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Cucurbit[7]uril complexations of Good's buffers†

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The cucurbit[7]uril (CB[7]) host–guest complexations of a series of zwitterionic “Good’s” biological pH buffers have been investigated in aqueous solution by means of ^1H NMR spectroscopy. The cyclohexylammonium buffers bind very strongly ($K_{\text{CB}[7]} = 10^7$ to $10^8 \text{ dm}^3 \text{ mol}^{-1}$), while the morpholinium (10^2 to $10^3 \text{ dm}^3 \text{ mol}^{-1}$) and piperazinium (10^3 to $10^4 \text{ dm}^3 \text{ mol}^{-1}$) buffers have binding constants several orders of magnitude smaller. The binding constants increase as the distance between the ammonium and sulfonate groups increases. The $\text{p}K_{\text{a}}$ of 2-(cyclohexylammonio)ethanesulfonate (CHES) increases by 3.1 units upon complexation by CB[7].

Introduction

In 1966, N. E. Good proposed criteria for an optimum biological buffer,¹ and his group^{2,3} and others⁴ reported the syntheses of several series of zwitterionic compounds (Fig. 1) which are now standard buffers employed in biochemistry. The criteria included easy synthetic preparation (with a resistance to degradation), $\text{p}K_{\text{a}}$ values between 6 and 8 (with a minimal dependence on temperature, concentration, or ionic strength), water solubility, impermeability to biological membranes, and a lack of binding to other biological cations. The majority of the Good’s buffers bear sulfonate groups, which are anionic at physiological pH and enhance the water solubility of the buffer, and tertiary amine groups, whose protonated forms have $\text{p}K_{\text{a}}$ values near to or above 7.

As a result of their zwitterionic nature, the Good’s buffers have recently been used as the anionic or cationic components of ionic liquids (GBILs), for use as self-buffering and biocompatible media for protein and antibody extractions.^{5–10} The Good’s buffers have also been employed to stabilize platinum and gold nanoparticles, affecting their structures and optical properties.^{11–16}

Despite the proposed criteria, a number of the buffers result in metal ion depletion (notably copper)^{17–23} and interactions with biological systems, leading to unwanted or beneficial physiological effects.^{24,25} The former problem has been remedied to some degree by “Better” pH buffers such as PIPES (Fig. 1), which are non-complexing towards metal ions.^{26,27} There has, however, been little in the way of studies on the interactions of these pH buffers with macrocyclic host molecules,^{28–30} such as cyclodextrins, calixarenes, and cucurbiturils,

which are of increasing interest in biological, biochemical, and medicinal applications.^{31–34}

In this study, the host–guest complexations of a series of Good’s (and related) pH buffers (Fig. 1) by the macrocyclic host molecule cucurbit[7]uril (CB[7]) have been studied using ^1H NMR spectroscopy. The cucurbit[n]urils (CB[n], $n = 5–8, 10$ and

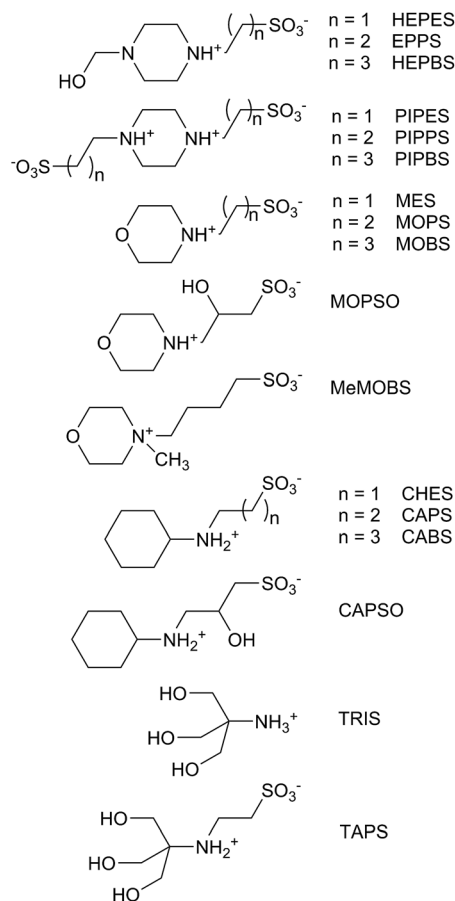


Fig. 1 Good’s type biological pH buffers used in this study.

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† Electronic supplementary information (ESI) available: ^1H NMR spectra of guest and host–guest complexes, ^1H NMR titration plots, and limiting chemical shift values. See DOI: 10.1039/c7ra08865e



13–15) are a family of host molecules comprised of cyclic oligomers of n glycoluril units bridged by $2n$ methylene groups.^{34–36} The CB[n] hosts form very stable ($K_{\text{CB}[n]}$ up to $10^{17} \text{ dm}^3 \text{ mol}^{-1}$)³⁷ host–guest complexes with organic cations in aqueous solution, utilizing dipole–dipole and ion–dipole interactions at the polar portals and hydrophobic interaction in the cavity. One feature of the binding of basic organic guests is the preferred complexation of their conjugate acid, resulting in an increase in the $\text{p}K_{\text{a}}$ value.^{38–41}

Cucurbiturils have recently gained interest in biological, biochemical and medicinal applications, as studies have suggested little evidence of toxicity.^{42–45} For pH control in these cucurbituril host–guest studies, acetate and phosphate buffers are most commonly used, although TRIS (tris(hydroxymethyl) methylamine) has been used in a few instances,^{46–49} including early studies of CB[7] with the methyl viologen dication (MV^{2+}).^{46,47} Ong and Kaifer⁴⁷ noted that binding constant for MV^{2+} , measured in 30 mmol dm^{-3} TRIS, was about 10% higher than previously reported in 50 mmol dm^{-3} TRIS,⁴⁵ and suggested that the buffer can compete for the host, as found for alkali metal and alkali earth cations.^{50–52}

The cucurbit[7]uril complexations of the four series of Good's buffers, based on cyclohexylammonium, piperazinium, N -2-hydroxyethylpiperazinium, morpholinium, and tris(hydroxymethyl)methylammonium, with alkanesulfonate group substituents, have been investigated using ^1H NMR spectroscopy. The host–guest stability constants and the complexation induced guest proton chemical shift values have been determined and related to the nature of cationic portion of the buffer and the length of the alkane chain separating it from the sulfonate group(s). The effect of CB[7] complexation of CHES on its $\text{p}K_{\text{a}}$ value is also reported.

Experimental

Materials

The cucurbit[7]uril was prepared and characterized by the method of Day *et al.*⁵³ The majority of the buffers were used as received (Sigma-Aldrich or GFS Chemicals (for PIPBS)). The MOBS, MeMOBS and CABS compounds were prepared by the method of Yi *et al.*,⁵⁴ employing 1,4-butane sultone with morpholine, N -methylmorpholine, or cyclohexylamine, respectively. The PIPBS buffer was prepared as described by Jermyn.⁵⁵

MOBS. 60% yield. ^1H NMR (300 MHz, D_2O) δ 4.01 (br, 2H), 3.75 (br, 2H), 3.44 (br, 2H), 3.14 (t, 2H, $J = 7.3$ ppm, overlapping br peak, 2H), 2.88 (t, 2H, 7.3 Hz), 1.90–1.65 (m, 6H) ppm. ^{13}C NMR (75 MHz, D_2O) δ 63.85, 56.55, 51.60, 49.95, 21.92, 21.19 ppm.

MeMOBS. 43% yield. ^1H NMR (300 MHz, D_2O) δ 3.97 (br, 4H), 3.42 (m, 8H), 3.11 (s, 3H), 2.90 (t, 2H, $J = 7.5$ Hz), 1.88 (m, 2H), 1.75 (qn, 2H, $J = 7.1$ Hz) ppm. ^{13}C NMR (75 MHz, D_2O) δ 64.33, 60.31, 59.59, 49.91, 46.79, 21.04, 19.85 ppm.

CABS. 48% yield. ^1H NMR (300 MHz, D_2O) δ 3.00 (m, 3H), 2.87 (m, 2H), 1.98 (br, 2H), 1.74 (br, 2H), 1.73 (m, 4H), 1.58 (d, $J = 12.3$ Hz, 1H), 1.23 (qn, $J = 12.3$ Hz, 4H), 1.11 (m, 1H) ppm. ^{13}C NMR (75 MHz, D_2O) δ 57.09, 50.07, 43.78, 28.86, 24.62, 23.89, 21.33 ppm.

PIPBS. 58% yield. ^1H NMR (300 MHz, D_2O) δ 3.66 (br, 8H), 3.25 (t, 4H, $J = 7.9$ Hz), 2.87 (t, 4H, $J = 7.3$ Hz), 1.86 (qn, 4H, $J = 7.9$ Hz), 1.75 (qn, 4H, $J = 7.3$ Hz) ppm. ^{13}C NMR (75 MHz, $\text{D}_2\text{O}/\text{DCl}$) δ 56.07, 49.35, 43.77, 29.33, 21.58 ppm.

Methods

The 1D and 2D ^1H and ^{13}C NMR spectra were recorded on Bruker Avance 300 and 400 instruments. The host–guest binding constants were determined from ^1H NMR titrations or by ^1H NMR competitive binding experiments in D_2O at $\text{pD} = 4.75$ ($0.050 \text{ mol dm}^{-3}$ NaOAc- d_3 /0.025 mol dm^{-3} DCl). The competitor guest used was 3-(trimethylsilyl)propionic-2,2,3,3- d_4 acid (TSP),⁵⁶ with a reported CB[7] binding constant of $(1.82 \pm 0.22) \times 10^7 \text{ dm}^3 \text{ mol}^{-1}$. The stability constants for weaker CB[7] complexes were determined from either non-linear least squares fitting of ^1H NMR chemical shift titrations or from Benesi–Hildebrand plots ($1/\Delta\delta$ vs. $1/\text{CB}[7]$).⁵⁷ For the PIPBS buffer, slow exchange behaviour is observed in the NMR titrations, and the value of $K_{\text{CB}[7]}$ was calculated by integrating the free and bound methine proton resonances of CB[7]. The limiting chemical shift changes, for the guest protons on weaker binding guests, were determined by extrapolation in fitting the chemical shift titrations.

Results and discussion

The ^1H NMR titrations of the Good's buffers with CB[7] (Fig. 2) provide a measurement of the host–guest stability constants (Table 1) and the limiting complexation-induced chemical shift changes for the guest proton resonances ($\Delta\delta_{\text{lim}} = \delta_{\text{bound}} - \delta_{\text{free}}$) give an indication of the average position of the guest within the

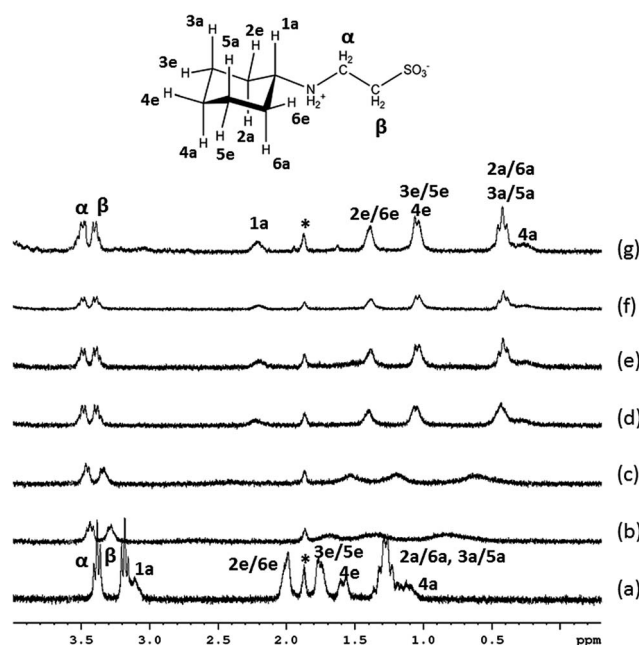


Fig. 2 ^1H NMR titration of CHES with (a) 0.00, (b) 0.57, (c) 0.80, (d) 0.96, (e) 1.05, (f) 1.73, and (g) 4.20 equivalents of CB[7] in D_2O ($\text{pD} = 4.75$, $0.050 \text{ mol dm}^{-3}$ NaOAc- d_3 (*)/0.025 mol dm^{-3} DCl).



Table 1 Host–guest stability constants for CB[7] with the Good's buffers in D₂O (pD = 4.75, 0.050 mol dm⁻³ NaOAc-*d*₃/0.025 mol dm⁻³ DCl)

Buffer	p <i>K</i> _a	Δ <i>δ</i> _{lim} (<i>H</i> _α), ppm	<i>K</i> _{CB[7]} , dm ³ mol ⁻¹
HEPES	7.55 ^a	+0.43	(2.3 ± 0.3) × 10 ³
EPPS	8.0 ^b	+0.27	(8.9 ± 1.3) × 10 ³
HEPBS	8.3 ^b	+0.04	(2.1 ± 0.4) × 10 ⁴
PIPES	2.67, 6.78 ^c	0	0
PIPPS	3.79, 7.97 ^c	0	0
PIPBS	4.29, 8.55 ^c	+0.15	(2.2 ± 0.6) × 10 ³
MES	6.06 ^c	-0.02	210 ± 30
MOPSO	6.95 ^d	-0.19, -0.32	330 ± 50
MOPS	7.09 ^c	-0.25	(2.2 ± 0.3) × 10 ³
MOBS	7.48 ^c	-0.40	(2.4 ± 0.3) × 10 ³
MeMOBS		-1.05	(4.2 ± 0.6) × 10 ⁴
CHES	9.27 ^e	+0.10	(3.6 ± 0.5) × 10 ⁷
CAPSO	9.6 ^f	+0.02, +0.13	(6.0 ± 0.9) × 10 ⁷
CAPS	10.35 ^e	-0.02	(1.0 ± 0.2) × 10 ⁸
CABS	10.7 ^b	+0.03	(1.3 ± 0.2) × 10 ⁸
TRIS	8.06 ^a		290 ± 50
TAPS	8.4 ^g	-0.01	110 ± 20

^a Ref. 58. ^b Ref. 4. ^c Ref. 27. ^d Ref. 3. ^e Ref. 59. ^f Ref. 60. ^g Ref. 61.

CB[7] cavity (Fig. 3, S1, S9, S20†). Upfield shifts (Δ*δ*_{lim} < 0 ppm) for the proton resonances indicate positioning within the cavity, while downfield shifts are indicative of the protons located outside of the cavity near the polar portal(s) of CB[7]. In the buffers which bind to the CB[7] in this study, the values of Δ*δ*_{lim} (Fig. 3, S1, S9, S20†) indicate that the host is binding primarily over the ring (and the hydroxyethyl group for HEPES, EPPS, and HEPBS, Fig. 3 and S20†), rather than the alkylsulfonate chain.

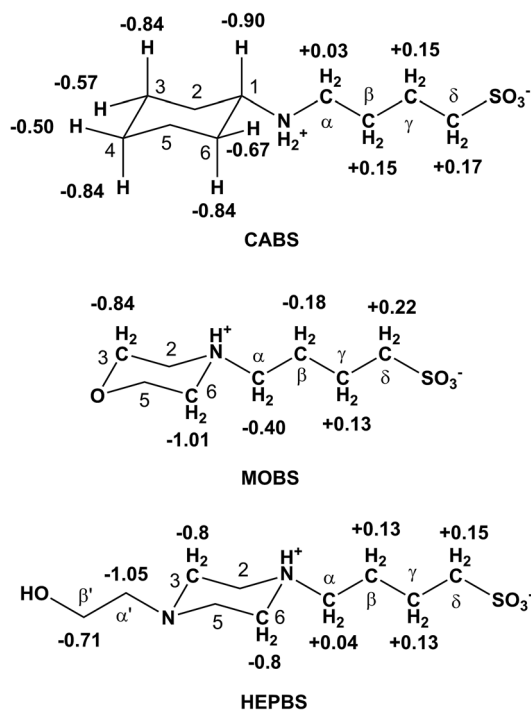


Fig. 3 CB[7] complexation-induced chemical shift changes, Δ*δ*_{lim} (ppm), for representative buffers.

The strongest binding by CB[7] was observed for the cyclohexylammonium buffers (Fig. 2, S1–S8†), with stability constants in the range of 10⁷ to 10⁸ dm³ mol⁻¹ (Table 1). It is known that the cyclohexylammonium cation binds strongly to cucurbit[7]uril, as a result of the hydrophobic interactions of the cyclohexyl ring with the cavity of CB[7] and the ion-dipole interactions of the ammonium group at the carbonyl portal. Values of *K*_{CB[7]} = (2.4 ± 0.4) × 10⁷ dm³ mol⁻¹ (0.10 mol dm⁻³ Na₃PO₄/DCl, pH = 7.4) and 1.1 × 10⁸ dm³ mol⁻¹ (0.050 mol dm⁻³ NaOAc, pH 4.0) been reported by Cao and Isaacs,⁶² and Li and Kaifer,⁶³ respectively. Yu *et al.* have reported a value of (3.1 ± 0.1) × 10⁶ dm³ mol⁻¹ (30 °C, pH = 4.0) from an ITC experiment.⁶⁴ Cao and Isaacs have also observed that increasing the size of the ring in cycloalkylammonium cations increases the value of *K*_{CB[7]},⁶² while Yu *et al.* observed a slight increase in the binding constant with the *N*-methylcyclohexylammonium cation.⁶⁴ The introduction of a hydroxyl group on the β-carbon in CAPSO results in a slight reduction in the binding constant compared to CAPS.

The values of *K*_{CB[7]} increase with an increase in the number of methylene groups between the protonated nitrogen and the sulfonate group (Table 1). This would reduce the repulsions between the sulfonate group and the polar carbonyl groups on the CB[7] portal and increase the hydrophobicity of the guest.

The morpholinium buffers (MES, MOPS, MOPSO, and MOBS) exhibit lower binding constants than the corresponding cyclohexylammonium buffers (Table 1), with *K*_{CB[7]} in the range of 200–2400 dm³ mol⁻¹, increasing with an increase in the alkylsulfonate chain length (Fig. 4, 5 and S10–S17†). This trend in higher *K*_{CB[7]} values is also accompanied by a deeper inclusion of the morpholine ring within the cavity, as exhibited by the greater upfield chemical shifts of the methylene groups which are attached to the protonated nitrogen (*H*_α: -0.01, -0.23, and -0.35 ppm for MES, MOPS, and MOBS, respectively (Fig. 3, S9†)). The MOPSO buffer has a slightly lower binding constant than MOPS, similar to that seen for CAPSO compared with CAPS.

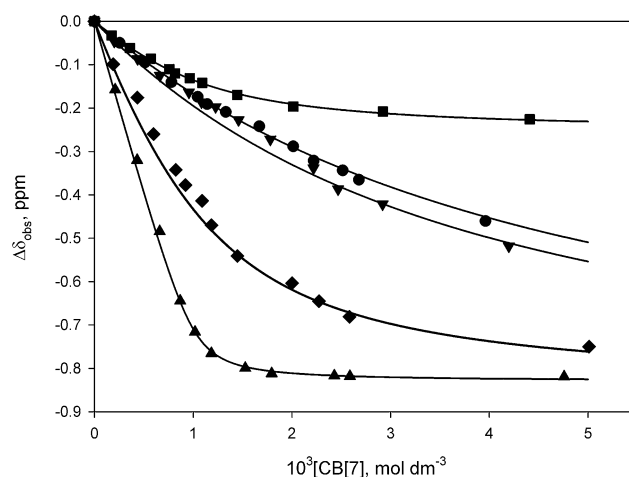


Fig. 4 ¹H NMR titrations of MES (● *H*_{3,5}), MOPS (■ *H*_{3,5}), MOPSO (▼ *H*_{3,5}), MOBS (◆ *H*_{3,5}), and MeMOBS (▲ *H*_{Me}) with CB[7] in D₂O (pD = 4.75, 0.050 mol dm⁻³ NaOAc-*d*₃/0.025 mol dm⁻³ DCl).



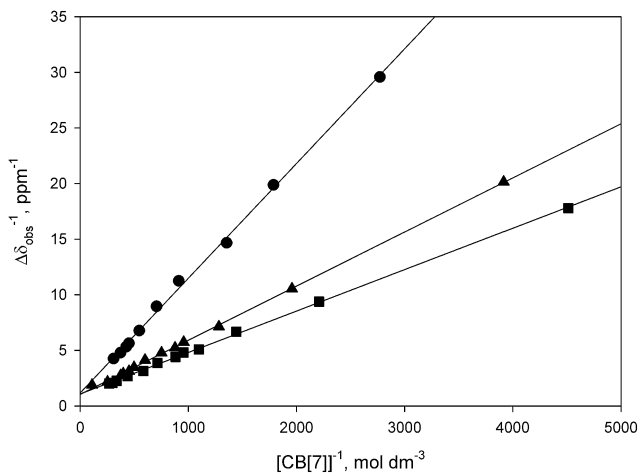


Fig. 5 Benesi–Hildebrand plots for the binding of TRIS (●), TAPS (■), and MES (▲) with CB[7] in D₂O (pD = 4.75, 0.050 mol dm⁻³ NaOAc-d₃(*)/0.025 mol dm⁻³ DCl).

Among the other zwitterionic guests investigated previously with CB[7], is carnitine which has also has a β-hydroxypropyl chain separating a trimethylammonium group (resides inside the CB[7] cavity) from a carboxylate group.⁶⁵ The binding constant of $K_{\text{CB}[7]} = 8.0 \times 10^4 \text{ dm}^3 \text{ mol}^{-1}$ for carnitine is between the values observed for MOPSO and CAPSO.

The *N*-methyl derivative of MOBS (MeMOBS) exhibited a higher binding constant of $(4.2 \pm 0.5) \times 10^4 \text{ dm}^3 \text{ mol}^{-1}$, a factor of 20 compared to that of MOBS, with a very large upfield shift of -1.05 ppm for H_z (Fig. S19 and S20[†]). We have reported a similar increase in CB[7] binding of the zwitterionic $N^{\epsilon}, N^{\epsilon}, N^{\epsilon}$ -trimethyllysine ($6.0 \times 10^4 \text{ dm}^3 \text{ mol}^{-1}$) compared with protonated dimethyllysine ($2.1 \times 10^3 \text{ dm}^3 \text{ mol}^{-1}$).⁶⁶ In both sets of guests, the replacement of a proton by a methyl group makes the ammonium center more hydrophobic and more included within the CB[7] cavity, as reflected in the values of $\Delta\delta_{\text{lim}}$.

The host–guest stability constants for this group of buffers are much lower than reported for protonated or alkylated morpholine guest molecules with CB[7], with $K_{\text{CB}[7]} = (2.3 \pm 0.4) \times 10^6$ and $(5.1 \pm 0.8) \times 10^6 \text{ dm}^3 \text{ mol}^{-1}$ for the protonated *N*-methylmorpholine and *N,N*-dimethylmorpholinium, respectively.⁶⁷ The sulfonate group on the alkyl substituent of the morpholinium buffers is likely responsible for their weaker binding to CB[7].

In the piperazinium family of buffers (HEPES, EPPS, and HEPBS), there is hydroxyethyl substituent on one of the N atoms, in addition to the alkylsulfonate substituents found in the cyclohexylammonium and morpholinium buffers. As with the other two sets of buffers, the magnitude of $K_{\text{CB}[7]}$ (Table 1, Fig. S20–S26[†]) increases with alkylsulfonate chain length, with values about an order of magnitude greater than the corresponding morpholinium buffers. Either N atom could be the site of protonation of these buffers in slightly acidic solution, and for HEPES, both protonation tautomers have been observed in the solid state.^{68–70} Upon complexation by CB[7], the large upfield shifts in the hydroxyethyl proton resonances is suggestive of inclusion of this group and the piperazine ring within the

host cavity, with the protonated nitrogen in proximity to one of the portals. The $\Delta\delta_{\text{lim}}$ values for the H_z and H_x resonances for HEPES and EPPS are larger in magnitude than normally seen for fully included guest protons (Fig. S20[†]). The chemical shift change for H_z is likely a combination of the shielding within the cavity plus an upfield shift as the buffer transitions between a tautomer equilibrium to a localized protonation on the N atom bearing the alkylsulfonate group. For the H_x resonance, this would lead to a greater downfield shift than strictly from deshielding effects of the CB[7] portal. The similarities in the resonance positions for the alkylsulfonate methylene protons, for the piperazinium buffers compared with those of the other two sets, suggests that the nitrogen bearing the alkylsulfonate group is the site of protonation on the complexed buffer. The piperazinium ring proton resonances become very broad (and difficult to locate) upon complexation (Fig. S21, S23, S25[†]), as the ring flipping rate approaches the NMR timescale.⁶⁷

The replacement of the hydroxyethyl group on the aforementioned piperazinium buffers with another alkylsulfonate group gives rise to the diprotic bases PIPES, PIPPS, and PIPBS. With PIPES and PIPPS, no change in the guest proton resonances are observed upon addition of 5-fold CB[7] at low pH (pH 1–2, both N atoms protonated), pH 5 (one N atom protonated), and high pH (pH 11–12, both N atoms nonprotonated) (Fig. S27 and S28[†]). The formation of a host–guest complex would require the host to pass over one of the sulfonate groups to bind to the central piperazinium group. This, combined with the presence of repulsive sulfonate groups near each of the two CB[7] polar portals, would not provide for very stable internal host–guest (“pseudorotaxane”) complexes. With PIPBS, the longer alkyl chains would reduce the ion-dipole repulsions, and changes in the guest proton resonances upon addition of CB[7] (Fig. S29[†]), with slow exchange behaviour, are observed for both the buffer and CB[7] proton resonances. A binding constant of $2.2 \times 10^3 \text{ dm}^3 \text{ mol}^{-1}$ for PIPBS was determined using the integrations of the free and bound methine proton resonances.

The tris(hydroxymethyl)methylammonium buffers TRIS and TAPS bind weakly to CB[7] (Table 1, Fig. 5, S30–S33[†]), with TAPS exhibiting a lower binding constant as a result of the pendant *N*-ethylsulfonate group. The binding constants were determined by Benesi–Hildebrand plots using the methylene proton resonance of the hydroxymethyl groups (Fig. 5). The methylene proton resonances for the encapsulated hydroxymethyl arms of TRIS and TAPS shift upfield by -0.92 and -0.85 ppm , respectively, while the ethylsulfonate group of the latter buffer remains outside the cavity of CB[7].

pK_a shift of CHES

The complexation of acidic guest molecules by cucurbit[7]uril is known to increase the guest pK_a value compared to that of the free acid. A pH titration of the CB[7]-complexed CHES buffer (pK_a of free CHES is 9.27)⁵⁸ was carried out in basic solution (Fig. 6). The pK_a of the bound CHES was found to be 12.37, representing an increase of 3.1 units, as the deprotonation of the quaternary nitrogen results in weaker binding to the anionic guest species. This may be compared to a shift of 1.3 pK_a units



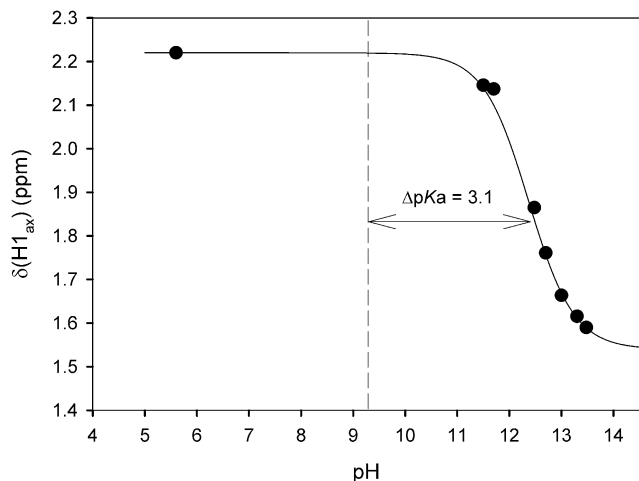


Fig. 6 pH titration of the chemical shift of the CB[7] complex of CHES. The dashed vertical line represents the pK_a of free CHES.

for cyclohexylmethylamine upon binding to CB[6].⁷¹ The larger shift for CHES with CB[7] would be consistent with a greater reduction in the binding of the conjugate base, because of the sulfonate group. The ΔpK_a value of 3.1 for CHES allows for a binding constant estimation ($K_{CB[7]}^b = K_{CB[7]}^a(10^{-\Delta pK_a})$)^{38,39} of $K_{CB[7]}^b = 3 \times 10^4 \text{ dm}^3 \text{ mol}^{-1}$ for the conjugate base of CHES. The binding constant for this form of CHES with the β -cyclodextrin (β -CD) host has been reported to be $490 \pm 20 \text{ dm}^3 \text{ mol}^{-1}$ (pH 10.5, with a similar value for CAPS)²⁸ and $440 \pm 30 \text{ dm}^3 \text{ mol}^{-1}$ (pH 11.6).²⁹ The zwitterionic form of CHES binds very much weaker (about $30 \text{ dm}^3 \text{ mol}^{-1}$) with β -CD, which would arise from a decrease in the pK_a value upon complexation.²⁹ The strong binding of the cyclohexylammonium buffers to CB[7], along with significant pK_a shifts, would make them inappropriate as biological pH buffers with cucurbit[7]uril.

Conclusions

The macrocyclic host molecule cucurbit[7]uril exhibits a wide range of binding constants with the Good's biological pH buffers in aqueous solution. The CB[7] binding constants are observed to increase as the anionic sulfonate group is placed further from the protonated nitrogen site. The morpholinium (MES, MOPS, MOPSO) and tris(hydroxymethyl)methylammonium (TRIS, TAPS) buffers exhibit binding constants of $<10^3 \text{ dm}^3 \text{ mol}^{-1}$ and could be used in studies with guests which exhibit significantly stronger binding with CB[7]. We are presently looking at reducing the CB[7] binding strength by changing the alkyl spacer between the nitrogen and sulfonate groups to methane.⁷² The dibasic PIPES and PIPBS buffers would be useful for buffering biological CB[7] solutions over a wide pH range as internal pH indicators for ^1H NMR spectroscopy⁷³ with CB[7]. A mixture of PIPES ($pK_{a1} = 2.67$ and $pK_{a2} = 6.78$) and acetic acid ($pK_a = 4.75$) would cover a significant pH range, and we are currently utilizing them in determining pK_a shifts of polyprotic drug molecules upon complexation by CB[7].

Conflicts of interest

There are no conflicts to declare.

Acknowledgements

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