



Cite this: *RSC Adv.*, 2017, 7, 39899

Molecular design of organic small molecules based on diindole–diimide with fused aromatic heterocycles as donors for organic solar cells†

Ruifa Jin *^{ab} and Ahmad Irfan^{cd}

A series of diindole–diimide based small molecules with different aromatic end groups have been designed and investigated theoretically as donor materials for organic solar cells (OSCs). Their optical, electronic, and charge transport properties have been investigated by density functional theory (DFT) and time-dependent density functional theory (TD-DFT). The calculation results showed that the designed molecules can lower the band gap and extend the absorption spectrum towards longer wavelengths. The optical, electronic, and charge transport properties of the designed molecules are affected by the introduction of different aromatic end groups. We have also predicted the mobility of molecules with benzo[*d*]thiazole and benzo[*c*]thiophene end groups. Their hole mobility values are 0.157 and $6.57 \times 10^{-2} \text{ cm}^2 \text{ V}^{-1} \text{ s}^{-1}$, respectively. Our results reveal that the designed molecules are promising hole transport materials and donors with intense broad absorption spectra for OSCs.

Received 24th June 2017
 Accepted 9th August 2017

DOI: 10.1039/c7ra07017a

rsc.li/rsc-advances

Introduction

Organic solar cells (OSCs) have been regarded as potential energy resources due to their promising features, including lightweight, mechanical flexibility, low manufacturing cost, and so on.¹ Unfortunately, the low power conversion efficiencies (PCEs) of OSCs remain the primary obstacle to their commercial application. To address the above issue, great efforts have been intensively devoted to improving the PCEs of OSCs.² PCEs of bulk heterojunction (BHJ) polymer OSCs close to 10% have been reported.³ In recent years, as another branch of OSCs, OSCs based on small molecules have also received much attention.⁴ The PCE of small-molecule OSCs has recently reached 8%.⁵ Although the PCEs of small-molecule OSCs are still lower than those of their polymer counterparts, small molecules possess prominent advantages over their polymer counterparts, such as

well-defined structures, higher charge carrier mobility, easier purification and synthesis, less batch to batch variation, and intrinsic monodispersity.⁶ These advantages make small molecules more suited to mass production than polymer-based ones. Especially, compared with polymer molecule, the structure of small molecule can be tuned more easily, which is beneficial for effective sunlight absorption.⁷ However, the overall performance of OSCs based on the small molecules still lagged behind that of their polymer counterparts. Therefore, a perspective means how to improve the PCEs of small-molecule OSCs is to design and synthesize high-performance and desirable donor novel small molecules. In order to achieve high-performance of small molecules based OSCs, a low band gap (E_g) and a low highest occupied molecular orbital (HOMO) energy level are the two most important factors. In addition, the material should possess high carrier mobility. Nowadays, push–pull structures comprising electron donors and acceptors are regarded as one of the most promising materials in OSCs research.⁸ This push–pull structure can enhance intramolecular charge transfer (ICT) and lower the E_g of material, which extend the absorption spectrum towards longer wavelengths.⁹ At the same time, the HOMO and the lowest unoccupied molecular orbital (LUMO) energy and E_g levels can be tuned effectively by adjusting the push and pull units.¹⁰ Among the various push–pull organic compounds, the indole-containing molecules have been proved to be a promising building block for small-molecule OSCs owing to their strong light-capturing, excellent charge carrier mobility, and good chemical and environmental stability properties.¹¹ Furthermore, the introduction of the electron-deficient aromatic imide group into this electron-rich π -system may lower the energy level of the π -conjugated

^aInner Mongolia Key Laboratory of Photoelectric Functional Materials, Chifeng University, Chifeng 024000, China

^bCollege of Chemistry and Chemical Engineering, Chifeng University, Chifeng 024000, China. E-mail: Ruifajin@163.com

^cResearch Center for Advanced Materials Science, King Khalid University, Abha 61413, P.O. Box 9004, Saudi Arabia

^dDepartment of Chemistry, Faculty of Science, King Khalid University, Abha 61413, P.O. Box 9004, Saudi Arabia

† Electronic supplementary information (ESI) available: Calculated the longest wavelength of absorption λ_{max} and corresponding oscillator strength f of molecule **1** by various methods with 6-31G(d,p) basis set; the λ_{abs} and corresponding oscillator strength f and absorption region R of the first fifteen excited states for **1–6** obtained by the TD-B3LYP/6-31G(d,p) level. The calculated total energies and crystal cell parameters of **5** and **6** in different space groups. Herringbone structures of **5** in $P2_1$ space group and **6** in $Pbcn$ space group. See DOI: 10.1039/c7ra07017a



semiconductors.¹² Recently, indole-containing small-molecules for OSCs have been reported, they showed good charge transport, film morphology, and optical properties.¹³

Herein, we investigated the relationship between topologic structures and optical as well as charge transporting properties for diindole–diimide based small molecules. Furthermore, with the aim to extend the UV-vis in NIR region, the fused aromatic heterocycles as the electron donor were introduced to the π system of molecules. A series of novel diindole–diimide-based small molecules have been designed (Scheme 1). The optical and electronic properties of designed molecules have been investigated. We also select two compounds of the designed compounds as representation to predict their mobility.

Computational methods

Ground neutral and charged state geometries of the designed compounds were optimized by using density functional theory (DFT) at the B3LYP/6-31G(d,p) level.¹⁴ The frontier molecular orbitals (FMOs) of the designed compounds were then calculated at this level of theory based on the optimized geometries. The HOMOs and LUMOs contributions of the individual fragments of the designed compounds to the respective FMO were also calculated. Additionally, the absorption spectra of the designed compounds were predicted using the time dependant density functional theory (TD-DFT) at the TD-B3LYP/6-31G(d,p) level of theory based on the optimized geometries.¹⁵ The molecule **1** was taken as an example to testify to the validity of the selected approach. The geometry optimizations of molecule **1** were carried out by the DFT method using the 6-31G(d,p) basis set. The absorptions were predicted using TD-DFT with the 6-31G(d,p) basis set based on the optimized geometries. The various functionals for all DFT and TD-DFT computations include B3LYP, PBE0, CAM-B3LYP, wB97XD, M062X, and LCWPBE. The longest λ_{abs} of molecule **1** are listed in ESI Table S1.† The results displayed in Table S1† showed that the TD-B3LYP/6-31G(d,p) method provided a better agreement with the reported experimental observations¹³ than those obtained with other methods, with the deviation being only 2 nm. In addition, our designed molecules are used as OSCs materials in the solid film. Therefore, the solvent effect for investigated system is negligible in this work. Thus, geometry optimizations,

band gaps E_g , and absorption properties of designed molecules were carried out by the B3LYP/6-31G(d,p) method.

Generally, the Marcus theory has been applied to estimate the charge transfer rate:^{16,17}

$$K = \left(\frac{V^2}{h}\right) \left(\frac{\pi}{\lambda k_B T}\right)^{\frac{1}{2}} \exp\left(-\frac{\lambda}{4k_B T}\right) \quad (1)$$

where T and k_B are the temperature and the Boltzmann constant, correspondingly, λ and V the reorganization energy and transfer integral, respectively. Clearly, the two key parameters λ and V determine the charge transfer rate. For the λ , it is composed of two parts: external reorganization energy (λ_{ext}) and internal reorganization energy. The λ_{ext} is the effect of the surrounding media in bulk materials, while the λ_{int} is the measure of relaxation in molecular geometry.¹⁸ In this work, we focus only on the λ_{int} and neglect the λ_{ext} because the λ_{ext} values are much smaller than their λ_{int} counterparts in pure organic condensed phases.¹⁹ The λ_{int} can be expressed as:

$$\lambda = \lambda_1 + \lambda_2 = (E_0^\pm - E_\pm^\pm) + (E_\pm^0 - E_0^0) \quad (2)$$

where E_0^0 and E_0^\pm represent the energy of the neutral and cation/anion species with the optimized structure of the neutral species, respectively, while E_\pm^0 and E_\pm^\pm are the energies of the neutral and cation/anion species with the optimized cation/anion structure, respectively. The λ_{int} for electron (λ_e) and hole (λ_h) were calculated at the B3LYP/6-31G(d,p) level.²⁰

The V can be calculated by:^{21,22}

$$V_{ij} = \langle \phi_1^0 | \hat{F}^0 | \phi_2^0 \rangle \quad (3)$$

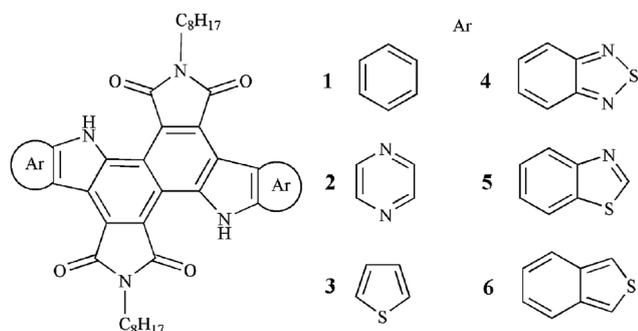
ϕ_1^0 and ϕ_2^0 mean the HOMOs or LUMOs of the two monomers in the dimer. \hat{F}^0 is the Kohn–Sham Hamiltonian of the dimer. The molecular orbitals and density matrix of the two individual monomers are applied to calculate the \hat{F}^0 . In this work, the V were calculated at the pw91pw91/6-31G(d) level.²³ The module polymorph of software package Materials Studio²⁴ was employed to predict the molecular crystal structure. The polymorph predictor calculations are restricted to the ten most popular space groups, $P2_1/c$, $P\bar{1}$, $P2_12_12_1$, $C2/c$, $P2_1$, $Pbca$, $Pna2_1$, Cc , $Pbcn$, and $C2$. The compass force field was used for the prediction. The geometry of the cluster models was taken from B3LYP/6-31G(d,p) level. van der Waals and Coulomb interactions were evaluated by using the Ewald summation method with a cutoff of 6 Å, and the Ewald accuracy tolerance was set to 0.0001 kcal mol⁻¹.

The carrier mobility μ can be evaluated by Einstein equation:

$$\mu = \frac{e}{k_B T} D \quad (4)$$

The diffusion constant D can be obtained by equation:²⁵

$$D = \lim_{t \rightarrow \infty} \frac{1}{2n} \frac{\langle x(t)^2 \rangle}{t} \approx \frac{1}{2n} \sum_i d_i^2 k_i p_i = \frac{1}{2n} \frac{\sum_i d_i^2 k_i^2}{\sum_i k_i} \quad (5)$$



Scheme 1 Molecular structures of the designed compounds.



here d is the intermolecular center-to-center distance, while k is the hopping rate due to charge transfer to neighbour. $n = 3$ means the spatial dimension of the crystal $P_i = k_i / \sum_i k_i$

represents the probability of the specific hopping route. This means that it is a three-dimension averaged diffusion process. All the calculations are carried out using the Gaussian 09 program.²⁶

Results and discussion

Frontier molecular orbital

With the aim to characterize the optical, electronic, and charge transport properties, the distribution of electron density in HOMOs and LUMOs for the designed compounds are plotted in Fig. 1. Table 1 presents the E_{HOMO} , E_{LUMO} , E_g and the HOMOs and LUMOs contributions of the 3,8-dihydroindolo[7,6-g]indole-diimide fragments (**IID**) and aromatic end groups (**Ar**) fragments (in %) to the FMO. As visualized in Fig. 1, both the HOMOs and LUMOs of the designed compounds exhibited π characteristics. Furthermore, the distribution patterns of HOMOs and LUMOs delocalized across the whole molecular backbone because of the planar and rigid conjugated molecular geometry. The results displayed in Table 1 reveal that the HOMOs of the designed compounds are mainly localized on the **IID** and **Ar** fragments. On the other hand, the LUMOs mainly reside at the **IID** fragments with minimal density on **Ar** fragments. The contributions of **IID** fragments of LUMOs are larger than 78.8%, while the contributions of **Ar** fragments are within 8% except the corresponding contribution of **Ar** fragment is 21.2% for 5. As expect, the designed compounds exhibited obvious the ICT character of the vertical $S_0 \rightarrow S_1$ transition due to the planar and rigid molecular structure. Analysis of the FMOs indicates that the HOMOs \rightarrow LUMOs excitations results in transfer of electron density from the **Ar** fragments to **IID** fragments for the designed compounds. The percentages of charge transfer from **Ar** to **IID** fragments for 1–6 are 23, 7.3, 60.8, 35.5, 26.4, and 56.6%, respectively. It suggests that the **Ar** fragments serve as donors and **IID** fragments serve as acceptors for the designed compounds.

To gain a deeper insight into the influence of the optical and electronic properties, the E_{HOMO} , E_{LUMO} , and E_g were analyzed. From Table 1, one can find that both the E_{HOMO} and E_{LUMO}

Table 1 The FMOs energy and band gaps E_g , and FMOs molecular orbital contribution (%) from individual fragments (in %) to the FMOs of the designed compounds at the B3LYP/6-31G(d,p) level

Species	HOMO			LUMO			E_g
	E_{HOMO}	Ar ^a	IID ^b	E_{LUMO}	Ar	IID	
1	−5.494	27.8	72.2	−2.764	4.8	95.2	2.730
2	−6.106	12.1	87.9	−2.973	4.8	95.2	3.133
3	−5.520	65.3	34.7	−2.808	4.5	95.5	2.712
4	−5.571	42.7	57.3	−3.098	7.2	92.8	2.473
5	−5.516	47.6	52.4	−2.900	21.2	78.8	2.616
6	−4.811	62.4	37.6	−2.724	5.8	94.2	2.087

^a **Ar**: aromatic groups. ^b **IID**: 3,8-dihydroindolo[7,6-g]indole-diimide fragments.

values of 2–5 decrease compared with those of parent compound 1, respectively. Interestingly, the E_{HOMO} and E_{LUMO} values of 6 increases compared with those of parent compound 1, respectively. The trend of the E_{HOMO} is $6 > 1 > 5 > 3 > 4 > 2$ while E_{LUMO} is $6 > 1 > 3 > 5 > 2 > 4$. On the other hand, the E_g is in the order of $2 > 1 > 3 > 5 > 4 > 6$. These results indicate that molecule with benzo[*c*]thiophene (6) end group possesses lower both E_{HOMO} and E_{LUMO} , while molecules with pyrazine (2), thiophene (3), benzo[*c*][1,2,5]thiadiazole (4), benzo[*d*]thiazole (5) end groups have higher E_{HOMO} and E_{LUMO} compared with molecule with benzene end group. The E_g value of molecule with pyrazine end group are larger, while the corresponding values of molecules with thiophene, benzo[*c*][1,2,5]thiadiazole, benzo[*d*]thiazole, benzo[*c*]thiophene end groups are smaller than that of molecules with benzene end group. These results suggest that the different end groups have effects on the E_{HOMO} , E_{LUMO} , and E_g for the compounds under investigation.

Absorption spectra

Table 2 presents the longest wavelength of absorption λ_{abs} , the oscillator strength f , main configurations, and the absorption region R of the designed molecules. R corresponds to the difference of the longest and shortest wavelength values with oscillator strength larger than 0.01 considering the first fifteen excited states. The corresponding λ_{abs} and f of the first fifteen

Table 2 Predicted λ_{abs} , corresponding f and R , and main configurations of 1–6 at the TD-B3LYP/6-31G(d,p)//B3LYP/6-31G(d,p) level

Species	λ_{abs}	f	Main assignment	R
1	542 (339) ^b	0.127 (0.321) ^c	H \rightarrow L (0.70)	236
2	473 (371)	0.084 (0.210)	H \rightarrow L (0.69)	140
3	559 (308)	0.088 (0.550)	H \rightarrow L (0.70)	259
4	595 (353)	0.175 (1.265)	H \rightarrow L (0.70)	242
5	565 (334)	0.188 (1.001)	H \rightarrow L (0.70)	244
6	708 (346)	0.195 (1.251)	H \rightarrow L (0.70)	362
Exp ^a	540 (340)			

^a Experimental data for 1 from ref. 13. ^b Data in the parentheses are the wavelength with the largest f values. ^c Data in the parentheses are the f values.

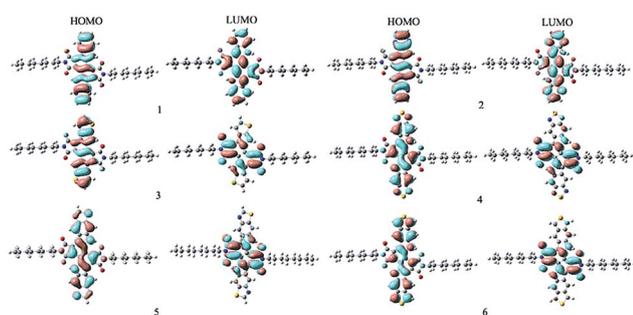


Fig. 1 The FMOs of the designed compounds at the B3LYP/6-31G(d,p) level.



excited states for the designed compounds are listed in Table SII in ESI.† We also present the wavelength with the largest f of absorption in Table 2. From the data in Table 2 and Table SII in ESI,† one can find that the **1** has three main absorption peaks at about 542 ($f = 0.127$), 339 ($f = 0.321$), and 318 ($f = 0.283$) nm. These results are in good agreement with the experimental results namely, absorption peaks at about 540, 340, and 320 nm, which are all observed in experiments.¹³ Furthermore, the designed molecules have two main absorption peaks at about 350 and 500–700 nm. We aim to obtain a broad and intense absorption region which can match the sunlight spectrum by the introduction of aromatic end groups. Therefore, we mainly focus on the UV-vis regions absorption of the designed molecules. The major assignments of the lowest $S_0 \rightarrow S_1$ electronic transitions for **1–6** are mainly as HOMOs \rightarrow LUMOs. It can be noted that the λ_{abs} of **3–6** exhibit bathochromic shifts, while the corresponding λ_{abs} of **2** has hypsochromic shift compared with that of the parent compound **1**. The λ_{abs} values are in the order of $6 > 4 > 5 > 3 > 1 > 2$, which is in excellent agreement with the corresponding reverse order of E_g values displayed in Table 1. It indicates that the introduction of pyrazine (**2**) end group decrease, while the introduction of thiophene (**3**), benzo[*c*][1,2,5]thiadiazole (**4**), benzo[*d*]thiazole (**5**), benzo[*c*]thiophene (**6**) end groups increase the λ_{abs} values compared with parent molecules with benzene end group. The sequence of R is $6 > 3 > 5 > 4 > 1 > 2$. It suggests that the introduction of pyrazine end group end results in smaller R value, while the introduction of thiophene, benzo[*c*][1,2,5]thiadiazole, benzo[*d*]thiazole, benzo[*c*]thiophene end groups lead to the increase of R values compared with parent molecules with benzene end group. In general, larger oscillator strength means larger experimental absorption coefficient. The order of the predicted f is in the decreasing order of $6 > 5 > 4 > 1 > 3 > 2$. It indicates that the introduction of benzo[*c*][1,2,5]thiadiazole, benzo[*d*]thiazole, benzo[*c*]thiophene end groups increases, while the introduction of pyrazine and thiophene end groups decreases slightly the f values compared with benzene end group for the designed molecules. Furthermore, from Table 2 one can find that the λ_{abs} and R values of molecules with benzo[*c*][1,2,5]thiadiazole, benzo[*d*]thiazole, benzo[*c*]thiophene end groups are larger than those of other molecules. It suggests that molecules with benzo[*c*][1,2,5]thiadiazole, benzo[*d*]thiazole, benzo[*c*]thiophene end groups can lower the material band gap and extend the absorption spectrum towards longer wavelengths. The λ_{abs} and R values of molecules with pyrazine and thiophene end groups are smaller slightly than those of parent molecule **1**. Therefore, the designed compounds possess the large λ_{abs} , f , and R values. The designed molecules could be used as donors with intense broad absorption spectra for solar cell material, particularly for **4–6**.

Reorganization energy

In order to design good candidates for OSCs devices, one key factor is to understand the relationship between molecular structure and charge transport property of material. As mentioned above, the reorganization energy is one major parameter determining the charge transfer rate.^{16,17} The

calculated reorganization energies for hole and electron are listed in Table 3. From Table 3, one can find that the λ_e of **1–6** are larger than that of the typical electron transport material ($\lambda_e = 0.276$ eV), tris(8-hydroxyquinolato)aluminum(III) (Alq3).²⁷ On the contrary, the λ_h values of **1–6** are smaller than that of the typical hole transport material ($\lambda_h = 0.290$ eV), *N,N'*-diphenyl-*N,N'*-bis(3-methylphenyl)-(1,1'-biphenyl)-4,4'-diamine (TPD).²⁸ This indicates that the hole transfer rates of **1–6** might be higher than that of TPD. Furthermore, λ_h values of **2** and **3** are larger, while the corresponding values of **4–6** are smaller than that of **1**. It suggests that the introduction of pyrazine and thiophene end groups decrease, while the introduction of benzo[*c*][1,2,5]thiadiazole, benzo[*d*]thiazole, benzo[*c*]thiophene end groups increase the and hole transfer rates compared with parent molecule **1**. It implies that the designed molecules can be used as promising hole transport materials from the stand point of the smaller reorganization energy.

Charge-carrier mobility

We select **5** and **6** as representation to study their charge transport property. The total energies and lattice constants of the predicted crystal structures for **5** and **6** in different space groups are summarized in Tables SIII and SIV in ESI,† respectively. The crystal structures of **5** and **6** with the lowest total energies belong to space groups $P2_1$ and $Pbcn$ respectively and shown in Fig. S1 in ESI.† Thus, we predict the mobility of **5** and **6** in these two space groups, respectively. The main carrier hopping pathways of **5** and **6** are selected based on their optimized crystal structures, respectively. The selected pathways (dimers) of **5** and **6** in space groups $P2_1$ and $Pbcn$ respectively are depicted in Fig. 2. The calculated transfer integrals and mobility of **5** and **6** for holes and electrons in space groups $P2_1$ and $Pbcn$ are listed in Table 4. The hole or electron coupling means the orbital coupling of the neighboring molecules. The FMO distributions and relative position of the interacting molecules are two major parameters that determine the electronic coupling.²⁹ From Table 4, it can be noted that **5** possesses the largest absolute electron and hole coupling values in pathways **1** and **2** in space group $P2_1$. For **6**, the largest absolute electron and hole coupling are located in pathways **4** and **7** in space group $Pbcn$. The orientation of the interacting molecules is the key factor of hole or electron coupling for **5** and **6**. The values of hole mobility of **5** and **6** for $P2_1$ and $Pbcn$ space groups (0.157 and $6.57 \times 10^{-3} \text{ cm}^2 \text{ V}^{-1} \text{ s}^{-1}$) are larger than that of TPD ($1.0 \times 10^{-3} \text{ cm}^2 \text{ V}^{-1} \text{ s}^{-1}$),³⁰ respectively. The values of electron mobility of **5** and **6** for $P2_1$ and $Pbcn$

Table 3 Calculated λ_e and λ_h (both in eV) of **1–6** at the B3LYP/6-31G(d,p) level

Species	λ_h	λ_e
1	0.173	0.329
2	0.261	0.355
3	0.266	0.340
4	0.104	0.280
5	0.129	0.322
6	0.068	0.288



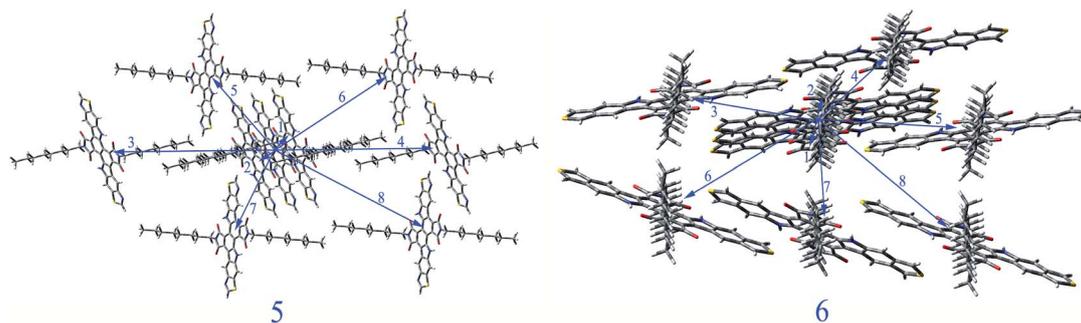


Fig. 2 Crystal structures and hopping routes of **5** in $P2_1$ space group and **6** in $Pbcn$ space group.

Table 4 The hole and electron coupling (in eV) in all of the nearest neighbor pathways and the distance (in Å) for **5** and **6** in $P2_1$ and $Pbcn$ space groups respectively

5	Pathway	Distance	Electron coupling	Hole coupling
$P2_1$	1	4.809	-7.70×10^{-3}	-2.80×10^{-2}
	2	4.809	-7.70×10^{-3}	-2.80×10^{-2}
	3	31.560	6.30×10^{-12}	-9.14×10^{-10}
	4	31.560	6.30×10^{-12}	-9.14×10^{-10}
	5	14.189	-3.20×10^{-5}	-3.85×10^{-5}
	6	27.553	3.10×10^{-14}	1.23×10^{-12}
	7	14.189	-3.20×10^{-5}	-3.85×10^{-5}
	8	27.553	3.10×10^{-14}	1.23×10^{-12}
Drift mobility			1.15×10^{-3}	0.157
$Pbcn$	1	30.038	1.47×10^{-9}	-3.67×10^{-8}
	2	30.038	1.47×10^{-9}	-3.67×10^{-8}
	3	18.369	-1.21×10^{-7}	2.37×10^{-6}
	4	7.972	1.90×10^{-3}	-1.80×10^{-3}
	5	18.369	-1.21×10^{-7}	2.37×10^{-6}
	6	20.643	3.46×10^{-10}	1.75×10^{-10}
	7	7.883	-8.16×10^{-4}	-7.20×10^{-3}
	8	20.643	3.46×10^{-10}	1.75×10^{-10}
Drift mobility			2.46×10^{-4}	6.57×10^{-2}

space groups are 1.15×10^{-2} and $2.46 \times 10^{-4} \text{ cm}^2 \text{ V}^{-1} \text{ s}^{-1}$, respectively. Furthermore, the values of both hole and electron mobility of **5** are larger than those of **6**, which is in agreement with the corresponding reverse order of reorganization energies for hole and electron values displayed in Table 3, respectively. It indicates shows that the stacking structure is the most important factor for molecular mobility property.

Considering the reorganization energy and charge-carrier mobility, the designed molecules can be made as hole transfer materials high hole mobility. It is well established that the mobility of charge carriers play a dominant role in the efficiency of OSCs. The high hole mobility is a basic requirement for effective photovoltaic active donors, to ensure effective charge carrier transport to the electrode and reduce the photocurrent loss in OSCs.³¹ Additionally, the high hole mobility for the donors as a hole transport layer is also crucial for the carrier transport to enhancement in both short-circuit current density (J_{sc}) and fill factor (FF) of the devices, and consequently the overall PCE of

OSCs.³² The calculation results suggest that the designed molecules can be made as donors material with intense broad absorption spectra and hole transfer materials using for OSCs.

Conclusions

In summary, a series of diindole–diimide-based small molecules with different aromatic end groups have been systematically investigated for OSCs applications. Analysis of the FMOs indicates that the vertical $S_0 \rightarrow S_1$ transition results in transfer of electron density from the **Ar** fragments to **IID** fragments for the designed compounds. The introduction of different aromatic end groups affects the optical, electronic, and charge transport properties of the designed compounds. The molecules can lower the material band gap and extend the absorption spectrum towards longer wavelengths. Our results show that the molecules under investigation own the longest wavelength of absorption spectra, oscillator strength, and absorption region values. In addition, the mobility values of **5** and **6** as representation are also investigated. The values of hole mobility of **5** and **6** are 0.157 and $6.57 \times 10^{-2} \text{ cm}^2 \text{ V}^{-1} \text{ s}^{-1}$, respectively, which are larger than that of TPD ($1.0 \times 10^{-3} \text{ cm}^2 \text{ V}^{-1} \text{ s}^{-1}$). On the basis of the investigated results, the designed molecules are expected to be promising candidates for donors with intense broad absorption spectra and hole transport materials for OSCs.

Conflicts of interest

There are no conflicts to declare.

Acknowledgements

Financial supports from the NSFC (No. 21563002), the Natural Science Foundation of Inner Mongolia Autonomous Region (No. 2015MS0201), and the Research Program of Sciences at Universities of Inner Mongolia Autonomous Region (No. NJZZ235) are gratefully acknowledged.

Notes and references

- (a) R. F. Service, *Science*, 2011, **332**, 293; (b) H. X. Shang, H. J. Fan, Y. Liu, W. P. Hu, Y. F. Li and X. W. Zhan, *Adv. Mater.*, 2011, **23**, 1554–1557; (c) Y. Zou, A. Najari,



- P. Berrouard, S. Beaupré, R. B. Aich, Y. Tao and M. Leclerc, *J. Am. Chem. Soc.*, 2010, **132**, 5330–5331; (d) Z. He, C. Zhong, X. Huang, W. Wong, H. Wu, L. Chen, S. Su and Y. Cao, *Adv. Mater.*, 2011, **23**, 4636–4643; (e) Y. Lin, Y. F. Li and X. W. Zhan, *Chem. Soc. Rev.*, 2012, **41**, 4245–4272; (f) B. Walker, C. Kim and T. Q. Nguyen, *Chem. Mater.*, 2011, **23**, 470–482.
- 2 (a) Y. Liang and L. Yu, *Polym. Rev.*, 2010, **50**, 454–473; (b) N. F. Montcada, B. Pelado, A. Viterisi, J. Albero, J. Coro, P. D. L. Cruz, F. Langa and E. Palomares, *Org. Electron.*, 2013, **14**, 2826–2832.
- 3 (a) J. You, L. Dou, K. Yoshimura, T. Kato, K. Ohya, T. Moriarty, K. Emery, C. C. Chen, J. Gao, G. Li and Y. Yang, *Nat. Commun.*, 2013, **4**, 1446–1456; (b) G. Li, R. Zhu and Y. Yang, *Nat. Photonics*, 2012, **6**, 153–161.
- 4 (a) B. Walker, C. Kim and T. Q. Nguyen, *Chem. Mater.*, 2011, **23**, 470–482; (b) D. Demeter, T. Rousseau, P. Leriche, T. Cauchy, R. Po and J. Roncali, *Adv. Funct. Mater.*, 2011, **21**, 4379–4387.
- 5 (a) A. K. K. Kyaw, D. H. Wang, V. Gupta, J. Zhang, S. Chand, G. C. Bazan and A. J. Heeger, *Adv. Mater.*, 2013, **25**, 2397–2402; (b) A. K. K. Kyaw, D. H. Wang, V. Gupta, W. L. Leong, L. Ke, G. C. Bazan and A. J. Heeger, *ACS Nano*, 2013, **7**, 4569–4577; (c) J. Y. Zhou, Y. Zuo, X. J. Wan, G. K. Long, Q. Zhang, N. Wang, Y. S. Liu, Z. Li, G. R. He, C. X. Li, B. Kan, M. Li and Y. S. Chen, *J. Am. Chem. Soc.*, 2013, **135**, 8484–8575.
- 6 (a) Y. Lin, Y. Li and X. Zhan, *Chem. Soc. Rev.*, 2012, **41**, 4245–4272; (b) A. Mishra and P. Bäuerle, *Angew. Chem., Int. Ed.*, 2012, **51**, 2020–2067; (c) Z. He, C. Zhong, X. Huang, W. Wong, H. Wu, L. Chen, S. Su and Y. Cao, *Adv. Mater.*, 2011, **23**, 4636–4643.
- 7 T. Zhang, M. Liu, Q. Zeng, Z. Wu, L. Piao and S. Zhao, *RSC Adv.*, 2013, **3**, 13259–13264.
- 8 (a) X. Li, W. C. H. Choy, L. Huo, F. Xie, W. E. I. Sha, B. Ding, X. Guo, Y. Li, J. Hou, J. You and Y. Yang, *Adv. Mater.*, 2012, **24**, 3046–3052; (b) X. Hu, M. Shi, J. Chen, L. Zuo, L. Fu, Y. Liu and H. Chen, *Macromol. Rapid Commun.*, 2011, **32**, 506–511; (c) Y. Huan, X. Guo, F. Liu, L. Huo, Y. Chen, T. P. Russell, C. C. Han, Y. Li and J. Hou, *Adv. Mater.*, 2012, **24**, 3383–3389; (d) J. Roncali, *Acc. Chem. Res.*, 2009, **42**, 1719–1730.
- 9 (a) J. Zhang, D. Den, C. He, M. Zhang, Z. Zhang, Z. Zhang and Y. Li, *Chem. Mater.*, 2011, **23**, 817–822; (b) Y. Sun, G. C. Welch, W. L. Leong, C. J. Takacs, G. C. Bazan and A. J. Heeger, *Nat. Mater.*, 2012, **11**, 44–48.
- 10 (a) B. Yin, L. Yang, Y. Liu, Y. Chen, Q. Qi, F. Zhang and S. Yin, *Appl. Phys. Lett.*, 2010, **97**, 023303; (b) Y. Liu, J. Zhou, X. Wan and Y. Chen, *Tetrahedron*, 2009, **65**, 5209–5215; (c) J. Zhang, D. Den, C. He, M. J. Zhang, Z. G. Zhang, Z. J. Zhang and Y. F. Li, *Chem. Mater.*, 2010, **23**, 817–822.
- 11 (a) P. F. Xia, X. J. Feng, J. Lu, S. W. Tsang, R. Movileanu, Y. Tao and M. S. Wong, *Adv. Mater.*, 2008, **20**, 4810–4815; (b) J. Sim, K. Do, K. Song, A. Sharma, S. Biswas, G. D. Sharma and J. Ko, *Org. Electron.*, 2016, **30**, 122–130; (c) D. Yang, Y. Jiao, Y. Huang, T. Zhuang, L. Yang, Z. Lu, X. Pu, H. Sasabe and J. Kido, *Org. Electron.*, 2016, **32**, 179–186; (d) G. D. Sharma, M. A. Reddy, K. Ganesh, S. P. Singh and M. Chandrasekharam, *RSC Adv.*, 2014, **4**, 732–742; (e) F. Baert, C. Cabanetos, M. Allain, V. Silvestre, P. Leriche and P. Blanchard, *Org. Lett.*, 2016, **18**, 1582–1585; (f) J. Sim, H. Lee, K. Song, S. Biswas, A. Sharma, G. D. Sharma and J. Ko, *J. Mater. Chem. C*, 2016, **4**, 3508–3516; (g) B. Liu, Q. Liu, D. You, X. Li, N. Y. Aruta and W. Zhu, *J. Mater. Chem.*, 2012, **22**, 13348–13356.
- 12 (a) J. Gao, Y. Li and Z. Wang, *Org. Lett.*, 2013, **15**, 1366–1369; (b) S. L. Suraru, U. Zschieschang, H. Klauk and F. Würthner, *Chem. Commun.*, 2011, **47**, 11504–11506; (c) Y. Dong and J. Xie, *J. Mol. Sci.*, 2014, **30**, 118–123.
- 13 Z. Zhao, Z. Wang, X. Zhang, S. Gao, X. Yang, Z. Duan and X. Gao, *ChemPlusChem*, 2015, **80**, 57–61.
- 14 (a) A. D. Becke, *J. Chem. Phys.*, 1993, **98**, 5648–5652; (b) C. Lee, W. Yang and R. G. Parr, *Phys. Rev. B: Condens. Matter Mater. Phys.*, 1988, **37**, 785–789; (c) B. J. Lynch, P. L. Fast, M. Harris and D. G. Truhlar, *J. Phys. Chem. A*, 2000, **104**, 4811–4815.
- 15 (a) C. R. Zhang, W. Z. Liang, H. S. Chen, Y. H. Chen, Z. Q. Wei and Y. Z. Wu, *J. Mol. Struct.*, 2008, **862**, 98–104; (b) J. Sun, J. Song, Y. Zhao and W. Z. Liang, *J. Chem. Phys.*, 2007, **127**, 234107–234113.
- 16 R. A. Marcus, *Rev. Mod. Phys.*, 1993, **65**, 599–610.
- 17 R. A. Marcus, *Annu. Rev. Phys. Chem.*, 1964, **15**, 155–196.
- 18 V. Lemaur, M. Steel, D. Beljonne, J. L. Brédas and J. Cornil, *J. Am. Chem. Soc.*, 2005, **127**, 6077–6086.
- 19 (a) D. L. Cheung and A. Troisi, *J. Phys. Chem. C*, 2010, **114**, 20479–20488; (b) D. P. McMahon and A. Trois, *J. Phys. Chem. Lett.*, 2010, **1**, 941–946; (c) S. D. Motta, E. D. Donato, F. Negri, G. Orlandi, D. Fazzi and C. Castiglioni, *J. Am. Chem. Soc.*, 2009, **131**, 6591–6598.
- 20 M. E. Köse, W. J. Mitchell, N. Kopidakis, C. H. Chang, S. E. Shaheen, K. Kim and G. Rumbles, *J. Am. Chem. Soc.*, 2007, **129**, 14257–14270.
- 21 T. Fujita, H. Nakai and H. Nakatsuji, *J. Chem. Phys.*, 1996, **104**, 2410–2417.
- 22 A. Troisi and G. Orlandi, *J. Phys. Chem. A*, 2006, **110**, 4065–4070.
- 23 B. Liu, X. Chen, Y. Zou, L. Xiao, X. Xu, Y. He, L. Li and Y. Li, *Macromolecules*, 2012, **45**, 6898–6905.
- 24 *Materials Studio*, Accelrys Inc., San Diego, CA, 2005.
- 25 W. Q. Deng and W. A. Goddard, *J. Phys. Chem. B*, 2004, **108**, 8614–8621.
- 26 M. J. Frisch, G. W. Trucks, H. B. Schlegel, G. E. Scuseria, M. A. Robb, J. R. Cheeseman, V. G. Zakrzewski, J. A. Montgomery Jr, R. E. Stratmann, J. C. Burant, S. Dapprich, J. M. Millam, A. D. Daniels, K. N. Kudin, M. C. Strain, O. Farkas, J. Tomasi, V. Barone, M. Cossi, R. Cammi, B. Mennucci, C. Pomelli, C. Adamo, S. Clifford, J. Ochterski, G. A. Petersson, P. Y. Ayala, Q. Cui, K. Morokuma, D. K. Malick, A. D. Rabuck, K. Raghavachari, J. B. Foresman, J. Cioslowski, J. V. Ortiz, B. B. Stefanov, G. Liu, A. Liashenko, P. Piskorz, I. Komaromi, R. Gomperts, R. L. Martin, D. J. Fox, T. Keith, M. A. Al-Laham, C. Y. Peng, A. Nanayakkara, C. Gonzalez, M. Challacombe, P. M. W. Gill, B. G. Johnson,



- W. Chen, M. W. Wong, J. L. Andres, M. Head-Gordon, E. S. Replogle and J. A. Pople, *Gaussian 09*, Gaussian, Inc, Wallingford, CT, 2009.
- 27 B. C. Lin, C. P. Cheng, Z. Q. You and C. P. Hsu, *J. Am. Chem. Soc.*, 2005, **127**, 66–67.
- 28 N. E. Gruhn, D. A. da Silva Filho, T. G. Bill, M. Malagoli, V. Coropceanu, A. Kahn and J. L. Brédas, *J. Am. Chem. Soc.*, 2002, **124**, 7918–7919.
- 29 C. Wang, F. Wang, X. Yang, Q. Li and Z. Shuai, *Org. Electron.*, 2008, **9**, 635–640.
- 30 K. Shizu, T. Sato, A. Ito, K. Tanaka and H. Kaji, *J. Mater. Chem.*, 2011, **21**, 6375–6382.
- 31 G. D. Sharma, M. Anil Reddy, K. Ganesh, S. P. Singhb and M. Chandrasekharam, *RSC Adv.*, 2014, **4**, 732–742.
- 32 Y. Li, *Acc. Chem. Res.*, 2012, **45**, 723–733.

