




Cite this: *RSC Adv.*, 2017, 7, 51426

Surface adsorption and thermodynamic properties of mixed system of ionic liquid surfactants with cetyltrimethyl ammonium bromide†

Li Qin and Xin-Hong Wang *

The effects of temperature and concentration of CTAB added into imidazolium-based ionic liquids ($C_n\text{mimBr}$, $n = 10, 12, 16$) solution were investigated, and surface active parameters and aggregation of these surfactants were explored by surface tension and conductivity measurements, which include critical micelle concentration (CMC), maximum surface excess concentration (Γ_{max}), minimum surface area per surfactant molecule (A_{min}), surface tension at the CMC (γ_{CMC}), adsorption efficiency ($\rho_{C_{20}}$), and surface pressure (Π_{CMC}), as well as thermodynamic parameters of micellization. The results were reported in the presence of CTAB at 0, 4, 8, 12, and 16 mM for $C_{10}\text{mimBr}$, 0, 0.2, 0.4, 0.8, and 1.2 mM for $C_{12}\text{mimBr}$, and 0, 0.04, 0.08, 0.12, and 0.16 mM for $C_{16}\text{mimBr}$ in mixed systems. The increase of the CMCs with temperature were observed, while $\rho_{C_{20}}$, Γ_{max} , and standard entropy of aggregation were decreased, indicating that the aggregation is an enthalpy-driven process within the temperature limits of the present study. FT-IR and ^1H NMR measurements revealed that the changes mentioned above were controlled by the electrostatic effect and hydrophobic interactions between CTAB and ILs molecules.

Received 12th August 2017
Accepted 17th October 2017

DOI: 10.1039/c7ra08915e

rsc.li/rsc-advances

1 Introduction

Ionic liquids (ILs) have gained great attention for their promising role as alternative media in catalysis, separation, and electrochemical processes owing to their unique chemical and physical properties, such as negligible vapor pressure, wide liquidous range, excellent solvation, high thermal stability, ionic conductivity and wide electrochemical windows. Especially, as an environmentally friendly solvent, also called a “green solvent”, ILs make some contribution towards preventing environmental pollution by virtue of their non-volatility. Based on their inherent amphiphilic structure consisting of a non-polar hydrophobic tail and a polar cationic head group, a typical long-chain ILs could possess surface active properties similar to those of conventional cationic surfactants, so they have been extensively studied in the field of colloid and interface science in recent years. For instance, the surface tension and critical micelle concentration (CMC) of long-chain ILs decrease with alkyl chain length. ILs applications in wide fields are closely dependent on such self-assembly behavior.¹ In addition, ILs with different cations and anions have been

reported to understand the effect of the ions on aggregation behavior.^{2–4} However, it is unclear how the surface active and micellization features depend upon their molecular characteristics, concentration, temperature and additives *etc.* In view of this, study of the interaction and microstructural characteristics of micellar systems is always an important subject area in the discipline of colloid and interface science.

Numerous pioneering studies on multicomponent mixtures of ionic liquid have widened the application field of ILs. For instance, phase behaviour of aqueous two-phase systems containing an ionic liquid,⁵ molecular interactions and preparation of nanoparticles,^{6,7} and effects on properties of nanocomposites.⁸ In these studies, the binary mixture is relatively fascinating from an academic point of view wherein including common ionic-ionic systems,^{9–11} and ionic–nonionic systems.^{12–14} The CTAB/bmim-octylSO₄ binary system in aqueous solution at two total concentrations and different CTAB mole fractions was investigated and found to form the CTA-octylSO₄ catanionic complex,⁹ which is considered analogue to the effect of the addition of inorganic salts in such similar systems.¹⁵ The aggregation behaviour of the alkyl triphenylphosphonium bromide ($C_n\text{TPPBr}$, $n = 10, 12, 14, 16$)/Triton X-100 system was studied by Thakkar *et al.*¹³ Furthermore, a generalized phase separation model was discussed in ionic–nonionic mixtures.¹⁶ The ionic liquid TEA(BF₄) and non-ionic surfactant TX-100/Brij-56/Brij-58/Tween-20/Tween-80 in aqueous media at different temperatures were studied, respectively.¹⁷ Some other reports have just been published regarding the IL surfactant applications.^{18,19} Different

Jiangsu Key Laboratory for Bioresources of Saline Solis, School of Chemistry and Environmental Engineering, Yancheng Teachers University, Xiwang Road, Yancheng 224051, Jiangsu Province, P. R. China. E-mail: wangxhong1227@163.com; Fax: +86 515 88233187; Tel: +86 515 88233187

† Electronic supplementary information (ESI) available. See DOI: 10.1039/c7ra08915e



measurement technologies were used to obtain the needed surface activities and thermodynamics functions, such as a small-angle neutron scattering (SANS),^{13,20} dynamic light scattering (DLS) and viscosity measurements,¹⁷ conductivity,^{12,21,22} surface tension measurements,²³ as well different analytical methods, FT-IR,³ UV-visible spectrum,^{22,24} fluorescence spectroscopy,^{4,22,25,26} isothermal titration calorimeter²⁵ and NMR technology.^{27,28}

Apart from the self-assemblies and applications of ILs surfactants in various fields, the aggregation behavior of 1-alkyl-3-methylimidazolium ($C_n\text{mim}^+$) in water has been presented by many research groups.^{13,15,29} The micellization behavior of the DTAB/bdmimCl system has been reported by Pal' group,²⁷ also including their related work on tetradecyltrimethylammonium bromide TTAB/ $C_5\text{mim}/\text{PF}_6$ system,³⁰ CTAB/bdmimBr system,³¹ odmimCl/DTAB mixed system,²¹ and TTAB/bdmimBr mixture.²⁸ However, in all these systems, all studied ILs as an additive were added into traditional surfactant solution to form IL-surfactant mixtures. The CMC values of the mixtures increase with increasing wt% IL content and temperature at the same IL concentration. So it would be interesting to observe the effect of CTAB as an additive on the micellization of IL as no reports have been found to date. Consequently the purpose of the present work is to study the effects of temperature and CTAB addition into ILs ($C_n\text{mimBr}$, $n = 10, 12, 16$) solution on the surface properties of the mixture system and the thermodynamic properties of micellization. This prompted us to carry out systematic investigations with a focus mainly on examining the surface activity and aggregate formation of ionic liquids based on these long-chain ILs and to assess the effect of chain length and head group nature on the CMC, pC_{20} , Γ_{\max} , A_{\min} , and thermodynamic parameters of aggregation by surface tension and electrical conductance measurements for mixed systems. Conductivity and surface tension measurements were used to obtain the CMC values for three IL/CTAB solutions, and the thermodynamic parameters were used to analyze the driving force of micellization. In addition, ^1H NMR and FT-IR measurements were done to understand the interactions between the IL and the surfactant molecules. The results obtained could be compared with similar systems for further applications of binary systems in some areas of separation, adsorption, and other surface research fields.

2 Materials and methods

2.1 Materials

Cetyltrimethyl ammonium bromide (CTAB, MW: 364.45, purity $\geq 99\%$, CAS number: 57-09-0) was obtained from Sinopharm Chemical Reagent Co., Ltd (Shanghai, China). 1-Hexadecyl-3-methylimidazolium bromide ($C_{16}\text{mimBr}$, MW: 387.44, purity $\geq 99.9\%$, CAS: 132361-22-9), 1-dodecyl-3-methylimidazolium bromide ($C_{12}\text{mimBr}$, MW: 331.33, purity $\geq 99.9\%$, CAS: 61546-00-7), and 1-decyl-3-methylimidazolium bromide ($C_{10}\text{mimBr}$, MW: 303.28, purity $\geq 99.9\%$, CAS: 188589-32-4) were purchased from Shanghai Cheng Jie Chemical Co., Ltd., China, and these ILs were dried at 343 K for 24 h in vacuum drying oven (DZF-6050, Shanghai Shupeil Laboratory Equipment

Co., Ltd., China) prior to use. The molecular structures of the ILs and CTAB are shown in Fig. 1. D_2O ($D > 99.96\%$) was obtained from Sigma-Aldrich and used as the solvent in ^1H NMR measurements.

2.2 Methods

2.2.1 Conductivity measurement. The electrical conductivities were measured at different temperatures using a digital conductivity meter (DDSJ-308A, Shanghai Leici Instrument Factory, China) equipped with a platinized platinum electrode. A stable temperature was maintained with a super-thermostatic water bath (ZC-18Q, Tianheng Instrument Factory, Ningbo, China). Before the measurements, the conductivity cell constant was calibrated with a known specific conductance aqueous solution of 0.01 M potassium chloride aqueous solutions. At least three measurements were made every time and the specific conductivity was calculated from the mean value. The conductance due to water ionization ($5.83 \mu\text{s cm}^{-1}$, $\text{pH} = 6.98$) was subtracted from the measured data before data analysis. The uncertainty in conductivity measurements was less than $\pm 0.4\%$. The cell assembly was dipped in glass vials kept in a water bath thermostatted at the studied temperature maintained to ± 0.1 K.

2.2.2 Surface tension measurement. The surface tension for aqueous solutions of single surfactants and binary mixtures were assessed using the maximum bubble pressure method with a tensiometer (DP-AW-1, Nanjing SangLi Electronic Equipment Factory, China) under atmospheric pressure. The calibration was performed by measuring the surface tension of doubly distilled de-ionized water (71.97 mN m^{-1}) before each experiment. The temperature was maintained within ± 0.1 K by circulating water. The stock solutions were then diluted with fresh distilled water to desired concentrations. The uncertainty in the measured surface tensions was within $\pm 0.03 \text{ mN m}^{-1}$.

The surface excess adsorption, Γ_{\max} , and the minimum area per surfactant molecule, A_{\min} , at the air-solvent interface were calculated using surface tension measurement values according to the Gibbs adsorption isotherm equation, given by eqn (1) and (2):

$$\Gamma_{\max} = -\frac{1}{RT} \left(\frac{\partial \gamma}{\partial \ln c} \right)_T \quad (1)$$

$$A_{\min} = \frac{1}{L\Gamma_{\max}} \quad (2)$$

where R is the gas constant, T is the absolute temperature, γ is the surface tension, and c is the IL concentration of mixed solutions. L is Avogadro's constant. The adsorption efficiency (pC_{20}) is an important surface activity parameter, which is defined by eqn (3):

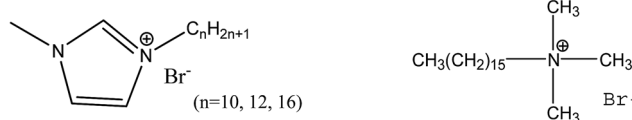


Fig. 1 Molecular structures of ILs and CTAB.



$$pC_{20} = -\log C_{20} \quad (3)$$

Surface pressure at the CMC, Π_{CMC} , is a measure of the efficiency of the surfactant to lower the surface tension of the water, and is obtained by means of eqn (4):

$$\Pi_{\text{CMC}} = \gamma_0 - \gamma_{\text{CMC}} \quad (4)$$

where C_{20} is the effective concentration of all surfactants in the mixed system required to reduce the surface tension of pure water by 20 mN m^{-1} . γ_0 is the surface tension of pure water for the respective temperature, and γ_{CMC} is the measured surface tension at the CMC.

The enthalpic contribution to the micellization process, ΔH_{m} , was determined based on ΔG_{m} , using the Gibbs-Helmholtz relation as shown in eqn (5) and (6). The standard entropy of the micellization, ΔS_{m} , was calculated from eqn (7):

$$\Delta G_{\text{m}} = (2 - \alpha)RT \ln X_{\text{CMC}} \quad (5)$$

$$\Delta H_{\text{m}} = -RT^2 \left[(2 - \alpha) \frac{d \ln X_{\text{CMC}}}{dT} + \ln X_{\text{CMC}} \frac{d(2 - \alpha)}{dT} \right] \quad (6)$$

$$\Delta S_{\text{m}} = \frac{\Delta H_{\text{m}} - \Delta G_{\text{m}}}{T} \quad (7)$$

where X_{CMC} is the CMC in mole fraction, the degree of dissociation, α , was associated to the degree of ionization, which represents the fraction of amphiphile ions in the aggregate neutralized by counter anions. R is the gas constant, and T is temperature. According to eqn (5), ΔG_{m} of micellization at different temperatures may be calculated using X_{CMC} . Thus, the curve of $\Delta G_{\text{m}}/T$ against $1/T$ can be drawn, and ΔH_{m} can be finally calculated from the slopes of tangential lines at various temperatures. The parameters thus obtained are presented in Table 2.

2.2.3 FT-IR spectra measurement. The spectra of the samples were measured using a Thermo Scientific Nicolet IS10 FT-IR spectrometer equipped with a standard KBr beam splitter and DTGS detector. The mixture spectra of the $[C_n\text{mimBr}]/\text{CTAB}$ solution were scanned on the Smart iTR™ diamond ATR reflection accessory for qualitative analysis in this study. The spectra from samples (pure IL or mixture of IL and TX-100 without water) were collected in 40 seconds over the full mid-infrared range (4000 to 400 cm^{-1}). The samples needed to be premixed (molar ratio 1 : 1) for a period of time before analysis. Instrument control, data collection and analysis were performed using the Thermo Scientific OMNIC software package.

2.2.4 ^1H NMR measurements. The NMR chemical shifts for various protons were observed with a Bruker AVANCE III HD 400 M NMR spectrometer at room temperature (about 298 K). Deuterium oxide (D_2O) as a solvent for all the NMR measurements was used to determine chemical shift δ for the $[C_n\text{mimBr}]/\text{CTAB}$ solution.

All the experiments were carried out in doubly distilled de-ionized water.

3 Results and discussion

3.1 Conductivity measurement

3.1.1 CMC determination. Conductivity measurements were carried out in solutions of ionic liquids over a wide concentration range at 298.15–338.5 K to obtain CMC values with the addition of the surfactant CTAB and without CTAB. Sufficient care was taken to avoid high IL concentrations to ensure that the solutions were not highly viscous and were free of air bubbles during the measurements.

For pure ionic liquids solutions, the plots of conductivity against concentration of IL surfactant (Fig. 2) displayed the difference of the curves well, namely the slope in the pre-micellar region was higher than one in post-micellar region regardless of the pure IL surfactants, or mixture solutions. The presence of turning point is attributable to the change of aggregation state from CTAB molecules and IL molecules. Thus, the abscissa value corresponding to the turning point is the CMC value of the studied system, as listed in Table 1. Reasonably, the CMC decreased with increasing length of alkyl chain and such a systematic decrease could be due to the enhanced hydrophobic character of higher alkyl chains.² Therefore, it can be concluded that the interfacial activity of the pure IL surfactants is in the order of $C_{16}\text{mimBr} > C_{12}\text{mimBr} > C_{10}\text{mimBr}$. It is therefore suggested that the aggregation process of long-chain ILs in water needs to be close to that of the micellization process of classical ionic surfactants in aqueous media. Furthermore, the degree of dissociation (α) for these systems was obtained from the ratio of the slopes of the two linear lines drawn through regions corresponding to above and below the CMC value, as listed in Table 1.

Considering the CMC of CTAB is about 0.90 mM,^{9,27} the concentrations values of CTAB that lie on both sides of the CMC of CTAB were selected to investigate the effect of adding CTAB monomer or micelles to IL aqueous solution on the CMC values and surface activities of mixed systems. The representative plots of the conductance of the mixed system ($C_{12}\text{mimBr}/\text{CTAB}$ system) versus IL concentration on addition of different concentrations of CTAB in aqueous medium are shown in Fig. 3, while conductivity measurement results for $C_{10}\text{mimBr}/\text{CTAB}$ and $C_{16}\text{mimBr}/\text{CTAB}$ are depicted in Fig. S1 and S2 (ESI†) and the measured data are given in Table 1. A similar change to that observed in Fig. 3 can be found and the CMCs of the mixed systems can be calculated.

According to the data from Fig. 3, S1 and S2,† the CMCs, the degree of dissociation (α) and the thermodynamic parameters of aggregation for ILs/CTAB in water were calculated at five different temperatures of 298.15, 308.15, 318.15, 328.15 and 338.15 K. The corresponding data are listed in Table 1. The data from Table 1 show that the CMCs value of the $C_n\text{mimBr}$ aqueous systems decreased with the increase in the alkyl chain length and increased with increasing temperature,^{27,32} which is consistent with the description in the literature³³ and the results for other IL solutions.²⁵ This can be interpreted as elevated temperature causes the disruption of the water structure leading to low solubility of the surfactant. Besides, the elevated



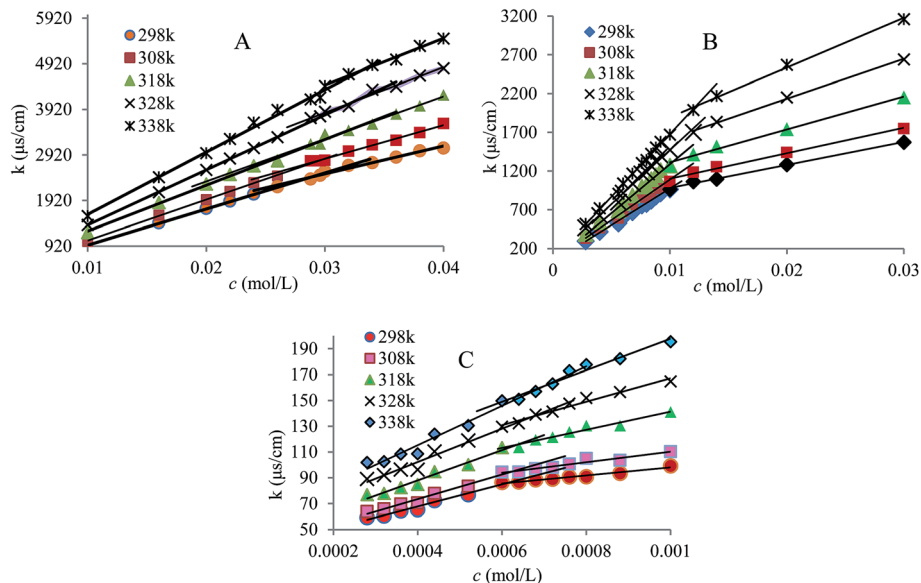


Fig. 2 Conductivity of pure C_{10} mimBr (A), C_{12} mimBr (B) and C_{16} mimBr (C) in aqueous solutions at different temperatures.

temperature causes difficulties in forming the micelles, mainly due to the hydration of the ion being greatly enhanced, which significantly weakens the electrostatic interaction of micellization.

In addition, the CMCs of the C_n mimBr/CTAB aqueous systems decline with increasing CTAB concentration at the

same temperature. Although the positive charges on the head group of CTAB and the imidazolium ring of the IL repel each other, in fact, electrostatic attraction effect between the counterion of the IL, Br^- , and the head group of the CTAB molecules may reinforce micelle formation to make the CMC value low or the surface activity high. At the same time, the hydrophobic

Table 1 Thermodynamic functions of C_n mimBr in the absence of CTAB (conductivity method)^a

T/K	C_{10} mimBr			C_{12} mimBr			C_{16} mimBr		
	CTAB (mM)	α	CMC (mM)	CTAB (mM)	α	CMC (mM)	CTAB (mM)	α	CMC (mM)
298.15	0	0.321	28.9, 29.3 ^b	0	0.321	10.1, 10.9 ^b	0	0.316	0.611, 0.55 ^c
308.15		0.417	29.8		0.415	10.2		0.406	0.626
318.15		0.481	31.4		0.483	10.3		0.478	0.637
328.15		0.504	32.4		0.509	11.6		0.499	0.673
338.15		0.511	33.1		0.514	12.2		0.507	0.703
298.15	4	0.86	28.1	0.2	0.362	9.6	0.04	0.322	0.561
308.15		0.825	29.4		0.437	10.2		0.354	0.575
318.15		0.872	30.8		0.481	10.6		0.441	0.605
328.15		0.859	31.4		0.523	10.9		0.549	0.641
338.15		0.881	31.7		0.547	11.4		0.721	0.652
298.15	8	0.971	26.6	0.4	0.520	9.4	0.08	0.347	0.545
308.15		0.914	27.1		0.568	9.9		0.461	0.556
318.15		0.884	28.2		0.617	10.4		0.502	0.583
328.15		0.823	29.6		0.624	10.8		0.528	0.625
338.15		0.927	30.6		0.657	11.2		0.661	0.641
298.15	12	0.823	25.7	0.8	0.605	9.3	0.12	0.376	0.534
308.15		0.865	26.4		0.642	9.8		0.524	0.544
318.15		0.906	27.3		0.680	10.1		0.418	0.558
328.15		0.964	28.5		0.715	10.5		0.498	0.574
338.15		0.819	30.3		0.744	10.8		0.601	0.601
298.15	16	0.747	24.2	1.2	0.662	8.7	0.16	0.407	0.529
308.15		0.796	25.1		0.689	9.1		0.351	0.533
318.15		0.828	25.3		0.699	9.3		0.443	0.55
328.15		0.827	25.4		0.716	10.3		0.495	0.568
338.15		0.869	25.8		0.735	10.7		0.587	0.587

^a Standard uncertainties s are $s(\text{CMC}) = \pm 0.03$ (mM). ^b The value is from the literature.²⁵ ^c The value is from the literature.⁴



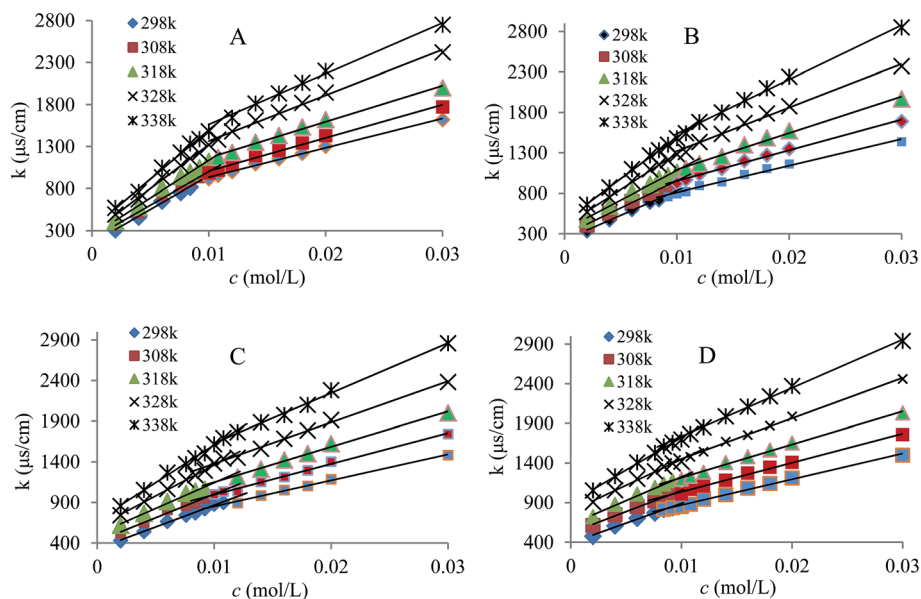


Fig. 3 Conductivity of the mixed system of $C_{12}\text{mimBr}/\text{CTAB}$ at different temperatures. c_{CTAB} (A–D): 0.2, 0.4, 0.8 and 1.2 mM.

interaction between the tail of the CTAB and the alkyl chain of the IL cation promote the procession of micellization. This finding differs from the CTAB/bmim-octyl SO_4 system reported by other authors,⁹ who thought the formation of the pseudo-double chain cationic surfactant (CTA-octyl SO_4) led to a much lower critical aggregation concentration (CAC) and surface tension value at the CAC for the mixtures than for the individual components, whereas similar results in CTAB/SDS system have been obtained showing that the electrostatic repulsion of the head groups decreases and consequently the CMC is reduced when SDS gets into the solution.³⁴ This behavior is also similar to the effect of electrolytes such as sodium chloride (NaCl) when added to surfactant solutions. Liu *et al.* found that the CMC values of $C_{12}\text{mimBr}$ aqueous solution decreased with increasing NaBr concentration, and illustrated that the phenomenon arises from the reduction of electrostatic repulsion between the surfactant head groups caused by the additional ions in the solution.¹⁵ Additionally, the degree of dissociation (α) decreased with the increase in the alkyl chain length. The small value of α in general indicates that the aggregate structures are compact for $C_n\text{mimBr}/\text{CTAB}$ aqueous systems in water.

3.1.2 Thermodynamic functions. The temperature dependence of the CMC has been employed to compute the thermodynamic parameters of micellization in order to obtain better knowledge about the behavior of the surfactant and intermolecular interactions present in such systems.³⁵ The thermodynamic parameters of the $C_n\text{mimBr}/\text{CTAB}$ mixed system at different CTAB concentrations are shown in Fig. 4. The standard free energy change of the micelle formation (ΔG_m) revealed a rather slight decrease with the increment of temperature in mixed systems. There is no doubt that the negative values of ΔG_m indicated the spontaneous micellization. The more negative value of ΔG_m with the increase in the length of the alkyl chain is attributed to the fact that the longer alkyl chains facilitate the

spontaneous aggregate formation. Furthermore, the negative ΔH_m implies that micellization formation remains mainly an exothermic process, which is owing to greater dissociation of counter-ions from the micellar surface leading to a more exothermic micellization process. However, ΔH_m firstly increased and then declined with temperature limits of 298.15–338.15 K with CTAB concentrations of 0.2, 0.4, 0.8 and 1.2 mM, as can be seen from Fig. 4.

Meanwhile, entropy change ($-T\Delta S_m$) was plotted against temperature to better understand the contribution of entropy to the free energy (Fig. 4). It is noteworthy that the trend in the change of $-T\Delta S_m$ with temperature in the mixed solution is analogous to ΔH_m . It can be seen from Fig. 4A–D that $-T\Delta S_m$ (ΔS_m) gradually decreases (increases) with temperature at the same low CTAB concentration of 0.2 and 0.4 mM, then

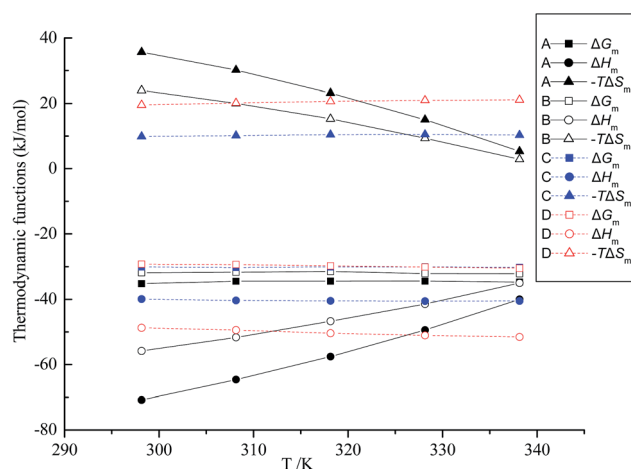


Fig. 4 Thermodynamic functions (kJ mol^{-1}) of $C_{12}\text{mimBr}/\text{CTAB}$ system against temperature. c_{CTAB} (A–D): 0.2, 0.4, 0.8 and 1.2 mM.



substantially keeps flat at 0.8 mM, and finally increases when reaching 1.2 mM. Obviously, this change may be dependent on the presence of the CTAB molecules in Fig. 4A–D with the same IL concentration.¹⁷ For CTAB monomer ($c_{\text{CTAB}} < 0.9$ mM) added into IL solution, the disruption of the water structure by the interaction between IL molecules and CTAB monomer makes the arrangement of the mixture system less and less ordered. However, ΔS_m decreases with temperature when CTAB micelles ($c_{\text{CTAB}} > 0.9$ mM) are added into IL solution. The essence of this phenomenon is that the higher temperature makes it more difficult to dissolve nonpolar substances in water, leading to enhancement of hydrophobic interactions between micelle molecules. In other words, hydrophobic alkyl chains come closer to each other in the micellar phase. As a result, the molecular arrangement is more ordered so that entropy is reduced. According to the data in Table 1, the CMCs of the two surfactant mixtures at elevated temperatures were lower than each surfactant, which also gives evidence of the strong synergism produced in mixed aggregates.³⁴ In general, Fig. 4 shows that the contribution of enthalpy term ΔH_m to ΔG_m is superior to $(-T\Delta S_m)$ at all studied temperatures. Therefore, it can also be concluded that the micelle formation of the CTAB/ C_n mimBr mixture is an enthalpy-driven process. The critical micelle concentration (CMC) determined from conductivity data was also found in the study of the CTAB/DTAB-glycyl dipeptide system,³⁶ which revealed a decrease in the CMC values for both the surfactants with [glycyl dipeptide] and demonstrated the entropy–enthalpy contribution to the micellization process.

3.2 Surface tension measurement

3.2.1 CMC determination. Surface tension measurements were performed at different temperatures. Based on the plots of surface tension *versus* ILs concentration depicted in Fig. 5–7, several different parameters related to the surface activities are listed in Table 2 wherein the CMC value of the pure ILs are from the previous work.³⁷ It was obvious that the surface tensions of the mixed systems in aqueous media exhibited a steep decrease

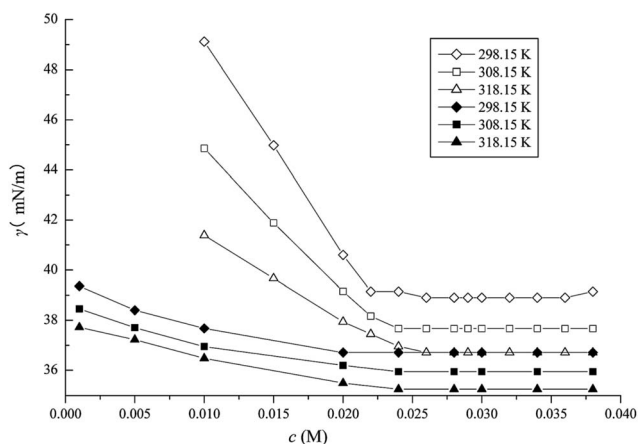


Fig. 5 Surface tension of C_{10} mimBr in the presence of CTAB (\diamond 14 mM, 298.15 K), (\square 14 mM, 308.15 K), (\triangle 14 mM, 318.15 K), (\blacklozenge 30 mM, 298.15 K), (\blacksquare 30 mM, 308.15 K) and (\blacktriangle 30 mM, 318.15 K).

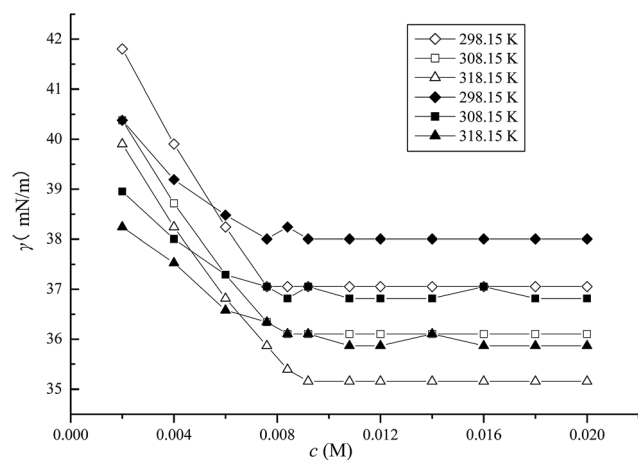


Fig. 6 Surface tension of C_{12} mimBr in the presence of CTAB (\diamond 2 mM, 298.15 K), (\square 2 mM, 308.15 K), (\triangle 2 mM, 318.15 K), (\blacklozenge 8 mM, 298.15 K), (\blacksquare 8 mM, 308.15 K) and (\blacktriangle 8 mM, 318.15 K).

with an increase of IL concentration up to a certain concentration and subsequently remained constant. The CMC value was obtained by the intersection between the rapidly decreasing line and the flat line of surface tension. As shown in Fig. 5–7, in the case of CTAB at 14 mM or 30 mM in the C_{10} mimBr/CTAB system, a dramatic decrease of the surface tension was observed compared with a moderate decrease at 2 mM or 8 mM CTAB in the C_{12} mimBr/CTAB system, and 0.02 mM or 0.12 mM in the C_{16} mimBr/CTAB system. These trends in the CMC value associated to temperature and CTAB concentration endorse the results drawn from the conductivity method. Chauhan *et al.* applied surface tension to investigate CM(A)C of a ternary (water + lactose + NaC/NaDC) system.²⁶ The decrease of the CM(A)C values of bile salts in aqueous lactose solution was found to be in excellent agreement with the mixture systems in this work. Moreover, they employed conductivity and UV-vis probe to examine the interactions of the ionic liquid

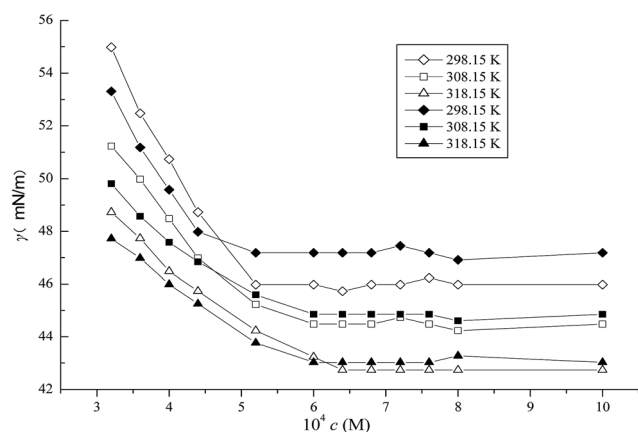


Fig. 7 Surface tension of C_{16} mimBr in the presence of CTAB (\diamond 0.02 mM, 298.15 K), (\square 0.02 mM, 308.15 K), (\triangle 0.02 mM, 318.15 K), (\blacklozenge 0.12 mM, 298.15 K), (\blacksquare 0.12 mM, 308.15 K) and (\blacktriangle 0.12 mM, 318.15 K).



Table 2 Thermodynamic functions of C_n mimBr in the absence of CTAB (surface tension method)^a

CTAB (mM)	<i>T</i> (K)	CMC (mM)	ΔG_m (kJ mol ⁻¹)	ΔH_m (kJ mol ⁻¹)	$-T\Delta S_m$ (kJ mol ⁻¹)	Γ_{max} (μmol m ⁻²)	A_{min} (Å ²)	γ_{CMC} (mN m ⁻¹)	Π_{CMC} (mN m ⁻¹)	pC_{20}
C₁₀mimBr										
0	298.15	27.24	-37.44	-8.28	-29.15	1.75	94.61	48.95	23.02	1.61
	308.15	30.78	-38.42	-3.93	-34.49	1.68	98.92	45.15	26.82	1.70
	318.15	31.60	-39.53	0.15	-39.68	1.57	105.8	41.29	30.68	1.63
14	298.15	22.32	-38.77	-9.22	-29.55	105.53	38.50	33.47	38.50	105.53
	308.15	24.25	-39.65	-7.04	-32.60	113.43	37.17	33.21	38.76	113.43
	318.15	25.85	-40.59	-5.00	-35.59	121.99	36.96	31.78	40.19	121.99
30	298.15	19.89	-39.34	-18.42	-20.92	108.31	36.02	35.95	36.02	108.31
	308.15	22.55	-40.02	-11.89	-28.13	114.96	36.00	34.38	37.59	114.96
	318.15	24.38	-40.91	-5.77	-35.13	121.99	34.68	34.06	37.91	121.99
C₁₂mimBr										
0	298.15	9.19	-43.17	-35.96	-7.21	1.80	92.49	46.71	25.26	2.35
	308.15	10.23	-44.07	-31.60	-12.46	1.78	93.49	43.52	28.45	2.29
	318.15	11.06	-45.08	-27.52	-17.56	1.70	97.59	39.36	32.61	2.26
2	298.15	9.07	-43.23	-34.71	-8.52	1.82	91.46	38.02	33.95	2.86
	308.15	9.80	-44.29	-30.36	-13.93	1.80	92.47	37.04	34.93	2.86
	318.15	10.31	-45.45	-26.28	-19.18	1.68	98.69	36.72	35.25	2.83
8	298.15	8.90	-43.33	-32.39	-10.94	1.86	89.47	37.41	34.56	2.92
	308.15	9.51	-44.44	-28.03	-16.41	1.83	90.50	34.78	37.19	2.92
	318.15	9.90	-45.67	-23.95	-21.72	1.81	91.49	34.53	37.44	2.90
C₁₆mimBr										
0	298.15	0.610	-56.61	-33.36	-23.24	2.00	83.14	44.53	27.44	3.42
	308.15	0.631	-58.34	-29.01	-29.33	1.83	90.50	42.29	29.68	3.40
	318.15	0.637	-60.18	-24.93	-35.25	1.61	103.33	41.74	30.23	3.34
0.02	298.15	0.601	-56.69	-25.33	-31.36	2.22	74.83	42.64	29.33	3.48
	308.15	0.624	-58.40	-20.97	-37.42	2.05	81.02	41.25	30.72	3.40
	318.15	0.632	-60.22	-16.89	-43.33	1.76	94.44	40.39	31.58	3.34
0.12	298.15	0.597	-56.72	-17.03	-39.69	2.24	91.51	40.15	31.82	3.83
	308.15	0.621	-58.42	-12.68	-45.74	2.12	96.67	38.71	33.26	3.62
	318.15	0.629	-60.25	-8.60	-51.65	1.87	102.13	35.45	36.52	3.47

^a Standard uncertainties *s* are $s(\text{CMC}) = \pm 0.03$ (mM), $s(\Delta G_m) = \pm 0.03$ (kJ mol⁻¹), $s(\Delta H_m) = \pm 0.02$ (kJ mol⁻¹), $s(\Delta S_m) = \pm 0.02$ (J mol⁻¹ K⁻¹).

tetrabutylammonium dodecylsulfate (TBADS) in aqueous solutions of glycine with sodium dodecylsulfate (SDS).²²

3.2.2 Surface properties of IL-surfactant mixtures. The molecules of the IL-surfactant mixture adsorbed at the interface, even at low surfactant concentration, cause the reduction of the surface tension of water. Possible interactions with the surfactant adsorption are hydrophobic effect, electrostatic interaction, van der Waals interaction, hydrogen bonding,^{12,14,38,39} etc.

It is known that the hydrophobicity of surfactants increases with increasing hydrophobic chain length. Thus the surfactant monomers are more likely to escape the solvent into the air-liquid interface, so that the hydrophilic head group in the water is pulled by the hydrophobic hydrocarbon chain away from the solvent. Therefore, the parameter of Π_{CMC} in Table 2 reflects that the surface tension of the C₁₆mimBr/CTAB mixture is reduced more than that of C₁₂mimBr/CTAB when CTAB monomers get into solution. While in the C₁₀mimBr/CTAB system, the CTAB micelles almost have no effect in reducing the surface tension of water according to the Π_{CMC} listed in Table 2.

The adsorption efficiency (pC_{20}) of the surfactant molecules at the air-solution interface can be obtained in terms of the

surface tension data. For C_n mimBr/CTAB mixtures, it can be easily observed that pC_{20} values increase when increasing the CTAB content, showing superior adsorption efficiency.¹⁴ However, pC_{20} decreasing with the increment of temperature suggests that elevated temperature is not favorable for micellization, corresponding to CMC trends. It is noted that the C₁₀mimBr/CTAB mixtures exhibit increased pC_{20} values in the case of CTAB micelles, different from the decrease in the C₁₆mimBr/CTAB mixtures with CTAB monomer added into the solution. So the self-assembled behaviors of surfactant molecules in mixture solution are influenced by the CTAB molecule state at different temperatures.

Additionally, the maximum surface excess concentration (Γ_{max}) is close to pC_{20} , and the minimum area of a single surfactant molecule (A_{min}) is also estimated in Table 1. Like the temperature effect of pC_{20} mentioned above, it is shown that the higher the temperature, the greater the value of A_{min} in the presence of CTAB monomer in solution, while A_{min} gradually decreases with temperature when CTAB micelles are present in solution. By comparing the data for the three ILs with the same counter ion, it was found that A_{min} per adsorbed ionic liquid at the air/water interface decreases with the length of the alkyl



chain and it implies that the longer the alkyl chains, the more compact the monolayers at the air/water interface.⁴⁰ Certainly, the values of Γ_{\max} have the corresponding opposite trend.¹⁴ All these parameter changes can be thought to be related to enhanced molecular motion at elevated temperature, thus leading to fewer surfactant molecules being adsorbed at the air-water interface.¹ With an excess of CTAB, more CTAB molecules may participate in surface adsorption of mixture solution, besides self-assembly behavior, for example in the C₁₀mimBr/CTAB mixtures.

3.2.3 Thermodynamic functions of micellization. By applying eqn (8)–(10), the standard free energy change (ΔG_m), standard enthalpy (ΔH_m) and standard entropy change (ΔS_m) associated with the micelle formation were also calculated. In the case of ionic surfactants, according to eqn (5)–(8), the standard free energy change as well as other thermodynamic functions of aggregate formation at different temperatures may be calculated using X_{CMC} ,⁴ because the measurement of the degree of dissociation (α) may be omitted. Using eqn (8)–(10), the thermodynamic parameters of the CTAB with C_nmimBr mixed system at different CTAB concentrations are given in Table 2. Comparison to the results from the conductivity measurement (Fig. 4) showed that three thermodynamic parameters with temperature and CTAB concentration displayed a similar change, and here an entropy-driven process can also be clearly observed from the data in Table 2, which indicates a CMC value a little smaller than that determined by the conductivity measurement. The difference in the two results is mainly based on applying different eqn (5) and (8).

$$\Delta G_m = 2RT \ln X_{\text{CMC}} \quad (8)$$

$$\Delta H_m = d\left(\frac{\Delta G_m}{T}\right) / d\left(\frac{1}{T}\right) \quad (9)$$

$$\Delta S_m = \frac{\Delta H_m - \Delta G_m}{T} \quad (10)$$

For the conductivity of the mixture solution, the degree of counter-ion dissociation implies that a fraction of the counterions on the micelle surface are confined, and the formed micelles can also carry less charge to transport than the free ions.⁴ Consequently, the α value should change with the number and state of the ionic surfactants in the mixture solution, so it is not a constant “2” in eqn (8), which is an extreme case of $\alpha \approx 0$. Theoretically, eqn (5) is more reasonable and accordingly thermodynamic functions shown in Fig. 4 seem to be more convincing.

3.3 FT-IR analysis

According to experimental method mentioned above, FT-IR spectra of C₁₀mimBr, CTAB and their mixture studied in the present work were recorded, as shown in Fig. S3.† The FT-IR spectra of the mixture showed the apparent change with comparison to the pure components. The C–H stretching (C₄H, C₅H and C₂H) of the imidazole ring of IL can be observed at 3050–3200 cm⁻¹,^{3,41} but becomes weaker after interaction with CTAB in the mixture. It is worth noting that a band located at 1572 cm⁻¹ (C–N stretching) for the IL was weaker in the mixture than in the pure IL. Moreover, the characteristic peak of C–H plane deformation vibration in the imidazole ring significantly moves from 1174 cm⁻¹ to 1168 cm⁻¹. This is evidence of the electrostatic effect between the head groups of IL and CTAB molecules. On the other hand, the band between 2800 and

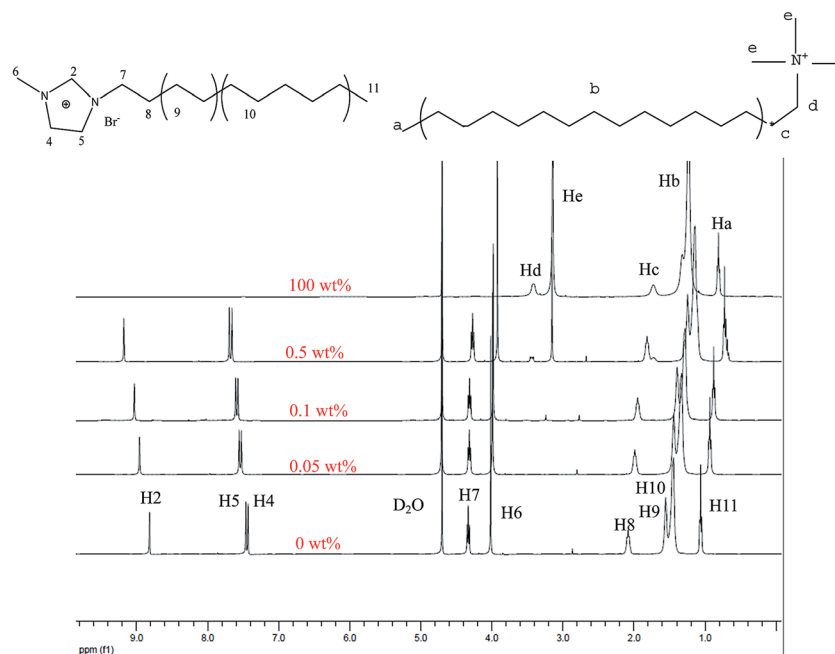


Fig. 8 ¹H NMR spectra of C₁₀mimBr in D₂O solution with different wt% of CTAB.



3050 cm^{-1} assigned to the asymmetric and symmetric C–H stretching vibration of the methyl groups was seen in $\text{C}_{10}\text{mimBr}$, CTAB and their mixture spectra,⁴² but the slight change corresponds to the hydrophobic interactions between the hydrocarbon chain of the surfactant CTAB and IL cation.

3.4 ^1H NMR

The ^1H NMR spectra were recorded to understand the various intermolecular interactions present among IL and CTAB molecules in aqueous solutions. In this part, the effect of CTAB on the NMR microenvironment of the C_nmimBr was investigated and all protons of C_nmimBr in the mixed system were analyzed. The ^1H NMR spectrum of the IL in Fig. 8 reveals nine characteristic peaks. It can be seen that the characteristic proton signals of Ha–He from CTAB gradually become higher and higher with increasing CTAB. It is well known that the chemical shifts of IL protons move upfield or downfield depend on the shielding and deshielding effects of surrounding groups. The greater the electronic density, the greater the shielding effect. Owing to addition of CTAB, all upfield shifts in the tail chain of the IL proton signals of H7–11 reflect the hydrophobic interactions present among the hydrophobic chain of the surfactant and alkyl chains of the IL cation, which led to shielding of the surrounding protons.¹²

In addition, the evident downfield shift of the head protons (H2, 4, 5) of the IL was observed. This may be interpreted that after the addition of CTAB, increased counterion dissociation from CTAB into solution leads to a stronger electrostatic attractive effect between the Br^- and the imidazolium cation than electrostatic repulsion, thus the lower electronic density of the imidazolium cation results in the signals of H2, 4, and 5 shifting toward lower field.²⁷ However, an upfield shift was also observed for protons of the imidazole ring (H6), but not significant in comparison with the tail region protons (H8–11) because the magnetic environment of proton H6 located at the micellar surface would not be changed so remarkably after addition of CTAB.

4 Conclusion

The present study presents the surface and thermodynamic properties of the $\text{C}_n\text{mimBr}/\text{CTAB}$ surfactants system in aqueous media using surface tension and conductivity methods. For all the systems, the surface activity and CMC values of the $\text{C}_n\text{mimBr}/\text{CTAB}$ ($n = 10, 12, 16$) systems are highly dependent on the CTAB concentration and temperature. The CMC values systematically decreased with the CTAB concentration and an inverse trend was observed when temperature was increased; such dependence is analogous with that of conventional ionic surfactants. It was further perceived that adsorption proficiency decreased at high CTAB monomer concentrations, while the opposite result was also observed in the presence of CTAB micelles. On the other hand, the minimum area per adsorbed molecule decreased sharply with CTAB concentration, indicating that compact mono-layers are formed at the air/water interface and facilitate the formation of well-defined

aggregates in the bulk. The value of surface excess concentration (Γ_{max}) implies that the adsorption of IL in the mixture system decreases with increasing CTAB content, while opposite trends were observed in other surfactant systems. For all the mixed systems, the negative ΔG_{m} presented a decrease with temperature and CTAB concentration, suggesting that all micellization processes were spontaneous. The ΔH_{m} and ΔS_{m} values for all the mixed systems were dependent on the CTAB molecule state, as the CTAB monomer and micelle led to an opposite change trend of ΔH_{m} and ΔS_{m} for the mixed system. FT-IR and NMR data provide evidence of the interaction between the IL and CTAB molecules in the mixture system to best understand the cause of the final results. By comparison of the traditional surfactants with CTAB addition into ionic liquid surfactants to form a surfactants mixture, it is important to reveal the higher surface activity. Therefore, similar to classical cationic surfactants, the active surface of such a surfactant mixture forms well defined aggregates at the air/water interface in bulk solution and has potential for the development of novel green surface-active agents. It is clear that the present investigations provide a fundamental understanding of the solution properties of the binary system. Our work herein will also be useful for designing separation technologies and would be further useful in developing new technologies for applying ILs to colloid and interfacial science chemical engineering fields as well.

Conflicts of interest

All authors declare no financial/commercial conflicts of interest.

Acknowledgements

The authors are grateful for the open project of Jiangsu Key Laboratory for Bioresources of Saline Solis (JKLBS2016012), and Jiangsu Province College Students' Innovative Training Project (201610324044x) and a Project Funded by the Excellent Specialties Program Development of Jiangsu Higher Education Institutions, Project (No. PPZY2015B113).

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