



Cite this: RSC Adv., 2017, 7, 44773

## Utility of 4-chloro-7-nitrobenzo-2-oxa-1,3-diazole for development of a highly sensitive stability indicating spectrofluorimetric method for determination of salmeterol xinafoate; application to human plasma

Mahmoud A. Omar, <sup>a</sup> Mohamed A. Hammad<sup>\*a</sup> and Mohamed Awad<sup>b</sup>

A new, simple and rapid spectrofluorimetric method was developed and validated for determination of salmeterol xinafoate in dosage form and spiked plasma. The method is considered as the first attempt for the spectrofluorimetric determination of the investigated drug. The proposed method is based on nucleophilic substitution reaction between salmeterol xinafoate and 4-chloro-7-nitrobenzo-2-oxa-1,3-diazole (NBD-Cl) at pH (9) to form a highly fluorescent product that can be measured at 543 nm after excitation at 473 nm. The fluorescence–concentration plot was rectilinear over the concentration range (100–1500 ng mL<sup>-1</sup>) with a detection limit of 19 ng mL<sup>-1</sup>. The suggested method was validated according to ICH guidelines and successfully applied for determination of salmeterol xinafoate in its pharmaceutical dosage form and spiked human plasma. The developed method was further extended for a stability study of salmeterol xinafoate under different stress circumstances including alkali, acid, oxidative and photodegradation conditions and the method was proved to be able to determine the intact drug in the presence of its degradation products.

Received 22nd July 2017

Accepted 13th September 2017

DOI: 10.1039/c7ra08106e

rsc.li/rsc-advances

### 1. Introduction

Salmeterol xinafoate (SAL), (RS)-1-[4-hydroxy-3-(hydroxymethyl)phenyl]-2-[[6-(4-phenylbutoxy)hexyl]amino]ethanol-1-hydroxynaphthalene-2-carboxylate (Fig. 1), is a direct-acting sympathomimetic agent with  $\beta$ 2-adrenoceptor stimulant activity and a selective action on  $\beta$ 2 receptors. When given by oral inhalation, it acts as a bronchodilator with a long acting action up to 12 hours. It is used for the treatment of asthma and chronic obstructive pulmonary disease.<sup>1</sup> Many analytical techniques in the literature have been reported for the determination of salmeterol xinafoate including spectrophotometric,<sup>2–8</sup> HPLC,<sup>9,10</sup> and HPTLC.<sup>11</sup> These methods were associated with some major drawbacks such as inadequate sensitivity, being time consuming and/or using sophisticated and expensive instruments. On the other hand, spectrofluorimetric methods have the advantages of being simple, highly sensitive and time saving technique.<sup>12–16</sup> To our knowledge, there is no reported spectrofluorimetric method for determination of salmeterol xinafoate up till now. In this study, we investigate the reaction of salmeterol xinafoate that contains a basic secondary

amine with NBD-Cl which has been widely used for the determination of many pharmaceuticals containing primary and secondary amino groups<sup>17–22</sup> in order to develop new, simple, sensitive and rapid spectrofluorimetric method for determination of the studied drug in its dosage form, spiked plasma and in the presence of its degradation products. Thus, the proposed method can be considered as a simple and fast alternative to the already existing stability-indicating HPLC procedures.

### 2. Experimental

#### 2.1. Apparatus

Fluorescence spectrometer FS-2 (Scinco, Korea), connected to Dell PC, equipped with 1 cm quartz cell, grating excitation and emission monochromators with slit widths set at 5 nm, PMT

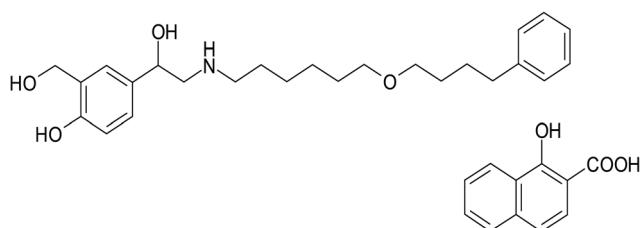


Fig. 1 Chemical structure of salmeterol xinafoate (SAL).

<sup>a</sup>Analytical Chemistry Department, Faculty of Pharmacy, Minia University, Minia, Egypt. Tel: +201098986660; +201063118597. E-mail: momar71g@yahoo.com; m\_abdelkhalek\_eg@mu.edu.eg

<sup>b</sup>Pharmaceutical Analytical Chemistry Department, Faculty of Pharmacy, Al-Azhar University, Assiut Branch, Assiut, Egypt

voltage 400 volt, Jenway pH meter model 350 (E.U), Laboratory centrifuge (Bremsen ECCO, Germany), A Shimadzu UV-1601 PC spectrophotometer (Tokyo, Japan) with 1 cm quartz cell, MLW type thermostatically controlled water bath (Memmert GmbH, Schwabach, Germany) and Digital analytical balance (AG 29, Mettler Toledo, Glattbrugg, Switzerland).

## 2.2. Materials and reagents

All materials and reagents used through this study were of analytical reagent grade salmeterol xinafoate (SAL) pure sample (99.8%) was obtained from The Arab Drug Company for pharmaceutical and chemical industries CO., ADCO, (Cairo, Egypt). Metrovent® inhaler (batch no. 490738) (Arab Drug Company for pharmaceutical and chemical industries CO., ADCO, Cairo, Egypt) labeled to contain 25 µg salmeterol xinafoate per actuation.

4-Chloro-7-nitrobenzo-2-oxa-1,3 diazole (NBD-Cl) was purchased from Sigma Aldrich Chemie GmbH (Steinheim, Germany) and NBD-Cl reagent ( $2.5 \times 10^{-3}$  M) was prepared in methanol. All other reagents used such as methanol, ethanol, acetone, acetonitrile, *N,N*-dimethylformamide (DMF), phosphoric acid, citric acid, boric acid, borax, acetic acid, sodium hydroxide and hydrochloric acid were purchased from El-Nasr chemical CO., (Cairo, Egypt). Different types of buffer solutions such as Teorell and Stenhagen, 0.1 M borate and phosphate buffers of different pH ranges were prepared in distilled water.<sup>23</sup> Human plasma samples were obtained from healthy volunteers at Assiut University Hospital according to institutional guidelines. In all cases, informed consent was obtained from all participants and plasma samples were kept frozen at  $-20^{\circ}\text{C}$  until assay after gentle thawing.

## 2.3. Standard drug solution

Stock standard solution ( $100.0 \mu\text{g mL}^{-1}$ ) of salmeterol xinafoate was prepared in methanol. Further dilutions with methanol were made as appropriately to obtain final concentration range of ( $100\text{--}1500$ )  $\text{ng mL}^{-1}$ .

## 2.4. General analytical procedure

Into a series of 10 mL stoppered test tubes, different volumes of (SAL) working solution were added followed by 2.0 mL of borate buffer and 0.5 mL of NBD-Cl reagent ( $2.5 \times 10^{-3}$  M), the tubes were heated at thermostatically controlled water bath fixed at  $80^{\circ}\text{C}$  for 20 min. After that the tubes were cooled in ice bath and 1.0 mL of 2 M HCl was added, the contents of the tubes were quantitatively transferred into 10 mL volumetric flasks and completed to the volume with methanol. The fluorescence intensity of the reaction product was measured at 543 nm after excitation at 473 nm; blank experiment was performed similarly omitting the drug.

## 2.5. Application of the proposed method for determination of SAL in its pharmaceutical formulation

Metrovent® inhaler labeled to contain 25.0 µg SAL per actuation was shaken well and 200 actuations (the whole content of the

inhaler) were actuated in 100 mL beaker containing 40 mL of methanol. The mouth piece of the inhaler was immersed beneath the methanol to ensure complete delivery of the drug without any waste. The contents of the beaker were gently warmed at  $50^{\circ}\text{C}$  in thermostatically controlled water bath to expel propellants.<sup>9</sup> The contents of the beaker were then transferred into 50 mL volumetric flask and completed to the volume with methanol to obtain final concentration of  $100 \mu\text{g mL}^{-1}$  and then general analytical procedure was followed.

## 2.6. Procedure for spiked human plasma

A sample of 5.0 mL of drug-free human blood sample was taken from healthy volunteer into a heparinized tube, centrifuged at 4000 rpm for 30 min. Into a 10 mL stoppered calibrated tube, 1.0 mL of the drug-free plasma (supernatant) was spiked with 1.0 mL of the drug standard working solution. Three milliliters of acetonitrile were added for precipitation of protein<sup>13,16,24</sup> and the resultant mixture was diluted to 10 mL with methanol and then centrifuged at 4000 rpm for 30 min, adequate volumes of the supernatant were taken to obtain solutions within the concentration range of the studied drug, and then general procedure was applied, and after that blank value was determined by treating the drug-free blood sample in the same manner without drug addition.

## 2.7. Procedure for stability indicating study

**2.7.1. Acid induced degradation.** Into 10 mL test tube, 1.0 mL of stock drug solution was transferred and heated at  $80^{\circ}\text{C}$  with 1.0 mL of 2 M HCl for 2 hours. After the specified time, the solution was cooled and neutralized to pH 7 using 2 M NaOH solution; the solution was transferred quantitatively into 10 mL volumetric flask and completed to the volume with methanol. Then 1.0 mL from the solution was taken and the general analytical procedure was applied to detect the remaining intact drug content.

**2.7.2. Base induced degradation.** Into 10 mL test tube, 1.0 mL of stock drug solution was transferred and 1.0 mL of 2 M NaOH was added. The solution was heated at  $80^{\circ}\text{C}$  for 2 hours and after the specified time the solution was cooled and neutralized for pH 7 using 1.0 mL of 2 M HCl. The solution was transferred quantitatively into 10 mL volumetric flask and completed to the volume with methanol. Then 1.0 mL from the solution was taken and the general analytical procedure was applied to detect the remaining intact drug content.

**2.7.3. Oxidative degradation.** Oxidative induced degradation with hydrogen peroxide was carried out by transferring 1.0 mL of stock drug solution into 10 mL test tube then 1.0 mL of hydrogen peroxide (30.0% v/v) was added. The solution was heated at  $80^{\circ}\text{C}$  for 2 hours. The volume was completed to 10 mL with methanol. Then 1.0 mL from the solution was taken and the general analytical procedure was applied to detect the remaining intact drug content.

**2.7.4. Photochemical degradation.** The photochemical stability of the drug was studied by exposing stock solution to direct artificial day light for 12 hours.



### 3. Results and discussion

NBD-Cl is known to form highly fluorescent derivatives with primary and secondary amines under relatively mild alkaline conditions.<sup>17–22</sup> In this work, we investigated the reaction between SAL & NBD-Cl at 80 °C pH (9) to develop the first spectrofluorimetric derivatizing method for determination of salmeterol xinafoate that gives yellow reaction product with high fluorescence intensity at 543 nm after excitation at 473 nm, Fig. 2. The derivatization reaction mechanism was supposed to proceed through secondary aliphatic amine,<sup>17–22</sup> Fig. 3.

#### 3.1. Optimization of variables

The different experimental parameters affecting the reaction development and stability of the fluorescent product were carefully studied and optimized. These factors were changed individually, while the others were kept unchanged. The studied factors include pH, type of buffers, volume of buffer, NBD-Cl volume, temperature, reaction time and diluting solvent.

**3.1.1. Effect of pH and buffer type.** The fluorescence intensity of reaction product of SAL with NBD-Cl was studied using different types of buffer solution as (borate, phosphate, Teorell and Stenhamer buffers) over pH range (7.0–10). Maximum fluorescence intensity obtained at pH (9 ± 0.2) using borate buffer after which the fluorescence intensity started to decline due to increase in the concentration of hydroxide ions which hold back the reaction between SAL and NBD-Cl,<sup>25</sup> Fig. 4. Type of buffer was also investigated. Borate buffer was found to be the best buffer that gave maximum fluorescence intensity. Other buffers such as phosphate, Teorell and Stenhamer buffer at same pH did not give high fluorescence intensity as borate buffer. This probably because of the slow rate of hydrolysis of NBD-Cl to NBD-OH. This result is in agreement with Miyano *et al.*<sup>25</sup>

**3.1.2. Effect of buffer volume.** The volume of borate buffer pH 9 and its influence on fluorescence intensity was also investigated. It was found that the increase in volume of buffer gradually increased the fluorescence intensity of the reaction product and maximum fluorescence intensity was obtained upon using 2 ± 0.5 mL borate buffer pH 9.

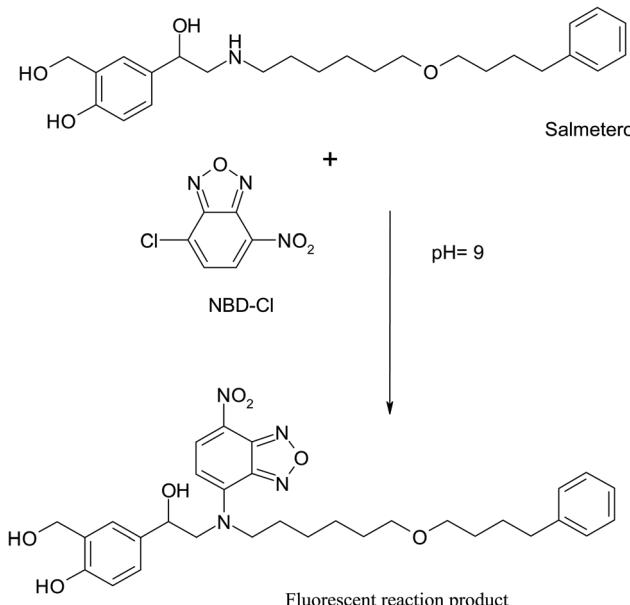


Fig. 3 Suggested reaction mechanism between SAL and NBD-Cl.

**3.1.3. Effect of volume of NBD-Cl.** The fluorescence intensity of reaction product of SAL was studied using different volumes of NBD-Cl ( $2.5 \times 10^{-3}$  M) in the range of (0.2–1.4 mL). Maximum fluorescence intensity obtained upon using  $0.5 \pm 0.1$  mL of NBD-Cl reagent (Fig. 5), after which the fluorescence intensity started to decline probably due to inner filter effect.<sup>25</sup>

**3.1.4. Effect of temperature and time.** The effect of temperature on fluorescence intensity of reaction product of SAL with NBD-Cl was studied, at different temperatures over different time intervals. It was observed that, the reaction product started to form at 50 °C and the fluorescence intensity reached its maximum at 80 °C ± 5 °C after 20 min.

**3.1.5. Effect of diluting solvent.** In order to select the most appropriate solvent for the reaction, many solvents were studied such as: water, methanol, ethanol, acetonitrile, acetone and DMF. Methanol was found to be the best solvent for dilution as it gave the highest fluorescence readings with good reproducibility and low blank readings; water gave poor readings, while

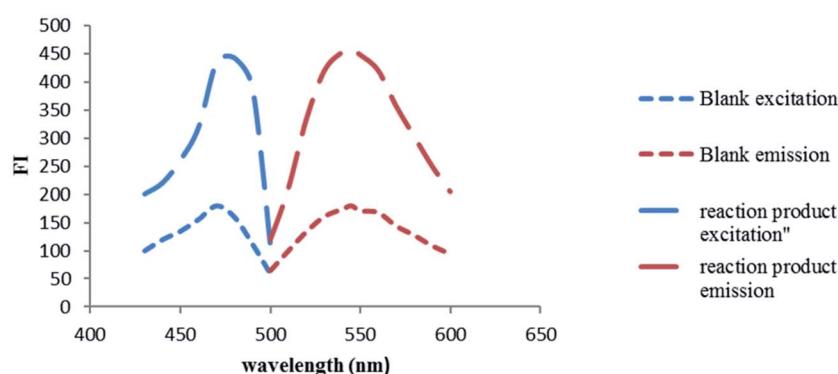


Fig. 2 Excitation and emission fluorescence spectra of the reaction product of SAL ( $1.0 \mu\text{g mL}^{-1}$ ) and NBD-Cl.

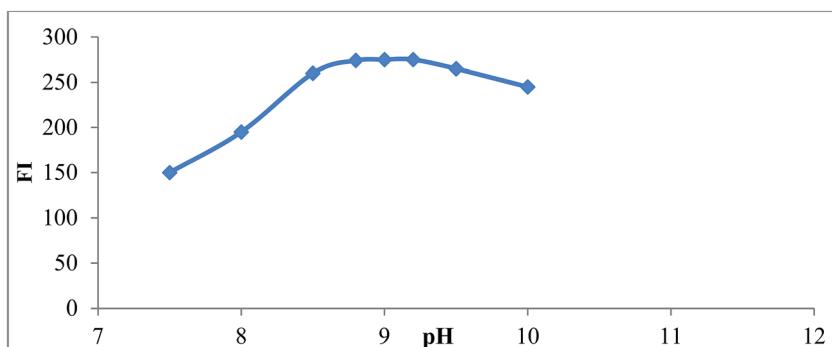


Fig. 4 Effect of pH on the reaction product of SAL ( $1.0 \mu\text{g mL}^{-1}$ ) and NBD-Cl, using borate buffer.

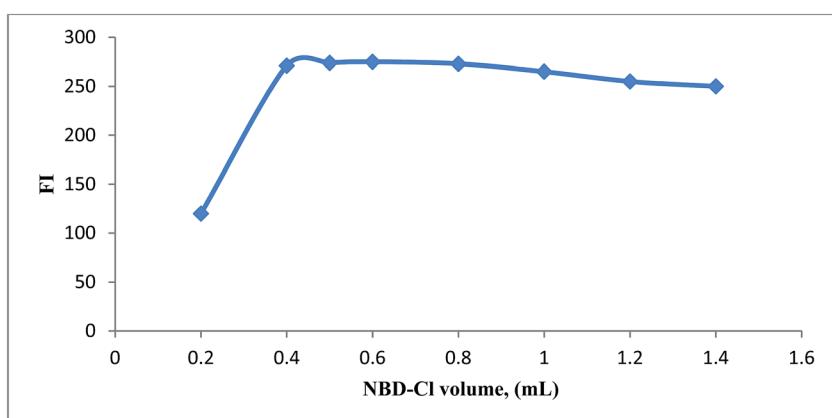


Fig. 5 Effect of NBD-Cl volume ( $2.5 \times 10^{-3} \text{ M}$ ) on the reaction with SAL ( $1.0 \mu\text{g mL}^{-1}$ ).

other organic solvents gave high blank readings compromising the sensitivity of the method or gave poor results relative to methanol.

**3.1.6. Effect of HCl concentration.** Under these reaction conditions, relatively high fluorescence background was observed. This was attributed to the bi product formed during the reaction namely 4-hydroxy-7-nitrobenzo-2-oxa-1,3-diazole (NBD-OH), which come from the hydrolysis of NBD-Cl.<sup>25</sup> The fluorescence of NBD-OH was found to be quenched by decreasing the pH of the reaction medium to less than one, this

was done by the addition of 1 mL of 2 M HCl prior to measure. The volume of 1 mL was sufficient to quench the fluorescence of the NBD-OH without affecting the fluorescence of the reaction product.

**3.1.7. Stoichiometry and mechanism of the reaction.** The stoichiometry of the derivatization reaction was studied through Job's method of continuous variation,<sup>26</sup> using equimolar solutions ( $1.0 \times 10^{-3} \text{ M}$ ) of both SAL and NBD-Cl. Then portions of mixture of master solutions of the drug and reagent were made up comprising different complementary proportions

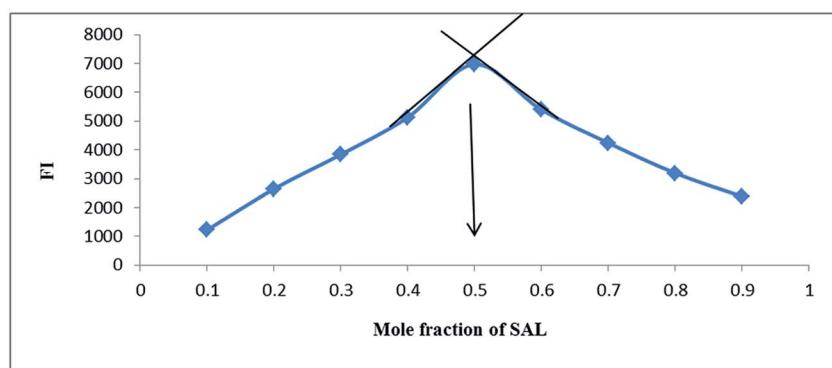


Fig. 6 Method of continuous variation for determination of stoichiometry of the reaction between SAL & NBD-Cl ( $5 \times 10^{-3} \text{ M}$ ).



(0.1 : 0.9, 0.2 : 0.8, 0.3 : 0.7 : 0.3, 0.8 : 0.2, 0.9 : 0.1). It was found that the reaction proceeds in 1 : 1 ratio and this result agrees with the suggested reaction mechanism,<sup>17–22</sup> Fig. 6.

### 3.2. Validation of the proposed procedure

The proposed spectrofluorimetric method was validated according to the ICH guide line<sup>27</sup> regarding accuracy, precision, limit of quantification, limit of detection and robustness.

**3.2.1. Linearity and concentration range.** In this work, the general analytical procedures were performed on a series of standard drug solutions having concentrations ranging from 100 to 1500 ng mL<sup>-1</sup> for SAL. The whole set of experiments were carried out within this range to ensure the validation of the proposed procedure. Linear calibration was obtained by plotting the fluorescence intensity (FI) of the reaction product between SAL and NBD-Cl, subtracted from blank reading and the concentration of SAL. Statistical treatment of the data was carried out using linear regression analysis and the analytical parameters were calculated. The correlation coefficient was 0.9998 over the concentration range (100–1500) ng mL<sup>-1</sup> indicating good linearity of the proposed method, Table 1.

**3.2.2. Accuracy and precision.** Accuracy was checked at five concentration levels within the specified range. Three replicate measurements were recorded at each concentration level. The results were presented as percentage recovery  $\pm$  standard deviation. The obtained results showed good agreement between the measured and true value indicating high accuracy of the proposed method, Table 2.

Precision was examined at three concentration levels; three replicate measurements were recorded at each concentration level; both inter-day and intra-day precision were evaluated and their results were summarized in Table 3. The calculated relative standard deviations were below 2% indicating excellent precision of the proposed procedure at both level of repeatability and intermediate precision.

**3.2.3. Limits of detection (LOD) and quantitation (LOQ).** The limits of detection (LOD) and quantification (LOQ) were calculated based on standard deviation of response and the slope of calibration curve using the equations; LOD = 3.3  $\sigma/S$

**Table 1** Analytical parameters of proposed spectrofluorimetric method between SAL and NBD-Cl<sup>a</sup>

Parameter	Value
$\lambda_{\text{ex}}$ (nm)	473
$\lambda_{\text{em}}$ (nm)	543
Concentration range (ng mL <sup>-1</sup> )	(100–1500)
Correlation coefficient ( $r$ )	0.9998
Determination coefficient ( $r^2$ )	0.9996
Slope	0.2649
Intercept	11.116
SD of the intercept (S <sub>a</sub> )	1.567
SD of slope (S <sub>b</sub> )	0.0021
LOD (ng mL <sup>-1</sup> )	19
LOQ (ng mL <sup>-1</sup> )	59

<sup>a</sup> LOD: limit of detection, LOQ: limit of quantitation.

**Table 2** Evaluation of the accuracy of the proposed spectrofluorimetric method<sup>b</sup>

Sample no.	Taken (ng mL <sup>-1</sup> )	Found (ng mL <sup>-1</sup> )	% recovery <sup>a</sup>
1	100	100.19	100.19
2	300	302.19	100.73
3	500	496.89	99.37
4	700	700.43	100.06
5	1000	1007.65	100.76
Mean			100.22
SD			0.56
RSD			0.55

<sup>a</sup> Mean of three determinations. <sup>b</sup> SD: standard deviation, RSD: relative standard deviation.

**Table 3** Evaluation of intraday and interday precision of the proposed spectrofluorimetric method<sup>b</sup>

Precision level	Conc. (ng mL <sup>-1</sup> )	% recovery <sup>a</sup> $\pm$ SD	RSD (%)
Intraday precision	500	99.49 $\pm$ 0.42	0.42
	700	100.63 $\pm$ 0.50	0.50
	1000	99.83 $\pm$ 0.21	0.21
Interday precision	500	99.59 $\pm$ 1.53	1.53
	700	100.51 $\pm$ 1.32	1.32
	1000	99.87 $\pm$ 1.53	1.53

<sup>a</sup> Mean of three determinations. <sup>b</sup> SD: standard deviation, RSD: relative standard deviation.

and LOQ = 10  $\sigma/S$ , where  $S$  is the slope of the calibration curve and  $\sigma$  is the standard deviation of intercept. The results were presented in Table 1. The limit of quantitation was 19 ng mL<sup>-1</sup> indicating high sensitivity of the proposed method compared with the reported HPLC method for SAL in dosage form [9]. The reversed phase HPLC method for determination of SAL in metered dose inhalers represent detection level of 2  $\mu$ g mL<sup>-1</sup> using expensive HPLC solvents. Therefore, our proposed spectrofluorimetric method is considered highly sensitive and economic compared to the reported one.

**3.2.4. Robustness.** Robustness of the suggested procedure was assessed by evaluating the influence of small variation in experimental variables (pH of buffer solution, volume of buffer, NBD-Cl volume, reaction time and heating temperature). The results presented in Table 4, indicated that small variations in any of these variables did not significantly affect the performance of the suggested procedure. This gave an indication for the reliability of the proposed method.

### 3.3. Application to pharmaceutical dosage forms

The proposed method was successfully applied for determination of SAL in its pharmaceutical dosage form. The selectivity of the method was investigated by observing any interference encountered from the excipients of the pharmaceutical preparation. It was shown that these compounds did not interfere with the proposed method and the results of the proposed method were statistically compared with these of the reported



**Table 4** Evaluation of robustness of the proposed spectrofluorimetric method

Variation	% recovery <sup>a</sup> ± SD
Optimum conditions	100.2 ± 0.42
<b>1-Effect of pH</b>	
pH 8.8	99.5 ± 0.44
pH 9.2	99.7 ± 0.53
<b>2-Volume of NBD-Cl</b>	
0.4 mL	98.8 ± 0.71
0.6 mL	99.9 ± 1.3
<b>3-Reaction time</b>	
15 min	99.1 ± 1.1
25 min	100.0 ± 0.49
<b>4-Heating temperature</b>	
75 °C	99.3 ± 1.1
85 °C	99.5 ± 1.3
<b>5-Volume of buffer</b>	
1.5 mL	99.8 ± 0.92
2.5 mL	100.2 ± 0.95

<sup>a</sup> Mean of five determinations.

method<sup>8</sup> using student's *t*-test and *F*-test. There was no significant difference between the proposed and the reported method as the calculated values did not exceed the theoretical values at 95% confidence level, Table 5. This indicates good level of precision and accuracy of the proposed spectrofluorimetric method.

#### 3.4. Application to spiked human plasma

The high sensitivity attained by the proposed method allowed the possible determination of salmeterol in spiked human plasma. Inter-day, intra-day precision and accuracy in spiked plasma samples were determined by analyzing the biological samples at three concentration levels (100, 200 and 300 (ng mL<sup>-1</sup>)) on three separated days. The precision was defined as the relative standard deviation of spiked sample concentrations determined at three replicates, whereas accuracy was assessed as the percentage to the nominal concentration (%).

Precision and accuracy data for inter- and intra-day human plasma samples are presented in Table 6. The results for all samples (inter- and intra-day) were found to be within the

**Table 5** Application of proposed spectrofluorimetric method to dosage form

Dosage form	% recovery <sup>a</sup> ± SD			
	Proposed	Reported [8]	<i>t</i> -Value <sup>b</sup>	<i>F</i> -Value <sup>b</sup>
Metrovent® inhaler (25 µg SAL/actuation)	100.6 ± 1.3	99.4 ± 0.94	1.7	1.8

<sup>a</sup> Mean of five determinations. <sup>b</sup> The tabulated values of *t*- and *F*-values at 95% confidence limit are 2.78 and 6.39, respectively.

**Table 6** Intra-day, inter-day precision and accuracy study of proposed method for determination the studied drug in spiked human plasma sample

Concentration (ng mL <sup>-1</sup> )	Intra-day precision		Inter-day precision	
	% recovery <sup>a</sup> ± SD	RSD	% recovery <sup>a</sup> ± SD	RSD
100	99.49 ± 0.61	0.61	99.44 ± 0.26	0.26
200	98.90 ± 0.22	0.22	100.5 ± 0.41	0.41
300	99.83 ± 0.35	0.35	99.8 ± 0.38	0.62

<sup>a</sup> Average of three determinations.

acceptance criteria for the method validation. The concentration of the studied drug was computed from its corresponding regression equation. The mean percentage recovery for SAL ranged from 98.9% to 100.5% at the three spiked concentrations applied. The relative standard deviation of the obtained recoveries did not exceed 2% which fall within the acceptable limits of analytical method variability arising from different matrix effects.

Although the proposed method has been successfully applied for determination of SAL in spiked plasma, it is recommended to be applied only in real plasma for those patients who are being treated with SAL and not receiving other drugs containing primary or secondary amine moieties due to the unselective reaction of NBD-Cl towards these drugs.

#### 3.5. Stability indicating study

Our proposed spectrofluorimetric procedure is one of few methods that could determine SAL in presence of its degradation products. The reported methods,<sup>28,29</sup> used high cost techniques such as HPLC & UPLC, while this method is cost effective technique. It was reported that the main degradation pathway of phenolic compounds that have the same nucleus as studied drug is through oxidation under different conditions, especially under alkaline conditions<sup>30,31</sup> and these results agree with that of the proposed method.

Application of analytical proposed method for stability indicating application using the previously mentioned degradation conditions indicated that 46%, 36% and 37% of SAL remained intact after acid, base induced and oxidative degradation, respectively, while no degradation was observed in case of exposing stock solution to direct artificial day light for 12 hours.

## 4. Conclusion

The present study describes the first spectrofluorimetric attempt to determine salmeterol xinafoate in its pharmaceutical dosage form as well as spiked human plasma without being affected by surrounding matrices depending on derivatization reaction with NBD-Cl. The proposed method is simple, rapid and accurate and it has been validated according to ICH guidelines. Furthermore, the developed method is extended to determine the intact drug in the presence of its degradation products. Therefore, the suggested method is suitable for

analysis of the investigated drug in quality control and clinical laboratories.

## Conflicts of interest

There are no conflicts to declare.

## References

- 1 S. Sweetman, *Martindale: the complete drug reference*, Royal Pharmaceutical Society of Great Britain, Pharmaceutical Press, London, 36th edn, 2009.
- 2 K. P. Chowdary and G. Devala Rao, *Indian Drugs*, 1997, **34**, 606–607.
- 3 S. Thea Sufernent, *Indian J. Pharm. Sci.*, 1998, **60**, 294–296.
- 4 K. P. Chowdary and G. Devala Rao, *Indian J. Pharm. Sci.*, 1999, **61**, 246–247.
- 5 M. N. Reddy, K. V. Rao Kanna, S. S. Rao and M. E. Rao, *Indian J. Pharm. Sci.*, 2000, **62**, 193–195.
- 6 M. N. Reddy, K. K. Rao, S. S. Rao and M. E. Rao, *Indian Drugs*, 2000, **37**, 105–107.
- 7 M. S. Kondawar, J. J. Waghmare, M. K. Malusare and N. D. Shah, *Int. J. PharmTech Res.*, 2011, **3**, 1801–1806.
- 8 A. Samir, H. Salem and M. Abdelkawy, *Bull. Fac. Pharm. (Cairo Univ.)*, 2012, **50**, 121–126.
- 9 V. G. Nayak, S. G. Belapure, C. D. Gaitonde and A. A. Sule, *J. Pharm. Biomed. Anal.*, 1996, **14**, 511–513.
- 10 P. S. Jain, A. P. Gorle, S. S. Patil, R. S. Chavan, P. R. Bari and S. J. Surana, *Int. J. Pharmaceut. Chem. Anal.*, 2015, **2**, 28–33.
- 11 L. Kasaye, A. Hymete and A. M. Mohamed, *Saudi Pharm. J.*, 2010, **18**, 153–159.
- 12 M. A. Omar, D. M. Nagy, M. A. Hammad and A. A. Aly, *AAPS PharmSciTech*, 2013, **14**, 828–837.
- 13 S. M. Derayea, M. A. Omar, I. M. Mostafa and M. A. Hammad, *RSC Adv.*, 2015, **5**, 78920–78926.
- 14 A. M. Mohamed, M. A. Omar, M. A. Hammad and A. A. Mohamed, *Spectrochim. Acta, Part A*, 2015, **149**, 934–940.
- 15 M. A. Omar, H. M. Ahmed, M. A. Hammad and S. M. Derayea, *Spectrochim. Acta, Part A*, 2015, **135**, 472–478.
- 16 M. A. Omar, M. A. Hammad, D. M. Nagy and A. A. Aly, *Spectrochim. Acta, Part A*, 2015, **136**, 1760–1766.
- 17 A. A. Al-Majed, *Anal. Chim. Acta*, 2000, **408**, 169–175.
- 18 A. A. Al-Majed and J. Al-Zehouri, *Il Farmaco*, 2001, **56**, 291–296.
- 19 I. A. Darwish, S. M. Amer, H. H. Abdine and L. I. Al-Rayes, *J. Fluoresc.*, 2009, **19**, 463.
- 20 T. Sevgi, *Chin. J. Chem.*, 2010, **28**, 2209–2215.
- 21 T. Sevgi and K. Elif, *J. Chromatogr. Sci.*, 2011, **49**, 417–421.
- 22 M. Walash, N. El-Enany and H. Askar, *Luminescence*, 2015, **30**, 1119–1124.
- 23 M. Pesez and J. Bartos, *Colorimetric and Fluorimetric Analysis of Organic Compounds and Drugs*, Marcel Dekker Inc, New York, 1974.
- 24 M. A. Omar, M. A. Hammad, B. I. Salman and S. M. Derayea, *Spectrochim. Acta, Part A*, 2016, **157**, 55–60.
- 25 H. Miyano, T. Toyooka and K. Imai, *Anal. Chim. Acta*, 1985, **170**, 81–87.
- 26 J. S. Renny, L. L. Tomasevich, E. H. Tallmadge and D. B. Collum, *Angew. Chem.*, 2013, **52**, 11998–12013.
- 27 ICH Harmonized Tripartite Guideline, Validation of Analytical Procedures: Text and Methodology, Q2(R1), Current Step 4 Version, Parent Guidelines on Methodology Dated November 6 1996, Incorporated in November 2005, <http://www.ich.org/LOB/media/MEDIA417.pdf>, accessed 15 February 2008.
- 28 P. Pratibha, S. Kapendra, K. Chandrabose, N. Narayana Moorthy and T. Piuesh, *Lat. Am. J. Pharm.*, 2011, **30**, 985–990.
- 29 B. Akmese, S. Sanli, N. Sanli and A. Asan, *J. Anal. Chem.*, 2014, **69**, 563–573.
- 30 G. A. Jacobson and G. M. Peterson, *J. Pharm. Biomed. Anal.*, 1994, **12**, 825–832.
- 31 H. Al-Malaq, A. Al-Majed and F. Belal, *Anal. Lett.*, 2000, **33**, 1961–1974.

