


 Cite this: *RSC Adv.*, 2017, 7, 42289

Dependence of electrochemical properties of spinel LiMn_2O_4 on Li_2CO_3 with micro-flaky, micro-flower and nanorod morphologies†

 Lang Li, Jinsong Sui, Rui Huang, Wei Xiang and Wei Qin *

Herein, the dependence of spinel LiMn_2O_4 on Li_2CO_3 with micro-flaky, micro-flower and nanorod morphologies is investigated. The results show that the as-synthesized LiMn_2O_4 with micron sized Li_2CO_3 as raw materials have a much higher discharge capacity than that of the one prepared with nano sized Li_2CO_3 . It delivers an initial charge capacity of 110.1, 105.2 and 104.9 mA h g^{-1} followed by a discharge capacity of 109.1, 103.9 and 104.2 mA h g^{-1} with the micro-flower, nanorod and micro-flaky Li_2CO_3 morphologies, respectively, at room temperature (about 99% of the charge capacity is discharged). The smaller specific surface area is found in the spinel LiMn_2O_4 with micron sized Li_2CO_3 , resulting in a better stable electrochemical performance in LiMn_2O_4 with micro-flower and micro-flaky Li_2CO_3 . Their capacities are maintained at 99.2 mA h g^{-1} and 94.2 mA h g^{-1} after 100 cycles at 1C rate. The capacity retention was more than 90% at the 100th cycle with the micron-sized Li_2CO_3 . Moreover, the as-synthesized spinel LiMn_2O_4 with micro-flower Li_2CO_3 retained more than 95% of its initial discharge capacity (92 mA h g^{-1}) after 200 cycles at 2C rate. The cubic spinel structure was detected after 200 cycles of LiMn_2O_4 at 2C rate.

 Received 26th June 2017
Accepted 9th August 2017

DOI: 10.1039/c7ra07088h

rsc.li/rsc-advances

Introduction

Rechargeable batteries are becoming increasingly important recently, particularly in consumer electronic devices such as cellular telephone, notebook computers, compact camcorders, and electronic vehicles.^{1–4} Spinel LiMn_2O_4 is regarded as one of the most promising cathode materials for lithium ion batteries⁵ by virtue of its obvious advantages such as abundant and cheap resources, environmental friendliness, safe handling and low toxicity.^{6,7} However, the disadvantages of structural instability, poor cycle efficiency and low energy density have held back its application for large scale production. In general, as for the battery application, LiMn_2O_4 powder should possess a single-phase, homogeneity, uniform particle morphology with a submicron size distribution, and a large surface area in order to achieve better electrode properties. The factors such as crystallinity, crystal structure, the deviation from theoretical stoichiometry of elemental composition as well as the grain size and its distribution play important and eventually, decisive roles in the electrochemical performance of spinel LiMn_2O_4 . The compound was typically obtained by reaction of a mixture of a lithium salt and manganese oxides subjected to high temperature in air for several hours.

However, the preparation of spinel LiMn_2O_4 involves the solid-state reactions starting from the raw materials, such as manganese oxides, nitrate or carbonate with lithium hydroxide, nitrate or carbonate at elevated temperatures as high as 700–900 °C. The final products usually contain larger irregular particles, with a broad size distribution, as well as impurity phases. Furthermore, this method could not provide a good control on the crystalline growth, compositional homogeneity, morphology, and microstructure. To overcome these problems, the purity and particle size distribution of Li_2CO_3 as the cathode material were optimized. Moreover, the reaction process of Li_2CO_3 and MnO_2 was studied using thermogravimetry (TG), derivative thermogravimetry (DTG) and differential scanning calorimetry (DSC). A uniform cubic nanostructured spinel LiMn_2O_4 fused together was obtained. The most suitable calcination temperature was programmed using an oven.

Recently, many studies have been conducted to improve the performance of rechargeable lithium ion battery by using cationic doping^{8,9} and the nanostructured electrode^{10,11} that have shown to enhance power performance due to the large surface-to-volume ratio that allows for a large electrode–electrolyte contact area. Nanocrystal,¹² nanoporous,¹³ nanorod,¹⁴ nanotube,^{15–17} nanowire,^{18,19} and nanotube active materials are particularly attractive. The disadvantage of nanosized materials is that they cannot be packed as dense as the micron-sized materials on the current collector, since the electrodes made of nanosized materials exhibit a high porosity, thus resulting in a decrease in the cell capacity. Therefore, the best way to

Department of Chemical Engineering, State Key Laboratory of Chemical Engineering, Tsinghua University, Beijing 100084, China

† Electronic supplementary information (ESI) available. See DOI: 10.1039/c7ra07088h



improve both the rate capacity and the electrode density would be to use the micron-sized particles that consist of aggregated nanoparticles.²⁰ Most of the nanostructured electrode materials are synthesized by low-temperature treatment processes such as soft chemical,²¹ sol-gel,^{22,23} and hydrothermal methods.²⁴ However, little attention has been paid to the effect of the purity and the morphology of the cathode material Li_2CO_3 .²⁵ In this paper, the cubic uniform nanostructured spinel LiMn_2O_4 fused together was synthesized by the solid state reaction with Li_2CO_3 in micro-flaky, micro-flower and nanorod forms. Moreover, the electrochemical characteristics of the as-synthesized spinel LiMn_2O_4 were compared with each other to determine the most suitable morphology and size of the raw materials.

Here, a novel strategic synthesis process of a battery grade lithium carbonate was carried out in a homogeneous-like organic phase.²⁶ Three types of crystal morphology were prepared: micro-flower, micro-flaky, and nanorod. The nanostructured spinel LiMn_2O_4 was then synthesized by the solid-state reaction with MnO_2 (EMD) with a calcination program at high temperature. The electrochemical performances of the prepared spinel LiMn_2O_4 were discussed. The result showed that the nanostructured spinel LiMn_2O_4 prepared with the micron sized Li_2CO_3 as the raw material had the best performance and maintained more than 90% of its initial capacity rate after 100 charge–discharge cycles at 1C rate. The as-synthesized LiMn_2O_4 delivered an initial discharge capacity of 109.1 and 104.2 mA h g^{-1} at 1C rate with the flower and flaky morphologies, respectively. Moreover, the as-synthesized spinel LiMn_2O_4 with micro-flower Li_2CO_3 exhibited an excellent cycling capability, maintaining 95% of its initial capacity after 200 cycles at a charge–discharge rate of 2C.

Experimental section

Li_2CO_3 synthesis

Battery grade Li_2CO_3 was synthesized through the reaction coupled with a solvent extraction process. First, LiCl (>97% pure, Beijing Yili, Fine, Chemical, Co., Beijing, China) was dissolved in deionized water. Then, 50 mL LiCl aqueous solution and 200 mL *n*-butanol (>97%, pure, Beijing, Chemical Works, Beijing, China), that served as the extracting agent and the diluent were added into a 500 mL separating funnel. The suspension was shaken for 15 min until the water phase and the alcohol phase reached equilibrium. Further, the upper alcohol phase was removed and added together with 100 mL tri-*n*-octyl amine (Alamine 336 – Sinopec, Beijing, China), which served as the extractant for HCl , to a three-necked flask. The reaction mixture was stirred at 150 rpm for approximately 45 min and then CO_2 (>95% pure, Tsinghua University, Beijing, China) was bubbled into the reactor at a pressure of 0.1 Mpa. The crystals obtained were then collected by filtration, washed with alcohol, and dried.

Synthesis of LiMn_2O_4

The as-synthesized Li_2CO_3 (0.50 g) along with 2.32 g MnO_2 were ground with high purity ethanol to produce a well-mixed

powder. The powder was then calcinated by a two-step calcination at 650 °C and 800 °C for 5 hours with the heating rate at 3 °C min^{-1} in air to produce spinel LiMn_2O_4 .

Characterization

The prepared Li_2CO_3 was mixed and reacted with MnO_2 (EMD, Xiangtan, Hunan) (molar ratio, $n(\text{Li}) : n(\text{Mn}) = 1.01 : 2$). The nanostructured spinel LiMn_2O_4 was prepared by the solid-state combustion synthesis which is a two-step calcination process at 650 °C and 800 °C with the heating rate at 3 °C min^{-1} in air for 5 hours. The spinel LiMn_2O_4 formation process was studied by thermogravimetry (TG), derivative thermogravimetry (DTG) and differential scanning calorimetry (DSC). The crystal structure and morphology were identified by using X-ray diffraction analysis (XRD; D8-Advance, Bruker, Germany), scanning electron microscope (SEM; JSM7401, Jeol, Japan), and Tabletop scanning electron microscope (TM3000, Hitachi, Japan). Particle size distribution was measured by Malvern particle size analyzer (Mastersizer2000, Malvern, England). The Li concentration was analysed by using atomic absorption spectrophotometer (Z-5000-AAS, Hitachi, Tokyo, Japan). The purity of the as-synthesized Li_2CO_3 was determined by ICP-AES (Varian, 700-ES, USA).

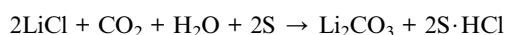
Electrochemical measurement

Electrochemical measurements were performed by the coin cell which was prepared in an argon atmosphere inside a Glove Box (MIKROUNA, GER). The electrodes for electrochemical studies were prepared by making a slurry containing 85 wt% active material of LiMn_2O_4 , 10 wt% conducting carbon black, and 5 wt% polyvinylidene fluoride (PVDF) binder in *N*-methyl-2-pyrrolidone (NMP) as a solvent. The slurry was applied onto an etched aluminum foil current collector using a doctor-blade and dried at 130 °C for 12 h in a vacuum oven. The coated cathode foil was then pressed to form a uniform layer and cut into a square sheet. Lithium metal foil was used as the anode. A 1 M solution of LiPF_6 in ethylene carbonate/dimethyl carbonate (EC/DMC, 1 : 1 v/v) was used as the electrolyte with a Cegrad 2400 membrane as a separator. The charge–discharge cycles were performed at different C rates between 3.0–4.3 V at room temperature using LAND battery testers.

Result and discussion

Synthesis, structure and morphology of Li_2CO_3

The synthesis of Li_2CO_3 is described in the Experimental section. The reaction process was coupled with the solvent extraction process, which is a new approach for producing Li_2CO_3 from LiCl and CO_2 . The equation for the overall reaction as follows:



where S is referred to the extractant alamine 336. Three types of morphologies of Li_2CO_3 were synthesized by changing the initial concentration of LiCl . The flaky, flower-like, and nanorod

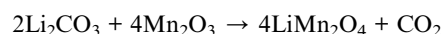
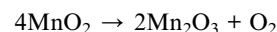


particles were prepared with the initial concentration of LiCl at 2, 3, 4, and 5 mol L⁻¹ (Fig. 1a–d, respectively). The purity of the as-synthesized Li₂CO₃ was tested by ICP-MS with the mass fraction higher than 99.95% (see Table S1 in the ESI†). By comparing the purity of commercial Li₂CO₃, the as-synthesized Li₂CO₃ has purity higher than that of the required battery Li₂CO₃ (99.5%). With such high purity of raw material, the as-synthesized LiMn₂O₄ also exhibited a higher purity, which can improve the electrochemical performance of all sides of the surface of the resultant cathode. SEM images showed that the higher was the concentration of LiCl solution, the smaller was the particle size obtained. Thus, at a low concentration of LiCl, the synthesized crystals typically consisted of the flaky and the flower-like morphologies with micron size, while the nanorod particles were obtained with a high concentration of LiCl solution. Particle size distribution showed that the crystals were in micron and nanometer range with an approximate size of 8 ± 1 μm and 80 ± 10 nm (Fig. 1f), which are consistent with SEM images. The XRD patterns of the reaction-extraction of as-synthesized Li₂CO₃ correspond to JCPDS data no. 87-0729, and no additional impurity peaks were detected (Fig. 1e). Three different crystal morphologies are involved in two types of crystallization mechanism (see the explanation in the ESI†).

Synthesis, structure and morphology of LiMn₂O₄

Spinel LiMn₂O₄ was synthesized by a traditional solid-state calcination reaction by using the prepared Li₂CO₃ of three different morphologies as the raw materials. To optimize the structure, morphology and size of the as-synthesized spinel

LiMn₂O₄, the formation process was studied by thermogravimetry (TG), derivative thermogravimetry (DTG) and differential scanning calorimetry (DSC) (Fig. 2a). The weight loss of the mixtures occurred at two discrete regions of 30–650 °C and 700–800 °C and terminated at 800 °C. The weight loss at the temperature range of 30–650 °C is attributed to the loss of water followed by the decomposition of MnO₂, whose exothermic peak concentrated at 450–550 °C in the DTG curve. The weight loss at the temperature range of 700–800 °C and a large exothermic peak at 800 °C in the DTG curve can be considered as the result of the formation of spinel LiMn₂O₄. Two exothermic peaks were observed in the DSC curve at the temperature of 480 °C and 820 °C. The possible chemical reactions occurring in the synthesis of spinel LiMn₂O₄ was as follows:



The calcination procedure was programmed following a two-step calcination at 650 °C and 800 °C for 5 hours with the heating rate at 3 °C min⁻¹ in air.

Fig. 2b shows the XRD patterns of the products. All the samples were synthesized under the same reaction conditions. The LiMn₂O₄ XRD diffraction diagram shows the features of the spinel structure with *Fd3m* space group (JCPDS card no. 35-0782), with no peaks of the Li₂CO₃ phase detected. The characteristic peaks of planes such as [111], [311], [222], [400], [331], [511], [440] and [531] can be clearly identified. Thus, the

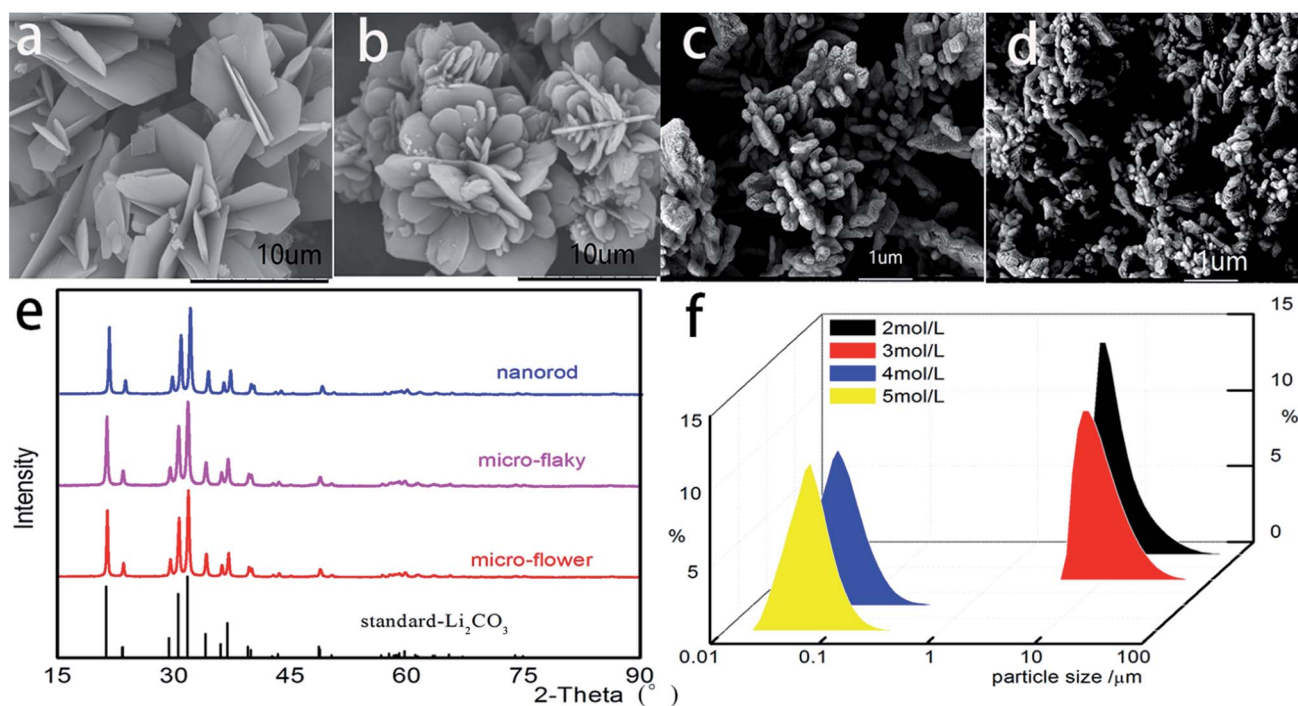


Fig. 1 SEM images, XRD pattern and particle-size distribution of the as-synthesized Li₂CO₃. (a) Micro-flaky, (b) micro-flower and (c, d) nanorod Li₂CO₃ with the initial concentration of LiCl at 2, 3, 4 and 5 mol L⁻¹, respectively, (e) XRD pattern, and (f) particle size distribution of the as-synthesized Li₂CO₃.



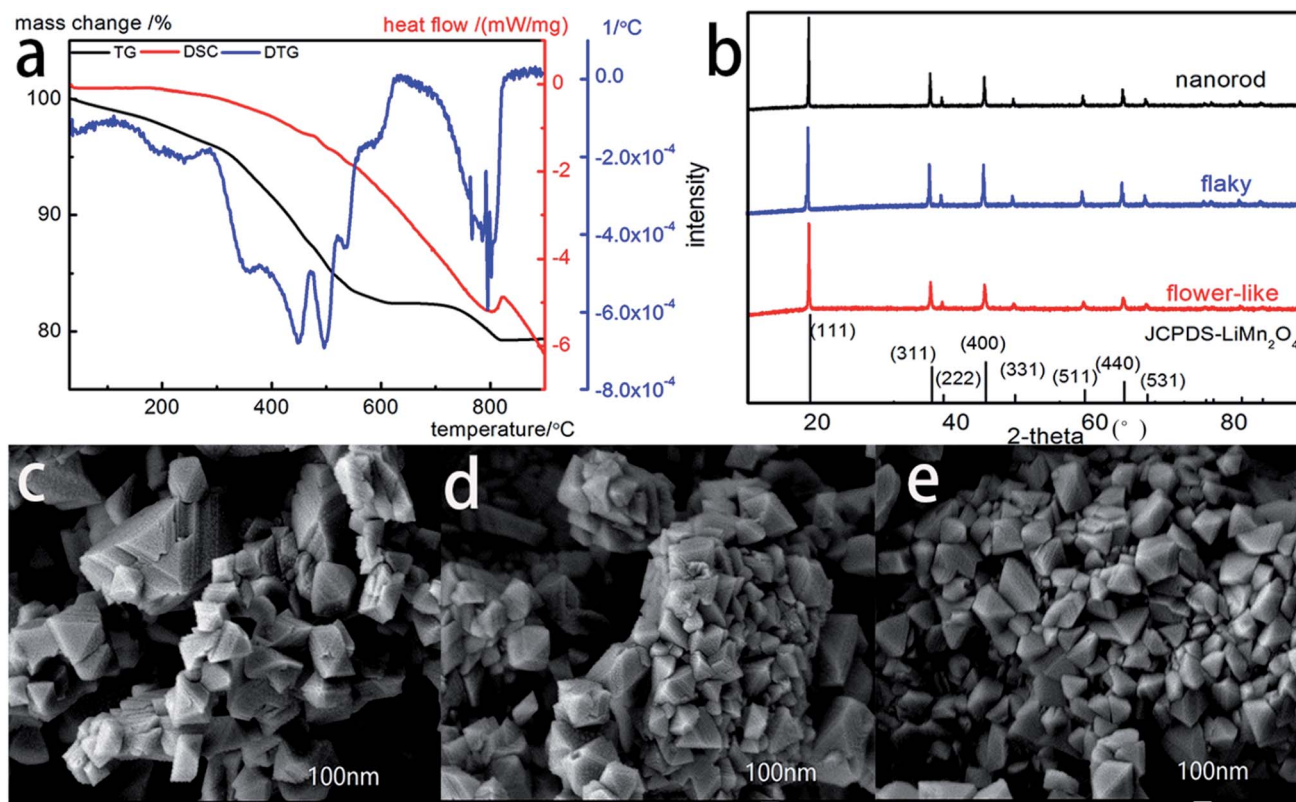


Fig. 2 (a) TG, DTG and DSC curves of the mixtures of Li_2CO_3 and MnO_2 with the ratio of $n(\text{Li})/n(\text{Mn}) = 1.01 : 2$, (b) XRD patterns, and (c–e) SEM images of the as-synthesized spinel LiMn_2O_4 with the cathode material Li_2CO_3 in micro-flaky, micro-flower, and nanorod morphologies, respectively.

reaction between MnO_2 and Li_2CO_3 produced the pure LiMn_2O_4 phase. To analyze the morphological differences of LiMn_2O_4 prepared with different morphologies of Li_2CO_3 , the SEM method was applied. Fig. 2c–e show that the synthesized products are highly crystalline and exhibit spinel structure along with the three types of Li_2CO_3 shown in Fig. 1a–c, respectively. It can be seen that the spinel structure displayed in Fig. 2c and d is more uniform than that shown in Fig. 2e. The distinctive, well-defined crystal faces exhibit the single spinel crystals of octahedral shape with well-developed planes. The as-synthesized spinel LiMn_2O_4 has an average size of 100 nm.

Electrochemical performance

The electrochemical measurements of three types of cathode materials were performed on the coin cell. The as-synthesized LiMn_2O_4 with flaky, flower-like and nanorod Li_2CO_3 forms is abbreviated as E1, E2, and E3, respectively. The galvanostatic charge/discharge tests were performed at room temperature and shown in Fig. 3a. The first charge and discharge curves were carried out at a current density of 1C. It delivered an initial charge capacity of 110.1, 105.2 and 104.9 mA h g^{-1} followed by a discharge capacity of 109.1, 103.9 and 104.2 mA h g^{-1} with Li_2CO_3 in micro-flower, nanorod and micro-flaky morphologies, respectively, at room temperature (about 99% of the charge capacity is discharged). The discharge specific capacity of the three types of cathode materials, which were prepared from the

same preparation route, as a function of cycle number has been compared with each other (Fig. 3b). The charge and discharge capacity was measured with the power rate at 1C (148 mA g^{-1}). It was observed that E1 and E2 had a much higher charge capacity than E3. Moreover, E1 and E2 showed better stability since their capacities were maintained at 94.2 mA h g^{-1} and 99.2 mA h g^{-1} after 100 cycles at the 1C rate, respectively. On the other hand, the discharge capacity decreased to 28.2 mA h g^{-1} in E3. The capacity retention was more than 90% at 100th cycle with the micron-sized Li_2CO_3 .

To better understand the correlation between the structure and the electrochemical performance of the spinel LiMn_2O_4 prepared with the raw material of micron-sized Li_2CO_3 , the TEM images and the ED patterns of LiMn_2O_4 E2 were analyzed. The results are shown in Fig. 4. TEM images and diffraction patterns showed that spinel LiMn_2O_4 is a cubic single crystal. To obtain excellent battery performance, the surface of the active materials is a very important factor. Based on the facet plane, one can conclude that the surface of this spinel LiMn_2O_4 is very pure and clean. The results showed that it is advantageous to obtain uniform cubic spinel nanostructured LiMn_2O_4 with micron-sized Li_2CO_3 as the raw material. The excellent cycle performance is possibly because the spinel LiMn_2O_4 prepared with the micron sized Li_2CO_3 and MnO_2 has a better particle size distribution and uniform morphology. Particle size distribution of the micron-sized Li_2CO_3 raw material was limited to a narrow



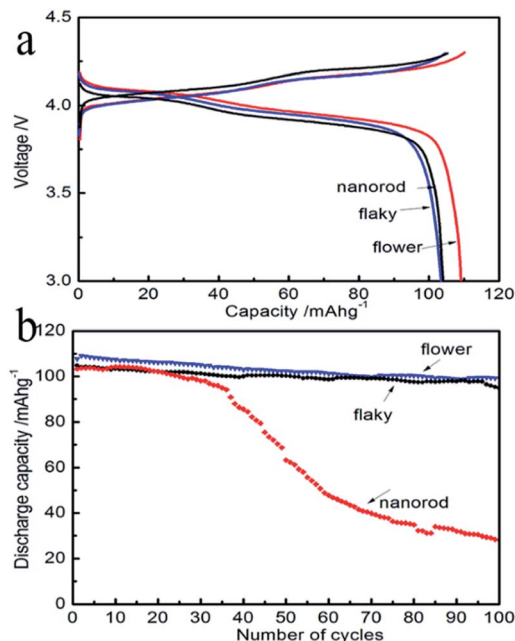


Fig. 3 (a) The charge and discharge curves at 1C rate of the as-synthesized spinel LiMn₂O₄ with the raw material Li₂CO₃ in micro-flower, micro-flaky and nanorod. (b) The discharge capacity of curves of the as-synthesized spinel LiMn₂O₄ with the cathode material Li₂CO₃ in micro-flower, micro-flaky and nanorod vs. number of cycles at 1C rate.

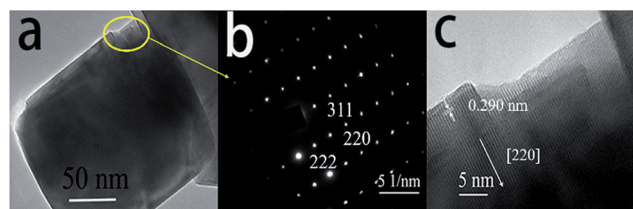


Fig. 4 (a and b) Low and high resolution TEM of LiMn₂O₄ with the cathode material Li₂CO₃ in micro-flower morphology, and (c) the electron diffraction pattern of LiMn₂O₄.

micron range; thus, it resulted in the uniform cubic spinel morphology of LiMn₂O₄. They were fused together forming a polyhedral structure which made the Li⁺ entering the surface of the spinel diffuse through the huge inner structure. Moreover, the high purity of the as-synthesized Li₂CO₃ also plays a key positive role on the electrochemical performance of the as-synthesized LiMn₂O₄. The TEM images and the electron diffraction patterns also confirmed the cubic spinel structure of the as-synthesized LiMn₂O₄.

The specific surface area has a significant effect on the electrochemical performance of the as-synthesized spinel LiMn₂O₄. Thus, the specific surface areas of the as-synthesized Li₂CO₃ and LiMn₂O₄ were analyzed. The results showed that the specific surface area of nanosized Li₂CO₃ was larger than that of the micron-sized Li₂CO₃ (Fig. 5a), which indicates that the particle size of micron-sized Li₂CO₃ was larger than that of nano-sized Li₂CO₃. However, the as-synthesized spinel LiMn₂O₄

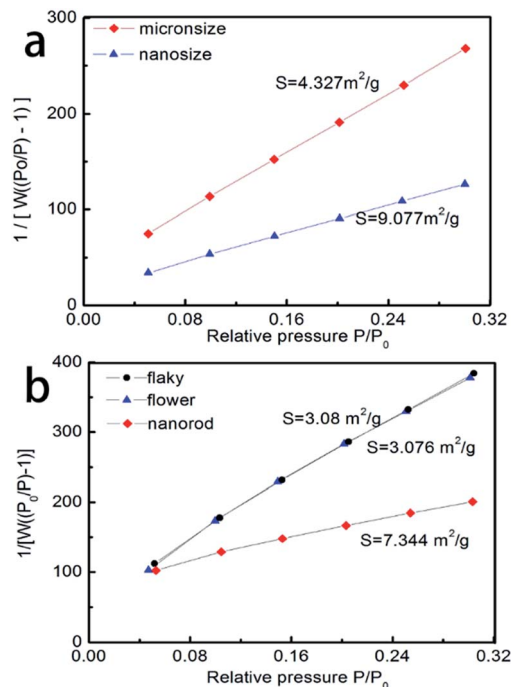


Fig. 5 (a) The specific surface area of the as-synthesized Li₂CO₃ and (b) the specific surface area of the as-synthesized LiMn₂O₄.

(Fig. 5b) with micron-sized Li₂CO₃ represented a smaller specific surface area. This exhibited an adverse result to the specific surface area of Li₂CO₃. The as-synthesized LiMn₂O₄ with micron-sized Li₂CO₃ showed a similar specific surface area (3.08 m² g⁻¹ and 3.076 m² g⁻¹, respectively) while the as-synthesized LiMn₂O₄ with nano-sized Li₂CO₃ exhibited a higher specific surface area (7.344 m² g⁻¹). Moreover, the results showed that the excellent electrochemical performance was probably because of the smaller specific surface area of the as-synthesized LiMn₂O₄.

Fig. 6a shows the discharge capacity of LiMn₂O₄ E2 with the power rate at 2C. Its capacity was maintained at 87.5 mA h g⁻¹ after 200 cycles at this rate. The capacity retention is about 95% at 200th cycle. The inset illustrates the charge–discharge curves of Li/LiMn₂O₄ cells. It can clearly be seen that the samples have two voltage plateaus at approximately 4.0 and 4.1 V, which is a typical feature of the spinel LiMn₂O₄. Two voltage plateaus specify that the insertion and extraction of lithium ion occurred in two stages. The first voltage plateau, at about 4.0 V, is attributed to the removal of lithium ions from half of the tetrahedral sites in which Li–Li interactions occur. The second voltage plateau observed at about 4.1 V is due to the removal of lithium ions from the other tetrahedral sites, in which lithium ions do not have Li–Li interactions. Fig. 6b illustrates the discharge specific capacity at various rates; the charge capacity was measured with the power rate from 0.2C (29.6 mA h) to 5C (740 mA h). It is observed that the as-synthesized spinel LiMn₂O₄ has a much higher discharge capacity than the commercial powders at different power rates. At the lowest value of current (from cycle 1 to 10), the sample has a specific capacity of 107.8 mA h g⁻¹. By increasing the rate to 1C, the

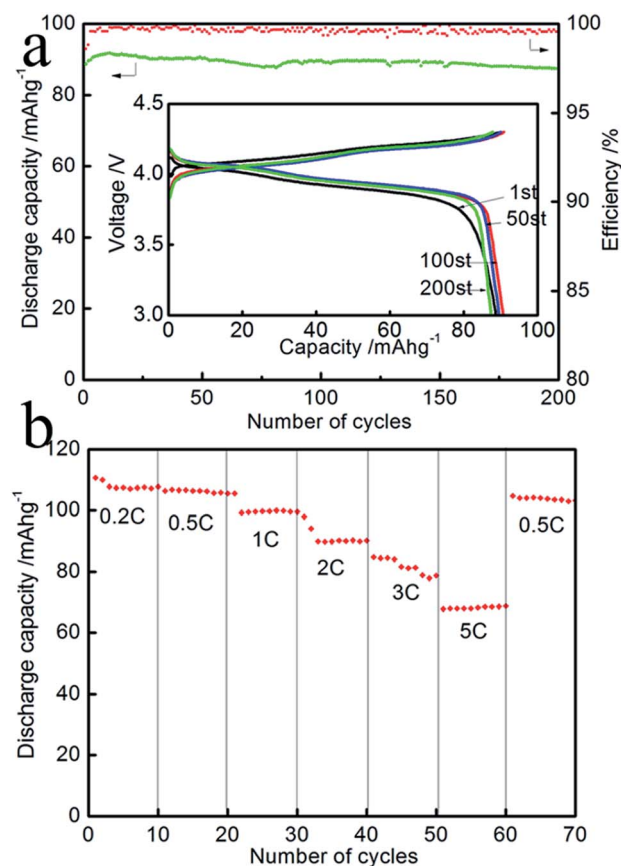


Fig. 6 (a) Discharge specific capacity and efficiency curves vs. number of cycles at the 2C rate. The inset image in (a) is the charge and discharge curves of LiMn_2O_4 at the 2C rate, and (b) the discharge capacity of spinel LiMn_2O_4 at various charge-discharge currents.

sample still has a discharge capacity of 100 mA h g^{-1} . The discharge capacity can return to 104 mA h g^{-1} at the 0.5C rate after the high rate discharge cycles, resulting in excellent cycle performance. The rate performance of E1 and E3 are shown in Fig. 7. The results show that E1 delivered a better rate performance. By analyzing the cycle performance and rate

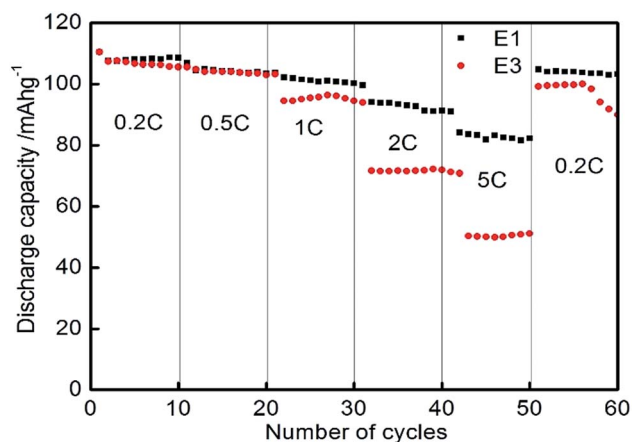


Fig. 7 Comparison of E1 and E3 at different charge-discharge currents.

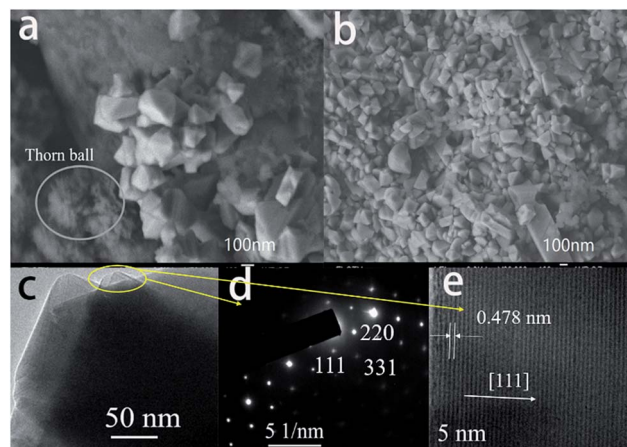


Fig. 8 (a and b) SEM images of spinel LiMn_2O_4 after 200 cycles, the thorn ball besides the spinel LiMn_2O_4 was detected to be impurities. (c) Low and (d), high resolution TEM of spinel LiMn_2O_4 after 200 cycles at the 2C rate. And (e) the electron diffraction pattern image of spinel LiMn_2O_4 after 200 cycles at the 2C rate.

performance results of E1, E2 and E3, it is observed that the as-synthesized LiMn_2O_4 with micron-level Li_2CO_3 shows a better electrochemical performance. The better cycle and rate performance are owing to the smaller specific surface area of the as-synthesized LiMn_2O_4 .

The anode lithium metal foil was immersed in the electrolyte for 6 hours, then washed with alcohol and dried in a vacuum dryer before the analysis methods were conducted. SEM images of spinel LiMn_2O_4 after 200 cycles are shown in Fig. 8a and b. It is seen that the spinel structure was maintained after 200 cycles. The impurity was detected at the same time as shown in Fig. 8a as a thorn ball was located besides the spinel LiMn_2O_4 . It is probably because LiMn_2O_4 had not been washed throughout, while as shown in Fig. 8b the pure spinel LiMn_2O_4 was detected without any impurity phases existing. Moreover, TEM was tested to determine the cubic spinel structure of LiMn_2O_4 . From the images shown in Fig. 8c–e, clear lattice images were observed. The electron diffraction patterns indicate that the single crystalline cubic spinel structure is maintained after 200 cycles. The plane [111] is observed in the TEM images with a lattice parameter of 0.478 nm which corresponds to one of the peaks of LiMn_2O_4 .

Conclusion

The cubic spinel LiMn_2O_4 was synthesized with three different morphologies of the raw material Li_2CO_3 by the solid-state calcination synthesis. The effects of Li_2CO_3 with micro-flaky, micro-flower and nanorod morphologies on the electrochemical properties of spinel LiMn_2O_4 were analyzed. The electrochemical measurements showed that the as-synthesized spinel LiMn_2O_4 (E1, E2 and E3) delivered a discharge capacity of 104.2, 109.1 and 103.9 mA h g^{-1} at the 1C charge-discharge rate. The results showed that the as-synthesized spinel LiMn_2O_4 with micron sized Li_2CO_3 exhibited excellent electrochemical



cyclability because of its well-crystalline nanostructure, and it maintained more than 90% of its initial capacity after 100 charge–discharge cycles at the 1C rate. Moreover, the as-synthesized spinel LiMn_2O_4 E2 retained more than 95% of its initial discharge capacity after 200 cycles at the 2C rate. The cubic spinel structure was detected after 200 cycles of LiMn_2O_4 at this rate. By comparing three types of morphologies of raw material Li_2CO_3 , the Li_2CO_3 particles with micron size were found to be more suitable than those of nano size for calcination with MnO_2 to produce the cubic spinel LiMn_2O_4 .

Conflicts of interest

There are no conflicts to declare.

Abbreviation

E1	LiMn_2O_4 with micro-flaky as raw material
E2	LiMn_2O_4 with flower as raw material
E3	LiMn_2O_4 with nanorod Li_2CO_3 as raw material

Acknowledgements

This work was supported by the National Natural Science Foundation of China under Grant (U1607118).

References

- 1 J. Y. Luo and Y. Y. Xia, *Adv. Funct. Mater.*, 2010, **17**, 3877–3884.
- 2 X. Liu, G. Zhu, K. Yang and J. Wang, *J. Power Sources*, 2007, **174**, 1126–1130.
- 3 P. G. Bruce, S. A. Freunberger, L. J. Hardwick and J. Tarascon, *Nat. Mater.*, 2012, **11**, 19–29.
- 4 J. Li, M. Wei, W. Chu and N. Wang, *Chem. Eng. J.*, 2017, **316**, 277–287.
- 5 A. S. Wills, N. P. Raju and J. E. Greedan, *Chem. Mater.*, 1999, **11**, 1510–1518.
- 6 Q. C. Zhuang, T. Wei, L. L. Du, Y. L. Cui and L. Fang, *J. Phys. Chem. C*, 2010, **114**, 8614–8621.
- 7 H. L. Zhu, Z. Y. Chen, S. Ji and V. Linkov, *Solid State Ionics*, 2008, **179**, 1788–1793.
- 8 R. Singhal, S. R. Das, O. Oviedo, M. S. Tomar and R. S. Katiyar, *J. Power Sources*, 2006, **160**, 651–656.
- 9 C. H. Lu, Y. Lin and H. C. Wang, *J. Mater. Sci. Lett.*, 2003, **22**, 615–618.
- 10 S. Nieto, S. B. Majumder and R. S. Katiyar, *J. Power Sources*, 2004, **136**, 88–98.
- 11 H. Uchiyama, E. Hosono, H. Zhou and H. Imai, *J. Mater. Chem.*, 2009, **19**, 4012–4016.
- 12 K. M. Shaju and P. G. Bruce, *Chem. Mater.*, 2008, **20**, 5557–5562.
- 13 Q. Qu, L. Fu, X. Zhan, D. Samuelis, J. Maier, L. Li, S. Tian, Z. Li and Y. Wu, *Energy Environ. Sci.*, 2011, **4**, 3985–3990.
- 14 D. K. Kim, P. Muralidharan, H. Lee, R. Ruffo, Y. Yang, C. K. Chan, H. Peng, R. A. Huggins and Y. Cui, *Nano Lett.*, 2008, **8**, 3948–3952.
- 15 W. Tang, Y. Hou, F. Wang, L. Liu and Y. Wu, *Nano Lett.*, 2013, **13**, 2036–2040.
- 16 H. Xia, K. R. Ragavendran, J. Xie and L. Lu, *J. Power Sources*, 2012, **212**, 28–34.
- 17 Y. Ding, J. Xie, G. Cao, T. Zhu, H. Yu and X. Zhao, *Adv. Funct. Mater.*, 2011, **21**, 348–355.
- 18 E. Hosono, T. Kudo, I. Honma, H. Matsuda and H. Zhou, *Nano Lett.*, 2009, **9**, 1045–1051.
- 19 H. W. Lee, P. Muralidharan, R. Ruffo, C. M. Mari and Y. Cui, *Nano Lett.*, 2010, **10**, 3852–3856.
- 20 S. Lee, Y. Cho, H. K. Song, K. T. Lee and J. Cho, *Angew. Chem., Int. Ed.*, 2012, **51**, 8748–8752.
- 21 J. Luo, X. Li and Y. Xia, *Electrochim. Acta*, 2007, **52**, 4525–4531.
- 22 X. He, L. Wang, W. Pu and G. Zhang, *Int. J. Electrochem. Sci.*, 2006, **1**, 12–16.
- 23 Y. K. Sun, I. H. Oh and K. Y. Kim, *Ind. Eng. Chem. Res.*, 1997, **36**, 4839–4846.
- 24 H. M. Wu, J. P. Tu, Y. F. Yuan, X. T. Chen and J. Y. Xiang, *J. Power Sources*, 2006, **161**, 1260–1263.
- 25 Y. Bi, T. Wang, M. Liu, R. Du and W. Yang, *RSC Adv.*, 2016, **6**, 19233–19237.
- 26 Z. Zhou, F. Liang, W. Qin and W. Fei, *AIChE J.*, 2014, **60**, 282–288.

