

## PAPER

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# Nickel tungstate–graphene nanocomposite for simultaneous electrochemical detection of heavy metal ions with application to complex aqueous media†

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We show for the first time that a composite of carbon and binary transition metal oxide, in the form of a reduced graphene oxide and nickel tungstate (RGO/NiWO<sub>4</sub>) nanocomposite, is an effective material for electrochemical heavy metal ions detection. The multivalent electronic states of this composite show well-defined peaks of Cd(II), Pb(II), Cu(II) and Hg(II) during simultaneous detection, which is otherwise not observed for NiWO<sub>4</sub> NPs and RGO sheets. Simultaneous and selective detection of heavy metal ions in drinking water as well as in complex aqueous media such as carbonated drinks, milk and fruit juices has been successfully demonstrated. Differential pulse anodic stripping voltammetric (DPASV) method was adopted for detection because it partially suppresses the background current and improves signal which leads to a low limit of detection (LOD) when compared to linear sweep voltammetry (LSV). LOD for Cd(II), Pb(II), Cu(II) and Hg(II) ions were found to be  $4.7 \times 10^{-10}$  M,  $3.8 \times 10^{-10}$  M,  $4.4 \times 10^{-10}$  M and  $2.8 \times 10^{-10}$  M for individual detection and  $1.0 \times 10^{-10}$ ,  $1.8 \times 10^{-10}$ ,  $2.3 \times 10^{-10}$  and  $2.8 \times 10^{-10}$  M, for simultaneous detection, respectively. The effect of deposition time and deposition potential on the sensing parameter was studied in acetate buffer (pH = 5.0). The better sensitivity with the high capacitive current along with individual and simultaneous electrochemical detection of RGO/NiWO<sub>4</sub> nanocomposite is mainly attributed to its large surface area, good electronic conductivity, and better electron transport properties which lead to better catalytic response towards the heavy metal ions detection.

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## Introduction

Trace amounts of heavy metal may cause serious health issues in living organisms because of their tendency to accumulate in the major organs in humans and animals.<sup>1,2</sup> Heavy metal ions such as Cu(II), Cd(II), Pb(II) and Hg(II) are a major threat as once absorbed, even in trace amounts, can cause serious adverse effects on the immune, central nervous and reproductive systems.<sup>3,4</sup> Hence, it is important to develop simple, fast, accurate and cost-effective methods to simultaneously detect these toxic metal ions present in trace amounts. In order to detect these heavy metal ions several spectrometric methods such as atomic absorption spectroscopy (AAS), inductively coupled plasma mass spectrometry (ICP-MS), UV-Vis spectroscopy X-ray fluorescence spectroscopy (XFS) and inductively

coupled plasma atomic emission spectrometry (ICP-AES) are employed routinely.<sup>1–3</sup> However, these methods require excessive amounts of time, expensive instrumentation and high operating cost. Quite the reverse, the electrochemical method<sup>5–7</sup> offers an exciting alternative due to its numerous advantages. It is simple, rapid, portable and cost effective in order to detect heavy metals at greater accuracy and high sensitivity and above all, it is user-friendly. Electrochemistry itself offers a host of techniques depending on detection range, such as for higher concentrations ( $>10^{-5}$  M) linear sweep voltammetry (LSV) is employed, however, for lower concentrations, the double-layer charging current is not negligible as the faradic current becomes smaller and thus pulse techniques, *i.e.*, normal pulse (NPV), square wave (SWV) and differential pulse (DPV) voltammetry are employed wherein partial suppression the background current occurs and thus improving LOD. One of the pulsed techniques, differential pulse anodic stripping voltammetry (DPASV) involves the sensing mechanism by accumulating the target metal ions on the surface of the electrode during the applied potential and the electron transportation between the electrode material and metal ions. This method also offers detection of a very low concentration of metal ions

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and simultaneous detection of various types of heavy metal ions with high sensitivity, thus this method is ideal.

Pure NiO has a cubic structure and is classified as a Mott–Hubbard insulator with room temperature conductivity less than  $10^{-13} \text{ S cm}^{-1}$ , which will greatly reduce its electrochemical properties.<sup>8–10</sup> The incorporation of W atoms can greatly enhance the conductivity of  $\text{NiWO}_4$  in comparison with pure NiO. Therefore, nickel tungstate ( $\text{NiWO}_4$ ) possesses high electrical conductivity ( $10^{-7}$  to  $10^{-2} \text{ S cm}^{-1}$ ) than many single transition metal oxides and multivalent binary oxides *e.g.*  $\text{NiMoO}_4$  (ref. 11)  $\sim 10^{-11}$  to  $10^{-4.5} \text{ S cm}^{-1}$   $\text{FeWO}_4$  (ref. 12)  $\sim 10^{-6} \text{ S cm}^{-1}$  and  $\text{Bi}_2\text{WO}_6 \sim 10^{-8} \text{ S cm}^{-1}$ .<sup>13</sup> This material has low synthesis cost, low corrosion, high stability and high electrocatalytic activity. Further, it is environmentally friendly and has been utilised in a wide variety of applications such as supercapacitors, lithium-ion battery, and electrocatalyst.<sup>14</sup> However, the agglomeration of individual nanoparticles leading to low surface area severely hinders the electro-catalytic performance. Additionally, the electrical conductivity of transition metal oxides is far less than the carbonaceous material, which further reduces the electro-catalytic performance. Therefore, to facilitate the better electrochemical, electro-catalytic performance and improving electrical conductivity, incorporation of carbonaceous materials is essential. Among all types of carbon sources, graphene is an appropriate 2-D material owing to its high surface area, good electrical conductivity, superior flexibility, electronic transport properties and excellent electro-catalytic activity. However, the agglomeration of individual graphene sheets due to restacking and consequent surface area reduction limited its applicability. Here, among other things, the problem of restacking of individual graphene sheets is overcome by incorporation of metal oxide nanoparticles.

Although there is a vast literature on electrochemical sensing of heavy metal ions, there are only a few reports on the simultaneous and selective electrochemical sensing of toxic metal ions. Recently, Wei *et al.*<sup>15</sup> demonstrated the electrochemical detection of heavy metal ions using  $\text{SnO}_2/\text{RGO}$  nanocomposite. Xie *et al.*<sup>16</sup> utilised  $\text{CeO}_2/\text{RGO}$  nanocomposite for the simultaneous detection of heavy metal ions. Further, Xiong *et al.*<sup>17</sup> demonstrated the use of  $\text{RGO-Fe}_3\text{O}_4$  nanocomposite and Lee *et al.*<sup>18</sup> employed RGO decorated with tin nanoparticles for heavy metal detection in soil and tap water respectively. Apart from metal oxide and graphene composites, numerous types of modified electrodes were tested for the detection of toxic metals such as carbon nanotubes,<sup>19</sup> gold nanoparticles,<sup>20</sup> mesoporous silica,<sup>21</sup> polypyrrole/carbonaceous nanospheres,<sup>22</sup> tin oxide nanoparticles,<sup>23</sup> thiacalixarene-functionalized graphene oxide,<sup>24</sup> and palladium nanoparticle incorporated porous activated carbon.<sup>25</sup>

Here, we present a simple hydrothermal synthesis of reduced graphene oxide and nickel tungstate ( $\text{RGO/NiWO}_4$ ) nanocomposite to investigate the simultaneous and effective detection of  $\text{Cu(II)}$ ,  $\text{Cd(II)}$ ,  $\text{Pb(II)}$  and  $\text{Hg(II)}$  metal ions by DPASV. The  $\text{RGO/NiWO}_4$  nanocomposite exhibited enhanced sensing performance as compared to pristine RGO and  $\text{NiWO}_4$  NPs. The binary metal oxide nanoparticles not only act as spacers to

reduce the restacking of individual graphene sheets but also behave as an electrochemical catalyst, which enhances the sensing performance. Our results suggest that the  $\text{RGO/NiWO}_4$  nanocomposite can indeed be used as an effective electro-catalyst for the simultaneous detection of toxic heavy metal ions. Not limited to testing in water, these fabricated electrodes exhibited excellent sensing properties even in complex aqueous media of carbonated drinks, milk and fruit juices, thus offering a great potential for wider application as a metal ion sensor.

## Experimental section

### Materials

Sodium tungstate dihydrate ( $\text{Na}_2\text{WO}_4 \cdot 2\text{H}_2\text{O}$ ), and nickel chloride ( $\text{NiCl}_2 \cdot 6\text{H}_2\text{O}$ ) were purchased from Fisher Scientific. Acetic acid glacial, cadmium nitrate ( $\text{Cd}(\text{NO}_3)_2$ ), copper nitrate ( $\text{Cu}(\text{NO}_3)_2$ ), mercury nitrate ( $\text{Hg}(\text{NO}_3)_2 \cdot \text{H}_2\text{O}$ ), lead nitrate ( $\text{Pb}(\text{NO}_3)_2$ ), graphite flakes, sulphuric acid ( $\text{H}_2\text{SO}_4$ ), sodium nitrate ( $\text{NaNO}_3$ ), sodium acetate ( $\text{CH}_3\text{COONa}$ ) and potassium permanganate ( $\text{KMnO}_4$ ) were purchased from Loba Chemie. DI water was used for solution preparation.

### Methods

**Synthesis of  $\text{RGO/NiWO}_4$  nanocomposite.** Modified Hummer's method was used to synthesise graphene oxide (GO) as reported previously.<sup>26</sup> Reduced graphene oxide/nickel tungstate ( $\text{RGO/NiWO}_4$ ) nanocomposite were synthesised *via* hydrothermal method by mixing GO with nickel and tungsten salts followed by calcination and pyrolysis in air and nitrogen atmosphere, respectively.<sup>14</sup> Briefly,  $2 \text{ mg ml}^{-1}$  aqueous suspension of GO was stirred with 2 mmol  $\text{NiCl}_2 \cdot 6\text{H}_2\text{O}$  for 1 h at room temperature. After complete mixing, 2 mmol of  $\text{Na}_2\text{WO}_4 \cdot 2\text{H}_2\text{O}$  was added and stirred for another 1 h. The resulting solution was transferred into 50 ml hydrothermal reactor and heated at  $180^\circ\text{C}$  for 12 h. After completion of the reaction, the reactor was cooled to room temperature and the product was collected by centrifugation. The collected product was washed several times with ethanol and water followed by drying at  $80^\circ\text{C}$  overnight. Finally, the dried product was heated in air at  $400^\circ\text{C}$  for 2 h followed by pyrolysis in a nitrogen atmosphere for 2 h at  $600^\circ\text{C}$ . Neat  $\text{NiWO}_4$  NPs were also prepared by adopting the same procedure but without the addition of GO.

**Characterisation  $\text{RGO/NiWO}_4$  nanocomposite.** The morphological structure of  $\text{RGO/NiWO}_4$  nanocomposite was visualised by field emission scanning electron microscopy (FESEM, ZEISS Supra 40VP, Germany) and transmission electron microscopy (TEM, FEI Tecnai G<sup>2</sup> 12 Twin, USA). Energy dispersive X-ray spectroscopy (EDX, Oxford Instruments, England) combined with elemental mapping was performed in FESEM to determine the elemental composition. X-ray diffraction (XRD, PAN analytical, Germany) of the nanocomposite was measured from  $5^\circ$  to  $80^\circ$  with  $\text{Cu K}\alpha$  radiation. Thermogravimetric analysis (TGA, TA instrument, USA) was performed in air with a heating rate of  $10^\circ\text{C min}^{-1}$ .

**Electrochemical measurements.** The electrodes were prepared by mixing the nanocomposite, acetylene black, and



poly(vinylidene fluoride) (PVDF) with the mass ratio of 80 : 10 : 10 and dispersed in *N*-methylpyrrolidone (NMP). The mixture was stirred till it formed a homogeneous slurry, then the slurry was dip-coated onto carbon paper and dried under vacuum at 120 °C for 24 h. The loading mass of each working electrode was about 2–3 mg, and each working electrode had a geometric surface area of about 1 cm<sup>2</sup>. In a three-electrode system, 0.1 M acetate buffer (pH 5.0) was used as the electrolyte, platinum rod and Ag/AgCl (3 M KCl-saturated) electrode was used as the counter and a reference electrode, respectively. Cyclic voltammetry (CV), electrochemical impedance spectroscopy (EIS) measurements and differential pulse anodic stripping voltammetry (DPASV) were performed using an electrochemical workstation (Autolab PGSTAT 302N). EIS was recorded under the following conditions: AC voltage amplitude of 10 mV, the frequency range of  $1 \times 10^5$  to 0.1 Hz, and open circuit potential.

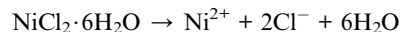
**Individual detection of Cd(II), Pb(II), Cu(II), and Hg(II) ions.** In order to detect heavy metal ions, differential pulse anodic stripping voltammetry (DPASV) was employed. The deposition potential of Cu(II), Cd(II), Pb(II), and Hg(II) was performed at −1.0 V for 120 s in 0.1 M acetate buffer (pH 5.0) by the reduction to their respective metallic state. The re-oxidation of metal to metal ions, *i.e.*, anodic stripping of electrodeposited metal was performed at various potentials ranging from 0.5 to 1 V with 25 mV amplitude, 15 Hz frequency, and 4 mV increment potential and similar experimental conditions were used to perform the selective/individual and simultaneous detection of Cu(II), Cd(II), Pb(II), and Hg(II) ions.

## Results and discussion

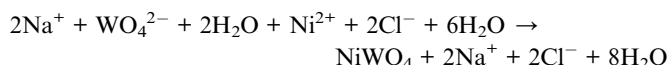
Scheme 1 shows the schematic of synthesis process of RGO/NiWO<sub>4</sub> nanocomposite. As stated before, negatively charged GO was dispersed in distilled water followed by addition of nickel and tungsten salts. During the addition of nickel salt, positively charged Ni<sup>2+</sup> cations were adsorbed on negatively charged GO by electrostatic interaction. After, completion of the process, the solution became positively charged which assisted in adsorption of WO<sub>4</sub><sup>2−</sup> anions during the addition of tungstate salt. The functional groups present on the surface of graphene oxide were reduced by hydrothermal treatment. The crystallisation of NiWO<sub>4</sub> occurred by high-temperature pyrolysis initially at 400 °C in air and then at 600 °C in nitrogen atmosphere.

Following are the governing reactions for the formation of NiWO<sub>4</sub>.

After dissolution of salt at room temperature:



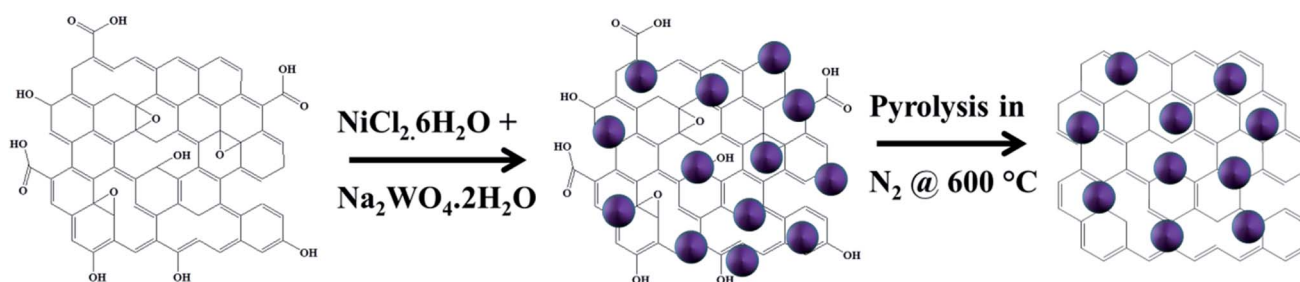
After hydrothermal treatment:



### Morphology of RGO/NiWO<sub>4</sub> nanocomposite

Initially, the composition of the nanocomposite was confirmed using XRD and thermogravimetric analysis and results are discussed in Fig. 1. X-ray diffraction pattern was recorded for GO, RGO, NiWO<sub>4</sub> and RGO/NiWO<sub>4</sub> nanocomposite and their respective diffraction patterns are shown in Fig. 1a. For GO, the diffraction peak observed at  $2\theta$  of 10.5° reflected (002) plane of exfoliated GO sheets with interlayer *d*-spacing of 0.8 nm.<sup>26</sup> After reduction of GO to RGO, the peak at 10.5° of GO disappeared completely and a new peak appeared at 26° indicating the formation of RGO. For NiWO<sub>4</sub> NPs, diffraction peaks were observed at 19.2°, 23.9°, 24.9°, 30.9°, 36.5°, 39.1°, 41.6°, 46.4°, 52.3°, 54.6°, 62.3°, 65.8°, 68.9°, 72.3° and 79.3° which corresponded to (100), (011), (110), (−111), (002), (200), (−102), (−211), (130), (−202), (−113), (−311), (041), (−141) and (−223) planes of wolframite monoclinic structure with *P2/c* space group of NiWO<sub>4</sub> (JCPDS file no. 15-0755).<sup>27</sup> The crystallite size of NiWO<sub>4</sub> NPs was measured as 21 nm at  $2\theta$  of 30.9° using Scherrer equation. In the case of RGO/NiWO<sub>4</sub> nanocomposite, diffraction peaks due to graphene sheets were not seen due to the complete coverage of graphene by NiWO<sub>4</sub> NPs and that resulted in the strong shielding of the characteristic peak of graphene. Therefore, for RGO/NiWO<sub>4</sub> nanocomposite, the diffraction peaks completely matched with the peaks observed for NiWO<sub>4</sub>.

In order to determine the amount of loading of NiWO<sub>4</sub> in RGO/NiWO<sub>4</sub> nanocomposite, TGA was performed in the air till 800 °C. TGA curve of GO, RGO, and RGO/NiWO<sub>4</sub> nanocomposite are shown in Fig. 1b. For GO, initial weight loss of 17% was observed below 100 °C due to the removal of surface adsorbed



Scheme 1 Schematic representation of synthesis process of RGO/NiWO<sub>4</sub> nanocomposite.





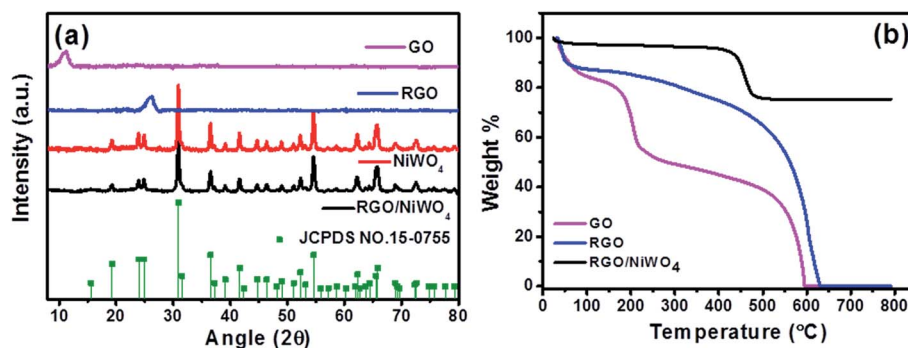


Fig. 1 (a) XRD of GO, RGO,  $\text{NiWO}_4$  and RGO/ $\text{NiWO}_4$  nanocomposite and (b) TGA of GO, RGO, and RGO/ $\text{NiWO}_4$  nanocomposite.

water. Around 200 °C functional groups degradation and complete decomposition around 600 °C was observed. For RGO, initial weight loss of 12% was observed below 100 °C, due to the removal of surface adsorbed water. The complete burning of carbon occurred at 600 °C confirming no impurity present in the sample. RGO/ $\text{NiWO}_4$  nanocomposite showed 26% weight loss at 450–600 °C mainly due to the burning of graphene in the air. Beyond 600 °C, there was no weight loss observed due to complete removal of carbon species. TGA result confirmed that ~74 wt%  $\text{NiWO}_4$  NPs was loaded in RGO/ $\text{NiWO}_4$  nanocomposite.

The morphology of GO,  $\text{NiWO}_4$  NPs, and RGO/ $\text{NiWO}_4$  nanocomposite was confirmed using electron microscopy. TEM and SEM were performed in order to visualise the samples and their respective images are shown in Fig. 2. TEM and SEM image of GO sheets, shown in Fig. 2a and d respectively, displayed transparent but wrinkled sheet morphology. The lateral

dimension of GO sheets is in microns and the sheet thickness is in nanometres.  $\text{NiWO}_4$  NPs with size ranging from 20 to 50 nm can be seen in Fig. 2b and e. The inset in Fig. 2b shows the SAED pattern with random and continuous bright spots confirming the polycrystalline nature of  $\text{NiWO}_4$  NPs. Further, the lattice fringes obtained from HRTEM indicated the inter-planar distance of 0.36 nm corresponding to (011) planes of  $\text{NiWO}_4$ . Homogenous distribution of  $\text{NiWO}_4$  NPs on RGO sheets in the nanocomposite of  $\text{NiWO}_4$  with RGO is clearly visualised from TEM and SEM images shown in Fig. 2c and f images. The inset of Fig. 2c showed the lattice fringes and SAED pattern of RGO/ $\text{NiWO}_4$  nanocomposite, which further confirmed the polycrystalline nature. Fig. S2 (see ESI†) showed the energy dispersive X-ray spectroscopy (EDX) confirming the composition and elemental mapping of RGO/ $\text{NiWO}_4$  nanocomposite, which confirmed the homogeneous distribution of C, O, Ni, and W elements. From these results, it is clear that RGO sheets

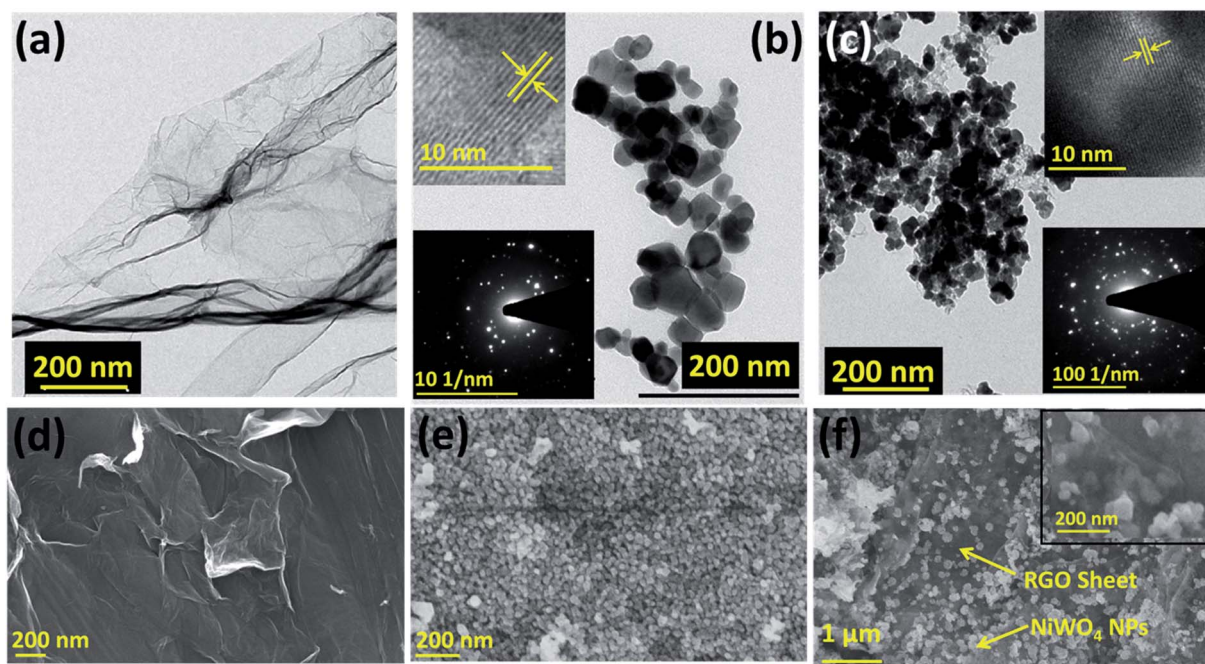


Fig. 2 TEM images of (a–c) GO,  $\text{NiWO}_4$  NPs and RGO/ $\text{NiWO}_4$  nanocomposite with inset showing SAED and lattice fringes. SEM images of (d–f) GO,  $\text{NiWO}_4$  NPs and RGO/ $\text{NiWO}_4$  nanocomposite.



provided as 3-dimensional conductive support and a basal plane for the appropriate attachment of  $\text{NiWO}_4$  NPs on the surface of RGO sheets.

### Electrochemical characterization of RGO/ $\text{NiWO}_4$ nanocomposite

**Cyclic voltammetry and DPASV.** Before performing electrochemical detection of toxic metal ions using the prepared electrodes, cyclic voltammetry was performed on bare CP, RGO,  $\text{NiWO}_4$  NPs and RGO/ $\text{NiWO}_4$  nanocomposite coated on CP in aqueous solution of 0.1 M KCl containing 5 mM  $\text{Fe}(\text{CN})_6^{3-}$  in the voltage range of  $-0.2$  V to  $0.6$  V at a scan rate of  $50 \text{ mV s}^{-1}$ . The CV recorded are shown in Fig. 3a, the anodic peak current drawn by the CV curve for bare CP is the least among all the electrodes and RGO seem to be slightly better than CP. For the  $\text{NiWO}_4$  NPs, the current response in the CV is equivalent to that of RGO but with no obvious redox peaks which is mainly attributed to the agglomeration of individual nanoparticles and poor electronic conductivity, thus causing the hindrance of electron transfer on the electrode surface. RGO/ $\text{NiWO}_4$  nanocomposite exhibited the highest current response among all the electrodes. This is mainly attributed to the larger surface area of the nanocomposite as the  $\text{NiWO}_4$  NPs acted as spacers and prevented the restacking of the RGO sheets and also an agglomeration of individual NPs, hence causing the better conducting pathway for the electron transport and enhanced electro-catalytic process at the surface of

nanocomposite electrode. Long cyclic stability of the electrode was studied by performing cyclic voltammetry. The electrode has been subjected to consecutive cycling in the potential range of  $-0.2$  to  $0.6$  V for 200 cycles in 0.1 M KCl solution containing 5 mM  $\text{Fe}(\text{CN})_6^{3-}$ . As shown in Fig. 3b, the peak current slightly increased up to 100 cycles and thereafter, no obvious changes were observed indicating good stability and reproducibility of the electrochemical performance of the electrode.

An electrochemical impedance spectroscopy (EIS) was employed to characterize the interfacial properties of the electrodes. Fig. 3c showed the Nyquist plot of the electrodes measured at the open-circuit potential in the frequency range from  $0.01$  to  $10^5$  Hz, where  $Z'$  and  $Z''$  are the real and imaginary parts of the impedance, respectively. The semicircle at higher frequency range corresponded to the electron-transfer resistance ( $R_{\text{et}}$ ), while a linear part at lower frequency range represented the diffusion limited process or Warburg resistance. The  $R_{\text{et}}$  value of bare CP is highest among all the electrodes, which signified the poor electron transport. After coating the CP with RGO,  $\text{NiWO}_4$  NPs and RGO/ $\text{NiWO}_4$  nanocomposite, the semi circle domain with  $R_{\text{et}}$  value decreased and the lowest value was observed for RGO/ $\text{NiWO}_4$  nanocomposite, suggesting better electron transfer process at nanocomposite electrode surface. The inset in Fig. 3c displayed the enlarged view of the Nyquist plot of RGO/ $\text{NiWO}_4$  nanocomposite for better visualization.

Further, DPASV was performed on bare CP, RGO,  $\text{NiWO}_4$  NPs and RGO/ $\text{NiWO}_4$  nanocomposite. The deposition process was

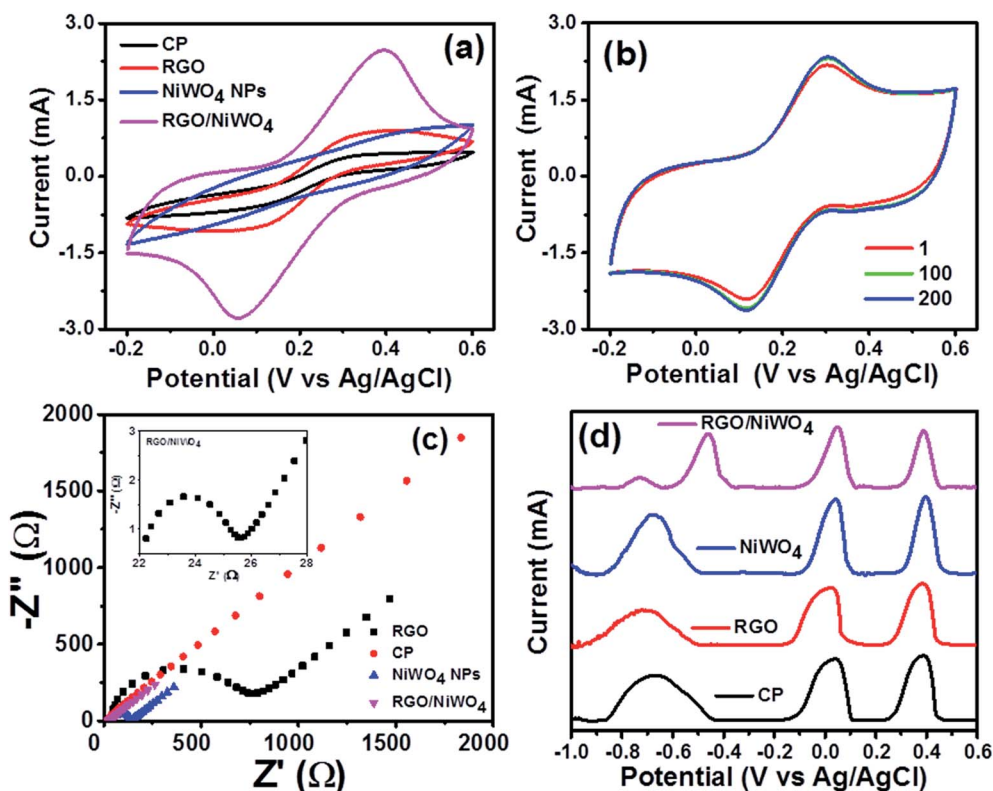


Fig. 3 (a) CV (b) CV with a different number of scanning cycles 1–200 with a scan rate of  $50 \text{ mV s}^{-1}$  (c) EIS with inset showing the zoomed-in image and (d) DPASV response of bare CP, RGO,  $\text{NiWO}_4$  NPs and RGO/ $\text{NiWO}_4$  nanocomposite. Electrolyte: 5 mM  $\text{Fe}(\text{CN})_6^{3-/4-}$  in 0.1 M KCl.



carried out in a solution containing 0.1  $\mu\text{M}$  each of  $\text{Cd(II)}$ ,  $\text{Pb(II)}$ ,  $\text{Cu(II)}$ , and  $\text{Hg(II)}$  ions at  $-1.0$  V for 120 seconds in 0.1 M acetate buffer (pH 5.0) and the results were shown in Fig. 3d. Three peaks were observed for bare CP, RGO and  $\text{NiWO}_4$  NPs in the range of  $-1.0$  V to 0.6 V, however, four sharp and distinct peaks were observed for RGO/ $\text{NiWO}_4$  nanocomposite. Individually, peaks due to  $\text{Cd(II)}$ ,  $\text{Pb(II)}$ ,  $\text{Cu(II)}$ , and  $\text{Hg(II)}$  ions can be identified at potentials of 0.774, 0.578, 0.109 and 0.24 V, respectively, which is not seen for the rest of the electrodes, wherein  $\text{Cd(II)}$  and  $\text{Pb(II)}$  peaks have merged together to form one broad peak thus indicating their inferior performances. Therefore, RGO/ $\text{NiWO}_4$  nanocomposite electrodes seem to be a suitable electrode system for heavy metal ions detection.

**Optimization of experimental conditions.** Two parameters, namely, deposition potential and deposition time play a key role in achieving best sensitivity while detection of the metal ions. Therefore, a study on the effect of the deposition potential on the peak current for a fixed time duration (120 s deposition time) was conducted in the voltage range from 0 to 1.4 V in 0.1 M acetate buffer (pH 5.0) for  $\text{Cu(II)}$  and  $\text{Pb(II)}$  ions and the results are shown in Fig. 4a. At lower deposition potentials, 0–0.4 V, hardly any peak current was observed, further increase in deposition potential led to increasing in peak current till 1.2 V. Although, at 1.2 V, higher peak current for  $\text{Cu(II)}$  and  $\text{Pb(II)}$  ions was observed, however,  $\text{H}_2$  evolution was also accompanied. In order to avoid the competitive generation of  $\text{H}_2$  and co-deposition of other metal ions while analysis of complex aqueous media, 1.0 V was chosen as the optimized deposition potential. The difference in peak current values observed for different metal ions may be attributed to the different standard potentials of the respective metals. Deposition time can play a crucial role in determining the detection limit and sensitivity; hence variation in deposition time was studied. The deposition time was varied from 20–140 s at 1.0 V for  $\text{Cu(II)}$  and  $\text{Pb(II)}$  ions and the results are shown in Fig. 4b. The peak current increased with increase in deposition time, this is due to the increased amount of ions on the electrode surface. Although increasing the deposition time improved the sensitivity, but due to saturation of electrode surface, the upper detection limit at high metal ion concentrations may be lowered.<sup>28</sup> Therefore, 120 s was chosen as optimized deposition time to achieve a lower detection limit.

**Stripping behaviour toward  $\text{Cd(II)}$ ,  $\text{Pb(II)}$ ,  $\text{Cu(II)}$ , and  $\text{Hg(II)}$  ions.** To understand the stripping behaviour of metal ions, individual detection of  $\text{Cd(II)}$ ,  $\text{Pb(II)}$ ,  $\text{Cu(II)}$  and  $\text{Hg(II)}$  ions at RGO/ $\text{NiWO}_4$  nanocomposite electrode using DPASV was performed. The limit of detection (LOD), sensitivity and correlation coefficients were determined for individual ions by varying the concentration of metal ions from 0.1 to 2.0  $\mu\text{M}$ . In Fig. 5, the peak current increased linearly with increasing the metal ion concentration regardless of the type of the metal ion. Fig. 5a showed the DPASV response of  $\text{Cd(II)}$  ion at different concentrations in the range of 0.1 to 2.0  $\mu\text{M}$ , where the well-defined peak was observed. From the calibration plot, the linear equation  $i/\text{mA} = 0.5419 + 0.1145C$  ( $\mu\text{M}$ ) with the correlation coefficient of 0.9978 (inset of Fig. 5a) was obtained and LOD was calculated to be  $4.7 \times 10^{-10}$  M ( $3\sigma$  method). The DPASV response of  $\text{Pb(II)}$  ions over a concentration range of 0.1 to 2.0  $\mu\text{M}$  was measured and the results are shown in Fig. 5b. From the calibration plot, the linear equation was  $i/\text{mA} = 0.5327 + 0.03412C$  ( $\mu\text{M}$ ), with the correlation coefficient of 0.9510 (inset of Fig. 5b) and LOD was calculated to be  $3.8 \times 10^{-10}$  M ( $3\sigma$  method). The DPASV response of  $\text{Cu(II)}$  ions over a concentration range from 0.1 to 2.0  $\mu\text{M}$  is shown in Fig. 5c. From the calibration plot, the linear equation was  $i/\text{mA} = 1.6169 + 0.0407C$  ( $\mu\text{M}$ ), with the correlation coefficient of 0.9986 (inset of Fig. 5c) and LOD was calculated to be  $4.4 \times 10^{-10}$  M ( $3\sigma$  method). The DPASV response of  $\text{Hg(II)}$  ions over a concentration range from 0.1 to 2.0  $\mu\text{M}$  is shown in Fig. 5d. From the calibration plot, the linear equation was  $i/\text{mA} = 1.0549 + 0.0683C$  ( $\mu\text{M}$ ), with the correlation coefficient of 0.9968 (inset of Fig. 5d) and LOD was calculated to be  $2.8 \times 10^{-10}$  M ( $3\sigma$  method). One common observation made for all the heavy metal ions detection is that with the increase in the metal ion concentration, the peak intensity increases along with a positive shift in the peak potential. This phenomenon of peak shift is common as adsorption product tends to shift the peak potential positively.<sup>29</sup> For further understanding, we performed cyclic voltammetry in an aqueous solution of 0.1 M KCl containing 5 mM  $\text{Fe(CN)}_6^{3-}$  in the voltage range of  $-0.1$  V to 0.6 V with scan rate varying from 10 to 100  $\text{mV s}^{-1}$ . The results are shown in Fig. S4 in ESI.† It is clear that as the scan rate increases, the anodic and cathodic peak current increases linearly (see Fig. S4a†). A plot of anodic and cathodic peak current *versus*

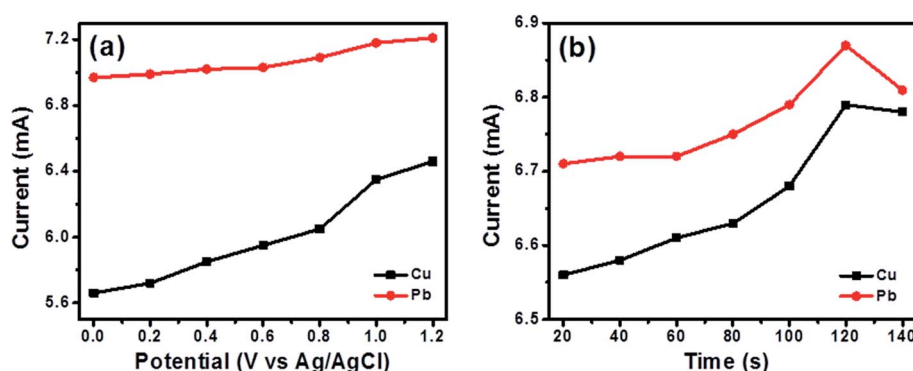


Fig. 4 Variation in (a) deposition potential and (b) deposition time for  $\text{Cu(II)}$  and  $\text{Pb(II)}$  ions using DPASV. Electrolyte: 0.1 M acetate buffer (pH 5.0).





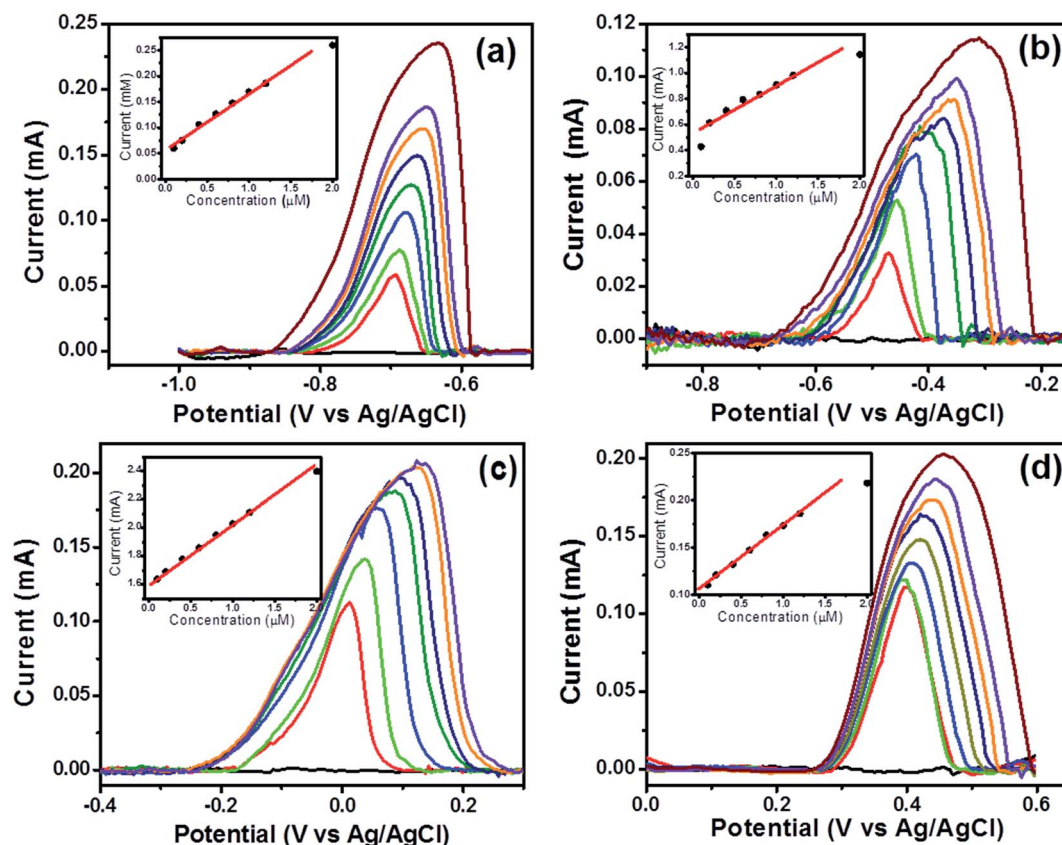


Fig. 5 DPASV response of RGO/NiWO<sub>4</sub> nanocomposite recorded at 1.0 V for 120 s for change in concentration of individual (a) Cd(II) (b) Pb(II) (c) Cu(II) and (d) Hg(II) ions with inset showing respective calibration plot. Electrolyte: 0.1 M acetate buffer (pH 5.0).

scan rate is also shown in Fig. S4b.† A linear trend indicates that this process is surface confined and thus adsorption controlled process. Further, the sensitivity, LOD and correlation coefficients suggested that RGO/NiWO<sub>4</sub> nanocomposite electrode can potentially serve as a good candidate for the determination of heavy metal ions.

**Evaluation of mutual interferences.** After evaluating the individual stripping behaviour of heavy metal ions, the simultaneous analysis was performed to understand their interference effect, if any. For simultaneous determination of Cu(II) and Hg(II) ions concentration was varied and results are shown in Fig. 6. Well-defined peaks of Cu(II) and Hg(II) ions were observed for all concentrations as shown in Fig. 6a and the peak currents of the both analyte ions increased with the increasing their respective concentrations. The resulting calibration plots for both the ions are shown in Fig. 6b, which is linear over the range from 0.2 to 0.7 μM. From the calibration plot, the linear equations where  $i/\text{mA} = 0.8899 + 0.1174C (\mu\text{M})$  and  $i/\text{mA} = 0.9794 + 0.1245 (\mu\text{M})$ , with the correlation coefficients of 0.9963 and 0.9971 and LOD was calculated to be  $1.1 \times 10^{-10}$  M and  $1.2 \times 10^{-10}$  M for Cu(II) and Hg(II) ions, respectively. During individual analysis of Cu(II) and Hg(II) ions, the sensitivity obtained was  $0.0407 \text{ mA } \mu\text{M}^{-1} \text{ cm}^{-2}$  for Cu(II) ions and  $0.0683 \text{ mA } \mu\text{M}^{-1} \text{ cm}^{-2}$  for Hg(II) ions, however, in simultaneous analysis, the sensitivity obtained was  $0.1174$  and  $0.1245 \text{ mA } \mu\text{M}^{-1} \text{ cm}^{-2}$ , respectively (see Table S1, ESI†). These

results are encouraging as there is a good separation between Cu(II) and Hg(II) ions peaks ruling out the formation of inter-metallic compounds from these metals which are the case with many electrode systems.<sup>15,30–32</sup> This result is attributed to the highly porous structure of RGO/NiWO<sub>4</sub> nanocomposite which lowered oxidation overpotential of each metal ions.<sup>29,31,33</sup>

Further, simultaneous analysis of three ions namely Cd(II), Cu(II) and Hg(II) ions was performed and results indicated existence of individual peaks for each ion in DPASV response as demonstrated in Fig. 7a and the peak currents of the three analyte ions increased with the increasing their respective concentrations. The correlation coefficients for Cd(II), Cu(II) and Hg(II) ions were obtained as 0.9905, 0.9954 and 0.9975, respectively from Fig. 7b. The sensitivity and LOD were obtained as  $0.0622$ ,  $0.1411$  and  $0.1652 \text{ mA } \mu\text{M}^{-1} \text{ cm}^{-2}$  and  $1.2 \times 10^{-10}$ ,  $1.16 \times 10^{-10}$  and  $1.36 \times 10^{-10}$  M for Cd(II), Cu(II) and Hg(II) ions respectively (see Table S1, ESI†). This result was likely due to the formation of Cd–Cu<sup>34</sup> and Cd–Hg<sup>33</sup> intermetallic alloy which increased the sensitivity for Cu(II) and Hg(II) ions. Similar study was performed for Pb(II), Cu(II) and Hg(II) ions and well-defined peaks were observed in DPASV response for each ion as shown in Fig. 7c and sensitivity and LOD from simultaneous detection were found as  $0.0857$ ,  $0.07432$ ,  $0.0809 \text{ mA } \mu\text{M}^{-1} \text{ cm}^{-2}$  and  $2.04 \times 10^{-10}$ ,  $1.11 \times 10^{-10}$  and  $1.92 \times 10^{-10}$  M for Pb(II), Cu(II) and Hg(II) ions respectively from Fig. 7d. The correlation coefficients for Pb(II), Cu(II) and Hg(II) ions were obtained as 0.9952,



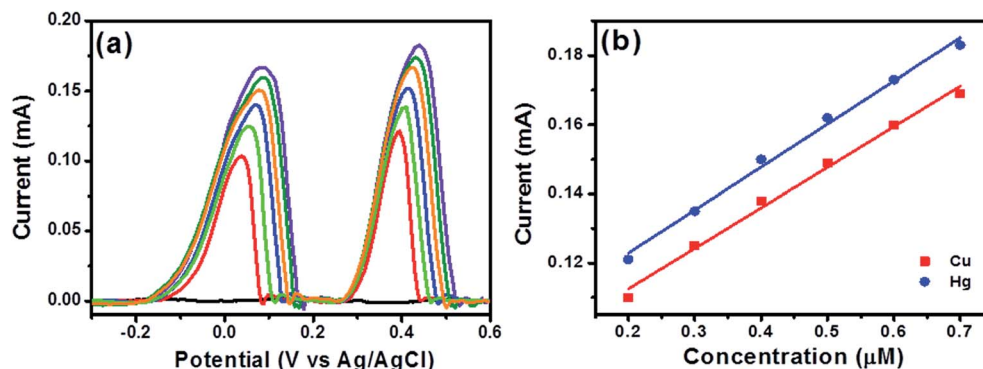


Fig. 6 DPASV response of RGO/NiWO<sub>4</sub> nanocomposite recorded at 1.0 V for 120 s for simultaneous change in concentration of (a) Cu(II) and Hg(II) ions from 0.2 to 0.7  $\mu$ M for each metal ion with their calibration plots (b). Electrolyte: 0.1 M acetate buffer (pH 5.0).

0.9983 and 0.9962, respectively. In presence of Pb(II), it is clear that the sensitivity of Cu(II) and Hg(II) has decreased probably due to the formation of an intermetallic alloy.

After simultaneous analysis of two and three metal ions, simultaneous analysis of all the four ions, Cd(II), Pb(II), Cu(II), and Hg(II) ions at increasing concentrations under the optimal experimental conditions were performed and results are shown in Fig. 8 and the peak currents of all the analyte ions increased with the increasing their respective concentrations. Coexistence of individual peaks in DPASV response for Cd(II), Pb(II), Cu(II),

and Hg(II) ions in the simultaneous detection of different concentrations is observed which is remarkable (see Fig. 8a) and this makes the simultaneous or the selective detection of metal ions on this electrode surface feasible. The correlation coefficients for Cd(II), Pb(II), Cu(II) and Hg(II) ions were obtained as 0.995, 0.984, 0.976 and 0.965, respectively from Fig. 8b. The LOD was found to be  $1.015 \times 10^{-10}$ ,  $1.839 \times 10^{-10}$ ,  $2.269 \times 10^{-10}$  and  $2.789 \times 10^{-10}$  M, respectively. These LOD values are very well below the guideline value given by the World Health Organization (WHO). Also, the sensitivity values for

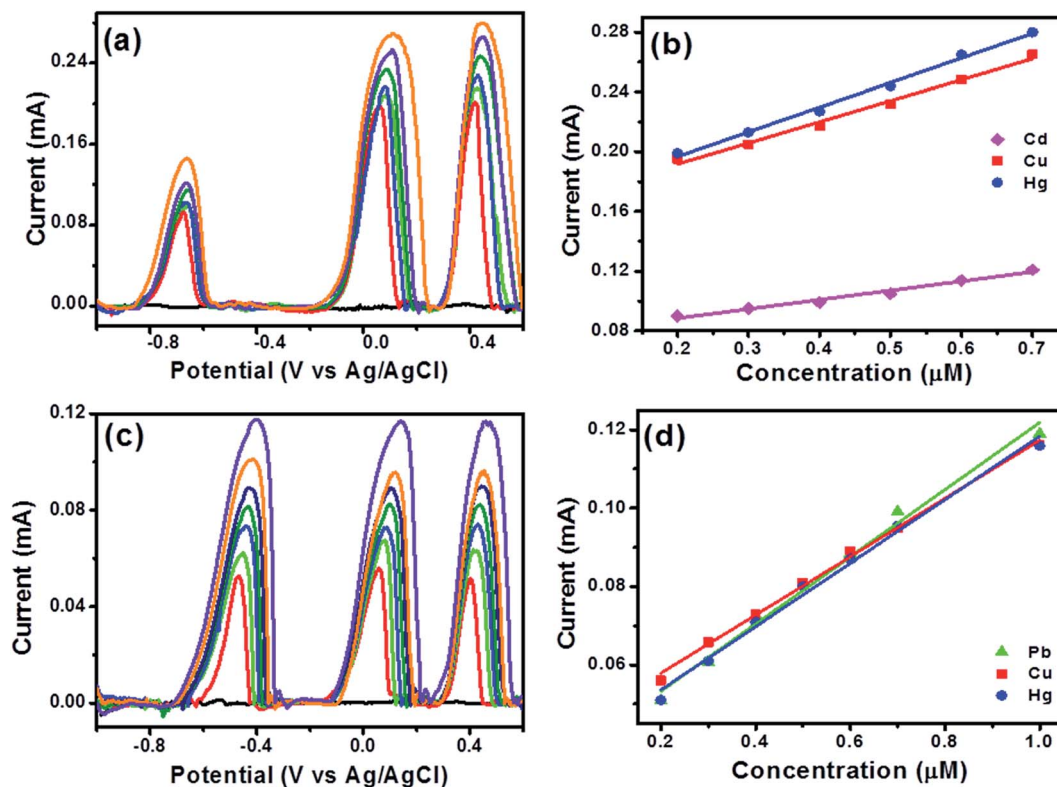


Fig. 7 DPASV response of RGO/NiWO<sub>4</sub> nanocomposite recorded at 1.0 V for 120 s for simultaneous change in concentration of (a) Cd(II), Cu(II) and Hg(II) ions (c) Pb(II), Cu(II) and Hg(II) ions from 0.2 to 0.7  $\mu$ M for each metal ion with their respective calibration plots (b) and (d). Electrolyte: 0.1 M acetate buffer (pH 5.0).





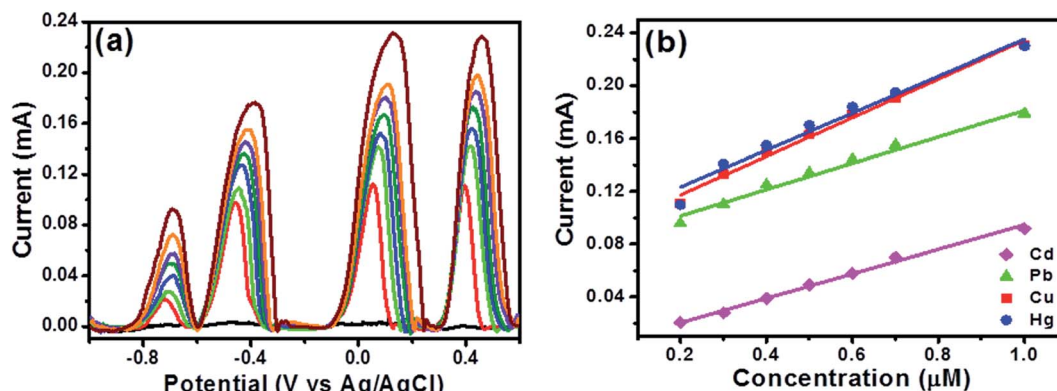


Fig. 8 DPASV response of RGO/NiWO<sub>4</sub> nanocomposite recorded at 1.0 V for 120 s for simultaneous change in concentration of (a) Cd(II), Pb(II), Cu(II) and Hg(II) ions from 0.2 to 1.0  $\mu\text{M}$  for each metal ion with their calibration plots (b). Electrolyte: 0.1 M acetate buffer (pH 5.0).

Cd(II), Pb(II), Cu(II) and Hg(II) ions is found to be 0.00998, 0.0098, 0.01465, 0.01401  $\text{mA } \mu\text{M}^{-1} \text{cm}^{-2}$ , respectively. It is clear that in presence of Pb(II), the sensitivity of Cd(II), Cu(II) and Hg(II) ions has decreased as seen in the Fig. 7c. These results indicated that the interference may occur due to the formation of intermetallic alloys for the four metal ions, which affects the sensitivity and similar observations are made in studies with coexistence of four target metal ions<sup>45</sup> in literature as well and another reason being to the competitive adsorption of the heavy metal ions on the electrode surface. A comparison of LOD values of the proposed electrode with previously reported different electrodes for the determination of Cd(II), Pb(II), Cu(II), Hg(II) ions is provided in Table S2, ESI†. However, the best aspect of the proposed electrode is demonstrated in the next section where it has been tested in complex aqueous media.

**Analysis of Cd(II), Pb(II), Cu(II), and Hg(II) ions in complex aqueous media.** In order to evaluate the real application performance of the proposed electrode, we illustrate the use of RGO/NiWO<sub>4</sub> nanocomposite electrode in complex aqueous media. The electrode is employed for detection of heavy metal ions in carbonated drinks, milk and fruit juice samples which normally contain traceable amounts of Cd(II), Pb(II), Cu(II) and Hg(II) ions due to contamination which goes unnoticed.<sup>25,35</sup> The DPASV responses of RGO/NiWO<sub>4</sub> nanocomposite electrode towards Cd(II), Pb(II), Cu(II) and Hg(II) ions in a carbonated drink, milk, and orange juice is shown in Fig. 9. These samples show the absence of toxic metal ions; however, upon addition of various metal ions, their signals were detected. As the concentrations of metal ions were increased, peak current also increased. Fig. 9a showed increase in peak current of four metal ions in carbonated drink. The correlation coefficients for Cd(II), Pb(II), Cu(II) and Hg(II) ions in carbonated drink were obtained as 0.9965, 0.9999, 0.9883 and 0.9929, respectively as shown in Fig. 9d. Also, LOD was found to be  $1.11 \times 10^{-10}$ ,  $2.04 \times 10^{-10}$ ,  $2.80 \times 10^{-10}$  and  $3.05 \times 10^{-10}$  M, respectively, whereas the sensitivity was found to be 0.0008, 0.0147, 0.0202 and 0.2192  $\text{mA } \mu\text{M}^{-1} \text{cm}^{-2}$ , respectively. The detection of Cd(II), Pb(II), Cu(II) and Hg(II) ions in milk by adding the known amount of toxic metal ions was also performed. All four metal ions were successfully detected as seen in Fig. 9b and the

correlation coefficients were found to be 0.9993, 0.9984, 0.9983 and 0.9927, respectively as shown in Fig. 9e. LOD was found to be  $3.83 \times 10^{-10}$ ,  $2.09 \times 10^{-10}$ ,  $1.74 \times 10^{-10}$  and  $9.34 \times 10^{-10}$  M, respectively, whereas, the sensitivity was found to be 0.0027, 0.0149, 0.0125 and 0.00671  $\text{mA } \mu\text{M}^{-1} \text{cm}^{-2}$ , respectively. Additionally, orange juice as medium for heavy metal ions detection displayed four distinct peaks as shown in Fig. 9c and the correlation coefficients were found to be 0.9905, 0.9999, 0.9870 and 0.9907 for Cd(II), Pb(II), Cu(II) and Hg(II) ions, respectively from the calibration plot (Fig. 9f). LOD was obtained as  $2.72 \times 10^{-10}$ ,  $2.26 \times 10^{-10}$ ,  $2.04 \times 10^{-10}$  and  $2.88 \times 10^{-10}$  M, respectively, whereas the sensitivity was found to be 0.0211, 0.01634, 0.1468 and 0.02073  $\text{mA } \mu\text{M}^{-1} \text{cm}^{-2}$ , respectively. The above results clearly demonstrate that RGO/NiWO<sub>4</sub> nanocomposite is suitable electrode material for detection of heavy metal ions in complex aqueous media as well. The experiments were repeated six times with good reproducibility as the obtained relative standard deviations (RSDs) were all less than 4%. Although the sensing properties of LOD and sensitivity at the RGO/NiWO<sub>4</sub> nanocomposite are comparable with some work (see Table S2, ESI†) reported previously,<sup>15,16,25,29</sup> the electrode described here can provide selective and simultaneous analysis of four toxic heavy metal ions in water and complex aqueous media. The obtained sensing performance is good enough for implication in practice and as demonstrated it is not limited to water and can be extended to many practical applications with some standardization. Thus, the proposed nanocomposite electrode can successfully be used for simultaneous detection of toxic metal ions without any further modifications.

**Stripping mechanism.** To further understand the stripping mechanism of the proposed electrode, XPS analysis on the RGO/NiWO<sub>4</sub> nanocomposite electrode was performed before and after stripping. Initially, the electrode was subjected to 50 cycles and XPS was recorded (before) and DPASV was performed with four heavy metal ions in the solution and following that XPS was recorded (after). XPS spectra of the electrode before and after DPASV are provided in the ESI (Fig. S5 and S6†). The results indicate that the proposed electrode is stable and does not undergo any chemical modification during the cycling and DPASV measurements. In Fig. 10, XPS spectra collected of Cd



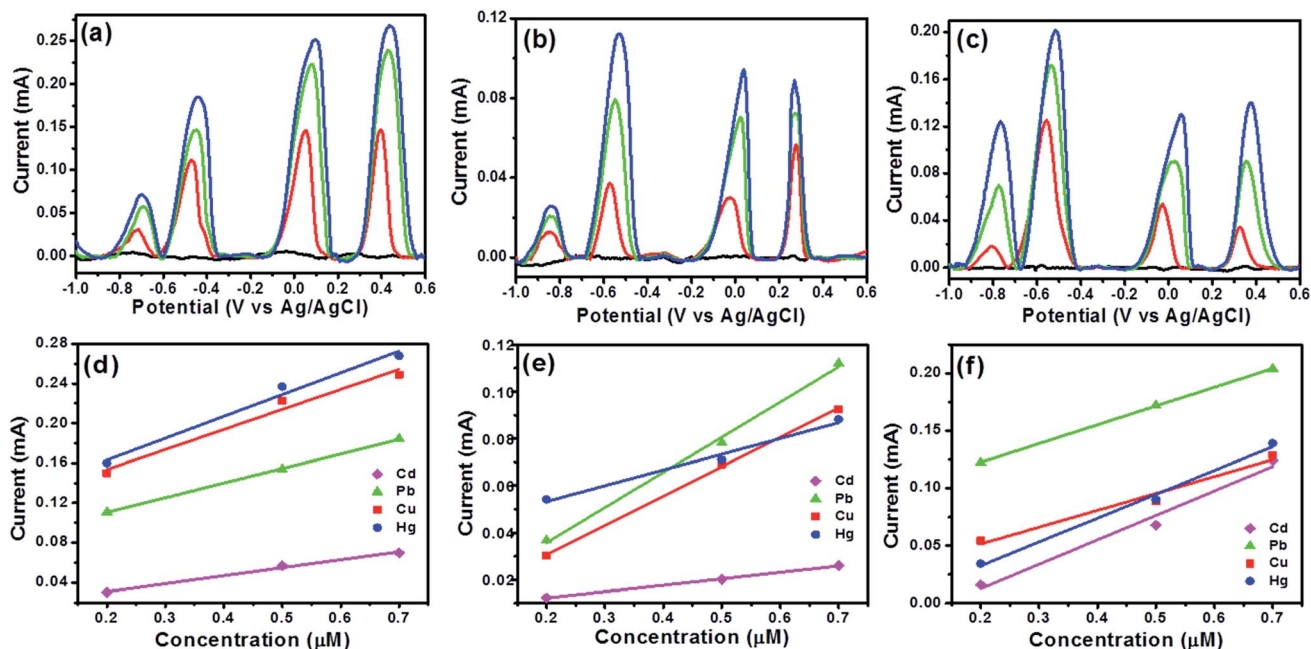


Fig. 9 DPASV response of RGO/NiWO<sub>4</sub> nanocomposite recorded at 1.0 V for 120 s for simultaneous detection of Cd(II), Pb(II), Cu(II) and Hg(II) ions in (a) carbonated drink (b) milk (c) orange juice from 0.2 to 0.7  $\mu\text{M}$  for each metal ion with their respective calibration curve (d–f). Electrolyte: 0.1 M acetate buffer (pH 5.0).

3d, Pb 4f, Cu 2p and Hg 4f levels are shown. The noise in the spectra is attributed to the low concentration of heavy metal ions. The XPS spectra for Cd 3d level (Fig. 10a) showed two broad peaks with binding energy values of 405.7 and 413.8 eV for Cd 3d<sub>5/2</sub> and Cd 3d<sub>3/2</sub> core energy levels, respectively and the binding energy values corresponded to Cd(II). The two peaks

(Fig. 10b) at 139.2 eV and 144.0 eV for Pb 4f<sub>7/2</sub> and Pb 4f<sub>5/2</sub>, respectively and the binding energy values corresponded to Pb(II). The XPS spectra for Cu 2p level (Fig. 10c) showed two broad peaks with binding energy values of 933.2 and 953.0 eV for Cu 2p<sub>3/2</sub> and Cu 2p<sub>1/2</sub> core energy levels, respectively along with a satellite around 944.4 eV confirming Cu(II). In Fig. 10d,

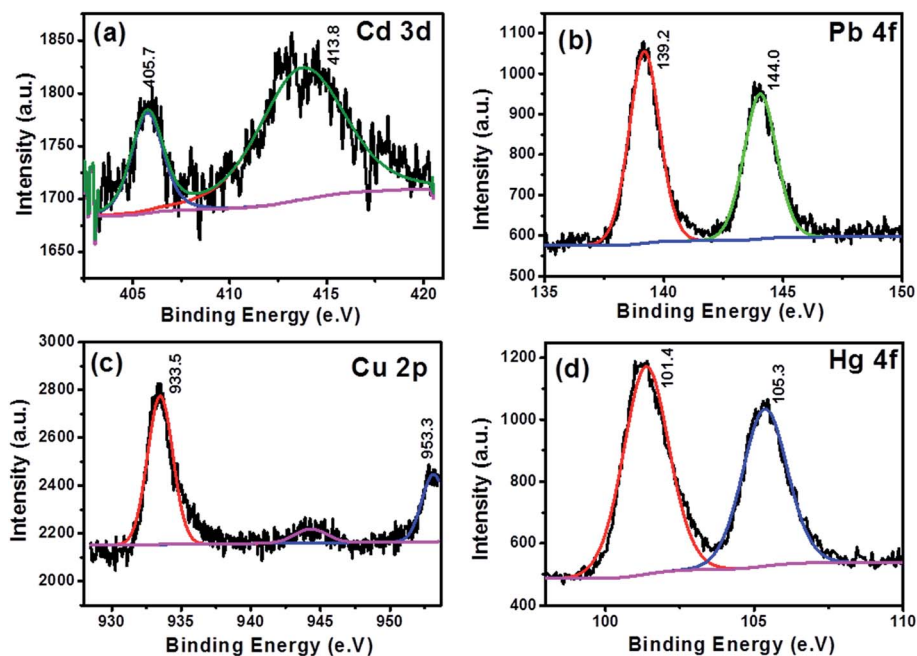


Fig. 10 XPS analysis of RGO/NiWO<sub>4</sub> nanocomposite electrode after cycling and DPASV with heavy metal ions: (a) Cd 3d, (b) Pb 4f, (c) Cu 2p and (d) Hg 4f.



XPS spectra of Hg 4f core level, Hg 4f<sub>7/2</sub> and Hg 4f<sub>5/2</sub> with their binding energy values as 101.4 eV and 105.3 eV, respectively are shown and the binding energy values corresponded to Hg(II). Thus the above results indicate that there is no metal reduction or complex formation between the heavy metal ions and the proposed electrode. The advantage of RGO/NiWO<sub>4</sub> nanocomposite electrode is it has large surface area and pore volume as indicated by BET measurements Fig. S3† that provided high adsorption capacity for the electrode surface. The heavy metal ions get adsorbed on to the surface of the electrode and provide electrochemical detection.<sup>36</sup> The adsorption mechanism is also supported by the CV study shown Fig. S4† wherein the peak current increases linearly as the scan rate is varied which support adsorption controlled process. XPS analysis along with the stripping peak current can provide some insight into the adsorption behaviour of the target heavy metal ions as the stripping peak current values are driven by the amount of target ions adsorbed onto the electrode surface. The basic principle for stripping being, the more target ion is adsorbed onto the surface of the electrode, the more will be released and thus increasing the stripping response. The signal from XPS analysis indicated Cu(II) > Hg(II) > Pb(II) > Cd(II) is the order of preferential adsorption of the heavy metal ion onto the proposed electrode. Though the peak intensities of Cu(II), Hg(II) and Pb(II) are comparable and Cd(II) with a very weak signal thus indicating the least adsorption capacity of the electrode which is consistent with the DPASV electrochemical measurements.

## Conclusions

The present study demonstrates a novel synthesis of RGO/NiWO<sub>4</sub> nanocomposite by the one-pot hydrothermal method and its application in selective and simultaneous detection of heavy metal ions by DPASV. The results suggest that the combination of RGO with binary metal oxide exhibits better and well-resolved peaks of Cd(II), Pb(II), Cu(II), and Hg(II) ions during simultaneous detection as compared to NiWO<sub>4</sub> NPs and RGO sheets. The incorporation of NiWO<sub>4</sub> NPs into RGO sheets not only enhances the specific surface area and the electrical conductivity of the nanocomposite but also improves the electrochemical catalytic activity toward the selectivity of heavy metal ion detection. There is no mutual inference of the metal ions observed in our study, which demonstrates the better catalytic property of RGO/NiWO<sub>4</sub> nanocomposite towards all the four heavy metal ions investigated. LOD was well below the recommended value provided by the World Health Organization (WHO). The nanocomposite developed here is not only effective for simultaneous determination of Cd(II), Pb(II), Cu(II), and Hg(II) ions in their aqueous solutions, but also in the complex media such as carbonated drinks, milk and fruit juices. XPS analysis of target metal ions on the electrode indicate preferential adsorption of heavy metal ions in the following order Cu(II), Hg(II), Pb(II) and least being Cd(II). The study thus suggests the optimal RGO/NiWO<sub>4</sub> nanocomposite developed here to be an ideal candidate for heavy metal ions detection by electrochemical stripping method.

## Conflicts of interest

There are no conflicts to declare.

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