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Electrochemical properties of NiCoO₂ synthesized by hydrothermal method

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NiCoO₂ microspheres were successfully synthesized *via* an easy hydrothermal method, followed by an annealing process at 350 °C under a nitrogen atmosphere. The X-ray diffraction results show that a single phase NiCoO₂ was synthesized. The results of electrochemical measurements show that the NiCoO₂ exhibits a high specific capacitance of 760 and 470 F g⁻¹ at current densities of 1 and 16 A g⁻¹, respectively, and the specific capacitance retention is 40.9% after 5000 cycles at the current density of 16 A g⁻¹. This reveals the great potential of NiCoO₂ as an electrode material for high-performance supercapacitors.

1. Introduction

Clean energy conversion and storage devices have attracted great attention. Currently, supercapacitors, as a clean energy conversion and storage technology, have drawn intensive research attention in the past several years due to their high power densities, short charging times, safe operation mode and long cycle life.^{1–6} The electrode material, as an important component of a supercapacitor, plays a key role to gain high supercapacitor performances.⁷ Many kinds of materials have been considered as the electrode materials for supercapacitors, and they can be divided into three main kinds of electrode materials:⁸ carbonaceous materials,⁹ transition metal oxides¹⁰ and conducting polymer materials.¹¹ However, many carbonaceous materials show low specific capacitance, and the conducting polymer materials exhibit poor electrical conductivity, leading to limited practical applications of these electrode materials for supercapacitors.^{12,13} Therefore, it is important to search for novel electrode materials, which show high electrical conductivity, large specific capacitance and high cyclic stability. Thus transition metal oxides as electrode materials have been widely investigated because of their alternative redox properties and high theoretical specific capacitance values.¹⁴ Especially, some transition metal oxides, such as MnO₂,¹⁵ Co₃O₄ (ref. 16) and NiO^{17,18} have attracted great interest for use as supercapacitor electrode materials because of their low cost, natural abundance and environmental friendliness. Moreover, the nickel-cobalt binary metal oxides, such as NiCo₂O₄ (ref. 19 and 20) and NiCoO₂ (ref. 21–23) have been also studied due to their

higher electrical conductivity and specific capacitance than those of the single component nickel or cobalt oxides.

Recently, much research on the preparation and electrochemical properties of NiCo₂O₄ has been reported.^{24–29} Especially, some artificial nanostructures show excellent properties.^{30–34} Two types of homogeneous NiCo₂O₄ nanosheet@NiCo₂O₄ nanorod and heterogeneous NiCo₂O₄ nanosheet@NiO nanoflake hierarchical core-shell arrays were synthesized *via* facile solution methods followed by a simple thermal treatment, and high specific capacitances of 1925 and 2210 F g⁻¹ were achieved by NiCo₂O₄@NiCo₂O₄ and NiCo₂O₄@NiO at 0.5 A g⁻¹, respectively.³⁵ However, research on NiCoO₂ as an electrode material for supercapacitors is rare.^{2,19,21,36–39} Du *et al.* prepared NiCoO₂ single-crystalline nanoparticles with a specific capacitance of 184 F g⁻¹ at 1 A g⁻¹.⁴⁰ NiCoO₂ was synthesized based on the templates and used for supercapacitors.^{2,21,36–39}

In this study, we successfully synthesized micro-spheres NiCoO₂ by a simple hydrothermal method, followed by calcining at 350 °C under a nitrogen atmosphere. The NiCoO₂ exhibited a specific capacitance of 760 F g⁻¹ at a current density of 1 A g⁻¹ in three electrode measurements. Especially, the specific capacitance retains approximately 40.9% of its maximum value after 5000 cycles at the current density of 16 A g⁻¹.

2. Experimental

2.1 Preparation of NiCoO₂

0.474 g NiCl₂·6H₂O, 0.474 g CoCl₂·6H₂O and 2.0 g urea were dissolved in 30 mL distilled water with magnetic stirring for 30 min at room temperature to form a clear solution. Then, the obtained solution was transferred into a 50 mL Teflon-lined stainless steel autoclave and maintained at 120 °C for 12 h.

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After cooling down to room temperature, the precipitate was centrifuged and washed with distilled water and ethanol. Then, the obtained precipitate was dried at 60 °C for 6 h. Finally, the dried precipitate was calcined at 350 °C for 3 h under nitrogen atmosphere.

2.2 Characterization

The crystal structure was characterized by X-ray diffraction (XRD, Bruker D8 discover) using Cu K α radiation at 40 kV and 40 mA. The morphology was examined by field emission scanning electron microscopy (FE-SEM, Zeiss SUPRA 40) and transmission electron microscopy (TEM, JEM-2100). X-ray photoelectron spectroscopy (XPS, ESCALAB250Xi) was used to characterize valence states of the sample. The specific surface area and pore size distribution of the sample were performed by Brunauer–Emmett–Teller (BET) nitrogen adsorption–desorption and Barrett–Joyner–Halenda (BJH) methods (Autosorb-iQ). The electrochemical characterizations were carried out in a CS 350 electrochemical workstation (CorrTest, Wuhan) at room temperature using 3 M KOH solution as electrolyte solution. A three-electrode system was used in the measurement. Nickel foam coated with electroactive materials was used as the working electrode. Platinum and Hg/HgO electrode (1 M NaOH solution) were served as the counter electrode and the reference electrode, respectively. The working electrode was fabricated by mixing the as-prepared products, carbon blacks and polytetrafluoroethylene (PTFE) with a weight ratio of 70 : 20 : 10 to form homogeneous paste, which was coated onto nickel foam with the area of 1 cm \times 1 cm. After drying in a vacuum oven at 50 °C for 5 h to remove the solvent, the nickel foam was then pressed at 10 MPa for 2 min by bead machine. The mass of active materials loading on the nickel foam with the area of 1 cm \times 1 cm was about 3–5 mg.

Cyclic voltammetry (CV) tests were performed from -0.1 V to 0.8 V at a scan rate of 5, 10, 20, 40, 60 and 100 mV s $^{-1}$. Galvanostatic charge–discharge (GCD) tests were performed at a current density of 1, 2, 4, 6, 10 and 16 A g $^{-1}$ with the potential window from 0 V to 0.55 V. Electrochemical impedance spectroscopy (EIS) measurements were measured at a frequency range from 10^5 Hz to 0.01 Hz with an amplitude of 3 mV.

3. Results and discussion

3.1 Structure characterization

Fig. 1 shows the XRD pattern of the samples calcined at 350 °C under nitrogen atmosphere. The diffraction pattern showed six peaks at 36.80°, 42.82°, 61.79°, 73.99°, 77.99° and 93.22°, corresponding to (111), (200), (220), (311), (222) and (400) of NiCoO $_2$ phase (JCPDS no. 10-0188), respectively. No impurity peaks were observed in the pattern, confirming the single pure phase of the obtained product.

The elemental composition and the valence states of the Ni and Co in the sample were investigated by XPS, and the results are presented in Fig. 2. Fig. 2(a) shows the survey spectrum of the sample. It clearly showed the peaks of Ni, Co, O and C elements. The C element can be attributed to the exposure of

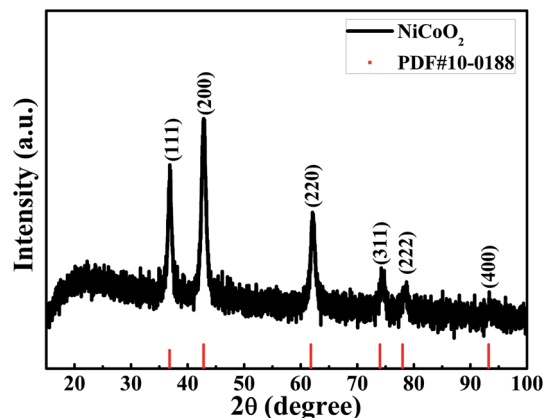


Fig. 1 XRD patterns of the NiCoO $_2$ samples.

air. Fig. 2(b) shows the spectrum of Ni 2p, which was composed of two spin–orbit doublets and two shakeup satellites. As reported in ref. 41 there are four easily discernible features in Ni 2p spectrum, they are the Ni 2p $_{3/2}$ main peak and its satellite, and the Ni 2p $_{1/2}$ main peak and its satellite, respectively. The fitting peaks at 854.2 and 856.0 eV are corresponding to Ni 2p $_{3/2}$, whereas the peaks at 871.8 and 873.7 eV are corresponding to Ni 2p $_{1/2}$. The peaks Sat.1 at 861.2 eV and Sat.2 at 879.1 eV are ascribed to the two shakeup satellites of Ni main peaks. The peak of Ni 2p $_{3/2}$ at 854.2 eV is ascribed to local screening from lattice oxygen adjacent to the Ni 2p core hole.⁴¹ And the existence of peak at 856.0 eV is prevalent, suggesting a higher hydroxylation of NiO in the composite oxide surface,⁴² which can be further confirmed by the O 1s spectrum. The binding energy peaks of Ni 2p $_{1/2}$ are consistent with that of Ni 2p $_{3/2}$. The fitting results show the appropriate binding energies of Ni 2p, indicating that the divalent oxide state of Ni was existed in sample.⁴¹ Fig. 2(c) shows the core level spectrum of Co 2p. The fitting peak at 780.1 eV and 797.0 eV is corresponding to Co 2p $_{3/2}$ and Co 2p $_{1/2}$, respectively, and these two peaks could be indexed to Co $^{2+}$.^{43,44} In addition, the shake-up satellite peak Sat.1 located at 785.8 eV has a separation energy of ~ 5.2 eV with Co 2p $_{3/2}$ at 780.1 eV, while the shake-up satellite peak Sat.2 located at 802.2 eV has a separation energy of ~ 5.2 eV with Co 2p $_{1/2}$ at 797.0 eV. The energy gap between the Co 2p main peak and the satellite peaks is related to the oxidation states, and the energy gaps are about 6.0 eV and 9–10 eV for Co $^{2+}$ and Co $^{3+}$.⁴⁵ The results prove that the divalent oxide state of Co was existed in sample. The high resolution spectrum for O 1s peak in Fig. 2(d) decomposes into two oxygen contributions. The fitting peak located at 529.4 eV (O $_1$) are typical of the metal–oxygen bonds (M = Ni or Co).⁴⁶ The component O $_2$ located at 531.1 eV indicates the presence of surface hydroxyl groups.^{41,42,47,48} According to the XPS analysis, these results show that the NiCoO $_2$ samples contain Ni $^{2+}$ and Co $^{2+}$, the atomic percentage of Ni, Co and O is 23.11, 23.49 and 53.4%, respectively. The calculated atomic ratio of Ni to Co and O is close to 1 : 1 : 2, which is on the verge of the stoichiometry of NiCoO $_2$.

The morphologies of NiCoO $_2$ sample were characterized by SEM and TEM. As shown in Fig. 3(a), NiCoO $_2$ sample showed



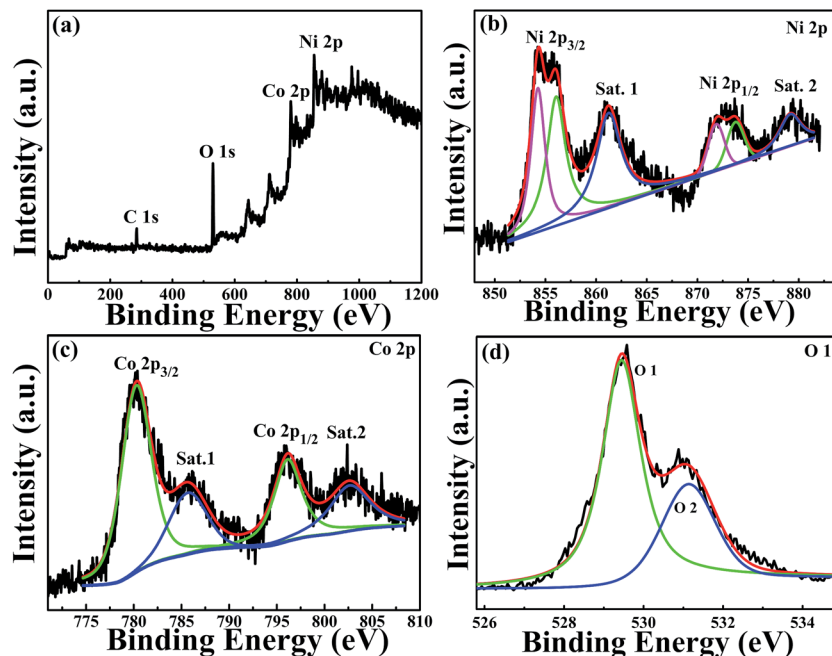


Fig. 2 XPS spectrum of (a) survey spectrum, (b) Ni 2p spectrum, (c) Co 2p spectrum and (d) O 1s spectrum for the NiCoO₂ samples.

the microspheres with different size in the low-magnification SEM image of sample. The diameter of the sphere varied from 1 μm to 6 μm . Furthermore, the microspheres are assembled by a large amount of nanowires according to the high-magnification SEM images as shown in Fig. 3(b). To identify the distribution of element in sample, the element mappings of Ni, Co and O were examined by energy dispersive spectroscopy (EDS). Fig. 3(c) is a typical microsphere of sample, and the element mapping of Ni, Co and O were included. It can be seen clearly that Ni, Co and O elements are uniformly distributed in the microsphere, and that the ratio of three elements is around 1 : 1 : 2, which corresponded with the composition of NiCoO₂. Fig. 3(d)–(f) show TEM images of the samples. Pompon shape was observed in samples, and this is consistent with the morphology results of SEM. The high-magnification TEM image revealed that the nanowires in pompon were composed of nanoparticles. Fig. 3(e) selected area electron diffraction (SAED) pattern showed ring patterns, which revealed that the nanowire in pompon has a polycrystalline structure. These rings could be assigned to the (111), (200), (220), (311), (222) and (400) crystal planes of NiCoO₂, which is consistent with the observation of XRD results. The lattice spacing of 0.208 nm is attributed to the (200) crystal plane of NiCoO₂ sample, as shown in the HRTEM image of Fig. 3(f).

The results of XRD, TEM, EDS and XPS confirm that single pure phase of NiCoO₂ with pompon shape was obtained in present research, and the pompon with different diameter is composed of nanowires, which were composed of NiCoO₂ nanoparticles.

The porous structure of sample was evaluated by Brunauer–Emmett–Teller (BET) N₂-adsorption–desorption analysis. The N₂ adsorption–desorption isotherm and pore size distribution

were shown in Fig. 4. The BET specific surface area is about 88.3 m² g⁻¹. As can be seen from Fig. 4, the typical type IV hysteresis loop can be observed in the region of P/P_0 from 0.6 to 1.0, which illustrates the existence of a mesoporous structure in sample.^{43,49} The pore size distributions, which were shown in insert of Fig. 4, were calculated by the Barrett–Joyner–Halenda (BJH) method. It shows that sharp peaks are located at 2.7 nm and 11.0 nm, illustrating a few mesoporous were presented in NiCoO₂ samples. Such results are consistent with the results of SEM and TEM results. The rich internal mesoporous structure could increase the contact area between electrode and electrolyte, leading to more sufficient active sites for redox reaction and efficient transport of electrons and ions.^{43,49} Therefore, the high surface area and mesoporous structure of NiCoO₂ sample are benefic for application in supercapacitor.^{43,49,50}

3.2 Electrochemical characterization

Fig. 5(a) shows the CV curves of the NiCoO₂ micro pompon at different scan rates from 5 mV s⁻¹ to 100 mV s⁻¹ within potential range from -0.1 V to 0.8 V (vs. Hg/HgO) in 3 M KOH solution. All the curves show typical faradic reaction peaks, corresponding to the M–O/M–O–OH (M represents Ni and Co ions).^{39,51,52} With the increase of scan rate, a shift toward higher and lower potentials was observed for the oxidation and reduction peaks, respectively. As shown in inset of Fig. 5(b), this shift leading to the increase of the potential difference between redox peaks. It indicated that the irreversible degree of the quasi-reversible reaction increased with the increase of scan rate. The shift is mainly related to the internal resistance of the electrode and the polarization caused by high scan rate.⁵³

The specific capacitance of samples can be calculated by using the following equation:



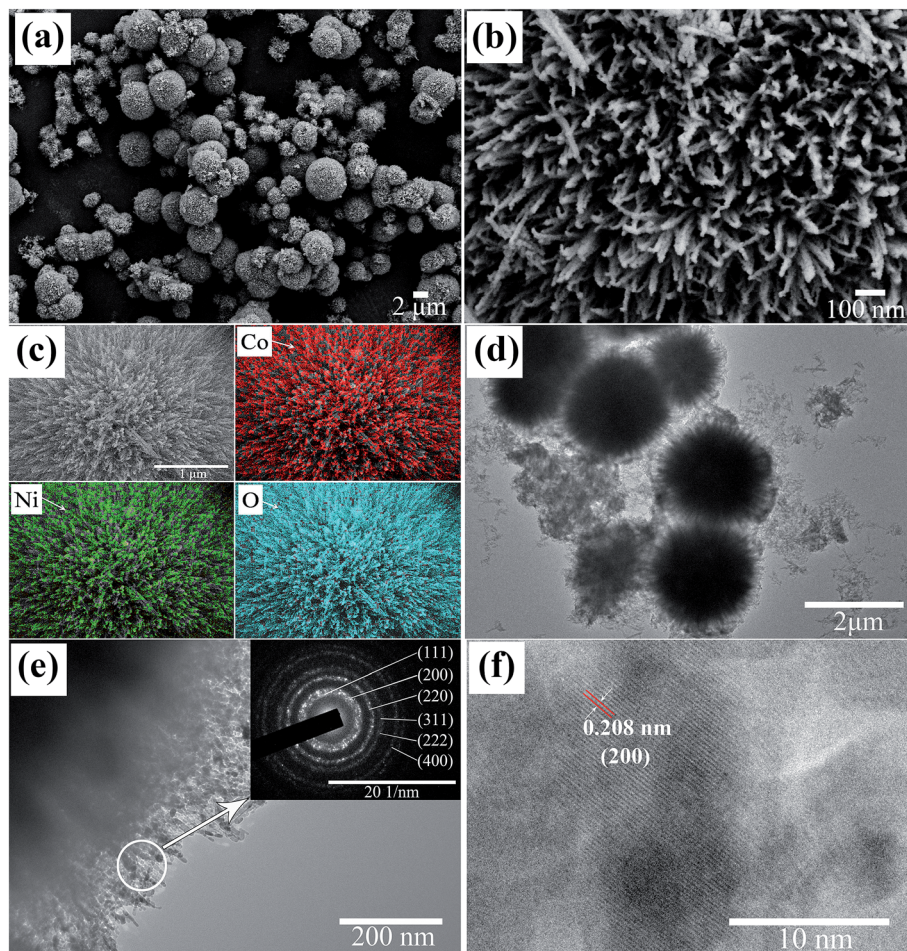


Fig. 3 (a) Low-magnification SEM, (b) high-magnification SEM, (c) energy dispersive spectroscopy (EDS) element mapping images of the NiCo₂O₄ sample and the Ni, Co and O distribution, (d) low-magnification TEM, (e) high-magnification TEM (inset showing the selected area electron diffraction ring) and (f) HRTEM image of NiCo₂O₄ sample.

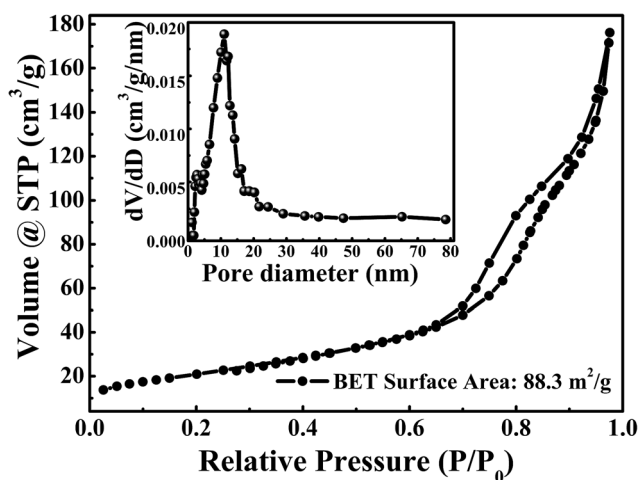


Fig. 4 The N₂ adsorption–desorption isotherms of the NiCo₂O₄. The inset is the pore size distributions are calculated by the Barrett–Joyner–Halenda (BJH) method.

$$C_{sp} = \frac{\int I_{(V)} dV}{2m \times v \times \Delta V}, \quad (1)$$

where C_{sp} (F g⁻¹) is the specific capacitance, $I_{(V)}$ (A) is cathodic current, m (g) is the mass of the active material used to prepare the electrode, v (mV s⁻¹) is the potential scan rate, and V (V) is the sweep potential window. Fig. 5(b) presents specific capacitance dependent on the potential scan rate. The specific capacitance decreased with the increase of scan rate. The specific capacitance was 1461 F g⁻¹ at a scan rate of 5 mV s⁻¹, and it decreases sharply to be 924 F g⁻¹ at a scan rate of 10 mV s⁻¹, then the capacitance decreases slightly with the continue increase of scan rates. Anion of OH⁻ could reach the electro-active sites adequately for slow scan rate, but only the electrode surface could be utilized to store charges during the electrochemical reaction process for fast scan rate, leading to the decrease of specific capacitance with increase of scan rate.⁴²

As shown in Fig. 6(a), the GCD curves of NiCo₂O₄ electrodes were measured at different current densities of 1, 2, 4, 6, 10 and 16 A g⁻¹ in 3 M KOH solution in the voltage range of 0–0.55 V (vs. Hg/HgO). The nonlinear curve indicated the



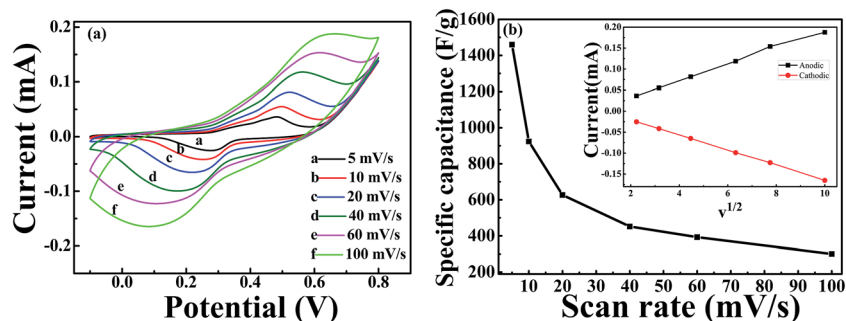


Fig. 5 (a) CV curves under different scan rates of 5, 10, 20, 40, 60, 100 mV s⁻¹ for NiCo₂O₄ samples in the potential window ranged from -0.1 to 0.8 V; (b) the specific capacitance of NiCo₂O₄ sample under different scan rates of 5, 10, 20, 40, 60 and 100 mV s⁻¹. The inset in (b) is the plots of anodic and cathodic peak current against the square root of the scan rate.

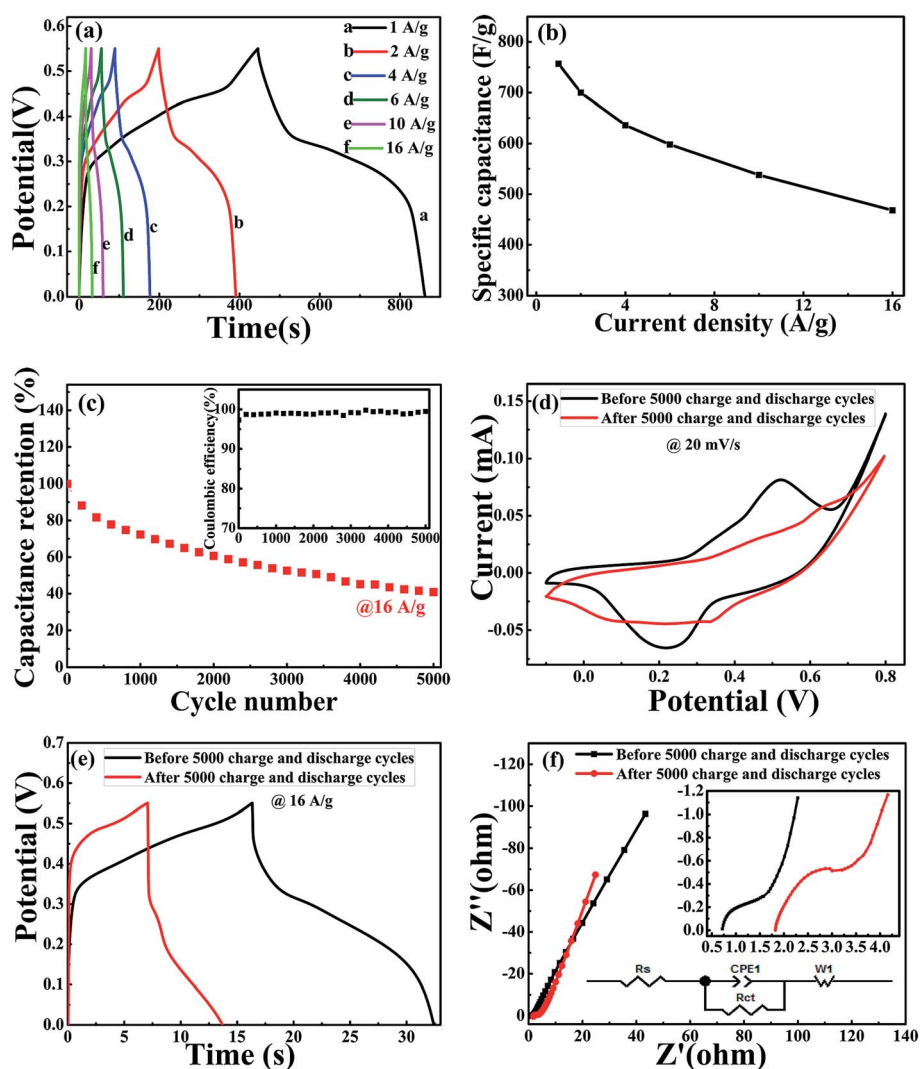


Fig. 6 (a) Galvanostatic charge–discharge curves are within the potential range of 0 to 0.55 V at various current densities from 1 A g⁻¹ to 16 A g⁻¹, (b) specific capacitance at various current densities from 1 A g⁻¹ to 16 A g⁻¹. (c) Cycling performance of the NiCo₂O₄ electrode at 16 A g⁻¹, the insert shows the corresponding coulombic efficiency of NiCo₂O₄ during cycling test. (d) CV and (e) GCD curves of NiCo₂O₄ electrode before and after 5000 charge–discharge cycles. (f) Nyquist plots obtained for NiCo₂O₄ samples at the frequency ranging from 0.01 Hz to 100 KHz in 3 M KOH. Insets: the local amplification of the Nyquist plots at high-frequency.



pseudocapacitance behavior, and the potential plateaus corresponded with the redox reaction, which was consistent with the CV curves in Fig. 5(a). The specific capacitance could be calculated from the galvanostatic charge–discharge curves according to the following equation:

$$C_m = \frac{I \times \Delta t}{m \times \Delta V}, \quad (2)$$

where C_m (F g^{-1}) is the specific capacitance, I (A) refers to the current density used for the charge–discharge measurements, Δt (s) refers to the time elapsed for the discharge cycle in the potential window, ΔV (V) represents the voltage interval of the discharge measurement, and m (g) denotes the mass of active material in the working electrode. Fig. 6(b) shows the dependence of the specific capacitance on the current density. The specific capacitance decreased with the increase of current density, which could be attributed to the reduce of effective contact between ions and electroactive sites of electrode at high current rate.⁵⁴ The specific capacitance could be calculated to 760, 700, 636, 598, 538 and 470 F g^{-1} at various current densities from 1, 2, 4, 6, 10 and 16 A g^{-1} , respectively. The specific capacitance showed small loss under the current density increasing, and it was only 38% when the current density increased from 1 to 16 A g^{-1} . The high performance at high current density could be attributed to the excellent micro-structure. The pompon assembled with nanowire facilitates ion permeability, and the pompon could be served as ion reservoir. The diffusion distance of electrolyte to electro active sites is decreased largely, as is beneficial for the rapid redox reaction at high current density.^{43,49} A cycle stability of NiCoO₂ electrode was also evaluated by the repeated charge–discharge measurement at 16 A g^{-1} for 5000 cycles. As shown in Fig. 6(c), the specific capacitance of the samples decreased slowly, and approximate 40.9% of the initial specific capacitance was retained after 5000 cycles, indicating fine long-cycle stability in NiCoO₂ sample with pompon structure. The insert in Fig. 6(c) is the coulombic efficiencies of NiCoO₂ during the 5000 charge–discharge cycles, it was remained around 97.3% to 99.8%. The fine long-term cycle properties might attributed to the low internal electron transport resistance during the charge–discharge process.⁵⁵ Furthermore, the CV curves at 20 mV s^{-1} and GCD curves at 16 A g^{-1} of NiCoO₂ electrode after 5000 charge–discharge cycles were investigated. Comparing to that before 5000 cycles, the electrode shows weaken electrochemical properties, which can be seen in Fig. 6(d) and (e). The CV and GCD curves all show around 40% retention of specific capacitance after the 5000th cycle.

The electrochemical impedance spectroscopy (EIS) was used to evaluate the electrical resistance responses of the NiCoO₂ electrode materials. Fig. 6(f) shows the Nyquist plots of NiCoO₂ before and after 5000 charge–discharge cycle measurement. The EIS spectra could be fitted according to the equivalent circuit diagram inserted in Fig. 6(f). R_s is the internal resistances of the electrode, including the total resistances of the electrode resistance, the bulk electrolyte resistance and the resistance at electrolyte–electrode interface,⁵⁶ and R_{ct} is the charge-transfer resistances, corresponding to the semicircle in the high-

frequency region.^{57,58} The fitted R_s and R_{ct} values are very small, and the low values are beneficial for long-cycle stability. After 5000 cycles, the R_s value increased from 0.71 Ω to 1.81 Ω , and the R_{ct} value increased from 0.45 Ω to 0.84 Ω . The linear part at low frequency region representative of the typical Warburg resistance (W_1),⁵⁹ and the CPE_1 represents double layer capacitance caused by the rough surface of NiCoO₂ samples and inhomogeneous distribution of active sites.³⁷ For an ideal supercapacitor, the Nyquist plot should be a vertical line at a low frequency, which is parallel to the imaginary impedance axis.³⁴ The slope of line increased from 2.32 to 2.67. The curve becomes more vertical in the low-frequency range after 5000 cycles, indicating that the electro active material becomes more capacitive after long-term cycling.⁶⁰ EIS results further indicate that NiCoO₂ sample with micro pompon structure show low internal and charge-transfer resistance, and stable electrochemical stability.

4. Conclusions

In summary, NiCoO₂ micro pompon sample were successfully prepared by a facile hydrothermal method. The micro pompon was assembled with nanowire, and the nanowire is composed with NiCoO₂ nanoparticles. Electrochemical measurement results display that the micro-sphere NiCoO₂ materials have the remarkable electrochemical behavior. The specific capacitance was 760 F g^{-1} at a current density of 1 A g^{-1} . The specific capacitance showed small loss under the current density increasing, and it was only 38% when the current density increased from 1 to 16 A g^{-1} . In addition, the NiCoO₂ micro pompon with low internal and charge-transfer resistance exhibited fine long-cycle stability, around 40.9% of its initial specific capacitance was maintained after 5000 charge–discharge cycles. The excellent electrochemical properties make the NiCoO₂ micro pompon sample be the promising candidate electrodes materials of supercapacitor in practical applications.

Conflicts of interest

There are no conflicts to declare.

Acknowledgements

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References

- 1 F. Béguin, V. Presser, A. Balducci and E. Frackowiak, *Adv. Mater.*, 2014, **26**, 2219–2251.
- 2 X. Xu, H. Zhao, J. K. Zhou, R. Xue and J. Gao, *J. Power Sources*, 2016, **329**, 238–246.



- 3 W. Chen, C. Xia and H. N. Alshareef, *ACS Nano*, 2014, **8**, 9531–9541.
- 4 M. Jana, S. Saha, P. Samanta, N. C. Murmu, N. H. Kim, T. Kuila and J. H. Lee, *J. Mater. Chem. A*, 2016, **4**, 2188–2197.
- 5 D. Li, Y. Gong, M. Wang and C. Pan, *Nano-Micro Lett.*, 2017, **9**, 16.
- 6 J. Ji, L. L. Zhang, H. Ji, Y. Li, X. Zhao, X. Bai, X. Fan, F. Zhang and R. S. Ruoff, *ACS Nano*, 2013, **7**, 6237–6243.
- 7 X. L. Guo, X. Y. Liu, X. D. Hao, S. J. Zhu, F. Dong, Z. Q. Wen and Y. X. Zhang, *Electrochim. Acta*, 2016, **194**, 179–186.
- 8 X. Yang, K. Xu, R. Zou and J. Hu, *Nano-Micro Lett.*, 2016, **8**, 1–8.
- 9 L. L. Zhang and X. S. Zhao, *Chem. Soc. Rev.*, 2009, **38**, 2520–2531.
- 10 H. B. Wu, H. Pang and X. W. Lou, *Energy Environ. Sci.*, 2013, **6**, 3619–3626.
- 11 L. Nyholm, G. Nyström, A. Mihranyan and M. Strømme, *Adv. Mater.*, 2011, **23**, 3751–3769.
- 12 M. Jana, S. Saha, P. Samanta, N. C. Murmu, J. H. Lee and T. Kuila, *Mater. Chem. Phys.*, 2015, **151**, 72–80.
- 13 P. Simon and Y. Gogotsi, *Nat. Mater.*, 2008, **7**, 845–854.
- 14 H. Gao, G. Wang, M. Yang, L. Tan and J. Yu, *Nanotechnology*, 2012, **23**, 015607.
- 15 Y. Sun, Z. Fang, C. Wang, K. R. Ariyawansa, A. Zhou and H. Duan, *Nanoscale*, 2015, **7**, 7790–7801.
- 16 L. Hou, C. Yuan, L. Yang, L. Shen, F. Zhang and X. Zhang, *RSC Adv.*, 2011, **1**, 1521–1526.
- 17 G. Chen, H. Guan, C. Dong, X. Xiao and Y. Wang, *J. Phys. Chem. Solids*, 2016, **98**, 209–219.
- 18 Q. Sun and S. Bao, *Nano-Micro Lett.*, 2013, **5**, 289–295.
- 19 Y. Zhang and Z. Guo, *Mater. Chem. Phys.*, 2016, **171**, 208–215.
- 20 M. Kuang, Y. X. Zhang, T. T. Li, K. F. Li, S. M. Zhang, G. Li and W. Zhang, *J. Power Sources*, 2015, **283**, 270–278.
- 21 X. Xu, H. Zhou, S. Ding, J. Li, B. Li and D. Yu, *J. Power Sources*, 2014, **267**, 641–647.
- 22 X. Xu, B. Dong, S. Ding, C. Xiao and D. Yu, *J. Mater. Chem. A*, 2014, **2**, 13069–13074.
- 23 X. Xu, S. Chen, C. Xiao, K. Xi, C. Guo, S. Guo, S. Ding, D. Yu and R. V. Kumar, *ACS Appl. Mater. Interfaces*, 2016, **8**, 6004–6010.
- 24 Q. Wang, X. Wang, B. Liu, G. Yu, X. Hou, D. Chen and G. Shen, *J. Mater. Chem. A*, 2013, **1**, 2468–2473.
- 25 Y. Li, M. Zhang, D. Pan, Y. Wang, J. Xie, Z. Yan and J. Jing, *J. Alloys Compd.*, 2017, **722**, 903–912.
- 26 W. Jiang, F. Hu, S. Yao, Z. Sun and X. Wu, *Mater. Res. Bull.*, 2017, **93**, 303–309.
- 27 M. Yao, N. Wang, J. Yin and W. Hu, *J. Mater. Sci.: Mater. Electron.*, 2017, **28**, 11119–11124.
- 28 K. Xu, X. Yang, J. Yang and J. Hu, *J. Alloys Compd.*, 2017, **700**, 247–251.
- 29 S. Han, L. Lin, K. Zhang, L. Luo, X. Peng and N. Hu, *Mater. Lett.*, 2017, **193**, 89–92.
- 30 M. Shahraki, S. Elyasi, H. Heydari and N. Dalir, *J. Electron. Mater.*, 2017, **46**, 1–7.
- 31 Y. Zhang, Y. Zhang, D. Zhang and L. Sun, *Dalton Trans.*, 2017, **46**, 9457–9465.
- 32 H. Xin, Z. Xu, Y. Liu, W. Li and Z. Hu, *J. Alloys Compd.*, 2017, **711**, 670–676.
- 33 T. Han, *Int. J. Electrochem. Sci.*, 2017, **12**, 4724–4732.
- 34 V. Venkatachalam, A. Alsalmeh, A. Alghamdi and R. Jayavel, *Ionics*, 2017, **23**, 977–984.
- 35 W. Zhou, D. Kong, X. Jia, C. Ding, C. Cheng and G. Wen, *J. Mater. Chem. A*, 2014, **2**, 6310–6315.
- 36 X. W. Wang, X. E. Wang, Y. P. Liu, Y. Y. Kong, L. Y. Sun, Y. C. Hu and Q. Q. Zhu, *J. Mater. Sci.: Mater. Electron.*, 2017, 1–7, DOI: 10.1007/s10854-017-7364-x.
- 37 Y. Chen, Y. Xu, J. Wei, L. Tan and J. Yu, *J. Mater. Chem. A*, 2015, **3**, 7121–7131.
- 38 Y. Ma, P. Liu, Q. Liu, Z. Chen, D. Qu and J. Shi, *RSC Adv.*, 2016, **6**, 74643–74648.
- 39 J. Liang, K. Xi, G. Tan, S. Chen, T. Zhao, P. R. Coxon, H. K. Kim, S. Ding, Y. Yang and R. V. Kumar, *Nano Energy*, 2016, **27**, 457–465.
- 40 W. Du, Y. Gao, Q. Tian, D. Li, Z. Zhang, J. Guo and X. Qian, *J. Nanopart. Res.*, 2015, **17**, 1–12.
- 41 M. A. Peck and M. A. Langell, *Chem. Mater.*, 2013, **24**, 4483–4490.
- 42 K. Wang, Z. Zhang, X. Shi, H. Wang, Y. Lu and X. Ma, *RSC Adv.*, 2014, **5**, 1943–1948.
- 43 Y. Lei, J. Li, Y. Wang, L. Gu, Y. Chang, H. Yuan and D. Xiao, *ACS Appl. Mater. Interfaces*, 2014, **6**, 1773–1780.
- 44 X. Leng, Y. Shao, L. Wu, S. Wei, Z. Jiang, G. Wang, Q. Jiang and J. Lian, *J. Mater. Chem. A*, 2016, **4**, 10304–10313.
- 45 B. Babakhani and D. G. Ivey, *Electrochim. Acta*, 2011, **56**, 4753–4762.
- 46 M. Li, J. P. Cheng, J. Wang, F. Liu and X. B. Zhang, *Electrochim. Acta*, 2016, **206**, 108–115.
- 47 Y. Yang, M. Zhou, W. Guo, X. Cui, Y. Li, F. Liu, P. Xiao and Y. Zhang, *Electrochim. Acta*, 2015, **174**, 246–253.
- 48 M. Li, J. P. Cheng, F. Liu and X. B. Zhang, *Electrochim. Acta*, 2015, **178**, 439–446.
- 49 X. Liu, J. Zhao, Y. Cao, W. Li, Y. Sun, J. Lu, Y. Men and J. Hu, *RSC Adv.*, 2015, **5**, 47506–47510.
- 50 X. W. Wang, D. L. Zheng, P. Z. Yang, X. E. Wang, Q. Q. Zhu, P. F. Ma and L. Y. Sun, *Chem. Phys. Lett.*, 2016, **667**, 260–266.
- 51 X. W. Wang, Q. Q. Zhu, X. E. Wang, H. C. Zhang, J. J. Zhang and L. F. Wang, *J. Alloys Compd.*, 2016, **675**, 195–200.
- 52 X. W. Wang, X. E. Wang, H. C. Zhang, Q. Q. Zhu, D. L. Zheng and L. Y. Sun, *Key Eng. Mater.*, 2017, **727**, 698–704.
- 53 G. Zhang and X. W. Lou, *Adv. Mater.*, 2013, **25**, 976–979.
- 54 G. Zhang, W. Li, K. Xie, F. Yu and H. Huang, *Adv. Funct. Mater.*, 2013, **23**, 3675–3681.
- 55 S. Peng, L. Li, C. Li, H. Tan, R. Cai, H. Yu, S. Mhaisalkar, M. Srinivasan, S. Ramakrishna and Q. Yan, *Chem. Commun.*, 2013, **49**, 10178–10180.
- 56 M. Fan, B. Ren, L. Yu, D. Song, Q. Liu, J. Liu, J. Wang, X. Jing and L. Liu, *Electrochim. Acta*, 2015, **166**, 168–173.
- 57 C. Yuan, X. Zhang, Q. Wu and G. Bo, *Solid State Ionics*, 2006, **177**, 1237–1242.
- 58 Q. Y. Shan, X. L. Guo, F. Dong and Y. X. Zhang, *Mater. Lett.*, 2017, **202**, 103–106.
- 59 D. Du, Z. Hu, Y. Liu, Y. Deng and J. Liu, *J. Alloys Compd.*, 2014, **589**, 82–87.
- 60 H. Chen, J. Jiang, Y. Zhao, L. Zhang, D. Guo and D. Xia, *J. Mater. Chem. A*, 2014, **3**, 428–437.

