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Synthesis of an oligomer ruthenium complex and its catalysis in the oxidation of alcohols†

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A ligand with both a terpyridine and a pyridine-2,6-dicarboxylate group (abbreviated as terpy-pydic) was designed and synthesized. This ligand reacted with $[\text{Ru}(p\text{-cymene})\text{Cl}_2]_2$ to afford a novel oligomer ruthenium complex named as oligomer-Ru(terpy)(pydic) which was characterized thoroughly. Under the catalysis of this oligomer ruthenium complex, different sorts of secondary alcohols were oxidized to the corresponding ketones by the oxidant *tert*-butyl hydroperoxide. Besides, this catalyst can be readily recovered and recycled several times without a large loss of its efficiency.

Introduction

Oxidation of alcohols to their corresponding carbonyl compounds is a fundamental transformation in the synthesis of various valuable chemicals such as fragrances, pharmaceuticals, dyes and food additives.^{1–4} Traditionally, this transformation is realized with a stoichiometric or even an over-stoichiometric amount of MnO_2 ,⁵ chromium(vi) oxide,^{1,6–8} and hypervalent iodine compounds^{9–13} as oxidants, which will bring about serious environmental problems. To overcome these problems, many efforts have been made to explore transition-metal-catalyzed processes using environmentally benign oxidants, especially molecular oxygen as the oxidant.^{14–27}

Due to the existence of redox ion pairs $\text{Ru}^{n+}/\text{Ru}^{(n+2)}$, ruthenium complexes have emerged as efficient catalysts for the oxidation of various substrates.^{28–42} Among all of them, some *N,N,N*-tridentate ligand ruthenium complexes exhibited good performances in the oxidation of alcohols in view of their high efficiency and selectivity.^{24–26,36} For instances, both Ru(terpyridine)(pydic) and its analogue Ru(bbp)(pydic) gave good to excellent yields of ketones in the oxidation of various secondary alcohols with hydrogen peroxide as oxidant.^{24,25} Another similar ruthenium complex Ru(pymieb)(pydic) reported by us also exhibited high efficiency in the oxidation of secondary alcohols to ketones with TBHP as oxidant under mild conditions.²⁶ Despite of their excellent performances in the reaction, it is difficult to separate the expensive homogeneous catalysts with products, which leads to impossible recycling in application. Heterogenization of the ruthenium complexes has been expected to resolve this problem. To this end, some ruthenium complexes were supported on organic or inorganic supports and the immobilized complexes

exhibited good recyclability in the oxidation of various alcohols to the corresponding ketones.^{27,40–43} Although promising results were obtained, the immobilization processes were tedious in some cases, and the immobilization of the complex generally led to decrease of its activity due to mass transfer resistance of the reactants in the matrix of the support. Very recently, we used a polymer with a tridentate 2,6-bis(benzimidazol-2-yl)pyridine (bbp) unit in combination with disodium pyridine-2,6-dicarboxylate (pydic) coordinating with RuCl_3 to form a polymer-ruthenium complex poly-Ru(bbp)(pydic).⁴⁴ This polymer-ruthenium complex showed high activity in the oxidation of various secondary alcohols to their corresponding ketones in high yields with TBHP as oxidant, and could be recycled for several times. Inspired by the catalytic results and center structures of the above catalysts, we designed and synthesized a ligand with both terpyridine and pydic moieties. This ligand coordinated with Ru^{2+} to afford an organo-ruthenium oligomer which showed good performances in the oxidation of alcohols with TBHP as oxidant in view of both activity and recyclability.

Experimental

Materials

Dimethyl 4-hydroxypyridine-2,6-dicarboxylate (**1**) was synthesized as reported method.^{45–47} 4'-(4-Bromomethylphenyl)-[2,2':6',2'']terpyridine (**2**) prepared as described in literature.^{48,49} *tert*-Butyl hydroperoxide (TBHP, 70% aqueous solution) was supplied by Energy Chemical. Dichloro(*p*-cymene)ruthenium(II) dimer $[\text{Ru}(p\text{-cymene})\text{Cl}_2]_2$ was provided by J&K Scientific Ltd. All the substrate alcohols were purchased from Alfa Aesar China Co. Ltd. All reagents were used as received without further purification.

Characterization

¹H NMR and ¹³C NMR spectra were recorded on a Bruker AC-P 400 type spectrometer (400 MHz for ¹H and 101 MHz for ¹³C).

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Chemical shifts (δ) were measured in ppm relative to TMS $\delta = 0$ for ^1H , or to CDCl_3 $\delta = 77.2$ or DMSO $\delta = 39.6$ for ^{13}C as internal standard. Mass spectra were measured using miorOTOF-QII mass spectrometer. Elemental analyses were carried out using Perkin-Elmer 2400 CHN analyzer. FT-IR spectra were recorded with a Bruker Vector 22 spectrophotometer with KBr pellets. Raman spectra were measured at room temperature using an inVia Reflex visible Raman spectrometer in backscattering geometry equipped with excitation argon ion lasers working at 532 nm. UV-vis DRS were recorded with a Varian Cary 300 UV-visible spectrophotometer. Powder X-ray diffraction (XRD) patterns were performed on a Bruker AXS GMBH D8 Discover X-ray diffractometer using with a Cu $K\alpha$ radiation at 40 kV and 40 mA ($2\theta = 10\text{--}90^\circ$). X-ray photoelectron spectroscopy (XPS) data were obtained on a PHI 1600 spectroscopy using Al $K\alpha$ radiation, using the carbon C 1s peak at 284.6 eV as a reference for the charge correction. SEM images and EDX spectra were taken with a Nova Nano SEM450 instrument and an Octane Pro Det. Transmission electron microscopy studies were carried out using a JEM-2100F Transmission Electron Microscope (TEM) at an accelerating voltage of 200 kV. The ruthenium content of the samples was determined using a Perkin Elmer Opfima 7300 V type ICP-AES instrument. Thermogravimetric analysis (TG-DTA) was evaluated using a TG/DTA 6300 thermogravimetric analyzer from 30 to 800 $^\circ\text{C}$ with a temperature rate of 10 $^\circ\text{C min}^{-1}$ in air. Oxidation reaction progress was monitored by a Gas Chromatograph (Shandong Lunan Ruihong SP-7800A) equipped with a SE 54 column of 30 m.

Catalyst preparation

Synthesis of dimethyl 4-(4-(2,2':6',2''-terpyridin-4'-yl)benzyloxy)pyridine-2,6-dicarboxylate (3). Into a solution of 2 (1.87 g, 4.66 mmol) in dry DMF (50 mL) was added K_2CO_3 (2.49 g, 18.0 mmol) and 1 (0.95 g, 4.5 mmol) successively. The mixture was stirred at 80 $^\circ\text{C}$ for 14 h. Then the reaction was stopped by adding water (200 mL), and the precipitate was collected by filtration. The precipitate was extracted with dichloromethane. The dichloromethane was removed by rotary evaporation to obtain 3 (1.5 g, 63%) as a white solid. $^1\text{H NMR}$ (400 MHz, DMSO-d_6): δ 3.92 (s, 6H), 5.51 (s, 2H), 7.53–7.56 (m, 2H), 7.72 (d, $J = 8.1$ Hz, 2H), 7.89 (s, 2H), 8.00–8.07 (m, 4H), 8.69 (d, $J = 8.0$ Hz, 2H), 8.75 (s, 2H), 8.78 (d, $J = 4.0$ Hz, 2H). $^{13}\text{C NMR}$ (101 MHz, CDCl_3) δ 53.3, 70.4, 114.9, 118.9, 121.4, 123.9, 127.9, 128.2, 135.5, 136.9, 139.0, 149.2, 149.5, 149.9, 156.0, 156.1, 165.1, 166.7.

Synthesis of potassium 4-(4-(2,2':6',2''-terpyridin-4'-yl)benzyloxy)pyridine-2,6-dicarboxylate (4). Compound 3 (3.40 g, 6.39 mmol) was added to a solution of KOH (8.2 g, 0.15 mol) in THF (150 mL) and water (200 mL). The mixture was stirred at 80 $^\circ\text{C}$ for 10 h, and then was cooled to room temperature, from which a white solid was precipitated. The solid was filtrated and washed with dichloromethane. The obtained solid was dried under reduced pressure to afford product 4 (3.1 g, 84%). $^1\text{H NMR}$ (400 MHz, D_2O): δ 4.87 (s, 2H), 7.11–7.16 (m, 4H), 7.30 (d, $J = 8.2$ Hz, 2H), 7.40 (s, 2H), 7.62–7.67 (m, 4H), 7.89 (d, $J = 7.9$ Hz, 2H), 8.19 (d, $J = 4.4$ Hz, 2H). $^{13}\text{C NMR}$ (101 MHz, DMSO-

d_6 : $\text{H}_2\text{O} = 1 : 3$) δ 69.4, 111.8, 118.1, 118.1, 122.2, 124.7, 126.8, 128.2, 135.9, 137.0, 138.2, 148.7, 148.8, 154.2, 155.2, 166.3, 172.2. HRMS (ESI): m/z calcd for $\text{C}_{29}\text{H}_{19}\text{K}_2\text{N}_4\text{O}_5$; 581.0630 $[\text{M} + \text{H}]^+$; found: 581.0625.

Synthesis of oligomer-Ru(terpy)(pydic). A solution of $[\text{Ru}(\text{p-cymene})\text{Cl}_2]_2$ (0.70 g, 1.14 mmol) in DMF (20 mL) was added to a solution of 4 (1.2 g, 2.07 mmol) in DMF– H_2O (1 : 1, 500 mL) under argon atmosphere. After the mixture was stirred at 70 $^\circ\text{C}$ for 5 h, a dark violet precipitate was formed. The precipitate collected by filtration was further purified by Soxhlet extraction for 10 h with ethanol and water to remove the impurities thoroughly and finally to give oligomer Ru(terpy)(pydic) (1.1 g, 88%). Anal. Calc. for $\text{C}_{116}\text{H}_{72}\text{N}_{16}\text{O}_{20}\text{Ru}_3\text{K}_2$: C, 58.19; H, 3.03; N, 9.37; found: C, 57.93; H, 2.92; N, 9.18.

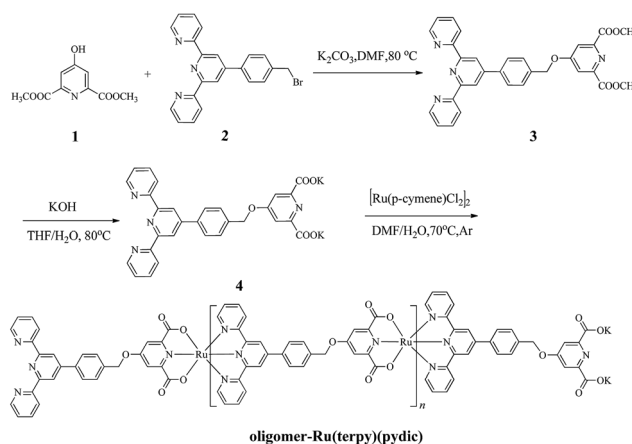
Typical procedure for oxidation of alcohols. Substrate alcohol (2.0 mmol), catalyst (0.1 mol%) and *n*-hexane (2 mL) were charged into a two necked flask. The mixture was heated to 50 $^\circ\text{C}$, then 70% TBHP (6 mmol) was added slowly into the mixture in 0.5 h under stirring. The reaction mixture was further vigorously stirred at 50 $^\circ\text{C}$ to complete the reaction. After reaction, the product was isolated by silica gel column chromatography with mixture of *n*-hexane and ethyl acetate as eluent.

The procedure for the recycling of the catalyst. 1-Phenylethanol (60 mmol), catalyst (0.06 mmol), and *n*-hexane (60 mL) were charged into a flask, successively. The mixture was heated to 50 $^\circ\text{C}$ under stirring. Then TBHP (70% aqueous solution) (180 mmol) was slowly dropped into the mixture in 0.5 h. After reaction (9.5 h), the catalyst was recovered by filtration and washed with ethanol, dried at 80 $^\circ\text{C}$ under vacuum, and then subjected to next run. In each subsequent cycle, the conversion of 1-phenylethanol was measured by GC after 9.5 h.

Results and discussion

Synthesis of oligomer-Ru(terpy)(pydic)

As shown in Scheme 1, the ligand precursor called dimethyl 4-(4-(2,2':6',2''-terpyridin-4'-yl)benzyloxy)pyridine-2,6-dicarboxylate (3) was smoothly obtained from 1 and 2 *via* Williamson reaction,



Scheme 1 Synthesis route for the oligomer-Ru(terpy)(pydic).



which was confirmed by ^1H NMR and ^{13}C NMR. Hydrolysis of ligand precursor **3** in the presence of potassium hydroxide readily gave the ligand called potassium 4-(4-(2,2':6',2''-terpyridin-4'-yl)benzyloxy)pyridine-2,6-dicarboxylate (**4**) which was thoroughly characterized by ^1H NMR, ^{13}C NMR, HR-MS, FT-IR and UV-vis spectroscopy.

The coordination of **4** with $[\text{Ru}(p\text{-cymene})\text{Cl}_2]_2$ under the conditions similar to those for preparing $\text{Ru}(\text{terpy})(\text{pydic})$ ^{50,51} immediately afforded a dark violet solid product which was thoroughly characterized, and regarded as oligomer- $\text{Ru}(\text{terpy})(\text{pydic})$.

Characterization

The synthesized solid product was firstly characterized by FT-IR. Fig. 1 shows the FT-IR spectra of ligand **4** (A) and oligomer- $\text{Ru}(\text{terpy})(\text{pydic})$ (B) in the range of 4000–400 cm^{-1} . A broad band around 3445 cm^{-1} is observed, which is indicative the presence of potassium carboxylate in the structure.⁵² The strong band at 1594 cm^{-1} is from the C=N and C=O stretching vibration,^{53,54} and the bands at 1460, 1404, and 1333 cm^{-1} are associated with C–O and C–N stretching.^{55–58} Compared with that of the ligand, the spectrum of oligomer- $\text{Ru}(\text{terpy})(\text{pydic})$ exhibits additional bands at 622 cm^{-1} and 468 cm^{-1} respectively, attributed to the Ru–O and Ru–N vibration.^{59,60} Besides, all the characteristic peaks in the IR spectrum of $\text{Ru}(\text{terpy})(\text{pydic})$ based complexes⁶¹ can be found in that of oligomer- $\text{Ru}(\text{terpy})(\text{pydic})$. To further confirm the formation of oligomer- $\text{Ru}(\text{terpy})(\text{pydic})$, the Raman spectra of oligomer- $\text{Ru}(\text{terpy})(\text{pydic})$ and its model complex $\text{Ru}(\text{terpy})(\text{pydic})$ were taken and compared. As shown in Fig. 2, the spectrum (A) of oligomer- $\text{Ru}(\text{terpy})(\text{pydic})$ is almost coincide with that (B) of $\text{Ru}(\text{terpy})(\text{pydic})$. In the spectra, the bands at 1604 cm^{-1} and 1470 cm^{-1} can be ascribed to the stretching vibration of the C=N from ligand pydic, and the bands at 1520 cm^{-1} and 1355 cm^{-1} are due to the stretching vibration of the C=N from ligand terpy. The band at 650 cm^{-1} and the weak band at 455 cm^{-1} can be attributed to the Ru–O and Ru–N vibrations, respectively. Both the FT-IR and Raman spectra of oligomer- $\text{Ru}(\text{terpy})(\text{pydic})$ indicate that the obtained ruthenium complex has a $\text{Ru}(\text{terpy})(\text{pydic})$ unit in its structure. Besides, the procedures for preparing oligomer- $\text{Ru}(\text{terpy})(\text{pydic})$ are identical to those for the synthesis of $\text{Ru}(\text{terpy})(\text{pydic})$. Therefore, we conclude that the

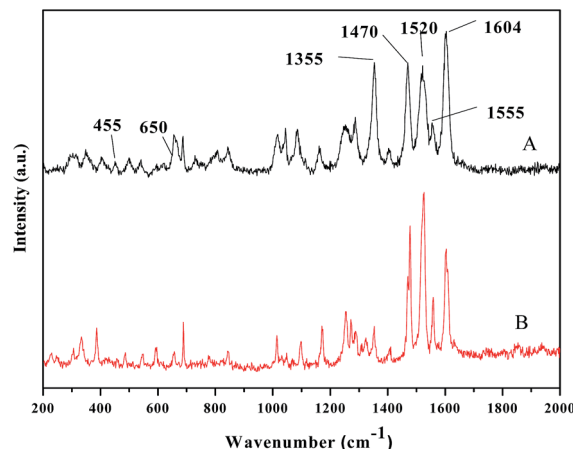


Fig. 2 Raman spectra of oligomer- $\text{Ru}(\text{terpy})(\text{pydic})$ (A) and $\text{Ru}(\text{terpy})(\text{pydic})$ (B).

oligomer- $\text{Ru}(\text{terpy})(\text{pydic})$ has the structure as shown in Scheme 1, in which both a $\text{Ru}(\text{terpy})(\text{pydic})$ unit and a potassium carboxylate moiety are involved. In the $\text{Ru}(\text{terpy})(\text{pydic})$ unit, the terpy coordinates to central ruthenium atom *via* three pyridine nitrogens, and the pydic binds the metal center at the pyridine nitrogen and two carboxylate oxygens. The ligand terpy coordinates to Ru on one plane and the ligand pydic acts as a dianionic *O,N,O*-tridentate ligand on another plane. The two planes are nearly perpendicular to each other.⁶²

ICP analysis revealed that the ruthenium content of the oligomer- $\text{Ru}(\text{terpy})(\text{pydic})$ was 12.7 wt%, from which it is calculated that the average polymerization degree of the obtained oligomer- $\text{Ru}(\text{terpy})(\text{pydic})$ is about 4.0, and the *n* in the structure shown in Scheme 1 is about 2.0. The elementary analysis results of C, H and N are in agreement with the calculated ones based on the structure determined from ruthenium content.

The UV-vis DRS spectra for **4** (A) and oligomer- $\text{Ru}(\text{terpy})(\text{pydic})$ (B) are showed in Fig. 3. In the spectrum of **4**, two bands maxima at 234 and 333 nm are derived from the π - π^* transitions of phenyl group and pyridine ring,^{63,64} and no absorption band above 400 nm is found. Nevertheless, the spectrum of oligomer-

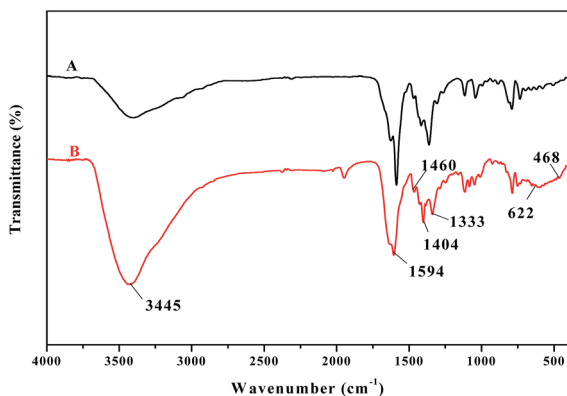


Fig. 1 FT-IR spectra of ligand **4** (A) and oligomer- $\text{Ru}(\text{terpy})(\text{pydic})$ (B).

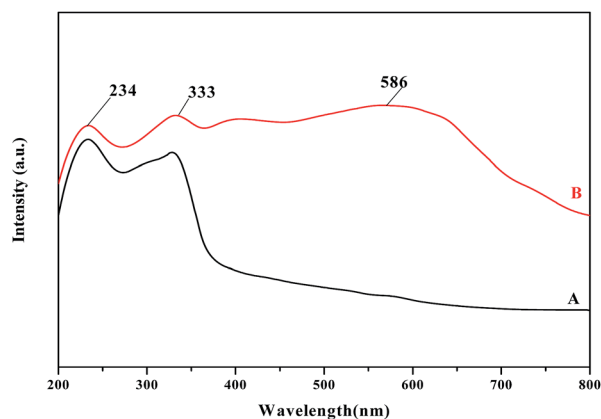


Fig. 3 UV-vis DRS spectra of **4** (A) and oligomer- $\text{Ru}(\text{terpy})(\text{pydic})$ (B).



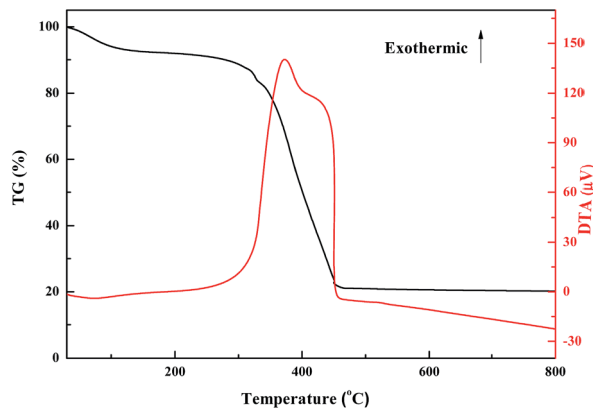


Fig. 4 TG and DTA curves of the thermal degradation of oligomer-Ru(terpy)(pydic).

Ru(terpy)(pydic) exhibits an additional broad band centered at 586 nm, assigned to the metal–ligand charge transfer transition [MLCT, $d(\text{Ru}) \rightarrow \pi^*(\text{pyridine})$], which is a characteristic transition band in ruthenium polypyridine complexes.^{65,66}

The thermal stability of oligomer-Ru(terpy)(pydic) was evaluated by thermal gravimetric analysis (Fig. 4). Two weight loss stages can be seen from the TG curve. The first endothermic stage occurring in the temperature range of 50–110 °C with a weight loss close to 7%, is due to the desorption of physically adsorbed water. The second weight loss occurring in the range of 250 to 460 °C can be ascribed to the exothermic oxidative decomposition of oligomer-Ru(terpy)(pydic). These results proved that oligomer-Ru(terpy)(pydic) is thermally stable and could be used under 250 °C.

Fig. 5(a) and (b) show the SEM images of the fresh and used samples of oligomer-Ru(terpy)(pydic), respectively. The images show that the morphologies of both the fresh and used samples composed of particles with irregular shapes and different

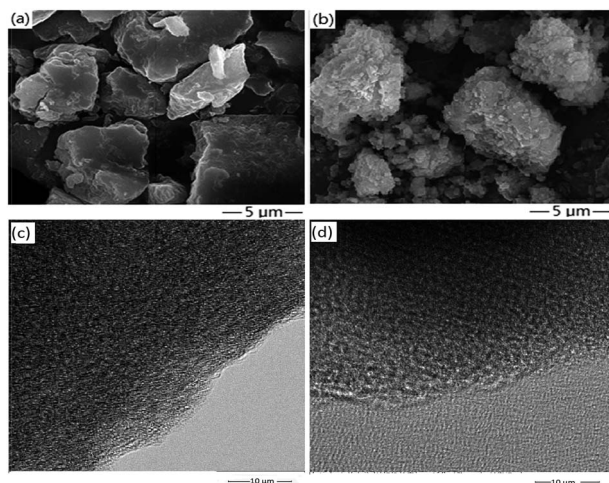


Fig. 5 (a) SEM image of the fresh oligomer-Ru(terpy)(pydic), (b) SEM image of the used oligomer-Ru(terpy)(pydic), (c) TEM image of the fresh oligomer-Ru(terpy)(pydic), (d) TEM image of the used oligomer-Ru(terpy)(pydic).

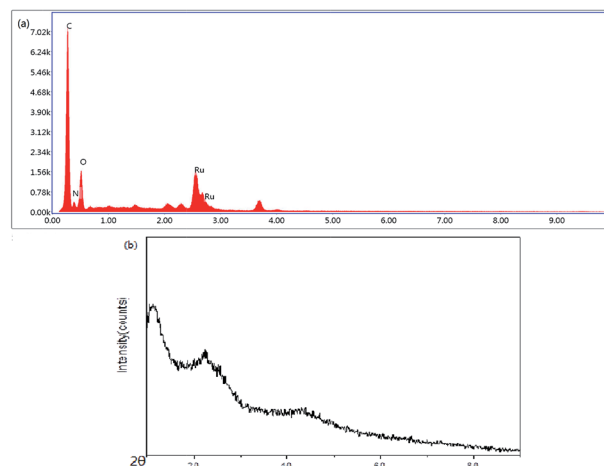


Fig. 6 (a) EDX analysis of oligomer-Ru(terpy)(pydic), (b) XRD analysis of oligomer-Ru(terpy)(pydic).

dimensions. The surfaces of the particles of the fresh sample are smooth, but the surfaces of the particles of the used one are rough. Besides, fragmentation of some particles is found in the used samples. These results indicated that some oligomer-Ru(terpy)(pydic) molecules decomposed and peeled off from the surface during catalytic runs.

Fig. 5(c) and (d) exhibit the TEM of the fresh and used samples of oligomer-Ru(terpy)(pydic), respectively. As can be seen that the fresh sample has a porous structure composed of small particles, and no ruthenium aggregates are found in the structure. After catalytic run, the used sample shows big pores compared to the fresh one, indicating the peeling off of the surface oligomer-Ru(terpy)(pydic) molecules on the particles due to their decomposition during catalytic runs. No ruthenium aggregates are found, too. The TEM results are in accordance with those from SEM analysis.

The chemical composition of oligomer-Ru(terpy)(pydic) was characterized by energy dispersive spectroscopy. As shown in Fig. 6(a), the EDX spectrum exhibits peaks for C, N, O and Ru, which are the component elements of oligomer Ru(terpy)(pydic). The EDX analysis of the fresh sample of oligomer-Ru(terpy)(pydic) results showed the molar ratios for N : O to be about 1 : 1.31, which is well matched with the result calculated from the structure of oligomer Ru(terpy)(pydic). From EDX results the Ru content was calculated to be 13.0 wt%, which is in agreement with the result from ICP analysis. Fig. 6(b) shows the XRD patterns of the sample of oligomer-Ru(terpy)(pydic) and no diffraction peaks representing ruthenium species are observed, which indicates that ruthenium does not exist in the form of inorganic particles in the sample.

The surface chemical composition and the oxidation state of ruthenium in oligomer-Ru(terpy)(pydic) were investigated by XPS analysis. The full-range XPS spectra (Fig. 7(a)) clearly showed characteristic peaks of O, N and Ru elements. The narrow scan spectra of elements C 1s, Ru 3d and Ru 3p are shown in Fig. 7(b and c), respectively. The appearance of Ru 3d_{5/2} (280.2 eV) and Ru 3p_{3/2} (462.2 eV) peaks makes it evident that the ruthenium is present in Ru²⁺ state in the oligomer complex.^{67,68}



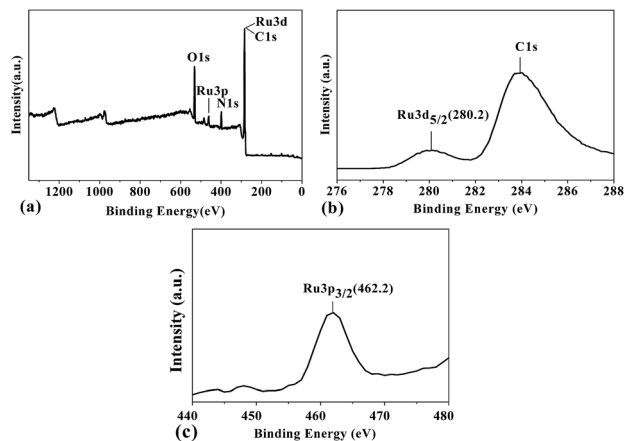


Fig. 7 XPS spectra of oligomer-Ru(terpy)(pydic): (a) wide view, (b) C 1s and Ru 3d, (c) Ru 3p spectra.

Catalytic properties of oligomer-Ru(terpy)(pydic)

The oligomer ruthenium complex was evaluated in the oxidation of secondary alcohols as a catalyst. Initially, several oxidants including molecular oxygen, H_2O_2 and TBHP were tested in the reaction with 1-phenylethanol as a model substrate in cyclohexane (see Table S1 in the ESI†). No reaction was observed with molecular oxygen as oxidant. In the case of H_2O_2 as oxidant, the yield of the target product acetophenone was only 32% for 16 h, which is different from that under the catalysis of Ru(terpy)(pydic) in literature.²⁵ In that case, H_2O_2 showed good performance in the oxidation of secondary alcohols to the corresponding ketones. Under the same reaction conditions, the yield of acetophenone reached up to 96% with TBHP as oxidant. Next, the solvent was screened. In viewpoint of the yield of acetophenone, *n*-hexane was the best one among the screened solvents (see Table S1 in the ESI†). After the determination of oxidant and solvent, the parameters such as the loadings of TBHP and catalyst, and the reaction temperature were optimized thoroughly (see Tables S2–S4 in the ESI†). Based on the experimental results, the optimal reaction conditions for the oxidation of alcohols were obtained, which are alcohol substrate 2 mmol, catalyst 0.1 mol%, molar ratio of TBHP to substrate 3 : 1, reaction temperature 50 °C, and *n*-hexane 2 mL.

The substrate tolerance of the catalytic oxidation system was then investigated under the optimal reaction conditions. The results are compiled in Table 1. As shown in Table 1, the catalytic oxidation system has good substrate tolerance. Various secondary alcohols including benzylic, linear and cyclic aliphatic ones were oxidized to their corresponding ketones. For the benzylic alcohols, the ones with whether an electron-donating or an electron-withdrawing group were almost oxidized to the corresponding ketones quantitatively, but the reaction time to finish the reaction was different (Table 1, entries 1–14). The substrate with an electron-withdrawing substituent *para* to the hydroxyalkyl group had high reactivity compared to the one with an electron-donating substituent, and the stronger the electron-withdrawing capacity of the substituent is, the shorter time is required to finish the reaction (Table

1, entries 2, 3, 6–8). The *ortho*-substituted substrates reacted slowly compared to the *para*- or *meta*-isomer due to the steric hindrance of between the substituent and the hydroxyalkyl group. For examples, long reaction time and high TBHP loading were required in the oxidation of 1-(2-chlorophenyl) ethanol and 1-(2-methoxyphenyl)ethanol compared to that of 1-(4-chlorophenyl) ethanol and 1-(4-methoxyphenyl) ethanol (Table 1, entries 3, 5, 8, 9). In the oxidation of 2-pentanol and cyclohexanol which are the representatives of linear and cyclic secondary aliphatic alcohols, excellent yields were also obtained though long reaction time and high TBHP loading were needed (Table 1, entries 15, 18, 19). However, in the oxidation of 2-isopropyl-5-methylcyclohexanol poor result was received due to iso-propyl *ortho* to the hydroxyl leading to big steric hindrance in the structure (Table 1, entries 16, 17). The catalytic oxidation system was also evaluated in the oxidation of some primary benzylic alcohols. In all the cases the primary alcohols were over oxidized to the corresponding carboxylic acids (Table 1, entries 20–23). Because the aldehydes from alcohol were more easily oxidized than alcohols, more TBHP was consumed in the reaction of aldehydes to acids leading to low conversion of alcohols. Meanwhile, it is difficult to control the selectivity either to aldehydes or carboxylic acids. High TBHP loading was in favor of the reaction to the acid completely. For instance, both the conversion of alcohol and the yield of acid can approach 100% at 10 mmol TBHP loading (Table 1, entry 24). Finally, 4-*tert*-butoxyphenethyl alcohol and 3-phenyl-1-propanol as representatives of the primary aliphatic alcohols were tested in the oxidation by the catalytic oxidative system. However, the expected products 4-*tert*-butoxyphenylacetic acid and 3-phenylpropionic acid were not obtained. Interestingly, the former was transformed to 4-*tert*-butoxybenzoic acid, and the latter was oxidized to 3-hydroxy-1-phenylpropan-1-one, respectively (Table 1, entries 25–26). Both of the products were derived from the oxidation of benzylic methylene of substrates.

The results in Table 1 indicate that this catalytic system is more suitable for the selective oxidation of the secondary alcohols to the corresponding ketones, which is complementary with the copper/TEMPO catalyzed aerobic oxidation of the primary alcohols to aldehydes.^{14,69,70} Generally, the copper/TEMPO based aerobic oxidation systems are not efficient in the oxidation of the secondary alcohols to ketones. This oligomer-Ru(terpy)(pydic) is only efficient in the case of TBHP as oxidant, however, the homogeneous ruthenium complexes with a similar structure unit to it are generally active in the oxidation of alcohols with H_2O_2 .^{24,25} It is obvious that H_2O_2 is more environmentally friendly than TBHP.

Additional experiments were conducted to evaluate the heterogeneity of the reaction catalyzed by oligomer-Ru(terpy)(pydic) with 1-phenylethanol as the substrate. First, oligomer-Ru(terpy)(pydic) (0.1 mol%) was suspended in the mixture of *n*-hexane (2 mL) and 1-phenylethanol (2 mmol) under stirring at 50 °C for eight hours being the same time in the oxidation of 1-phenylethanol by TBHP. The solid catalyst was removed from the mixture by hot filtration. To the filtrate was added TBHP (6 mmol) in dropwise and the mixture was stirred at 50 °C for eight hours. Only 24% of the conversion of 1-



Table 1 Oxidation of various alcohols with TBHP catalyzed by oligomer-Ru(terpy)(pydic)^a

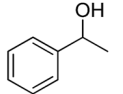
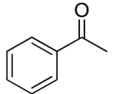
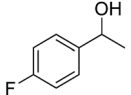
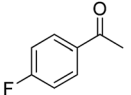
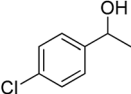
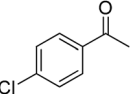
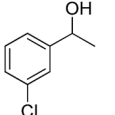
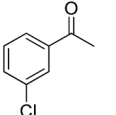
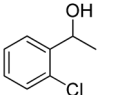
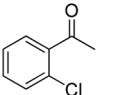
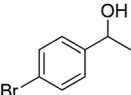
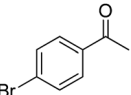
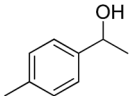
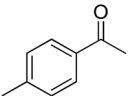
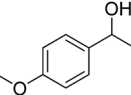
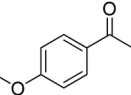
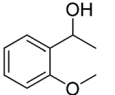
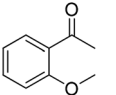
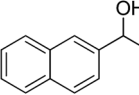
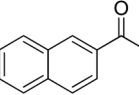
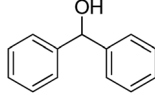
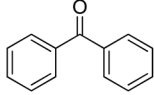
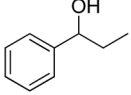
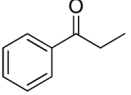
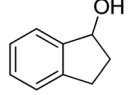
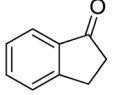
$\text{R}_1\text{-CH(OH)-R}_2 \xrightarrow[\text{TBHP, } n\text{-hexane, } 50\text{ }^\circ\text{C}]{\text{oligomer-Ru(terpy)(pydic) (0.1 mo l\%)}} \text{R}_1\text{-C(=O)-R}_2 \quad \text{or} \quad \text{R}_1\text{-C(=O)-OH} \quad (\text{R}_2 = \text{H})$						
Entry	Substrate	Product	Time (h)	Conv. ^b (%)	Select. ^b (%)	Yield ^c (%)
1			8.5	>99	>99	94
2			4.0	>99	>99	91
3			5.5	>99	>99	93
4			8.0	>99	>99	94
5 ^d			20.0	96	>99	89
6			6.5	>99	>99	94
7			10.0	>99	>99	93
8			13.0	>99	>99	94
9 ^e			23.0	>99	>99	90
10			4.5	>99	>99	90
11			3.0	>99	>99	94
12			10.0	>99	>99	89
13			5.0	>99	>99	94

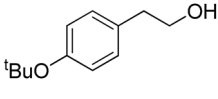
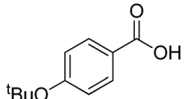
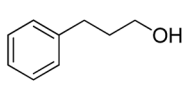
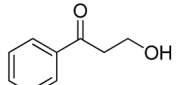


Table 1 (Contd.)

$\text{R}_1\text{CH}(\text{OH})\text{R}_2 \xrightarrow[\text{TBHP, } n\text{-hexane, } 50\text{ }^\circ\text{C}]{\text{oligomer-Ru(terpy)(pydic) (0.1 mo l\%)}} \text{R}_1\text{C}(=\text{O})\text{R}_2 \quad \text{or} \quad \text{R}_1\text{C}(=\text{O})\text{OH} \quad (\text{R}_2 = \text{H})$						
Entry	Substrate	Product	Time (h)	Conv. ^b (%)	Select. ^b (%)	Yield ^c (%)
14			12.5	>99	>99	90
15 ^f			16.0	>99	>99	85
16			12.0	22	>99	17
17 ^d			16.0	39	>99	28
18			17.5	67	>99	56
19 ^d			20.0	>99	>99	83
20 ^f			10.0	99	62	59
					37	30
21 ^f			9.0	98	70	62
					29	22
22 ^d			12.0	98	75	65
					24	19
23 ^f			11.0	98	80	64
					19	15
24 ^g			10.5	>99	>99	95



Table 1 (Contd.)

Entry	Substrate	Product	Time (h)	Conv. ^b (%)	Select. ^b (%)	Yield ^c (%)
25 ^d			16.0	65	95	51
26 ^d			16.0	65	60	32

^a Reaction condition: 1-phenylethanol (2 mmol), oxidant (6 mmol), catalyst (0.1 mol%), *n*-hexane (2 mL), reaction temperature 50 °C. ^b Conversions and selectivity were determined by GC (area normalization method). ^c Isolated yields, the product was purified by column chromatography over silica gel (eluent: *n*-hexane/ethyl acetate), all products were determined by ¹H NMR. ^d Oxidant: 8 mmol. ^e Oxidant: 7 mmol. ^f Oxidant: 7.5 mmol. ^g Oxidant: 10 mmol.

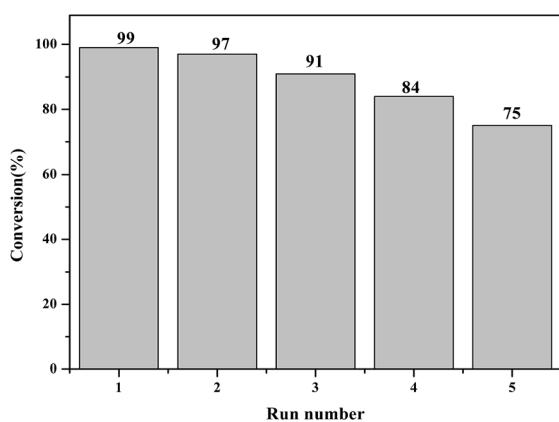


Fig. 8 Recycling of oligomer-Ru(terpy)(pydic) in the oxidation of 1-phenylethanol. Reaction condition: 1-phenylethanol (60 mmol), catalyst (0.06 mmol), oxidant (180 mmol), *n*-hexane (60 mL), temperature 50 °C, reaction time 9.5 h.

phenylethanol was observed. In another experiment, to the mixture of *n*-hexane (2 mL) and 1-phenylethanol (2 mmol) without catalyst was added TBHP (6 mmol) in dropwise under stirring at 50 °C for eight hours, too. 1-Phenylethanol was converted to acetophenone in a conversion of 22% in the absence of catalyst, which is close to the above result (24%) of the oxidation of the hot filtrate. The results indicated that oligomer-Ru(terpy)(pydic) is almost insoluble in *n*-hexane, and oligomer-Ru(terpy)(pydic) is belong to a heterogeneous catalyst.

The recyclability of the catalyst

Further experiments were performed to examine the reusability of oligomer-Ru(terpy)(pydic) with 1-phenylethanol as a model substrate. The procedures have been given in the experimental section. The recycle test results are displayed in Fig. 8. As shown

in the figure, the conversion decreased gradually with recycle times but still reached 84% in the fourth run. It can be concluded that the catalyst could be recycled at least four cycles without a large reduction in its activity and selectivity. The oligomer-Ru(terpy)(pydic) showed better performance than the organic polymer supported ruthenium complexes but poorer performance than the inorganic immobilized ruthenium complexes.^{27,40–43} The ruthenium content of the samples of the fresh catalyst and recovered one after the fifth run was determined by ICP to be 12.7 wt% and 10.5 wt%, respectively. The ruthenium loss is approximately 17.3%. This result in combination with the SEM images of the fresh and used samples indicates that the decrease in activity of the oligomer ruthenium complex was mainly due to its light decomposition during catalytic run.

Conclusions

A ligand with both terpyridine and pyridine-2,6-dicarboxylate functional groups were designed and synthesized. The ligand can coordinate with Ru²⁺ to give an oligomer ruthenium complex with an average polymerization degree of 4.0 named as oligomer-Ru(terpy)(pydic). This oligomer ruthenium complex as a heterogeneous catalyst showed high activity and selectivity in the oxidation of various secondary alcohols to the corresponding ketones with TBHP as oxidant. The oligomer ruthenium complex can be easily separated and recycled several times. The gradual deactivation of this complex is due to its slight decomposition during the catalytic run.

Conflicts of interest

There are no conflicts to declare.



Acknowledgements

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