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# Palladium-catalyzed three-component tandem reaction of sulfonyl hydrazones, aryl iodides and allenes: highly stereoselective synthesis of $(Z)-\alpha$ -hydroxymethyl allylic sulfones†

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Sulfonyl hydrazones have been identified as an excellent sulfonyl anion surrogate in the base and Pd<sup>0</sup>-catalyzed three-component tandem reaction with aryl iodides and allenes for the synthesis of functionalized allylic sulfones. By forming a stabilized six-membered palladacycle intermediate, the reaction led to the desired higher substituted allylic sulfones with excellent Z selectivities and high yields.

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Selective construction of carbon-heteroatom bonds is an ongoing challenge in organic synthesis. Although a variety of methods for forming C–O and C–N bonds have been developed, similar reactions to form C–S bonds have received comparably less attention.¹ However, the formation of C–S bonds is a fundamental transformation in the synthesis of organosulfur compounds. Sulfone, the most attractive sulfur-containing fragment, has emerged as a building block in several bioactive natural products, pharmaceuticals and agrochemicals (Fig. 1, I–IV).² For example, intepirdine (Fig. 1, III), a selective 5-HT6 receptor antagonist, is currently undergoing phase III clinical trials for application in different types of Alzheimer's disease (AD),³ while rigosertib (Fig. 1, IV) is a potent, polo-like kinase 1 (PLK1) pathway

modulator for treating lymphocytic leukemia.<sup>4</sup> In addition, sulfones are widely used as polymeric compounds (Fig. 1, **V**) and intermediates in organic synthesis.<sup>5</sup> Accordingly, the efficient incorporation of a sulfonyl group into organic molecules has attracted considerable interest among organic chemists.

Traditional sulfone syntheses are mainly based on the oxidation of the corresponding sulfide,<sup>6</sup> or the reaction of either organomagnesium halides or organolithiums with sulfonate esters (Scheme 1, eqn (1)).<sup>7</sup> However, these methods have certain drawbacks, including limited substrate sources, toxic by-products, and harsh reaction conditions. Recently, sulfonyl

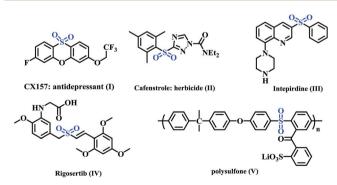
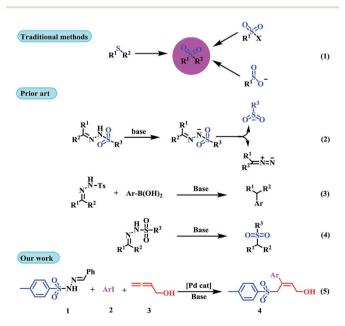


Fig. 1 Representative medicinally relevant agents, and polymeric agent containing sulfones.

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**Scheme 1** Strategies for the synthesis of sulfones and studies on the sulfonyl hydrazones.

hydrazones are well-known as useful synthons in organic synthesis, and they can be easily prepared in solid form through condensation of ketones or aldehydes with sulfonyl hydrazide in high yields. Sulfonyl hydrazones are in general stable, and in many cases they are crystalline compounds and easy to purify, and used as an *in situ* source of diazo compounds in a variety of transition-metal-catalyzed and transition-metal-free reactions, which affords novel methodologies for the formation of carboncarbon and carbon-heteroatom bonds (Scheme 1, eqn (2)).8 Valdés and co-workers have reported the metal-free coupling of tosylhydrazones with boronic acids to form C-C bonds (Scheme 1, eqn (3)).9 More recently, Yu et al. and Zhao et al. explored a new organocatalytic strategy to synthesize sulfones via loss of intramolecular nitrogen from sulfonyl hydrazones (Scheme 1, egn (4)). 10 Despite the fact that it is well-known in the literature that sulfonyl hydrazones undergo dissociation in the presence of a base to generate a sulfinate anion and a diazo compound, all of the reported methodologies have used this reactivity for the generation of carbene species from the released diazoalkane reagent. In this sense, literature precedent illustrating the possibility of using this behavior for the generation of an Sbased nucleophile (sulfonyl anion) that can be used in a subsequent transformation are rare.11

Multi-component reactions (MCRs) involving domino processes have emerged as powerful tools for the efficient and atom-economic formation of a diverse range of carbonheteroatom bonds. In addition, functionalized allenes have been established as starting materials for the efficient synthesis of potentially useful carbo- and heteroatom bonds. Drawing from our recent experience in the field of sulfonyl hydrazones and allenes, If and with regard to the importance of sulfones in organic synthesis, we herein report the ability of sulfonyl hydrazones to participate as a source of a sulfonyl anion equivalent towards (Z)- $\alpha$ -hydroxymethyl allylic sulfones by the assembly of aryl iodides (2) and buta-2,3-dien-1-ol (3) (Scheme 1, eqn (5)).

Initially, we investigated the model reaction of sulfonyl hydrazones (1) with iodobenzene (2) and buta-2,3-dien-1-ol (3) in the presence of 5 mol% of [Pd(OAc)<sub>2</sub>] at 80 °C in THF (Table 1). Using these conditions, and with  $K_2CO_3$  as the base, allylic sulfone 4 was obtained in 23% yield. The effect of the ligand on the catalytic activity was also investigated, and the presence of PPh<sub>3</sub> was found to promote the catalytic reaction (Table 1, entries 1 vs. 2-5). In order to improve the yield, different catalysts (Pd(PPh<sub>3</sub>)<sub>4</sub>, Pd(dppf)Cl<sub>2</sub>, and Pd(PPh<sub>3</sub>)<sub>2</sub>Cl<sub>2</sub>) were tested (Table 1, entries 6–8). Use of Pd(PPh<sub>3</sub>)<sub>4</sub> (5 mol%) significantly improved the yield to 68%. No product was detected in the absence of a catalyst (Table 1, entry 9). To further improve the reaction efficiency, several bases were evaluated (Table 1, entries 10-14). As shown, Na<sub>2</sub>CO<sub>3</sub>, CsF, or Cs<sub>2</sub>CO<sub>3</sub> showed no improvement compared with K<sub>2</sub>CO<sub>3</sub>. The use of stronger bases (e.g., t-BuOLi and NaOH) did not result in any allylic sulfone 4 at all. Based on previous studies, different additives were investigated. Using either the phase-transfer catalyst n-Bu<sub>4</sub>NBr (Table 1, entry 15) or the Lewis acid (Cu(OTf)<sub>2</sub>) (Table 1, entry 16) did not increase the yield, nor did increasing the number of equivalents of allene 3 from 1.2 to 3.0 (Table 1, entries 5 vs. 17).

Table 1 Optimization of reaction conditions

Entry	Cat.	Base	Additive	Yield <sup>b</sup> (%)
1	Pd(OAc) <sub>2</sub>	$K_2CO_3$	_	23
2	Pd(OAc) <sub>2</sub>	$K_2CO_3$	$PPh_3$	41
3	Pd(OAc) <sub>2</sub>	$K_2CO_3$	dppf	28
4	Pd(OAc) <sub>2</sub>	$K_2CO_3$	XantPhos	20
5	Pd(OAc) <sub>2</sub>	$K_2CO_3$	BINAP	25
6	$Pd(PPh_3)_4$	$K_2CO_3$	_	68
7	$Pd(PPh_3)_2Cl_2$	$K_2CO_3$	_	42
8	Pd(dppf)Cl <sub>2</sub>	$K_2CO_3$	_	44
9		$K_2CO_3$	_	n.d.
10	$Pd(PPh_3)_4$	$Na_2CO_3$	_	n.d.
11	$Pd(PPh_3)_4$	CsF	_	48
12	$Pd(PPh_3)_4$	$Cs_2CO_3$	_	39
13	$Pd(PPh_3)_4$	t-BuOLi	_	Trace
14	$Pd(PPh_3)_4$	NaOH	_	n.d.
15	$Pd(PPh_3)_4$	$K_2CO_3$	<i>n</i> -Bu₄NBr	63
16	$Pd(PPh_3)_4$	$K_2CO_3$	$Cu(OTf)_2$	55
17 <sup>c</sup>	$Pd(PPh_3)_4$	$K_2CO_3$		69
$18^d$	$Pd(PPh_3)_4$	$K_2CO_3$	_	23

 $<sup>^</sup>a$  Reaction conditions: under  $\rm N_2$  atmosphere, 1 (1 equiv.), 2 (1.2 equiv.), 3 (1.2 equiv.), [Pd Cat.] (5 mol%), base (3 equiv.), THF (5.0 mL).  $^b$  Isolated yield.  $^c$  Allenes 3 (3.0 equiv.) was used.  $^d$  Under air.

When the reaction was conducted in air instead of nitrogen, a relatively lower yield was obtained (Table 1, entry 18).

The effect of the solvent and reaction temperature was also examined (Table 2). It was found that the solvent played a critical role in this process (Table 2, entries 1–6). The tested solvents, including DMSO, DMF, toluene, ethyl acetate, and DCM, did not show any improvement in the transformation yield. However, when the reaction was conducted in dioxane, 4 was obtained in a relatively higher yield of 75% (Table 2, entry 7). When the reaction temperature was raised to 100 °C, the rate and conversion were not significantly affected (Table 2, entry 8). However, conversion decreased upon lowering the temperature

Table 2 Solvent and temperature effect on the  $Pd^0$ -catalyzed three-component tandem reaction of 1, 2, and  $3^a$ 

Entry	Solvent	$T$ [ $^{\mathrm{o}}$ C]	Time (h)	Yield <sup>b</sup> (%)
1	THF	Reflux	12	68
2	DMF	80	12	41
3	DMSO	80	12	48
4	Ethyl acetate	Reflux	12	25
5	Toluene	80	12	33
6	DCM	Reflux	24	n.d.
7	Dioxane	80	12	75
8	Dioxane	100	10	76
9	Dioxane	40	48	28

<sup>&</sup>lt;sup>a</sup> Reaction conditions: under  $N_2$  atmosphere, **1** (1 equiv, **2** (1.2 equiv.), **3** (1.2 equiv.), [Pd(PPh<sub>3</sub>)<sub>4</sub>] (5 mol%),  $K_2CO_3$  (3 equiv.), solvent (5.0 mL). <sup>b</sup> Isolated yield.

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to 40 °C, even when the reaction time was extended to 48 h (Table 2, entry 9). Consequently, the standard reaction conditions used in further investigations were 1 (1 equiv.), 2 (1.2 equiv.), and 3 (1.2 equiv.) with K<sub>2</sub>CO<sub>3</sub> (3 equiv.) in dioxane at 80 °C, using [Pd(PPh<sub>3</sub>)<sub>4</sub>] (5 mol%) as the catalyst.

With the optimal conditions in hand, the scope of this protocol was explored for a series of sulfonyl hydrazones, accessible from the corresponding substituted aldehydes with p-toluenesulfonyl hydrazide as a standard substrate (Table 3). Reaction with an electronically diverse range of aryl hydrazones provided the corresponding allylic sulfones 4 in moderate to good yields. Electron-rich substrates provided higher yields, and the reaction time could be reduced to 8 h (Table 3, entries 5 vs. 3). Use of a heteroaryl group (Table 3, entry 7,  $R^2 = 2$ -thienyl), slightly reduced the yield to 65%, perhaps due to the electronic properties of the thiophene ring. Thus, sulfonyl hydrazone 1a, derived from p-methoxybenzaldehyde, was used as the standard substrate for further investigations. Furthermore, sulfonyl hydrazones 1b-1h, derived from aryl sulfonyl hydrazide and pmethoxybenzaldehyde, were also suitable substrates for the assembling reaction, and furnished the corresponding allylic sulfones in 65–85% yields (Table 3, entries 8–14). However, pnitrophenyl sulfonyl hydrazone 1i and methyl sulfonyl hydrazone 1j gave only trace amounts of products 4ai and 4aj, respectively, making it unsuitable for further investigation (Table 3, entries 15-16). In order to investigate the scope of allenes, ethyl 2,3-butadienoate and 1,2-pentadiene were evaluated as representative compounds. As presented in Table 3 (entries 17-18), no product was observed, which suggested that the terminal substituent group of the allenes had a great effect on the three-component tandem reaction. The configuration of the product was supported by <sup>1</sup>H NOESY analysis of **4aa** as the Zisomer (ESI, SI†), and those of the other products were assigned by analogy.

The scope and limitations of the transformation were also evaluated for the reaction of 1a with various aryl iodides 2 (Table 4), under the optimized reaction conditions. Electronwithdrawing groups on the phenyl ring decreased the reaction efficiency, suggesting the importance of an electron-rich phenyl ring (Table 4, 4ea, 4fa, 4ja). Further studies showed that steric hindrance of the aryl iodides had a negative effect on the reaction. For example, p-iodoanisole afforded the desired product 4ca in 88% yield under the optimized conditions. However, the use of o-iodoanisole dramatically decreased the conversion, perhaps due to the steric hindrance of the methoxy group (Table 4, 4da). Interestingly, the heterocyclic compound 2-iodothiophene reacted smoothly, affording the corresponding allylic sulfone 4la in 80% yield. 1-Iodonaphthalene gave only trace amounts of product 4na, probably because its rigid and bulky structure disfavors the transformation. However, diversely substituted aliphatic iodides were not suitable substrates (Table 4, 40a).

In order to investigate the limitations of aryl halides, phenyl bromide and phenyl chloride were evaluated as representative compounds. As shown in Table 5, phenyl bromide 2b also reacted with 1a and 3a to give the expected products 4aa, although the yields were slightly lower than the corresponding aryl iodides

Table 3 Pd<sup>0</sup>-catalyzed three-component tandem reaction of iodobenzene 2a, 3a, and different sulfonyl hydrazones<sup>a</sup>

(Pd(PPh.).) (5 mol%)

_	O N=/ S-NH + PhI +	OH [Pd(PPh <sub>3</sub> ) <sub>4</sub> ] (5 mol% K <sub>2</sub> CO <sub>3</sub> (3.0 equiv.) Dioxane, 80 °C	$\xrightarrow{\circ}$	ОН
	1 2a 3	Bioxane, 80°C	4	
Entry	1	4	Time (h)	Yield <sup>b</sup> (%)
		Ph		
1	$(R^2 = Ph)$	OH 4aa	12	75
2	$(R^2 = p\text{-NO}_2\text{C}_6\text{H}_4)$	4aa	16	70
3	$(R^2 = p - CF_3C_6H_4)$	4aa	16	65
4	$(R^2 = p\text{-MeC}_6H_4)$	4aa	12	73
5	$(R^2 = p\text{-MeOC}_6H_4)$	4aa	8	78
6	$(R^2 = p\text{-BrC}_6H_4)$	4aa	16	70
7	$(R^2 = 2\text{-thienyl})$	4aa	18	65
8	O-NH	OH 4ab	8	65
9	O-NH	O-OH 4ac	8	81
10	H,COCHN—	H <sub>3</sub> COCHN OH	8	79
11	F-O-NH	F—OH	8	78
12	CI-NH	CI—OH	8	80
13		CI—CI—Ph OH	8	73
14	O- O-NH	OH OH	8	85
15	O <sub>2</sub> N — O N	O <sub>2</sub> N OH	24	Trace
16	O-NH O-NH	OH OH 4aj	24	Trace
17 <sup>c</sup>	$(R^2 = p\text{-MeOC}_6H_4)$	4ak COOEs	24	Trace
18 <sup>d</sup>	$(R^2 = p\text{-MeOC}_6H_4)$	4al O Ph	24	Trace

<sup>&</sup>lt;sup>a</sup> Reaction conditions: under  $N_2$  atmosphere, 1 (1 equiv.), 2 (1.2 equiv.), 3 (1.2 equiv.),  $[Pd(PPh_3)_4]$  (5 mol%),  $K_2CO_3$  (3 equiv.), dioxane (5.0 mL). b Isolated yield. c 3a = ethyl 2,3-butadienoate. d 3a = 1,2-pentadiene.

(Table 5, entry 2). Phenyl chloride 2c and phenyl fluoride 2d with 1a and 3a gave irreproducible results (Table 5, entries 3 and 4). These results indicate that the cleavage of the C-I and C-Br bonds is much faster than that of the C-Cl and C-F bonds.

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Table 4 Pd<sup>0</sup>-catalyzed three-component tandem reaction of 1a, 3a, and different aryl iodides<sup>a</sup>

<sup>a</sup> Reaction conditions: under N<sub>2</sub> atmosphere, 1 (1 equiv.), 2 (1.2 equiv.), 3 (1.2 equiv.), [Pd(PPh<sub>3</sub>)<sub>4</sub>] (5 mol%), K<sub>2</sub>CO<sub>3</sub> (3 equiv.), dioxane (5.0 mL).

Table 5 Pd<sup>0</sup>-catalyzed three-component tandem reaction of 1, 3, and different aryl halides

Entry	2	4	Time (h)	Yield <sup>b</sup> (%)
1	2a	4aa	8	78
2	Br 2b	4aa	16	65
3	CI 2c	4aa	48	Trace
4	F	4aa	48	Trace

<sup>&</sup>lt;sup>a</sup> Reaction conditions: under N<sub>2</sub> atmosphere, 1 (1 equiv.), 2 (1.2 equiv.), 3 (1.2 equiv.), [Pd(PPh<sub>3</sub>)<sub>4</sub>] (5 mol%), K<sub>2</sub>CO<sub>3</sub> (3 equiv.), dioxane (5.0 mL). Isolated yield.

Encouraged by this three-component tandem reaction and sustainable synthesis, we conducted scale-up experiments to examine the synthetic utility of the reaction. When 3.6 mmol of 4-methylphenyl-sulfonylhydrazide 1a was treated with 4.4 mmol

aryl iodides (2a) and 4.4 mmol buta-2,3-dien-1-ol (3a), the corresponding product 4aa was obtained with a satisfactory yield of 75%, although an extended reaction time was required (Scheme 2).

To get an insight into the mechanism of the sulfonylation process, several controlled experiments were conducted. Firstly, radical trapping experiments were conducted to elucidate whether the reaction involves radical species. When the radical scavenger (2,2,6,6-tetramethyl-piperidinyloxy, TEMPO; butylated hydroxytoluene, BHT) was employed under standard conditions, the reaction afforded 4aa in 75% and 72% yields, which indicated that the transformation did not proceed via a free-radical pathway (Scheme 3, eqn (1) and (2). Subsequently, when using p-toluene sulfinic acid sodium salt as the substrate instead of 1a under standard conditions, the reaction proceeded with a 58% yield (Scheme 3, eqn (3)), which indicated that the sulfinate anion should be an intermediate in this transformation. Finally, the corresponding reaction of 4methoxybuta-1,2-diene 3a', which can be viewed as the hydroxy group in 3a was protected as the methoxy group, with PhI (2a) and sulfonyl hydrazine (1a) did not afford the corresponding (Z)-allylic sulfones 4aa' (Scheme 3, eqn (4)).

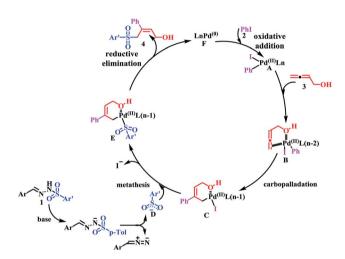
On the basis of the aforementioned investigations and previous reports, 10,14 a plausible mechanism for the crosscoupling reaction is proposed in Scheme 4. First, oxidative addition of iodobenzene (2) gives PdII intermediate A. Presumably, buta-2,3-dien-1-ol (3) plays two roles: (1) enhancing the reactivity of the allene component by coordinating to the palladium-aryl species A prior to insertion; and (2) controlling the regioselectivity of the allene insertion by forming a stabilized six-membered palladacycle intermediate C. Then, N-tosylhydrazones (1) undergo thermal decomposition in the presence of a base to generate diazo compound and sulfinate anion D. Metathesis of C with sulfinate anion D furnishes intermediate E. Finally, intermediate E converts to product 4 via reductive elimination, regenerating the Pd<sup>0</sup> catalyst F. The high stereoselectivity for the (Z)- $\alpha$ -hydroxymethyl allylic sulfone construction might be due to the formation of a stabilized, sixmembered palladacycle intermediate C. In order to investigate this hypothesis, ethyl 2,3-butadienoate, 1,2-pentadiene and 4methoxybuta-1,2-diene instead of buta-2,3-dien-1-ol (3) were evaluated. As presented in Table 3 (entries 17-18) and Scheme 3 (eqn (4)), no product was observed, which suggested that the terminal hydroxy group of the allenes played an important role on the three-component tandem reaction.

Scheme 2 Large scale experiment: under a N2 atmosphere, 1a (3.6 mmol), 2a (4.4 mmol), 3a (4.4 mmol), [Pd(PPh<sub>3</sub>)<sub>4</sub>] (5 mol%), K<sub>2</sub>CO<sub>3</sub> (10.8 mmol) in 20 mL of dioxane, 80 °C, 12 h. Product 4aa was isolated in 75% yield

Ethyl, 1-Butyl

Scheme 3 Control experiments.

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Scheme 4 Plausible reaction mechanism

### Conclusions

An efficient,  $Pd^0$ -catalyzed MCR for the synthesis of substituted (Z)- $\alpha$ -hydroxymethyl allylic sulfone derivatives has been developed. This reaction tolerates a broad range of functional groups in sulfonyl hydrazones with aryl iodides and buta-2,3-dien-1-ol to afford various allylic sulfones. In addition, we have demonstrated that sulfonyl hydrazones can be used as an excellent sulfonyl anion surrogate for performing tandem reactions. As all three building blocks are readily available, this study is expected to advance the transition-metal-catalyzed chemistry of allenes. Further studies regarding the scope and synthetic applications of this reaction are being pursued in our laboratory and will be reported in due course.

## **Experimental section**

Representative procedure for the synthesis of (Z)- $\alpha$ -hydroxymethyl allylic sulfone products: buta-2,3-dien-1-ol **3a** (28 mg, 0.39 mmol) and iodobenzene **2a** (81 mg, 0.39 mmol, 1.2 equiv.)

were consecutively added to a sealed tube charged with a mixture of  $K_2CO_3$  (136 mg, 0.99 mmol, 3.0 equiv.),  $[Pd(PPh_3)_4]$  (19 mg, 0.018 mmol, 5 mol%), and sulfonyl hydrazone 1a (100 mg, 0.33 mmol, 1 equiv.) in dioxane (5 mL), under an atmosphere of nitrogen. The reaction mixture was stirred at 80 °C for 8 h and analyzed by TLC. After the reaction was complete, water (10 mL) was added, and the solution was extracted with dichloromethane. The organic phase was separated, washed with brine, dried (MgSO<sub>4</sub>), filtered, and concentrated *in vacuo* to give the crude product, which was purified by column chromatography on silica gel with a mixture of ethyl acetate/petroleum (10:10, v/v) to afford the desired product 4aa.

#### Conflicts of interest

There are no conflicts to declare.

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