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Synthesis and characterization of a ZnMn_2O_4 nanostructure as a chemical nanosensor: a facile and new approach for colorimetric determination of omeprazole and lansoprazole drugs

Mostafa Y. Nassar, ^{*,a} Eman A. El-Moety^b and M. F. El-Shahat^{*,c}

We have developed a facile method for preparation of ZnMn_2O_4 nanostructures *via* an auto-combustion method using various fuels: urea, glycine and L-alanine. The type of fuel and pH of the combustion media have a significant effect on the combustion products. Glycine fuel generated spinel ZnMn_2O_4 on combustion, and calcination of the sample produced pure spinel ZnMn_2O_4 nanoparticles with an average crystallite size of 19 nm. Whereas, other fuels generated multiphase compounds on combustion. The products were elucidated by means of XRD, FE-SEM, EDS, TEM, FT-IR, and UV-Vis diffuse reflectance spectra. The as-fabricated product was applied to construct a novel chemical nanosensor for the determination of omeprazole and lansoprazole drugs. Different factors influencing the colorimetric determination of the drugs were examined such as contact time, temperature, initial drug concentration, and ZnMn_2O_4 dose. The proposed chemical nanosensor revealed high sensitivity, low detection limit, and a relatively wide linear range (0.80–8.0 and 0.80–8.8 $\mu\text{g mL}^{-1}$ at λ_{max} 454 nm for omeprazole and lansoprazole, respectively).

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1. Introduction

A growing interest in spinel metal oxide nanostructures; AB_2O_4 (A and B = metal), has been observed in the past few years.^{1–4} This intensive interest is due to the unique physical and chemical properties that these oxides possess, which make them suitable for various applications.^{5–9} Among these spinel nanostructures, nano-sized zinc manganite (ZnMn_2O_4) has received a considerable attention due to its various applications such as in sensors, electrodes, lithium rechargeable batteries, specific memory devices, thermistors, and catalysis.^{10–14} Therefore, many research groups have devoted their efforts to synthesizing spinel ZnMn_2O_4 nanostructures. In this light, ZnMn_2O_4 nanostructures with different morphologies and particle sizes have been prepared using various routes such as pyrolysis, solvothermal, electrospinning, hydrothermal, sol–gel, and the solid state method.^{15–18} However, some of these techniques are time-consuming, relatively complicated and expensive, as well as they include sophisticated instruments. Searching for a facile, inexpensive, and less time-consuming process for the production of ZnMn_2O_4 nanostructures is still

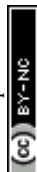
a challenge. Besides, combustion process can fulfill these requirements since it is an inexpensive, simple, short time-consuming, and scalable procedure.^{19–21} In addition, this method produces porous materials with high surface areas owing to the evolution of gasses during the combustion process.^{19–21} In this respect, reports on the synthesis of ZnMn_2O_4 nanostructures using combustion methods are still limited.^{22,23} However, the combustion processes were not extensively investigated in those reports. And the reported preparations applied higher temperatures, during the calcination steps, and produced nanoparticles with larger crystallite sizes in comparison to our current investigation, as will be explained later in the Results and discussion section.

On the other hand, lansoprazole; $\text{C}_{16}\text{H}_{14}\text{F}_3\text{N}_3\text{O}_2\text{S}$, 2-[[[3-methyl-4-(2,2,2-trifluoroethoxy)-2-pyridinyl]methyl]sulphonyl]-1H-benzimidazole, is a proton pump inhibitor, and it is effectively used for treatment of gastric ulcers and duodenal.²⁴ This drug is a chemosensitizing cytotoxic, and it can treat various human tumor cells.²⁵ In addition, omeprazole; $\text{C}_{17}\text{H}_{19}\text{N}_3\text{O}_3\text{S}$, 6-methoxy-2-[[[4-methoxy-3,5-dimethyl-2-pyridinyl]methyl]sulphonyl]-1H-benzimidazole, is a proton pump inhibitor which can be used for treatment of symptomatic gastro-esophageal reflux, gastroduodenal ulcers, and Zollinger–Ellison syndrome.^{26,27} This drug can enhance the oral digoxin bioavailability by decreasing the acid production in the stomach and increasing the gastric permeability to digoxin.²⁶ Some methods have been developed for quantitative determination of

^aChemistry Department, Faculty of Science, Benha University, Benha 13815, Egypt. E-mail: m_y_nassar@yahoo.com; m_y_nassar@fsc.bu.edu.eg; Tel: +20 1068727555

^bNational Cancer Institute, Cairo, Egypt

^cChemistry Department, Faculty of Science, Ain Shams University, Cairo, Egypt. E-mail: elshahatmf@hotmail.com; Tel: +20 1128323115



lansoprazole and omeprazole drugs in formulations and biological fluids such as HPLC,^{28,29} capillary electrophoresis,^{30,31} spectrophotometry,³² and voltammetry.^{33,34} However, the demand for a simple, reliable, and inexpensive analytical assay of the aforementioned drugs is growing due to their importance. Additionally, it is reported that organic sulfide compounds can be readily converted into sulfoxide (R_2SO) and/or sulfone (R_2SO_2) compounds by oxidation.^{35,36}

Based on the approach presented, we have herein developed a facile synthetic method for the preparation of $ZnMn_2O_4$ nanoparticles *via* an auto-combustion method. Therefore, we have investigated different parameters influencing the product species and their crystallite sizes such as fuel type and pH of the combustion media. Moreover, we have developed, for the first time, a simple, accurate, and sensitive colorimetric determination method for omeprazole and lansoprazole drugs in their pharmaceutical formulations based on the as-prepared $ZnMn_2O_4$ nanostructure as a chemical sensor.

2. Experimental

2.1. Materials and reagents

Zinc nitrate ($Zn(NO_3)_2 \cdot 6H_2O$), manganese nitrate ($Mn(NO_3)_2 \cdot 4H_2O$), potassium chloride (KCl), L-alanine ($CH_3CH(NH_2)COOH$), glycine (NH_2CH_2COOH), urea (NH_2CONH_2), were provided by Sigma-Aldrich Company. Nitric acid (HNO_3) was supplied by Carlo Erba Company, France. Ammonia solution 33% (NH_4OH) was purchased from El Nasr Pharmaceutical Chemicals Company (ADWIC) – Company, Egypt. Lansoprazole and omeprazole (Scheme 1) were kindly provided by El Arabeya Company for Pharmaceutical and Chemical Industries, Egypt. All other reagents and chemicals used, as received without any further purification, were of analytical grade.

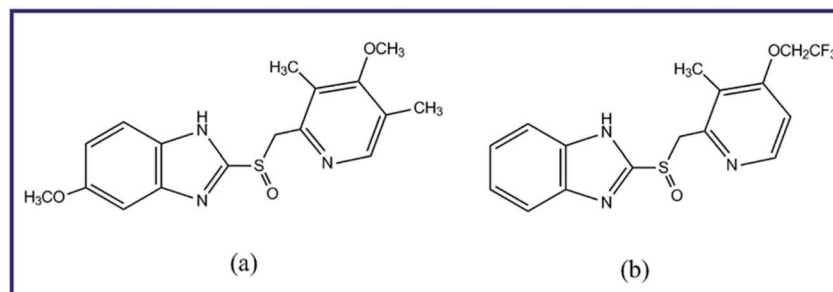
2.2. Fabrication of $ZnMn_2O_4$ nanostructures (chemo-sensors) using different fuels

We have proposed an auto-combustion method to prepare $ZnMn_2O_4$ nanoparticles employing three different fuels: L-alanine, glycine and urea, and the as-prepared $ZnMn_2O_4$ products will be denoted as A, G, and U, respectively. It is noteworthy that in the present combustion method, the stoichiometric compositions of the redox mixtures are calculated based on that the total reducing (F) valency of the fuel and the

total oxidizing (O) valency of the mixed metal nitrate salts verify that $\Phi_c = (F/O) = 1$ (where, Φ_c is the equivalence ratio).^{21,37} This was carried out to release the maximum quantity of heat from the combustion process. In a typical procedure: zinc nitrate (3.00 g, 10.08 mmol, 1.00 eq.) and manganese nitrate (5.06 g, 20.02 mmol, 2.00 eq.) were dissolved in 40 mL bi-distilled water. Then, a glycine aqueous solution; 40 mL, (2.52 g, 33.57 mmol, 3.33 eq.) was added to the stirring mixed zinc/manganese nitrate aqueous solution. The reaction blend pH was adjusted using ammonia solution (1 : 1) to pH 6. The reaction blend was then stirred at 80 °C until it turned into a viscous liquid. After that, the temperature of the reaction mixture was increased to 300 °C, while a vigorous self-propagating combustion reaction took place and lasted for *ca.* 5 min. The as-burnt material (G_6) was ground then calcined at 500 °C for 2 h producing $ZnMn_2O_4$ products denoted as G_{6-500} . This procedure was repeated using glycine fuel at different pH values such as 3 and 10 using nitric acid or ammonia solution (1 : 1), producing products labelled G_3 and G_{10} , respectively. Besides, we synthesized $ZnMn_2O_4$ products using similar combustion procedure and calcination temperature, while different fuels such as urea and L-alanine were utilized and the corresponding products were labelled as U and A, respectively. In addition, the aforementioned combustion and calcination steps were repeated while pH of the combustion media was adjusted at 3, 6 or 10 using nitric acid or ammonia solution, respectively. And the produced $ZnMn_2O_4$ products were referred to as U_3 , U_6 , U_{10} , A_3 , A_6 and A_{10} , according to the corresponding fuel and pH.

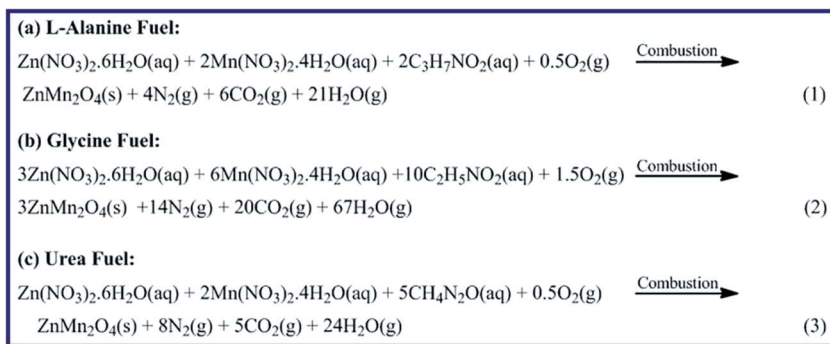
2.3. Materials characterization

The composition and phase purity of the as-fabricated products were analyzed utilizing powder X-ray diffraction (XRD) patterns measured by 18 kW diffractometer (Bruker; model D8 Advance) equipped with a monochromated Cu-K α radiation ($\lambda = 1.54178$ Å). Morphologies of the as-fabricated $ZnMn_2O_4$ products were examined using a field emission scanning electron microscope (FE-SEM) connected with a microscope (FE-SEM) (JEOL JSM-6500F). And TEM images of the as-fabricated $ZnMn_2O_4$ products were taken using a high-resolution transmission electron microscope (HR-TEM) (JEM-2100) at 200 kV accelerating voltage. The functional groups of the as-fabricated $ZnMn_2O_4$ products were determined by collecting their FT-IR spectra in the range of 4000–400 cm^{-1} employing an FT-IR spectrometer



Scheme 1 Chemical structures of omeprazole (a) and lansoprazole (b) drugs.





Scheme 2 The proposed combustion reactions of zinc and manganese nitrates with L-alanine, glycine, and urea fuels.

(Thermo Scientific; model Nicolet iS10) utilizing KBr pellets. The UV-Visible spectra of the drug solutions, under investigation, were measured on an UV-Vis spectrophotometer (Jasco, model v670).

2.4. Colorimetric determination of lansoprazole and omeprazole drugs

For the determination of lansoprazole and omeprazole drugs with an initial concentration of 10 mg L^{-1} (dissolved in ethanol), 0.05 g of the as-prepared ZnMn_2O_4 chemosensor was added to 20 mL of the prepared drug solution, and the pH value of the solution was adjusted using 0.5 M HNO_3 or NaOH solutions. After that, the volume of the reaction mixture was adjusted to 25 mL by the addition of bi-distilled water, and the mixture was then allowed to stir for a specific time. Notably, the color of the drug solution after addition of the chemosensor nanomaterial turned into yellow. Different parameters influencing the developing of that yellow color have been investigated such as contact time, pH, ZnMn_2O_4 dose, initial drug concentration, and temperature. The absorbance spectra of the supernatant were measured employing an UV-Vis spectrophotometer at $\lambda_{\text{max}} = 454 \text{ nm}$, after separation of the suspension by centrifugation. Interestingly, the order of addition does not influence the accuracy, sensitivity, and reproducibility of the present developed method.

3. Results and discussion

3.1. Auto-combustion fabrication and characterization of ZnMn_2O_4 chemosensor

In our earlier study, we have reported on the preparation of MgAl_2O_4 , MgO , and TiO_2 nanostructures *via* a combustion procedure using metal nitrate salts and different fuels.^{19–21} And we concluded that the fuel type and equivalence ratio (Φ_c) value influenced the crystallite sizes of the products. However, we have not examined the effect of pH of the combustion media on the combustion process. Therefore, in this light, we have explored the influence of fuel type and combustion media pH on the combustion of zinc/manganese nitrate (oxidant). As we will explain later, the type and amount of the fuel, as well as the pH value of the combustion media have a significant impact on the combustion products, their morphologies, and their crystal

sizes. Hence, the combustion reactions between zinc nitrate/manganese nitrate and a fuel such as glycine, L-alanine, or urea, were carried out at different pH values and at fuel-to-oxidant equivalence ratio (Φ_c) of 1. The produced ZnMn_2O_4 nanostructure obtained at the optimum conditions was then applied to colorimetric determination of omeprazole and lansoprazole drugs. The proposed combustion reactions are presented in Scheme 2. The combustion and calcined products were analyzed using XRD, FR-IR, SEM, UV-Vis reflectance, and TEM techniques.

3.1.1. XRD study. The crystal structures of the combustion products were investigated using the X-ray diffraction technique. Fig. 1 through 3 present the XRD patterns of the combustion and calcined products using L-alanine, glycine, and urea fuels at different pH values (3, 6, and 10) at $\Phi_c = 1$. Inspection of the figures revealed that the fuel type and pH of the combustion media have a significant impact on the phase composition and crystallite size of the products. It is reported that pH of the starting solutions of the combustion media can prevent selective precipitation and/or phase segregation during

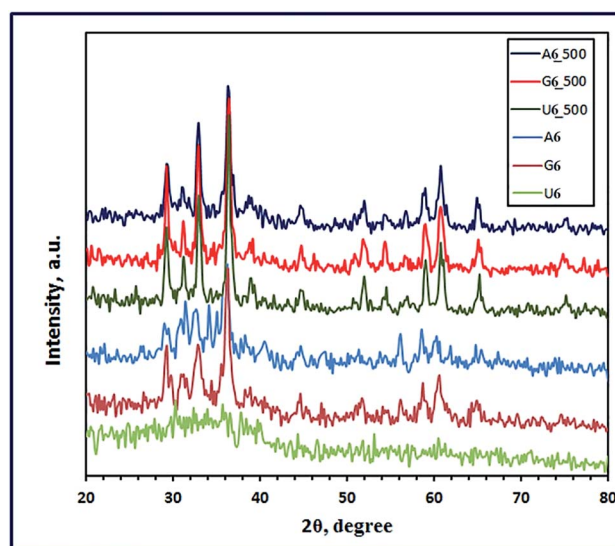


Fig. 1 XRD patterns of the as-fabricated ZnMn_2O_4 products after combustion (A_6 , G_6 , and U_6) and calcination ($A_{6,500}$, $G_{6,500}$, and $U_{6,500}$) using L-alanine, glycine and urea fuels, respectively.



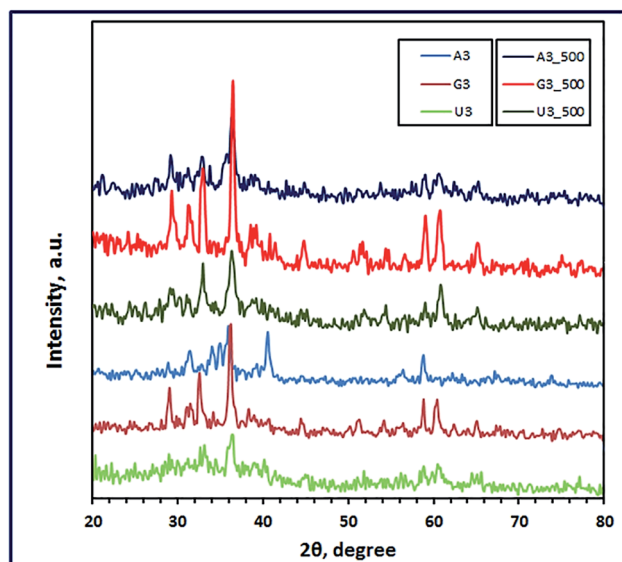


Fig. 2 XRD patterns of the as-fabricated ZnMn_2O_4 products (at pH 3) after combustion (A_3 , G_3 , and U_3) and calcination (A_{3_500} , G_{3_500} , and U_{3_500}) using L-alanine, glycine and urea fuels, respectively.

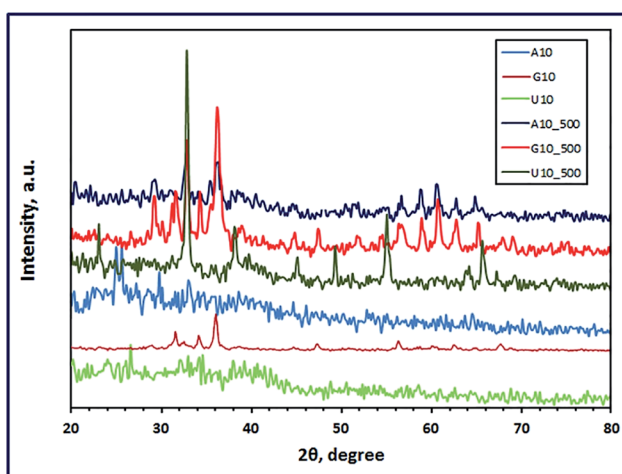


Fig. 3 XRD patterns of the as-fabricated ZnMn_2O_4 products (at pH 10) after combustion (A_{10} , G_{10} , and U_{10}) and calcination (A_{10_500} , G_{10_500} , and U_{10_500}) using L-alanine, glycine and urea fuels, respectively.

the combustion process which may cause phase inhomogeneity of the reaction blend.³⁸ In the current investigation, we used different fuels such as urea, L-alanine, and glycine, which might act simultaneously as chelating agents as well. The formation of metal coordination or metal chelation complexes is a pH dependent. And the produced soluble chelation complexes are an essential point in preventing the selective precipitation of the cations of interest and this can support the homogeneity of the reaction media during solvents evaporation throughout the combustion process.^{38,39} Moreover, hydrolysis of the cation compounds and their complexes can be controlled by the pH value of the reaction media.⁴⁰ Therefore, during the combustion process, there will be a competition between two main reactions: (1) formation of metal chelations and (2) hydrolysis of the

formed metal chelations and other metal compounds. And these two main mechanisms can tune the phase, morphology, and/or crystallite size of the products of the combustion process. Based on this discussion, we can explain our following results. Fig. 1 revealed that L-alanine fuel resulted in a mixture (A_6) of ZnMnO_3 , ZnO , and Mn_3O_4 through the combustion process at pH 6. Calcination of this mixture at 500 °C for 2 h produced pure product (A_{6_500}), corresponding to tetragonal ZnMn_2O_4 (JCPDS card no. 24-1133, space group: $I4_1/amd$). The diffraction peaks of the ZnMn_2O_4 product appeared at 2θ values of ca. 29.32°, 31.23°, 33.00°, 36.41°, 44.63°, 58.92°, 60.82° and 64.82° corresponding to the reflection planes values of (112), (200), (103), (211), (220), (321), (224), and (400), respectively. Fig. 2 and 3 revealed that combustion at pH value of 3 and 10 brought about a mixture of ZnMn_2O_4 , ZnO , and MnO (A_3), and a mixture of $\text{Zn}(\text{OH})(\text{NO}_3)(\text{H}_2\text{O})$ and $(\text{Mn}_2\text{Zn})_7(\text{OH})_{10}(\text{CO}_3)_2\text{Zn}_5(\text{OH})_8$ (A_{10}), respectively. And calcination of the combustion products did not give a pure ZnMn_2O_4 product, but it resulted in different mixed oxides (A_{3_500} and A_{10_500}), as shown in Fig. 2 and 3. L-Alanine is an amino acid which can exist in three forms: $^+\text{NH}_3\text{CH}(\text{CH}_3)\text{COOH}$, $^+\text{NH}_3\text{CH}(\text{CH}_3)\text{COO}^-$ (zwitterionic form), and $\text{NH}_2\text{CH}(\text{CH}_3)\text{COO}^-$, as dominant species. This will depend on the pH values of the media: 3, 6, and 10, since the isoelectric point value of this amino acid is 6. The rate of formation of L-alanine–metal complexes is low at acidic pH values, and it enhances with increasing the pH value of the media. Besides, in alkaline media the formed chelate complexes may undergo hydrolysis forming hydroxide species, and this may result in formation of a mixture of different metal oxide species on combustion, as shown in Fig. 3.

Interestingly, the combustion reaction using glycine fuel at pH 6 and 3 gave pure tetragonal ZnMn_2O_4 (JCPDS card no. 24-1133, space group: $I4_1/amd$) with crystallite size values of 11 (G_6) and 27 nm (G_3), respectively, as shown in Fig. 1 and 2. No other peaks were observed for other crystalline products in the XRD patterns. In addition, calcination of these combustion products gave ZnMn_2O_4 (G_{6_500}) with crystallite size of 19 nm, and a mixture of ZnMn_2O_4 , Mn_3O_4 , and Mn_5O_8 (G_{3_500}), as depicted in Fig. 1 and 2, respectively. Glycine is also an amino acid prevailing specific glycine species, based on pH values of the starting solutions of the combustion media. At pH 3, glycine may exist mainly as $^+\text{NH}_3\text{CH}_2\text{COOH}$ and the coordination ability of this species with zinc and manganese cations is weak; moreover, glycine may exist mainly as $^+\text{NH}_3\text{CH}_2\text{COO}^-$ (zwitterion, isoelectric point is ~ 6) at pH 6 and the complexing ability of this ion is higher than the previous one.⁴¹ However, both pH values resulted in pure ZnMn_2O_4 nanostructures on combustion. The crystallite size of the burnt nanoparticles at pH 6 was smaller than that at pH 3. This may be attributed to that zwitterionic glycine–metal complexes are less soluble at pH 6. And the formed nucleates may prefer to coagulate forming gel quickly during the evaporation of the solvent in the combustion process. Therefore, the primary small particles (crystallites) do not have much time in the combustion process to grow into larger crystallites producing smaller crystallites.⁴⁰ Increasing the pH of the combustion to 10 enhanced the chelation ability of the glycine species at that pH ($\text{NH}_2\text{CH}_2\text{COO}^-$); however, the



rate of hydrolysis also increased forming different metal hydroxide species as precipitate along with various glycine-metal chelates. And this might explain producing of a mixture of different manganese-zinc oxides on combustion at pH 10.

On the other hand, using urea as a fuel produced a mixture of ZnMnO_3 and Mn_3O_4 , or a mixture of ZnMn_2O_4 and Mn_3O_4 at pH 6 (U_6) or at pH 3 (U_3), respectively. Calcination of these combustion products at 500 °C for 2 h produced pure tetragonal ZnMn_2O_4 (U_{6-500} and U_{3-500}) (JCPDS card no. 24-1133, space group: $I4_1/amd$) with a crystallite size of 25 and 20 nm, respectively, as shown in Fig. 1 and 2. We have not observed any other peaks for other crystalline products in the XRD patterns. Moreover, adjusting the pH of the combustion media to 10 gave $(\text{Mn}_2\text{Zn})_7(\text{OH})_{10}(\text{CO}_3)_2$ precursor (U_{10}) on combustion, as shown in Fig. 3. However, calcination of that precursor produced a mixture of ZnMn_2O_4 and Mn_2O_3 (U_{10-500}). The obtained results exhibited that these pH values do not have a remarkable influence on the combustion products. This may be attributed to that during heating the solutions prior to the combustion process, urea decomposes into ammonia and CO_2 which can increase the alkalinity of the media enhancing hydrolysis of the cations into different hydroxides.⁸ These metal hydroxides might be converted into various metal oxides on combustion. However, on sintering the burnt oxides, produced at pH 3 and 6, may react in the solid state forming pure ZnMn_2O_4 nanostructures.

Based on the obtained results, we have arrived at a conclusion that pure ZnMn_2O_4 nanoparticles can be synthesized using L-alanine and urea fuels at pH 6 on combustion followed by calcination at 500 °C for 2 h. However, the combustion process resulted in directly pure ZnMn_2O_4 nanoparticles when glycine fuel was used and the combustion media pH was 6. And calcination of the combustion product in this case did not influence the product composition but it increased the crystallite size only. On the other side, adjusting pH to 3 produced pure ZnMn_2O_4 product only after calcination of the combustion products at 500 °C for 2 h when glycine and urea were used as fuels. Consequently, the present combustion method, in which $\Phi_c = 1$, gave pure spinel ZnMn_2O_4 nanoparticles with much smaller crystallite size compared to the reported one.²² In that reported method, the applied Φ_c was not equal to 1; therefore,

the authors used higher calcination temperature, 600 °C, for a longer time (6 h) so that the produced crystallite size was ca. 50 nm.²² Notably, the average crystallite sizes of the combustion and calcined pure products were estimated employing the Debye-Scherrer equation (eqn (4)) and listed in Table 1.⁴²

$$D = 0.9\lambda/\beta \cos\theta_B \quad (4)$$

where, β , λ , and θ_B are the full width of the diffraction peak at half maximum (FWHM), X-ray radiation wavelength (nm), and Bragg diffraction angle, respectively.

3.1.2. FT-IR study. The chemical compositions of the combustion and calcination products were elucidated by using the FT-IR spectra. In this light, Fig. 4 revealed the FT-IR spectra of the combustion products generated using L-alanine, glycine, and urea fuels at fuel-to-oxidant equivalence ratio (Φ_c) of 1 and at different pH values of the combustion media. The FT-IR results exhibited that glycine fuel was the optimum fuel for the current combustion procedure because this fuel generated ZnMn_2O_4 phase on combustion, at pH 3 (G_3) and pH 6 (G_6) of the combustion media (Fig. 4). This conclusion is based on the two characteristic vibrational frequencies appeared at ca. 496 and 615 cm^{-1} which can be attributed to ZnMn_2O_4 .⁴³ However, this fuel at the aforementioned pH conditions produced some organic residue, and this was attributed to appearance of some vibrational frequencies at ca. 1050 (C–O stretching) and 1426 cm^{-1} (CH_2 stretching). This is consistent with the published results.^{43,44} Notably, increasing the pH of the combustion media to pH 10, resulted in different phases (G_{10}) on combustion as shown in Table 2. Nevertheless, the other used fuels (urea and alanine) did not give pure ZnMn_2O_4 phase on combustion, as shown in Fig. 4 and Table 2. However, alanine fuel produced phases (A_3): ZnMn_2O_4 , ZnO, and MnO (at pH 3, A_3), phases (A_6): ZnMnO_3 , ZnO, and Mn_3O_4 (at pH 6), and phases (A_{10}): $\text{Zn}(\text{OH})(\text{NO}_3)(\text{H}_2\text{O})$ and $(\text{Mn}_2\text{Zn})_7(\text{OH})_{10}(\text{CO}_3)_2 \cdot \text{Zn}_5(\text{OH})_8$ (at pH 10), along with some organic residues,^{2,43–49} as shown in Fig. 4 and Table 2. Urea fuel gave also different phases (U_3 , U_6 , U_{10}) on combustion at different pH values. Based on glycine fuel results, we focused on the combustion products (A_6 , G_6 , U_6) generated using the three fuels at pH 6, for comparison purpose. These combustion products were calcinated at 500 °C

Table 1 Influence of fuel, pH and calcination temperature on the produced phases of the combustion processes

Fuel	Equivalence ratio, Φ_c	pH	Product phases after combustion	Product phases after calcination at 500 °C for 1 h	Crystallite size (after combustion, after calcination) (nm)	Product label (after combustion, calcination)
L-Alanine	1	3	ZnMn_2O_4 , ZnO, MnO	ZnMn_2O_4 , Mn_3O_4	—, —	A_3 , A_{3-500}
	1	6	ZnMnO_3 , ZnO, Mn_3O_4	ZnMn_2O_4	—, 28	A_6 , A_{6-500}
	1	10	$\text{Zn}(\text{OH})(\text{NO}_3)(\text{H}_2\text{O})$, $(\text{Mn}_2\text{Zn})_7(\text{OH})_{10}(\text{CO}_3)_2 \cdot \text{Zn}_5(\text{OH})_8$	ZnMn_2O_4 , Mn_3O_4 , Mn_2O_3	—, —	A_{10} , A_{10-500}
Glycine	1	3	ZnMn_2O_4	ZnMn_2O_4 , Mn_3O_4 , Mn_5O_8	27, —	G_3 , G_{3-500}
	1	6	ZnMn_2O_4	ZnMn_2O_4	11, 19	G_6 , G_{6-500}
	1	10	ZnMn_2O_4 , ZnO, Mn_3O_4 , Mn_5O_8	ZnMn_2O_4 , ZnO	—	G_{10} , G_{10-500}
Urea	1	3	Mn_3O_4 , ZnMn_2O_4	ZnMn_2O_4	—, 20	U_3 , U_{3-500}
	1	6	Mn_3O_4 , ZnMnO_3	ZnMn_2O_4	—, 25	U_6 , U_{6-500}
	1	10	$(\text{Mn}_2\text{Zn})_7(\text{OH})_{10}(\text{CO}_3)_2$	ZnMn_2O_4 , Mn_2O_3	—, —	U_{10} , U_{10-500}



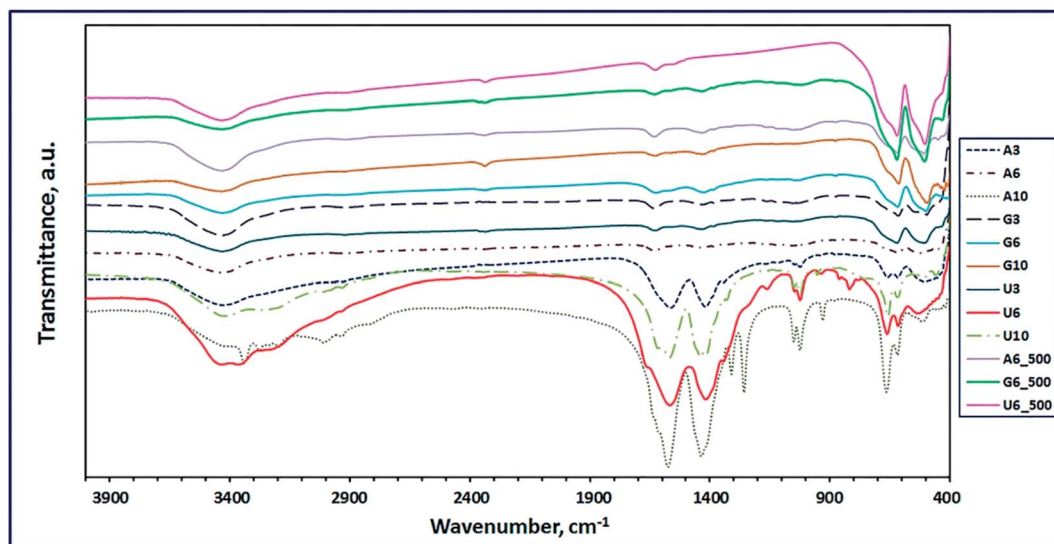


Fig. 4 FT-IR spectra of the as-fabricated ZnMn_2O_4 products after combustion and calcination (A_{6_500} , G_{6_500} , and U_{6_500}) using L-alanine, glycine and urea fuels, respectively.

Table 2 FT-IR vibrational frequencies of the as-synthesized products

Product	Label	$\nu_{\text{Mn-O}}$, cm^{-1}	$\nu_{\text{Zn-O}}$, cm^{-1}	ν_{MnCO_3} , cm^{-1}	ν_{ZnCO_3} , cm^{-1}	$\nu_{\text{C-O}}$, cm^{-1}	ν_{CH_2} , cm^{-1}	ν_{OH} , cm^{-1}	δ_{OH} , cm^{-1}
ZnMn_2O_4	A ₃	600, 617	510	—	—	1024	1419	3430	1563
ZnO		—	510	—	—				
MnO		600	—	—	—				
ZnMnO_3	A ₆	617	500 ^{sha}	—	—	1054	1432	3433	1613
ZnO		—	500 ^{sha}	—	—				
Mn_3O_4		520, 617	—	—	—				
$\text{Zn}(\text{OH})(\text{NO}_3)(\text{H}_2\text{O})$	A ₁₀	—	463	—	—	1024	1436	3339	1575
$(\text{Mn}_2\text{Zn})_7(\text{OH})_{10}(\text{CO}_3)_2\text{Zn}_5(\text{OH})_8$		522, 665	463	1253, 105, 930, 665, 463	1253, 1050, 930, 665, 463				
ZnMn_2O_4	G ₃	615	496	—	—	1052	1431	3432	1631
ZnMn_2O_4	G ₆	618	501	—	—	1050 ^{sha}	1426	3432	1623
ZnMn_2O_4	G ₁₀	613	494	—	—	—	1436	3434	1626
ZnO		—	458	—	—				
Mn_3O_4		613, 494	—	—	—				
Mn_5O_8		877, 613	—	—	—				
Mn_3O_4	U ₃	508, 621	508	—	—	1053	1433	3433	1626
ZnMn_2O_4		508, 621	—	—	—				
Mn_3O_4	U ₆	534, 616	534	—	—	1024	1417	3432	1566
ZnMnO_3		534, 616	—	—	—				
$(\text{Mn}_2\text{Zn})_7(\text{OH})_{10}(\text{CO}_3)_2$	U ₁₀	658	460	1416, 952, 658	1416, 952, 460	1049	1440	3427	1567
ZnMn_2O_4	A _{6_500}	620	508	—	—	1052	1431	3434	1631
ZnMn_2O_4	G _{6_500}	621	505	—	—	—	—	3434	1632
ZnMn_2O_4	U _{6_500}	620	504	—	—	—	—	3434	1626

^a sh = shoulder.

for 2 h to eliminate the organic residues and converting some multi-phases products into single-phase products. The FT-IR spectra of the calcined products (A_{6_500} , G_{6_500} , U_{6_500}) are depicted in Fig. 4. Interestingly, the calcined products; G_{6_500} and U_{6_500} , of the combustion samples generated from glycine and urea fuels, respectively, are composed of pure ZnMn_2O_4 phase. This was clear by the presence of only the characteristic vibrational frequencies of ZnMn_2O_4 product (ca. 508 and

620 cm^{-1}). However, calcination of the combustion products generated using alanine fuel did not produce pure product, and it generated organic residue along with ZnMn_2O_4 product (A_{6_500}), as listed in Table 2. Fig. 4 also shows two characteristic vibrational absorptions corresponding to bending and stretching vibrations of OH groups of adsorbed water on all products at ca. 1600 and 3430 cm^{-1} , respectively. These results are



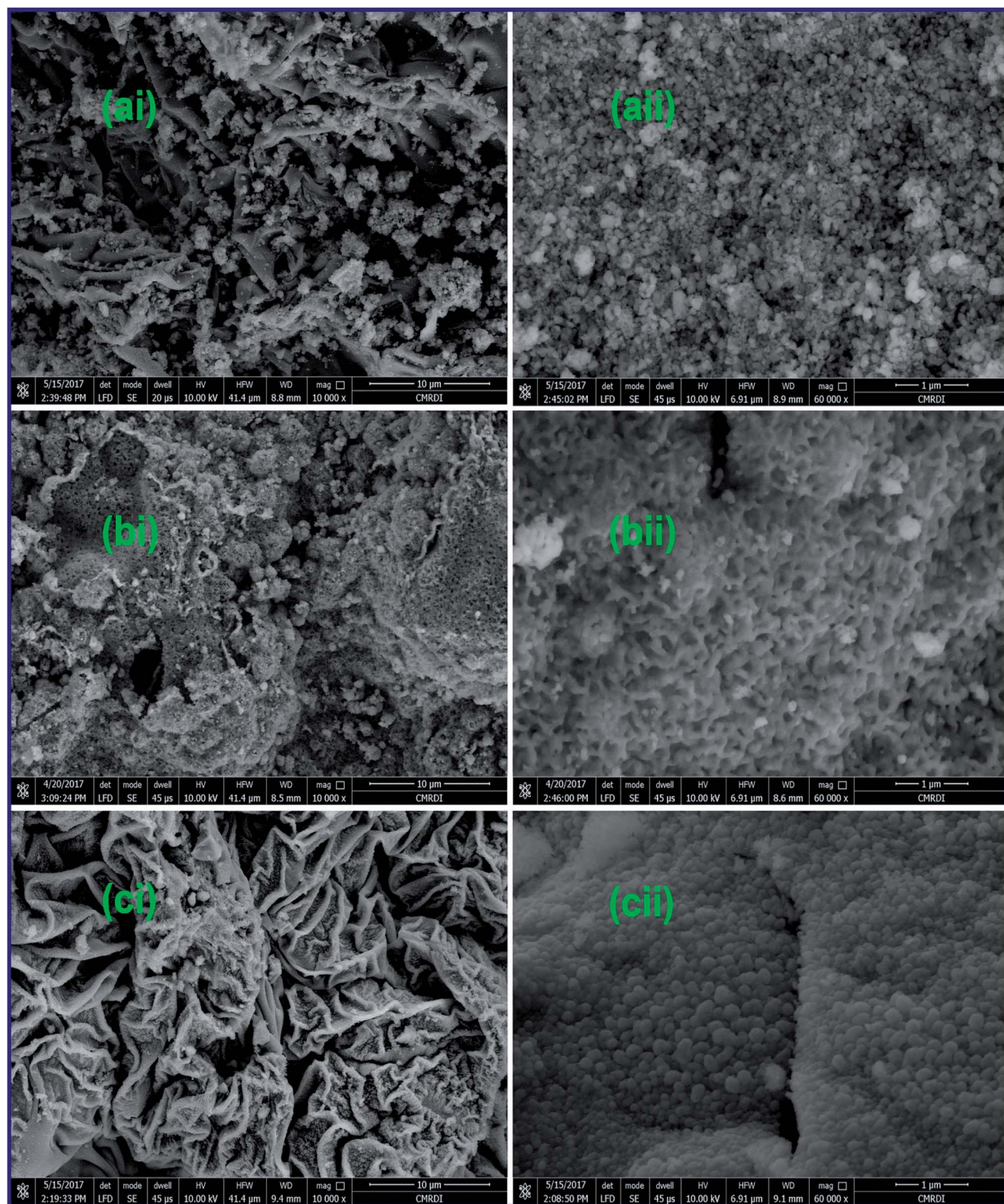
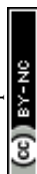


Fig. 5 FE-SEM images of the as-fabricated ZnMn_2O_4 products after calcination (A_{6-500} (a), G_{6-500} (b), and U_{6-500} (c)) by L-alanine, glycine and urea fuels, respectively.

consistent with the reported data.^{6,20,50–54} The vibrational frequencies of the produced products are presented in Table 2.

3.1.3. Morphology investigation. The morphology of the as-fabricated ZnMn_2O_4 products: A_{6-500} , G_{6-500} , and U_{6-500} ,

calcined at 500 °C was investigated by using a field-emission scanning electron microscope (FE-SEM). As shown in Fig. 5(ai), L-alanine fuel generated ZnMn_2O_4 nanostructure (A_{6-500}) with aggregates in the form of irregular and curled



layers. However, a high-magnification FE-SEM image (Fig. 5(aii)) exhibits that those layers are composed of sphere-like particles with an average diameter of *ca.* 80 nm. Fig. 5(aii) also shows the existence of in-plane pores within the ZnMn_2O_4

layers. However, there are some dark spots appeared in those FE-SEM images (Fig. 5(ai and ii)) indicating the presence of some carbon residues along with the ZnMn_2O_4 product. Fig. 5(bi) revealed that glycine fuel produced ZnMn_2O_4 product (G_{6_500}) with porous irregular agglomerates composed of porous peanut-like particles with an average size of *ca.* 300 nm, as shown in a high-magnification FE-SEM image (Fig. 5(bii)). Whereas, the calcined ZnMn_2O_4 product particles (U_{6_500}), produced from combustion with urea fuel, were aggregated in thick curled layers and these layers were agglomerated in flower-like structures as shown in a low-magnification FE-SEM image (Fig. 5(cii)). Close inspection of the high-magnification FE-SEM image (Fig. 5(cii)) of these layers, revealed that the layers consisted of uniform sphere-like particles with an average diameter of *ca.* 100 nm. Notably, the porous structures of the ZnMn_2O_4 products are produced during the combustion process by action of the released gases (*e.g.* CO_2 and N_2) at elevated temperatures.^{20,21,53} The detailed structures of the ZnMn_2O_4 products calcined at 500 °C are displayed in TEM images (Fig. 6). Fig. 6(a–c) showed that the products generated by the three different fuels have similar morphologies. And the produced ZnMn_2O_4 particles have a hexagon, sphere, and cube-like structures with an average particle size of *ca.* 30, 20, 25, for

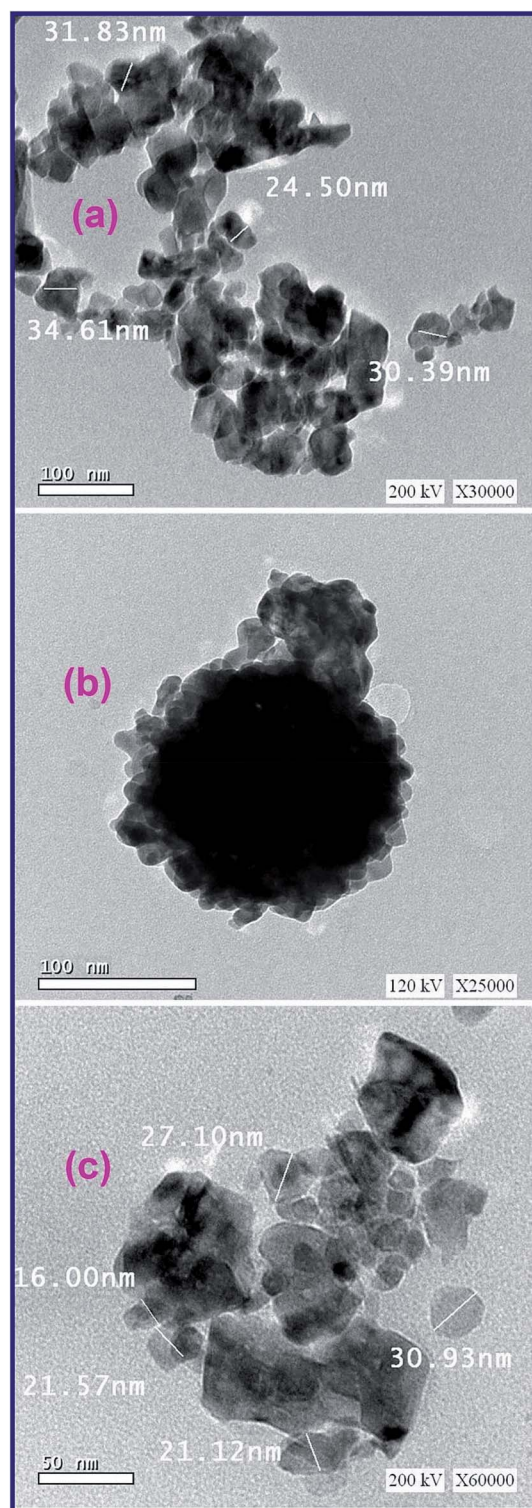


Fig. 6 TEM images of the as-fabricated ZnMn_2O_4 products after calcination (A_{6_500} (a), G_{6_500} (b), and U_{6_500} (c)) generated by L-alanine, glycine and urea fuels, respectively.

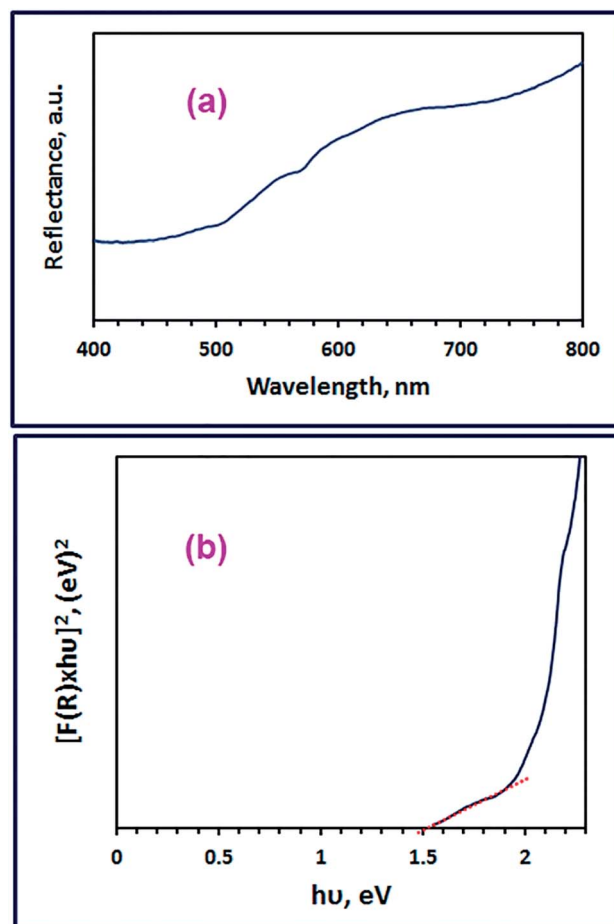


Fig. 7 Diffuse reflectance spectrum (a), and plot $[F(R)xhu]^2$ against $h\nu$ (b) of the as-fabricated ZnMn_2O_4 product (G_{6_500}).



A₆₋₅₀₀, G₆₋₅₀₀, and U₆₋₅₀₀ ZnMn₂O₄ products, respectively. The obtained results are compatible with the XRD ones. Eventually, the EDS analysis results (not shown here) support that the as-synthesized spinel ZnMn₂O₄ is composed solely of zinc, manganese, and oxygen elements.

3.1.4. Optical properties investigation. We focused in our subsequent investigation on the spinel ZnMn₂O₄ product (G₆₋₅₀₀) fabricated using glycine fuel owing to its small crystallite size and high purity compared to the other products (A₆₋₅₀₀ and U₆₋₅₀₀). The optical band gap energy (E_g) of the spinel ZnMn₂O₄ product (G₆₋₅₀₀) can be determined employing its UV-Vis diffuse reflectance spectrum (Fig. 7(a)) by utilizing the Kubelka–Munk function, $F(R)$, (eqn (5)).^{55,56}

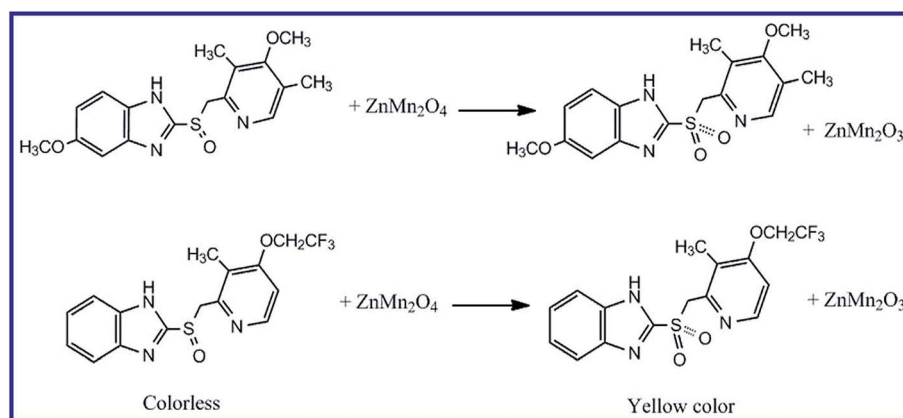
$$F(R) = \frac{(1 - R)^2}{2R} \quad (5)$$

where, R corresponds to the percentage reflectance. Fig. 7(a) reveals an absorption band in the visible region at ca. 656 nm which is consistent with the published results.¹⁰ The spinel ZnMn₂O₄ product (G₆₋₅₀₀) direct band gap energy was calculated by plotting $[F(R)h\nu]^2$ versus $h\nu$ (Fig. 7(b)) where $h\nu$ is the incident photon energy (eV). Then, the band gap energy value (E_g) of the as-fabricated spinel ZnMn₂O₄ product (G₆₋₅₀₀) was determined to be 1.51 eV, utilizing the extrapolation of the

linear section of the plot at $[F(R)h\nu]^2 = 0$. And the calculated band gap energy value is compatible with the reported ones.^{10,14} The estimated band gap energy value indicates the semi-conducting properties of the as-fabricated spinel ZnMn₂O₄ product (G₆₋₅₀₀).^{2,8}

3.2. Colorimetric determination of omeprazole and lansoprazole drugs

It was reported that organic sulfide compounds can be oxidized to sulfoxide (R₂SO) and/or sulfone compounds (R₂SO₂), or it can even be brominated utilizing *N*-bromosuccinimide.^{36,57,58} The oxidation of drugs containing sulfide groups can be also performed using iodate in an acidic medium and by following the liberated iodine spectrophotometrically, the concentration of the drugs can be determined.^{59,60} However, these methods are multistep and tedious procedures. Consequently, we have developed a simple and accurate method for determination of such type of drugs based on their oxidation using semi-conducting nanomaterials. This determination method was based on that when the drug containing sulfide functional group was stirred with the as-prepared ZnMn₂O₄ nanoparticles, the color of the solution was turned into yellow. In order to be able to interpret this result, we repeated these experiments in



Scheme 3 Proposed mechanism for the oxidation of omeprazole and lansoprazole drugs with ZnMn₂O₄ nanoparticles.

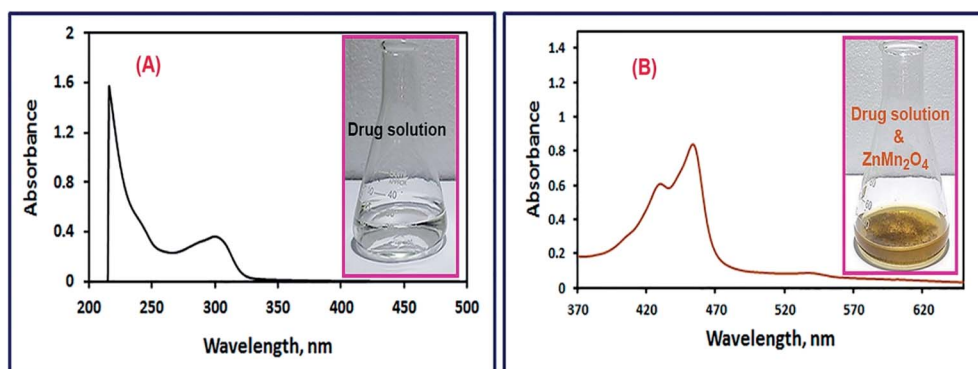


Fig. 8 UV-Vis spectra of omeprazole drug solutions in the absence (A), and presence of ZnMn₂O₄ nanoparticles (B). The images of the drug solutions before and after adding ZnMn₂O₄ nanoparticles are presented in the inset of the corresponding figures.



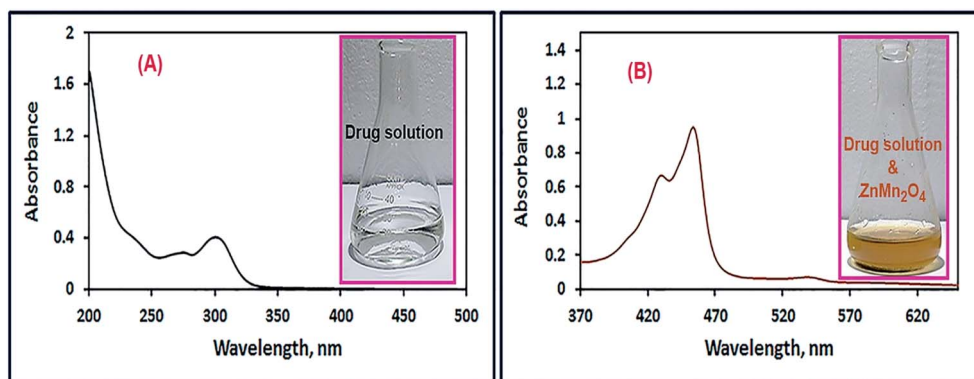


Fig. 9 UV-Vis spectra of lansoprazole drug solutions in the absence (A), and presence of ZnMn_2O_4 nanoparticles (B). The images of the drug solutions before and after adding ZnMn_2O_4 nanoparticles are presented in the inset of the corresponding figures.

the presence of an oxidizing agent such as H_2O_2 and we did not add any amount of ZnMn_2O_4 to the reaction. The same yellow color with similar absorbance was obtained. Based on this, we concluded that the produced yellow color upon addition of ZnMn_2O_4 to the drugs of interest might be owing to the oxidation of the drugs with that semiconducting material (ZnMn_2O_4). Moreover, this oxidation may be due to one or two of the following reasons: the first reason is: under visible light, the electrons absorb light and transfer through excitation from valence band to the conduction band of ZnMn_2O_4 leaving holes in the valence band. These holes of the valence band may abstract electrons from the sulfide groups of the drugs leading to their oxidation. The second reason is: Mn^{3+} of the spinel product (ZnMn_2O_4) may oxidize the drugs, in the presence of H_2O molecules, from R_2SO forms to sulfone derivatives (R_2SO_2), leaving Mn^{2+} . Hence, during the oxidation process of the drugs, some of the spinel ZnMn_2O_4 product may be converted into ZnMn_2O_3 due to its own reduction, as presented in Scheme 3. Hence, we have investigated various factors influencing the determination procedure such as selecting absorption wavelength of the oxidized drugs, reaction time, reaction

temperature, pH of the media, volume of the drug solution, concentration of the drugs, and dose of the nanomaterial.

3.2.1. Selection of the appropriate absorption wavelength.

A solution of the drug of interest (omeprazole or lansoprazole) was stirred with 50 mg of ZnMn_2O_4 nanoparticles, as explained

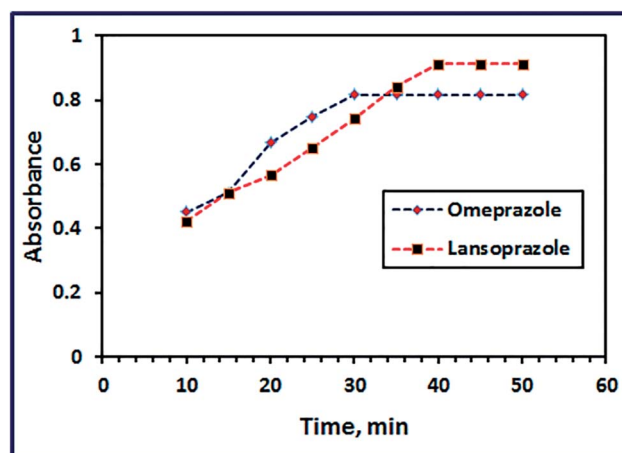


Fig. 11 Effect of reaction time on color development of the products.

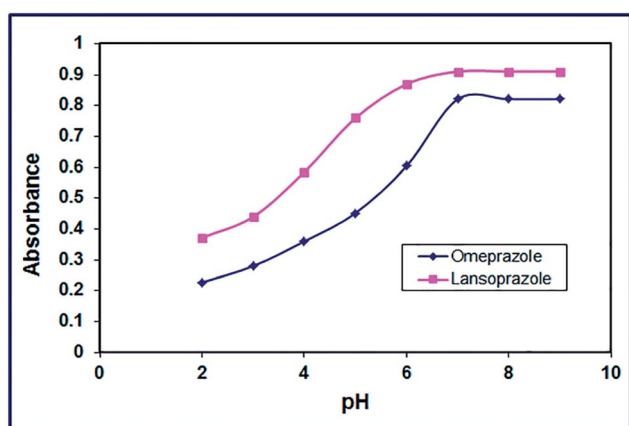


Fig. 10 Effect of pH of the drug solution on color development of the products.

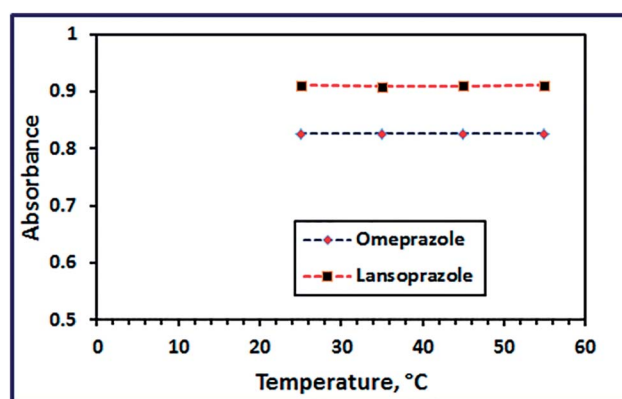


Fig. 12 Effect of reaction temperature on color development of the products.



in the Experimental section, for 1 h. The produced yellow color was scanned using an UV-Vis spectrophotometer in the wavelength range of 200–800 nm to detect the λ_{\max} of the oxidized

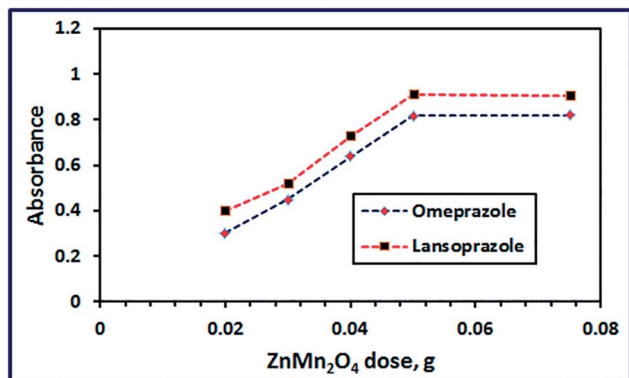


Fig. 13 Effect of ZnMn_2O_4 dose on color development of the products.

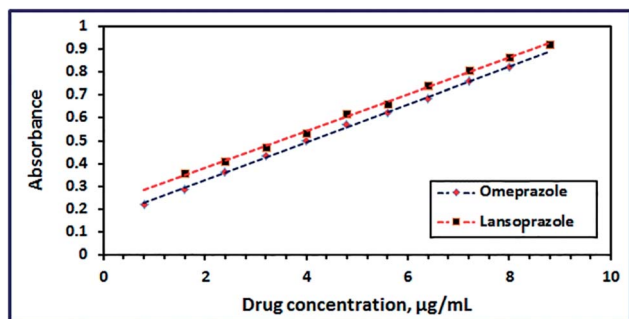


Fig. 14 Calibration plots of omeprazole and lansoprazole drugs using ZnMn_2O_4 nanoparticles.

drugs. Fig. 8 and 9 exhibited spectra of the omeprazole and lansoprazole drugs, respectively, before and after oxidation with ZnMn_2O_4 nanomaterial. The images of the drug solutions before and after adding ZnMn_2O_4 nanoparticles are presented in the inset of the corresponding figures (Fig. 8 and 9). The images clearly show turning the colorless color of the drug solutions into yellow color upon addition of the ZnMn_2O_4 nanoparticles. The spectra of the oxidized drugs revealed an absorption maximum at *ca.* 454 nm.

3.2.2. Effect of pH. The influence of reaction media pH on the color development was examined in the pH range of 2–9 at ambient temperature ($25 \pm 1^\circ\text{C}$). The results exhibited that the intensity of the yellow enhanced gradually with increasing the pH value until it reached its maximum at pH 7 and it remained constant as displayed in Fig. 10. Hence, the optimum pH value for this color development was opted to be 7 in the subsequent study.

3.2.3. Effect of time. Solutions of the drugs (omeprazole or lansoprazole) were blended with 50 mg of ZnMn_2O_4 nanoparticles, as explained in the Experimental section. The optimum reaction time was estimated by following the color development with time elapsed at pH 7 and ambient temperature ($25 \pm 1^\circ\text{C}$). It is clear from Fig. 11 that complete color intensity of the oxidized drugs was obtained after 30 and 40 min for omeprazole and lansoprazole drugs, respectively, and the color remained stable for 24 h.

3.2.4. Effect of temperature. Studying the effect of temperature (in the range of $25\text{--}55^\circ\text{C}$) on the color intensity of the oxidized drugs revealed that temperature has not a remarkable effect, as shown in Fig. 12.

3.2.5. Effect of ZnMn_2O_4 dose. Effect of ZnMn_2O_4 nanostructure dose on the color intensity of the oxidized drugs (at λ_{\max} 454 nm) has been examined at constant drug concentration and at pH 7. The results depicted in Fig. 13 exhibited that

Table 3 Comparison of the performance of the proposed method with those of the published methods for the determination of lansoprazole and omeprazole drugs^a

Item no.	Used reagents	λ_{\max} (nm)	Linear range ($\mu\text{g mL}^{-1}$)	Main remarks	Ref.
1 ^L	Bromate–bromide mixture and iron(II) with thiocyanate	470	0.5–4	It includes multistep procedure, many reagents, and narrower linear range	61
2 ^L	NBS–chloranilic acid utilizing FIA system	530	5–150	It is less sensitive and it includes many reagents	62
3 ^L	Ceric ammonium sulphate with indigo carmine dye	610	0.25–3	Sensitive but it includes multistep procedure, many reagents, and narrower linear range	63
4 ^{L–Om}	Fe(III) with ferricyanide	730 ^L , 720 ^{Om}	0.2–3.6, 0.08–2.8	Sensitive but it includes many used reagents and slow reactions	64
5 ^{Om}	Celestine blue dye and NBS in presence of HCl solution	540	0.4–2.4	Sensitive but it includes many used reagents, multistep reactions, and narrower linear range	65
6 ^{Om}	MBTH reagent and Fe(III)	660	0.1–10	Less sensitive and it includes multistep reactions	65
7 ^{L–Om}	ZnMn_2O_4	454, 454	0.8–8.8 ^L , 0.8–8 ^{Om}	It is a single step reaction, and sensitive. It does not require heating, extraction, or pH adjustment	Present procedure

^a Where, L = lansoprazole, Om = omeprazole, MBTH = 3-methyl-2-benzothiazolinone hydrazine, and NBS = *N*-bromosuccinimide.



the color intensity of the reaction mixture enhanced with increasing the nano-material quantity until it reached maximum at 50 mg. Afterward, increasing the ZnMn_2O_4 nano-sensor dose did not have any significant influence on the color intensity.

3.2.6. Calibration and precision of the ZnMn_2O_4 chemical nanosensor. Typical calibration curves for omeprazole and lansoprazole with the as-prepared ZnMn_2O_4 nanosensor obtained from linear regression analysis of absorbance against concentration of the drugs ($\mu\text{g mL}^{-1}$) were made. Plotting the concentration of each drug *versus* its absorbance revealed that Beer's law limits were 0.80–8.0 and 0.80–8.8 $\mu\text{g mL}^{-1}$ at λ_{max} 454 nm for omeprazole and lansoprazole, respectively, as shown in Fig. 14. Moreover, the corresponding correlation coefficients are 0.999 and 0.998 for omeprazole and lansoprazole drugs, respectively. The Sandell sensitivity was estimated to be 0.012 and 0.0125 $\mu\text{g cm}^{-2}$, for omeprazole and lansoprazole drugs, respectively. Moreover, the limit of detection was found to be 0.264 and 0.093 $\mu\text{g mL}^{-1}$ for omeprazole and lansoprazole, respectively. And the limit of quantification was determined to be 0.88 and 0.98 $\mu\text{g mL}^{-1}$ for omeprazole and lansoprazole, respectively. Therefore, it can be concluded from the above discussion that ZnMn_2O_4 nanostructure can be applied as a chemical nanosensor to detect omeprazole and lansoprazole drugs precisely in ethanolic solutions. On the other hand, some of the published spectrophotometric methods have one or more disadvantages, such as using extraction or heating during determination, low sensitivity, or determination narrow range, as tabulated in Table 3.

4. Conclusions

In this investigation, a novel approach has been adopted to use ZnMn_2O_4 nanostructure as a chemical nanosensor. Different ZnMn_2O_4 nanostructures have been synthesized using different fuels such as L-alanine, glycine, and urea *via* an auto-combustion process. Various factors influencing the combustion process have been explored. Glycine fuel produced the purest spinel ZnMn_2O_4 product on combustion while the other fuels generated multi-phases. The as-prepared spinel ZnMn_2O_4 nanostructure proved its efficiency as a chemical nanosensor for the sensitive determination of omeprazole and lansoprazole drugs. The developed chemical nanosensor showed high sensitivity, low detection limit, and a relatively wide linear range (0.80–8.0 and 0.80–8.8 $\mu\text{g mL}^{-1}$ at λ_{max} 454 nm for omeprazole and lansoprazole, respectively). This study may open a new approach for developing new chemical sensors for colorimetric determination of various drugs.

Conflicts of interest

There are no conflicts to declare.

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