


 Cite this: *RSC Adv.*, 2017, 7, 44619

Nanocasting synthesis of an iron nitride-ordered mesopore carbon composite as a novel electrode material for supercapacitors†

 Liyang Xu,^{id} Li Sun,* Jian Feng, Linlin Qi, Ishaq Muhammad, Jabeen Maher, Xiaoyu Cheng and Weiming Song*

Herein, a series of novel iron nitride (Fe₂N)/cubic-ordered mesoporous carbon (OMC) composites (Fe₂N@OMC) were synthesized *via* a facile nanocasting route and ammonia calcination. Physical characterization results indicate that small non-aggregated Fe₂N nanoparticles were evenly distributed on the external surface and in the pores of the carbon material with a linear array of mesopores in a regular interval arrangement. An electrochemical test was carried out using Fe₂N@OMC as a negative electrode of a supercapacitor in 6 mol L⁻¹ KOH aqueous electrolytes. The results indicated that the obtained Fe₂N@OMC-2 exhibited a superior specific capacitance of 547 F g⁻¹ at 1 mV s⁻¹ and 520 F g⁻¹ at 0.5 A g⁻¹, an excellent rate capability (398 F g⁻¹ at 20 A g⁻¹ with capacitance retention of 76%), and an outstanding cycling stability. After 1000 cycles, the specific capacitance retention of Fe₂N@OMC-2 remained 85%, which was much higher than 28% of the bare Fe₂N.

 Received 7th August 2017
Accepted 1st September 2017

DOI: 10.1039/c7ra08704g

rsc.li/rsc-advances

1. Introduction

In recent years, as there is an increasing requirement for the storage and utilization of a mass of energy, electrochemical devices have been developing rapidly. Among them, supercapacitors have attracted significant attention due to their high energy density, great rate capacity, and cycling stability.^{1–3} Thus, many researchers are engaged in the investigation and research of electrode materials for supercapacitors. Transition metal oxides (Fe₂O₃, Co₃O₄, and MnO₂), bimetallic oxides, and hydroxides are always used as electrode materials because of high capacitance of transition metal compounds.^{4–8} For instance, Yu *et al.* presented the design and development of novel Ag/PEDOT:PSS-nanopillar/MnO₂ layers by a layer structure for high-performance flexible supercapacitors. This electrode provided excellent specific capacitances of 862 F g⁻¹ (based on MnO₂) at a current density of 2.5 A g⁻¹.⁹ Thomas *et al.* have reported an all-solid-state, energy harvesting and storing ribbon that integrates a solar cell and a supercapacitor. In this device, MnO₂ deposited on Cu(OH)₂ nanotube arrays was used as an electrode, and the capacitance of the final electrodes was increased to 145.15 mF cm⁻².¹⁰ However, the high electrical resistivity of transition metal compounds affects the power density; this results in a significant reduction in capacitance

under operation at high-intensity current, which is not expected for electrode materials used in energy utilization and storage systems.^{11–14} Previous reports have proved that transition metal nitrides, such as chromium nitride, cobalt nitride, vanadium nitride, and iron nitride, used as electrode materials for supercapacitors have more advantages than transition metal oxides due to their superior electrical conductivity (4000–55 500 S cm⁻¹) and outstanding pseudocapacitive behaviour.^{15–18} For example, Lu *et al.* demonstrated a novel core-shell structure in which TiN nanowires were covered with ultrathin amorphous carbon protective layers. This TiN solid-state supercapacitor demonstrated a remarkable 91.3% retention in aqueous electrolytes after 15 000 cycles without any oxidation.¹⁹ Yu *et al.* synthesized tungsten oxynitride (WON) nanowires on carbon cloths, and the WON asymmetric supercapacitor anode showed a high volumetric capacitance of 4.95 F cm⁻³ at 12.5 mA cm⁻³.²⁰ Metal nitrides are considered to be the most promising representatives of new generation of electrode materials for electrochemical storage device such as lithium-ion batteries (LIBs), fuel cells, sodium-ion batteries, and supercapacitors.^{21–24} The unique nature of M (metal)-N (nitride) bonding can provide simultaneous multiple model contributions from metallic, covalent, and ionic bonding.²⁵ Moreover, transition metal nitrides may be able to provide similar properties to the noble metals due to their similar electronic structures.²⁶

Among many kinds of metal nitrides, Fe₂N is recognized as the best choice as a pseudocapacitive electrode material due to its outstanding electrical conductivity, low cost, low resistance, and environmental friendliness.^{27,28} However, in the synthesis of transition metal nitrides, products always form a large

College of Chemistry and Chemical Engineering, Qiqihar University, Qiqihar, Heilongjiang 161006, P. R. China. E-mail: qdsongweiming@163.com; sunli04070215@163.com

† Electronic supplementary information (ESI) available. See DOI: 10.1039/c7ra08704g



number of bulks of transition metal nitrides; this leads to low utilization rate of their high capacitances in a strong current system. To overcome this problem and enhance the electrochemical performance, carbon material and transition metal nitride composites have been synthesized by researchers. The carbon materials such as ordered mesoporous carbons (OMC), activated carbons, carbon nanofibers, and carbon nanotubes^{29–32} with a high specific surface area can increase the electric conductivity of composite materials and provide a carbon skeleton to support the uniform deposition of metal nitride particles, thus increasing the utilization efficiency of transition metal nitrides.^{33–36}

There are only a few investigations and research on the use of iron nitride as an electrode material for supercapacitors. Zhu *et al.* reported that Fe₂N/graphene was synthesized *via* deposition of ZnO on a graphene nanosheet surface. Then, ZnO and Fe³⁺ were subjected to ion-exchange and finally annealed under an atmosphere of ammonia. The synthetic composite tested as an electrode in an electrolyte of 1 M LiCl showed a specific capacitance of 170 F g⁻¹ at a scan rate of 10 mV s⁻¹ and 120 F g⁻¹ at 100 mV s⁻¹ in a three-electrode system.³⁷ A Fe₂N/activated carbon composite was fabricated *via* hydrothermal and ammonia treatments by Agata Śliwaka *et al.* The specific capacitance of this composite used as an electrode active material for supercapacitors operating in 6 M KOH is 498 and 371 F g⁻¹ at the scan rates of 10 and 100 mV s⁻¹, respectively.³⁸ In this study, we first report the synthesis of a series of novel Fe₂N@ordered cubic mesoporous carbon composites (Fe₂N@OMC), and the resultant best composite as an active material of the electrode for supercapacitors operating in 6 M KOH shows a superior capacitance values of 506 F g⁻¹ at the scan rates of 10 mV s⁻¹ and 435 F g⁻¹ at the scan rates of 100 mV s⁻¹.

2. Experimental

2.1 Materials

Triblock copolymer Pluronic P123 ($M_w = 5800$, PEG–PPG–PEG) was purchased from Macklin Biochemical Co. Ltd (Shanghai, China). Iron(III) chloride hexahydrate (FeCl₃·6H₂O), HCl water solution (AR, 36.0–38.0%), and other chemicals were supplied by Kemiou Chemical Reagent Co. Ltd (Tianjin, China).

2.2 Synthesis of the KIT-6 template and OMC

The cubic three-dimensional (3D) ordered mesoporous silica template KIT-6 was fabricated using the synthesis procedures reported in the literature,²⁹ and the 3D-ordered mesoporous carbon (OMC) materials were prepared using sucrose as a carbon source and KIT-6-ordered mesoporous silica as a hard template. As shown in Fig. 1, 1.00 g of sucrose precursor and 1.00 g of KIT-6 were added to 20 mL DI water and stirred to obtain a homogeneous mixture. The sucrose precursor was made to continuously flow into a mesoporous silica template by vacuum impregnation at room temperature. The obtained mixture was dried at 80 °C, and then, the dried solid was ground to form a fine powder. This powder was then calcined at 800 °C

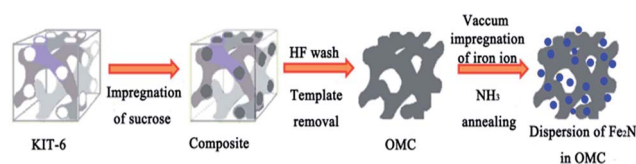


Fig. 1 Schematic for the nanocasting synthesis of OMC and Fe₂N@OMC composites.

under an inert atmosphere at a temperature ramp rate of 2 °C min⁻¹ and kept for 3 h. The calcined product was then removed out of the silica template using 10 wt% HF solution, and pure OMC was obtained.

2.3 Synthesis of Fe₂N, Fe₂N@OMC-1, Fe₂N@OMC-2, and Fe₂N@OMC-3

Fe₂N was prepared by a straightforward route. First, 1.00 g FeCl₃·6H₂O was dried at 80 °C in a vacuum oven, and then anhydrous FeCl₃ was ground to form a fine powder. Subsequently, the powder sample was tiled into a thin layer at the bottom of a quartz boat and calcined at 800 °C for 2 h at a heating rate of 5 °C min⁻¹ under the flow of ammonia at a rate of 100 mL min⁻¹. The calcined product was Fe₂N. Fe₂N@OMC-1, Fe₂N@OMC-2, and Fe₂N@OMC-3 were prepared using a vacuum nanocasting method, as shown in Fig. 1. OMC weighing 0.10 g was first dispersed in 20 mL of DI water under sonication for 2 h, and different amounts (0.25 g, 0.50 g, and 1.00 g) of FeCl₃·6H₂O were added to 20 mL DI water under stirring for 8 h to obtain a homogeneous mixture. The subsequent steps (drying, grinding, and calcination) are the same as those in the procedure for preparing Fe₂N. The obtained products were Fe₂N@OMC-1, Fe₂N@OMC-2, and Fe₂N@OMC-3.

2.4 Material characterization

The sample materials were characterized by X-ray diffraction (XRD, D8 FOCUS, Bruker, Germany) with Cu K α 1.5406 Å radiation. The microstructure of the sample materials was observed using a transmission electron microscope (TEM, H-7650, Hitachi, Japan) at a voltage of 100 kV, a high-resolution transmission electron microscope (HRTEM, Tecnai G2 F20, FEI, USA) at a voltage of 200 kV, and a scanning electron microscope (SEM, S-3400, Hitachi, Japan). Thermogravimetric analysis (TGA) was performed using a Pyris1TGA (PerkinElmer Instruments) under an air atmosphere at a heating rate of 5 °C min⁻¹. Nitrogen sorption analysis was conducted at 77 K using a physical adsorption instrument (AUTOSORB-1, Quantachrome, USA). The specific surface area was determined by the Brunauer–Emmett–Teller (BET) method, and the pore size distribution calculation was derived from the adsorption branches of the isotherms by the Barrett–Joyner–Halenda (BJH) theory. X-ray photoelectron spectroscopy (XPS) was performed using a spectrometer (ESCALAB 250Xi, Thermo Fisher, USA) with Al K α radiation (1486.6 eV). The shifts in energy (charging) of the XPS spectra were calibrated using the C 1s peak at 284.6 eV as a reference peak.



2.5 Electrochemical characterization

The electrochemical measurements were carried out using a standard three-electrode cell, where Pt foil and Hg/HgO electrode were used as the counter and reference electrodes, respectively, in a 6 M KOH aqueous solution. The working electrodes were prepared as follows:

The clean nickel foams ($1 \times 10 \text{ cm}^2$) were prepared in advance. Then, $1 \times 1 \text{ cm}^2$ nickel foams were coated by mixing 80 wt% active material (OMC, Fe_2N , $\text{Fe}_2\text{N}@$ OMC-1, $\text{Fe}_2\text{N}@$ OMC-2, or $\text{Fe}_2\text{N}@$ OMC-3) with 10 wt% polytetrafluoroethylene (PTFE) as a binding agent and 10 wt% acetylene black conductor. The mass of the active materials on each electrode was approximately 4.00 mg. Cyclic voltammetry (CV) and galvanostatic charge/discharge measurements were carried out by an electrochemical work station (CHI660E). The electrochemical impedance spectroscopy (EIS) measurement was performed using an electrochemical work station (PAR VersaSTAT 4) at an open circuit potential from 100 kHz to 0.01 Hz with an alternating current amplitude of 5 mV. The specific capacitance (C , F g^{-1}) of the electrode materials can be calculated from the galvanostatic discharged curves and the cyclic voltammograms according to the following equations:

$$C = \frac{I\Delta t}{Um} \quad (1)$$

$$C = \frac{\int Idt}{vm} \quad (2)$$

where I is the current (A), U is the operating cell voltage (V), t is the time (s), v is the scan rate (V s^{-1}), and m is the mass of the active material in the electrode (g).

3. Results and discussion

X-ray diffraction (XRD) analysis confirmed the presence of Fe_2N in all the samples except OMC, as shown in Fig. 2. The peaks at $2\theta = 37.4$ (021), 40.9 (200), 43.0 (121), 56.9 (221), 68.0 (023), and 76.0° (321) are consistent with the orthorhombic Fe_2N phase (JC-PDS #50-0958). No iron oxide phase was observed. These

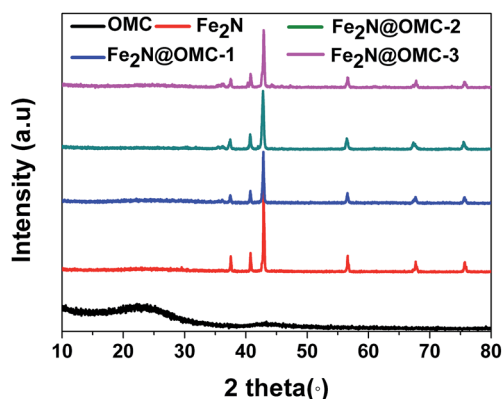
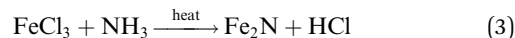


Fig. 2 The XRD patterns of OMC, Fe_2N , $\text{Fe}_2\text{N}@$ OMC-1, $\text{Fe}_2\text{N}@$ OMC-2, and $\text{Fe}_2\text{N}@$ OMC-3.

results support the complete conversion of FeCl_3 into Fe_2N by NH_3 annealing, as shown in eqn (3). The two broad peaks at approximately 23.8 and 43.6° in the XRD pattern of OMC arise from the (002) and (100) planes of graphite-type carbon, respectively.



Transmission electron microscopy (TEM) was performed to further observe the microstructure and the distribution of Fe_2N particles for the synthesized products. The TEM image of OMC shows a highly ordered mesoporous structure with a linear array of mesopores in a regular interval arrangement (Fig. 3a).³⁹ The TEM image of $\text{Fe}_2\text{N}@$ OMC-2 indicates a homogeneous distribution of small and non-agglomerative Fe_2N nanoparticles (with a diameter of 10–30 nm) on the surface of OMC, as shown in Fig. 3c. By contrast, the TEM image of $\text{Fe}_2\text{N}@$ OMC-1 shows that the well-ordered structure is still retained and the presence of Fe_2N particles almost cannot be observed in Fig. 3b. On the other hand, the image of $\text{Fe}_2\text{N}@$ OMC-3 revealed an uneven distribution of big Fe_2N particles on the surface of OMC, and the ordered structure was severely damaged, as shown in Fig. 3d. The high-resolution TEM (HRTEM) image (Fig. 3e) exhibits well-resolved lattice fringes, again confirming the high-crystalline nature of the Fe_2N nanoparticles for $\text{Fe}_2\text{N}@$ OMC-2. The lattice fringe spacings are measured to be 0.21 and

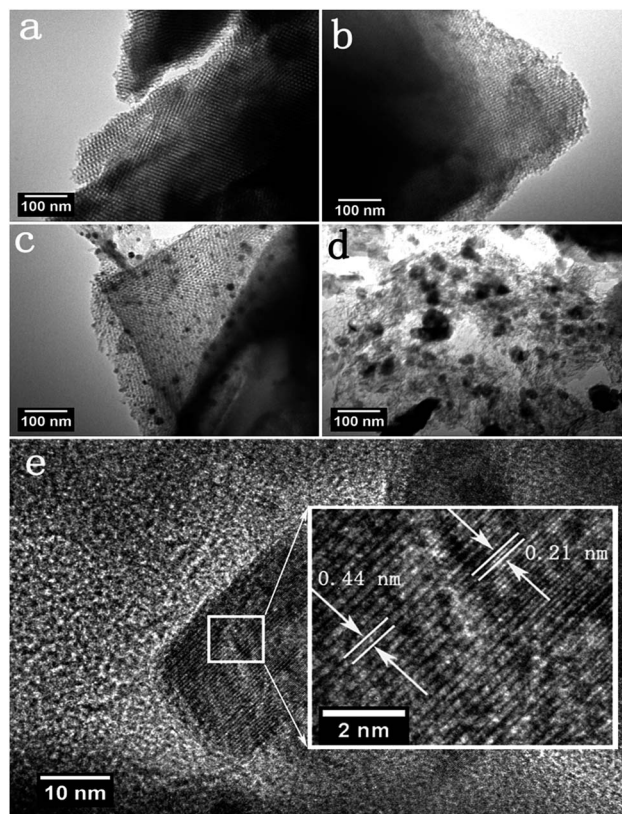


Fig. 3 TEM images of OMC (a), $\text{Fe}_2\text{N}@$ OMC-1 (b), $\text{Fe}_2\text{N}@$ OMC-2 (c), and $\text{Fe}_2\text{N}@$ OMC-3 (d). HRTEM image of $\text{Fe}_2\text{N}@$ OMC-2 (e).



0.44 nm, which are assigned to the *d*-spacing of (102) and (200) planes of the orthorhombic Fe₂N phase (JC-PDS #50-0958), respectively, and the result is consistent with that of the analysis of XRD.⁴⁰ X-ray energy-dispersive spectrometer (EDS) elemental mapping was performed to further confirm the element distribution. The SEM image (Fig. S1a ESI†) of Fe₂N@OMC-2 and its EDS elemental mapping data (Fig. S1c and d ESI†) show that the particles are mainly C (red), Fe (purple), and N (green), which are uniformly distributed in Fe₂N@OMC-2.

The content of Fe₂N calculated from thermogravimetric analysis (Fig. S2 and Table S1, ESI†) was 28.1, 40.3, and 57.1 wt% in Fe₂N@OMC-1, Fe₂N@OMC-2, and Fe₂N@OMC-3, respectively. For further characterization of the porous structure of the prepared composites, the N₂ adsorption-desorption isotherm (Fig. 4a) and pore size distribution curves (Fig. 4b) were obtained by nitrogen adsorption-desorption analysis. The samples OMC, Fe₂N@OMC-1, and Fe₂N@OMC-2 exhibit type IV curves with an H1 hysteresis loop, which indicates the mesoporous nature of the synthesized composite materials due to the presence of OMC. In comparison, the H1 hysteresis loop cannot be observed in the N₂ adsorption-desorption isotherm of Fe₂N and Fe₂N@OMC-3 (Fig. 4a). The textural properties and parameters, such as the BET surface area (*S*_{BET}), the total pore volume (*V*_T), and the average pore diameter (*D*_{BJH}), are estimated, and the data value are summarized in Table 1. The BET surface area, total pore volume, and average pore diameter of initial OMC are 1609 m² g⁻¹, 2.13 cm³ g⁻¹, and 5.29 nm, respectively. The porosity development in the composite materials is gradually reduced with the increasing deposition of Fe₂N on the carbon surface. Particularly, when the content of Fe₂N increases to 57 wt% of Fe₂N@OMC-3, the BET surface area significantly decreases from 1609 to 583 m² g⁻¹ from OMC to Fe₂N@OMC-3 due to the blockage of the pore system by the big Fe₂N particles deposited on the external carbon surface, which can be detected in the TEM image in Fig. 3d. This finding is also supported by a significant decrease in the total pore volume from 2.13 cm³ g⁻¹ for OMC to 0.44 cm³ g⁻¹ for Fe₂N@OMC-3. For Fe₂N@OMC-1, the porosity of OMC is preserved to a greater extent (*S*_{BET}: 1232 m² g⁻¹, *V*_T: 1.55 cm³ g⁻¹), but the content of Fe₂N is only 28 wt% of Fe₂N@OMC-1. The BET surface area of Fe₂N@OMC-2 is much higher than that of Fe₂N@OMC-3 (919 vs. 583 m² g⁻¹). The total pore volume of Fe₂N@OMC-2 is 0.99 cm³ g⁻¹, which is only two-fold lower than

Table 1 Porous property of composite materials identified by BET and BJH

Sample name	<i>S</i> _{BET} (m ² g ⁻¹)	<i>V</i> _T (cm ³ g ⁻¹)	<i>D</i> _{BJH} (nm)
OMC	1609	2.13	5.29
Fe ₂ N	62	0.02	14.2
Fe ₂ N@OMC-1	1232	1.55	5.03
Fe ₂ N@OMC-2	919	0.99	4.88
Fe ₂ N@OMC-3	583	0.44	2.62

that of OMC, 2.13 cm³ g⁻¹. Moreover, the content of Fe₂N can reach 40.3 wt% of Fe₂N@OMC-2. The pore size distribution calculated from *D*_{BJH} reveals that the widths of pores are 5.29, 14.2, 5.03, 4.88, and 2.62 nm in OMC, Fe₂N, Fe₂N@OMC-1, Fe₂N@OMC-2, and Fe₂N@OMC-3, respectively.

X-ray photoelectron spectroscopy (XPS) measurements were performed to analyze the surface information of the synthesized products. The surface iron concentration of Fe₂N@OMC-1, Fe₂N@OMC-2, and Fe₂N@OMC-3 is 1.7, 6.2, and 14.6 at%, respectively (Table S1, ESI†). The N 1s XPS spectra of Fe₂N@OMC-2 (Fig. 5c) and Fe₂N@OMC-3 (Fig. 5e) were divided into five component peaks, which corresponded to nitrogen bonded with metal nitride (~397.6 eV), pyridinic-N (~398.4 eV), pyrrolic-N (~399.8 eV), graphitic-N (~400.8 eV), and nitrogen oxide (~401.9 eV) functionalities.^{29,38} The N 1s XPS spectrum of

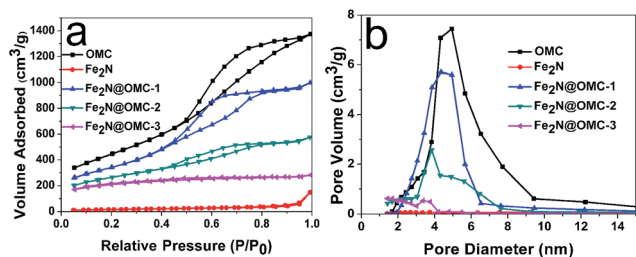


Fig. 4 The nitrogen adsorption-desorption isotherms (a) and pore diameter distribution (b) of OMC, Fe₂N, Fe₂N@OMC-1, Fe₂N@OMC-2, and Fe₂N@OMC-3.

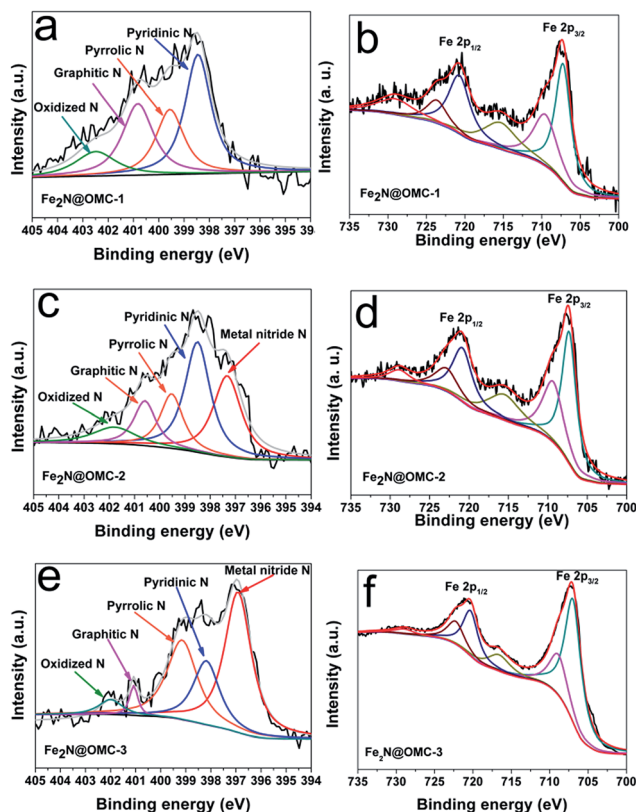


Fig. 5 High-resolution XPS spectra of N 1s of Fe₂N@OMC-1 (a), Fe₂N@OMC-2 (c), and Fe₂N@OMC-3 (e). Fe 2p spectra of Fe₂N@OMC-1 (b), Fe₂N@OMC-2 (d), and Fe₂N@OMC-3 (f).



Fe₂N@OMC-1 indicates no metal nitride functionality (Fig. 5a) as compared to that of Fe₂N@OMC-2 and Fe₂N@OMC-3, whereas Fe₂N really exists in the Fe₂N@OMC-1, as demonstrated by XRD analysis. This finding indicated that iron nitride also exists in the pores of OMC and this part of iron nitride cannot be detected *via* XPS, which also applies to Fe₂N@OMC-2 and Fe₂N@OMC-3. The high-resolution XPS spectrum of the as-made samples for the Fe 2p level exhibits main peaks at around 707.5 and 721.4 eV that correspond to the Fe 2p_{3/2} and Fe 2p_{1/2} levels, respectively. They are assigned to Fe(II) in iron nitride. Moreover, the peaks appear at 710.0, 715.6, 723.8, and 729.1 eV, which are attributed to the presence of Fe–O bonds.^{41,42} The electrochemical performance of the obtained samples was analyzed in a standard three electrode system. The cyclic voltammogram of OMC shows a rectangular and symmetric shape, indicating a typical electric double-layer (EDL) behaviour, and all cyclic voltammograms of Fe₂N-containing samples exhibit undoubtedly the reversible faradaic redox reactions attributing to Fe₂N, as shown in Fig. 6a. All the composites of Fe₂N@OMC-1, Fe₂N@OMC-2, and Fe₂N@OMC-3 have sharper and higher redox peaks as compared to bare Fe₂N; this indicates that the presence of OMC can effectively enhance the electrochemical performance of these composites due to better utilization of Fe₂N during charging and discharging when Fe₂N is evenly

distributed on the surface of OMC. The higher specific surface area of Fe₂N@OMC-2, which implies a better distribution of Fe₂N particles on the surface of OMC as compared to that of Fe₂N@OMC-3, is a reason for the stronger pseudocapacitive behavior reflected by the voltammetry curve. However, excessive specific surface area of Fe₂N@OMC-1 will sacrifice the content of Fe₂N, thereby decreasing the electrochemical pseudocapacity, which is not expected. For Fe₂N@OMC-1, Fe₂N@OMC-2, and Fe₂N@OMC-3 composites, both the charge storage mechanism of the electric double-layer behavior and faradaic redox reactions derived from OMC and Fe₂N, respectively, must be investigated and studied. In Fig. 6b, a symmetrical triangular can be observed remarkably in the galvanostatic charge–discharge curves of OMC, which demonstrates that OMC has an outstanding double-layer capacitance and the platforms of the triangular shape shown by the galvanostatic charge–discharge curves of Fe₂N, Fe₂N@OMC-1, Fe₂N@OMC-2, and Fe₂N@OMC-3 correspond to the redox peaks of cyclic voltammogram at $-0.75 \div -0.62$ V and $-1.05 \div -0.94$ V. Moreover, it was obvious that the discharge time curves (Fig. S3 ESI†) of synthetic products were in the order of OMC < Fe₂N < Fe₂N@OMC-3 < Fe₂N@OMC-1 < Fe₂N@OMC-2. For Fe₂N@OMC-2, excellent capacitance values of 547 F g⁻¹ at 1 mV s⁻¹ and 520 F g⁻¹ at 0.5 A g⁻¹ can be calculated by voltammetry and galvanostatic charge–discharge. Fe₂N@OMC-2 has better specific capacitance as compared to those reported in other studies,^{37,38} as shown in Table S2.† In the application of supercapacitors, the material of the OMC-based Fe₂N composite with very high specific capacitance has never been synthesized before. The capacitance of Fe₂N@OMC-2 is more than two times that of bare Fe₂N at 1 mV s⁻¹ and 0.5 A g⁻¹ and dramatically decreases in the strong current and high scan rate regions (Fig. 6c and d). The specific capacitances of Fe₂N@OMC-1, Fe₂N@OMC-2, and Fe₂N@OMC-3 were 310, 398, and 245 F g⁻¹ at a high current density of 20 A g⁻¹, maintaining 75, 76, and 61% of the initial capacitance at 0.5 A g⁻¹. For comparison, bare Fe₂N exhibits a specific capacitance of only 50 F g⁻¹ at a current density of 20 A g⁻¹; this fully illustrates the significance of OMC of carbon-based Fe₂N composite in improving the energy density. In addition, for supercapacitors, the cyclic stability is another indispensable performance indicator. All the prepared products were run for 1000 charge–discharge cycles at a current density of 1 A g⁻¹, as shown in Fig. 6e. It can be seen that all carbon-based Fe₂N composites show excellent cyclic stability after 1000 cycles due to a high specific surface area and the appropriate porous size. Particularly, Fe₂N@OMC-2 shows a much higher retention of 85% of the initial capacitance as compared to bare Fe₂N, which exhibits only 28% retention. Moreover, it was clearly observed that only a very small amount of Fe₂N nanoparticles accumulated after 1000 cycles (Fig. S4b ESI†) compared to those before 1000 cycles (Fig. S4a ESI†). The stable nanostructure made the Fe₂N@OMC-2 electrode have an outstanding cyclic stability. The nanocasting route can achieve a homogeneous dispersion of small non-aggregated Fe₂N nanoparticles on the external surface and in the pores of OMC; this has been demonstrated by nitrogen sorption analysis, SEM, TEM, HRTEM, and XPS results. This feature can improve the

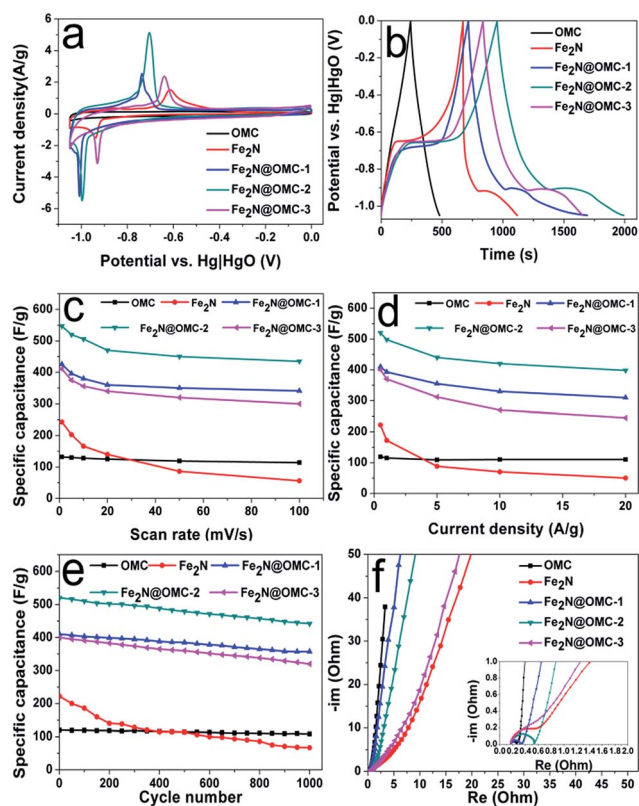


Fig. 6 Cyclic voltammograms of the synthetic products at 1 mV s⁻¹ (a), galvanostatic charge/discharge curves performed at 0.5 A g⁻¹ (b), specific capacitance as a function of the scan rate (c) and current density (d), cycle stability at a current density of 1 A g⁻¹ (e), and Nyquist plots with an inset representing the high-frequency region (f).



transport of ions from the bulk electrolyte to the composite material. The collective effect of the two factors, OMC providing more active sites due to the high specific surface area and the high pseudocapacitance of Fe₂N, accounts for the outstanding electrochemical performance of the Fe₂N@OMC-2 composite as an electrode material for supercapacitors.³⁸

The resistive characteristics of electrode materials were analyzed by EIS (Fig. 6f). In the Nyquist plots of OMC, Fe₂N@OMC-1, and Fe₂N@OMC-2, a small semicircular line corresponding to the interfacial charge-transfer resistance is noticed in the high-frequency region, which indicates that all the samples OMC, Fe₂N@OMC-1, and Fe₂N@OMC-2 have a small charge-transfer resistance. The slope lines at an angle of 80–90° are observed markedly in the lower frequency region of OMC, Fe₂N@OMC-1, and Fe₂N@OMC-2, which show an excellent charge-discharge capacitive behavior because of the fast charge-transfer rate from the active electrode material to the current collector.^{40,43} By contrast, no obvious semicircle is detected for bare Fe₂N and Fe₂N@OMC-3; this demonstrates the presence of charge-transfer resistance, resulting in a poor rate capability.⁴⁴ In addition, the slope lines at an angle near 45° for Fe₂N and Fe₂N@OMC-3 reveal the occurrence of the Warburg impedance, standing for ion-diffusion resistance.^{40,45} The ion-diffusion resistance of Fe₂N and Fe₂N@OMC-3 is higher than that of OMC, Fe₂N@OMC-1, and Fe₂N@OMC-2; this may be caused by the aggregated and bulky particles of Fe₂N, making ion diffusion in the electrolyte become more difficult.⁴⁶ The EIS measurements can further prove that the composite of OMC and small non-aggregated Fe₂N particles is advantageous in ion-transport, which plays a significant role in rate performance and specific capacitance.

Different electrolytes can better test the performance of the electrodes.⁴⁷ The Fe₂N@OMC-2 electrode was tested in 1 M LiOH, NaOH, and KOH electrolytes, as shown in Fig. S5 ESI.† It is clear that the redox peaks of CV curves increase in turn as the size of the cations in the electrolytes decreases (K⁺ > Na⁺ > Li⁺). This observation indicates enhanced pseudocapacitive reactions with smaller cations as they are easier to get intercalated into the pores of Fe₂N@OMC-2. Moreover, we tested the cyclic stability in different electrolytes (LiOH, NaOH, and KOH). After 1000 cycles, the specific capacitance retention of Fe₂N@OMC-2 remained to be 87% in KOH, which was much higher than that in LiOH (67%) and NaOH (72%). This is mainly because the smaller cation is difficult to desorb from the pore of the as-made electrode materials.

4. Conclusions

A novel Fe₂N@OMC-2 composite was successfully synthesized via a nanocasting route and NH₃ annealing. The vacuum impregnation of OMC with iron chloride resulted in the deposition of small and non-aggregated Fe₂N nanoparticles both on the external surface and in the pores of OMC. Owing to the synergistic effect between the mesoporous property of OMC and an excellent pseudocapacitance of Fe₂N, Fe₂N@OMC-2 exhibits an eye-catching high capacitance of 547 F g⁻¹ at 1 mV s⁻¹. A great rate capability of 76% was retained when the current

densities increased from 0.5 to 20 A g⁻¹, and a superior cyclic stability of 85% was retained after 1000 cycles. This study constitutes the first demonstration of employing the composite of Fe₂N and OMC as an electrode material for supercapacitors, which can effectively enhance the performance of energy storage devices.

Conflicts of interest

There are no conflicts to declare.

Acknowledgements

We gratefully acknowledge the support for this study provided by the National Natural Science Foundation of China (No. 21501104), the Natural Science Foundation of Heilongjiang Province (B2015014), and the University Nursing Program for Young Scholars with Creative Talents in Heilongjiang Province (UNPYSCT-2016088).

Notes and references

- S. L. Chiam, H. N. Lim, C. Y. Foo, A. Pandikumar and N. M. Huang, *Electrochim. Acta*, 2017, **246**, 1141–1146.
- Y. Liu and X. Peng, *Applied Materials Today*, 2017, **8**, 104–115.
- H. Yang and Y. Zhang, *J. Power Sources*, 2016, **312**, 165–171.
- K.-T. Lee, C.-B. Tsai, W.-H. Ho and N.-L. Wu, *Electrochem. Commun.*, 2010, **12**, 886–889.
- X. Meng and R. Wang, *Energy Procedia*, 2012, **17**, 1585–1590.
- J. Lv, Z. Wang, T. Liang, Y. Meng, K. Suzuki and H. Miura, *J. Electroanal. Chem.*, 2017, **799**, 595–601.
- L. Liu, H. Zhang, L. Fang, Y. Mu and Y. Wang, *J. Power Sources*, 2016, **327**, 135–144.
- X. Liu, A. Zhou, T. Pan, Y. Dou, M. Shao, J. Han and M. Wei, *J. Mater. Chem. A*, 2016, **4**, 8421–8427.
- Z. Yu, C. Li, D. Abbitt and J. Thomas, *J. Mater. Chem. A*, 2014, **2**, 10923–10929.
- C. Li, M. M. Islam, J. Moore, J. Sleppy, C. Morrison, K. Konstantinov, S. X. Dou, C. Renduchintala and J. Thomas, *Nat. Commun.*, 2016, **7**, 13319.
- D. Yang, *J. Power Sources*, 2011, **196**, 8843–8849.
- H. Chang and H. Wu, *Energy Environ. Sci.*, 2013, **6**, 3483–3507.
- M. S. Javed, C. Zhang, L. Chen, Y. Xi and C. Hu, *J. Mater. Chem. A*, 2016, **4**, 8851–8859.
- L. Ma, L. Su, J. Zhang, D. Zhao, C. Qin, Z. Jin and K. Zhao, *J. Electroanal. Chem.*, 2016, **777**, 75–84.
- M. Idrees, S. M. Abbas, R. Ata Ur, N. Ahmad, M. W. Mushtaq, R. A. Naqvi, K.-W. Nam, B. Muhammad and Z. Iqbal, *J. Chem. Eng.*, 2017, **327**, 361–370.
- B. Das, M. V. Reddy and B. V. Chowdari, *Nanoscale*, 2013, **5**, 1961–1966.
- P. J. Hanumantha, M. K. Datta, K. Kadakia, C. Okoli, P. Patel and P. N. Kumta, *Electrochim. Acta*, 2016, **207**, 37–47.
- N. Choudhary, C. Li, J. Moore, N. Nagaiah, L. Zhai, Y. Jung and J. Thomas, *Adv. Mater.*, 2017, **29**, 1605336.



- 19 X. Lu, G. Wang, T. Zhai, M. Yu, S. Xie, Y. Ling, C. Liang, Y. Tong and Y. Li, *Nano Lett.*, 2012, **12**, 5376–5381.
- 20 M. Yu, Y. Han, X. Cheng, L. Hu, Y. Zeng, M. Chen, F. Cheng, X. Lu and Y. Tong, *Adv. Mater.*, 2015, **27**, 3085–3091.
- 21 N. Mosavati, S. O. Salley and K. Y. S. Ng, *J. Power Sources*, 2017, **340**, 210–216.
- 22 S. Dong, X. Chen, X. Zhang and G. Cui, *Coord. Chem. Rev.*, 2013, **257**, 1946–1956.
- 23 S. Liu, J. Liu, W. Wang, L. Yang, K. Zhu and H. Wang, *RSC Adv.*, 2016, **6**, 86131–86136.
- 24 Y. Yu, W. Gao, Z. Shen, Q. Zheng, H. Wu, X. Wang, W. Song and K. Ding, *J. Mater. Chem. A*, 2015, **3**, 16633–16641.
- 25 L. Jiang, L. Lin, F. Yan, B. Fan, Y. Chen and Y. Qiu, *Ceram. Int.*, 2017, **43**, 9226–9234.
- 26 G. R. Li, J. Song, G. L. Pan and X. P. Gao, *Energy Environ. Sci.*, 2011, **4**, 1680.
- 27 H. Huang, S. Gao, A.-M. Wu, K. Cheng, X.-N. Li, X.-X. Gao, J.-J. Zhao, X.-L. Dong and G.-Z. Cao, *Nano Energy*, 2017, **31**, 74–83.
- 28 D. Di Lecce, R. Verrelli and J. Hassoun, *Green Chem.*, 2017, **19**, 3442–3467.
- 29 T. Panja, D. Bhattacharjya and J.-S. Yu, *J. Mater. Chem. A*, 2015, **3**, 18001–18009.
- 30 Y. Teng, E. Liu, R. Ding, K. Liu, R. Liu, L. Wang, Z. Yang and H. Jiang, *Electrochim. Acta*, 2016, **194**, 394–404.
- 31 J.-G. Wang, Y. Yang, Z.-H. Huang and F. Kang, *Carbon*, 2013, **61**, 190–199.
- 32 Y. Xie, L. Lu, Y. Tang, F. Zhang, C. Shen, X. Zang, X. Ding, W. Cai and L. Lin, *Mater. Lett.*, 2017, **186**, 70–73.
- 33 L. Feng, G. Li, S. Zhang and Y. X. Zhang, *Ceram. Int.*, 2017, **43**, 8321–8328.
- 34 X. Zhang, J. Luo, P. Tang, X. Ye, X. Peng, H. Tang, S.-G. Sun and J. Fransaer, *Nano Energy*, 2017, **31**, 311–321.
- 35 V. C. Lokhande, A. C. Lokhande, C. D. Lokhande, J. H. Kim and T. Ji, *J. Alloys Compd.*, 2016, **682**, 381–403.
- 36 L.-Q. Fan, G.-J. Liu, J.-H. Wu, L. Liu, J.-M. Lin and Y.-L. Wei, *Electrochim. Acta*, 2014, **137**, 26–33.
- 37 C. Zhu, P. Yang, D. Chao, X. Wang, X. Zhang, S. Chen, B. K. Tay, H. Huang, H. Zhang, W. Mai and H. J. Fan, *Adv. Mater.*, 2015, **27**, 4566–4571.
- 38 A. Sliwak, A. Moyseowicz and G. Gryglewicz, *J. Mater. Chem. A*, 2017, **5**, 5680–5684.
- 39 J. Feng, W. Song, L. Sun and L. Xu, *RSC Adv.*, 2016, **6**, 110337–110343.
- 40 M.-S. Balogun, M. Yu, Y. Huang, C. Li, P. Fang, Y. Liu, X. Lu and Y. Tong, *Nano Energy*, 2015, **11**, 348–355.
- 41 P. Yu, L. Wang, F. Sun, D. Zhao, C. Tian, L. Zhao, X. Liu, J. Wang and H. Fu, *Chem.–Eur. J.*, 2015, **21**, 3249–3256.
- 42 B. P. Vinayan, T. Diemant, R. J. Behm and S. Ramaprabhu, *RSC Adv.*, 2015, **5**, 66494–66501.
- 43 L. Ren, G. Zhang, Z. Yan, L. Kang, H. Xu, F. Shi, Z. Lei and Z. H. Liu, *ACS Appl. Mater. Interfaces*, 2015, **7**, 28294–28302.
- 44 F. Miao, C. Shao, X. Li, K. Wang and Y. Liu, *J. Mater. Chem. A*, 2016, **4**, 4180–4187.
- 45 J. Huang, S. Yang, Y. Xu, X. Zhou, X. Jiang, N. Shi, D. Cao, J. Yin and G. Wang, *J. Electroanal. Chem.*, 2014, **713**, 98–102.
- 46 B. Liang, Z. Qin, T. Li, Z. Dou, F. Zeng, Y. Cai, M. Zhu and Z. Zhou, *Electrochim. Acta*, 2015, **177**, 335–342.
- 47 N. Choudhary, C. Li, H. S. Chung, J. Moore, J. Thomas and Y. Jung, *ACS Nano*, 2016, **10**, 10726–10735.

