


Cite this: *RSC Adv.*, 2017, 7, 51993

Synthesis of core–shell structured magnetic mesoporous silica microspheres with accessible carboxyl functionalized surfaces and radially oriented large mesopores as adsorbents for the removal of heavy metal ions†

Shushan Hou,  Xiaoju Li, Hongli Wang, Minggang Wang, Ying Zhang, Yue Chi* and Zhankui Zhao*

In this work, we report the fabrication of novel multifunctional microspheres which are composed of ordered mesoporous silica shells with accessible carboxyl functionalized surfaces and radially oriented large mesopores, and nonporous silica-coated magnetite cores (Fe_3O_4), aiming to remove heavy metal ions from aqueous media. The well-designed multifunctional microspheres were thoroughly characterized with transmission electron microscopy, scanning electron microscopy, infrared spectroscopy, N_2 adsorption–desorption, X-ray diffraction as well as magnetization measurements. The as-prepared microspheres possess unique properties including stably grafted and accessible carboxyl groups, highly open mesopores (11.03 nm), high magnetization (34.5 emu g^{-1}), and large BET surface areas ($165 \text{ m}^2 \text{ g}^{-1}$), and as a result, the as-prepared microspheres exhibit an enhanced performance for the removal of $\text{Cd}(\text{II})$, $\text{Cu}(\text{II})$, and $\text{Pb}(\text{II})$ from wastewater, with a high adsorption capacity, a rapid adsorption rate, and an easy magnetically separable process.

Received 12th August 2017
Accepted 13th October 2017

DOI: 10.1039/c7ra08937f

rsc.li/rsc-advances

1. Introduction

Rapid industrialization has generated heavy metal contamination and intensified the presence of heavy metals in water. Accumulated heavy metals do great harm to the environment and to human health even at quite low concentrations. Many methods have been applied for treating heavy metals in wastewater, such as chemical precipitation,¹ solvent extraction,² ion-exchange,³ adsorption^{4–6} and nanofiltration.⁷ Among these methods, adsorption has received more attention due to its high efficiency, simplicity, and low cost.^{8–10}

Owing to its excellent features (such as high surface area, large pore volume, regular and tunable pore size, and chemical stability), mesoporous silica has been extensively employed as an adsorbent to remove different heavy metal ions.^{11,12} Up to now, although mesoporous silica adsorbents have been developed, their practical applications in treating wastewater are still limited by inconvenient recycling. Recently, significant efforts have been devoted to the

application of core–shell structured magnetic mesoporous silica in the adsorption of heavy metals due to convenient magnetic separation and recycling processes. To get enhanced adsorption performance, a further step beyond the preparation of mesoporous silica is the construction of core–shell structured magnetic mesoporous silica by combining the advantages of mesoporosity with the magnetic properties, which could lead to excellent candidates for treating heavy metals in wastewater.^{13,14}

On the other hand, functional groups such as amino or thiol groups can be immobilized on mesoporous silica by a grafting process to fine-tune the interactions with heavy metal ions, through which the adsorption performance can be elevated to the next level.^{15,16} However, the uncontrollable post-synthesis grafting process may give rise to pore plugging, which is disadvantageous to the adsorption capacity and kinetics. The construction of mesoporous silica integrating functionalized surfaces and accessible mesopores has long been one of the most attractive targets for various groups, aiming to make the surface active for the immobilization of heavy metal ions and to improve the efficiency of mesopore usage and rapid mass transfer.^{17,18}

On the basis of comprehensive considerations, we prepared core–shell structured magnetic mesoporous silica microspheres with a carboxyl functionalized surface and radially oriented large mesopores, which combined the advantages of good

College of Material Science and Engineering, Key Laboratory of Advanced Structural Materials, Ministry of Education, Changchun University of Technology, Changchun, 130012, China. E-mail: yuechi@ccut.edu.cn; zhaozk@ccut.edu.cn; Fax: +86 431 85716644; Tel: +86 431 85716644

† Electronic supplementary information (ESI) available. See DOI: 10.1039/c7ra08937f



magnetic separability, suitable surface properties and a mesoporous structure.¹⁹ The radially oriented large mesopores on the magnetic microspheres are sufficiently large avoiding the plugging caused by surface modification, so that the carboxyl groups immobilized inside the mesopores could maintain accessible formation. The obtained core-shell structured magnetic mesoporous silica microspheres showed fast adsorption speed, high capacity and good recycling stability for trapping Cd(II), Cu(II), and Pb(II) from wastewater.

2. Experimental

2.1 Chemicals

Cetyltrimethylammonium bromide (CTAB) and (3-aminopropyl)trimethoxysilane (APTMS) were purchased from Aladdin Industrial Corporation. Other chemicals were purchased from Beijing Chemical Corp. All chemicals were directly used as received without any further purification. Deionized water was used in all experiments.

2.2 Methods

2.2.1 Synthesis of the Fe₃O₄@SiO₂@mSiO₂ microspheres.

The Fe₃O₄ and Fe₃O₄@SiO₂ microspheres were prepared as previously described.²⁰ Briefly, 0.6 g of CTAB and 0.2 g of triethanolamine (TEA) were dissolved in deionized water (60 mL) by vigorous mechanical stirring to obtain a clear dispersion. Afterwards, 0.15 g of the Fe₃O₄@SiO₂ microspheres was added and treated ultrasonically for 30 min. Into the resulting dispersion, 0.2 mL of tetraethyl orthosilicate (TEOS) and 0.18 mL of cyclohexane were added consecutively. The mixture was ultrasonicated for 15 min and was then mechanically stirred at 60 °C for 12 h. The product was collected with the help of a magnetic field, followed by washing repeatedly with deionized water and ethanol. Finally, the product was dried at 60 °C overnight, then calcined at 550 °C for 6 h (here, it was named Fe₃O₄@SiO₂@mSiO₂).

2.2.2 Synthesis of the amino functionalized Fe₃O₄@SiO₂@mSiO₂ microspheres. In a typical procedure, 0.1 g of the Fe₃O₄@SiO₂@mSiO₂ microspheres was added into the mixed solution containing 50 mL of toluene and 4 mL of APTMS, then the mixture was refluxed at 110 °C with continuous stirring for 24 h. The product was washed with water and ethanol several times, and dried at 60 °C for 6 h (here, it was named Fe₃O₄@SiO₂@mSiO₂-NH₂).

2.2.3 Synthesis of the carboxyl functionalized Fe₃O₄@SiO₂@mSiO₂ microspheres. 0.1 g of the Fe₃O₄@SiO₂@mSiO₂-NH₂ microspheres was dispersed in a 30 mL glutaral solution (5 wt%) and ultrasonicated for 1 h. The recovered solid was washed with deionized water to remove the physically adsorbed species, and re-dispersed in 30 mL of a H₂O₂ solution (30 wt%) by mechanically stirring at 40 °C for 6 h. The product was magnetically separated, washed several times with deionized water, and dried at 60 °C for 6 h (here, it was named Fe₃O₄@SiO₂@mSiO₂-COOH).

2.3 Characterizations

Transmission electron microscopy (TEM) was performed on a JEM-2000. Scanning electron microscopy (SEM) was measured on a JSM-5600. Infrared spectroscopy was measured on a Perkin-Elmer series using the KBr pellet technique. The specific surface area was measured at 77 K using N₂ adsorption-desorption isotherms and BJH methods on a Micromeritics ASAP 2020 HD88 analyzer (USA). X-ray diffraction (XRD) was carried out using a Bruker D8 Advance X-ray diffractometer with a Cu K α X-ray source operating at 9 kV and 200 mA. Magnetization measurements of the samples were measured with a Quantum Design vibrating sample magnetometer (VSM) at room temperature in an applied magnetic field sweeping from -15 to 15 kOe. The concentrations of the heavy metal ions in the solutions were measured by atomic absorption spectroscopy.

2.4 Regeneration studies

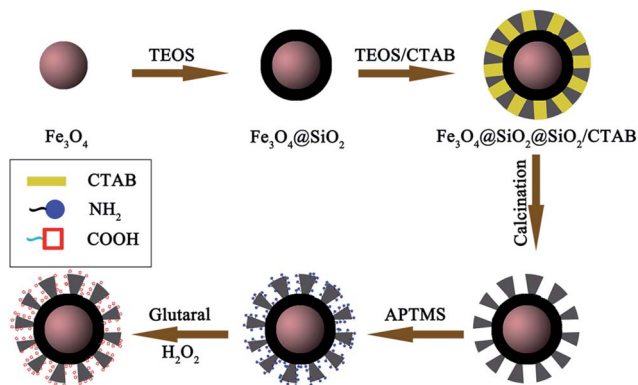
For the potential practical application of the Fe₃O₄@SiO₂@mSiO₂-COOH microspheres, adsorbent reusability is desirable. In this work, the used Fe₃O₄@SiO₂@mSiO₂-COOH microspheres were transferred to 1 M HCl solution and shaken by ultrasonication for 1 h, then washed with deionized water several times in order to reuse them in the next cycle.

3. Results and discussion

3.1 Controlled fabrication of multifunctional core-shell structured magnetic mesoporous silica microspheres

The core-shell structured Fe₃O₄@SiO₂@mSiO₂-COOH microspheres can be constructed through a multistep approach (Scheme 1). First, the Fe₃O₄ microspheres are prepared *via* a solvothermal reaction based on the reduction of Fe(III) salts. The Fe₃O₄ microspheres have a nearly spherical shape with a mean diameter of about 350 nm (Fig. S1a-c, see ESI†). By using the Stöber method, a thin silica layer can coat the Fe₃O₄ microspheres to form the Fe₃O₄@SiO₂ microspheres (Fig. S1d-f†). The SiO₂ shell can effectively prevent the chemical degradation and aggregation of the Fe₃O₄ microspheres for application in harsh liquid environments. The Fe₃O₄@SiO₂ microspheres still keep the spherical shape and exhibit a core-shell structure. It also can be clearly seen that the Fe₃O₄ cores are black spheres and that the SiO₂ shell shows a gray color with an average thickness of about 30 nm. The Fe₃O₄@SiO₂ microspheres are subsequently coated with a layer of mesoporous SiO₂ through a surfactant templating approach involving the controlled hydrolysis of TEOS with CTAB as the surfactant, and then calcined in air to form the radially oriented large mesopores (Fig. S1g-i†). The Fe₃O₄@SiO₂@mSiO₂ microspheres are nearly monodispersed with a clearly distinguishable three-layer structure including an Fe₃O₄ core, a SiO₂ interlayer, and a mesoporous SiO₂ outerlayer with a ~50 nm thickness. Finally, in order to introduce the carboxyl functional groups, the Fe₃O₄@SiO₂@mSiO₂ microspheres are first activated using APTMS, resulting in amino-functionalized Fe₃O₄@SiO₂@mSiO₂. Then glutaral is grafted on these amino groups and





Scheme 1 Schematic illustration of the formation process of the $\text{Fe}_3\text{O}_4@ \text{SiO}_2@m\text{SiO}_2\text{-COOH}$ microspheres.

treated with H_2O_2 (for details, see Experimental section), through which the $\text{Fe}_3\text{O}_4@ \text{SiO}_2@m\text{SiO}_2\text{-COOH}$ microspheres are prepared (Fig. 1a). The $\text{Fe}_3\text{O}_4@ \text{SiO}_2@m\text{SiO}_2\text{-COOH}$ (Fig. 1b and c) microspheres with a mean diameter of about 510 nm still keep the morphological properties of the $\text{Fe}_3\text{O}_4@ \text{SiO}_2@m\text{SiO}_2$ microspheres. It can be seen that the $\text{Fe}_3\text{O}_4@ \text{SiO}_2@m\text{SiO}_2\text{-COOH}$ microspheres are well dispersed without the occurrence of aggregation and the radially oriented large mesopores are well preserved after the multistep modification, indicating the uniform properties of the sample.

FTIR spectroscopy was carried out throughout the whole surface modification process of the $\text{Fe}_3\text{O}_4@ \text{SiO}_2@m\text{SiO}_2$, $\text{Fe}_3\text{O}_4@ \text{SiO}_2@m\text{SiO}_2\text{-NH}_2$ and $\text{Fe}_3\text{O}_4@ \text{SiO}_2@m\text{SiO}_2\text{-COOH}$ microspheres (Fig. 2). All microspheres exhibit absorption bands centered around 575 cm^{-1} (Fe-O vibration modes), 799 cm^{-1} (Si-O-Si symmetric stretching), 1081 cm^{-1} (Si-O-Si asymmetric stretching), 1635 cm^{-1} (the $\delta_{\text{O-H}}$ of adsorbed water and $\delta_{\text{N-H}}$ of free or protonated amino groups), and 3400 cm^{-1} (O-H and N-H stretching bands).¹⁶ Compared with the $\text{Fe}_3\text{O}_4@ \text{SiO}_2@m\text{SiO}_2$ microspheres (Fig. 2a), the absorption peak of the Si-OH bands at 949 cm^{-1} in the $\text{Fe}_3\text{O}_4@ \text{SiO}_2@m\text{SiO}_2\text{-NH}_2$ microspheres (Fig. 2b) disappeared upon functionalization, indicating that aminosilane groups are grafted on surface silanols. Moreover, two new absorption bands at 1589 cm^{-1} (N-H asymmetric stretching) and 691 cm^{-1} (N-H symmetric stretching) can be observed, further indicating the introduction of amino groups. After the $\text{Fe}_3\text{O}_4@ \text{SiO}_2@m\text{SiO}_2\text{-NH}_2$

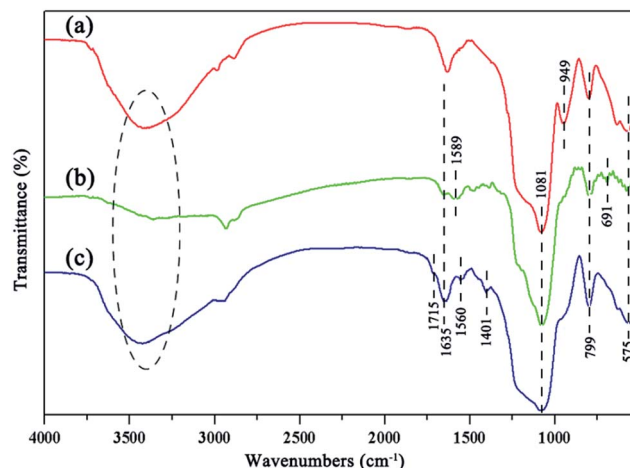


Fig. 2 FTIR spectra of the $\text{Fe}_3\text{O}_4@ \text{SiO}_2@m\text{SiO}_2$ (a), $\text{Fe}_3\text{O}_4@ \text{SiO}_2@m\text{SiO}_2\text{-NH}_2$ (b), and $\text{Fe}_3\text{O}_4@ \text{SiO}_2@m\text{SiO}_2\text{-COOH}$ (c) microspheres.

microspheres react with glutaral and hydrogen peroxide (Fig. 2c), the absorption bands at 1589 cm^{-1} (N-H asymmetric stretching) and 691 cm^{-1} (N-H symmetric stretching) are nearly invisible, indicating new species grafting on the amino groups. The new absorption at 1715 cm^{-1} assigned to the -COOH groups is clearly visible, which suggests that the carboxylic functional groups are successfully introduced upon multi-step surface functionalization. In addition, the new bands at 1560 cm^{-1} and 1401 cm^{-1} can be attributed to the symmetric COO^- vibration and to the carboxyl-carbonate structures, respectively. By reaction with NaOH, when the pH is equal to 7, the determined COOH loading is *ca.* 0.05 mmol g^{-1} for the $\text{Fe}_3\text{O}_4@ \text{SiO}_2@m\text{SiO}_2\text{-COOH}$ microspheres.

N_2 adsorption-desorption measurements of the $\text{Fe}_3\text{O}_4@ \text{SiO}_2@m\text{SiO}_2$ and $\text{Fe}_3\text{O}_4@ \text{SiO}_2@m\text{SiO}_2\text{-COOH}$ microspheres show type IV curves with H1 hysteresis loops (Fig. 3A), indicating well-uniform mesopores. In the pore size distribution curves, the samples exhibit a similar narrow pore size distribution in the range of 5–20 nm (Fig. 3B). The pore size, surface area and pore volumes slightly reduced after surface modification (Table 1), which can be attributed to the introduction of functional groups in the mesopores. According to previous studies, surface modification would lead to a remarkable perturbation of the pore size, surface area, and pore volume.

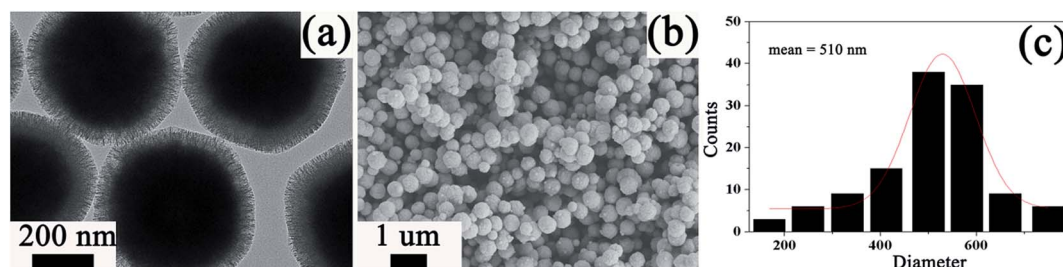


Fig. 1 TEM and SEM of the $\text{Fe}_3\text{O}_4@ \text{SiO}_2@m\text{SiO}_2\text{-COOH}$ microspheres (a and b), and size distribution histogram of the $\text{Fe}_3\text{O}_4@ \text{SiO}_2@m\text{SiO}_2\text{-COOH}$ microspheres calculated from the SEM images (c).



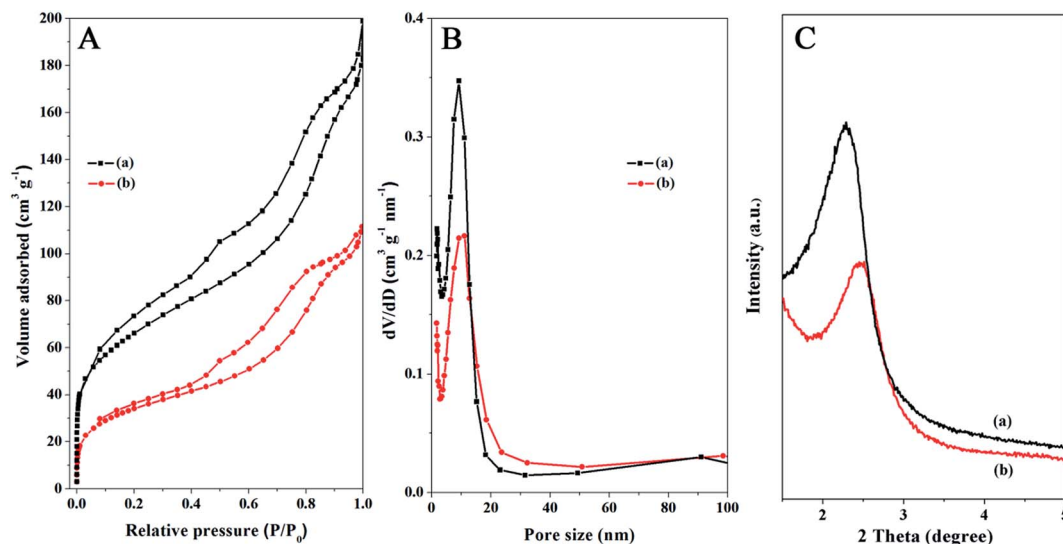


Fig. 3 N₂ adsorption-desorption isotherms (A), the pore size distributions (B), and the small-angle XRD patterns (C) of the Fe₃O₄@SiO₂@mSiO₂ (a) and Fe₃O₄@SiO₂@mSiO₂-COOH (b) microspheres.

However, no such phenomenon occurred in this work, suggesting that the surface modification method adopted in this work is suitable and the highly open mesopores can be preserved. The small angle XRD patterns of the Fe₃O₄@SiO₂@mSiO₂ and Fe₃O₄@SiO₂@mSiO₂-COOH microspheres show a strong diffraction peak between $2\theta = 2^\circ$ and 3° , indicating an ordered mesostructure (Fig. 3C). After surface modification, the mesoporous nature is intact, which is consistent with the TEM characterization. However, a gradual decrease in the intensities and a slight shift to higher angles for the peak can be observed, suggesting that the functional groups have been grafted inside the pores, which is similar to the previous case. The wide angle XRD pattern of the Fe₃O₄@SiO₂@mSiO₂-COOH microspheres exhibits almost the same feature as that of pure Fe₃O₄ except for a broad peak centered at $2\theta = 22^\circ$ (Fig. S2†), indicating that the coated SiO₂ is amorphous.

The magnetic properties of the Fe₃O₄@SiO₂@mSiO₂ and Fe₃O₄@SiO₂@mSiO₂-COOH microspheres are characterized at room temperature as shown in Fig. 4. The saturation magnetizations of the Fe₃O₄@SiO₂@mSiO₂ and Fe₃O₄@SiO₂@mSiO₂-COOH microspheres are 37.6 and 34.5 emu g⁻¹, respectively. Although the magnetization saturation value of the Fe₃O₄@SiO₂@mSiO₂-COOH microspheres is lower than that of the Fe₃O₄@SiO₂@mSiO₂ microspheres due to the grafting of the carboxyl group, the Fe₃O₄@SiO₂@mSiO₂-COOH microspheres

still show strong magnetization. It can be seen that the Fe₃O₄@SiO₂@mSiO₂-COOH microspheres are quickly attracted to the side of the vial with the aid of an external magnetic field, and the microspheres can redisperse well again with shaking (Fig. 4, inset).

3.2 Adsorption behavior of the Fe₃O₄@SiO₂@mSiO₂-COOH microspheres

Fig. 5a shows the adsorption isotherms of the heavy metal ions (Cd(II), Cu(II), and Pb(II)) onto the prepared adsorbent Fe₃O₄@SiO₂@mSiO₂-COOH microspheres at room temperature, which are expressed as plots of the solid-phase pollutant concentration vs. the liquid-phase pollutant concentration. It can be observed that the amount of heavy metal ions adsorbed gradually increases with the concentration of heavy metal ions. The maximum adsorption capacities for Cd(II), Cu(II), and Pb(II) are 395, 476, and 357 mg g⁻¹, respectively. The adsorption of heavy metal ions on the Fe₃O₄@SiO₂@mSiO₂-COOH microspheres quickly reaches 88%, 77%, and 88% (349, 366 and 316 mg g⁻¹)

Table 1 Textural properties of the Fe₃O₄@SiO₂@mSiO₂ and Fe₃O₄@SiO₂@mSiO₂-COOH microspheres

Sample	Pore size (nm)	BET surface area (m ² g ⁻¹)	Pore volume (cm ³ g ⁻¹)
Fe ₃ O ₄ @SiO ₂ @mSiO ₂	11.08	193	0.23
Fe ₃ O ₄ @SiO ₂ @mSiO ₂ -COOH	11.03	165	0.17

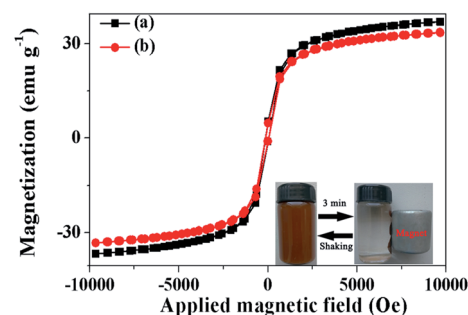


Fig. 4 Magnetic curves of the Fe₃O₄@SiO₂@mSiO₂ (a) and the Fe₃O₄@SiO₂@mSiO₂-COOH (b) microspheres, where the inset shows the separation process of the Fe₃O₄@SiO₂@mSiO₂-COOH microspheres using a magnet.



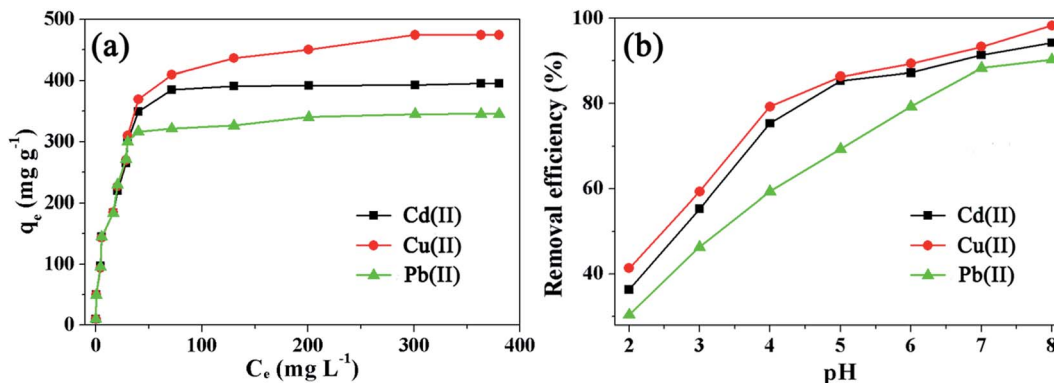


Fig. 5 Adsorption isotherms (a) and the pH effect (b) of the $\text{Fe}_3\text{O}_4@\text{SiO}_2@\text{mSiO}_2\text{-COOH}$ microspheres with different heavy metal ions.

of the total capacity at low concentrations below 420 mg L^{-1} for Cd(II), Cu(II), and Pb(II), respectively. $\text{Fe}_3\text{O}_4@\text{SiO}_2@\text{mSiO}_2\text{-COOH}$ microspheres exhibit a high adsorption capacity at low concentration, that may be derived from the complexation between the metal ions and the surface carboxyl groups.

The adsorption capacity of a given porous adsorbent strongly depends on the pore texture and surface properties. $\text{Fe}_3\text{O}_4@\text{SiO}_2@\text{mSiO}_2\text{-COOH}$ microspheres combine the highly open large mesopores and the accessible carboxyl functionalized surfaces, which lead to a high adsorption capacity compared with that of other similar magnetic adsorbents (Table S1†). However, because of the substantial increase in mass caused by the large-sized Fe_3O_4 , the $\text{Fe}_3\text{O}_4@\text{SiO}_2@\text{mSiO}_2\text{-COOH}$ microspheres show a smaller BET surface area value, thus leading to a lower adsorption capacity compared with that of some mesoporous adsorbents.

The adsorption of heavy metal ions is significantly affected by the solution pH. As shown in Fig. 5b, when the pH increases from 2.0 to 8.0, the percentage of removed heavy metal ions (Cd(II), Cu(II), and Pb(II)) increases gradually. The adsorption capacity of the $\text{Fe}_3\text{O}_4@\text{SiO}_2@\text{mSiO}_2\text{-COOH}$ microspheres is strong under near-neutral and alkaline conditions while it is poor under strong acidic conditions. Under near-neutral and alkaline conditions, the negatively charged carboxylate ions exhibit a much stronger coordinative affinity towards heavy metal ions, and they can capture the target metal ions by forming chelate complexes, which leads to the high removal efficiency of Cd(II), Cu(II), and Pb(II).

Three adsorption models named Langmuir, Freundlich and Temkin are employed to describe the adsorption, and are

presented in eqn S(2)–S(4)†. The comparisons between the empirical adsorption isotherms and their fitting curves obtained from the above three models are shown in Fig. S3† and the corresponding parameters obtained from nonlinear regression are shown in Table 2. Based on the value of the correlation coefficients (R^2), it can be seen that the Langmuir isotherm is a more suitable isotherm for the experimental data, illustrating that the form of the heavy metal ions on the surface of the $\text{Fe}_3\text{O}_4@\text{SiO}_2@\text{mSiO}_2\text{-COOH}$ microspheres is a monolayer adsorption and the major binding sites in the sorbent are homogeneously distributed. A high value of K_L indicates higher affinity.

The kinetics of Cd(II), Cu(II), and Pb(II) onto the $\text{Fe}_3\text{O}_4@\text{SiO}_2@\text{mSiO}_2\text{-COOH}$ microspheres are presented in Fig. 6a. It can be seen that an adsorption equilibrium is attained only within ~ 30 min. Notably, the adsorption of the heavy metal ions is very rapid in the first 20 min, then the adsorption rate reduces and the adsorption amount of the heavy metal ions is almost constant from 20–30 min and gradually increases with prolonged contact time. The initial rapid adsorption within the first 20 min can be ascribed to a slight resistance for the diffusion and adsorption of the heavy metal ions into the highly open and carboxyl functionalized mesopores. After that, heavy metal ions have to traverse farther and deeper into the available adsorption sites, thus slowing down the adsorption rate from 20–30 min. The intra-particle diffusion model describes the movement of species from the bulk of the solution to the solid phase. It can be defined as follows:

$$q_t = k_{\text{id}} t^{1/2} + c \quad (1)$$

Table 2 Adsorption isotherm models and corresponding parameters for the removal of Cd(II), Cu(II), and Pb(II) on the $\text{Fe}_3\text{O}_4@\text{SiO}_2@\text{mSiO}_2\text{-COOH}$ microspheres

Metal ions	Langmuir isotherm			Freundlich isotherm			Temkin isotherm		
	K_L (L g ⁻¹)	a_L (L mmol ⁻¹)	R^2	K_F (L g ⁻¹)	n_F	R^2	A	B	R^2
Cd(II)	412	0.059	0.96	110	0.219	0.89	46	12	0.91
Cu(II)	513	0.039	0.96	100	0.278	0.93	38	9	0.63
Pb(II)	359	0.102	0.96	116	0.202	0.83	40	8	0.69



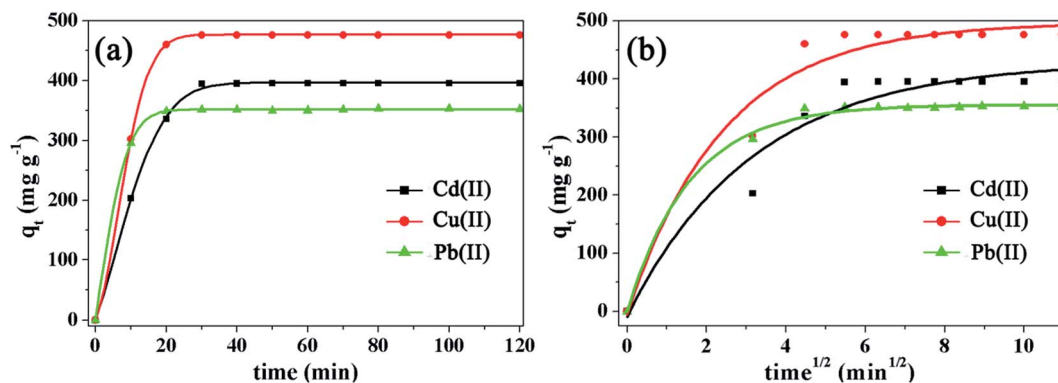


Fig. 6 Adsorption rate (a) and intra-particle diffusion model (b) of the $\text{Fe}_3\text{O}_4@\text{SiO}_2@\text{mSiO}_2\text{-COOH}$ microspheres with different heavy metal ions.

Table 3 Kinetic adsorption parameters of heavy metal ions on the $\text{Fe}_3\text{O}_4@\text{SiO}_2@\text{mSiO}_2\text{-COOH}$ microspheres obtained using a pseudo-first-order kinetic model and a pseudo-second-order kinetic model

Metal ions	Pseudo-first-order			Pseudo-second-order		
	k_1 (min^{-1})	q_{e1} (mg g^{-1})	R^2	k_2 (min^{-1})	q_{e2} (mg g^{-1})	R^2
Cd(II)	69	370	0.76	0.12	445	0.95
Cu(II)	85	457	0.86	0.20	515	0.96
Pb(II)	65	346	0.97	0.55	363	0.99

where q_t is the amount of the adsorbates at time t (mg g^{-1}), k_{id} is the intra-particle diffusion equation constant ($\text{mg g}^{-1} \text{min}^{-1/2}$), c is the film thickness, and t is the time (min). As seen in Fig. 6b, the plots of q_t against $t^{1/2}$ first fit the linear behavior according to the intra-particle diffusion model, then a deviation occurs, indicating that the adsorption is at least a two-step process. The first step is controlled by the intra-particle diffusion and the second step can be ascribed to the pore diffusion.

In order to further understand the mechanisms of adsorption and the heterogeneity of the adsorbent surface, the kinetic experimental data have been evaluated using a pseudo-first-order kinetic model and a pseudo-second-order kinetic model,²⁵ respectively (Fig. S3†). Due to diffusion and adsorption being often experimentally inseparable, the heavy metal ion adsorption onto the $\text{Fe}_3\text{O}_4@\text{SiO}_2@\text{mSiO}_2\text{-COOH}$ microspheres may be a complicated process that includes diffusion, coordinate bond formation and chemical reactions simultaneously. From the results obtained (Table 3), the value of R^2 of the pseudo-second-order model is closer to 1 compared with that of the pseudo-first-order model, suggesting that the adsorption process can be accurately described by the pseudo-second-order model, thus the adsorption of heavy metal ions onto the $\text{Fe}_3\text{O}_4@\text{SiO}_2@\text{mSiO}_2\text{-COOH}$ microspheres may take place based on a chemical process through the sharing or the exchange of electrons. Additionally, it is found that repeated washing of the used $\text{Fe}_3\text{O}_4@\text{SiO}_2@\text{mSiO}_2\text{-COOH}$ microspheres with an acid solution and with deionized water can lead to its regeneration. Through repeated adsorption-desorption cycles as shown in Fig. 7, the adsorption efficiency decreased slightly in sequential cycles. And $\text{Fe}_3\text{O}_4@\text{SiO}_2@\text{mSiO}_2\text{-COOH}$ still shows absorption

bands at 1715 cm^{-1} , 1560 cm^{-1} and 1401 cm^{-1} in the FTIR spectrum after multiple reuses (Fig. S4†), proving the existence of the carboxylic functional groups. However, the decrease in the absorption intensity indicates the leaching of adsorption sites, thus resulting in the adsorption efficiency decreasing in sequential cycles. After multiple reuses $\text{Fe}_3\text{O}_4@\text{SiO}_2@\text{mSiO}_2\text{-COOH}$ still maintains a high adsorption capacity, and the pore size, BET surface area and pore volume structure of $\text{Fe}_3\text{O}_4@\text{SiO}_2@\text{mSiO}_2\text{-COOH}$ only exhibit a slight decrease (Table S2†), suggesting the stability and reusability of the $\text{Fe}_3\text{O}_4@\text{SiO}_2@\text{mSiO}_2\text{-COOH}$ microspheres, which is owing to their unique textural structure and surface properties.

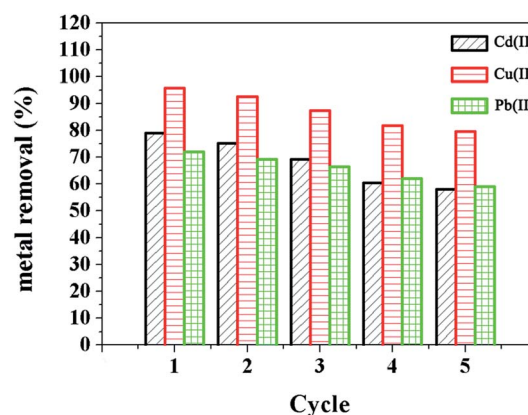


Fig. 7 Sorption-desorption cycles of the $\text{Fe}_3\text{O}_4@\text{SiO}_2@\text{mSiO}_2\text{-COOH}$ microspheres with different heavy metal ions.



4. Conclusions

In conclusion, we describe the synthesis of novel core-shell structured magnetic mesoporous silica microspheres combining well designed surface properties and a mesoporous structure and we also describe their enhanced adsorption performance. The obtained $\text{Fe}_3\text{O}_4@\text{SiO}_2@\text{mSiO}_2\text{-COOH}$ microspheres possess a number of important features as a recyclable adsorbent: they contain a mesoporous SiO_2 shell with radially oriented large mesopores (11.03 nm) and accessible carboxyl functionalized surfaces for an enhanced adsorption performance, a high field-responsive magnetic Fe_3O_4 core for efficient magnetic separation, and a SiO_2 inter-layer for protecting the Fe_3O_4 core from chemical dissolution. Owing to the well-defined nanostructure and the highly integrated functions, the prepared material is very suitable as an adsorbent for the removal of heavy metal ions (Cd(II) , Cu(II) , and Pb(II)), and exhibits a high adsorption capacity, a rapid adsorption rate, and an easy separation process. In addition, the adsorbent can be regenerated by acid treatment and reused in the next cycle with good stability. To get enhanced adsorption performance, this contribution takes further steps to coordinate mesoporous structures with surface properties, which could possibly inspire a lot of future research.

Conflicts of interest

There are no conflicts to declare.

Acknowledgements

This work was supported by the Jilin Province Science & Technology Department Project (Grant no. 20150520020JH, 20160519020JH and 20170520122JH), the National Natural Science Foundation of China (Grant no. 51608050, 51601018 and 51671035), and the Science and Technology Research Project of the Education Department of Jilin Province (Grant no: 2016327).

References

- 1 M. M. Matlock, B. S. Howerton and D. A. Atwood, Chemical precipitation of heavy metals from acid mine drainage, *Water Res.*, 2002, **36**(19), 4757–4764.
- 2 S. Otsuki, T. Nonaka, N. Takashima, *et al.*, Oxidative desulfurization of light gas oil and vacuum gas oil by oxidation and solvent extraction, *Energy Fuels*, 2000, **14**(6), 1232–1239.
- 3 P. B. Hamilton, Ion exchange chromatography of amino acids. A single column, high resolving, fully automatic procedure, *Anal. Chem.*, 1963, **35**, 2055–2064.
- 4 S. H. Huang and D. H. Chen, Rapid removal of heavy metal cations and anions from aqueous solutions by an amino-functionalized magnetic nano-adsorbent, *J. Hazard. Mater.*, 2009, **163**(1), 174–179.
- 5 Y. Gao, X. Chen, J. Zhang, *et al.*, Chitin-Derived Mesoporous, Nitrogen-Containing Carbon for Heavy-Metal Removal and Styrene Epoxidation, *ChemPlusChem*, 2015, **80**(10), 1556–1564.
- 6 J. Song, H. Kong and J. Jang, Adsorption of heavy metal ions from aqueous solution by polyrhodanine-encapsulated magnetic nanoparticles, *J. Colloid Interface Sci.*, 2011, **359**(2), 505–511.
- 7 B. A. M. Al-Rashdi, D. J. Johnson and N. Hilal, Removal of heavy metal ions by nanofiltration, *Desalination*, 2013, **315**, 2–17.
- 8 S. Babel and T. A. Kurniawan, Low-cost adsorbents for heavy metals uptake from contaminated water: a review, *J. Hazard. Mater.*, 2003, **97**(1), 219–243.
- 9 G. Crini, Non-conventional low-cost adsorbents for dye removal: a review, *Bioresour. Technol.*, 2006, **97**(9), 1061–1085.
- 10 M. Rafatullah, O. Sulaiman, R. Hashim, *et al.*, Adsorption of methylene blue on low-cost adsorbents: a review, *J. Hazard. Mater.*, 2010, **177**(1), 70–80.
- 11 L. Mercier and T. J. Pinnavaia, Access in mesoporous materials: advantages of a uniform pore structure in the design of a heavy metal ion adsorbent for environmental remediation, *Adv. Mater.*, 1997, **9**(6), 500–503.
- 12 X. Feng, G. E. Fryxell, L. Q. Wang, *et al.*, Functionalized monolayers on ordered mesoporous supports, *Science*, 1997, **276**(5314), 923–926.
- 13 J. P. K. Reynhardt, Y. Yang, A. Sayari, *et al.*, Periodic mesoporous silica-supported recyclable rhodium-complexed dendrimer catalysts, *Chem. Mater.*, 2004, **16**(21), 4095–4102.
- 14 N. Baccile and F. Babonneau, Organo-modified mesoporous silicas for organic pollutant removal in water: Solid-state NMR study of the organic/silica interactions, *Microporous Mesoporous Mater.*, 2008, **110**(2), 534–542.
- 15 S. Wu, F. Li, R. Xu, *et al.*, Synthesis of thiol-functionalized MCM-41 mesoporous silicas and its application in Cu(II) , Pb(II) , Ag(I) , and Cr(III) removal, *J. Nanopart. Res.*, 2010, **12**(6), 2111–2124.
- 16 Q. Yuan, N. Li, Y. Chi, *et al.*, Effect of large pore size of multifunctional mesoporous microsphere on removal of heavy metal ions, *J. Hazard. Mater.*, 2013, **254**, 157–165.
- 17 J. Sun, H. Zhang, R. Tian, *et al.*, Ultrafast enzyme immobilization over large-pore nanoscale mesoporous silica particles, *Chem. Commun.*, 2006, **12**, 1322–1324.
- 18 J. Yang, D. Shen, Y. Wei, *et al.*, Monodisperse core-shell structured magnetic mesoporous aluminosilicate nanospheres with large dendritic mesochannels, *Nano Res.*, 2015, **8**(8), 2503–2514.
- 19 H. Wang, Y. F. Yu, Q. W. Chen, *et al.*, Carboxyl-functionalized nanoparticles with magnetic core and mesopore carbon shell as adsorbents for the removal of heavy metal ions from aqueous solution, *Dalton Trans.*, 2011, **40**(3), 559–563.
- 20 Y. Chi, Q. Yuan, Y. Li, *et al.*, Synthesis of $\text{Fe}_3\text{O}_4@\text{SiO}_2\text{-Ag}$ magnetic nanocomposite based on small-sized and highly dispersed silver nanoparticles for catalytic reduction of 4-nitrophenol, *J. Colloid Interface Sci.*, 2012, **383**(1), 96–102.



- 21 N. Yeddou and A. Bensmaili, Equilibrium and kinetic modelling of iron adsorption by eggshells in a batch system: effect of temperature, *Desalination*, 2007, **206**(1–3), 127–134.
- 22 A. Mittal, L. Kurup and J. Mittal, Freundlich and Langmuir adsorption isotherms and kinetics for the removal of Tartrazine from aqueous solutions using hen feathers, *J. Hazard. Mater.*, 2007, **146**(1), 243–248.
- 23 M. M. Saeed, A. Rusheed and N. Ahmed, Modeling of iron adsorption on HTTA-loaded polyurethane foam using Freundlich, Langmuir and DR isotherm expressions, *J. Radioanal. Nucl. Chem.*, 1996, **211**(2), 283–292.
- 24 A. O. Dada, A. P. Olalekan, A. M. Olatunya, *et al.*, Langmuir, Freundlich, Temkin and Dubinin–Radushkevich isotherms studies of equilibrium sorption of Zn^{2+} unto phosphoric acid modified rice husk, *IOSR J. Appl. Chem.*, 2012, **3**(1), 38–45.
- 25 Y. S. Ho and G. McKay, Pseudo-second order model for sorption processes, *Process Biochem.*, 1999, **34**(5), 451–465.

