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(4-Methoxy)-phenoxy pentafluorocyclotriphosphazene as a novel flame retardant and overcharge protection additive for lithium-ion batteries†

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We evaluate (4-methoxy)-phenoxy pentafluorocyclotriphosphazene (4-MPPFPP) as a novel flame retardant and overcharge protection additive for lithium-ion batteries. The flammability of electrolytes containing 4-MPPFPP was studied by measuring the self-extinguishing time of the electrolytes. We found that 4-MPPFPP suppresses the flammability of the electrolyte. Moreover, 4-MPPFPP can be electrochemically polymerized on the surface of an electrode. This will protect batteries from voltage runaway. The influence of proper 4-MPPFPP addition on the cycling performance and capacity is negligible. The results may be helpful for the further design of advanced lithium ion battery electrolytes. We discuss the effect of the additive on the electrode/electrolyte interface and propose possible reaction paths.

Introduction

Lithium ion batteries are now widely used in various portable electronics and are promising candidates for energy storage and electric vehicles because of their good gravimetric and volumetric energy densities. The safety of lithium ion batteries is critical to their commercial use. Lithium ion batteries use highly flammable and carbonate-based organic electrolytes. These may be ignited and possibly cause serious hazards such as fires and explosion under abused conditions (heat, overcharge, short-circuit, *etc.*). To solve these problems, much effort has been focused on developing safety additives such as flame-retardants and overcharge protection additives.

Many flame retardant compounds have been investigated including phosphates,^{1–5} phosphite,⁶ phosphazenes,⁷ fluorinated ethers,^{8–10} fluorinated carbonates,¹¹ ionic liquids,^{12,13} *etc.* Of these, most flame retardants are based on phosphate or P(v) compounds because they suppress flame propagation.

As an overcharge protection additive, various compounds have been tested to protect the cell by redox shuttles. They consume current *via* a reversible redox reaction that occurs between the anode and cathode. Most of these are only practical for the overcharge protection of LiFePO₄ due to the low redox potential.^{13–19} Additives such as dimethoxydiphenylsilane (DDS), resorcinol bis(diphenyl phosphate) (RDP), and *N*-ethyl-2-pyrrolidone (NEP) have also been investigated for overcharge

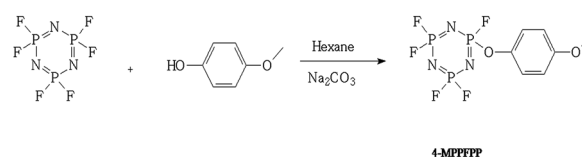
protection.^{20–22} These compounds can be electrochemically polymerized on the cathode during overcharging. The resulting polymer film acts as a passivation layer to increase the internal impedance and will eventually shut down the LIB.

If an additive such as resorcinol bis(diphenyl phosphate) (RDP)²¹ or tri-(4-methoxythphenyl) phosphate (TMPP)²³ can both be a flame retardant and offer overcharge protection, then it would provide more reliable protection. In this paper, (4-methoxy)-phenoxy pentafluorocyclotriphosphazene (4-MPPFPP) was synthesized as a flame retardant and overcharge protection additive in LIB electrolytes. This additive can suppress the ignitability of the electrolyte and polymerize to form a conductive polymer, bypassing the reaction current. The possible reaction paths and the electrochemical compatibility of the electrolyte with the additive are also discussed.

Experimental

4-MPPFPP synthesis

The N₃P₃F₅OC₆H₄OCH₃ (4-MPPFPP) was synthesized as shown in Scheme 1. In a 500 ml three-mouth round bottom flask equipped with an electric mixer, a thermometer and a reflux



Scheme 1 Synthetic route of (4-methoxy)-phenoxy pentafluorocyclotriphosphazene.

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condenser, 250 ml hexane, 23.8 g (0.225 mol) sodium carbonate, 50 g (0.20 mol) hexafluorocyclotriphosphazene, and 27.5 g (0.22 mol) 4-methoxyphenol were added. These materials were stirred at 60 °C for 6 hours, and the resulting mixture was allowed to cool to room temperature. The mixture was then poured into 1000 ml cold water, and the organic layer was separated, washed with 10% (wt) sodium hydroxide aqueous solution and water, dried over anhydrous magnesium sulfate, and filtered. The solvent was distilled under ambient pressure. The product was distilled under vacuum to obtain 56.2 g (92–94 °C/18 Torr) 4-MPPFP. ¹H-NMR (400 MHz, CDCl₃, ppm): 7.12 (d, *J* = 8.0 Hz, 2H), 6.89 (d, *J* = 8.0 Hz, 2H), 3.80 (s, 3H). ¹³C-NMR (400 MHz, CDCl₃, ppm): 157.95 (d), 142.15 (d), 121.54 (d), 114.99 (d), 55.60. ¹⁹F-NMR (400 MHz, CDCl₃, ppm): –65.75 to –66.00 (m), –67.61 to –68.47 (m), –70.17 to –70.76 (m). ³¹P-NMR (400 MHz, CDCl₃, ppm): 14.57–16.46 (m), 11.82–13.82 (m), 8.69–10.91 (m), 6.15–8.22 (m), 3.13–5.03 (m). See Fig. S1–S4 in the ESI.†

Electrolyte and electrode preparation

1 M LiPF₆/EC + DMC + EMC (1 : 1 : 1, wt%) was selected as a base electrolyte. The 4-MPPFP was added directly to the 1 M LiPF₆/EC + DMC + EMC electrolyte. This was prepared in an argon-filled glove box (the oxygen and water content were less than 5 ppm).

The LiMn₂O₄ electrode consisted of LiMn₂O₄ (80 wt%), acetylene black (10 wt%) and poly(vinylidene fluoride) (PVDF) (10 wt%). These were made on aluminum. The separators consisted of 2025-type coin cells of LiMn₂O₄/Li with Celgard 2325 microporous membrane as the separator and were assembled in the glove box. The MCMB/Li cells were fabricated similarly to the LiMn₂O₄/Li cells and were made on copper.

Characterization

The procedure for self-extinguishing measurements is to use fiber glass (–13 mm in length) soaked in 15 ml of electrolyte for 24 h. The fiber glass was then exposed to a burner for 15 s to allow ignition and the burning time of the electrolyte was recorded. Each test was repeated four times, and the burning times were averaged.

The conductivity of the electrolyte with and without 4-MPPFP was measured and calibrated with reference to the conductivity of 0.1 M KCl solution, using a conductivity measuring meter (DDS-307, Shanghai, China).

The overcharge behavior and cycling measurements of the LiMn₂O₄/Li cells were tested using the LAND cell test system (Land CT 2001A, China). The initial twelve formation cycles were as follows: 0.1C (1C = 148 mA g^{–1}) three times, 0.2C three times, 0.5C three times, and 1.0C three times at 25 °C under constant current conditions from 3.0–4.3 V. The cells were charged and discharged at 0.5C between 3.0 and 4.3 V and then overcharged to 5.0 V at 1C. To evaluate the influence of 4-MPPFP on the cycling performance of the LiMn₂O₄/Li and MCMB/Li cells, cycling tests were carried out with a constant current of 1C from 3.0–4.3 V at 55 °C. The MCMB/Li cells were first cycled at 0.2C (1C = 372 mA g^{–1}) five times for formation and then at 0.5C at 25 °C for the cycling performance test. Rate

performances of the LiMn₂O₄/Li cells were obtained by cycling the cells for three cycles of 0.1C, 0.2C, 0.5C, 1C, 2C and 5C for charge and discharge, respectively.

The electrochemical impedance spectroscopy of the cells before and after the overcharge test was performed *via* a frequency response analyzer (VSP, Bio-logic). The frequency range was from 100 kHz to 10 mHz, and the amplitude was 5 mV.

The morphological changes at the electrodes after overcharging were observed. The electrodes were rinsed with dimethyl carbonate (DMC) three times prior to surface analysis. Field emission scanning electron microscopy (SU-8010, Hitachi) was used to investigate the morphology of the electrodes. The chemical composition of the surface layer on the LiMn₂O₄ was analyzed by X-ray photoelectron spectroscopy (ESCALAB 250, Thermo Scientific) with Al K α line as an X-ray source. The binding energies were calibrated using containment carbon (C 1s = 284.6 eV).

The cells were disassembled in an argon-filled glove box to recover the electrolyte. The electrolytes before/after the overcharge test were characterized by ¹H nuclear magnetic resonance spectroscopy (NMR, AVANCE III, Bruker Bopspin) using CD₃CN as a solvent.

Results and discussion

Flammability and conductivity

In this study, the self-extinguishing time (SET) was used to evaluate the flammability of the electrolytes. Fig. 1(a) shows the results of the SET tests for solutions containing different concentrations of 4-MPPFP in the electrolyte. The burning time of the electrolyte decreases with increasing 4-MPPFP concentration. The SET value of the base electrolyte is 15.08 s cm^{–1}, and only 10.7 s cm^{–1} for the electrolyte containing 8 vol% 4-MPPFP. Although the SET value of the electrolyte containing 5 vol% 4-MPPFP is 13.7 s cm^{–1}, the flame propagation time increases from 0.62 s cm^{–1} to 9.23 s cm^{–1} (Fig. 1(b)). The flammability of the electrolytes with and without the 4-MPPFP additive can be visualized in Fig. 2. Obviously, for the base electrolyte the fire spreads very quickly, while the electrolyte with 5 vol% 4-MPPFP additive postpones the spread of fire under these testing conditions. The flame-retardant performance of 4-MPPFP is attributed to the introduction of F, N, and P, and due to the radical adsorption and oxygen isolation mechanism.²⁴ Additionally, the influence of 4-MPPFP on the conductivity was also evaluated as shown in Fig. 1(c). The conductivity of the electrolyte decreases as the concentration of the additive increased due to its high viscosity and low dielectric constant.

Overcharge behavior of 4-MPPFP in the LiMn₂O₄/Li cells

Fig. 3(a) compares the LiMn₂O₄/Li cells during charging using the electrolyte without an additive, with 5 vol% 4-MPPFP, and with 5 vol% PFPN. All of the cells were charged to 4.3 V at 0.5C, and then overcharged to 5.0 V at 1C. There is not much difference between these cells during normal charging. However, the



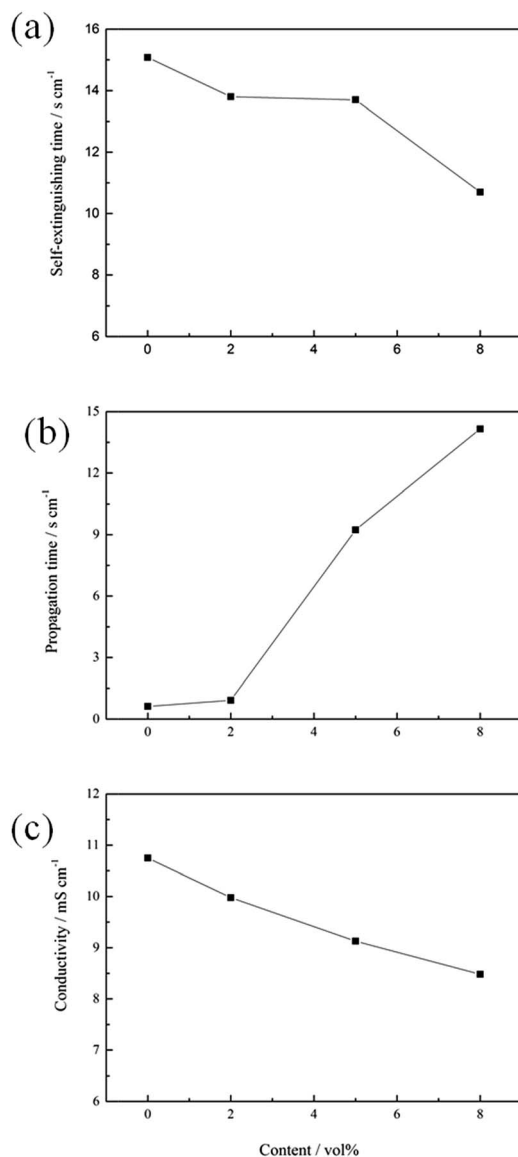


Fig. 1 The self-extinguishing time (a), flame propagation time (b) and conductivity (c) of the electrolytes containing different concentrations of 4-MPPFPP.

voltage at the overcharged stage rises up sharply to 5 V for the cells without an additive and with 5 vol% PFPN, while a charging voltage plateau of 4.45–5.0 V was observed for the cell with 4-MPPFPP. This suggests that the incorporation of 5 vol% 4-MPPFPP can dramatically postpone the growth of potential during the overcharge process. At the same time, we believe that the structures of A and C can't protect the LiMn₂O₄/Li cells from voltage runaway at overcharge.

Fig. 3(b) shows the difference in the differential capacity plots of LiMn₂O₄/Li cells with and without an additive. We find there are two oxidation peaks with the first sharp peak around 4.45 V and the second broad peak ranging from 4.6 to 5 V due to the addition of 4-MPPFPP. It is known that most anisoles act as redox shuttle agents.^{25,26} So, we believe that the first sharp peak is due to the oxidation of 4-MPPFPP as a redox shuttle reaction

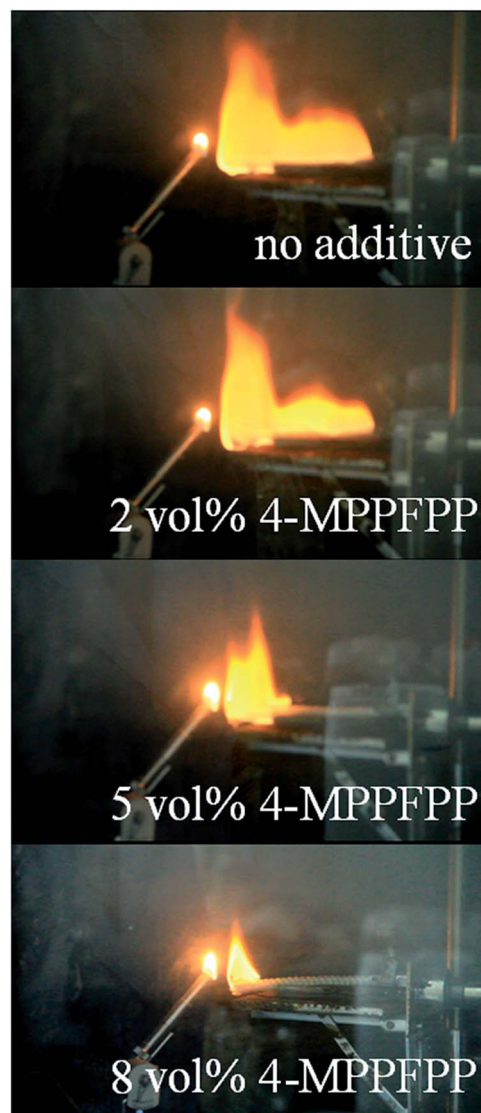


Fig. 2 The flammability of the electrolytes with and without 4-MPPFPP.

and the second broad peak may originate from the decomposition of 4-MPPFPP to achieve electrochemical polymerization.

Fig. 4 presents the results of the LiMn₂O₄/Li cells with and without 5 vol% 4-MPPFPP. The typical EIS of the Li-ion cells is composed of two partially overlapped semicircles and a straight slopping line at the low frequency end.²⁷ R_f is attributed to the solid-state interphase layer (SEI) formed on the surface of the electrodes, and R_{ct} is the faradic charge-transfer resistance. Before overcharge, the impedances are similar for both electrodes with and without 4-MPPFPP. The LiMn₂O₄/Li cell without an additive showed an interfacial impedance of about 18.76 Ω and a charge-transfer impedance of about 26.36 Ω . In contrast, the cell with 5 vol% 4-MPPFPP showed an interfacial impedance of about 17.94 Ω and a charge-transfer impedance of about 25.26 Ω . For the cell with 4-MPPFPP, the interfacial resistance after overcharge (94.76 Ω) is slightly bigger than that of the cell without the additive (24.49 Ω) indicating a thick interface



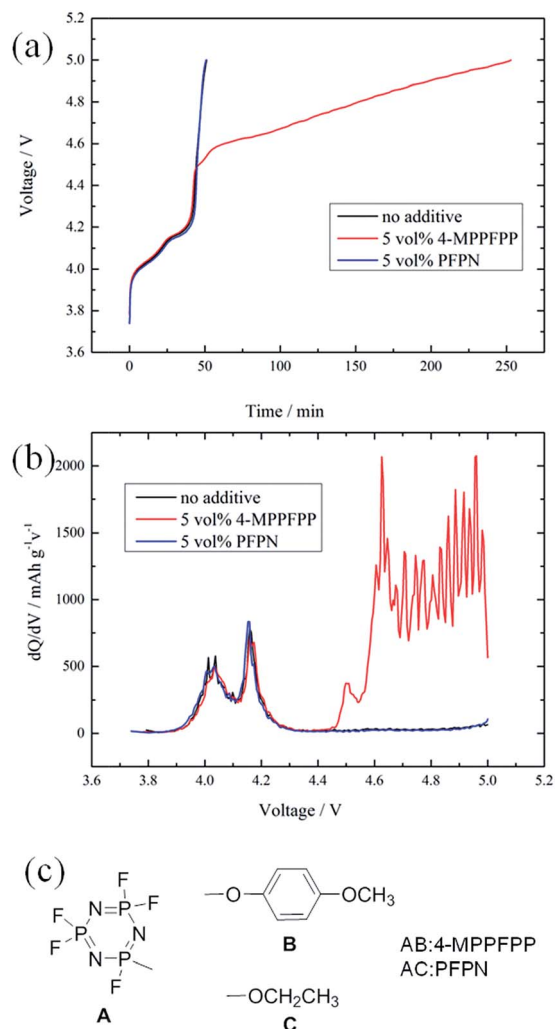


Fig. 3 V-t profiles (a) and the differential capacity plots (b) of LiMn₂O₄/Li cells with the electrolyte without an additive, with 5 vol% 4-MPPFPP, and with 5 vol% PFPN. The structures of 4-MPPFPP and PFPN (c).

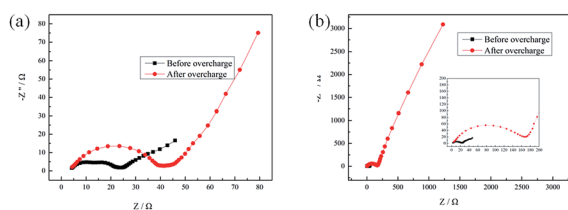


Fig. 4 Electrochemical impedance spectra of LiMn₂O₄/Li cells (a) without and (b) with 5 vol% 4-MPPFPP before and after the overcharge test.

between the electrode and the electrolyte. The charge transfer resistance rises up to about 118.7 Ω , which is higher than the value of the cell without the additive (37.08 Ω). The sharp increase in the charge transfer resistance may be due to the thick film formed on the cathode by 4-MPPFPP, which is nonconductive and inhibits the charge transfer process.

The cells were disassembled in a glove box under Ar atmosphere to analyze the effect of 4-MPPFPP on the LiMn₂O₄ surface after overcharge. SEM images show that the electrode exhibits a porous structure (Fig. 5(a)). No distinct surface layer was observed. The bulk structure of primary particles of the LiMn₂O₄ electrode was maintained well after the overcharge test (Fig. 5(c)). However, after being overcharged to 5 V in the electrolyte with 5 vol% 4-MPPFPP, the electrode surface was fully covered with a thick film (Fig. 5(b)). The particle edge of the electrode with the additive electrolyte becomes indistinct (Fig. 5(d)).

The surfaces of the LiMn₂O₄ electrodes after being overcharged with different electrolytes were analyzed by XPS. The Mn 2p, O 1s, and N 1s spectra of the LiMn₂O₄ electrodes are shown in Fig. 6. In the Mn 2p spectra, there are main peaks dominated by Mn 2p_{3/2} (642.9 eV (ref. 28) and 641.7 eV (ref. 28)) and Mn 2p_{1/2} (653.6 eV (ref. 28)). The intensity of the peaks for the surface of the LiMn₂O₄ electrode without the additive is much stronger than that with the additive. In the O 1s spectra, the intensity of the Mn-O peak (529.6 eV (ref. 29)) observed on the surface of the LiMn₂O₄ electrode without the additive is stronger than that with 5 vol% 4-MPPFPP. This is due to a thick film which comes from the 4-MPPFPP on the electrode surface. In the N 1s spectra, the surface of the LiMn₂O₄ is covered with some nitrogen compounds. Based on the XPS results, we have assumed that the 4-MPPFPP additive helps form the thick film.

Fig. 7 shows the ¹H NMR data for the electrolytes with 5 vol% 4-MPPFPP before/after the overcharge test. We can find that there are two groups of chemical shifts attributed to 4-MPPFPP in Fig. 7. One is located at 7.12 ppm (H_A) and has two splitting peaks. The other is located at 6.89 ppm (H_B) and also has two splitting peaks. The intensities of the 7.12 ppm and 6.89 ppm peaks are smaller after being overcharged, as shown in Fig. 7(b), and the ratio was approximately 1 : 1. For redox additives, the additive concentration should be maintained after the cell is overcharged. We assume that this decomposition reaction could be strongly related to the polymerization reaction of 4-MPPFPP. The possible reaction paths are presented in Fig. 8.^{30,31}

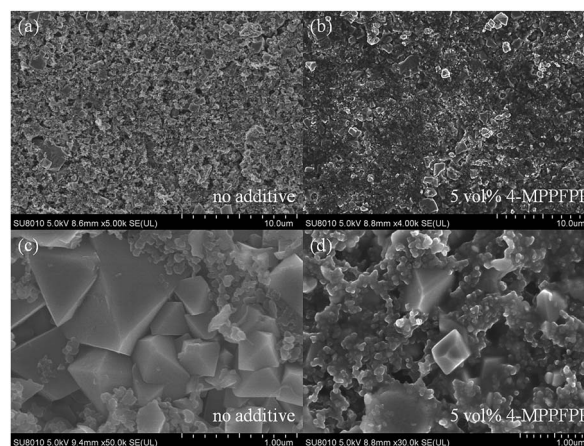


Fig. 5 SEM images of the LiMn₂O₄ electrode taken from the overcharged cells (a and c) without and (b and d) with 5 vol% 4-MPPFPP electrolyte.



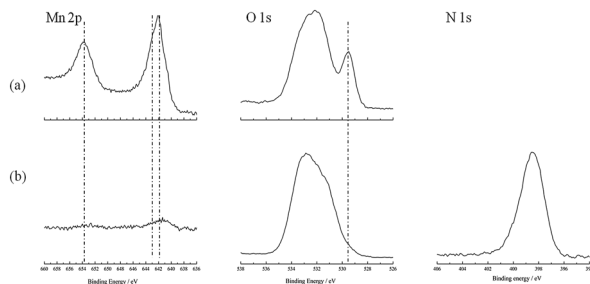


Fig. 6 The Mn 2p, O 1s, and N 1s spectra of the LiMn_2O_4 electrodes taken from the overcharged cells without (a) and with (b) 5 vol% 4-MPPFPP.

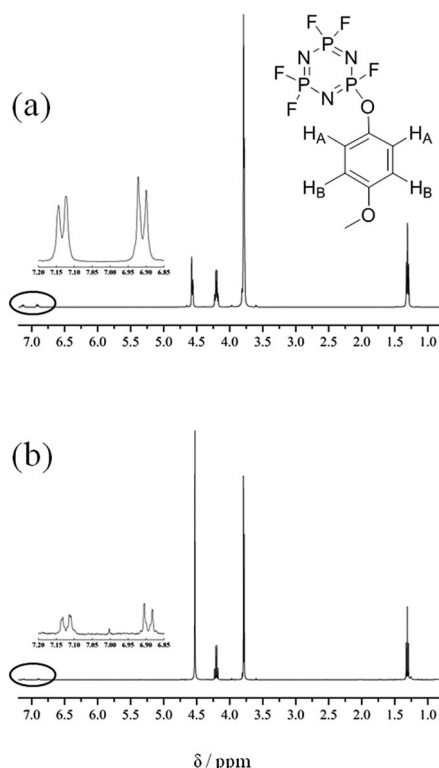


Fig. 7 ^1H NMR data for the electrolyte with 5 vol% 4-MPPFPP before overcharge (a) and the electrolyte taken from the cell with 5 vol% 4-MPPFPP after the overcharge test (b).

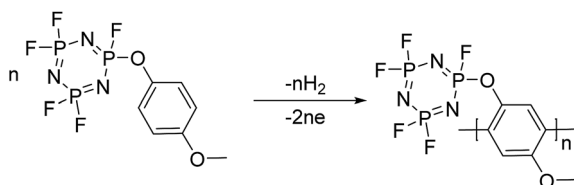


Fig. 8 The possible reaction paths of 4-MPPFPP on the cathode.

Cycling performance of 4-MPPFPP in the $\text{LiMn}_2\text{O}_4/\text{Li}$ and MCMB/Li cells

To evaluate the influence of 4-MPPFPP on the electrochemical performance of the $\text{LiMn}_2\text{O}_4/\text{Li}$ cells and MCMB/Li cells, cycling

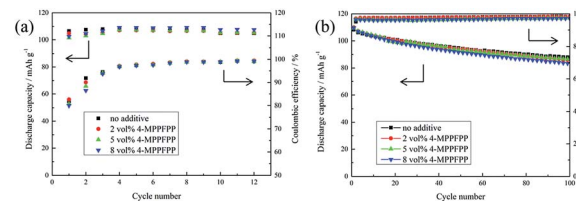


Fig. 9 Formation cycles of the $\text{LiMn}_2\text{O}_4/\text{Li}$ cells at 25 °C (a) and the cycling performance of the $\text{LiMn}_2\text{O}_4/\text{Li}$ cells at 55 °C (b).

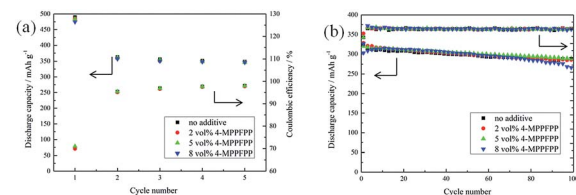


Fig. 10 Formation cycles of the MCMB/Li cells at 25 °C (a) and the cycling performance of the MCMB/Li cells at 25 °C (b).

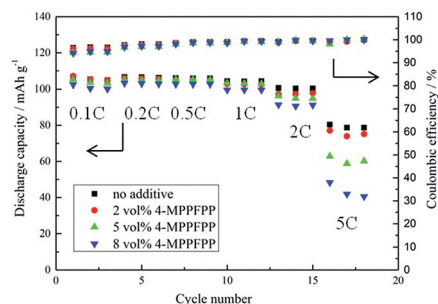


Fig. 11 Rate performance of the $\text{LiMn}_2\text{O}_4/\text{Li}$ cells at 25 °C.

tests were carried out. Fig. 9(a) shows that the discharge capacity and coulombic efficiency of the cells exhibited no obvious differences during the formation cycles. In Fig. 9(b), it can be found that the $\text{LiMn}_2\text{O}_4/\text{Li}$ cell without 4-MPPFPP keeps 81.2% of its initial discharge capacity (after 100 cycles at 55 °C), while the $\text{LiMn}_2\text{O}_4/\text{Li}$ cell with 5 vol% 4-MPPFPP keeps 79.2% (after 100 cycles at 55 °C). The $\text{LiMn}_2\text{O}_4/\text{Li}$ cell with 8 vol% 4-MPPFPP retains 76.1% of its initial discharge capacity after 100 cycles at 55 °C. This indicates that a higher ratio of 4-MPPFPP might cause a lower cycling capability and coulombic efficiency. The capacity only slightly decreases due probably to the dilution of LiPF_6 concentrations and trace impurities.

Fig. 10 shows the formation cycles and cycling performance of the MCMB/Li cells at 25 °C with different electrolytes. These results confirm that 5 vol% 4-MPPFPP has no negative effect on the MCMB/Li cells.

Fig. 11 presents the rate performance of the $\text{LiMn}_2\text{O}_4/\text{Li}$ cells at 25 °C. The cells with the 4-MPPFPP electrolyte additive generally exhibited only slightly lower discharge capacities than that of the electrolyte without the additive at lower charge/discharge rates.



Conclusions

In summary, 4-MPPFPP was synthesized as a bifunctional additive for a carbonate-based electrolyte for lithium ion batteries. The addition of 4-MPPFPP can shorten the flame extinguishing time of the electrolyte. Besides, 4-MPPFPP can also be polymerized to form a thick polymer film, which can protect the batteries from voltage runaway at overcharge. The influence of 5 vol% 4-MPPFPP on the cycling performance and capacity is negligible.

Conflicts of interest

There are no conflicts to declare.

Acknowledgements

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