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Significantly improved cyclability of lithium manganese oxide, simultaneously inhibiting electrochemical and thermal decomposition of the electrolyte by the use of an additive†

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Lithium manganese oxide (LiMn₂O₄) is one of the most promising cathodes for lithium ion batteries because of its abundant resources and easy preparation. However, its poor cyclability, especially under elevated temperature, limits its application on a large scale. In this work, it is reported that the cyclability of LiMn₂O₄ can be significantly improved by applying 4-(trifluoromethyl)benzonitrile (4-TB) as an electrolyte additive. Charge/discharge tests indicate that the capacity retention of LiMn₂O₄ after 450 cycles at 1C and 55 °C in a standard electrolyte, 1 M LiPF₆ in EC/EMC/DEC (3 : 5 : 2, in weight), is improved from 19% to 69%. Further electrochemical and physical characterization demonstrates that 4-TB can, on the one hand, be electrochemically oxidized preferentially compared to the standard electrolyte, which generates a protective interphase film on LiMn₂O₄. On the other hand, 4-TB can effectively combine with protonic impurities, which inhibits the thermal decomposition of the electrolyte. This dual-functionality of 4-TB contributes to the significantly improved cyclability of LiMn₂O₄.

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Introduction

Lithium ion batteries are the most used power sources for electric vehicles due to their high energy density and environmental friendliness compared with other secondary batteries.^{1–3} Cobalt-containing oxides such as LiCoO₂ and ferrous phosphates such as LiFePO₄ are the mainstream cathodes of commercial lithium ion batteries. However, cobalt is rare on earth and the preparation of ferrous phosphates is complicated, leading to the high cost of currently commercialized lithium ion batteries. To apply lithium ion batteries more widely, it is necessary to reduce the cost of manufacturing lithium ion batteries. Spinel LiMn₂O₄ is one of the most promising cathodes for lowering the cost of lithium ion batteries, owing to the abundance of manganese and the easy preparation of LiMn₂O₄.^{4,5} However, the poor cyclability of LiMn₂O₄, especially under elevated temperature, limits its application in practice,^{6–8} which is caused by the poor compatibility of LiMn₂O₄ with carbonate-based electrolytes.^{9,10}

The carbonate-based electrolytes might be decomposed electrochemically-oxidatively on charged cathode or thermally under elevated temperature.^{11,12} The electrochemical-oxidation decomposition of electrolyte takes place in the form of complexes of solvents with anion PF₆[−], yielding HF.¹³ On the other hand, the electrolyte also suffers thermal decomposition when it is stored under elevated temperature. HF is also generated during the thermal decomposition due to the reactions of LiPF₆ with trace protonic impurities in the electrolyte such as H₂O.^{14,15} The resulting HF from the electrochemical-oxidative or thermal decompositions causes manganese dissolution of LiMn₂O₄ or even the corrosion of aluminum (current collector), leading to the poor cyclability of LiMn₂O₄.

Several methods have been reported to mitigate the electrochemical-oxidation decomposition of the electrolyte, including partially substituting transition metal elements such as Co and Ni for Mn or non-metal elements such as F for O in LiMn₂O₄,^{16,17} coating LiMn₂O₄ with inert materials,^{18,19} designing LiMn₂O₄ with special crystal plane,²⁰ and using electrolyte additives that can help build protective interphase film on LiMn₂O₄.^{21–23} Comparatively, less attention has been paid to the inhibition of the thermal decomposition of the electrolyte.

Previously, we developed a new additive, 4-(trifluoromethyl) benzonitrile (4-TB), to enhance the cyclability of LiNi_{0.5}Mn_{1.5}O₄ under 4.9 V (ref. 24) and LiCoO₂ under 4.5 V (ref. 25). It was found that 4-TB could combine transition metal ions in cathode and then be oxidized preferentially compared to the solvents.

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This preference helps build a protective interphase film that can suppress the electrochemical-oxidation decomposition of the electrolyte. In this work, we report that 4-TB can not only be electrochemical-oxidized preferentially, which helps build a protective interphase film on LiMn_2O_4 , but also coordinates protonic impurities in electrolyte, which mitigates the thermal stability of the electrolyte. This dual-functionality of electrolyte additive has not been reported before in literature and provides a new approach for improving the cyclability of cathodes of lithium ion batteries, not limited to LiMn_2O_4 .

Experimental and calculation

Sample preparation

4-(Trifluoromethyl)benzotrile (4-TB, >99%) was purchased from Alfa Aesar Technology Co., Ltd. A standard (STD) electrolyte with water content controlled to less than 10 ppm, 1 mol L^{-1} lithium hexafluorophosphate (LiPF_6) in a mixed solvent of ethylene carbonate (EC)/ethyl methyl carbonate (EMC)/diethyl carbonate (DEC) (3 : 5 : 2, in weight) was provided by Guangzhou Tinci Materials Technology Co., Ltd., China. These chemicals were used without further purification. The electrolyte containing 4-TB was prepared in an argon-filled glove box (MBraun, Germany), by adding 4-TB into the STD electrolyte. 0.5 wt% 4-TB, which was found to be the optimal content for the cyclability improvement of $\text{LiNi}_{0.5}\text{Mn}_{1.5}\text{O}_4$ under 4.9 V (ref. 24) and LiCoO_2 under 4.5 V (ref. 25) in our previous reports, was considered in this work.

LiMn_2O_4 working electrode consists of 80 wt% LiMn_2O_4 (Hunan Dahua New Energy Co., Ltd., China), 10 wt% conductive carbon (TIMCAL Ltd., Switzerland), and 10 wt% PVDF (Oflourine Chemical Tech Co., Ltd., China). The mixture was coated on aluminum current collector, dried at 120 °C for 8 h, and punched into disks of 1.13 cm^2 . To ensure the consistency of all the electrodes, the loading of LiMn_2O_4 in the electrodes were controlled at the level of $1.93 \pm 0.01 \text{ mg cm}^{-2}$. Coin cells (2025-type) were assembled with lithium foil as counter electrode and microporous membrane (Celgard 2400) as separator in the glove box mentioned above.

Electrochemical measurements

Charge/discharge performances of $\text{Li}/\text{LiMn}_2\text{O}_4$ cells were tested on a LAND system (CT2001A, Wuhan, China) under 25 °C and 55 °C at 0.5C (1C = 148 mA h g^{-1}) rate for the initial three formation cycles and at 1C rate for the remaining cycles between 3.0 and 4.5 V. Chronoamperometry and electrochemical impedance spectroscopy (EIS) were performed on Autolab (Metrohm, Netherlands) at room temperature. In chronoamperometry, $\text{Li}/\text{LiMn}_2\text{O}_4$ cells were charged to 4.5 V at 0.1C and then kept at this voltage for 13.5 h. EIS was carried out at discharged state with a sinusoidal modulation of 5 mV from 100 kHz to 0.01 Hz. To ensure the accuracy of the experimental results, all the electrochemical measurements were performed with five cells for every sample and the reported results in this work were taken from the cells that represented the mean values of five cells.

Chemical analysis and physical characterization

The thermal stability of electrolyte was evaluated by storing the electrolyte under elevated temperature. The electrolyte was injected to polytetrafluoroethylene (PTFE) bottle, which was packed in aluminum-plastic bag and stored in vacuum oven at 55 °C for 72 h. The hydrofluoric acid (HF) content in electrolytes was analyzed by NaOH titration with bromothymol blue as indicator.

GC-MS was performed using Agilent 7890A and 7895C (Agilent Technologies Inc., America) for analyzing the change in electrolyte composition after storage under elevated temperature and confirming the interaction of additive with protonic impurities. Morphology of LiMn_2O_4 electrodes was observed with SEM (ZEISS ULTRA 55, Germany) and TEM (JEM-2100, JOEL, Japan). Surface composition on LiMn_2O_4 electrodes was determined by XPS (Thermo Scientific, K-Alpha, US). The contents of manganese and aluminum on lithium electrodes of the cycled cell were determined by ICP-AES (Optima 8300, US). Crystal structure of LiMn_2O_4 electrodes was analyzed by XRD (Bruker D8 Advance, Germany). The electrodes taken from the cycled cells were cleaned with DMC to remove the residual electrolyte and dried before physical characterizations.

Theoretical calculation

The highest occupied molecular orbital (HOMO) energy of EC, EMC, DEC and 4-TB was calculated with Gaussian 09 software package. The equilibrium state structures were optimized by B3LYP method at 6-311++G (d) basis set.

Results and discussion

Cycle performances of LiMn_2O_4 under elevated temperature

Fig. 1 presents the cycle performance of LiMn_2O_4 electrode in the electrolytes with and without 0.5 wt% 4-TB at 1C under 55 °C. As shown in Fig. 1a, which presents the representative one of five samples for each electrode (Fig. S1†), LiMn_2O_4 exhibits poor cyclability after cycling in the STD electrolyte: retaining a discharge capacity of only 21 mA h g^{-1} and providing a capacity retention of only 19%. Correspondingly, its coulombic efficiency is low (Fig. 1b) and its discharge profile after deep cycling becomes different completely from its initial one (Fig. 1c). This is the main issue that restricts the utilization of LiMn_2O_4 in practice. The electrochemical-oxidative and thermal decompositions are responsible for this poor cyclability. The electrolyte might be decomposed electrochemically on LiMn_2O_4 during charge process and thermally under elevated temperature. Both of these decompositions yield HF that causes the manganese dissolution from LiMn_2O_4 , leading to the final structural collapse. The continuous discharge capacity decaying up to the 280th cycle for the LiMn_2O_4 cycled in the STD electrolyte can be attributed to the increased polarization resulting from the electrolyte decomposition products,²⁶ while the abrupt discharge capacity decay from the 280th cycle can be ascribed to the structural collapse LiMn_2O_4 .²⁷

When 4-TB is employed, the cyclability of LiMn_2O_4 is improved significantly, as shown in Fig. 1a. The capacity retention is enhanced to 69%. Compared to the LiMn_2O_4 cycled in the electrolyte without additive, the LiMn_2O_4 cycled in the additive-



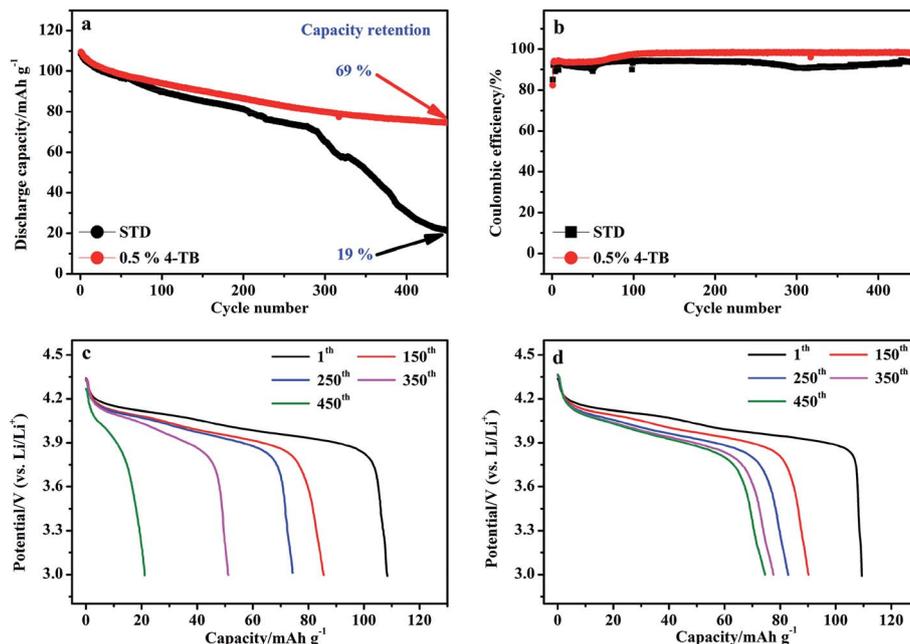


Fig. 1 Cyclability (a), coulombic efficiency (b), and discharge curves at selected cycles in STD (c) and 4-TB-containing (d) electrolytes for LiMn_2O_4 electrodes at 1C under 55 °C. The electrodes were performed with three cycles at 0.5C under room temperature before cycling.

containing electrolyte shows less capacity decay before the 280th cycle. Most importantly, no abrupt discharge capacity decay is observed for the LiMn_2O_4 cycled in the additive-containing electrolyte. Correspondingly, the coulombic efficiency of LiMn_2O_4 cycled in the additive-containing electrolyte remains higher (Fig. 1b) and its discharge profile after deep cycling changes insignificant (Fig. 1d). It is apparent that 4-TB can inhibit the decomposition of electrolyte and the structural collapse of LiMn_2O_4 .

Various electrolyte additives have been proposed for the cyclability improvement of LiMn_2O_4 , which are usually based on the interphase film generated from the additives.^{21–23} Compared with these additives, 4-TB behaves better for the cyclability improvement of LiMn_2O_4 . For example, applying 3 wt% tripropyl borate yields a capacity retention of only 74% for LiMn_2O_4 after 200 cycles at 1C under 55 °C.²⁷ Comparatively, the capacity retention is 80% by using only 0.5 wt% 4-TB for the LiMn_2O_4 cycled under the same rate, same temperature and same cycle number. Since the capacity decay of LiMn_2O_4 under elevated temperature is related to not only the electrochemical-oxidation decomposition but also the thermal decomposition of the electrolyte, the improved cyclability of LiMn_2O_4 is also attributed to the thermal stability improvement of electrolyte by 4-TB besides of its contribution to the inhibition of the electrochemical-oxidation decomposition of the electrolyte. To confirm these contributions, the interphase film formed from 4-TB and the interaction between 4-TB and electrolyte were identified separately.

Interphase film formed from 4-TB

As an effective additive for forming an interphase film that can suppress electrolyte decomposition, it should be oxidized

electrochemically on the cathode surface before STD electrolyte. The preferential electrochemical-oxidation of 4-TB can be indicated by its higher highest occupied molecular orbit (HOMO) energy than those of solvents, as shown in Fig. 2a. To further confirm the preferential electrochemical-oxidation of 4-TB, the initial charge/discharge process of LiMn_2O_4 was analyzed further. Fig. 2b presents the differential capacity (dQ/dV) curve of LiMn_2O_4 in electrolytes with and without 4-TB during the initial cycle of Fig. 1a. Both electrodes show two apparent redox peaks, which are characteristic of spinel LiMn_2O_4 for typical lithium insertion/extraction.⁶ Compared with the electrode in the STD electrolyte, the LiMn_2O_4 in the 4-TB-containing electrolyte shows an additional small oxidation peak before lithium extraction, as shown in Fig. 2b, suggesting that 4-TB can be oxidized on LiMn_2O_4 during charge process. This oxidation peak potential is about 3.85 V, suggesting that 4-TB is far more easily oxidized during charge process than the solvents in the STD electrolyte, which is usually stable at the potential lower than 4.2 V. The preferential electrochemical-oxidation of 4-TB might help build a protective interphase film that inhibits the electrolyte electrochemical-oxidation decomposition and the structural collapse of LiMn_2O_4 .

Fig. 2c presents the chronoamperometric responses of LiMn_2O_4 electrodes, which were charged to 4.5 V at 0.1C and then kept at this voltage in the electrolytes with and without 4-TB. As shown in Fig. 2c, the electrode in the STD electrolyte has a large residue current during constant voltage process. Since the electrode is charged at a low rate current and the lithium ions in LiMn_2O_4 should be extracted completely, the residue current should be attributed to the serious decomposition of the electrolyte. Differently, the LiMn_2O_4 electrode in the



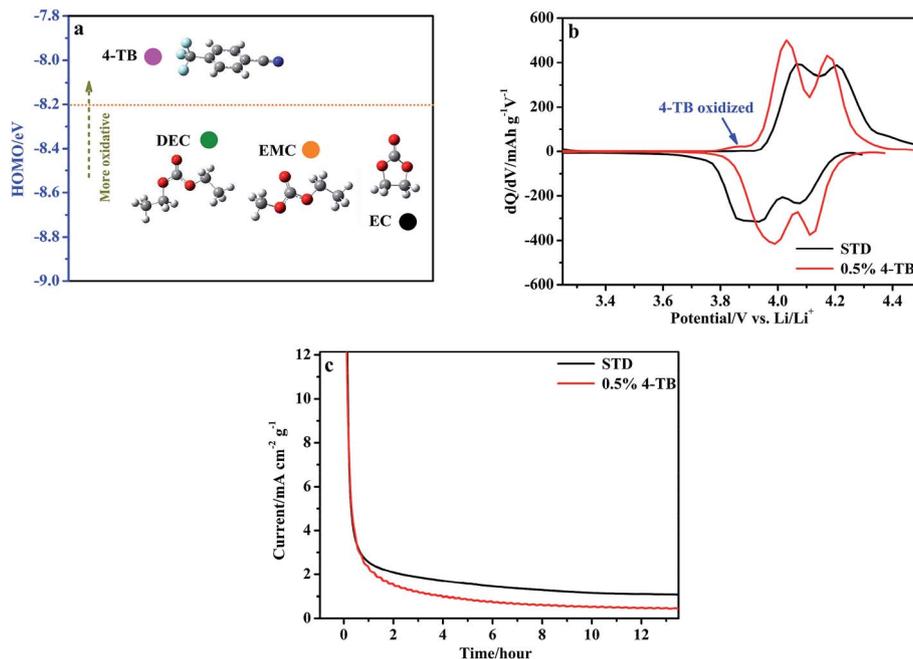


Fig. 2 Calculated HOMO energy (eV) levels of 4-TB, EC, EMC and DEC (a) initial differential charge/discharge capacity curves at 0.5C (b) and chronoamperograms (c) at 4.5 V after charging at 0.1C for LiMn_2O_4 electrodes in STD and 4-TB-containing electrolytes.

additive-containing electrolyte has a far smaller residue current, suggesting that the interphase film resulting from the preferential electrochemical-oxidation of 4-TB can effectively inhibit the electrolyte electrochemical-oxidation decomposition.

It can be noted from the peak positions of Fig. 2b that the electrode in the electrolyte without additive is more polarized for lithium insertion/extraction than that in the additive-containing electrolyte. This difference suggests that the interphase film formed from 4-TB is beneficial for the lithium insertion/extraction, which can be indicated by the smaller interfacial impedance of the electrode cycled in the additive-containing electrolyte. Fig. 3 presents the electrochemical impedance spectra of LiMn_2O_4 electrodes in the electrolytes with and without 4-TB after 1 and 10, 450 cycles. The impedance spectra show a linear relation at low frequencies, associated with Li-ion diffusion, and one or more semicircles at high frequencies, representing interfacial properties including charge transfer and interphase film impedance.^{27,28} It can be found, by comparing the semicircles at high frequencies between two electrodes cycled initially (at the first and 10th cycles), that the electrode cycled in the additive-containing electrolyte has smaller interfacial impedance than that in the STD electrolyte, indicative of the less polarization for lithium insertion/extraction on LiMn_2O_4 with the interphase film formed by 4-TB. The interfacial impedance of the cells with STD electrolyte (Fig. 3a) increases significantly after cycling, showing the interfacial instability of electrode/electrolyte. This instability results from the products of electrolyte decomposition on LiMn_2O_4 and the manganese dissolution that causes the structural collapse of LiMn_2O_4 . Differently, the impedance spectra of the cells with 4-TB show less change from initial cycle

to the last cycle, as shown in Fig. 3b. Obviously, the application of 4-TB can maintain the interfacial stability of LiMn_2O_4 /electrolyte during cycling under elevated temperature. The interphase film originated from the preferential electrochemical-

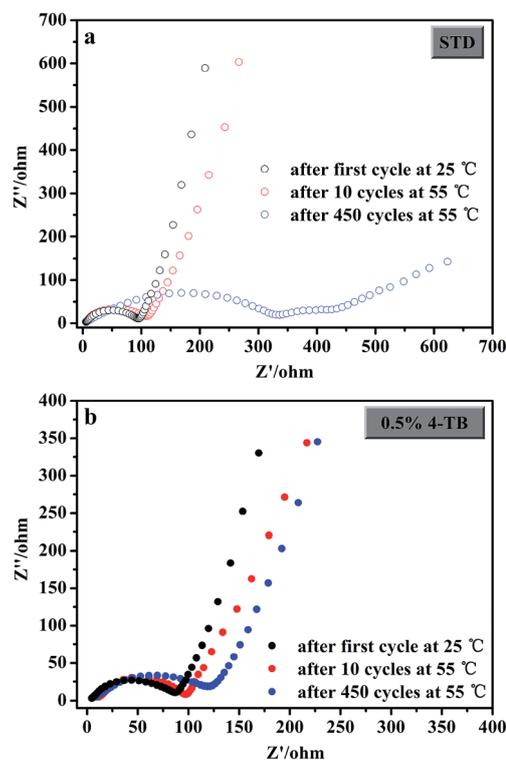


Fig. 3 Electrochemical impedance spectra of LiMn_2O_4 electrodes in the STD (a) and 4-TB-containing (b) electrolytes after cycling.



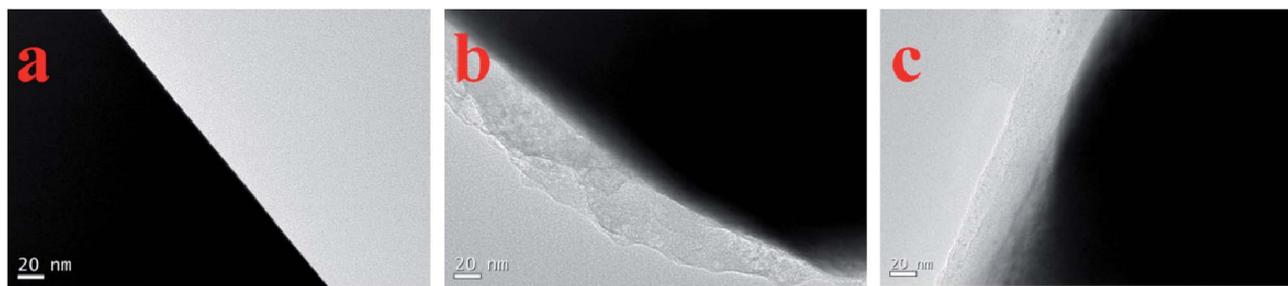


Fig. 4 TEM images of fresh LiMn_2O_4 electrode (a) and the cycled ones in STD (b) and 4-TB-containing (c) electrolytes.

oxidation of 4-TB not only suppresses the electrolyte electrochemical-decomposition but also prevents LiMn_2O_4 from structural collapse. The interphase film formed from 4-TB can be confirmed by TEM. Fig. 4 presents the TEM images of the LiMn_2O_4 particles from the cycled electrodes in the electrolytes with and without additive, with a comparison of fresh one.

Smooth and clean surface can be observed on the fresh particle (Fig. 4a). After cycling in the electrolyte without additive, the particle is covered with thick and inhomogeneous deposit (Fig. 4b), indicative of the electrolyte decomposition during cycling. In the electrolyte with additive, however, a thin and

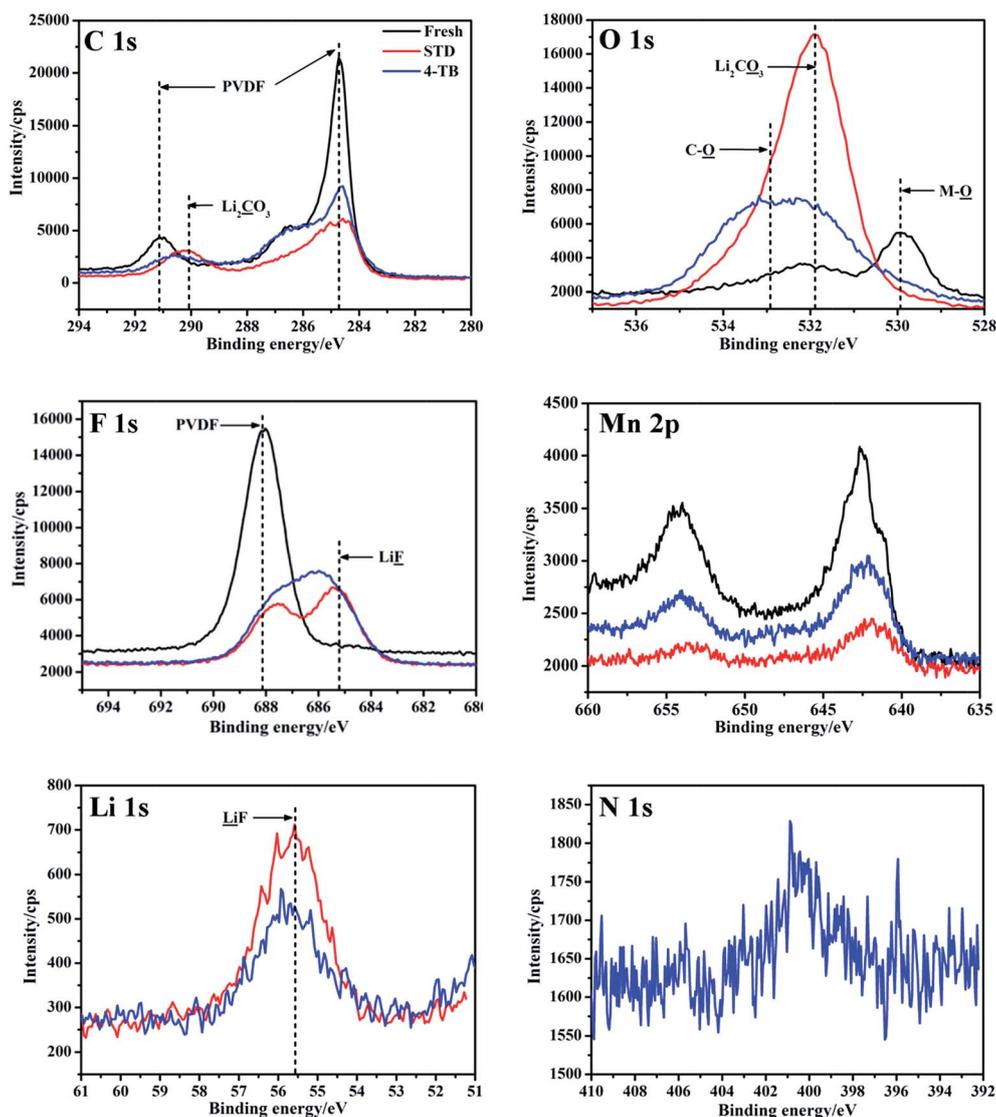


Fig. 5 XPS profiles of the cycled LiMn_2O_4 electrodes in STD and 4-TB-containing electrolytes at 55°C , with a comparison of the fresh one.



uniform interphase film can be identified on the particle, as shown in Fig. 4c.

Fig. 5 presents XPS profiles of the LiMn_2O_4 electrodes in the electrolytes with and without 4-TB after 450 cycles at 1C under 55 °C, with a comparison of the fresh one. The cycled electrodes show their XPS profiles differently from the fresh one. The main compositions of the electrode including PVDF as binder, Mn and O in LiMn_2O_4 , can be obviously identified on the fresh electrode: the significant peaks at 285 and 291 eV in C 1s and at 688 eV in F 1s correspond to PVDF; while those at 530 eV in O 1s and at 643 and 653 eV in Mn 2p correspond to metal-oxide bonds.^{29,30} After cycling, however, the intensities of these peaks are reduced in the STD and additive-containing electrolytes, suggesting that the cycled electrodes have been covered by the electrolyte decomposition products. Comparatively, the reduced magnitude of these peaks is more significant for the electrode cycled in the electrolyte without additive than that cycled in the electrolyte with additive, suggesting that there are

more electrolyte decomposition products existing on the electrode cycled in the STD electrolyte. Obviously, electrolyte decomposes continuously on the LiMn_2O_4 electrode in the electrolyte without additive during cycling process, resulting in the accumulation of the electrolyte decomposition products. In the electrolyte with additive, however, an interphase film is generated from the preferential electrochemical-oxidation of 4-TB and the electrolyte electrochemical-oxidation decomposition is inhibited, resulting in the stronger peak intensities of PVDF and metal-oxide bonds compared with those in the electrolyte without additive.

In fact, the peaks at 290 eV in C 1s and at 532 eV in O 1s, both of which correspond to Li_2CO_3 , and the peak at 56 eV in Li 1s, which corresponds to LiF,³¹ can be observed more clearly for the electrode cycled in the electrolyte without additive than that in the electrolyte with additive. Since Li_2CO_3 and LiF are the main products of electrolyte decomposition,³² the weaker peaks for Li_2CO_3 and LiF on the electrode cycled in 4-TB-containing

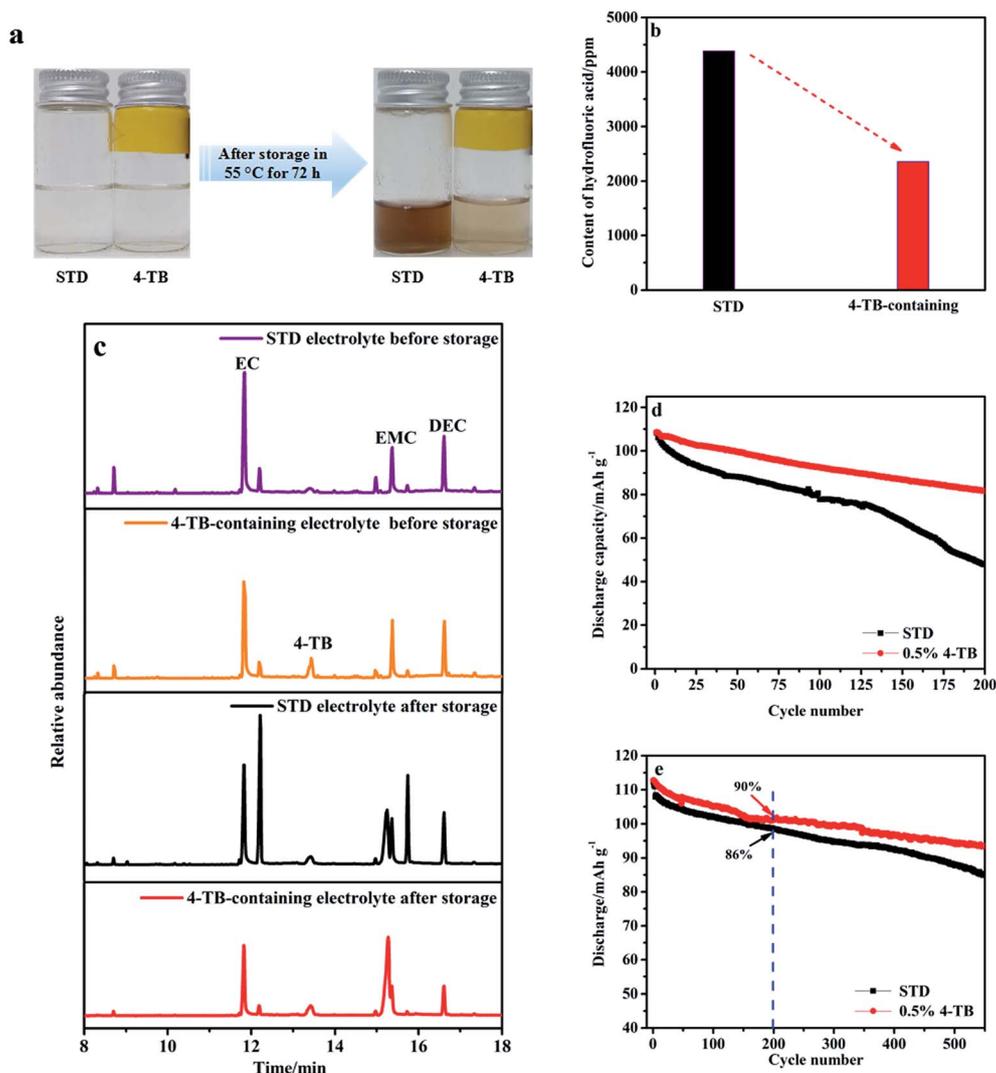


Fig. 6 Changes in color (a), HF content (b) and GC-MS spectrum (c) of the electrolytes after storage at 55 °C for 72 h. Cyclability of LiMn_2O_4 electrodes at 0.5C for the first three cycles and at 1C for the subsequent cycles under 25 °C, using the electrolytes after (d) and before (e) storage.



electrolyte confirm that the electrolyte decomposition can be inhibited by applying 4-TB. It should be noted that an unidentified peak near LiF in F 1s spectrum appears for the electrode cycled in the electrolyte with additive, suggesting that F in 4-TB has been incorporated into the interphase film. The element N existing in 4-TB can only be detected in N 1s for the electrode cycled in the electrolyte with additive, confirming that the products from the preferential electrochemical-oxidation of 4-TB has been incorporated into the protective interphase film that inhibits the electrolyte electrochemical-oxidation decomposition and prevents LiMn_2O_4 from structural collapse.

Improved thermal stability of electrolyte by 4-TB

To confirm the effect of 4-TB on the thermal stability of electrolyte, the STD and 0.5 wt% 4-TB-containing electrolytes were stored at 55 °C for 72 h. The photos of the electrolytes before and after storage tests are shown in Fig. 6a. It can be seen from Fig. 6a that the color of the STD electrolyte turns to far brown than that of the electrolyte containing 4-TB, confirming that the electrolyte without additive suffers serious thermal decomposition that can be inhibited to a great extent by applying 4-TB. The composition change of the electrolytes after storage was monitored by acid–base titration and GC-MS analysis, and the effect of the electrolytes after storage on the cyclability of LiMn_2O_4 was evaluated in coin cell under room temperature. Fig. 6b presents the increased HF concentrations of the electrolytes after storage. Correspondingly to the color change observed in Fig. 6a, there is a significantly increased concentration of HF (4386 ppm) in the STD electrolyte after storage, about two times that in the 4-TB-containing electrolyte after storage (2356 ppm). HF will erode LiMn_2O_4 , leading to the capacity decay of LiMn_2O_4 .

Fig. 6c presents the GC-MS spectra of the electrolytes with and without 4-TB before and after storage at 55 °C for 72 h. Both electrolytes have similar spectra: three main peaks at 11.8, 15.3 and 16.6 min, representing EC, EMC and DEC, respectively, and other minor peaks at 8.7, 12.2, 13.4, 14.9 and 15.7 min, which should be ascribed to the coordination of solvents with ions in LiPF_6 or trace impurities in the electrolyte. Differently, the additional peak at 13.4 min is stronger in the electrolyte containing 4-TB. This peak might come from the contribution of 4-TB, which can be confirmed later. In the STD electrolyte after storage, the peaks representing EC and EMC become weaker, while the peaks at 12.2, 14.9 and 15.7 min become stronger, indicating that the STD electrolyte is unstable thermally. Under elevated temperature, LiPF_6 will decompose into PF_5 , which reacts rapidly with protonic impurities such as ROH or H_2O forming HF and POF_3 and initiates an autocatalytic decomposition of the electrolyte.¹¹ The resulting HF has been confirmed as discussed above. Interestingly, the GC-MS spectrum of the 4-TB-containing electrolyte after storage has less changed compared to its fresh one, confirming that the thermal stability of the STD electrolyte has been improved by applying 4-TB.

The resulting HF from the thermal decomposition of electrolyte might deteriorate the cyclability of LiMn_2O_4 . Fig. 6d presents the cycle performance of LiMn_2O_4 electrode in the

electrolyte with and without additive after storage at 55 °C for 72 h, the data in which are the representative ones of five samples for each electrode (Fig. S2†). The cycling tests were performed in Li/ LiMn_2O_4 coin cells under 25 °C at 0.5C for the initial three formation cycles and at 1C for the remaining cycles. As shown in Fig. 6d, LiMn_2O_4 experiences fast capacity decay in the STD electrolyte after storage, showing a capacity retention of only 44% after 200 cycles compared to the 76% of that in the electrolyte containing 4-TB. Comparatively, the cycle performance of LiMn_2O_4 in the electrolytes before storage (Fig. 6e), which was obtained under the same cycling conditions as Fig. 6d, shows less difference between the STD electrolyte and the 4-TB-containing one. The capacity retention of LiMn_2O_4 is 90% and 86% after 200 cycles, and 83% and 75% after 550 cycles, for the electrolytes with and without 4-TB, respectively. These analyses indicate that the electrolyte after thermal decomposition deteriorates the cyclability of LiMn_2O_4 . It should be noted that there still exists a difference in cyclability of LiMn_2O_4 between the electrolytes with and without 4-TB under room temperature (Fig. 6e), which emphasizes the contribution of interphase film formed from 4-TB to the cyclability improvement of LiMn_2O_4 . Therefore, the improved cyclability of LiMn_2O_4 comes from dual contributions of 4-TB: one is forming a protective interphase film from the preferential electrochemical-oxidation of 4-TB; the other is stabilizing the electrolyte by 4-TB. The latter contribution has never been reported before, and therefore, 4-TB provides a more effective solution to the cyclability issue of LiMn_2O_4 under elevated temperature than those additives that have been reported before.

Since the thermal instability of the electrolyte results from trace protonic impurities, the improved thermal stability of the electrolyte should be related to the interaction of 4-TB with these protonic impurities. To confirm this interaction, H_2O was added into DEC with and without 4-TB and the resulting mixtures were analyzed with GC-MS. The obtained results are presented in Fig. 7. Pure DEC shows its characteristic peak at 16.6 min, as shown in Fig. 7a. When 1 wt% deionized water is added into pure DEC, no significant change in GC-MS spectrum is observed (Fig. 6b), suggesting that there is no interaction between DEC and H_2O . When 0.5 wt% 4-TB is added into pure DEC, the characteristic peak of 4-TB appears at 13.4 min, as shown in Fig. 7c. No other peak appears in Fig. 7c, suggesting that there is no interaction between 4-TB and DEC. When 0.5 wt% 4-TB is added into the DEC containing 1 wt% deionized water, however, there appears a new peak at 14.3 min except for the characteristic peaks of 4-TB and DEC, as shown in Fig. 7d. Apparently, there exists an interaction between 4-TB and H_2O , which are related to –CN group. The –CN combines proton forming $-\text{CNH}^+$, which then combines H_2O forming $-\text{C}(\text{NH})\text{OH}_2^+$ that is finally transformed into $-\text{C}(\text{NH}_2)\text{O}$, as shown in Fig. 7e.¹⁴ The peak at 14.3 min in Fig. 7d corresponds to the coordinate of 4-TB with H_2O .

With the dual functionalities of 4-TB, the dissolution of manganese from LiMn_2O_4 and aluminum from current collector will be inhibited to a great extent. To confirm these contributions, the LiMn_2O_4 and lithium electrodes, taken from



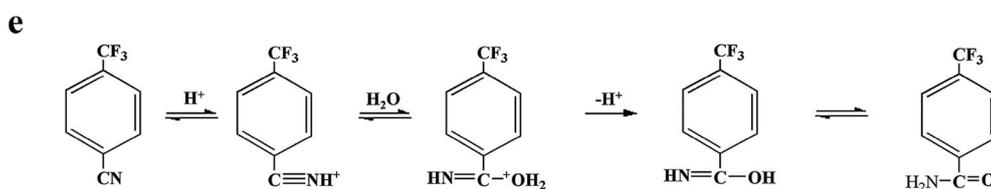
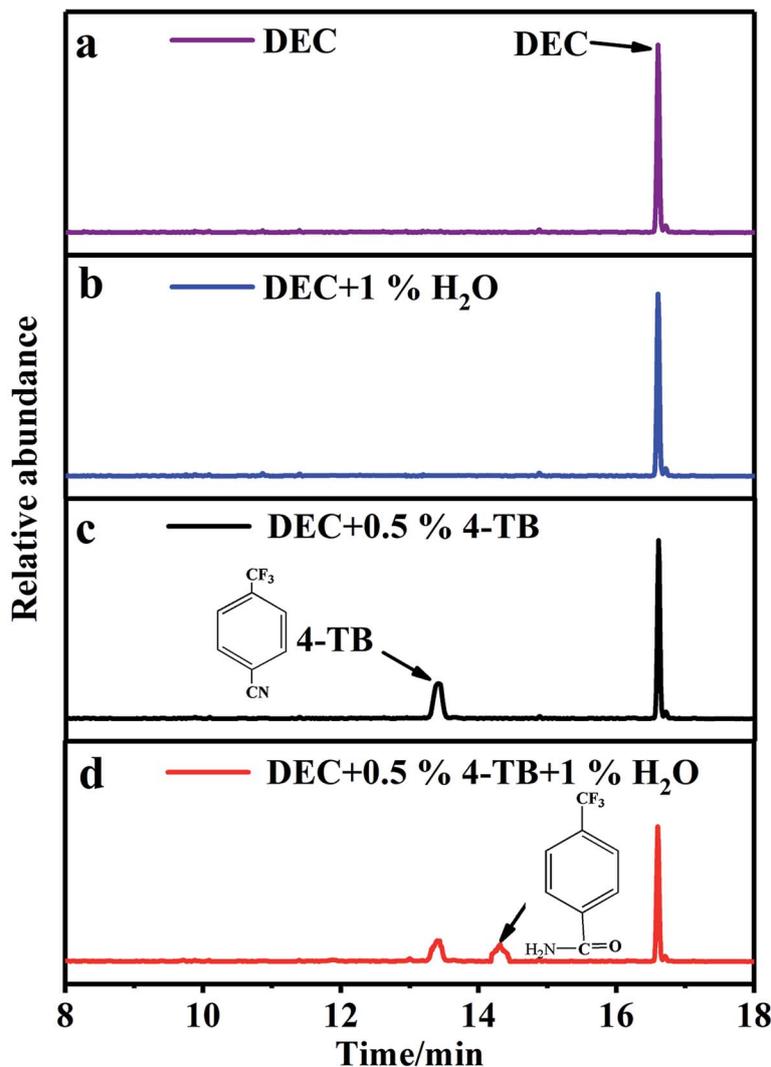


Fig. 7 GC-MS spectra of DEC containing H₂O and/or 4-TB (a–d) and illustration on the coordination of 4-TB with H⁺ and H₂O (e).

the cells after cycling in Fig. 1a, were performed with morphological observations and element analyses. The obtained results are presented in Fig. 8. LiMn₂O₄ particles with spinel morphology and clean surface can be observed in the fresh electrode, as shown in Fig. 8a. After deep cycling in the electrolyte without additive under elevated temperature, however, LiMn₂O₄ particles have been covered by thick deposits (Fig. 8b), which mainly result from the electrochemical-oxidation decomposition of the electrolyte. During charge process, the solvents together with PF₆[−] in the electrolyte will be electrochemical-oxidatively decomposed, forming deposits

such as carbonates and LiF, gases such as carbon oxide and acidic species such as HF.^{11,13} The deposits accumulate on LiMn₂O₄ as the cycling proceeds, accounting for the thick deposits observed in Fig. 8b. The resulting HF together with that from thermal decomposition of the electrolyte will cause the subsequent dissolution of manganese from LiMn₂O₄ and final structural collapse of LiMn₂O₄, as indicated by the arrow in Fig. 8b. Differently, the LiMn₂O₄ particle on the electrode cycled in the electrolyte containing 4-TB maintains the fresh morphology in Fig. 8a, as shown in Fig. 8c, confirming that a protective interphase film has been generated on LiMn₂O₄



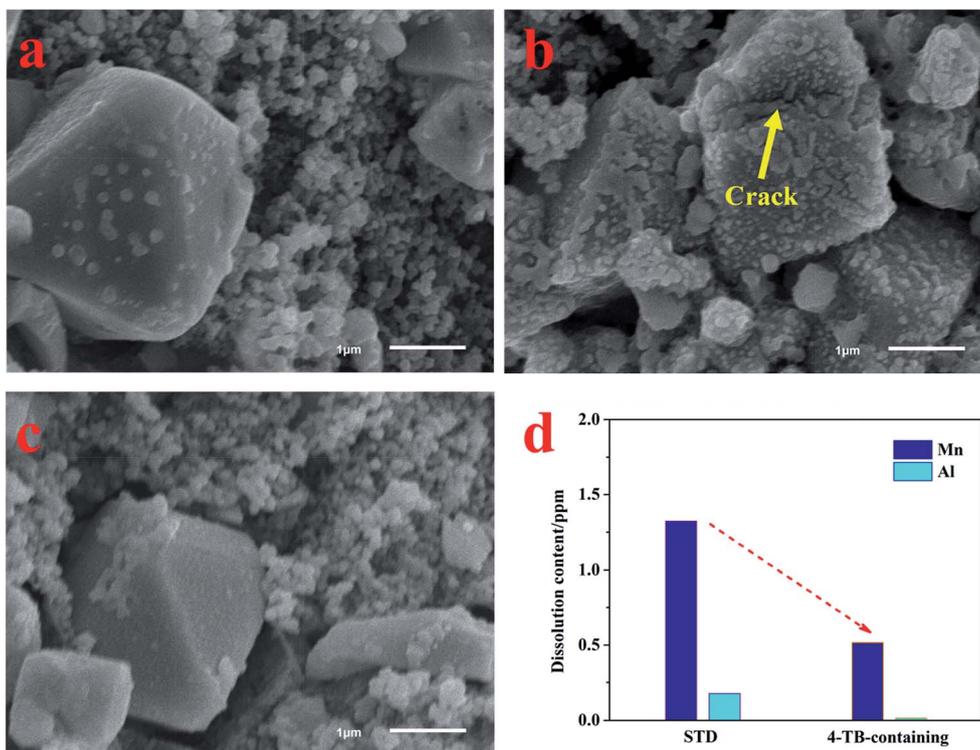


Fig. 8 SEM images of fresh LiMn_2O_4 electrode (a) and the cycled ones in STD (b) and 4-TB-containing (c) electrolytes. Contents of Mn and Al deposited on the cycled lithium electrodes (d). The cycled LiMn_2O_4 and lithium electrodes are taken from the cells after cycling test of Fig. 1a.

from the preferential oxidation of 4-TB during initial cycling, which inhibits the electrolyte oxidation decomposition and the manganese dissolution. Fig. 8d presents the contents of manganese and aluminum deposited on lithium electrode of the cycled cells. The lithium electrodes were dissolved into 25 mL dilute nitric acid solution, which was then performed with ICP-AES analyses. As shown in Fig. 8d, besides manganese, aluminum can also be detected on the lithium electrode, suggesting that the HF also causes the corrosion of the current collector in LiMn_2O_4 electrode. The concentration of Mn and Al

is 1.3 mg L^{-1} and 0.2 mg L^{-1} for the electrolyte without 4-TB but only 0.5 mg L^{-1} and 0.02 mg L^{-1} for the electrolyte with 4-TB, respectively. It is apparent that serious dissolutions of manganese and aluminum happen from LiMn_2O_4 electrode in the cycled cell using the electrolyte without additive, which can be inhibited significantly by applying 4-TB.

Fig. 9 shows XRD patterns of the cycled LiMn_2O_4 electrodes in STD and 4-TB-containing electrolytes at 55°C , with a comparison of fresh one. The diffractions of fresh electrode are characteristic of spinel LiMn_2O_4 structure together with the current collector aluminum. After 450 cycles at 55°C in the STD electrolyte, all the diffraction peaks for LiMn_2O_4 almost disappear, confirming that LiMn_2O_4 suffers structural deterioration. By contrast, the LiMn_2O_4 electrode cycled in 4-TB-containing electrolyte maintains the characteristic diffractions of spinel structure, confirming that the interphase film formed from 4-TB yields an effective protection for the structural integrity of LiMn_2O_4 .

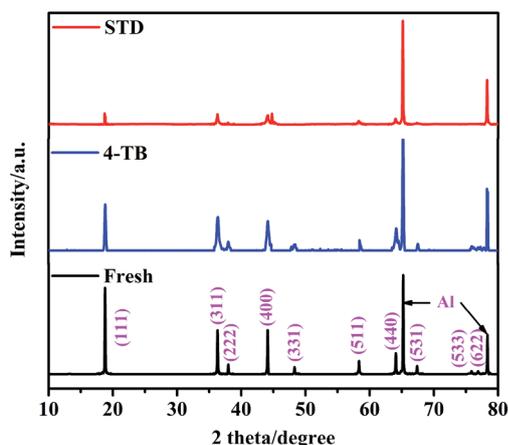


Fig. 9 XRD patterns of the cycled LiMn_2O_4 electrodes in STD and 4-TB-containing electrolytes at 55°C , with a comparison of fresh one.

Conclusions

The cyclability of LiMn_2O_4 electrode under elevated temperature can be significantly improved by applying 4-TB as an electrolyte additive. This improvement is attributed to the dual functionalities of 4-TB. On the one hand, 4-TB can be oxidized electrochemically before the electrolyte oxidation decomposition, forming a protective interphase film on LiMn_2O_4 . On the other hand, 4-TB can coordinate with trace impurities such as



small alcohol and water that are detrimental to the thermal stability of electrolyte. Therefore, the application of 4-TB can inhibit the formation of HF to a great extent by simultaneously inhibiting the electrochemical oxidation and thermal decompositions of electrolyte, and thus reduce the manganese dissolution from LiMn_2O_4 and aluminum from current collector, leading to the significantly improved cyclability of LiMn_2O_4 .

Conflicts of interest

The authors declare that there is no conflict of interest.

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Notes and references

- 1 Y.-M. Song, J.-G. Han, S. Park, K. T. Lee and N.-S. Choi, *J. Mater. Chem. A*, 2014, **2**, 9506–9513.
- 2 A. Väyrynen and J. Salminen, *J. Chem. Thermodyn.*, 2012, **46**, 80–85.
- 3 G. H. Yang, J. L. Shi, C. Shen, S. W. Wang, L. Xia, H. S. Hu, H. Luo, Y. G. Xia and Z. P. Liu, *RSC Adv.*, 2017, **7**, 26052–26059.
- 4 J. B. Goodenough, *Solid State Ionics*, 1994, **69**, 184–198.
- 5 I. Scivetti and G. Teobaldi, *J. Phys. Chem. C*, 2015, **119**, 21358–21368.
- 6 L. Yang, M. Takahashi and B. Wang, *Electrochim. Acta*, 2006, **51**, 3228–3234.
- 7 N.-S. Choi, J.-G. Han and S. Y. Ha, *RSC Adv.*, 2015, **5**, 2732–2748.
- 8 D. C. Tang, Y. Sun, Z. Z. Yang, L. B. Ben, L. Gu and X. J. Huang, *Chem. Mater.*, 2014, **26**, 3535–3543.
- 9 M. S. Whittingham, *Chem. Rev.*, 2004, **104**, 4271–4302.
- 10 J. C. Hunter, *J. Solid State Chem.*, 1981, **39**, 142–147.
- 11 X. Sun, H. S. Lee, X. Q. Yang and J. Mcbreen, *Electrochim. Solid-State Lett.*, 2002, **5**, A248–A251.
- 12 X. L. Liao, P. Y. Sun, M. Q. Xu, L. D. Xing, Y. H. Liao, L. P. Zhang, L. Yu, W. Z. Fan and W. S. Li, *Appl. Energy*, 2016, **175**, 505–511.
- 13 L. D. Xing, W. S. Li, C. Y. Wang, F. L. Gu, M. Q. Xu, C. L. Tan and J. Yi, *J. Phys. Chem. B*, 2009, **113**, 16596–16602.
- 14 R. J. Chen, F. Liu, Y. Chen, Y. S. Ye, F. Wu and L. Li, *J. Power Sources*, 2016, **306**, 70–77.
- 15 G. J. Xu, Z. H. Liu, C. J. Zhang, G. L. Cui and L. Q. Chen, *J. Mater. Chem. A*, 2015, **3**, 4092–4123.
- 16 C. Zhan, J. Lu, A. J. Kropf, T. P. Wu, A. Jansen, Y.-K. Sun, X. P. Qiu and K. Amine, *Nat. Commun.*, 2013, **4**, 1–8.
- 17 C. L. Tan, H. J. Zhou, X. H. Hou, D. S. Lu, M. Q. Xu and Q. M. Huang, *J. Power Sources*, 2008, **184**, 408–413.
- 18 L. Sunhhye and C. Jaephil, *Chem. Commun.*, 2008, **37**, 4472–4474.
- 19 W. W. Sun, H. Q. Liu, Y. M. Liu, G. X. Bai, W. Liu, S. S. Guo and X. Z. Zhao, *Nanoscale*, 2015, **7**, 13173–13180.
- 20 H. B. Lin, J. N. Hu, H. B. Rong, Y. M. Zhang, S. W. Mai, L. D. Xing, M. Q. Xu, X. P. Li and W. S. Li, *J. Mater. Chem. A*, 2014, **2**, 9272–9279.
- 21 B. Li, Y. Q. Wang, H. B. Rong, Y. T. Wang, J. S. Liu, L. D. Xing, M. Q. Xu and W. S. Li, *J. Mater. Chem. A*, 2013, **1**, 12954–12961.
- 22 J. C. Burns, A. Kassam, N. N. Sinha, L. E. Downie, L. Solnickova, B. M. Way and J. R. Dahn, *J. Electrochem. Soc.*, 2013, **160**, A1451–A1456.
- 23 Y. M. Zhu, X. Y. Luo, M. Q. Xu, L. P. Zhang, L. Yu, W. Z. Fan and W. S. Li, *J. Power Sources*, 2016, **317**, 65–73.
- 24 W. N. Huang, L. D. Xing, Y. T. Wang, M. Q. Xu, W. S. Li, F. C. Xie and S. G. Xia, *J. Power Sources*, 2014, **267**, 560–565.
- 25 X. S. Wang, X. L. Liao, W. N. Huang, L. D. Xing, Y. H. Liao, Q. M. Huang, M. Q. Xu and W. S. Li, *Electrochim. Acta*, 2015, **184**, 94–101.
- 26 W.-K. Shin, J.-H. Yoo and D.-W. Kim, *J. Power Sources*, 2012, **279**, 737–744.
- 27 X. R. Yang, J. H. Li, L. D. Xing, Y. H. Liao, M. Q. Xu, Q. M. Huang and W. S. Li, *Electrochim. Acta*, 2017, **227**, 24–32.
- 28 F. Wu, M. Wang, Y. Su, S. Chen and B. Xu, *J. Power Sources*, 2009, **191**, 628–632.
- 29 J. Q. Liu, Q. C. Zhuang, Y. L. Shi, X. D. Yan, X. Zhao and X. B. Chen, *RSC Adv.*, 2016, **6**, 42885–42891.
- 30 Y. Kusachi, Z. C. Zhang, J. Dong and K. Amine, *J. Phys. Chem. C*, 2011, **115**, 24013–24020.
- 31 S. W. Mai, M. Q. Xu, X. L. Liao, L. D. Xing and W. S. Li, *J. Power Sources*, 2015, **273**, 816–822.
- 32 M. Q. Xu, L. Zhou, Y. N. Dong, Y. J. Chen, J. Demeaux, A. D. MacIntosh, A. Garsuch and B. L. Lucht, *Energy Environ. Sci.*, 2016, **9**, 1308–1319.

