



Cite this: *RSC Adv.*, 2017, 7, 41962

Bi-functional reduced graphene oxide/AgCo composite nanosheets: an efficient catalyst and SERS substrate for monitoring the catalytic reactions†

Xiaowei Ma, Yue Guo, Jing Jin, Bing Zhao  and Wei Song *

The preparation of supported monodispersed bimetallic nanoparticles is important in the fields of catalysis and surface-enhanced Raman scattering (SERS) detection. In this work, a simple *in situ* co-reduction approach has been demonstrated for the synthesis of AgCo alloy nanoparticles with a size of 20–100 nm supported on the surface of reduced graphene oxide (rGO) nanosheets. The resulting rGO/AgCo composite nanosheets are well characterized using scanning electron microscopy (SEM), transmission electron microscopy (TEM), Fourier transform infrared spectroscopy (FTIR), Raman spectroscopy, X-ray diffraction (XRD), and X-ray photoelectron microscopy (XPS) measurements. The as-prepared rGO/AgCo composite nanosheets show a superior catalytic activity for the reduction of methylene blue (MB) in the presence of NaBH₄ and the degradation of MB under ultraviolet light over rGO/Ag, rGO/Co, rGO/Ag₃Co₁ and rGO/Ag₁Co₃ composite nanosheets. In addition, the rGO/AgCo composite nanosheets have also been proved to be efficient surface-enhanced Raman scattering (SERS) substrates. The unique two-dimensional structure and strong interactions between rGO and AgCo alloy nanoparticles contribute to the electromagnetic field SERS enhancement. Owing to the excellent catalytic properties and unique SERS sensitivity, a facile approach to monitor and determine the reaction kinetics of the catalytic reduction or photocatalytic degradation of MB using rGO/AgCo composite nanosheets as both catalyst or photocatalyst and SERS substrate has been demonstrated. In addition, the as-prepared rGO/AgCo composite nanosheets have also been used as efficient SERS substrates to monitor the plasmon-driven catalytic reaction of *p*-nitrothiophenol (PNTp) dimerizing into 4,4'-dimercaptoazobenzene (DMAB). Therefore, it is anticipated that this approach will be used to fabricate many other kinds of supported nanocatalysts, which could be applicable to study more general catalytic reaction processes and provide more information on the catalytic intermediates and final products.

Received 29th June 2017
Accepted 24th August 2017

DOI: 10.1039/c7ra07216c

rsc.li/rsc-advances

1. Introduction

During the past decades, the fabrication of transition metal nanomaterials has attracted considerable interest due to their exceptional properties and broad applications in the fields of catalysis, electrocatalysis, electronics, chemical sensing, and biomedicine.^{1–5} Among the target materials, bimetallic nanomaterials have been especially interesting due to their synergistic enhanced properties.⁶ For example, Xia and co-workers reported the preparation of Pd–Pt bimetallic nanodendrites toward the oxygen reduction reaction (ORR).⁷ It was found the Pd–Pt bimetallic nanodendrites showed five times more activity than the Pt-black catalyst. However, the naked bimetallic

nanoparticles with small size are usually easily aggregated due to their high surface energies, resulting in a decrease or loss of catalytic activity. Therefore, it is desirable to prepare bimetallic nanoparticles supported on some appropriate matrices such as carbon, molecular sieves, and polymers.^{8–10} For example, Jiang and co-workers synthesized bimetallic AgCo nanoparticles supported on loofah sponge carbon (LSC) for catalyzing the hydrolysis of NaBH₄ toward hydrogen generation.⁸ It was found that the prepared LSC/AgCo composite with an appropriate ratio of Ag and Co showed a much higher catalytic activity than that of individual LSC/Ag and LSC/Co composites, demonstrating a synergistic effect between Ag and Co counterparts.

Graphene is an excellent support to deposit bimetallic nanomaterials due to their unique two-dimensional (2D) structure and large surface area, which offers them promising applications in catalysis, photocatalysis and electrocatalysis.^{11–13} For example, Feng and co-workers prepared porous Pt–Au nanodendrites supported on reduced graphene oxide (rGO)

State Key Laboratory of Supramolecular Structure and Materials, Jilin University, Changchun 130012, P. R. China. E-mail: weisong@jlu.edu.cn; Fax: +86-431-85168473; Tel: +86-431-85168473

† Electronic supplementary information (ESI) available. See DOI: 10.1039/c7ra07216c



nanosheets for catalytic reduction of 4-nitrophenol by NaBH_4 , which exhibited an enhanced catalytic activity compared to Pt black and Au nanocrystals.¹⁴ Kidwai and co-workers also demonstrated that AgNi alloy nanoparticles supported on rGO showed a superior catalytic activity over rGO/Ag, rGO/Ni and individual AgNi nanoparticles toward the reduction of 4-nitrophenol in the presence of NaBH_4 .¹⁵ Recently, rGO has also been reported to act as a support to immobilize Ag@CoNi nanoparticles, which displayed much higher catalytic activities toward the hydrolytic dehydrogenation of ammonia borane (AB) than the monometallic, bimetallic and individual Ag@CoNi counterparts.¹⁶

On the other hand, due to the unique 2D structure and excellent adsorption properties, graphene has been proved to be a good candidate as an efficient surface-enhanced Raman scattering (SERS) substrate.^{17,18} Recently, graphene supported metal nanoparticles have also been widely reported as SERS substrates with a high sensitivity.^{19–21} For instance, Ag or Au nanoparticles have been fabricated on the surface of reduced graphene oxide (rGO) nanosheets for SERS substrate to detect adsorbed aromatic molecules with a detection limit of only 10 nM. Recently, it has been attracted more and more attention for the fabrication of SERS-active substrates for the monitoring of the catalytic or photocatalytic reactions, which can provide some new insights into the reaction kinetics and mechanism. For example, Kang, Shao, and coworkers prepared ZnO–rGO–Au hybrid nanomaterials to be used as a both photocatalyst and SERS substrate, which can quantitatively monitor the catalytic degradation of Rhodamine 6G.²² Recently, we have demonstrated a facile approach for the fabrication of Ag–Cu₂O/rGO nanocomposites, which can be employed as not only a peroxidase-like nanocatalyst but also good SERS substrate.²³ On the basis of the high performance of bifunctional Ag–Cu₂O/rGO nanocomposites, the detection of glucose with a high sensitivity and high selectivity has been achieved.

In this work, we have demonstrated a facile *in situ* co-reduction approach for the preparation of AgCo alloy nanoparticles supported on the surface of rGO nanosheets. The as-prepared rGO/AgCo composite nanosheets are not only good catalysts for the reduction and photocatalysts for the degradation of methylene blue (MB), but also efficient SERS substrate due to the electromagnetic field contributions. More importantly, the as-prepared rGO/AgCo composite nanosheets can be used to monitor and determine the reaction kinetics of the catalytic reduction and photocatalytic degradation of methylene blue (MB). In addition, the as-prepared rGO/AgCo composite nanosheets has also been demonstrated to be used as SERS substrate to monitor of plasmon-driven catalytic reaction of *p*-nitrothiophenol (PNTp) dimerizing into 4,4'-dimercaptoazobenzene (DMAB). Our work provides a facile and general approach for the modification of metal alloy nanoparticles on the surface of rGO as bifunctional nanocatalyst and SERS substrate. Importantly, the hybrid nanostructures offer an opportunity for the monitoring of the catalytic and photocatalytic reaction process and determination of the reaction kinetics. This work will, therefore, be of great significance to study more general catalytic reaction

processes and provide more information of the catalytic kinetics and mechanism.

2. Experiment

2.1 Chemicals and materials

Silver nitrate (AgNO_3) was purchased from Beijing Chemical Works. Cobalt nitrate ($\text{Co}(\text{NO}_3)_2 \cdot 6\text{H}_2\text{O}$) was obtained from Xilong Chemical Co., Ltd. Natural graphite powder and sodium borohydride (NaBH_4) was bought from Sinopharm chemical reagent Beijing Co., Ltd. MB was purchased from Tianjin Guangfu Fine Chemical Research Institute. *p*-Aminothiophenol (PATP), 4-mercaptobenzoic acid (MBA) and PNTp was acquired from Sigma-Aldrich. All the chemicals were used as received without purification. Deionized water was used throughout the study.

2.2 Synthesis of rGO/AgCo composite nanosheets

First, GO was synthesized from natural graphite powder *via* a modified Hummer and Offeman method.^{24,25} Then, rGO/AgCo composite nanosheets were prepared *via* an *in situ* co-reduction approach. In a typical procedure, 40 mg of GO was added into 30 mL of deionized water under sonication for about 1 h, then a clear uniform GO aqueous solution was prepared. In the following, 0.2 mmol $\text{Co}(\text{NO}_3)_2 \cdot 6\text{H}_2\text{O}$ and 0.2 mmol AgNO_3 were added into the above GO aqueous solution, then 12 mL of NaBH_4 (1.0 M) was added. The pH value of the solution was about 9.0 at this time. After stirring for about 30 min, the color of the solution changed from dark brown to black. After that, the above solution was transferred into a 50 mL round-bottom flask and heated at 110 °C under stirring for another 2.5 h. After cooling down, the as-prepared product was collected by centrifuge and washed thoroughly with water and ethanol for several times. Finally, the prepared rGO/AgCo composite nanosheets were dried in air. The rGO/Ag and rGO/Co composite nanosheets with the same content of metal were also prepared *via* a similar procedure. The rGO/AgCo composite nanosheets with different Ag/Co molar ratios were synthesized by varying the concentration of AgNO_3 and $\text{Co}(\text{NO}_3)_2 \cdot 6\text{H}_2\text{O}$ in initial solution with the total molar amount of (Ag + Co) was kept constant. To find the influence of pH values of the solution for the reduction process, the pH values of solution were changed to 3 or 14 by adding a certain amount of HNO_3 or NaOH after the addition of NaBH_4 .

2.3 Catalytic reduction and photocatalytic degradation of MB

In a typical procedure for the catalytic reduction of MB by NaBH_4 , 30 μL of MB solution (10^{-3} M) was added into 2.67 mL of water. Then 200 μL of rGO/AgCo composite nanosheets aqueous dispersions (1 mg mL^{-1}) was added into the above solution. After that, 100 μL of NaBH_4 solution (10^{-1} M) was added into the above solution. For the photocatalytic degradation of MB under ultraviolet (UV) light, 30 μL of MB solution (10^{-3} M) was added into 2.77 mL of water. Then 200 μL of rGO/AgCo composite nanosheets aqueous dispersions (1 mg mL^{-1})



was added into the above solution. After that, the above solution was illuminated under UV light. UV-vis spectral measurement was used to evaluate the catalytic and photocatalytic activities of the prepared rGO/AgCo composite nanosheets.

2.4 SERS measurements

In a typical experiment, PATP and MBA were used as Raman probes to estimate the response of the rGO/AgCo composite nanosheets substrate for SERS measurements. For the SERS detection of PATP, 180 μL of PATP solution with different concentrations was mixed with 20 μL of rGO/AgCo composite nanosheets aqueous dispersions (1 mg mL^{-1}). After standing for about 2 h, 60 μL of the above solution was transferred for Raman measurements. The similar procedure was performed for the SERS detection of MBA under continuous 633 nm laser excitation.

2.5 SERS monitoring of the catalytic reduction and photocatalytic degradation of MB

In a typical procedure for the SERS monitoring of the catalytic reduction of MB by NaBH_4 , 30 μL of MB aqueous solution (10^{-3} M) was diluted to 2 mL, then 180 μL of the diluted MB aqueous solution was mixed with 20 μL of rGO/AgCo composite nanosheets aqueous dispersions (1 mg mL^{-1}). After that, 20 μL of NaBH_4 solution (10^{-2} M) was added into 40 μL of the above solution. The catalytic reduction process was measured by monitoring the SERS spectra of the final product at different reaction time. For the SERS monitoring of the photocatalytic degradation of MB under UV light, 30 μL of MB aqueous solution (10^{-3} M) was diluted to 2 mL, then 180 μL of the diluted MB aqueous solution was mixed with 20 μL of rGO/AgCo composite nanosheets aqueous dispersions (1 mg mL^{-1}). After that, the above solution was performed under UV light. SERS measurement was performed at different reaction time under continuous 532 nm laser excitation.

2.6 SERS *in situ* monitoring of plasmon-enhanced driven catalytic reaction of PNTP dimerizing into DMAB

In a typical experiment, PNTP ethanol solution (10^{-3} M) was diluted with water to prepare PNTP solution with a concentration of 10^{-5} M . Then 45 μL of the PNTP aqueous solution was mixed with 5 μL of rGO/AgCo composite nanosheets aqueous dispersions (1 mg mL^{-1}) for 30 min to ensure sufficient adsorption. After that, time-dependent SERS spectra of the plasmon-driven catalytic reaction were acquired directly from the above solution under continuous 532 nm laser excitation.

2.7 Characterization

The morphology of the rGO/AgCo composite nanosheets was operated on a scanning electron microscopy (SEM, Nova NanoSEM 450) and transmission electron microscopy (TEM, JEOL JEM-1200 EX) operated at 15 and 100 kV, respectively. High-resolution TEM (HRTEM) images and energy dispersive X-ray (EDX) analysis were performed on a JEOL JEM-2100F TEM operated at 200 kV. X-ray diffractometer (XRD, Empyrean

PANalytical B.V.) with Cu-K α radiation was used to measure the crystal structure of the samples. Raman spectra were characterized by Raman spectrometer (Renishaw-1000 spectrometer (He/Ne laser 532 nm) and LabRAM ARAMIS Smart Raman Spectrometer (air-cooled He/Ne laser 633 nm)). The chemical composition of the synthesized rGO/AgCo composite nanosheets was observed through X-ray photoelectron spectrometer (XPS, Thermo Scientific ESCALAB250). Ultraviolet-visible-near infrared (UV-VIS-NIR) spectra were performed on a SHIMADZU UV-3600 UV-Via-NIR spectrophotometer. Fourier transform infrared (FTIR) spectra were recorded on a BRUCKER VERTEX 80V FTIR Spectrometer.

3. Results and discussion

The rGO/AgCo composite nanosheets have been prepared *via an in situ* co-reduction approach. During the reaction process, the GO was reduced to rGO by NaBH_4 , and at the same time, AgCo nanoparticles were formed on the surface of rGO due to the same reduction process. Fig. 1a showed a typical SEM image of the as-prepared rGO/AgCo composite nanosheets. It was clearly observed that the rGO/AgCo composite nanosheets possessed a 2D layer structure, and some of these nanosheets were wrinkled together. The TEM image showed that AgCo nanoparticles were well dispersed on the surface of rGO nanosheets (Fig. 1b). The size of AgCo nanoparticles was from 20 to 100 nm. HRTEM image was further performed to reveal the crystalline structure of the as-prepared rGO/AgCo composite nanosheets, demonstrating high crystallinity of AgCo nanoparticles was calculated to be around 0.22 nm, which is indexed to the (111) plane of face-centered cubic (fcc) of AgCo alloy nanocrystals (Fig. 1c). The as-prepared rGO/AgCo composite nanosheets have also been characterized by energy-dispersive X-ray (EDX) spectroscopy (Fig. 1d), which displayed the existence of C, O, Co, Ag, Cu, and Si elements in the spectrum. The Cu and Si elements originate from carbon coated copper grid and instrument. These results demonstrated the evidence of the successfully preparation of rGO/AgCo composite nanosheets.

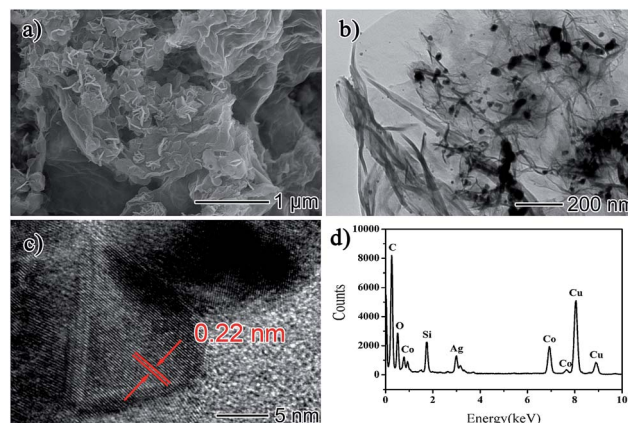


Fig. 1 (a) SEM and (b) TEM image of the rGO/CoAg composite nanosheets; (c) HRTEM image of the synthesized rGO/CoAg composite nanosheets; (d) EDX images of the synthesized rGO/CoAg composite nanosheets.



FTIR and Raman spectra were also used to characterize the chemical structure of the as-prepared rGO/AgCo composite nanosheets. Fig. 2a shows the FTIR spectra of GO and rGO/AgCo composite nanosheets, respectively. It can be found the presence of numerous oxygen containing groups in the FTIR spectrum of GO product. For instance, the broad absorption band at 3394 cm^{-1} is ascribed to O–H stretching vibration. The typical peaks at 1729 , 1395 and 1047 cm^{-1} are assigned to C–O stretching mode, O–H bending mode and C–O stretching mode, respectively. The characteristic peak at 1625 cm^{-1} is related to the vibration of adsorbed water molecules on GO. After the modification of AgCo nanoparticles, the intensity of the absorption bands at 3394 , 1729 , 1395 and 1047 cm^{-1} decreased or disappeared, demonstrating that GO has been reduced up to a great extent. The appearance of the characteristic new absorption peak at 1564 cm^{-1} is ascribed to the skeletal vibration of rGO nanosheets. These results are in accordance with the previous reports.^{15,26} The chemical structure of the as-prepared rGO/AgCo composite nanosheets has also been revealed by Raman spectroscopy. As shown in Fig. 2b, two dominant D and G bands at 1345 and 1592 cm^{-1} were observed in all of the GO, rGO/AgCo, rGO/Ag and rGO/Co composite

nanosheets. However, the intensity ratios between D and G band for rGO/AgCo ($I_D/I_G = 0.964$), rGO/Ag ($I_D/I_G = 1.056$) and rGO/Co composite nanosheets ($I_D/I_G = 0.978$) are higher than that of individual GO ($I_D/I_G = 0.754$), suggesting the formation of a conjugated graphene network after the reduction of GO.²⁷

The composition and phase purity of the as-prepared rGO/AgCo composite nanosheets has been further examined by a XRD measurement. As shown in Fig. 3, the diffraction pattern of the rGO/AgCo composite nanosheets display five characteristic bands centered at 38.5 , 44.5 , 64.7 , 77.8 and 81.9° , which showed a slight deviation from the (111), (200), (220), (311) and (222) planes of face centered cubic (fcc) structures of the Ag bulk state (JCPDS no. 04-0783). In addition, the (200) plane of rGO/AgCo composite nanosheets lied between pure Ag (200) and pure Ni (111) peak, indicating the formation of AgCo alloy nanoparticles. As the redox potential of Ag^+/Ag (0.8 V) is much higher than that of Co^{2+}/Co (-0.28 V), therefore, the reduction rate of Ag^+ is higher than that of Co^{2+} . In the presence of NaBH_4 , Ag is formed in the initial phase followed with sequential deposition of Co. The metallic Co atoms are stabilized by the Ag lattice, resulting in the formation of stable alloy nanoparticles. This result was similar with the previous report on the formation of AgNi alloy nanoparticles.¹⁵

XPS spectroscopy was used to characterize the chemical composition and valence state of the prepared rGO/AgCo composite nanosheets. As shown in Fig. 4a, the survey spectrum clearly displayed C, O, Ag and Co elements in the rGO/AgCo composite nanosheets product. Fig. 4b represented the high resolution XPS spectrum of C 1s, which could be the high resolution XPS spectrum of C 1s, which could be deconvoluted into four peaks centered at 283.4 , 284.7 , 286.6 and 288.8 eV , which are ascribed to C–Co, C–C/C=C, C=O, and O=C–O bands of rGO, respectively. The high resolution XPS spectrum of Ag element exhibited two typical bands at 368.7 and 374.7 eV , which are attributed to the signals for Ag $3d_{5/2}$ and Ag $3d_{3/2}$, respectively, revealing the formation of metallic Ag (Fig. 4c).²⁸

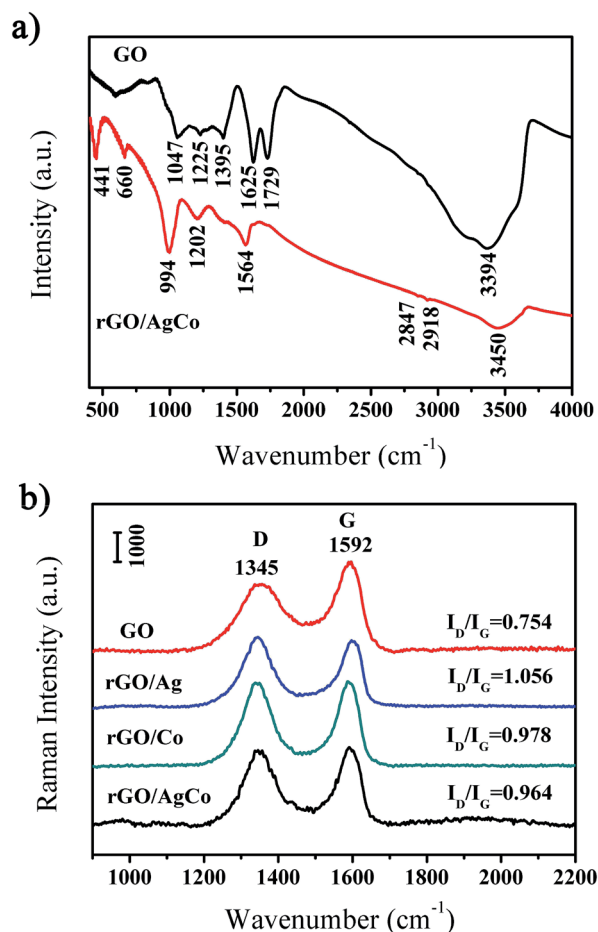


Fig. 2 (a) FTIR spectra of GO and rGO/AgCo composite nanosheet samples; (b) Raman spectra of GO, rGO/Ag, rGO/Co and rGO/AgCo composite nanosheet samples.

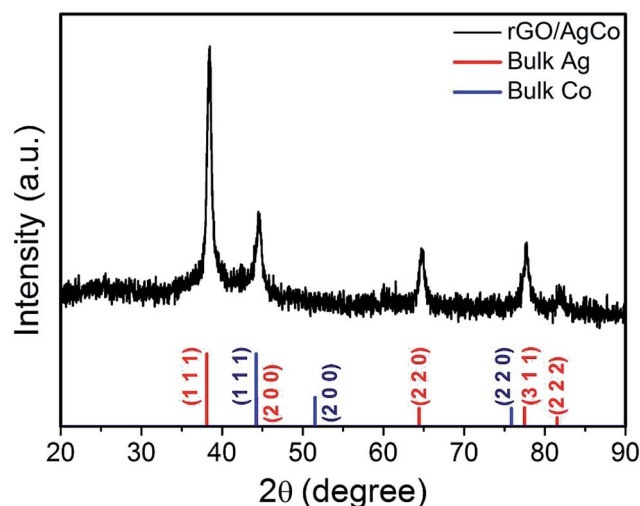


Fig. 3 XRD pattern of the synthesized rGO/AgCo composite nanosheets. The XRD patterns of bulk Ag in red (JCPDS 04-0783) and Co in blue (JCPDS 15-0806) has also shown at the bottom of the image.



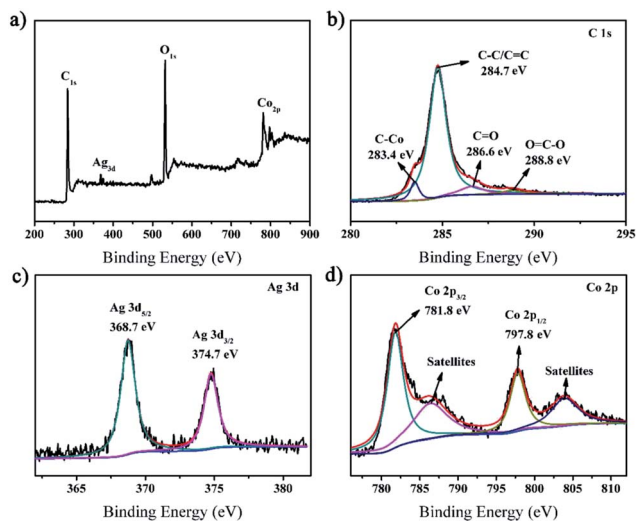


Fig. 4 XPS spectra of the synthesized rGO/AgCo composite nanosheets. (a) Survey spectrum; (b) C 1s; (c) Ag 3d; (d) Co 2p.

From Fig. 4d, it can be observed that four characteristic peaks appeared in the Co 2p spectrum. The first two peaks at about 781.8 and 797.8 eV are related to Co 2p_{3/2} and Co 2p_{1/2}, and the other two peaks with higher binding energies at around 786.4 and 804.0 eV are assigned to the shake-up satellites.²⁹

The as-prepared rGO/AgCo composite nanosheets has been proven to be efficient catalyst toward the reduction of MB in the presence of NaBH₄. The catalytic process can be monitored by time dependent UV-vis spectroscopy. As shown in Fig. S1a,[†] the aqueous solution of MB is blue in color and exhibits a strong absorption at 665 nm. After the addition of NaBH₄, the absorption peak decreases obviously and almost disappears in 8 min, indicating the completely reduction of MB. The reduction of MB by NaBH₄ in the presence of rGO/AgCo composite nanosheets can be explained using Langmuir–Hinshelwood mechanism. The BH₄[−] from NaBH₄ and MB are simultaneously adsorbed on the surface of rGO/AgCo composite nanosheets. After adsorption, the BH₄[−] ions transfer the electrons to MB through rGO/AgCo composite nanosheets. Herein, rGO/AgCo composite nanosheets act as electron transfer mediator, which reduce the energy barrier between the reactant and product. We have also compared the catalytic activity of the as-prepared rGO/AgCo composite nanosheets with that of rGO/Ag and rGO/Co composite nanosheets. It was found that the rGO/AgCo composite nanosheets possessed a higher catalytic activity for the reduction of MB by NaBH₄ than individual rGO/Ag and rGO/Co composite nanosheets. In addition, we have also investigated the influence of Ag/Co molar ratio to the catalytic reduction of MB in the presence of NaBH₄. It is found that the catalytic activity of rGO/AgCo is higher than that of rGO/Ag₃Co₁ and rGO/Ag₁Co₃. The reaction kinetics toward the reduction of MB has also been studied. Fig. 5a showed the linear relationship between $\ln(I_t/I_0)$ and the reaction time for the catalytic reduction of MB by rGO/AgCo composite nanosheets, where in I_t and I_0 stand for the absorption intensities at 665 nm at time t and 0, respectively. From Fig. 5a, it can be clearly observed that

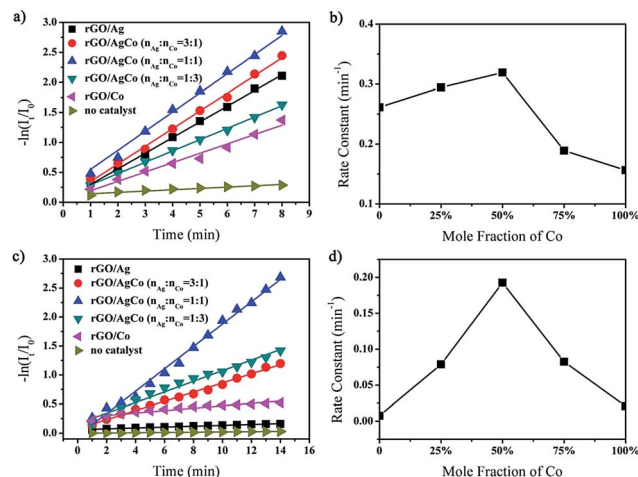


Fig. 5 (a) Plots of $\ln(I_t/I_0)$ versus time for rate constant calculation of the reduction of MB by NaBH₄ in the presence of as-synthesized samples with different Ag/Co molar ratio and blank control; (b) the relationship between the rate constant and the mole fraction of Co for the reduction of MB by NaBH₄; (c) plots of $\ln(I_t/I_0)$ versus time for rate constant calculation of the degradation of MB under UV light in the presence of as-synthesized samples with different Ag/Co molar ratio and blank control; (d) the relationship between the rate constant and the mole fraction of Co for the photodegradation of MB under UV light.

MB molecules are almost not reduced by only NaBH₄ in the absence of rGO/AgCo composite nanosheets. Based on the rate law of the fitting equation of $\ln(I_t/I_0) = 0.22994 + 0.31906t$, the rate constant can be calculated to be about 0.32 min^{−1} for rGO/AgCo composite nanosheets as catalyst, which was higher than that of rGO/Ag (0.26 min^{−1}), rGO/Co (0.15 min^{−1}), rGO/Ag₃Co₁ (0.29 min^{−1}), and rGO/Ag₁Co₃ (0.19 min^{−1}) as catalysts, respectively (Fig. 5b). This result indicates that rGO/AgCo composite nanosheets with 1 : 1 Ag/Co molar ratio possess the highest catalytic activity for reduction of MB by NaBH₄. The as-prepared rGO/AgCo composite nanosheets are also good photocatalyst for the degradation of MB under UV light. As shown in Fig. 5c, the individual rGO/Ag composite nanosheets almost did not show photocatalytic activity toward the degradation of MB under UV light, indicating that rGO/Ag is inert toward the photocatalytic degradation of MB. On the other hand, rGO/Co composite nanosheets exhibit much lower catalytic activity than that of rGO/AgCo composite nanosheets. From Fig. 5d, it is found that the rate constant for the catalytic degradation of MB by rGO/AgCo composite nanosheets under UV light is calculated to be about 0.19 min^{−1}, which is much higher than those of rGO/Ag (0.01 min^{−1}), rGO/Co (0.02 min^{−1}), rGO/Ag₃Co₁ (0.08 min^{−1}), and rGO/Ag₁Co₃ (0.09 min^{−1}) as the photocatalysts. Generally, the catalytic and photocatalytic activities are related to the surface area of the catalysts. Thus we have estimated the surface areas of the prepared catalysts. As shown in Fig. S2,[†] the Brunauer–Emmett–Teller (BET) surface area of rGO/AgCo composite nanosheets is reckoned to be about 99.4 m² g^{−1}, which is higher than that of rGO/Ag (67.3 m² g^{−1}) and rGO/Ag₃Co₁ composite nanosheets (78.9 m² g^{−1}), but lower than that of rGO/Co (101.9 m² g^{−1}) and rGO/Ag₁Co₃ composite



nanosheets ($177.6 \text{ m}^2 \text{ g}^{-1}$). Therefore, the enhanced catalytic and photocatalytic activity of rGO/AgCo composite nanosheets should also be attributed to the synergistic effect between rGO and AgCo alloys. On the one hand, the rGO nanosheets with a high conductivity facilitate a fast electron transfer between substrate molecules and AgCo alloy nanoparticles. In addition, the rGO nanosheets also prevent the aggregation of AgCo alloy nanoparticles. On the other hand, electron transfer from Co to Ag will take place due to their different electronegativities, resulting in a partial charge separation on the CoAg alloy surface. The prepared rGO/AgCo composite nanosheets display a balanced Ag/Co ratio and generate the best charge separation, generating highest catalytic and photocatalytic activities.

The as-prepared rGO/AgCo composite nanosheets are not only excellent catalyst and photocatalyst for the reduction and degradation of MB, but also efficient SERS substrate for the detection of various target molecules. In this work, we have evaluated the SERS properties of rGO/AgCo composite nanosheets using PATP and MBA as model probes. Fig. 6a showed the SERS spectra of PATP molecules with different concentrations from 10^{-3} to 10^{-9} M on the surface of rGO/AgCo composite nanosheets. It can be clearly seen that very strong Raman bands of PATP molecules appear in the SERS spectrum when the concentration of PATP is 10^{-3} M. The characteristic peaks at 1576 and 1072 cm^{-1} are attributed to the C–C and C–S

stretching mode of PATP molecules, which are related to the a_1 modes of PATP (Fig. 6a). In addition, some other typical bands at 1435, 1389, 1187, 1142 cm^{-1} are also clearly observed in the SERS spectrum, which can be assigned to b_2 symmetry of PATP molecules. The strong resonance enhancement of these bands are due to the charge transfer process.³⁰ It is also found that the intensities of the typical a_1 and b_2 bands of PATP molecules decrease with the reducing of their concentrations. However, the typical bands of PATP molecules were still distinguished when the concentration of PATP molecules decreased to 10^{-9} M, suggesting the high SERS sensitivity of rGO/AgCo composite nanosheets. On the other hand, Fig. 6b showed the SERS spectra of MBA molecules with different concentrations from 10^{-5} to 10^{-9} M on the surface of rGO/AgCo composite nanosheets. All Raman peaks were characteristic SERS signals of MBA molecules. For instance, the strong band at 1582 cm^{-1} is assigned to the aromatic $\nu(\text{C}=\text{C})$ mode, and the typical peak at 1073 cm^{-1} is ascribed to the in-plane ring breathing mode coupled with $\nu(\text{C}=\text{S})$. The relative weak band at 1182 cm^{-1} is attributed to the C–H deformation mode.^{31,32} Similarly, the intensities of the characteristic bands of MBA molecules decrease with the decreasing of their concentrations, however, the typical bands at 1582 cm^{-1} are still observed when the concentration of MBA decreased to 10^{-9} M. These results demonstrated that the as-prepared rGO/AgCo composite nanosheets are efficient SERS substrates. Furthermore, we have also investigated the SERS sensitivity of the rGO/AgCo composite nanosheets that synthesized in the solution with different pH values. As shown in Fig. S3a,† the rGO/AgCo composite nanosheets that synthesized in a solution with pH value of 9.0 show the highest SERS intensity compared to other substrates using PATP molecules as probes. The similar results have also been obtained by using MBA molecules as probes. Therefore, the pH = 9.0 is an optimal condition for the preparation of rGO/AgCo composite nanosheets as SERS substrates.

Owing to the excellent catalytic activities and efficient SERS properties, the as-prepared rGO/AgCo composite nanosheets can be used as SERS substrate for *in situ* monitoring the catalytic reaction process and determine the reaction kinetics of the reduction of MB by NaBH_4 . Fig. 7a shows the SERS spectra of MB molecules on the surface of rGO/AgCo composite nanosheets in the presence of NaBH_4 at different time intervals. It was clearly observed that three typical peaks at 1623, 1475, and 1400 cm^{-1}

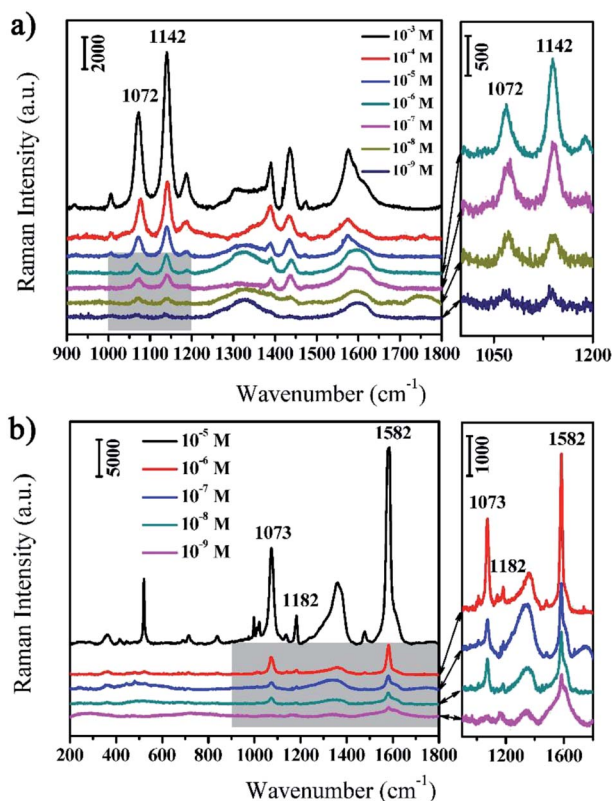


Fig. 6 (a) SERS spectra of PATP molecules with different concentrations (10^{-5} – 10^{-9} M) on the surface of rGO/AgCo composite nanosheets; (b) SERS spectra of MBA molecules with different concentrations (10^{-5} – 10^{-9} M) on the surface of rGO/AgCo composite nanosheets.

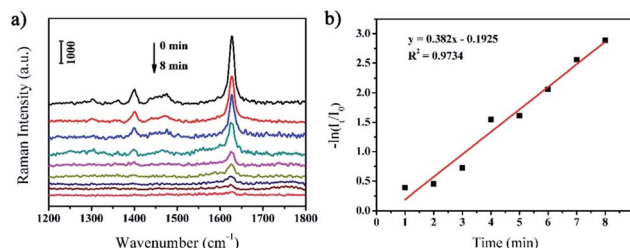


Fig. 7 (a) The change of SERS spectra for the catalytic reduction of MB by NaBH_4 on the surface of rGO/AgCo composite nanosheets. (b) The curve of the linear relationship between $\ln(I_t/I_0)$ and the reaction time according to the bands at 1623 cm^{-1} .



appeared in the SERS spectrum before the addition of NaBH_4 in the system. These three peaks are ascribed to C–C ring stretching, C–N stretching and N–C–H in-plane bending vibrations, respectively.³³ However, the SERS intensity of the three peaks of MB gradually decreases with increasing time after the addition of NaBH_4 . And these peaks almost disappear in 8 min, demonstrating completely reduction of MB on the surface of rGO/AgCo composite nanosheets. The catalytic reduction of MB by rGO/AgCo composite nanosheets can be indexed by a Langmuir–Hinshelwood model. Fig. 7b shows the linear relationship between $\ln(I_t/I_0)$ and the reaction time toward the catalytic reduction of MB by rGO/AgCo composite nanosheets, wherein I_t and I_0 stand for the SERS intensities at 1623 cm^{-1} at time t and 0, respectively. Then a fitting equation of $\ln(I_t/I_0) = -0.1925 + 0.382t$ was obtained, which showed the pseudo-first order reaction. The rate constant of the catalytic reaction can be calculated to be about 0.382 min^{-1} . In addition, the rate constant at the beginning of the catalytic reaction has also been studied. As shown in Fig. S4a,[†] the rate constant is about 0.0062 s^{-1} , which is much lower than the rate constant calculated from the whole reaction, indicating an induction period for the catalytic reaction.

Similar with the *in situ* monitoring the catalytic degradation of MB by NaBH_4 , the as-prepared rGO/AgCo composite nanosheets can also be used as SERS substrate for monitoring of the photocatalytic degradation of MB under UV light. Fig. 8a shows the SERS spectra of MB molecules on the surface of rGO/AgCo composite nanosheets under UV light at different time intervals. Similar with the Fig. 7a, it is clearly seen that the intensity of the characteristic peaks of MB molecules gradually decreased with increasing time under UV light, and these peaks almost disappeared in 13 min. The linear relationship between $\ln(I_t/I_0)$ and the reaction time toward the photocatalytic degradation of MB by rGO/AgCo composite nanosheets under UV light has also been shown in Fig. 8b. The rate law can be expressed as a fitting equation of $\ln(I_t/I_0) = -0.4304 + 0.2115t$. Based on this equation, the photocatalytic degradation of MB follows a pseudo-first order reaction. And the rate constant of the photocatalytic reaction can be calculated to be about 0.2115 min^{-1} . The rate constant at the beginning of the photocatalytic reaction has also been studied. As shown in Fig. S4b,[†] the rate constant is about 0.0007 s^{-1} , which is much lower than the rate constant calculated from the whole reaction, demonstrating an induction period for the photocatalytic reaction.

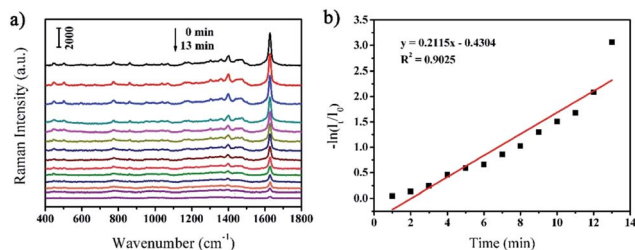


Fig. 8 (a) The change of SERS spectra for the catalytic degradation of MB under UV light on the surface of rGO/AgCo composite nanosheets. (b) The curve of the linear relationship between $\ln(I_t/I_0)$ and the reaction time according to the bands at 1623 cm^{-1} .

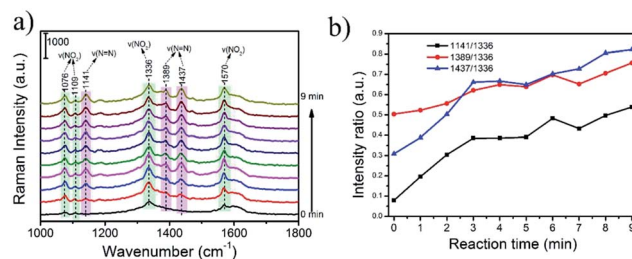


Fig. 9 (a) Time-dependent SERS spectra of plasmon-enhanced driven catalytic reaction of PNTP dimerizing into DMAB on the surface of rGO/AgCo composite nanosheets; (b) the curve of the relationship between relative SERS intensity and reaction time.

In this work, we have also performed the prepared rGO/AgCo composite nanosheets as SERS substrate for *in situ* monitoring of the plasmon-driven catalytic reaction of PNTP dimerizing into DMAB. As shown in Fig. 9a and b, an excellent SERS activities of PNTP has been observed on the surface of the rGO/AgCo composite nanosheets. At the beginning SERS measurement, the characteristic peaks at 1336 and 1570 cm^{-1} are clearly observed, which are attributed to the $-\text{NO}_2$ stretching mode of PNTP.³⁴ After continuous laser excitation, some new bands at 1141 , 1389 and 1437 cm^{-1} gradually emerged. And the intensity of these peaks increase with the increasing reaction time. These three typical peaks are ascribed to C–N symmetric stretching, N=N stretching, and C–H in-plane bending modes of DMAB, respectively. This result suggests that PNTP has been dimerized into DMAB due to the plasmon driven catalytic reaction. However, the intensities of the peaks at 1336 and 1570 cm^{-1} did not obviously decrease, indicating that only a part of PNTP molecules were dimerized into DMAB, which is similar with the previous reports.^{34,35} We proposed a possible mechanism for the dimerization of PNTP into DMAB (Fig. S5[†]). The dimerization process requires $4e^-$, which can be produced from hot electrons arising from surface plasmon.^{35–39} Therefore, the conversion of PNTP into DMAB can be attributed to the reduction process by the hot electrons that are generated by surface plasmon decay on the surface of rGO/AgCo composite nanosheets. The hot electrons generated by plasmon decay possess high kinetic energy, which can drive the surface-catalyzed reaction.

4. Conclusions

In summary, we have demonstrated a facile co-reduction approach to prepare rGO/AgCo composite nanosheets. The prepared AgCo nanoparticles are uniformly distributed on the surface of rGO nanosheets. The synthesized rGO/AgCo composite nanosheets have been employed as catalyst and photocatalyst toward the reduction and degradation of MB and show superior catalytic and photocatalytic activities over individual rGO/Ag and rGO/Co composite nanosheets. The rGO/AgCo composite nanosheets are also used as efficient SERS substrate for monitoring the catalytic and photocatalytic reaction toward the reduction and degradation of MB, and the plasmon-driven of PNTP dimerizing into DMAB. It is



anticipated that this study offers a facile approach for the fabrication of multi-functional nanocomposites for a large variety of applications in the fields of catalysts, biosensing, and environmental science and technology.

Conflicts of interest

There are no conflicts to declare.

Acknowledgements

This work was supported by the research grants from the National Natural Science Foundation of China (21473068, 21327803).

References

- 1 Y. Xia, Y. Xiong, B. Lim and S. E. Skrabalak, *Angew. Chem., Int. Ed.*, 2009, **48**, 60–103.
- 2 Z. Quan, Y. Wang and J. Fang, *Acc. Chem. Res.*, 2013, **46**, 191–202.
- 3 S. Gong and W. Cheng, *Adv. Electron. Mater.*, 2017, **3**, 1600314.
- 4 T. A. Taton, C. A. Mirkin and R. L. Letsinger, *Science*, 2000, **289**, 1757–1760.
- 5 Y. W. Jun, J. W. Seo and A. Cheon, *Acc. Chem. Res.*, 2008, **41**, 179–189.
- 6 K. D. Gilroy, A. Ruditskiy, H. C. Peng, D. Qin and Y. N. Xia, *Chem. Rev.*, 2016, **116**, 10414–10472.
- 7 B. Lim, M. Jiang, P. H. C. Camargo, E. C. Cho, J. Tao, X. Lu, Y. Zhu and Y. Xia, *Science*, 2009, **324**, 1302–1305.
- 8 L. Ai, X. Liu and J. Jiang, *J. Alloys Compd.*, 2015, **625**, 164–170.
- 9 Z. Qu, D. Chen, Y. Sun and Y. Wang, *Appl. Catal., A*, 2014, **487**, 100–109.
- 10 Y. Tong, X. Lu, W. Sun, G. Nie, L. Yang and C. Wang, *J. Power Sources*, 2014, **261**, 221–226.
- 11 Y. Cheng, Y. Fan, Y. Pei and M. Qiao, *Catal. Sci. Technol.*, 2015, **5**, 3903–3916.
- 12 Q. J. Xiang, J. G. Yu and M. Jaroniec, *Chem. Soc. Rev.*, 2012, **41**, 782–796.
- 13 E. Antolini, *Appl. Catal., B*, 2012, **123–124**, 52–68.
- 14 J. Lv, A. Wang, X. Ma, R. Xiang, J. Chen and J. Feng, *J. Mater. Chem. A*, 2015, **3**, 290–295.
- 15 R. Dhanda and M. Kidwai, *J. Mater. Chem. A*, 2015, **3**, 19563–19574.
- 16 L. Yang, J. Su, X. Meng, W. Luo and G. Cheng, *J. Mater. Chem. A*, 2013, **1**, 10016–10023.
- 17 X. Ling, L. M. Xie, Y. Fang, H. Xu, H. L. Zhang, J. Kong, M. S. Dresselhaus, J. Zhang and Z. F. Liu, *Nano Lett.*, 2010, **10**, 553–561.
- 18 W. G. Xu, N. N. Mao and J. Zhang, *Small*, 2013, **9**, 1206–1224.
- 19 Z. Zhang, F. G. Xu, W. S. Yang, M. Y. Gao, X. D. Wang, B. L. Zhang and J. L. Tang, *Chem. Commun.*, 2011, **47**, 6440–6442.
- 20 J. Lee, S. Shim, B. Kim and H. S. Shin, *Chem.–Eur. J.*, 2011, **17**, 2381–2387.
- 21 G. Lu, H. Li, C. Liusman, Z. Y. Yin, S. X. Wu and H. Zhang, *Chem. Sci.*, 2011, **2**, 1817–1821.
- 22 C. Y. Wen, F. Liao, S. S. Liu, Y. Zhao, Z. H. Kang, X. L. Zhang and M. W. Shao, *Chem. Commun.*, 2013, **49**, 3049–3051.
- 23 Y. Guo, H. Wang, X. W. Ma, J. Jin, W. Ji, X. Wang, W. Song, B. Zhao and C. Y. He, *ACS Appl. Mater. Interfaces*, 2017, **9**, 19074–19081.
- 24 W. S. Hummers and R. E. Offeman, *J. Am. Chem. Soc.*, 1958, **80**, 1339.
- 25 E. Jin, X. F. Lu, L. L. Cui, D. M. Chao and C. Wang, *Electrochim. Acta*, 2010, **55**, 7230–7234.
- 26 X. Lu, L. Yang, X. Bian, D. Chao and C. Wang, *Part. Part. Syst. Charact.*, 2014, **31**, 245–251.
- 27 Z. Ji, X. Shen, G. Zhu, H. Zhou and A. Yuan, *J. Mater. Chem.*, 2012, **22**, 3471–3477.
- 28 W. Song, W. Ji, S. Vantasin, I. Tanabe, B. Zhao and Y. Ozaki, *J. Mater. Chem. A*, 2015, **3**, 13556–13562.
- 29 Z. Z. Yang, Z. Zhang, Y. Z. Jiang, M. Q. Chi, G. D. Nie, X. F. Lu and C. Wang, *RSC Adv.*, 2016, **6**, 33636–33642.
- 30 Z. Mao, W. Song, X. X. Xue, W. Ji, Z. S. Li, L. Chen, H. J. Mao, H. M. Lv, X. Wang, J. R. Lombardi and B. Zhao, *J. Phys. Chem. C*, 2012, **116**, 14701–14710.
- 31 Y. Wang, W. Ji, H. M. Sui, Y. Kitahama, W. D. Ruan, Y. Ozaki and B. Zhao, *J. Phys. Chem. C*, 2014, **118**, 10191–10197.
- 32 M. D. Gong, X. Jiang, J. Du, X. L. Li, X. X. Han, L. B. Yang and B. Zhao, *RSC Adv.*, 2015, **5**, 80269–80275.
- 33 S. Dutta Roy, M. Ghosh and J. Chowdhury, *J. Raman Spectrosc.*, 2015, **46**, 451–461.
- 34 J. Jin, S. J. Zhu, Y. B. Song, H. Y. Zhao, Z. Zhang, Y. Guo, J. B. Li, W. Song, B. Yang and B. Zhao, *ACS Appl. Mater. Interfaces*, 2016, **8**, 27956–27965.
- 35 P. Li, B. B. Ma, L. B. Yang and J. H. Liu, *Chem. Commun.*, 2015, **51**, 11394–11397.
- 36 M. W. Knight, H. Sobhani, P. Nordlander and N. J. Halas, *Science*, 2011, **332**, 702–704.
- 37 B. Dong, Y. Fang, X. Chen, H. Xu and M. Sun, *Langmuir*, 2011, **27**, 10677–10682.
- 38 M. Sun and H. Xu, *Small*, 2012, **8**, 2777–2786.
- 39 W. Song, C. J. Querebillo, R. Götz, S. Katz, U. Kuhlmann, U. Gernert, I. M. Weidinger and P. Hildebrandt, *Nanoscale*, 2017, **9**, 8380–8387.

