactions (13) and (14). These oxidation reactions generate a large amount of heat; as a result, the oxygen carrier in the air reactor has a higher temperature

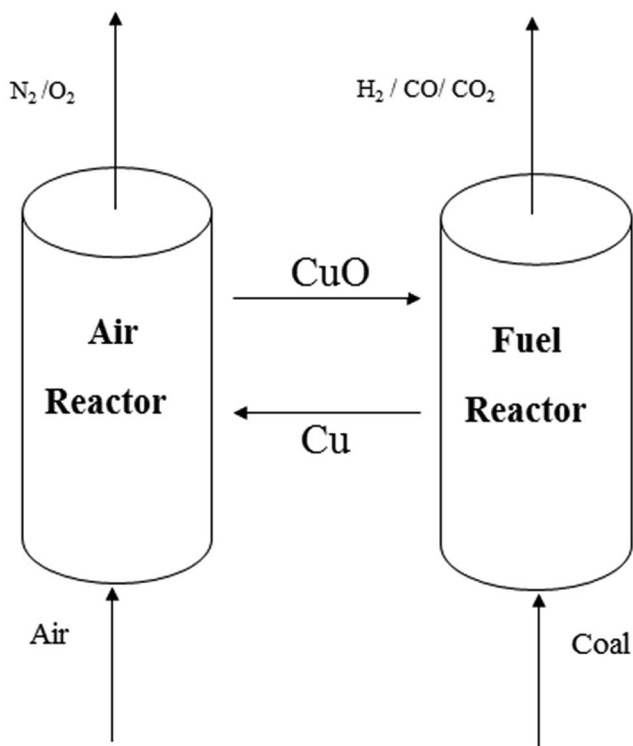


Fig. 1 A simplified flow diagram for the coal-direct chemical looping process.

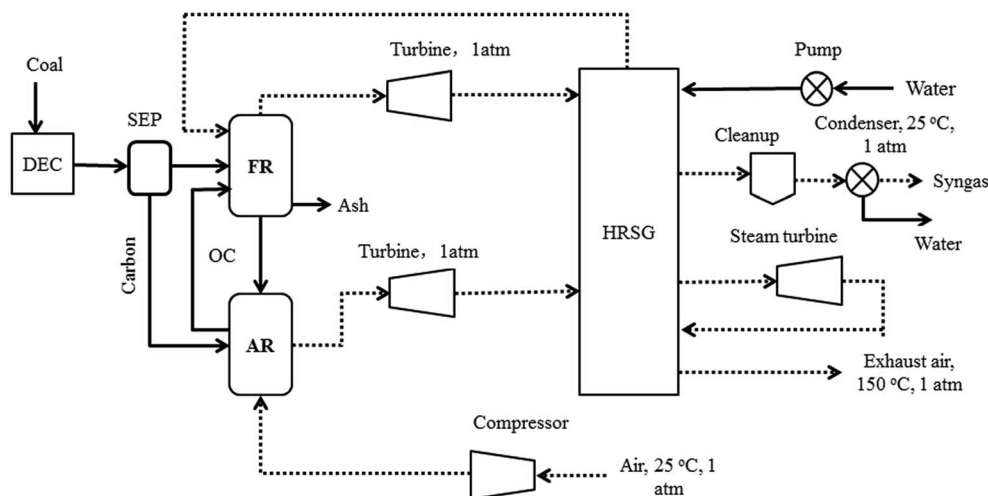


Fig. 2 A process flow diagram of the CDCLG process for syngas production.



Table 1 Simulation models for the key units

Unit operation	Model
Pyrolyzer	RYield
Gasifier	RGibbs
Air reactor	RGibbs
Pressure changers	Pump, compr, Mcompr
Heat exchangers	Heater, MheatX
Mixers/splitters/separators	Mixer/cyclone/sep

$$y = \frac{\eta}{\text{CO}_2 \text{ exhaust from the air reactor (mol kg}^{-1}\text{)}} \quad (18)$$

3. Operating conditions

Proximate and ultimate analyses of the coal used in this simulation are summarized in Table 2.

Char gasification has been reported to be a rate-determining process. Due to its kinetic limit, it is difficult to keep the char gasification process going in the final stages. The kinetics of char-CO₂ gasification are analyzed consecutively using a pressurized thermo-gravimetric analyzer (TGA-HP150S) at 0.1–2.0 MPa, 900 °C, and under 100% (mole ratio) CO₂. The TGA-HP150S model is specially configured for gasification and related studies.^{23,24} It includes an integrated steam generator and a double-walled reaction chamber that allows for higher temperatures. The maximum weight capacity is 25 g. The weighing accuracy is ±0.0002%. The weighing precision is ±30 µg. The sensitivity is 10 µg. The upper temperature and pressure limit are 1200 °C and 5.0 MPa respectively. The relationship between the gasification rate and carbon conversion efficiency is shown in Fig. 4.

It is observed that the gasification rate experienced an initially slow increase, then a rapid increase and finally a decrease, corresponding to the carbon conversion efficiency. The kinetics of char gasification finally dominated the determination of the threshold values for carbon conversion efficiencies between the fuel reactor and air reactor. Here, the optimal threshold value of the carbon conversion efficiency for the selected coal in this study is around 0.9, where further gasification of char residue is not suggested. In the following study on the effects of temperature, pressure and steam ratios on the process performance, the carbon conversion in the fuel reactor is set at 0.9, and the unconverted carbon is burned in the air reactor.

The coal processing capacity of the plant is specified to be 13 267 kg h⁻¹ (100 MW, HHV (Higher Heat Value)). In the simulation process, the oxygen carrier is composed of 60 wt%

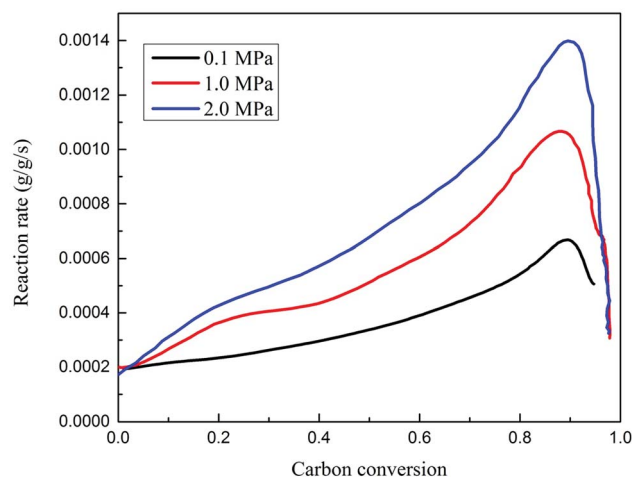


Fig. 4 The reaction rate versus carbon conversion during char gasification, at 100% mole CO₂ and 950 °C, between 0.1 and 2.0 MPa.

CuO and 40 wt% Al₂O₃. The operating conditions for the fuel reactor are set at a temperature of 750–950 °C and a pressure of 0.1–3.0 MPa, and the coal is dried (5% moisture content before entering the fuel reactor). For the air reactor, the temperature is set at 1000 °C and the pressure is 0.1–3.0 MPa. The gasifier is an isothermal reactor, and there are several approaches to balance the heat generation from the gasifier ($\Delta H < 0$), such as: (1) the injection of a reasonable amount of steam into the gasifier; (2) combusting a portion of the coal; or (3) sending a certain fraction of reduced copper-oxides from the air reactor to the gasifier. The air reactor is an adiabatic reactor; in order to maintain the temperature, an amount of air, which is not only used for reacting with oxygen carriers and carbon, but also for carrying extra heat generated from the air reactor, should be fed into the air reactor.

4. Results and discussion

Baseline performance of the fuel reactor

The simulation software sensitivity analysis is used to evaluate the fuel reactor performance under the baseline conditions, *i.e.*, at 3.0 MPa and 850 °C. The ratio (*R*) between oxygen carrier and coal flow rate is varied. The simulation results are summarized in Fig. 5. It can be seen that the minimum ratio (*R*) for the coal to be fully oxidized into CO₂ and H₂O in the fuel reactor is 2.17. A higher *R* value means higher carbon conversion, as well as a higher CO₂ concentration in the fuel gas, which is not preferred, because the fuel reactor is supposed to generate fuel

Table 2 Proximate and ultimate analysis of the coal used

	Proximate analysis				Ultimate (as dry)				
	Moisture	Fixed carbon	Volatiles	Ash	C	H	N	O	S
Coal	10.20	44.65	35.35	9.80	71.72	5.06	1.46	7.75	2.82
	HHV (MJ kg ⁻¹)				27.134				



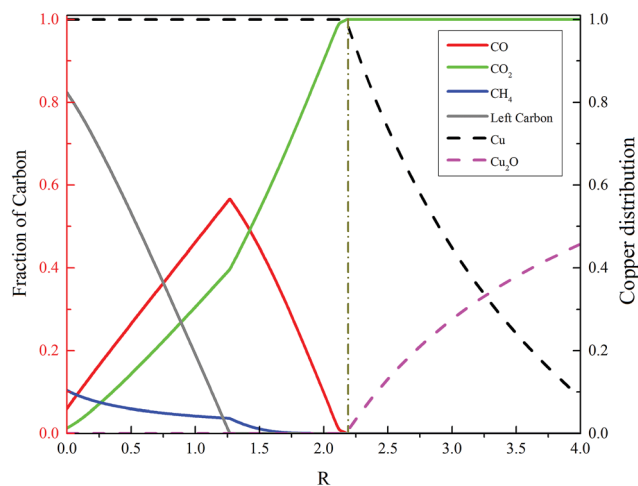


Fig. 5 Carbon distribution and copper oxide conversion under various CuO/coal flow rate ratios at 850 °C and 3.0 MPa in the fuel reactor.

gas. Thus, a reasonable oxygen carrier feed ratio must be maintained to keep an appropriate value of CO₂ concentration in the fuel gas.

The effects of temperature on the process performance

The fuel reactor temperature is a crucial factor for char gasification in interconnected fluidized beds.²⁵ In the cases when the air reactor temperature is 1000 °C, the fuel reactor temperature is varied from 750 to 950 °C, the operating pressure is 2.0 MPa, the steam/carbon ratio is 0.8 and the value of R is 1.20. All these parameters are presented to demonstrate the effect of the fuel reactor temperature, and the simulation results are shown in Fig. 6.

The results in Fig. 6a show that the flow rates of H₂ and CO increase upon increasing the temperature, CO₂ and CH₄ decrease as the temperature rises, and the ratio between the fuel gas (CO, H₂ and CH₄) and CO₂ increases with the elevating of the temperature. They also indicate that the concentrations of H₂ and CO increase as the fuel reactor temperature rises, whilst CO₂ and CH₄ decrease as the temperature increases. Moreover,

the ratio between H₂ and CO decreases with the temperature and the value of H₂/CO is around 1.0 at 850 °C.

The composition of the gas yield from the fuel reactor is based on the combination of a series of complex and competing reactions, as given in reactions (1)–(10). The major reactions (1) and (2) of char gasification, and reactions (3) and (4) in the fuel reactor, are intensive endothermic process, while reactions (5)–(12) are exothermic reactions. A higher temperature is conducive to enrich the reactants in exothermic reactions and the products in endothermic reactions.²⁵ Therefore, reactions (1)–(4) are strengthened with an increase in the fuel reactor temperature, which results in an increase of H₂ and CO, and a decrease of CO₂, and this trend is similar to the effects of temperature on the gas content, as described earlier.²⁵ At a lower fuel reactor temperature, the H₂ concentration is greater than that of CO, and the difference between the gases reduces as the fuel reactor temperature increases. Considering the importance of reactions (1) and (6), the presence of steam favors that reaction, leading to an increase in H₂. Although reaction (6) also releases CO₂ and consumes CO, the concentration of CO₂ decreased and CO increased with a rise in temperature. A reasonable explanation is that reaction (2) consumes CO₂ and becomes the dominant reaction, resulting in a CO increase and a CO₂ decrease as the temperature rises.

There are two paths to using the gas generated from the fuel reactor: a gas turbine electricity generator and chemical production. A high CO₂ concentration is not good for either of the paths; on the other hand, CH₄ is advantageous for a gas turbine electricity generator, but not for chemical production. Additionally, the concentration of CO₂ should be controlled at less than 30% in the gas generation from the fuel reactor, and for chemical production, the CH₄ concentration should be small.

The effect of the fuel reactor temperature on the air feed ratio is shown in Fig. 6b. The air reactor is an adiabatic reactor, and the temperature is controlled *via* the air feed ratio, the temperature of the oxygen carriers from the fuel reactor, and the reaction heat generated from the reactions between Cu, carbon and O₂. The higher operational temperature in the fuel reactor means the oxygen carriers fed into the air reactor have a higher temperature, so more air should be added into the air reactor to

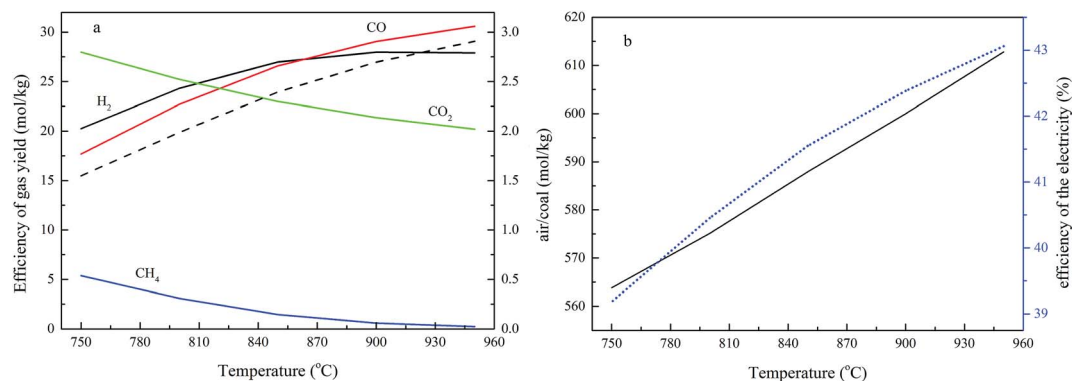


Fig. 6 The effects of temperature on the process performance (a) the effect of temperature on gas yield, dashed line: (CO + H₂ + CH₄)/CO₂, and (b) the effect of temperature on the air ratio and efficiency of electricity generation.



take away the heat from the air reactor. In this study, heat exchangers, heat recovery steam generators, and optional high pressure gas expanders are configured to extract energy from the high temperature exhaust gas streams produced by the CDCLG reactors.

Fig. 6b also shows that the efficiency of electricity generation (η) decreases with an increase in the operational temperature. A higher temperature leads to more fuel gas (CO, H₂ and CH₄) generation, and more energy in the coal being converted to the higher heating value of the fuel gas product.

The effects of pressure on the process performance

The operational pressure is another crucial factor for char gasification in interconnected fluidized beds. In this case, the air reactor temperature is 1000 °C, the fuel reactor temperature is 800 °C, the operating pressure varied from 0.1 MPa to 3.0 MPa, the ratio between steam and carbon is 0.6, R is 1.11%, and the carbon conversion in the fuel reactor is set at 0.9. The results are presented to demonstrate the effect of pressure on the process performance, and are given in Fig. 7.

The results in Fig. 7a show that the flow rates of H₂ and CO decrease as the pressure increases, while CO₂ and CH₄ increase with an increase in pressure, and the ratio between the fuel gas (CO, H₂ and CH₄) and CO₂ increases with rising pressure. As described above, the gas composition of the gas yield from the fuel reactor results from a combination of a series of complex and competing reactions, as given in reactions (1)–(12). In the reactions in which there is an increase in the molecular number of a gas, a higher pressure leads to a lower conversion from the viewpoint of thermodynamics. Thus, in the RGibbs chemical-equilibrium reactor, a higher pressure is negative for the forward reactions of reactions (1)–(4), (7) and (10), which results in a decrease in H₂ and CO, and an increase of CO₂ and CH₄. High pressure is not conducive to both paths described above for utilizing the gas generation from the fuel reactor, because of the higher CO₂ and CH₄ concentrations in the fuel gas. These results show good agreement with Shen's work.²⁵ Higher pressure is not good for the chemical looping process in thermodynamics; higher pressure operation is also found to favor

higher capital and operating costs in the looping process, but higher pressure is propitious to gas compression with lower energy requirements, and favors the kinetics of carbon gasification. Therefore, optimal pressurized conditions are necessary for the chemical looping process. The effects of fuel reactor pressure on the air feed ratio are shown in Fig. 7b. It shows that elevating the pressure requires more air for a constant air reactor temperature. A credible reason is that there is an increase in the compressor discharge temperature as the compressor discharge pressure increases. More air is required for the same coal and oxygen carrier flow rate to keep a constant air reactor temperature. The results in Fig. 7b also show that the efficiency of electricity generation increases first and then decreases upon increasing the operational pressure, and the efficiency of electricity generation reaches a maximum rate when the operational pressure is 1.0 MPa.

The effects of the steam/carbon ratio on the process performance

The effects of the steam/carbon ratio on the gas composition are measured when the air reactor temperature is 1000 °C, the fuel reactor temperature is 800 °C, the operating pressure is 2.0 MPa, the value of R is equal to 1.10, the carbon conversion in the fuel reactor is set at 0.9, and the steam/carbon ratio is varied from 0.4–1.0, and the results are shown in Fig. 8. The effects of the steam/carbon ratio on the efficiency of the gas yield are shown in Fig. 8a. This shows that with an increase in the steam/carbon ratio, the H₂ and CO₂ flow ratios increased, while CO and CH₄ dropped continuously. The results show good agreement with other literature.^{25,26} The performance described above can be explained through the chemical equilibrium of reactions happening in the fuel reactor. The increase in steam means that a higher concentration of reactant in reactions (1), (3) and (6) will promote the gasification reactions in the forward direction. A higher steam/carbon ratio leads to more fuel gas (CO and H₂) generation in the fuel reactor, and more energy in the coal being converted to a higher heat value in the fuel gas product.

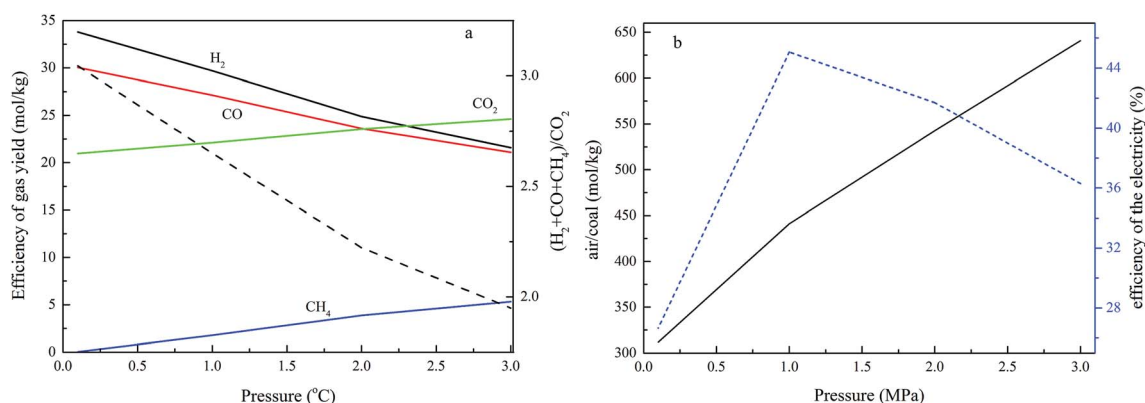


Fig. 7 The effect of pressure on the process performance (a) the effect of pressure on the gas yield, dashed line: (CO + H₂ + CH₄)/CO₂, and (b) the effect of pressure on the air ratio and efficiency of electricity generation.



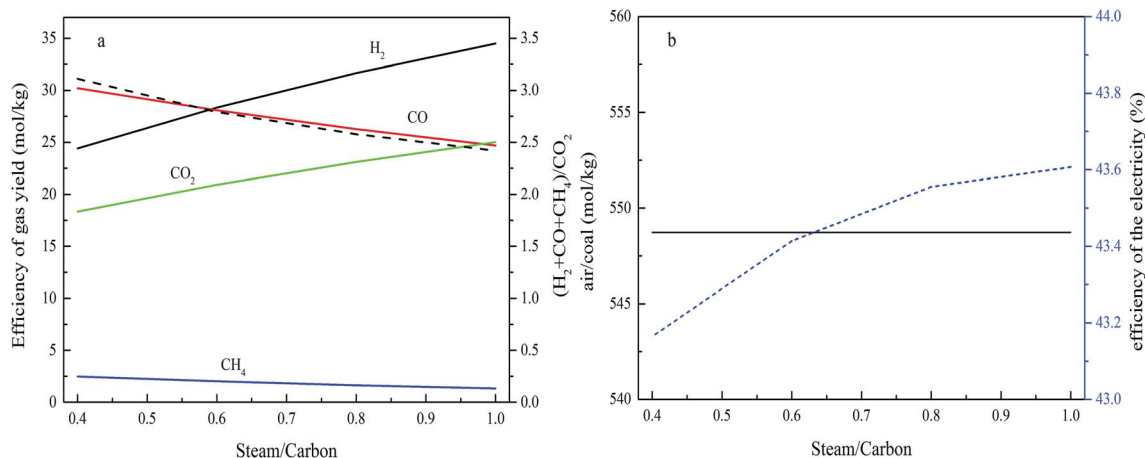


Fig. 8 The effect of the steam ratio on the process performance (a) the effect of pressure on gas yield, dashed line: $(CO + H_2 + CH_4)/CO_2$, and (b) the effect of pressure on the air ratio and efficiency of electricity generation.

Fig. 8b shows that the air feed ratio remains constant as the steam/carbon ratio changes. The reason for this performance is the lack of change of the temperature, pressure and oxygen carrier feed ratio, and the efficiency shows a slight increasing trend with the steam/carbon ratio.

The effects of carbon conversion in the fuel reactor on the process performance

Due to the sudden drop in the gasification kinetics of the char, complete carbon conversion in the fuel reactor is not suggested during the optimal operation of a practical fuel reactor. As for the coal selected in this study, the threshold carbon conversion efficiency is 0.9, based on the gasification kinetics of char, so the gasification conversion of carbon is not suggested to be larger than 0.9, and the exact value of the conversion needs to be determined using other factors. In this study, the basic constant parameters in the simulation are a temperature of 1000 °C in the air reactor and 800 °C in the fuel reactor, an operating pressure of 2.0 MPa, a steam/carbon ratio of 0.6, and an oxygen carrier feed ratio of 1.10. The investigated range of the carbon

conversion efficiency is between 0.8 and 0.9. The effects of carbon conversion in the fuel reactor on process performance are shown in Fig. 9. Fig. 9a shows that the total dry syngas productivity (the sum of H_2 , CO, CO_2 and CH_4) from the fuel reactor increased and the emission leakage of CO_2 from the air reactor decreased when the threshold of carbon conversion efficiency was controlled to show an increase. In this case, less carbon residue is burned out inside the air reactor. Fig. 9b shows that the overall air supply decreased as the carbon conversion efficiency in the fuel reactor increased. Less carbon being converted in the fuel reactor led to more residual carbon being burnt in the air reactor and more being heat produced, thus, more air should be added to the air reactor to react with carbon and carry heat out from the air reactor to maintain the temperature. The efficiency of electricity generation increases upon increasing the carbon conversion in the fuel reactor, because less carbon residue being burned in the air reactor leads to more energy in the coal being converted to chemical energy in the syngas.

Based on the aforementioned analysis, the parameters shown in Fig. 9 are converted to the parameters, x and y , which are based on carbon leakage from the air reactor, and the syngas

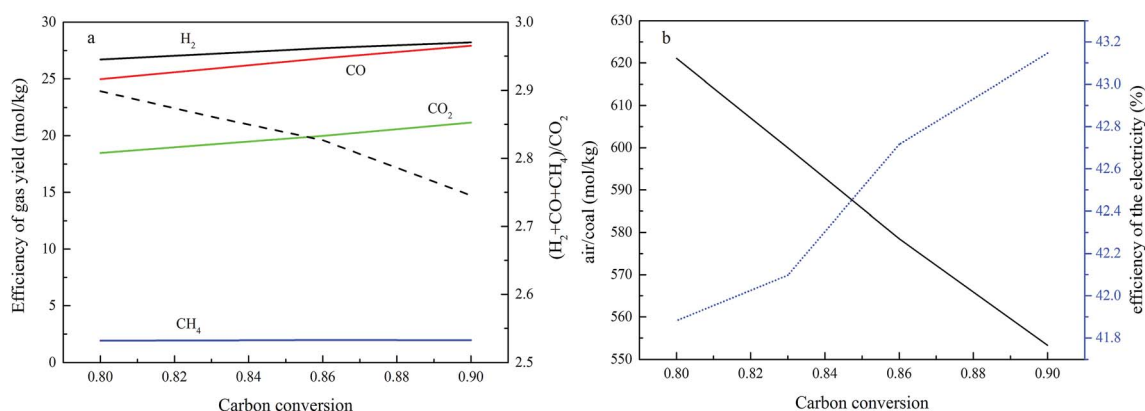


Fig. 9 The effect of carbon conversion on the process performance (a) the effect of carbon conversion on gas yield, dashed line: $(CO + H_2 + CH_4)/CO_2$, and (b) the effect of carbon conversion on the air ratio and efficiency of electricity generation.



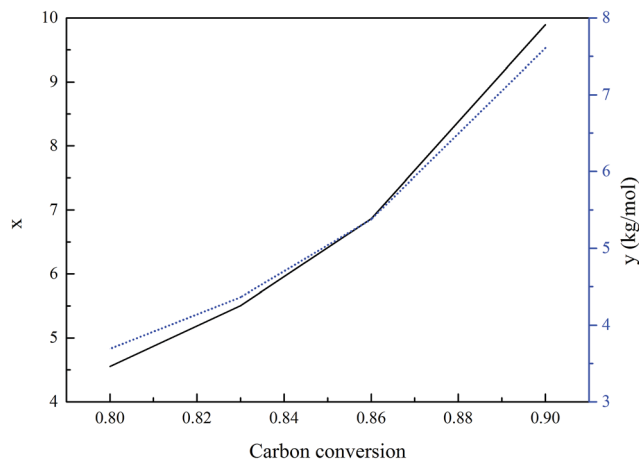


Fig. 10 The effect of carbon conversion on x and y .

yield refers to the total amount of H_2 plus CO . As shown in Fig. 10, a decrease in the carbon conversion efficiency in the fuel reactor decreased both the syngas productivity and the electricity generation efficiency significantly within the range between 0.8 and 0.9, while the required air supply was only decreased slightly in the range between 0.8 and 0.9. An increase in the carbon conversion efficiency inside the fuel reactor is always suggested.

5. Conclusions

A novel process for fuel syngas production from a chemical looping process, in which the carbon is not fully converted in the fuel reactor for fuel gas production and residual carbon is burnt in the air reactor using copper oxide as the oxygen carrier, is proposed. The results from the kinetics of char gasification suggest that the optimal threshold value for carbon conversion efficiency for the selected coal in this study is around 0.9, where further gasification of char residue is not suggested. A simulation of the processes, including chemical reactions and heat/mass balance, is carried out with software. The main objective is to study the effects of various conditions on the energy output parameters, such as energy efficiency and syngas quality, the carbon leakage from the air reactor, and the efficiency of electricity generation. The operating variables considered as part of the resource optimization analysis include: (i) the carbon conversion in the fuel reactor, (ii) the inlet temperature and pressure to the fuel reactor, and (iii) the steam-to-carbon ratio. The results suggest that higher carbon conversion in the fuel reactor is not good for the quality of the syngas, but favors the lower emission of CO_2 and the efficiency of electricity generation. Meanwhile, the coal-direct chemical looping gasification process has the potential to convert coal to syngas and electricity, capturing 90% of CO_2 ; the syngas can be adjusted to a reasonable value by varying temperature, pressure, and the oxygen carrier (OC) and steam feed ratio. The H_2/CO ratio varied from 1.05 to 1.28, the CO_2 concentration varied from 30.74% to 33.74%, the CH_4 was between 4.10% and 5.21%, and the efficiency of the electricity was arranged between 26.6% and 45.5%.

Conflicts of interest

There are no conflicts to declare.

Acknowledgements

This study was supported by the Youth Science and Technology Talents Growth project of the Guizhou Provincial Education Department (Qianjiaohe KY Zi [2017]220), and the Energy Chemistry Laboratory in Guizhou Province (Qianjiaohe KY Zi [2017]009).

References

- 1 I. Ahmed and A. Gupta, Pyrolysis and gasification of food waste: syngas characteristics and char gasification kinetics, *Appl. Energy*, 2010, **87**, 101–108.
- 2 E. Cetin, B. Moghtaderi, R. Gupta and T. Wall, Biomass gasification kinetics: influences of pressure and char structure, *Combust. Sci. Technol.*, 2005, **177**, 765–791.
- 3 H. L. Chum and R. P. Overend, Biomass and renewable fuels, *Fuel Process. Technol.*, 2001, **71**, 187–195.
- 4 K. Mitsuoka, S. Hayashi, H. Amano, K. Kayahara, E. Sasaoaka and M. A. Uddin, Gasification of woody biomass char with CO_2 : The catalytic effects of K and Ca species on char gasification reactivity, *Fuel Process. Technol.*, 2011, **92**, 26–31.
- 5 J. A. Onwudili and P. T. Williams, Enhanced methane and hydrogen yields from catalytic supercritical water gasification of pine wood sawdust *via* pre-processing in subcritical water, *RSC Adv.*, 2013, **3**, 12432–12442.
- 6 T. Mattisson, A. Lyngfelt and H. Leion, Chemical-looping with oxygen uncoupling for combustion of solid fuels, *Int. J. Greenhouse Gas Control*, 2009, **3**, 11–19.
- 7 T. Mattisson, H. Leion and A. Lyngfelt, Chemical-looping with oxygen uncoupling using CuO/ZrO_2 with petroleum coke, *Fuel*, 2009, **88**, 683–690.
- 8 E. M. Eyring, G. Konya, J. S. Lighty and A. H. Sahir, Chemical Looping with Copper Oxide as Carrier and Coal as Fuel Boucle chimique pour la combustion du charbon avec un transporteur d'oxygène à base d'oxyde de cuivre, *Oil Gas Sci. Technol.*, 2011, **66**, 209–221.
- 9 R. A. Gaggioli, Second law analysis for process and energy engineering, in *ACS symposium series*, Oxford University Press, 1983, pp. 3–50.
- 10 C. Yan, W. Yang, J. T. Riley and W. P. Pan, A novel biomass air gasification process for producing tar-free higher heating value fuel gas, *Fuel Process. Technol.*, 2006, **87**, 343–353.
- 11 Q. Song, R. Xiao, Z. Deng, L. Shen, J. Xiao and M. Zhang, Effect of Temperature on Reduction of $CaSO_4$ Oxygen Carrier in Chemical-Looping Combustion of Simulated Coal Gas in a Fluidized Bed Reactor, *Ind. Eng. Chem. Res.*, 2008, **47**, 8148–8159.
- 12 W. Liu, M. Ismail, M. T. Dunstan, W. Hu, Z. Zhang, P. S. Fennell, S. A. Scott and J. S. Dennis, Inhibiting the interaction between FeO and Al_2O_3 during chemical looping production of hydrogen, *RSC Adv.*, 2014, **5**, 1759–1771.



- 13 T. Andersen, H. M. Kvamsdal and O. Bolland, Gas turbine combined cycle with CO₂ capture using auto-thermal reforming of natural gas, *ASME Pap.*, 2000, 0162.
- 14 I. S. Ertesvåg, H. M. Kvamsdal and O. Bolland, Exergy analysis of a gas-turbine combined-cycle power plant with precombustion CO₂ capture, *Energy*, 2005, **30**, 5–39.
- 15 S. Consonni and F. Viganò, Decarbonized hydrogen and electricity from natural gas, *Int. J. Hydrogen Energy*, 2005, **30**, 701–718.
- 16 E. Eyring, G. Konya, J. Lighty, A. Sahir, A. Sarofim and K. Whitty, Chemical looping with copper oxide as carrier and coal as fuel, *Oil Gas Sci. Technol.*, 2011, **66**, 209–221.
- 17 F. Li, Z. Liang, L. G. Velazquez-Vargas, Y. Zachary and L. S. Fan, Syngas chemical looping gasification process: Bench-scale studies and reactor simulations, *AIChE J.*, 2010, **56**, 2186–2199.
- 18 S. G. Gopaul, A. Dutta and R. Clemmer, Chemical looping gasification for hydrogen production: A comparison of two unique processes simulated using ASPEN Plus, *Int. J. Hydrogen Energy*, 2014, **39**, 5804–5817.
- 19 L. Zeng, F. He, F. Li and L.-S. Fan, Coal-Direct Chemical Looping Gasification for Hydrogen Production: Reactor Modeling and Process Simulation, *Energy Fuels*, 2012, **26**, 3680–3690.
- 20 Z. Liang, H. Feng, F. Li and L. S. Fan, Coal-Direct Chemical Looping Gasification for Hydrogen Production: Reactor Modeling and Process Simulation, *Energy Fuels*, 2012, **26**, 3680–3690.
- 21 R. Porrazzo, G. White and R. Ocone, Aspen Plus simulations of fluidised beds for chemical looping combustion, *Fuel*, 2014, **136**, 46–56.
- 22 S. Ilaiyah, D. V. Sasikanth and B. Satyavathi, Process simulation of a entrained fluidized bed biomass gasification using aspen plus, *3rd International Conference on Recent Trends in Engineering Science and Management (ICRTESSM-16)*, Vedant College of Engineering and Technology, Bundi, Rajasthan, 10th April 2016, ISBN: 978-81-932074-4-4.
- 23 L. Liu, Y. Cao and Q. Liu, Kinetics studies and structure characteristics of coal char under pressurized CO₂ gasification conditions, *Fuel*, 2015, **146**, 103–110.
- 24 L. Liu, Y. Cao, Q. Liu and J. Yang, Experimental and kinetic studies of coal-CO₂ gasification in isothermal and pressurized conditions, *RSC Adv.*, 2017, **7**, 2193–2201.
- 25 L. Shen, Y. Gao and J. Xiao, Simulation of hydrogen production from biomass gasification in interconnected fluidized beds, *Biomass Bioenergy*, 2008, **32**, 120–127.
- 26 S. Rapagna, N. Jand, A. Kiennemann and P. Foscolo, Steam-gasification of biomass in a fluidised-bed of olivine particles, *Biomass Bioenergy*, 2000, **19**, 187–197.

