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# Synthesis and characterization of metal–organic frameworks fabricated by microwave-assisted ball milling for adsorptive removal of Congo red from aqueous solutions

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In this study, four metal–organic frameworks (MOFs) were prepared using a simple, low-cost, and high-efficiency technique utilizing simple carboxylic acids (1,3,5-benzenetricarboxylic acid and terephthalic acid) and metal salts (copper(II) acetate tetrahydrate and cobalt(II) acetate tetrahydrate) and microwave-assisted ball milling. The MOFs were characterized by infrared spectrometry, X-ray diffraction, scanning electron microscopy, and thermogravimetry. They were investigated for use as adsorbents for the adsorption of Congo red (CR) from aqueous solutions. The results showed that the adsorption capacity of MOF-2—synthesized from trimesic acid and cobalt(II) acetate tetrahydrate—was the highest, reaching 85.54% after 300 min. Electrostatic and  $\pi$ – $\pi$  stacking interactions are thought to play an important role in the adsorption of CR onto the MOFs. This MOF material reported in this work demonstrates a superior dye adsorption capacity when compared to other adsorbents.

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## 1. Introduction

Porous materials have come to constitute one of the studied focus areas of chemists, materials scientists, and physicists. Metal–organic frameworks (MOFs), a new class of porous materials, have quickly progressed into a fruitful research field.<sup>1–3</sup> In the past twenty years, MOFs have garnered considerable interest as a novel type of functional porous material and have been tested for applications involving energy storage,<sup>4,5</sup> adsorption and separation,<sup>6,7</sup> catalysis,<sup>8,9</sup> drug delivery,<sup>10,11</sup> CO<sub>2</sub> capture,<sup>12,13</sup> chemical sensors,<sup>14</sup> and others.<sup>15,16</sup> Since their first report, many studies have described the use of MOFs for filtration of wastewater. For instance, in hexavalent chromium reduction, various MOFs including MIL-125(Ti),<sup>17</sup> MIL-53(Fe),<sup>18</sup> and Zn(Ce)-MOFs<sup>19</sup> have been used. MOFs, namely MOF-235<sup>20</sup> and g-C<sub>3</sub>N<sub>4</sub>/MIL-125(Ti),<sup>21</sup> have also been utilized for the degradation of organic dyes in wastewater. MOFs are usually synthesized by solvothermal, microwave-assisted, and mechanochemical processes.

Mechanochemical processes have been long established and their usefulness stems from the capacity to rapidly and quantitatively promote solid-phase reactions using only ball milling or liquid-assisted grinding. In traditional chemical

synthesis, the solvent often plays a key role in solvation, transportation, and energy dispersion of chemicals. Energy and mass transport may be reduced in solvent-free reactions. The efficient mixing process under ball milling can overcome this drawback, enabling reactions between solids to take place even under solvent-free conditions.<sup>22–24</sup> In mechanochemical procedures, the energy needed for the activation of chemical reactions is usually provided by mechanical force, similarly to how it is provided by light, heat, and electrical potential in photochemistry, thermochemistry, and electrochemistry, respectively. Developments concerning mechanochemical procedures have mostly taken place in the last two decades, and usually focused on their application in synthetic organic chemistry. Recent research efforts have involved exploring the concept of mechanochemistry to obtain advanced materials with a wide range of applications and uses.<sup>25–28</sup> In addition to classical organic synthesis, therefore, mechanochemistry has been applied to the synthesis of MOFs,<sup>29–31</sup> metal complexes,<sup>25,26</sup> and hydrogen storage materials.<sup>32</sup>

Organic dyes are pollutants, and can include chemicals that are toxic or carcinogenic to mammals and other living organisms.<sup>33,34</sup> Dye molecules are often stable and resistant to biodegradation under normal conditions as a result of their complex molecular structures. Therefore, removal of dyes from wastewater by biological or chemical degradation methods can be very challenging. Nevertheless, several groups have explored the possibility of applying polymeric and MOF materials for the removal or degradation of various dye

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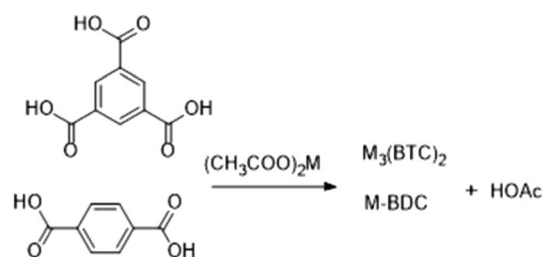
species. For example, Wang *et al.* have reported four cobalt(II) coordination polymers for the degradative removal of an azo dye Congo red (CR).<sup>35</sup> Degradation efficiency of complex 1 or 2 on Congo red ended up to 89 or 98% after 130 min, respectively. Complex 3 or 4 was introduced into the system, only about 56 or 52% of the dye was decolorized, respectively. Hu *et al.* have synthesized a MOF for use in the degradative removal of methyl orange.<sup>36</sup> Degradation efficiency can reach 76.1% for 1 within 30 min and reaching up to 92.8% for 1 in 120 min; Cui *et al.* have prepared two copper(I) cyanide coordination polymers and utilized them for the degradation and removal of methylene blue.<sup>37</sup> Photodegradation efficiency (using CP 1 is 90.8% and using CP 2 is 87.2%) for the degradation of methylene blue (MB) under UV light irradiation. Masoomi *et al.* have synthesized TMU-5 and TMU-6 materials for the degradation and removal of RhB.<sup>38</sup> Zhao *et al.* have reported the synthesis of Ni-MOF/GO, which was applied to the degradation and removal of CR.<sup>39</sup> The adsorption capacity of graphene oxide/metal-organic frameworks (GO/MOFs) for CR reached 2489 mg g<sup>-1</sup>. Racles *et al.* have synthesized a siloxane-based MOF and applied it to the degradative removal of CR<sup>40</sup> *etc.*

Microwave-assisted ball milling—first reported by our laboratory—is based on a solid-liquid ball milling approach where a ball milling machine is placed in a microwave oven.<sup>41–43</sup> This method improves the speed of chemical reactions and alter their routes.<sup>44–47</sup> Herein, four MOFs were obtained by microwave-assisted ball milling from trimesic acid, terephthalic acid, Cu(CH<sub>3</sub>COO)<sub>2</sub>·H<sub>2</sub>O, and Co(CH<sub>3</sub>COO)<sub>2</sub>·4H<sub>2</sub>O as the starting materials, and applied successfully for the reduction of aqueous CR.

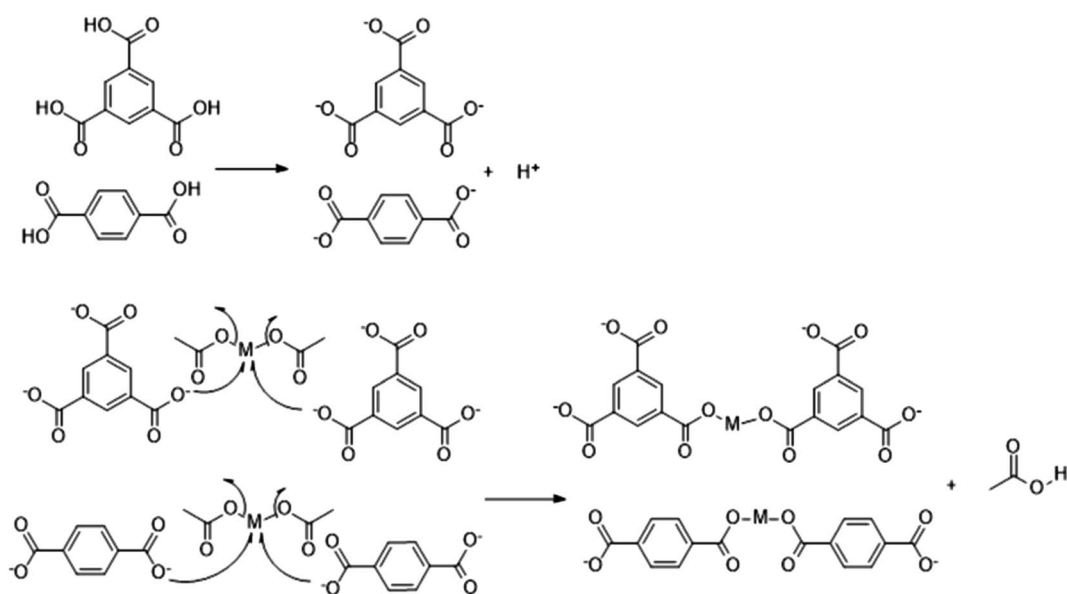
## 2. Results and discussion

### Preparation of MOFs

Crystals are most often isolated under microwave-assisted heating because of the fast kinetics of crystal growth and nucleation. Therefore, the greatest challenge with microwave heating is restricting the crystallite size. Crystals of large sizes are usually more difficult to acquire under microwave heating than in typical solvothermal processes. In this work, acid is produced using a post-treatment process according to the following reaction:



Under the combined action of microwave irradiation and ball milling that provides mechanical energy, trimesic acid is easily deionized to form a  $\pi_3^-$  negative ion in aqueous solutions. Cu and Co could each lose two electrons from their s orbital, leading to an empty orbital in each case. The trimesic acid anion can provide electrons to the metal ions and release acetic acid molecules, which react with the metal ions in a proportion of 1 : 1.5 or 1 : 1. The specific reaction mechanism has been previously reported.<sup>48</sup> The reaction mechanism for the reaction of a metal salt with H<sub>3</sub>BTC or H<sub>2</sub>BDC can be described as follows:



Preparation of the MOFs was fast and achieved without the need for an organic solvent, unlike the previously reported preparation of MOFs shown in Table 1.

### Structure characterization

**Fourier-transform infrared spectroscopy.** Fourier-transform infrared (FTIR) spectra of the MOFs were acquired using KBr pellets in the 500–3000  $\text{cm}^{-1}$  region and are shown in Fig. 1. The spectra displayed the peaks characteristic for the main functional groups of MOFs-1–4 at 1558  $\text{cm}^{-1}$ , 1400  $\text{cm}^{-1}$ , 1570  $\text{cm}^{-1}$ , 1373  $\text{cm}^{-1}$ ; 1558  $\text{cm}^{-1}$ , 1398  $\text{cm}^{-1}$ ; and 1558  $\text{cm}^{-1}$ , 1361  $\text{cm}^{-1}$ , respectively. The prepared materials displayed no strong carbonyl peak at  $\sim 1710 \text{ cm}^{-1}$ . This observation was mainly attributed to the fact that the carboxyl anions can form  $\pi_3^5$  after the formation of the carboxylate, and the two resulting oxygen atoms on the carboxyl group tend to be equivalent, thus leading to an delocalized electronic cloud and appearance of characteristic carboxylate peaks in the ranges of 1610–1550  $\text{cm}^{-1}$  and 1420–1300  $\text{cm}^{-1}$ .<sup>49</sup> Thus, these results confirmed that the metal salts and organic acids formed the target coordination compounds.

**X-ray diffraction.** The synthesized MOF structures were elucidated by powder X-ray diffraction (XRD). The main peak observed in the XRD pattern of MOF-1 matched well with literature data.<sup>50</sup> The XRD patterns for the various MOFs are shown in Fig. 2. It can be seen that MOFs-1–4 exhibited main diffraction peaks at  $2\theta$  values of 20.2°, 11.6°, 9.5°, 29.2°, 27.8°, 8.7°, 27.8°, 18.9°, 11.7°; and 27.9°, 14.3°, 9.7°, respectively. The patterns showed fairly high intensities, indicating the high crystallinity of the samples.

**Scanning electron microscopy.** The particle size and morphology of each MOF was observed by scanning electron microscopy (SEM), as shown in Fig. 3. The SEM images demonstrated the presence of MOF crystals with high dispersity and unusual particle morphology. The morphology of MOF-1 was quite similar to that reported previously.<sup>50,51</sup> Currently,

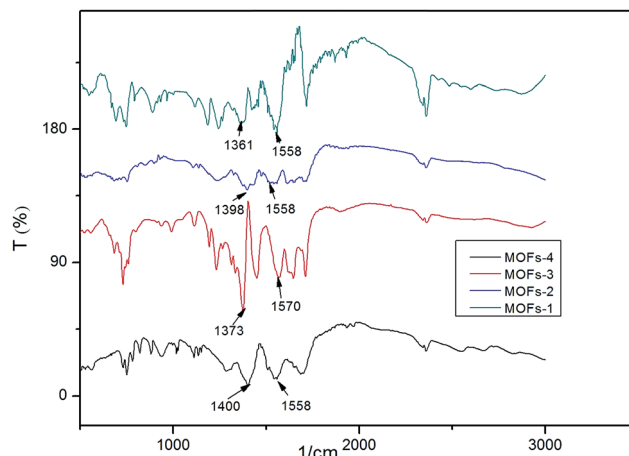


Fig. 1 FTIR spectra for MOFs.

the molecular interactions of the organic ligands will be weak or even absent, and the deprotonation of the organic ligands can be enhanced in order to promote the growth of crystals in aqueous solution.

**Thermogravimetric analysis.** Thermogravimetric analysis (TGA) was performed to determine the composition of this membrane (Fig. 4), demonstrating that metal salts of the same type exhibited similar thermal decomposition temperatures in Ar atmosphere. The TG curves showed three different areas for each MOF. For MOF-1: the first mass loss (8.3%) area between 106 °C and 189 °C was attributed to the evaporation of water molecules from the surface of the sample; the area between 189 °C and 310 °C, exhibiting slow mass loss of 7.6%, was related to the loss of water from the MOF together with oxidation<sup>48</sup> of  $\text{Cu}^{2+}$ ; and the third region starting at 310 °C was attributed to the collapse of the MOF and fracturing of the organic linkers, thus indicated that the MOF is thermally stable up to 310 °C. The loss of mass was completed by 310 °C and the

Table 1 Comparison of different procedures for preparation of MOFs

MOF	Reagent	Solvent	Temperature (°C)	Time (h)	BET ( $\text{m}^2 \text{g}^{-1}$ )	Ref.
MOFs-1	$\text{H}_3\text{BTC}$ , $\text{Cu}(\text{OAc})_2 \cdot \text{H}_2\text{O}$	$\text{H}_2\text{O}$	Room temperature	0.5	32.0708	This work
MOFs-2	$\text{H}_3\text{BTC}$ , $\text{Co}(\text{OAc})_2 \cdot 4\text{H}_2\text{O}$				7.7682	
MOFs-3	$\text{H}_2\text{BDC}$ , $\text{Cu}(\text{OAc})_2 \cdot \text{H}_2\text{O}$				82.991	
MOFs-4	$\text{H}_2\text{BDC}$ , $\text{Co}(\text{OAc})_2 \cdot 4\text{H}_2\text{O}$				0.9113	
MIL-125(Ti)	$\text{H}_2\text{BDC}$ , $\text{Ti}(\text{OC}_4\text{H}_9)_4$	$\text{CH}_3\text{OH}$ , DMF	150	48	1548.3	17
$\text{NH}_2\text{-MIL-125(Ti)}$	$\text{NH}_2\text{-BDC}$ , $\text{Ti}(\text{OC}_4\text{H}_9)_4$	$\text{CH}_3\text{OH}$ , DMF	150	48	1343.9	17
MIL-53(Fe)	$\text{H}_2\text{BDC}$ , $\text{FeCl}_3 \cdot 6\text{H}_2\text{O}$	DMF	150	15	—	18
HUT-091	$\text{H}_3\text{BTC}$ , $\text{Ce}(\text{NO}_3)_3 \cdot 6\text{H}_2\text{O}$		170	240	—	19
MIL-68(In)	$\text{H}_2\text{BDC}$ , $\text{In}(\text{NO}_3)_3 \cdot x\text{H}_2\text{O}$	DMF	100	48	681.4	55
$[\text{Co}(\text{L1})(\text{bpdc})]_n$	1,4-Bis(5,6-dimethylbenzimidazol-1-ylmethyl)benzene, $\text{H}_2\text{bpdc}$ , $\text{CoCl}_2 \cdot 6\text{H}_2\text{O}$	$\text{H}_2\text{O}$	140	36	0.912	35
$\{[\text{Co}(\text{L1})(\text{nph})] \cdot 0.5\text{H}_2\text{O}\}_n$	1,4-Bis(5,6-dimethylbenzimidazol-1-ylmethyl)benzene, $\text{H}_2\text{nph}$ , $\text{CoCl}_2 \cdot 6\text{H}_2\text{O}$				0.783	
$[\text{Co}(\text{L2})(\text{bpdc})]_n$	1,3-Bis(5,6-dimethylbenzimidazol-1-ylmethyl)benzene, $\text{H}_2\text{bpdc}$ , $\text{CoCl}_2 \cdot 6\text{H}_2\text{O}$				0.852	
$[\text{Co}(\text{L3})(\text{bpdc})(\text{H}_2\text{O})]_n$	1,1'-Bis(5,6-dimethylbenzimidazole)methane, $\text{H}_2\text{bpdc}$ , $\text{CoCl}_2 \cdot 6\text{H}_2\text{O}$				0.813	



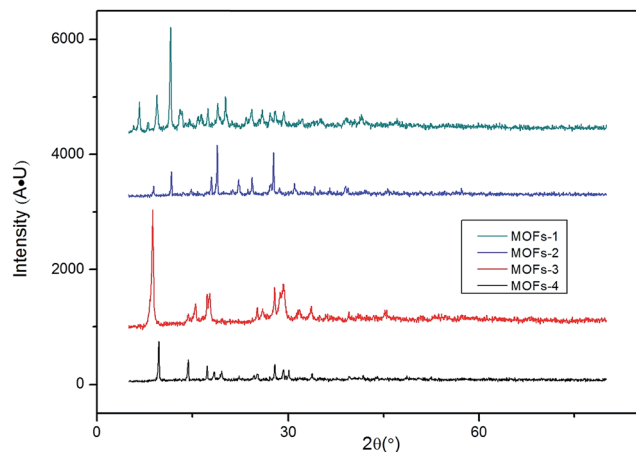


Fig. 2 XRD patterns for MOFs.

mass subsequently remained stable up to 800 °C. The residue measured at this temperature represented 35.03% of the original mass.

For MOF-2: the first mass loss (16.5%) area between 100 °C and 223 °C indicated the loss of moisture; the area between 223 °C and 366 °C with a slow loss of 3.3% mass was related to the oxidation of  $\text{Co}^{2+}$ ; the third region starting at 366 °C represented again the temperature at which the structure of the MOF collapsed and the organic linker was fractured. The loss of mass was deemed as complete at 366 °C, after which the MOF remained stable up to 800 °C. The residue measured at this temperature was 33.17% of the original mass.

For MOF-3: the first mass loss (7.63%) area between 74 °C and 330 °C indicated the loss of moisture and oxidation<sup>52</sup> of  $\text{Cu}^{2+}$ ; the third region started at 330 °C, a temperature at which the structure of the MOF collapsed and the organic linker was fractured. This loss of mass was completed by 330 °C and the MOF was stable up to 800 °C. The final residue was 43.5% of the original mass.

For MOF-4: the first mass loss (7%) area between 152 °C and 229 °C indicated the loss of moisture; the area between 229 °C and 346 °C with a slow mass loss of 5.6% was related to the loss of water from the MOF together with the oxidation of  $\text{Co}^{2+}$ ; the third region started at 346 °C, a temperature at which the structure of the MOF collapsed and the organic linker was fractured. This loss of mass was completed by 346 °C and the MOF remained stable up to 800 °C. The final residue represented 29.7% of the original mass.

### Removal of Congo red

The specific surface areas of MOFs-1–4 were determined to be  $32.0708 \text{ m}^2 \text{ g}^{-1}$ ,  $7.7682 \text{ m}^2 \text{ g}^{-1}$ ,  $82.9910 \text{ m}^2 \text{ g}^{-1}$ , and  $0.9113 \text{ m}^2 \text{ g}^{-1}$ , respectively, as shown Table 1, using the nitrogen adsorption isotherms. The adsorption performance of the four MOFs toward CR was studied. Fig. 5 shows the adsorbed quantities ( $q_t$ ) of CR on the MOFs as a function of time  $t$ , as determined using the following equations:

$$q_t = (C_0 - C_t)V/m \quad (1)$$

$$q_e = (C_0 - C_e)V/m \quad (2)$$

where  $q_t$  and  $q_e$  are the amounts ( $\text{mg g}^{-1}$ ) of dye adsorbed on the adsorbents at time  $t$  and equilibrium, respectively;  $C_0$ ,  $C_t$ , and  $C_e$  are the liquid-phase concentrations ( $\text{mg L}^{-1}$ ) of CR initially, at time  $t$ , and at equilibrium, respectively; and  $m$  (g) and  $V$  (L) represent the quantity of the adsorbent and the volume (L) of CR solution, respectively.

As seen in Fig. 5, initially, the adsorption capacities of the MOFs for CR increased rapidly and then gradually decreased until the adsorption equilibria were attained upon prolonged contact. The  $q_e$  values of the four MOFs for CR were quite different. Specifically, these values were 129.95, 120.15, 77.05, and  $64.56 \text{ mg g}^{-1}$  for MOF-2, MOF-1, MOF-3, and MOF-4, respectively. Clearly, the different results were arising due to the introduction of different organic ligands and metal ions. The electron energy arrangements of Cu and Co are  $3d^9 4s^2$  and  $3d^7 4s^2$ , respectively. For these metals, the loss of two electrons from their s orbital gives rise to an empty orbital in each case. Different metal ions and different organic ligands have different spatial structures, resulting in different adsorption results.

It is possible that the original concentrations of CR were sufficiently high to provide the driving force to surpass the resistance of mass transfer between the liquidoid and the solidoid.<sup>53</sup>

The adsorption isotherms were acquired after adsorption for 5 h. The adsorption capacities of the MOFs under investigation in the removal of CR were in the following order: MOF-4 > MOF-3 > MOF-2 > MOF-1 (as shown in Fig. 6). The maximum adsorption of CR was 85.45% after 300 min, determined for MOF-2 in natural light.

### Kinetics for the adsorption of Congo red by MOFs

The adsorption capacity was obtained to study the kinetics of adsorption of CR by the prepared MOFs. The adsorption model that describes the adsorption of a solute onto a solid surface can be expressed as:

$$\frac{dq}{dt} = k_1(q_e - q_t)$$

where  $k_1$  is the apparent pseudo-first-order rate constant ( $\text{min}^{-1}$ ),  $q_t$  is the adsorption capacity at time  $t$  ( $\text{mg g}^{-1}$ ), and  $q_e$  is the adsorption capacity at equilibrium ( $\text{mg g}^{-1}$ ).

The applicability of the pseudo-first-order kinetic model to the adsorption of CR by MOFs was tested by fitting of the experimental data to the model *via* the application of the least squares regression analysis (as shown in Fig. 8 and 9). The analysis revealed a ratio of 1 : 2 for the quantity of the MOF to the volume of the dye that afforded a suitable description of the adsorption process, as shown in Table 2. This result indicates that the investigated MOFs are excellent materials for the removal of CR. CR is a cationic dye that exists in the form of positively charged ions in aqueous solutions (as shown in Fig. 7). The CR adsorption may not only result from simple physical adsorption and chemical adsorption but also from the conjugation between CR and MOFs. CR contains  $\text{C}=\text{C}$  double bonds and  $\pi$  electrons, and these  $\pi$  electrons can easily interact





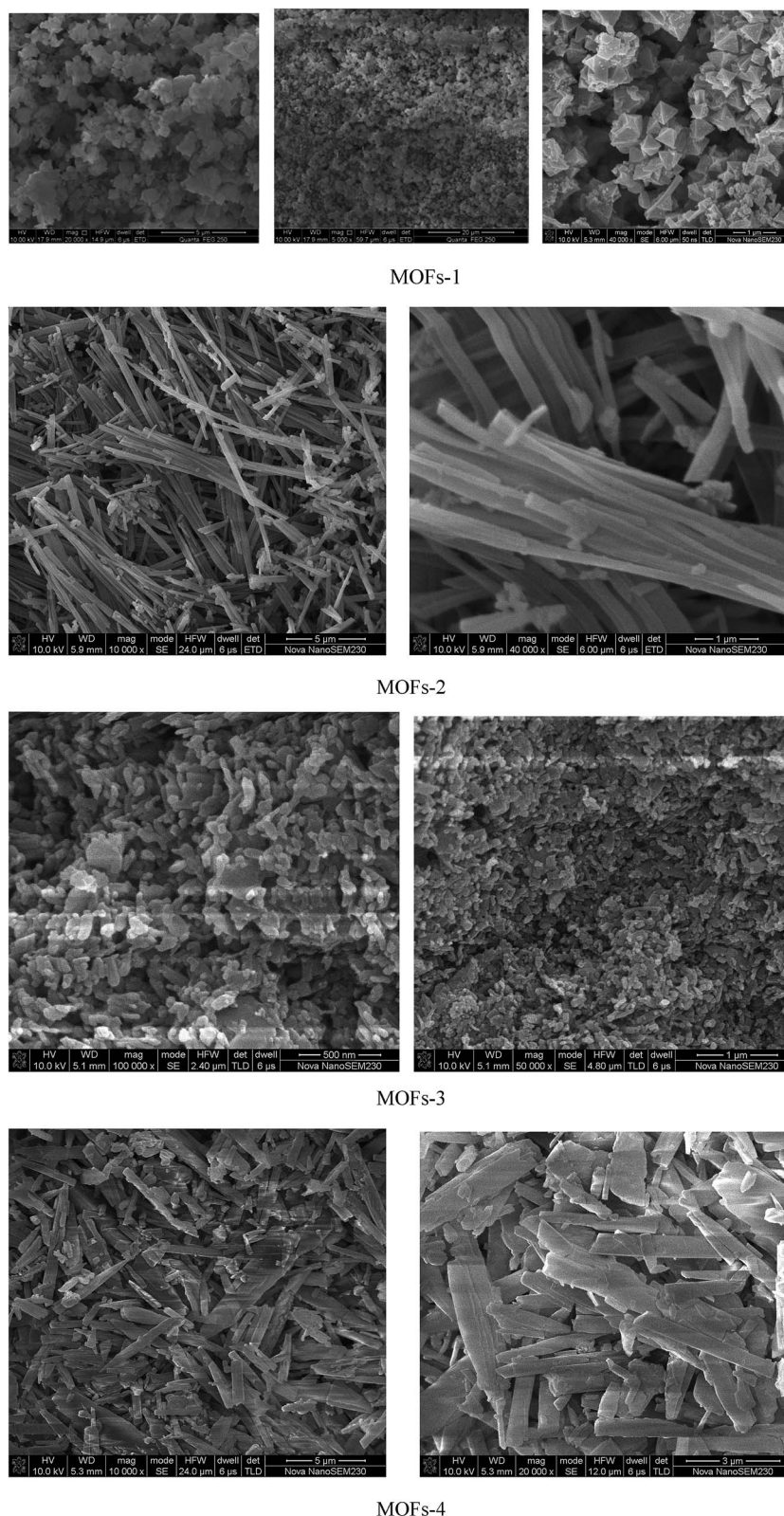


Fig. 3 SEM images of MOFs.

with the  $\pi$  electrons present on the benzene rings of the MOFs through  $\pi$ - $\pi$  interactions. CR could be adsorbed on the MOF surface with a face-to-face orientation by  $\pi$ - $\pi$  conjugation until an adsorption-desorption equilibrium is achieved. The

Zeta potential characterization showed that the adsorbents were charged, as shown in Table 3. Therefore, electrostatic forces are also beneficial for the adsorption of CR by the MOFs.<sup>54</sup>



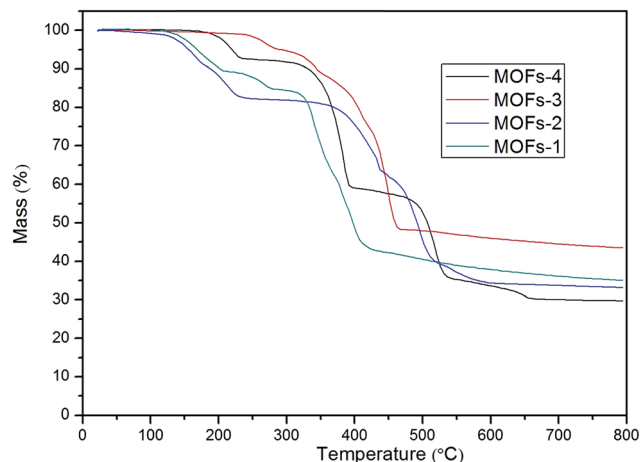


Fig. 4 TG curves for the four MOFs.

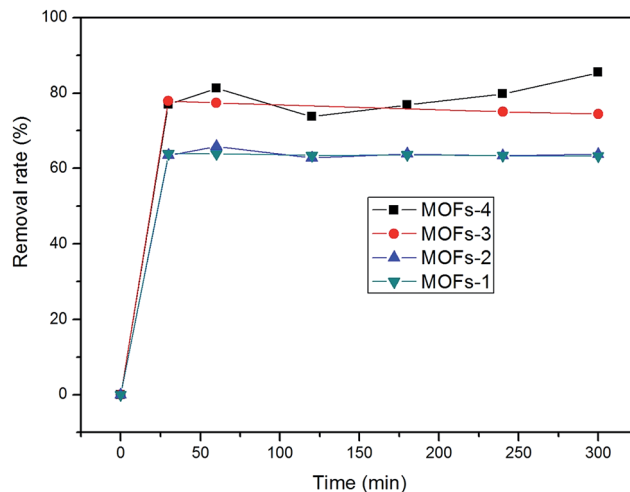


Fig. 6 Removal rate of MOFs.

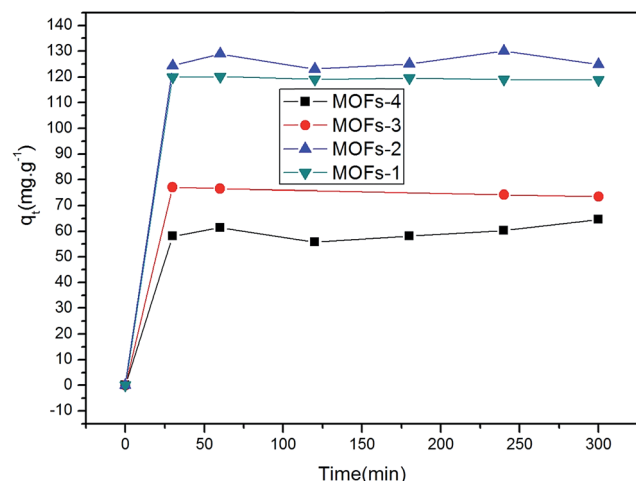
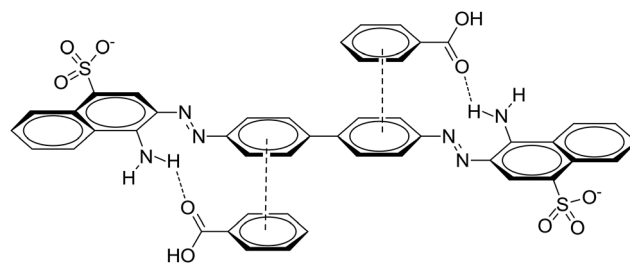
Fig. 5 Relationship between adsorbed quantities ( $q_t$ ) of CR and time.

Fig. 7 Binding mechanism between CR and MOFs.

### 3. Experimental

#### Materials

All materials were obtained from commercial suppliers and were used without further purification. 1,3,5-Benzenetricarboxylic acid (trimesic acid,  $C_6H_3O_6$ ,  $H_3BTC$ , 98%) and copper(II) acetate tetrahydrate ( $Cu(CH_3COO)_2 \cdot 4H_2O$ ,  $\geq 99.0\%$ ) were purchased from Aladdin Biological Technology Co., Ltd., Shanghai, China. Terephthalic acid ( $C_8H_6O_4$ ,  $H_2BDC$ , 99%) and cobalt(II) acetate tetrahydrate ( $Co(CH_3COO)_2 \cdot 4H_2O$ ,  $\geq 99.5\%$ ) were purchased from Sinopharm Chemical Reagent Co., Ltd., Shanghai, China. Congo red (CR) was purchased from Aladdin Biological Technology Co., Ltd., Shanghai, China. Trimesic acid and terephthalic acid were used as organic linkers. Copper(II) acetate tetrahydrate and cobalt(II) acetate tetrahydrate were used as the metal ion sources in the synthesis of MOFs.

#### Synthesis of MOFs

**Synthesis of MOF-1.** Trimesic acid (5.0158 g, 0.0238 mol), copper(II) acetate tetrahydrate (3.2242 g, 0.0159 mol), a stainless

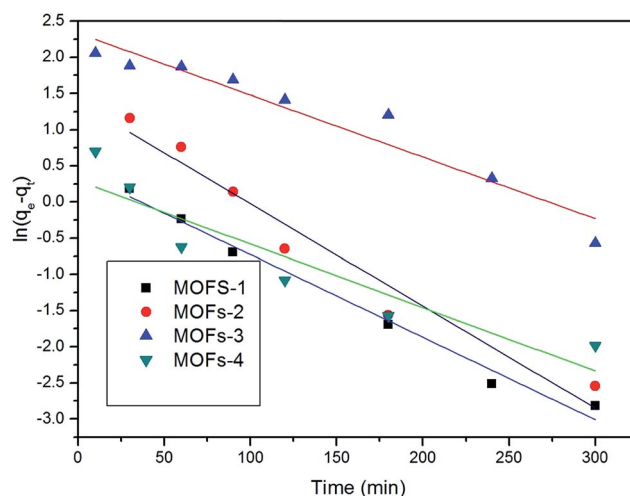


Fig. 8 Kinetic model of absorption when using 200 mg of MOFs.

steel ball (660 g), and 700 mL of deionized water were placed in a tetrafluoroethylene milling pot (diameter 140 mm, height 90 mm, volume 1.38 L). Stir milling and microwave irradiation were introduced simultaneously with a constant stirring speed of 200 rpm. After 40 min, the solution turned from colorless to blue and did not subsequently undergo any further changes in color. This change was accompanied by a significant evolution



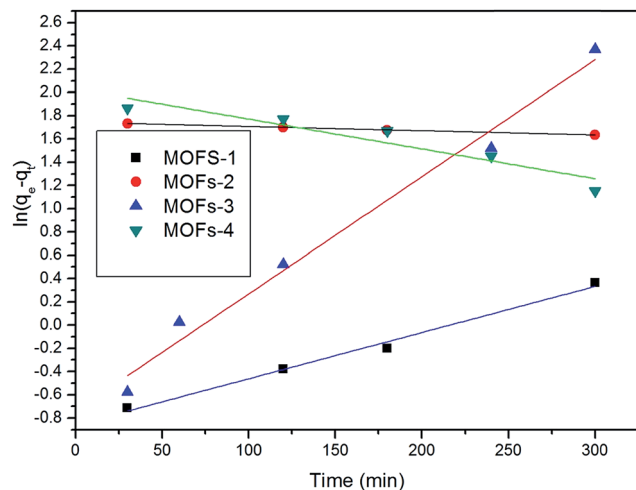


Fig. 9 Kinetic model of absorption when using 100 mg of MOFs.

Table 2 Calculated kinetic parameters for pseudo-first-order models for the adsorption of CR using MOFs as adsorbents

	Mass of MOFs (200 mg)		Mass of MOFs (100 mg)	
	<i>k</i>	<i>R</i> <sup>2</sup>	<i>k</i>	<i>R</i> <sup>2</sup>
MOFs-1	−0.01143	0.98408	0.00398	0.98814
MOFs-2	−0.01414	0.94257	−0.000364926	0.9986
MOFs-3	−0.00855	0.92131	0.01007	0.9803
MOFs-4	−0.00878	0.82756	−0.00257	0.86008

of sour gas. The compound was filtered and washed with water. The filtered solid was added to a beaker with ethanol and stirred for 3 h using a magnetic stirrer. The precipitate was filtered under suction and dried. The final product was collected for subsequent characterization and experiments. The MOF-1 was obtained as a blue solid in 63.3% yield (based on copper(II) acetate tetrahydrate and trimesic acid).

**Synthesis of MOF-2.** The MOF-2 was prepared using the same procedure as described for MOF-1, starting from trimesic acid (5.0076 g, 0.0238 mol), cobalt(II) acetate tetrahydrate (3.9533 g, 0.0159 mol), a stainless steel ball (450 g), and 700 mL of deionized water. After 40 min of reaction under the same conditions as stated above, the solution turned from colorless to damask, with no further changes in color observed subsequently. The color change was accompanied by a significant evolution of sour gas. The compound was filtered and washed with water. The filtered solid was added to a beaker with ethanol and stirred for 3 h using a magnetic stirrer. The precipitate was filtered under suction and dried. The final product was collected for subsequent characterization and absorption

experiments. The MOF-2 was obtained as a light red solid in 56.7% yield (based on cobalt(II) acetate tetrahydrate and trimesic acid).

**Synthesis of MOF-3.** This compound was synthesized using the same procedure as described for MOF-1 and MOF-2, commencing from terephthalic acid (5.0101 g, 0.0301 mol), copper(II) acetate tetrahydrate (7.3593 g, 0.0370 mol), a stainless steel ball (620 g), and 700 mL of deionized water. After 40 min of reaction, the solution turned from colorless to blue. This change was accompanied by a significant evolution of sour gas. The compound was filtered and washed with water. The filtered solid was added to a beaker with ethanol and stirred for 3 h using a magnetic stirrer. The precipitate was filtered under suction and dried. The final product was collected for subsequent characterization and experiments. The MOF-3 was obtained as a blue solid in 62.47% yield (based on copper(II) acetate tetrahydrate and terephthalic acid).

**Synthesis of MOF-4.** This compound was synthesized using the same procedure as stated above for MOFs-1–3, starting from terephthalic acid (5.0135 g, 0.0302 mol), cobalt(II) acetate tetrahydrate (7.6097 g, 0.0306 mol), a stainless steel ball (630 g), and 700 mL of deionized water. After 40 min of reaction, the solution turned from colorless to damask color and this change was accompanied by a significant evolution of sour gas. The compound was filtered and washed with water. The filtered solid was added to a beaker with ethanol and stirred for 3 h using a magnetic stirrer. The precipitate was filtered under suction and dried. The final product was collected for subsequent characterization and experiments. The MOF-4 was obtained as a light red solid in 44.7% yield (based on cobalt(II) acetate tetrahydrate and terephthalic acid).

### Characterization of MOFs

Structural and morphological characterizations were performed by XRD (D-5000, Siemens, Cu K $\alpha$ ), FTIR spectroscopy (IRTracer-100; Shimadzu), and field-emission SEM (FESEM, JSM-6700F, Japan). Thermogravimetry (TG) curves of samples were recorded by a NETZSCH STA 449C thermal analyzer (Germany) in Ar atmosphere in the range from 0 °C to 800 °C at a heating rate of 5 °C min<sup>−1</sup>.

### Removal of Congo red

Removal of Congo red was conducted at room temperature (~25 °C) in a 250 mL beaker. 50 mg of a MOF sample was added to 200 mL of an organic dye solution (~50 mg L<sup>−1</sup>), under natural light and magnetic stirring; every 10 to 30 min, a 10 mL aliquot of the sample was analyzed by ultraviolet spectrometry (UV-2550, 220 V, Shimadzu Instruments (Suzhou) Co., Ltd.). The removal rate was calculated by  $C = (C_0 - C_t)/C_0 \times 100\%$ , where  $C_0$  is the initial concentration and  $C_t$  is the concentration of after  $t$  minute.

## 4. Conclusions

In summary, the microwave-assisted ball milling approach was exploited successfully for the effective and reliable synthesis of

Table 3 The Zeta potentials of the four MOFs

	MOFs-1	MOFs-2	MOFs-3	MOFs-4
Zeta potentials (mV)	208	413	316	339





four robust metal–organic frameworks, namely MOFs-1–4, in aqueous solutions, whose structures were confirmed by FTIR, XRD, SEM, and TGA. The MOFs showed good capacity to adsorb CR, and could therefore be employed for the removal of CR from wastewater in the future. MOF-2—prepared from trimesic acid and cobalt(II) acetate tetrahydrate—exhibited the highest adsorption capacity for the removal of CR. These results demonstrate that the microwave-assisted ball milling process is an effective method for fast and simple preparation of MOFs for practical applications.

## Conflicts of interest

There are no conflicts to declare.

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