


Cite this: *RSC Adv.*, 2017, 7, 44827

MoO₃ subnanoclusters on ultrasmall mesoporous silica nanoparticles: an efficient catalyst for oxidative desulfurization†

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Subnano-MoO₃ supported on ultrasmall mesoporous silica nanoparticles (UMSN, ca. 14 nm) has been successfully synthesized by a reverse microemulsion method. The obtained subnano-MoO₃/UMSN exhibited superior catalytic activity for the oxidative desulfurization of model diesel. The dibenzothiophene can be oxidized to dibenzothiophene sulfone and removed simultaneously by adsorbing on the catalyst. The superiority in catalytic performance of subnano-MoO₃/UMSN can be attributed to the subnanometer size and the mesoporosity of the ultrasmall support. The origin of the unusual catalytic properties of the subnanoclusters was also explored preliminarily.

Received 3rd August 2017
Accepted 13th September 2017

DOI: 10.1039/c7ra08566d

rsc.li/rsc-advances

Introduction

Subnanoclusters are particles with diameters smaller than 1.0 nm (>0.1 nm).¹ As an intermediate size regime between nanoparticles and molecules, subnanoclusters have exhibited some new properties which are different from their nano-counterparts.² The metal subnanoclusters have been widely studied in the past decades.^{3,4} However, metal oxide subnanoclusters have just been emerging very recently.^{5,6} Both CoO_x (ref. 5) and FeO_x (ref. 6) subnanoclusters have exhibited superiority in catalytic performance for Fischer-Tropsch synthesis due to their featured atomic scale size compared to common nanoparticles. The promising catalytic performance of metal oxide subnanoclusters has evoked us to develop more active catalysts for more reaction types.

Ultra-deep desulfurization of fuels has attracted great attention because of the serious environmental problems and harsh legal requirements. Usually high temperature (>300 °C) and high pressure (>3 MPa) are needed for the traditional hydrodesulfurization (HDS).^{7,8} While the sulfur compounds which are difficult to remove in HDS, such as benzothiophene (BT), dibenzothiophene (DBT), and 4,6-dimethyldibenzothiophene (4,6-DMDBT), can be removed efficiently under mild conditions (low temperature and ambient pressure) without hydrogen through oxidative desulfurization (ODS).^{9,10} Therefore, the ODS has become a promising desulfurization method at present.^{11–16} However, the catalytic activity of the catalysts

now available cannot meet the needs for the industrialization of ODS.

MoO₃ is a commonly used ODS catalyst.^{17,18} However, as far as we know, though some MoO₃ clusters in sub nanoscale have been studied,^{19,20} not much amazing activity has been reported. And no subnano-MoO₃ has been tested for ODS catalysts. Therefore, the research for the application of subnano-MoO₃ for ODS is fueled.

On the other hand, mesoporous silica materials as excellent supports have been widely applied due to their high specific surface area and tunable channel.²¹ It is well known that decreasing the particle size of porous support to shorten the channel length could reduce the mass transfer resistance and improve the catalytic activity.²² Ultrasmall mesoporous silica nanoparticles (UMSN, <25 nm),²³ have more advantages over the conventional mesoporous silica nanoparticles (MSN) as support because the decrease of their particle size shortens the access path and increases the specific area. In addition, the mesopores of UMSN make the internal active sites accessible to some big molecules, like DBT, which are difficult to enter microporous channels due to size limitation. Thus we chose to prepare UMSN as support to improve the efficiency of mass transfer.

In this work, we synthesized subnano-MoO₃ supported on UMSN (ca. 14 nm) with “raisin-bun structure”²² by reverse microemulsion. The UMSN is among the minimal sizes of the reported UMSNs.²³ Then we used the subnano-MoO₃/UMSN for catalyzing oxidative desulfurization. The subnano-MoO₃/UMSN exhibited better catalytic activity compared to nano-MoO₃/UMSN and subnano-MoO₃/ultrasmall microporous silica. The sulfur-containing products can also be removed simultaneously *via* sorption and separation.

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† Electronic supplementary information (ESI) available: XRD pattern, ¹H NMR, and N₂ sorption isotherms. See DOI: 10.1039/c7ra08566d

Experimental section

Materials

n-Octadecyltrimethoxysilane (C₁₈TMS) was purchased from Sigma-Aldrich. Dibenzothiophene (DBT), dibenzothiophene sulfone (DBTO₂), biphenyl and decalin were purchased from Shanghai Macklin Biochemical Co., Ltd. Isopropanol and methanol were purchased from Tianjin Fuyu Fine Chemical Co., Ltd. Cyclohexane, *n*-butanol, ammonia, ammonium molybdate, tetraethyl orthosilicate (TEOS), and *tert*-butyl hydroperoxide (TBHP) were all purchased from Sinopharm Chemical Reagent Co., Ltd. of analytical grade. All chemicals were used as received without further purification.

Preparation of subnano-MoO₃/UMSN

In a typical experiment, 0.01 mol of CTAB was added into a 50 mL two-necked round-bottomed flask with 17 mL of cyclohexane and 5 mL of *n*-butanol at 35 °C under stirring. Then 0.8 mL of NH₃·H₂O and 50 μL of 0.1 mol L⁻¹ (NH₄)₆Mo₇O₂₄ aqueous solution were added to the system with stirring until the solution was transparent to form a reverse microemulsion. 30 min later, 1.8 mL of TEOS : C₁₈TMS (mole ratio 30 : 1) mixture was added. The hydrolysis process was kept for 2 h, after which 15 mL of IPA was added to demulsify the reverse microemulsion. After ultrasonication for 3 min, the solution was centrifuged at 6000 rpm for 10 min. The precipitate was washed with IPA twice, dried at 100 °C for 10 h, and calcined at 500 °C for 2 h under air stream to get the subnano-MoO₃/UMSN, designated as C-1.

To obtain nano-MoO₃/UMSN (designated as C-2), the protocol was the same as above. The only difference was the replacement of the 50 μL of 0.1 mol L⁻¹ (NH₄)₆Mo₇O₂₄ aqueous solution with 200 μL of 0.2 mol L⁻¹ (NH₄)₆Mo₇O₂₄ ammonia solution.

The subnano-MoO₃/ultrasmall microporous silica (designated as C-3) was also prepared under the same procedure of C-1 except the absence of C₁₈TMS, *i.e.*, pure TEOS was used instead of the TEOS : C₁₈TMS mixture.

Pure SiO₂ without Mo were made by removing MoO₃ with aqua regia. Typically, 150 mg of C-1 or C-3 were added into a beaker. 20 mL of aqua regia was added and stirred for 10 h. After centrifugation at 6000 rpm for 10 min, the precipitate was washed to neutral with deionized water and dried at 100 °C for 10 h. The obtained SiO₂ from C-1 and C-3 were designated as meso-SiO₂ and micro-SiO₂, respectively.

Catalysis of oxidative desulfurization

The model diesel with DBT content of 1000 ppm was prepared by dissolving 100 mg of DBT in 100 mL of decalin (1 ppm is defined as 1 mg of DBT molecular per liter of decalin). 50 mg of biphenyl was added as the internal standard of GC-FID. The mole ratios of the catalyst (MoO₃) and the oxidant (TBHP) to the substrate (DBT) were abbreviated as [cat.]/[S] and [O]/[S], respectively.

In a typical run, 2 mL of model diesel was added into a 10 mL reaction tube and heated to 70 °C under stirring. Then a certain

amount of catalyst was added to the system to make the [cat.]/[S] was 0.075 (*e.g.*, 38 mg of C-1 with a Mo content of 0.20 wt%). After stirring for 10 min, 2 mL of decalin containing 485 μL of TBHP ([O]/[S] = 6) was added. The reaction was run at 70 °C for 15 minutes. Then the reaction system was filtered to separate the catalyst. The filtrate was analyzed by GC. The conversion of DBT was investigated by GC chromatography (Agilent Technologies, 7820A GC system FID) with capillary column (HP-5, 30 m × 0.320 mm × 0.25 μm). TOF was calculated as mole of DBT converted per mole of MoO₃ per hour at the end of the reaction.

The separated catalyst can be recycled by methanol washing and drying at 100 °C for 10 h. DBTO₂ can be washed down by methanol and crystallized out when methanol was removed through rotary evaporation.

Different reaction conditions (catalyst, [cat.]/[S], [O]/[S], temperature, and time) were tested and listed in Table 2.

Characterization

The morphology and size of the solid samples were attained from field-emission transmission electron microscopy (FETEM, FEI Tecnai G2 F30). Aberration-corrected high-angle annular dark-field scanning transmission electron microscopy (HAADF-STEM) images were obtained on a JEOL JEM-ARM200F equipped with a CEOS probe corrector. XRD was performed on the X-ray diffractometer (XRD, Shimadzu Corp. XRD – 7000S) with Cu-Kα radiation ($\lambda = 1.542 \text{ \AA}$), scanning from 10° to 70° at a speed of 5° min⁻¹. The pore channel information was gained from N₂ sorption by automatic physical adsorption analyzer (Quantachrome Instruments, Autosorb-iQ-C) at 77 K. The pore size was studied by Barrett–Joyner–Halenda (BJH) method and Horvath–Kawazoe (HK) method for mesopores and micropores, respectively. The composition was detected by X-ray photoelectron spectroscopy (XPS, Thermo Fisher, ESCALAB™ 250Xi) with Al-Kα X-ray source. The binding energies were calibrated using the adventitious carbon contamination C 1s peak at 284.6 eV. And the curve fitting was performed with the software CasaXPS. The distributions of elements were attained from mapping on field-emission scanning electron microscope (FESEM, FEI, Nova NanoSEM 450). The dried samples were directly used for observation without any further treatment. The content of Mo was detected by an inductively coupled plasma atomic emission spectrometry (ICP-AES, PerkinElmer, Optima 2000 DV). ¹H NMR spectra were recorded on a Bruker Avance II 400 spectrometer using tetramethylsilane as an internal standard.

Results and discussion

Under the aforementioned synthesis process, subnano-MoO₃/UMSN was successfully prepared with uniform morphology by reverse microemulsion. The sizes of MoO₃/UMSN nanocomposites and MoO₃ clusters were shown in Fig. 1 and Table 1. According to the results of TEM image and particle size distribution histogram, the particle diameter of UMSN of C-1 was only 14.2 ± 4.9 nm and MoO₃ was unable to be seen under HRTEM. Since MoO₃ particles of 1 nm were observable,²² the



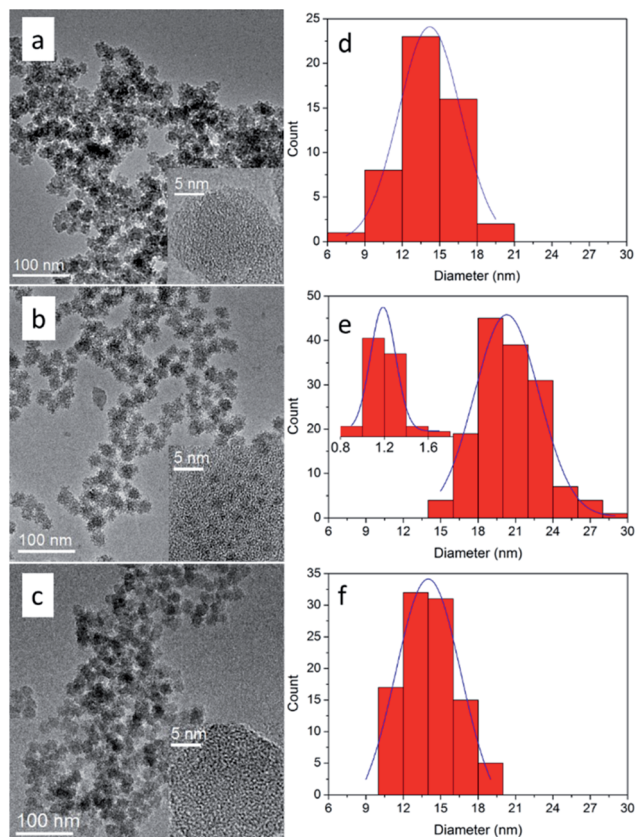


Fig. 1 Left: The TEM images of C-1 (a), C-2 (b) and C-3 (c), and the insets show corresponding HRTEM images. Right: The size distribution histograms of C-1 (d), C-2 (e) and C-3 (f), and the inset of (e) shows that of nano-MoO₃.

MoO₃ size of C-1 was supposed to be less than 1 nm, *i.e.*, subnanometer. With the increasing amount of ammonium molybdate, the MoO₃ size was raised from subnanometer to nanometer (1.2 ± 0.2 nm, C-2).

For direct observation of the subnanoclusters, we characterized C-1 and C-3 with aberration-corrected high-angle annular dark-field scanning transmission electron microscopy (HAADF-STEM). The results were shown in Fig. 2. As can be seen, there are subnanoclusters in both samples. The size of the bright dots, as marked by red circles, is 0.6 ± 0.1 nm. Thus the existence of subnano-MoO₃ was confirmed directly.

The missing of the characteristic peaks of MoO₃ in XRD (Fig. S1†) can further support the tiny size of MoO₃. However, it is still doubtful that whether the MoO₃ does exist or not. Therefore, element mapping was performed for C-1 (Fig. 3). The results showed the distributions of O, Si, and Mo elements. The

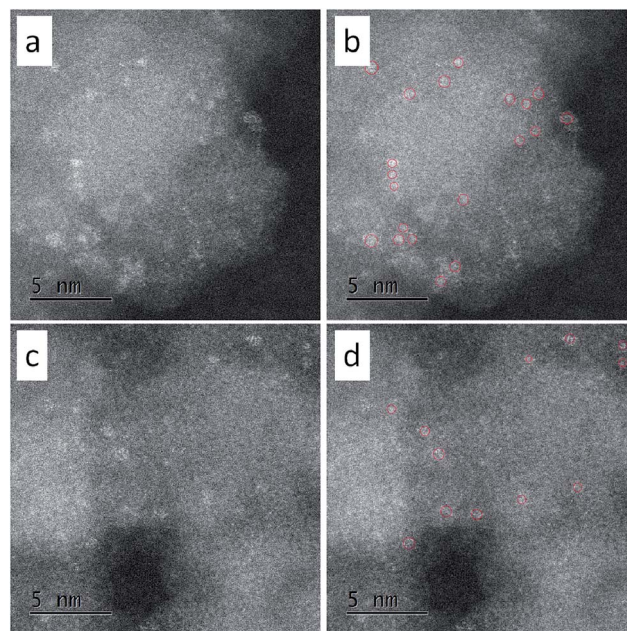


Fig. 2 The HAADF-STEM images of (a, b) C-1 and (c, d) C-3. (a) and (c) were original, (b) and (d) showed the subnanoclusters marked with red circles.

high dispersity of Mo on silica was thus confirmed undoubtedly.

But whether Mo does exist as MoO₃ is still uncertain. To prove our assumption, XPS was performed and shown in Fig. 4. In accordance with the mapping results (Fig. 3), Mo, Si, and O elements were detected again in the survey spectrum. The appearance of carbon was caused by air contamination. Fig. 4b showed the narrow scan spectrum of Si 2p with 102.9 eV which belonged to the binding energy of SiO₂. The spectrum of Mo (Fig. 4c) had two peaks of 237.1 eV and 234.0 eV, which were fitted to the binding energy of Mo 3d_{3/2} and 3d_{5/2} of MoO₃, respectively. The O 1s peaks (Fig. 4d) with 531.9 eV and 534.3 eV were matched with the O 1s of MoO₃ and SiO₂, respectively. Hence the composition of the subnano-MoO₃/UMSN can be ascertained.

The pore size is quite important for heterogeneous catalysts. So the N₂ sorption measurements were carried out. The N₂ sorption isotherms of C-1/2/3 were shown in Fig. S2† and the pore size distribution curves were shown in Fig. 5. The specific surface area (*S*_{BET}) and pore size data were listed in Table 1. The subnano-MoO₃/UMSN (C-1) owns highest specific surface area of 525 m² g^{−1} and the pore size is 2.6 nm. C-2 is also mesoporous with pore size at 2.4 nm and its specific surface area decreases to 480 m² g^{−1}. While C-3 has no mesopores, its pore sizes of 0.7 nm and 0.8 nm belong to micropore range. Meanwhile the specific surface area decreases to 404 m² g^{−1}. Obviously the specific surface area decreases with the reduction of the pore size.

The results of the N₂ sorption measurements can also help us solve another question which cannot be answered by TEM: are the MoO₃ subnanoclusters dispersed inside the UMSN or on the surface of UMSN? When subnano-MoO₃ of C-1 was removed

Table 1 The particle size, MoO₃ size, specific surface area, and pore size of the three catalysts

Catalyst	Particle size/nm	MoO ₃ size/nm	<i>S</i> _{BET} /m ² g ^{−1}	Pore size/nm
C-1	14.2 ± 4.9	0.6 ± 0.1	525	2.6
C-2	18.8 ± 5.3	1.2 ± 0.2	480	2.4
C-3	14.0 ± 5.2	0.6 ± 0.1	404	0.7, 0.8



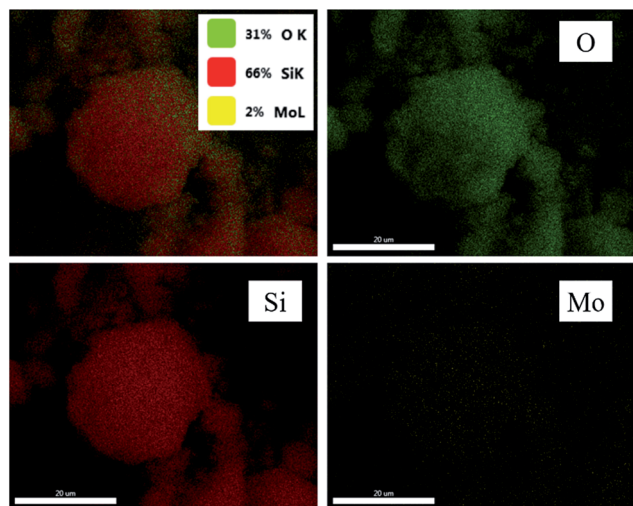


Fig. 3 The mapping results of subnano-MoO₃/UMSN.

(i.e., only UMSN left), the surface area and pore size increased to 550 m² g⁻¹ and 2.7 nm, respectively (Fig. S3†). From the comparison of the specific surface area and the pore size of the UMSN, C-1, and C-2, we can conclude that the MoO₃ subnanoclusters are dispersed inside the UMSN instead of the surface. If the MoO₃ dispersed on the surface, their appearance and growth would not cause obvious shrinking of the pore size and decrease of the specific area. Besides, the synthesis principle is similar to our former works where the same structure with multi-cores dispersed inside SiO₂ was obtained.^{22,24}

The formation mechanism of the subnano-MoO₃/UMSN was shown in Scheme 1. The spherical reverse micelles consist of surfactant (CTAB), cosurfactant (*n*-butanol), oil phase (cyclohexane), and aqueous phase (ammonia and ammonium molybdate solution).²⁵ When TEOS was added, it started hydrolyzing at the oil/water interface, forming Si-OH. Si-OH is

hydrophilic so the hydrolyzed TEOS could enter the core of the reverse micelles. Then TEOS condensed in the core to form Si-O-Si framework and the spherical silica particles emerged. After the removal of CTAB, porous silica was formed. However, without a porogen, no mesopores can be formed in silica because the alkyl chains of CTAB lie outside the core. While TEOS and C₁₈TMS (porogen) were added together, the silica particles formed based on their co-hydrolysis and co-condensation process. The hydrolyzed C₁₈TMS with three hydrophilic Si-OH can pull the hydrophobic long alkyl chains into the core of reverse micelles. After calcination, the alkyl chains were removed, leaving the wormhole-like mesopores. Based on this procedure, it seemed that a porogen is indispensable to form mesoporous silica in reverse microemulsion. The mechanism also explains why we cannot reproduce the paper forming mesopores without porogen.²⁶

The MoO₃ was formed *in situ* from the pyrolysis of the ammonium molybdate encapsulated in silica during the sol-gel process.²² The size of MoO₃ can be easily controlled through changing the precursor concentration. Subnanoclusters would form with low precursor concentration. With the increase of the addition amount of ammonium molybdate, nanoclusters would form. The size control principle is similar with our previous study on CuO in silica.²⁴

The catalytic mechanism of oxidative reaction of DBT could be described as Scheme 2. DBT is oxidized by TBHP to DBT sulfoxide then to DBTO₂. During the process, the formation of metal peroxide intermediate state is involved. According to the crystal structure analysis, the distance between the Mo atoms

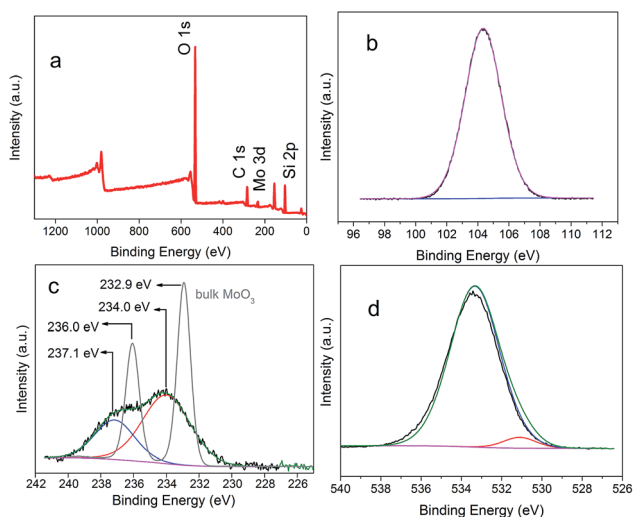


Fig. 4 Survey XPS spectrum of nano-MoO₃/UMSN (a), and narrow scan spectra of the elements of Si (b), Mo (c), and O (d). The grey line in (c) was the Mo spectra of bulk MoO₃.

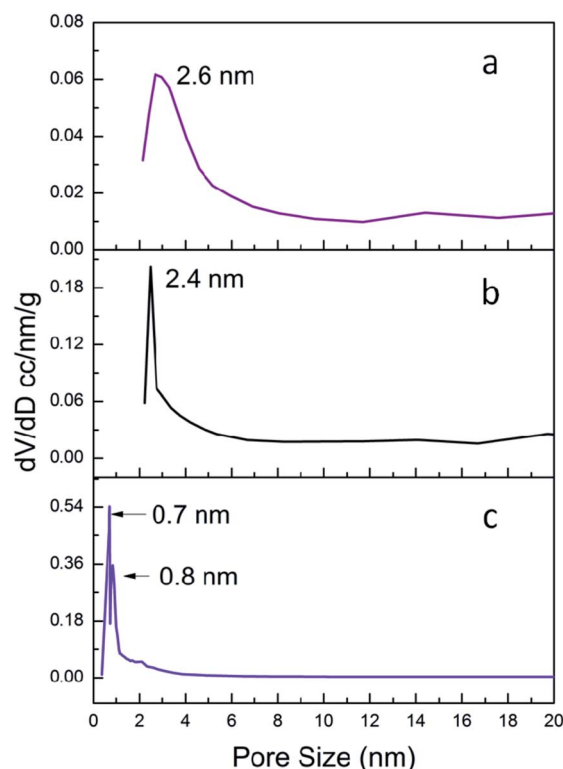
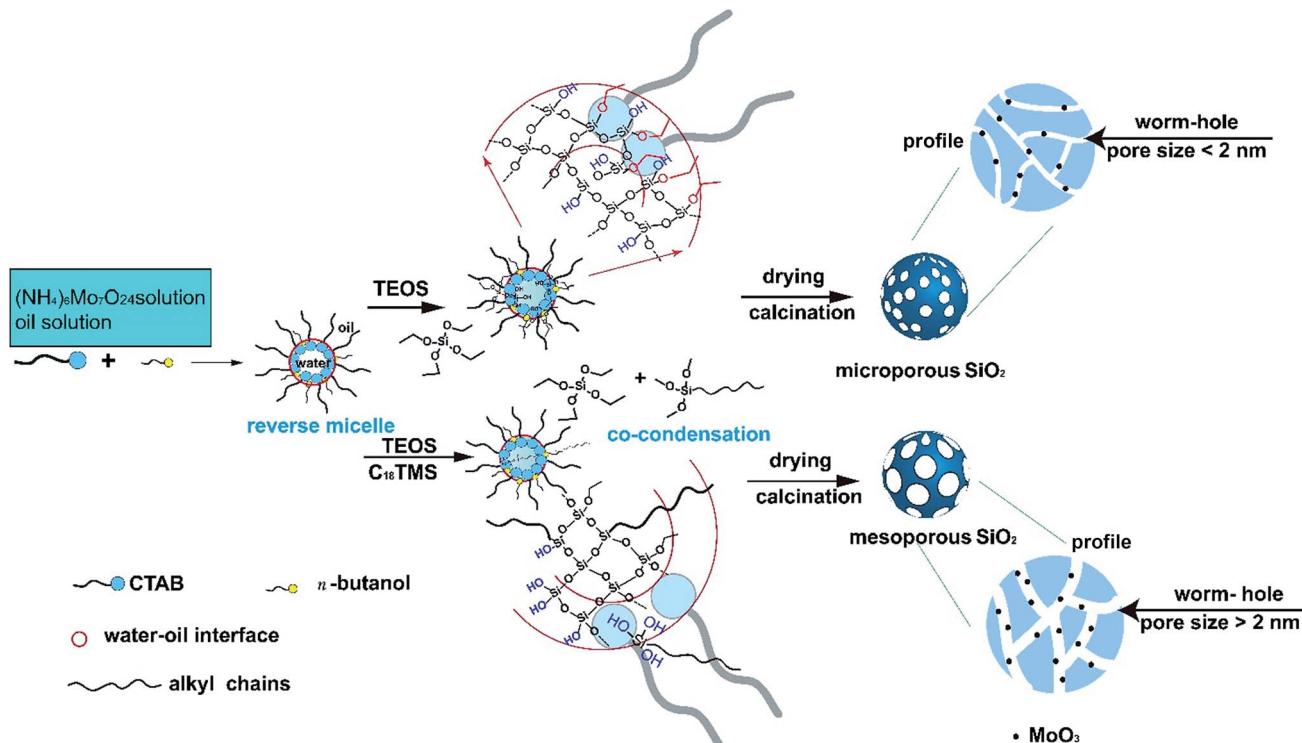


Fig. 5 Pore size distribution of C-1 (a), C-2 (b), and C-3 (c).





Scheme 1 The formation mechanism of the subnano-MoO₃/UMSN.

and the O atoms of MoO₃ is exactly matched with the distance between the tertibutyl group and the terminal hydrogen atom of TBHP.²⁷ Therefore, the metal peroxides intermediate state is easily formed with the help of MoO₃ and the oxidation of DBT could quickly happen. The polarization of Mo–O bond also promotes the formation of metal peroxide intermediate state. Moreover, the empty 3d orbital and high electronic density of S atom of DBT also made the S atom easily attacked by the O atom of hydroxyl group of TBHP.

The subnano-MoO₃/UMSN was used to catalyze oxidative desulfurization of model diesel. The results were shown in Table 2. As we can see, the reaction will not happen at all without catalyst (entry 10). The DBT conversion increased significantly with increasing catalyst amount, reaction time, and reaction temperature. While the amount of oxidant was not very critical. Based on the results of entries 1–7, an optimum condition ([cat.]/[S] = 0.075, [O]/[S] = 6, 70 °C, 15 min) was established. Under the optimum condition (entry 7), the conversion of DBT could reach 100% and the TOF was 53.3 h^{−1}.

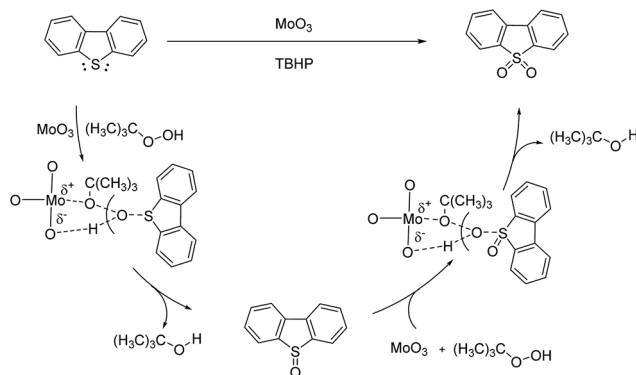
The catalytic activity of nano-MoO₃/UMSN (C-2, entry 8) was tested under the optimum condition for comparison. The conversion of DBT was 25.7% and the TOF was 10.2 h^{−1} (decreased by 80%). Since the surface area was of less difference (decreased by 8%), it can be concluded that subnano-MoO₃ is more active than nano-MoO₃ (>1 nm).

Then the catalytic activity of subnano-MoO₃/ultrasmall microporous silica (C-3, entry 9) was also tested under the optimum condition. The conversion of DBT was only 12.3% and the TOF was 6.6 h^{−1}, obviously inferior to entry 7 (C-1). Though the particle size of MoO₃ were both on subnanometer level, the

micropores may significantly block the mass transfer and thus disable the activity of the subnanoclusters.

It is worth noting that without MoO₃ subnanoclusters, both meso-SiO₂ and micro-SiO₂ exhibited low activity (entry 11, 12) even under the optimum condition. The DBT conversion data were 11.6% and 7.6% for the mesoporous silica and microporous silica, respectively. Though mesoporous silica has been reported to be active for ODS,²⁸ most of the surface silanol groups (reported active sites) in our meso-SiO₂ have been removed during the calcination process. Thus we can safely attribute the superior catalytic activity of C-1 to the existence of MoO₃ subnanoclusters.

Besides DBT, more different sulfur-containing compounds, including thiophene, benzothiophene (BT), and 4,6-



Scheme 2 The catalytic mechanism of MoO₃ catalyzing DBT oxidation with TBHP.



Table 2 Results of DBT conversion of model diesel under various reaction conditions

Entry	Catalyst	[cat.]/[S]	[O]/[S]	<i>T</i> /°C	<i>t</i> /min	DBT conversion%	TOF/h ⁻¹
1	C-1	0.05	4	70	30	82.6	33
2	C-1	0.075	4	70	30	100	27
3	C-1	0.075	2	70	15	94.0	50
4	C-1	0.075	4	70	15	94.1	50
5	C-1	0.075	6	50	15	60.2	32
6	C-1	0.075	6	60	15	96.8	52
7	C-1	0.075	6	70	15	100	53
8	C-2	0.075	6	70	15	25.7	10
9	C-3	0.075	6	70	15	12.3	7
10	—	—	6	70	15	0	—
11	Meso-SiO ₂	—	6	70	15	11.6	—
12	Micro-SiO ₂	—	6	70	15	7.6	—

dimethyldibenzothiophene (4,6-DMDBT) were also tested with C-1 under the optimum condition. The results were shown in Table S1.† The catalyst C-1 performed well in all cases.

Why is the subnano-MoO₃ more active than nano-MoO₃? It is well accepted that the catalytic activity usually increases with decreasing size of the particles.²⁹ The smaller the size is, the higher ratio of surface-to-bulk atoms (coordinatively unsaturated and normally function as active sites) becomes. However, such a significant difference between C-1 and C-2 cannot be just attributed to the little size discrepancy. The electron structure variation should be taken into account. Let us revisit the XPS narrow scan of Mo in Fig. 3c. While the nano-MoO₃ (≥1 nm) has a binding energy very close to the bulk,²² the binding energy of subnano-MoO₃ showed a shift of 1.1 eV higher than that of the bulk MoO₃ (the grey line), the same trend as that observed for small gold clusters.³⁰ This phenomenon might be attributed to the reduced core-hole screening in smaller clusters.³¹ It seems that the electronic properties of the subnanoclusters are significantly different from those of the nanoclusters. Thus the size-dependent alteration of electronic structure may be the primary reason for the unusual catalytic properties.^{17,32} The exact mechanism remains unclear and will likely fuel new research activities.

Then the system after the reaction was analyzed. To our surprise, neither DBT nor DBTO₂ could be detected in the solution after the separation of the catalyst. Then we found that the resultant DBTO₂ was adsorbed on the catalyst surface, just like what have been reported in Dou and Zeng's work.³³ Then methanol was used to wash the spent catalyst adequately. After removing methanol by rotary evaporation, white crystal was obtained, which was identified as DBTO₂ by ¹H-NMR (Fig. S4†). Isolated yield was 98.3% for C-1 under the optimal conditions, which means almost all of the DBTO₂ had been removed with the separation of the catalyst. This could be explained by the similarity–intermiscibility theory. While DBT as a nonpolar substance is easily soluble in nonpolar decalin, polar DBTO₂ tends to separate from decalin and adsorb on silica, which has some polarity due to the existence of residual silanol groups. While methanol with strong polarity can wash DBTO₂ out easily.

Finally, the stability of C-1 was studied. If the subnano-MoO₃ is not structured but supported in silica, which means the MoO₃

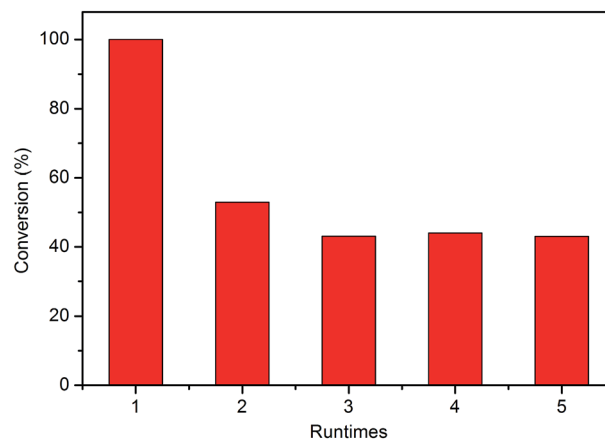


Fig. 6 The variation of DBT conversion with runtimes.

subnanoclusters were held by van der Waals forces, MoO₃ may have leached out into solution during the reaction. The catalyst was recycled by methanol washing and drying. Totally 5 cycles were performed and the results were shown in Fig. 6. As expected, the conversion decreased significantly from nearly 100% to 52.9% when regenerated once. The Mo content decreased from 0.20 wt% to 0.09 wt% according to ICP-AES for the first regeneration. This could also confirm that the subnano-MoO₃ is not structured in silica. The poor stability of C-1 remains a problem to be solved, which is underway in our lab and some preliminary results have been obtained. Anyhow, the high activity of subnanoclusters still presents promising perspectives.

For comparison, the stability properties of C-2 and C-3 were also studied. The results were shown in Fig. S5.† Both catalysts exhibited the same trend as C-1, indicating the existence of leaching.

Conclusions

To sum up, the subnano-MoO₃/UMSN was synthesized successfully by reverse microemulsion. The size of MoO₃ was affected by the amount of ammonium molybdate. The subnano-MoO₃/UMSN was used to catalyze the oxidative desulfurization



of DBT. The DBT conversion to DBTO₂ was up to 100% within 15 min and the TOF of oxidative desulfurization was 53.3 h⁻¹. The subnano-MoO₃/UMSN can also serve as effective adsorbents for removal of DBTO₂. The subnano-MoO₃/UMSN exhibited higher activity than nano-MoO₃/UMSN and subnano-MoO₃/ultrasmall microporous silica, which demonstrated the superiority of both the subnano size of MoO₃ and the mesoporosity of the ultrasmall support. The size-dependent alteration of electronic structure was believed to be the primary reason for the unusual catalytic properties.

Conflicts of interest

There are no conflicts of interest to declare.

Acknowledgements

The authors gratefully acknowledge the financial support of National Natural Science Foundation of China (No. 21506026), Doctoral Scientific Research Foundation of Liaoning Province (No. 201501171), and the China Postdoctoral Science Foundation funded project (No. 2016M601311).

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