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Role of lanthanum vacancy on the structural, magnetic and magnetocaloric properties in the lacunar perovskite manganites



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Lacunar $\text{La}_{0.8-x}\square_x\text{Na}_{0.2}\text{MnO}_3 (0 \leq x \leq 0.15)$ ceramics where \square is a lanthanum deficiency were synthesized via a sol–gel method. Magnetic phase transitions and magnetocaloric effects of the ceramics have been systematically studied. Structural studies using X-ray diffraction show that all compounds crystallize in the rhombohedral structure with an $\bar{R}\bar{3}c$ space group. The lanthanum deficiency does not modify the crystal structure of the pristine compound ($x = 0$) but results in a slight change of the lattice parameters. The unit cell volume and Mn–O–Mn angle decrease with increasing deficiency content whereas the Mn–O bond length increases. This situation weakens the double-exchange interaction and hence reduces the Curie temperature (T_C). By analyzing the temperature and field dependence of magnetization, we find that the $\text{La}_{0.8-x}\square_x\text{Na}_{0.2}\text{MnO}_3$ family exhibits a second order magnetic phase transition whose critical temperature is tunable near room temperature. A considerable magnetic entropy change is observed in $\text{La}_{0.8-x}\square_x\text{Na}_{0.2}\text{MnO}_3$ near room temperature. The maximum value of the relative cooling power (RCP) is found to be $\sim 268 \text{ J kg}^{-1}$ in $\text{La}_{0.65}\square_{0.15}\text{Na}_{0.2}\text{MnO}_3$ at 5 T. This value of RCP is about $\sim 65.3\%$ of that obtained in gadolinium metal, known as one of the most important materials for magnetic refrigeration, at the same magnetic field change of 5 T. With increasing vacancy our manganites exhibit a stable magnetocaloric effect in a wide temperature range. The results suggest that $\text{La}_{0.8-x}\square_x\text{Na}_{0.2}\text{MnO}_3$ and its composite materials could be expected to have effective applications for magnetic refrigeration near room temperature.

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1. Introduction

Reversible magnetic-field-driven thermal changes are known as magnetocaloric effects (MCE) and have been proposed for magnetic refrigeration (MR) technology.¹ The latter is a technology used to warm and cool in response to the application and removal of an external magnetic field, and it receives comprehensive attention due to its several apparent advantages over traditional vapor compression based cooling technologies: high cooling efficiency, environmentally friendly technology and convenient for miniaturization.^{2–4} The MCE refers to an induced temperature change (ΔT_{ad}) and/or entropy change (ΔS) caused by introducing a magnetic field change to a magnetic material.^{5,6} This effect originates from the coupling of

a magnetic field with magnetic moments in a solid. Magnetic entropy change is known to achieve relatively high values at finite temperatures near ferromagnetic and antiferromagnetic phase transition. Theoretical predictions and experimental results for various classes of material like manganites confirm that the magnetic entropy change is larger for the first than for the second order phase transition. However, physical properties in manganites are strongly affected by the modification of the average cationic size,⁷ cationic disorder,⁸ doping level,⁹ grain boundary engineered¹⁰ particle size¹¹ and oxygen stoichiometry.¹² The induce distortion of the MnO_6 octahedron thereby changes the distance and angle between the Mn–O–Mn ions. Moreover, a strong electron–phonon interaction arising from the Jahn–Teller splitting of the outer Mn d level¹³ and the oxygen deficiency plays a role in the electrical and magnetic properties of lanthanum manganites. Also, it is well known that any magnetic transitions, or even changes in magnetic properties with temperature or magnetic field, are accompanied by transformation or, at least, some spontaneous deformation of the crystal lattice. These distortions are evidenced by several techniques such as Raman spectroscopy.¹⁴ The observation of

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structural changes in doped manganites in the presence of a magnetic stimulus can be explained by the orthorhombic distortions, which can be tuned by an applied magnetic field. Based on the interplay between the double exchange and Jahn-Teller electron-phonon coupling, this phenomenon is interpreted as spectroscopic evidence of the strong coupling between the magnetic field and the lattice leading to an additional tunable charge carrier localization in the double exchange model. In the context of MCE compounds, the changes in the microstructure of the material can produce local inhomogeneities around the Mn atoms, and a distribution of the Mn-Mn interatomic distances affecting the exchange interactions. These alterations would change the value of the Curie temperature (T_C) in the sense that instead of a unique value for T_C , a more reliable picture considers also a distribution function for this magnitude in the prepared sample. The latter distribution of T_C values can make a modification to the nature of the transition and induces a broadening of the magnetic entropy change plots.¹⁵ The technical parameter called the relative cooling power (RCP) also depends on the width of the transition temperature interval¹⁶ and determines the applicability of materials for solid-state cooling applications. In developing MCE materials with large magnetocaloric properties, many community efforts are currently devoted to MCE materials that can generate a large adiabatic temperature change (ΔT_{ad}) and isothermal entropy change (ΔS); however, these MCE materials can require very high magnetic fields, which further limits practical purposes. There are a number of different magnetocaloric materials available for use in magnetic refrigeration such as pure Gd and manganites; these are systematically described in several reviews in the literature.^{1,5,17,18} Recently, Bahl *et al.*¹⁹ have demonstrated the potential of manganites as working magnetocaloric substances for application in devices. For this purpose, lanthanum manganites and their derivatives have secured a prominent position in the MCE phenomenon as they possess high magnetic moments and are ferromagnetic in nature.^{5,20,21} This class of material has good wear resistance, high thermal and chemical stability, uncomplicated preparation, low cost and energy-efficiency, grain growth to a desired size *via* heat treatment and, more importantly, the ability to control the magnetic transition temperatures close to room temperature by substitutions which make the material useful for solid-state cooling applications.^{11,22} These oxides, inexpensive and abundant, are known only to a limited extent, and a search for new refrigerant materials and new synthesis routes leading to a stronger magnetocaloric effect is still desired. However, only a small number of investigations have been proposed to discuss the deficiency effect in MCE-manganite materials. The presence of a cationic vacancy on the A-site of the perovskite lattice cell is an alternative way to introduce a modification in the average ionic radius (r_A) of the A site and also an additional change in the Mn⁴⁺ content. This change affects the lattice parameters and in particular the Mn-O-Mn networks and can thus enhance the super-exchange interactions and consequently, T_C may be diminished.^{23,24} In continuation of previous work from our laboratory, we have investigated the effect of a cationic vacancy in the A-site on the

physical properties of manganites.²⁵⁻²⁷ This investigation is also of interest to the magnetocaloric behaviors in lacunar systems. In the present study, we synthesized novel lacunar manganite $\text{La}_{0.8-x}\square_x\text{Na}_{0.2}\text{MnO}_3$ with different A-site vacancy amounts by the Pechini-type sol-gel process and studied the structural, magnetic and magnetocaloric effects due to the presence of La³⁺ concentration deficiencies.

2. Materials production

Poly-crystalline $\text{La}_{0.8-x}\square_x\text{Na}_{0.2}\text{MnO}_3$ samples where \square is a lanthanum vacancy were prepared by a modified Pechini-type sol-gel method. High purity (99.9%) precursors La_2O_3 , Na_2CO_3 and MnO_2 were taken as starting materials in an appropriate stoichiometric ratio. The precursor solution was prepared by dissolving the starting materials in a concentrated nitric acid solution in order to convert them into nitrates. During this procedure, the solution was continuously heated and stirred using a magnetic stirrer at 60 °C in order to accelerate the dissolution. After that, one adds citric acid (CA) and ethylene glycol (EG) as chelating agents until a completely homogeneous and transparent solution is obtained. These chelators serve to adjust the viscosity of the complex solution and control the moving velocity of the metal cations, resulting in the gelation of the reaction mixture. The product mixture was heated at 80 °C under constant stirring and the polyesterification reaction between citric acid and ethylene glycol takes place. After slowly drying the solution at 130 °C, a residue of high viscosity is formed and a transparent gel is developed during the heating process. In a later step, the obtained gel is heated to 300 °C with a heating rate of 10 °C per minute with the purpose of having the propagation of a combustion which transforms the freezing gel into fine powder. The resulting powders were then calcined at 450 °C in a furnace for 6 h, in order to expel the organics from the material. The final black powders were then pressed into a pellet (of about 1 mm thickness and 13 mm diameter) followed by sintering at 900 and 950 °C for 24 hours in air with intermediate regrinding and repelling.

3. Physical measurements

To understand the structural properties of the samples under study, XRD measurements were performed at room temperature. The lattice structure and cell parameters of the $\text{La}_{0.8-x}\square_x\text{Na}_{0.2}\text{MnO}_3$ system were characterized by diffraction (XRD) data, recorded on a PANalytical X'PERT Pro MPD diffractometer, using $\theta/2\theta$ Bragg-Brentano geometry with diffracted beam monochromatized CuK α radiation ($\lambda = 1.5406 \text{ \AA}$). The diffraction patterns were collected by steps of 0.017° over the angle range 10–100°/5–90° for $x = 0$ and $x = 0.05/x = 0.1$ and $x = 0.15$. Structural analysis was carried out using the standard Rietveld procedure²⁸ built-in FULLPROF software²⁹ which minimizes the difference between the observed and simulated powder diffraction patterns. Magnetization measurements *versus* temperature and *versus* the magnetic applied field up to 5 T were carried out using BS1 and BS2 extraction type magnetometers. The used magnetometers are employed



respectively for magnetic measurements at high and low temperatures equipped with a super conducting coil developed at the NEEL Institute. These magnetometers use an extraction technique and can produce a field of 7 T for BS1 and 11 T for BS2. The sample temperature is controlled by circulating helium gas. These two instruments are automated by a computer system that allows the registration of digital data for each successive measurement. $M(T)$ data were obtained under a magnetic applied field of 0.05 T in the field cooled mode (FC). Isothermal $M(\mu_0H)$ data were measured up to 5 T. The MCE results were deduced from the magnetization measurements *versus* magnetic applied field up to 5 T at several temperatures.

4. Results and discussion

4.1. Structural study

The X-ray powder diffraction patterns registered at 300 K for $\text{La}_{0.8-x}\square_x\text{Na}_{0.2}\text{MnO}_3$ compounds are depicted in Fig. 1 and indicate that all the samples have the same perovskite structure. The diffraction peaks can be indexed in the rhombohedral setting of the $\bar{R}\bar{3}c$ space group ($Z = 6$) in which La/Na atoms are located at the 6a (0, 0, 1/4) position, Mn atoms at the 6b (0, 0, 0) position, and O atoms at the 18e ($x, 0, 1/4$) position. Fig. 1 shows typical XRD patterns for $\text{La}_{0.8-x}\square_x\text{Na}_{0.2}\text{MnO}_3$ samples including the observed and calculated profiles as well as the difference profile. The data were refined by the Rietveld technique using FULLPROF software.³⁰ The refinement produced satisfactory agreement factors and lattice parameters which are listed in Table 1. In fact, we noticed the existence of diffraction peaks with very small intensities in the XRD patterns for the lacunar samples. This is due to the presence of small amounts of Mn_3O_4 , which is frequently encountered after the final step in the synthesis of mixed-valence manganites. Given the small concentration of Mn_3O_4 , we assume that the secondary phase does not have any significant effect on the perovskite stoichiometry according to ref. 31.

Lanthanum vacancies did not affect the space group of the parent compound $\text{La}_{0.8}\text{Na}_{0.2}\text{MnO}_3$. With increasing La-deficiency amount, the room temperature unit cell volume is systematically decreased (see Fig. 2). Basically, the introduction of the vacancy (x) in the A site implies a partial conversion of Mn^{3+} to Mn^{4+} ions according to the formula $\text{La}_{0.8-x}^{3+}\square_x\text{Na}_{0.4}^{+}\text{Mn}_{0.8-3x}^{3+}\text{Mn}_{0.2+3x}^{4+}\text{O}_3^{2-}$ and surely a change in the average ionic radius $\langle A \rangle$ of this site. In our case, the increase of the lanthanum-deficiency content (x) leads to an increase of the Mn tetravalent ion number with a smaller radius ($r_{\text{Mn}^{4+}} = 0.53 \text{ \AA}$ and $r_{\text{Mn}^{3+}} = 0.65 \text{ \AA}$ (ref. 32)). This increase in Mn^{4+} content induced a decrease in the average size of the B-site cation. Also, and for electrostatic considerations, a vacancy has an average radius $\langle r_v \rangle \neq 0$ and leads to a change in the average ionic radius. These two effects explain the decrease of the unit cell volume. A similar result has been observed in other lacunar manganite systems such as $\text{Pr}_{0.7-x}\square_x\text{Sr}_{0.3}\text{MnO}_3$ (ref. 33) and $\text{La}_{0.65-x}\square_x\text{Ca}_{0.35}\text{MnO}_3$.³⁴

It is known that the magnetic state of a manganite system is very sensitive to changes in the Mn–O–Mn bond lengths and angles. For this reason, the average Mn–O bond length and

Mn–O–Mn bond angle are extracted from the Rietveld refinement at room temperature and are depicted in Fig. 2. It is clear that the decrease of unit cell volume with the increase of La-deficiency is accompanied by an increase of $\langle \text{Mn–O} \rangle$ length and a decrease of $\langle \text{Mn–O–Mn} \rangle$ bond angle of MnO_6 (corresponding to the small tilts of oxygen octahedra which accommodate the elongation of the bond). Both variations compensate one another to diminish the internal strain induced by the La-deficiency amount, thereby enlarging the octahedral distortion. Since the exchange interaction between Mn–Mn depends on both the bond angle and the bond distance, the decrease in bond angle and the increase in bond length decrease the Mn–Mn exchange interaction which leads to a lower magnetic ordering temperature T_C and produces a change in the thermomagnetic properties as evident in the magnetic data which follows.

4.2. Magnetic properties

The normalized magnetization–temperature curves, measured at a constant magnetic field equal to 0.05 T, after field cooling (FC) mode for all samples, are depicted in Fig. 3-a. The latter show clearly a transition from a paramagnetic (PM) to a ferromagnetic (FM) state with decreasing temperature. This type of transition has already been observed in other lacunar manganites such as $\text{La}_{0.6}\text{Sr}_{0.2}\text{Ba}_{0.2-x}\square_x\text{MnO}_3$ (ref. 27) and $\text{Pr}_{0.7-x}\square_x\text{Sr}_{0.3}\text{MnO}_3$.³³ A reasonable estimation of the Curie temperature T_C can be obtained by determining the temperature of the minimum differential quotient of magnetization ($dM(T)/dT$). Fig. 3-b shows that the magnitudes and shapes of the dM/dT curves are varied depending on the vacancy level, which renders some indications regarding the sample homogeneity. These curves reveal a strong variation of magnetization around the Curie temperature T_C , defined as the temperature at which the $(dM/dT - T)$ curve reaches a minimum. This indicates that there is a possible large magnetic entropy change around T_C . A slight shift in T_C towards lower temperatures is also observed with increasing the lanthanum-deficiency content (x). The inset of Fig. 3-b shows the Curie temperatures T_C of $\text{La}_{0.8-x}\square_x\text{Na}_{0.2}\text{MnO}_3$ for different x amounts. The decrease in T_C is obviously related to an increase in the rate of Mn^{4+} , since the antiferromagnetic super-exchange interaction $\text{Mn}^{4+}\text{O}^{\text{2-}}\text{Mn}^{4+}$ is strengthened with the ferro-double exchange interaction $\text{Mn}^{3+}\text{O}^{\text{2-}}\text{Mn}^{4+}$. However, Table 1 contains an evaluation of the charge carrier bandwidth W through the empirical formula $W \propto \cos(\theta/2)/\text{Mn–O}^{3.5}$ (where $\theta = 180^\circ$ is the tilt angle in the Mn–O–Mn plane). It is found that the bandwidth W decreases with increasing lanthanum-deficiency. This decrease reduces the overlap between the O-2p and the Mn-3d orbitals, which in turn decreases the exchange coupling of $\text{Mn}^{3+}\text{–Mn}^{4+}$ as well as the magnetic ordering temperature T_C .¹³

The inverse of the susceptibility $\chi^{-1} \sim 1/M$ (calculated from the data of Fig. 3-a) is shown in Fig. 3-c, as a function of temperature in the PM region ($T > T_C$) for $x = 0.05$ and $x = 0.15$. It can be fitted by a Curie Weiss law³⁴ $\chi = C/(T - \Theta_p)$ where C is the Curie constant and Θ_p is the paramagnetic Weiss temperature. From the linearity of the χ^{-1} curve for all of our materials,



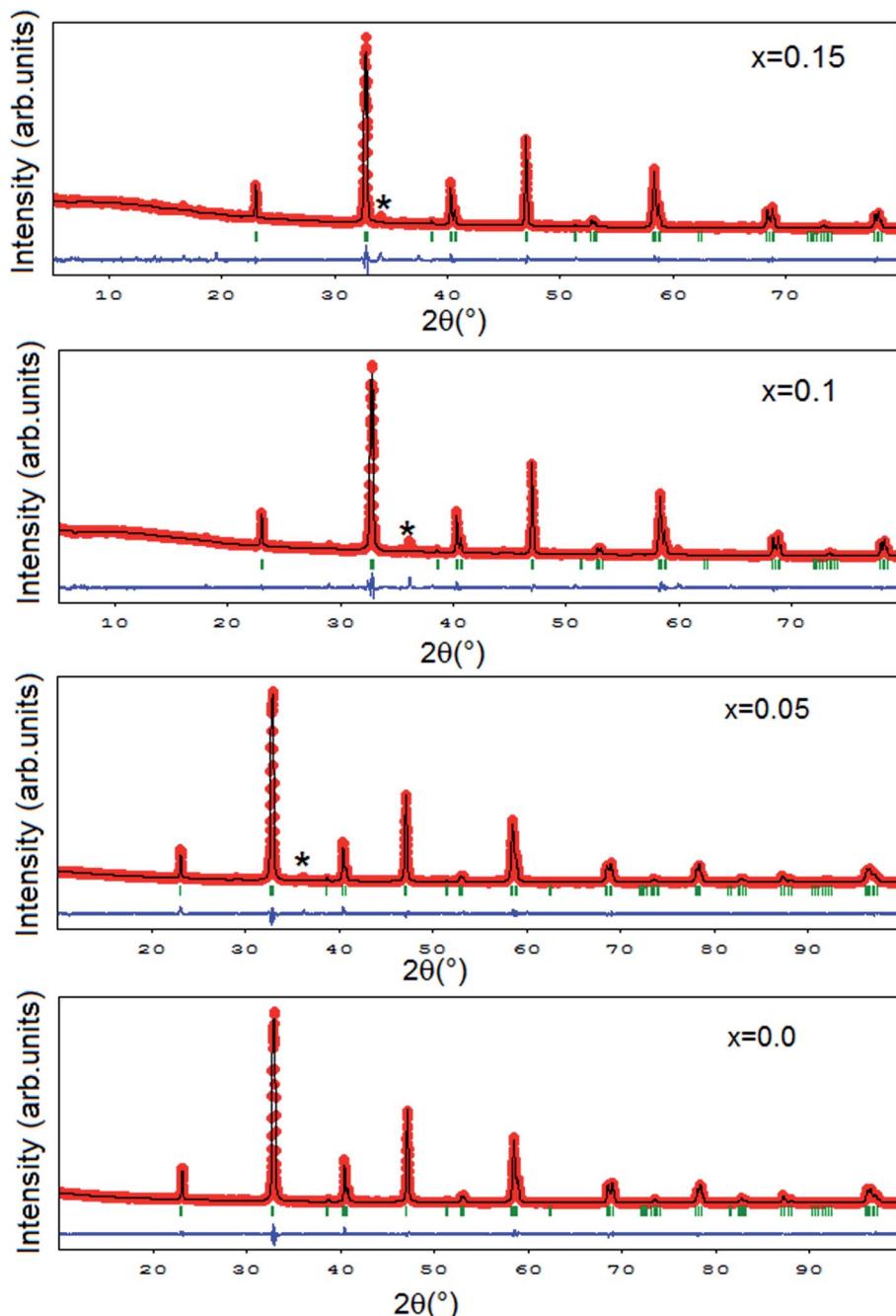


Fig. 1 XRD patterns of the XRD Rietveld refinement results for $\text{La}_{0.8-x}\square_x\text{Na}_{0.2}\text{MnO}_3$ ($0 \leq x \leq 0.15$) at RT showing the presence of a manganite as the majority phase and a minor secondary phase. Small extra peaks of the secondary phase are marked by black asterisks. The difference between the measured (red color) and calculated data (black color) is plotted in the bottom (blue color). Green bars correspond to Bragg positions.

the paramagnetic Curie temperature θ_p is found to be ~ 335 , 328, 326 and 323 K for $x = 0$, 0.05, 0.1 and 0.15, respectively. θ_p decreases with x content following the same trend as T_C . The positive values of θ_p reveal the existence of a ferromagnetic exchange interaction between the nearest neighbors in the paramagnetic region. Moreover, the obtained values are slightly higher than the corresponding T_C . This may be due to the existence of short-range FM correlation in the PM state which is related to the presence of a magnetic inhomogeneity.³⁵ Fitting

the linear $1/\chi(T)$ data to the Curie Weiss law introduces C values for all samples. Using these values, we obtained the experimental effective paramagnetic moment $\mu_{\text{eff}}^{\text{exp}}$. It is noted that the $\mu_{\text{eff}}^{\text{exp}}$ values for $x = 0$, 0.05, 0.1 and 0.15 are 4.44, 4.23, 5.085 and 5.107 μ_B respectively. Assuming orbital momentum to be quenched in Mn^{3+} and Mn^{4+} , commonly, the theoretical effective moment for each case is determined by the equation: $\mu_{\text{eff}}^{\text{theo}} = g\sqrt{S(S+1)} \mu_B$ where $g = 2$ is the gyromagnetic factor and S is the spin of the cation (the S values are 3/2 for Mn^{4+} and



2 for Mn^{3+}). The moment of the $La_{0.8-x}\square_xNa_{0.2}MnO_3$ system is expected to originate from the moments of the Mn ion only; thus, the calculated effective paramagnetic moment per formula unit can be written as:

$$\mu_{\text{eff}}^{\text{theo}} = \sqrt{(0.8 - 3x) [\mu_{\text{eff}}^{\text{th}}(Mn^{3+})]^2 + (0.2 + 3x) [\mu_{\text{eff}}^{\text{th}}(Mn^{4+})]^2} \mu_B \quad (1)$$

Considering the Mn^{3+} and Mn^{4+} magnetic moments ($\mu_{\text{eff}}^{\text{th}}(Mn^{4+}) \approx 3.87 \mu_B$ and $\mu_{\text{eff}}^{\text{th}}(Mn^{3+}) \approx 4.9 \mu_B$), $\mu_{\text{eff}}^{\text{theo}}$ was calculated for all compounds as 4.712, 4.565, 4.415 and 4.217 μ_B for $x = 0.00, 0.05, 0.1$ and 0.15, respectively. These values are not consistent with the experimental results suggesting the formation of FM clusters of Mn^{3+} – Mn^{4+} double exchange pairs in the paramagnetic region.³⁶ The discrepancy between the experimental and theoretical values implies that some short-range FM couplings might have been developed in the PM region contributing to the additional magnetic moments.

Table 1 Refined structural parameters at room temperature, W/W_0 , Curie constant C , the paramagnetic Weiss temperature Θ_p , and theoretical $\mu_{\text{eff}}^{\text{th}}$ and experimental $\mu_{\text{eff}}^{\text{exp}}$ as a function of the x content for $La_{0.8-x}\square_xNa_{0.2}MnO_3$ ($0 \leq x \leq 0.15$) samples

	$x = 0$	$x = 0.05$	$x = 0.1$	$x = 0.15$
a (Å)	5.488 (3)	5.483 (2)	5.481 (9)	5.480 (9)
b (Å)	5.488 (3)	5.483 (2)	5.481 (9)	5.480 (9)
c (Å)	13.330 (4)	13.310 (4)	13.319 (1)	13.317 (8)
Volume (Å ³)	347 740 (9)	347 180 (1)	346 628 (1)	346 399 (7)
$d_{\text{Mn-O}}$ (Å)	1.949	1.950	1.951	1.951
$\theta_{\text{Mn-O-Mn}}$ (°)	166.09	164.79	164.31	164.19
χ^2	2.85	2.53	2.80	3.05
W/W_0 (10 ⁻²)	9.6	9.571	9.548	9.54
C (μ_B)	2.465	2.237	3.233	3.261
Θ_p (K)	335	328	326	323
$\mu_{\text{eff}}^{\text{th}}$ (μ_B)	4.712	4.565	4.415	4.217
$\mu_{\text{eff}}^{\text{exp}}$ (μ_B)	4.44	4.23	5.085	5.107

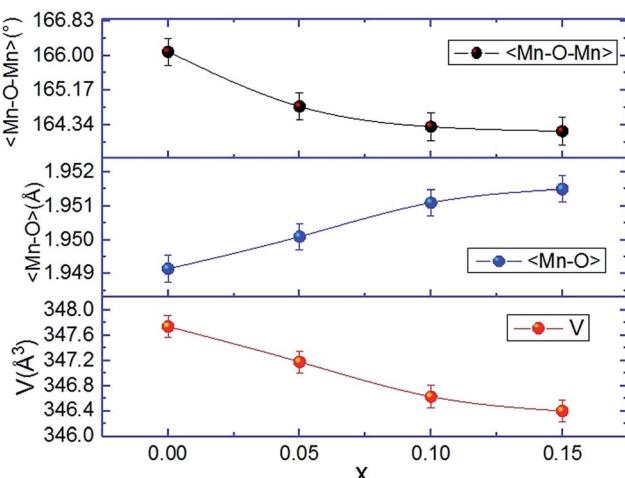


Fig. 2 Volume, $\langle Mn-O \rangle$ bond lengths and $\langle Mn-O-Mn \rangle$ bond angles versus x for all samples.

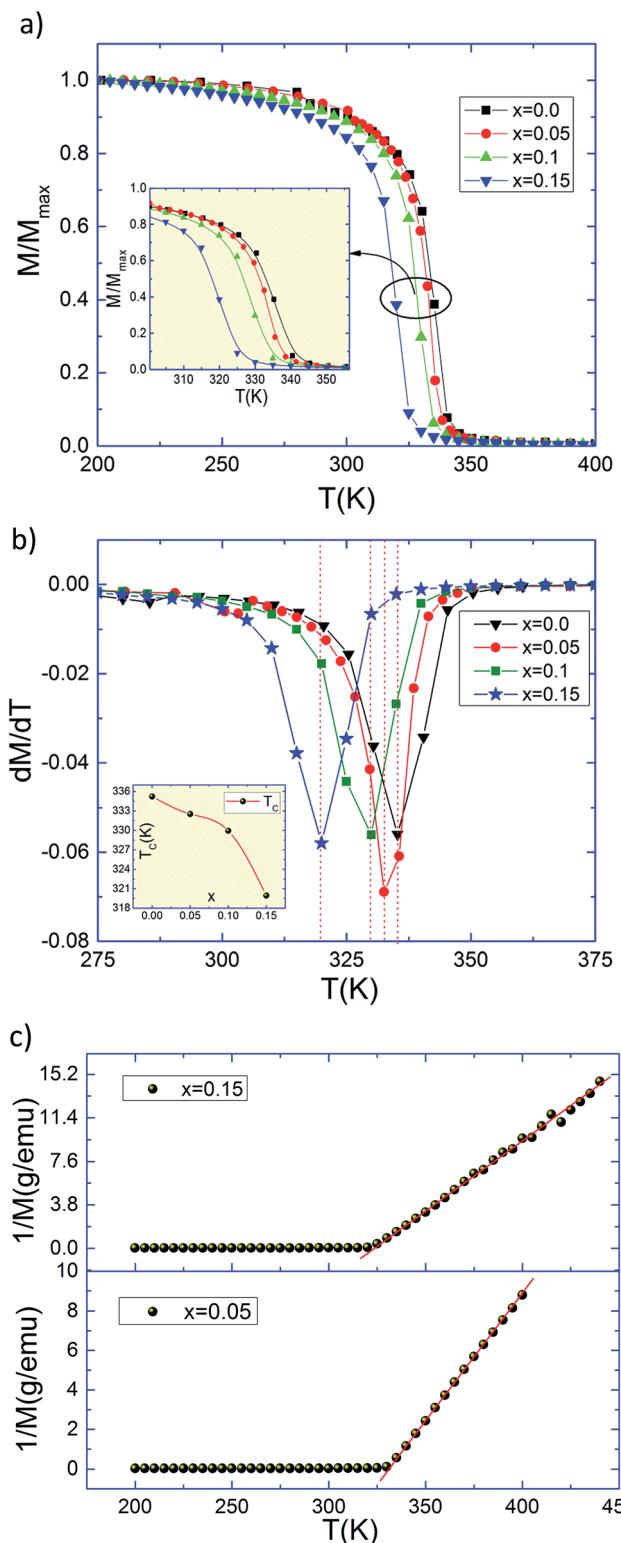


Fig. 3 (a) Magnetization measurements as a function of temperature for $La_{0.8-x}\square_xNa_{0.2}MnO_3$ ($0 \leq x \leq 0.15$) samples under 0.05 T. (b) dM/dT as a function of temperature of $La_{0.8-x}\square_xNa_{0.2}MnO_3$ ($0 \leq x \leq 0.15$). The inset shows the Curie temperature evolution as a function of La content for $La_{0.8-x}\square_xNa_{0.2}MnO_3$ ($0 \leq x \leq 0.15$). (c) Temperature dependence of the inverse of magnetic susceptibility for $La_{0.8-x}\square_xNa_{0.2}MnO_3$ ($x = 0.0, 0.05, 0.1$ and 0.15). The red line presents the linear fit at high temperature.

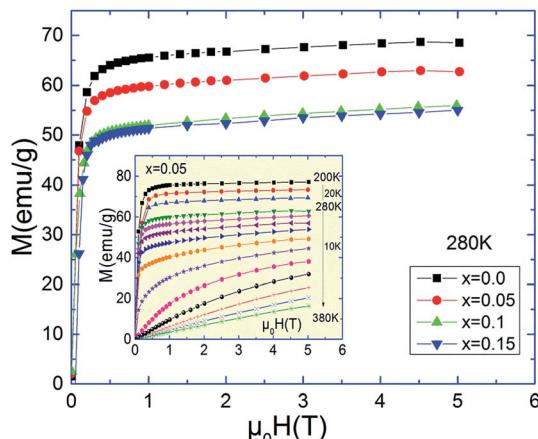


Fig. 4 Variation of the magnetization as a function of applied magnetic field at different temperatures for $\text{La}_{0.8}\text{Na}_{0.2-x}\square_x\text{MnO}_3$ ($0 \leq x \leq 0.15$) samples at 280 K. Inset shows isothermal magnetization curves at various temperatures for $x = 0.05$.

Fig. 4 presents the isothermal magnetization curves for $\text{La}_{0.8-x}\square_x\text{Na}_{0.2}\text{MnO}_3$ ($0 \leq x \leq 0.15$). These $M(\mu_0 H)$ measurements are measured near the magnetic transition with increasing magnetic field in a wide temperature range. It can be seen that $M(\mu_0 H)$ versus magnetic applied field up to 5 T recorded at 280 K shows a typical ferromagnetic nature for $\text{La}_{0.8-x}\square_x\text{Na}_{0.2}\text{MnO}_3$. The magnetization increases sharply with magnetic applied field for $\mu_0 H = 1$ T and then saturates. It can be seen that magnetization decreased as the lanthanum-deficiency in the $\text{La}_{0.8-x}\square_x\text{Na}_{0.2}\text{MnO}_3$ sample increased. This verifies that a vacancy gets incorporated into each sample proportionally to the amount of the x deficiency. This should be related to the observation of the enlarged Mn–O bond length and reduced Mn–O–Mn bond angle for our system. The inset of Fig. 4 shows isothermal magnetization curves for the $\text{La}_{0.75}\square_{0.05}\text{Na}_{0.2}\text{MnO}_3$ sample, measured under an applied magnetic field ranging from 0–5 T and at a temperature ranging from 200–380 K. At $T < T_c$, the $M(\mu_0 H)$ curves are non-linear corresponding to the FM state and become linear at $T > T_c$ corresponding to the PM state.

4.3. Magnetocaloric properties

According to thermodynamics, magnetic entropy change ΔS caused by the variation of applied magnetic field from 0 to $\mu_0 H_{\max}$ is given by:

$$\Delta S(T, \mu_0 H_{\max}) = S(T, \mu_0 H_{\max}) - S(T, 0) = \int_0^{\mu_0 H_{\max}} \left(\frac{\partial S}{\partial H} \right)_T d\mu_0 H \quad (2)$$

From the magnetization measurements made at discrete field and temperature intervals ΔS can be approximately evaluated by the following expression:

$$\Delta S_M(T, \mu_0 H_{\max}) = \sum_i \frac{M_{i+1}(T_{i+1}, \mu_0 H_{\max}) - M_i(T_i, \mu_0 H_{\max})}{T_{i+1} - T_i} \Delta \mu_0 H \quad (3)$$

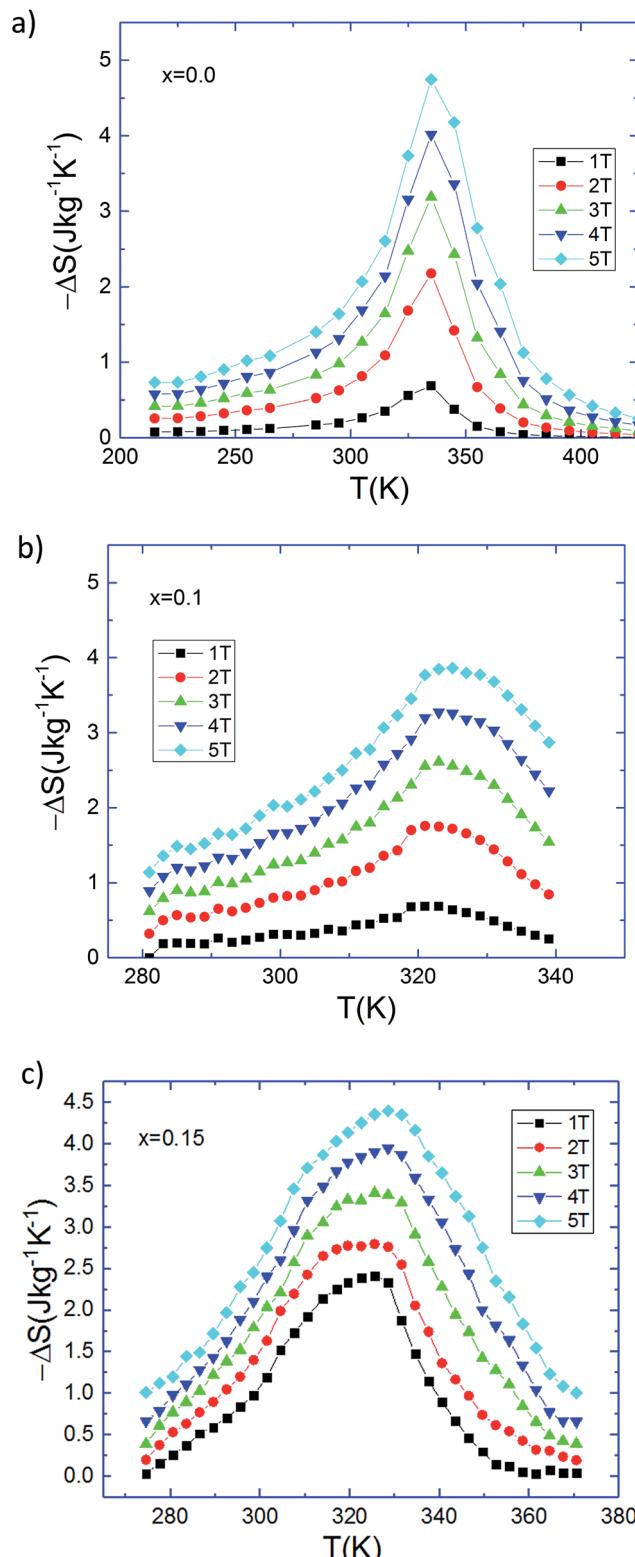


Fig. 5 Magnetic entropy change versus temperature for the $\text{La}_{0.8}\text{Na}_{0.2-x}\square_x\text{MnO}_3$ samples under several magnetic applied field changes: (a) $x = 0.00$, (b) $x = 0.10$, and (c) $x = 0.15$.

In this expression, M_i and M_{i+1} are the experimental values of magnetization measured, under an applied magnetic field $\mu_0 H$, at the temperatures T_i and T_{i+1} , respectively.



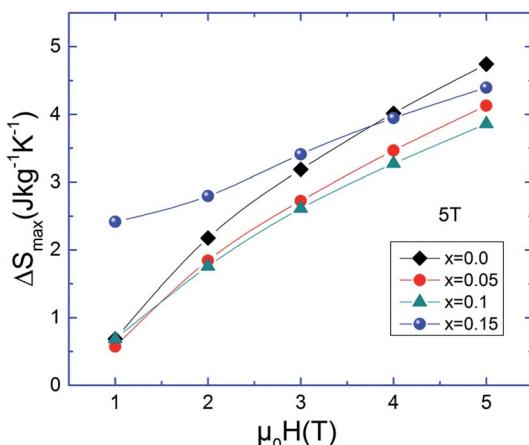


Fig. 6 The magnetic field dependence of ΔS_{\max} for $\text{La}_{0.8}\text{Na}_{0.2-x}\square_x\text{Na}_{0.2}\text{MnO}_3$ ($0 \leq x \leq 0.15$).

The magnetic entropy change, $-\Delta S$, induced by the magnetic field change was calculated using the $M(\mu_0H)$ data. In the neighborhood of room temperature, the $\text{La}_{0.8-x}\square_x\text{Na}_{0.2}\text{MnO}_3$ system is significant compared with alkaline-earth-doped

lanthanide manganites such as La-A (Ca, Ba and Sr)- MnO_3 .³⁷ Fig. 5 shows the temperature dependence of $-\Delta S$ for different applied magnetic field changes for $x = 0, 0.1$ and 0.15 . The obtained negative sign of the magnetic entropy change means that heat is released by the adiabatic change of the magnetic field confirming the ferromagnetic behavior of our lacunar samples. It can be seen that ΔS depends on both T and μ_0H . The magnetic entropy change increases to a maximum value (ΔS_{\max}) when the temperature approaches Curie temperature and reaches a maximum value around T_C . There is, however, a notable change in the shape of the $-\Delta S(T)$ curves, specifically, increasing lanthanum-deficiency content gives an increase in the width (see Fig. 5) and alters the overall peak structure to have a relatively flat shape more typical of magnetocaloric materials with second order phase transitions. The largest changes in magnetic entropy take place near T_C , which is a property of simple ferromagnets due to the efficient ordering of magnetic moments induced by the magnetic field at the ordering temperature. Basically, a broad peak on the $\Delta S(T)$ curves appears near the respective T_C , unlike the sharp peak typical of $\text{La}_{2/3}\text{Ca}_{1/3}\text{MnO}_3$,³⁸ where the phase transition is of the first order accompanied by hysteresis.³⁹

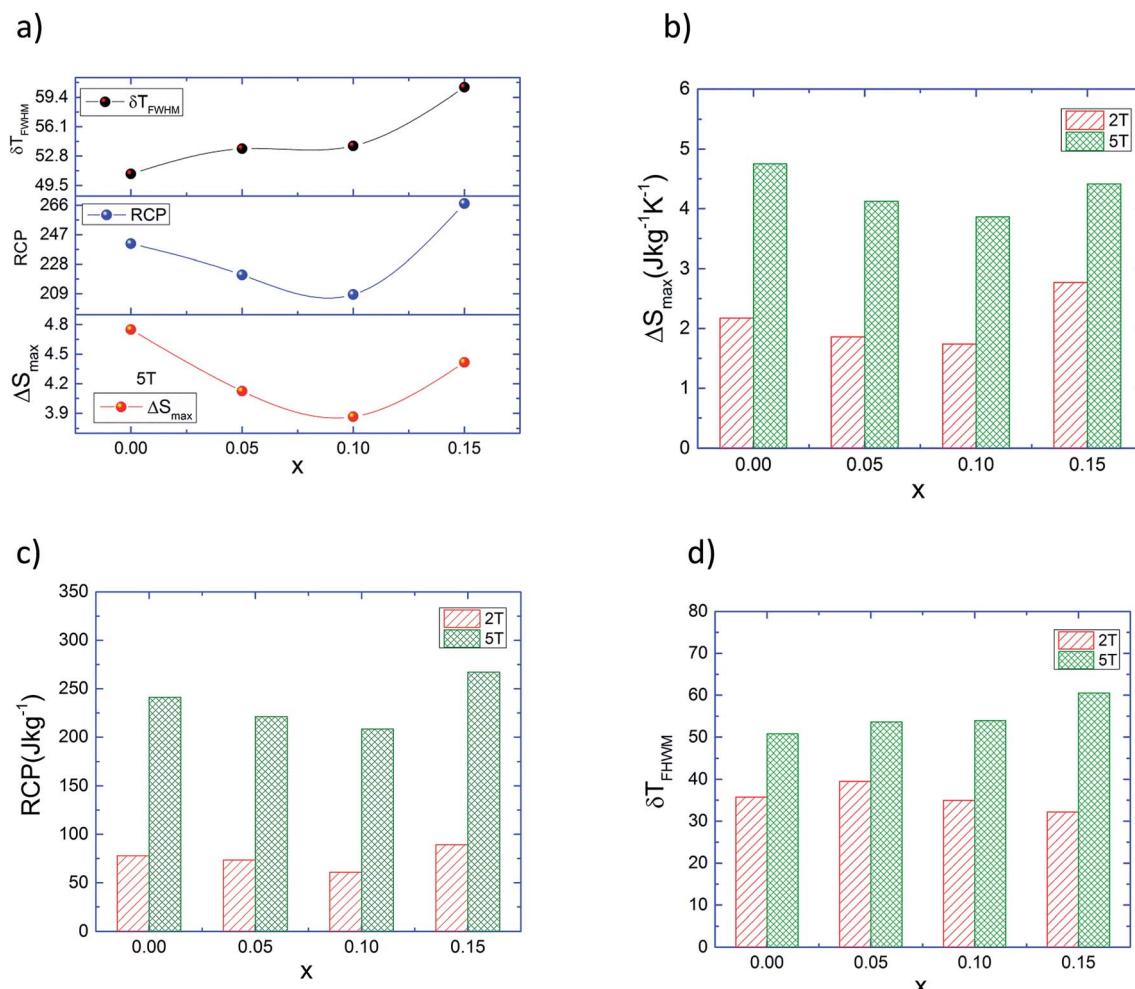


Fig. 7 Variation of ΔS_{\max} , δT_{FWHM} and RCP with La-deficiency. (a) for 5 T and (b-d) for 2 and 5 T.

The magnitude of change in the magnetic entropy of the $\text{La}_{0.8-x}\square_x\text{Na}_{0.2}\text{MnO}_3$ system is significant but lower than that with pure gadolinium (Gd) and intermetallic $\text{Gd}_5\text{Si}_2\text{Ge}_2$, which are considered the most conspicuous magnetocaloric materials at room temperature. Unfortunately, the Gd and the $\text{Gd}_5\text{Si}_2\text{Ge}_2$ materials are more expensive and oxidizable in air. Compared to Gd, our manganites present some advantages including low production cost, chemical stability, high resistivity, and corrosion-free effects. It is clear from Fig. 6 that the magnitude of the peak increases with increasing value of magnetic field for each composition. The main panel of Fig. 7a shows the variations in ΔS , RCP and ΔT_{FWHM} under $\Delta(\mu_0 H) = 5$ T as functions of the La-deficiency amount. While x increases monotonically, ΔT_{FWHM} increases; however, the relative cooling power (RCP) shows a maximum value of ~ 268 J kg^{-1} for $x = 0.15$ and a minimum value of ~ 210 J kg^{-1} for $x = 0.1$. The obtained values are comparable to about $\sim 65.3\%$ for $x = 0.15$ and $\sim 51.2\%$ for $x = 0.1$ of that of pure Gd (RCP = 410 J kg^{-1} (ref. 40)) which is considered the most conspicuous magnetocaloric material at room temperature. Thus, this indicates that the RCP value is large revealing the applicability of $\text{La}_{0.8-x}\square_x\text{Na}_{0.2}\text{MnO}_3$ in cooling devices and suggests that these compounds can thus be used in active magnetic refrigeration as suggested by Barclay.⁴¹

A comparison of the thermomagnetic properties in the $\text{La}_{0.8-x}\square_x\text{Na}_{0.2}\text{MnO}_3$ series at 5 T and 2 T magnetic fields is depicted in Fig. 7b, c and d respectively. However, it is interesting to notice that RCP reaches a maximum for the $x = 0.15$ lanthanum-vacancy level; for this composition, with T_C still around room temperature, the sharpness of the ferromagnetic transition makes it a potential candidate competitive with state-of-the art conventional refrigerant materials. Earlier investigations explain that the large MCE in perovskite manganites such as lacunar systems can originate from the spin-lattice coupling related to the magnetic ordering process. Because of this coupling mechanism, the lattice structural change in the Mn–O bond distances and Mn–O–Mn bond angles with temperature which exhibit variation in volume can cause an additional change in the magnetic properties of the magnetic substance.⁴²

For practical MCE use, it is better to have some information about the change in specific heat ΔC_p of the compounds, which is desirable for the fulfillment of a magnetic refrigerator. In our case, the ΔC_p associated with magnetic field changes from zero to $\mu_0 H$ can be deduced from ΔS induced in the material, by the following expression:

$$-\Delta C_p = T \frac{\partial \Delta S}{\partial T} \quad (4)$$

Through this relation, ΔC_p values for several maximum applied magnetic fields ranging from 1 to 5 T of $\text{La}_{0.8-x}\square_x\text{Na}_{0.2}\text{MnO}_3$ with $x = 0, 0.1$ and 0.15 are presented in Fig. 8. One can see clear anomalies around T_C in all curves, which are due to the magnetic phase transition.^{43,44} The magnitude of ΔC_p varies suddenly from (+) values to (–) values near T_C and quickly decreases with decreasing temperature. The sum of the two fractions is the magnetic contribution to the total specific heat

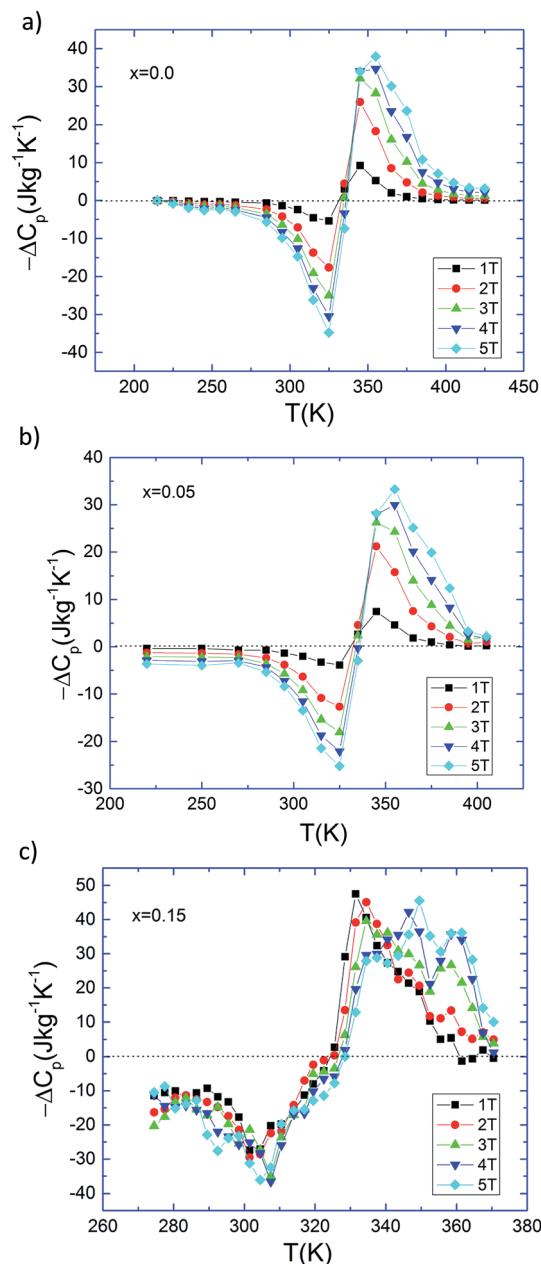


Fig. 8 The heat capacity changes $-\Delta C_p$ as a function of temperature for $\text{La}_{0.8}\text{Na}_{0.2-x}\square_x\text{MnO}_3$ (a) $x = 0.00$, (b) $x = 0.05$, and (c) $x = 0.15$ compounds.

which involves the cooling or heating power of the magnetic heat pumps and refrigerators.⁴⁵ Specific heat presents the advantage of delivering values required for further pump or refrigerator design, should the sample in question be selected.⁴⁶

Fig. 9 presents the temperature dependence of the exponent n for $\text{La}_{0.8-x}\square_x\text{Na}_{0.2}\text{MnO}_3$ ($x = 0$ and $x = 0.15$) describing the field dependence of the maximum magnetic entropy changes calculated locally from $n(T, \mu_0 H) = \frac{\text{dln}(|\Delta S|)}{\text{dln}(\mu_0 H)}$ where n is related to the magnetic state of ferromagnetic materials. The value of n depends on the values of the applied field and temperature. For our samples, it can be noted from the obtained curves that the



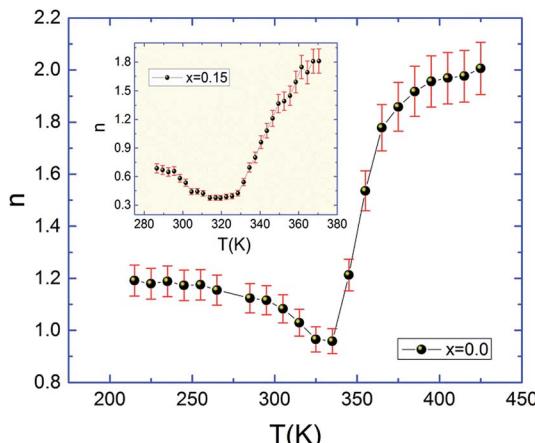


Fig. 9 Temperature dependence of exponent n for $x = 0.0$ and $x = 0.15$.

value of n reaches 2 above the Curie temperature and 1 far below the Curie temperature. The n exponent exhibits a moderate decrease with increasing temperature, with a minimum value in the vicinity of the transition temperature, sharply increasing above T_C . This exponent takes a minimum value at T_C . The obtained values of n are significantly different than $2/3$, as predicted by the mean field theory. Thus, this variation of $n(T)$ indicates the inhomogeneous character for all samples, and among the several forms of the inhomogeneous phase is the existence of local inhomogeneities or superparamagnetic clusters in the vicinity of the Curie temperature in the $\text{La}_{0.8-x}\square_x\text{Na}_{0.2}\text{MnO}_3$ compounds. A similar behavior has been observed in other perovskite materials.^{47–49}

According to the suggestion by Franco and Conde,⁵⁰ universal scaling analysis helps to determine the nature of the magnetic phase transition (first or second order) in our samples.⁵¹ This model can be confirmed by normalizing maximum magnetic entropy changes and temperatures, as defined by the following relation:⁵²

$$\Delta S'(T, \Delta\mu_0 H = \mu_0 H_{\max}) = \frac{\Delta S(T, \Delta\mu_0 H = \mu_0 H_{\max})}{\Delta S^{\text{peak}}(T, \Delta\mu_0 H = \mu_0 H_{\max})} \quad (5)$$

The temperature axis was rescaled in a different way below and above T_C , just by indicating that the position of the reference temperatures T_{r1} and T_{r2} of each curve corresponds to $\theta = \pm 1$:⁵³

$$\theta = \begin{cases} -(T - T_C)/(T_{r1} - T_C); & T \leq T_C \\ (T - T_C)/(T_{r2} - T_C); & T > T_C \end{cases} \quad (6)$$

Here, the two additional reference points are chosen such that $\Delta S(T_{r1}) = \Delta S(T_{r2}) = \Delta S^{\text{peak}}/2$. This method should remove the temperature and field dependence of the set of $\Delta S(\mu_0 H, T)$ plots so that all curves processed with the same scaling protocol collapse onto a single universal curve. This procedure has been successfully applied to several families of magnetic materials.^{54–58} Fig. 10 shows that the rescaled ΔS data for several

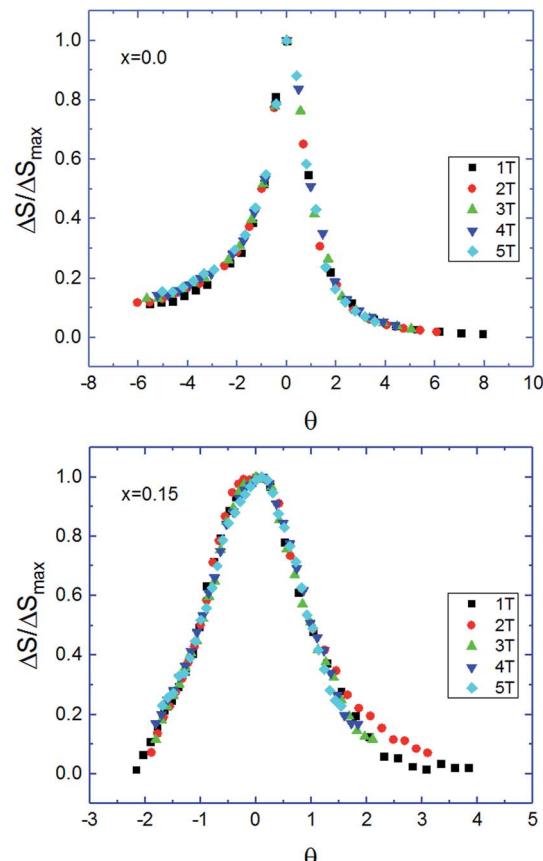


Fig. 10 Universal behavior of the scaled entropy change curves at different fields for $x = 0.0$ and $x = 0.15$.

applied fields and temperatures can collapse onto a single curve, confirming the master curve behavior in this system and clearly indicating that the magnetic phase transition in $\text{La}_{0.8-x}\square_x\text{Na}_{0.2}\text{MnO}_3$ is second order. Therefore, materials undergoing second-order transition are preferred for magnetic cooling applications.

5. Conclusion

To summarize, we have investigated the effect of lanthanum vacancy on the structural, magnetic, and magnetocaloric properties of $\text{La}_{0.8-x}\square_x\text{Na}_{0.2}\text{MnO}_3$ ($0 \leq x \leq 0.15$). The rhombohedral samples are produced from cheap and abundant elements using a simple and straight forward process making them economically attractive for room temperature applications. The increase of lanthanum deficiency induces local distortions in the Mn–O–Mn bonds and consequently causes a random distribution in the magnetic exchange interactions. The competition between magnetic interactions leads to the appearance of an inhomogeneous magnetic state in our lacunar system. Our samples exhibit second order PM–FM transition with decreasing temperature and their Curie temperatures decrease. Consequently, the ability to tune the temperature transition close to room temperature is revealed to be possible by changing the lanthanum-deficiency content as well. The

investigation of the magnetocaloric properties reveals a broad peak of $-\Delta S$ around T_C . As a main result from a practical point of view, for $x = 0.15$, the relative cooling power shows a maximum value of $\sim 268 \text{ J kg}^{-1}$ under an applied magnetic field of 5 T, which is sufficient for potential applications in magnetic cooling near room temperature due to the easy synthesis and processing, high chemical stability, very low cost and high chemical stability of these compounds. The ΔS data for different compositions distribute on a single master curve, further confirming the nature of second-order magnetic transition needed for practical applications.

Conflicts of interest

There are no conflicts to declare.

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