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Improved corrosion resistance and antibacterial properties of composite arch-wires by N-doped TiO₂ coating†

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In this paper, composite arch-wires (CAWs) coated with TiO₂ and N-doped TiO₂ nanocrystal thin films were fabricated. We found that the corrosion resistance and antibacterial properties of CAWs can be significantly improved by coating with the TiO₂ nanocrystal thin film, especially using the N-doped TiO₂ nanocrystal thin film, which is extremely important for their clinical application. The improved performances were confirmed by measuring potentiodynamic polarization curves and electrochemical impedance spectroscopy (EIS). The results revealed that N-doped TiO₂ coated CAWs can prolong the wearing time and improve the security of the arch wire in the oral cavity. In addition, the antibacterial activities of the samples were tested against *Streptococcus mutans* by bacterial culture. It was found that the CAWs coated with N-doped TiO₂ showed the highest antibacterial activity. Finally, the cytotoxicity of the samples was tested and evaluated, and the results showed that the CAWs modified with the N-doped TiO₂ nanocrystal thin film have great potential in practical dental treatment.

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1. Introduction

In modern society, people pay more and more attention to oral beauty, and orthodontic treatment is an effective way to improve the arrangement of teeth. So far, several efficient and multi-functional orthodontic materials have been well developed. The composite arch wire (CAW) is one of the most important orthodontic materials which was developed by Daqian Sun *et al.* in 2011.¹ CAW consists of a nickel–titanium shape memory alloy (TiNi SMA), and stainless steel (SS) arch wire which are laser-welded *via* a copper inter layer.² CAW is a new type of orthodontic arch wire. At present, NiTi and SS arch wires are widely used in different stages of orthodontic treatment. Nickel titanium wire is used first, then the stainless steel wire is gradually utilized. In this way, the orthodontic treatment period is prolonged, which decreases the efficiency of the orthodontic treatment. Especially, the different stiffness and

performance of the arch wires are needed in the different parts and different periods of the dental arch teeth. Thus, it cannot meet the clinical needs by using a single arch wire, and CAW can solve this problem. CAW can move the teeth with nickel titanium arch wires and provide anchorage with stainless steel wire. The power of the soft nickel titanium arch wires are directly used on the dislocation tooth to move tooth effectively, while the rigidity of stainless steel wires are used to strengthen anchorage teeth stability and prevent the negative moving. Then, CAW combines the advantages of each material to correct the malposition of the tooth which can improve the treatment efficiency and shorten the course of treatment.

Recent research demonstrated that copper (Cu) can obviously improve the mechanical property of the CAW when it is used as the welding transition layer.^{3,4} Compared with NiTi shape memory alloy and SS, Cu has a lower melting point and a decent strength; thus Cu exhibits excellent ductility and deformation capacity, which makes the joint flexible and ductile. The joint tensile strength and the shape recovery ratio of NiTi alloy can be greater than 520 MPa and 98%, respectively, which can meet the requirements of clinical performance.⁵

The selection of materials for dental applications depends upon a number of factors such as corrosion behavior, mechanical properties, biocompatibility and aesthetic value. The corrosion behavior of metallic materials is usually the most important property, because biocompatibility and cytotoxicity are dependent on the products of the corrosion process. It has been reported that 46% of the orthodontic patients will occur demineralization within 12 months who are treated with fixed

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appliance.⁶ The main reason of demineralization during orthodontic treatment with fixed appliance is attached dental plaque. It has been reported that after using fixed appliance for 1 to 3 weeks, the proportion of streptococci mutans in dental plaque increased one time. Note that the difference is significant, and the corrosion of the metals in which played an important role, which had been found to induce specific changes in the oral environment, such as decreasing pH and increasing dental plaque accumulation. The increase in the number of lactobacilli and *Streptococcus mutans* will decrease the pH value in the mouth, which is contributed to tooth demineralization when pathological changes occurred.⁷ Traditional preventive methods of enamel demineralization include proper tooth brushing, fluoride rinse or varnishes,⁸ a fluoride-releasing bonding system,⁹ acidulated phosphate fluoride (APF) foam or gel,¹⁰ and using acid etching area.¹¹ However, these approaches require good cooperation with orthodontic patients, and the drug concentration is difficult to maintain. The complicated procedures would make clinical treatment more difficult. Our group has performed a large number of experimental studies on CAW, and we found that CAW has no obvious antimicrobial properties and its corrosion resistance is also slightly lower than those of two base materials. In order to improve the corrosion resistance and antibacterial properties of CAW, we attempt to deposit TiO₂ and N-doped TiO₂ nanocrystal thin films on the surface of CAWs to improve their performance.

Until now, the nanotechnology has been widely used in the biomedical field. Titanium dioxide, as a photo-catalytic antibacterial material, has considerable beneficial properties: chemical stability, biocompatibility, high potential for self-cleaning, and high antibacterial activity.^{12–17} But TiO₂ is a wide-band gap semiconductor (3.2 eV for anatase TiO₂), and it can only absorb UV light (≤ 387 nm), so only a small part (~5%) of solar light can be utilized.¹⁸ Asahi *et al.* reported that N-doped TiO₂ has obvious catalytic activity under visible light irradiation compared with pure TiO₂.¹⁹ It was confirmed that N-doping can extend the absorption edge of TiO₂ to the visible region, and narrow the band gap,²⁰ thus enhancing the photo-catalytic antibacterial ability of TiO₂.²¹ In general, since understanding of new materials is often inadequate, medical workers often need to carry out *in vitro* experiments to select the appropriate material. For dental materials, the saliva in oral cavity is the basic corrosion environment which we simulated for studying corrosion resistance and antibacterial properties of CAWs *in vitro*. Radio frequency (RF) magnetron sputtering is an effective method due to its inherent multi functionality, uniform surface coverage and low-temperature deposition.^{22,23} The main advantage of magnetron sputtering is able to deposit inorganic nanoparticle thin films with controllable size and shape by tuning preparation parameters.²⁴

The newly developed CAWs with excellent mechanical properties have great potential for orthodontic applications,²⁵ but the corrosion resistance and antibacterial performance of CAW are still not satisfactory. In this study, in order to further improve the anticorrosion ability and antibacterial activity of CAWs, TiO₂ and N-doped TiO₂ nanocrystal thin films are respectively deposited on the CAW surface. The modified

samples were then exposed to artificial saliva to simulate common conditions in the oral cavity in order to investigate their anticorrosion ability. The antibacterial activity was also studied by potentiodynamic polarization curves and EIS. *Streptococcus mutans* culture in the samples' extract was studied to evaluate the antibacterial properties of the samples.

2. Experimental

2.1 Materials and sample preparation

Ti–44.73 wt% NiTi SMA wire (Smart Co., Beijing, China), Fe–18Cr–8Ni (SS), and pure Cu were used as the base metals in this paper. The dimensions of the wires were 15 mm (length) \times 0.64 mm (width) \times 0.48 mm (thickness). The thickness of the pure Cu interlayer was 0.2 mm. The base metal was ground using silicon carbide (SiC) papers of 600, 1000, 1500, 2000 and 5000 grit to remove the oxide layer and was then ultrasonically degreased in acetone. The NiTi SMA and SS wires were fixed end-to-end in a self-constructed apparatus with a pure Cu interlayer. A Nd:YAG laser welding system (HKW-1050) with a wavelength of 1064 nm was used for welding. The optimized laser parameters used in the study were as follows: current, 75 A; pulse width, 5.0 ms; rate, 1.0 Hz; defocusing amount, 2.3 cm.

2.2 Film preparation

The CAWs were used as sputtering substrate which were ultrasonically cleaned in 99.5% acetone, 99% ethanol and deionized water for 30 min, and were dried under nitrogen flow. TiO₂ nanoparticle thin films were deposited by RF magnetron sputtering method, with 99.99% ceramic TiO₂ as the target. A mixture of Ar and N₂ (gas flow ratio of Ar to N₂ was 30 : 1) was used as the sputtering gas for N-doped TiO₂ film and pure Ar was used as the sputtering gas for pristine TiO₂ film. The RF sputtering power was 120 W. The background temperature was 300 °C and the sputtering time was 180 min. The chamber pressure was pumped to 1.0×10^{-4} Pa, and the working pressure was maintained at 1.0 to 1.2 Pa during the film growth. The target was pre-sputtered for 10 min, in order to remove any contamination. After sputtering, all samples were annealed in air at 450 °C for 2 h.

2.3 Characterizations

X-ray diffractometer (D8, Bruker, Germany) with Cu K α radiation ($\lambda = 1.5406$ Å) was employed to detect the crystal structure of the coating materials. The X-ray photoelectron spectroscopy (XPS) was measured by a Thermo ESCALAB 250 instrument, coupled with Al K α radiation (VG Scientific, UK). The surface morphology of the sample was observed by scanning electron microscope (SEM S-4800 EDAX AMETEK, USA) with an energy dispersive spectrometer (EDS) analyzer (Genesis-2000 AMETEK, USA).

2.4 Electrochemical measurements

Electrochemical tests were performed by CHI820 electrochemical workstation (CH Instruments, Shanghai, China) with artificial saliva (AS) as the electrolyte. The composition of AS



Table 1 Composition of the modified Fusayama AS

Component	Concentration (mg L ⁻¹)
NaCl	400
KCl	400
CaCl ₂ ·2H ₂ O	795
KSCN	300
NaHPO ₄ ·H ₂ O	690
Na ₂ S·9H ₂ O	5
Urea	1000

electrolyte solution is listed in Table 1. The counter electrode was a rectangular platinum plate, and the reference electrode was a saturated calomel electrode and the CAW samples were used as working electrode. Rubbing down all the samples with SiC papers to guarantee consistent surface roughness, then they were coated with epoxy resin, with an exposed active surface area of $20 \times 0.64 \text{ mm}^2$. After an initial delay of 60 min to achieve the steady state, the scanning rate was 1 mV s^{-1} , starting from -1 V per SCE. The potentiodynamic polarization measurements were taken between -1000 and $+1000 \text{ mV}$. Electrochemical impedance spectroscopy (EIS) was tested by an AUTOLAB PGSTAT302N workstation.

2.5 Antibacterial test

Streptococcus mutans were chosen as standard bacteria. Separated suspensions of the bacteria were prepared at a concentration of 1.5×10^6 colony-forming units (CFU) per mL each. $100 \mu\text{L}$ of each suspension was added to 3 mL of liquid medium in tube to make the suspension of bacteria. The suspensions were divided into four groups. The first test substrate (coated with N-doped TiO₂ thin film), the second test substrate (coated with the TiO₂ thin film) and control substrate (pristine substrate) were placed into these suspensions, which were incubated at 37°C for 24 h under visible light irradiation a thermostatic light incubator and a xenon lamp (MICROSOLAR 300, Beijing Perfect light Co., Ltd) was used as the visible light source in our research, with all the UV light below 420 nm being removed with a glass filter. And the substrates were kept 20 cm away from the light source. After incubation and vigorously shaking, $100 \mu\text{L}$ of each suspension was taken and added to 10 mL of sterile saline for dilution. Then, $100 \mu\text{L}$ of the diluted solution was added to cooled (55°C) solid medium, which was vigorously shaken and then poured into a Petri dish. The solid medium was cultured at 37°C for 24 h, after which the numbers of colonies on each plate were counted. This procedure was repeated three times.

3. Results and discussion

3.1 Structural characterization

The CAW consists of SS and NiTi arch wire by laser welding *via* a pure copper interlayer. Fig. 1a shows SEM image of the microstructure of CAW welding zone. The surface of the welding zone is very flat and smooth without obvious defects and welding marks. According to EDS elemental mapping images in

Fig. 1b–f, the Ni and Ti contents from top to bottom gradually reduced, whereas Fe and Cr element increased. Cu is the main element in the welding area; thus Cu element spreads homogeneously, while Ni, Ti, Fe and Cr are distributed gradually and there is no obvious boundary between them which indicated that five types of metals were fused together under the action of laser, as shown in Fig. 1b–f. So the laser-welded composite arch wire is sturdy and durable in this paper.

XRD was used to detect the crystal structure and calculate the crystal size of the coating materials, *i.e.* TiO₂ and N-doped TiO₂ nanoparticle thin films. After calcination at 450°C for 2 h, their XRD patterns were measured and shown in Fig. 2. The diffraction peaks of TiO₂ and N-doped TiO₂ are in good agreement with those of the standard anatase TiO₂, indicating TiO₂ and N-doped TiO₂ nanoparticle thin films have high crystallinity and an anatase structure. According to the Scherrer equation, the crystal sizes of TiO₂ and N-doped TiO₂ thin films were calculated to be around 16 nm, confirming that TiO₂ and N-doped TiO₂ thin films are composed of TiO₂ nanoparticles.

The XPS was used to identify the various elements in the thin films. Fig. 3a shows two survey XPS spectra of TiO₂ and N-doped TiO₂ nanoparticle thin films. The XPS spectrum of pristine TiO₂ is composed of Ti, O and a contaminated carbon. From high-resolution XPS spectra of Ti 2p (see Fig. 3b), the full width at half maximum of Ti 2p for N-TiO₂ was remarkably greater than that of pure TiO₂ which indicated that the chemical composition and structure of N-TiO₂ are more complex than those of pure TiO₂. The nitrogen doped into TiO₂ may contribute it. But chemical shift of Ti 2p is not significant, indicated that the main component of thin film was still TiO₂. Fig. 3c shows the Ti 2p XPS peaks of the N-doped TiO₂ films. The Ti 2p peak was located at 458.7 eV and 464.4 eV, which resulted from Ti⁴⁺ in TiO₂.^{26,27} Two other small peaks of TiO₂ was observed at 457.8 eV and 463.5 eV which are attributed to the formation of Ti³⁺, because the outer electron of Ti is changed and oxygen atoms are replaced, indicating that oxygen vacancies existed in the sample.^{28,29} This result also means that oxygen vacancies may exist in N-doped TiO₂ thin film. Fig. 3d shows the chemical binding state of N 1s for N-doped TiO₂ thin film. The XPS peak is observed at 400.1 eV, which can be attributed to interstitial N.³⁰ According to the recent literature reports, the XPS peak locating at between 396 and 397.5 eV is ascribed to Ti–N–Ti linkage which is due to doping of nitrogen atoms into a TiO₂ lattice.^{31–34} Therefore, the XPS peak at 396.1 eV can prove that the nitrogen element has been successfully incorporated into the TiO₂ thin film by Ti–N–Ti linkage. In addition, the nitrogen content in the samples is about 1.34%, as calculated by XPS.

SEM surface morphologies of TiO₂ and N-doped TiO₂ coated CAWs are shown in Fig. 4a and b. TiO₂ coated and N-doped TiO₂ coated CAWs exhibited different film morphologies, but both of them are very uniform. Their thicknesses are about 300 nm. The morphology of N-doped TiO₂ coated CAW was changed significantly after N doping. Fig. 4a shows TiO₂ coating, and its surface revealed a clear smooth polyhedral structure. The N-doped TiO₂ coating has significantly different membrane structure, and TiO₂ particles have “blooming flower-like” appearance,³⁵ as shown in Fig. 4b. N-Doped TiO₂ coated CAW is more uniform and dense than pristine TiO₂ coated CAW.



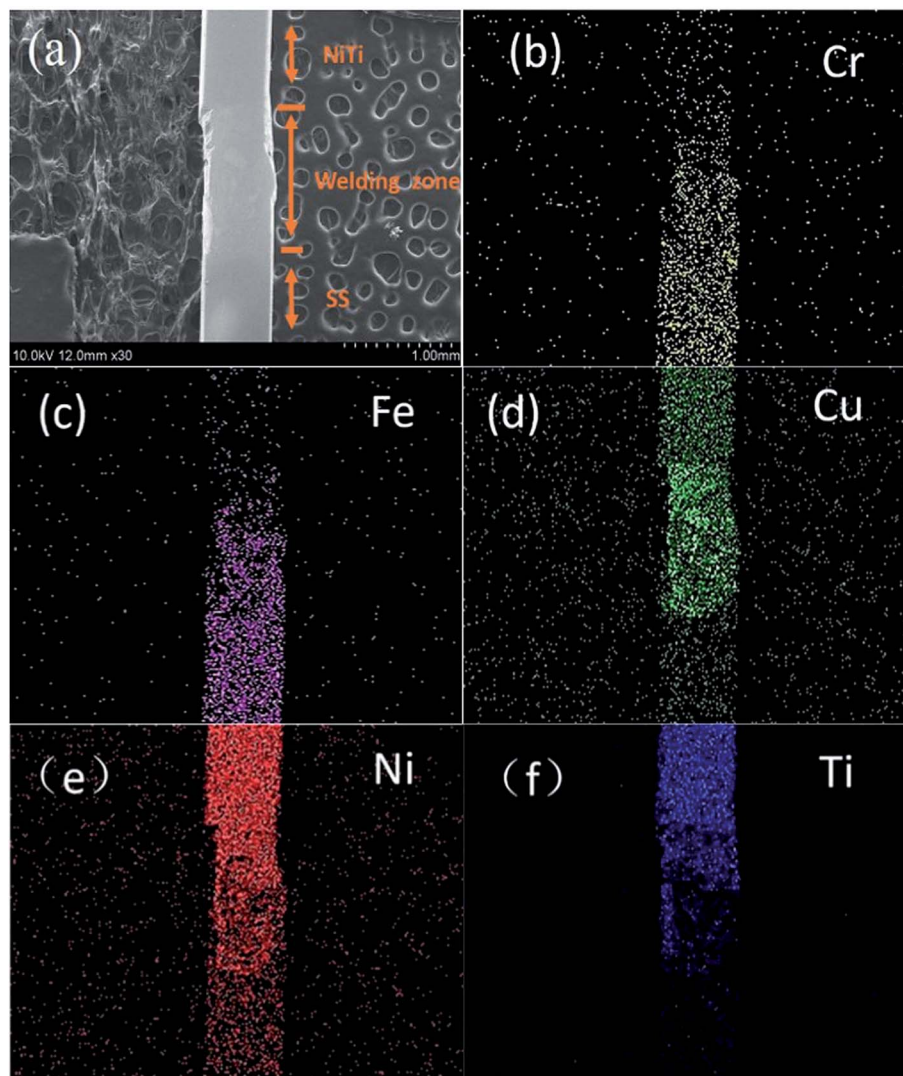


Fig. 1 SEM image (a) and the corresponding EDS mapping images of laser-welded composite arch wire (b–f).

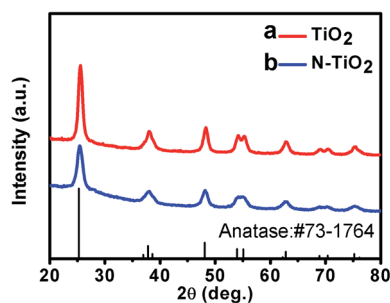


Fig. 2 XRD patterns of TiO_2 (a) and N-doped TiO_2 (b) coated CAW samples.

3.2 Corrosion resistance text

The electrochemical performance of CAW is one of the most important performances when it is used as biomedical orthodontic material.³⁶ So the corrosion behaviors of the different samples in the artificial saliva environment are evaluated. Fig. 5

presents three typical polarization curves of unmodified CAW, TiO_2 coated CAW and N- TiO_2 coated CAW in AS solutions. In the lower potential of the polarization curve for pristine CAW in AS, vertical region represents the generation of hydrogen, and the smaller part of the slope represents the oxygen consumption. During the process of electrochemical experiments, the most of the metal will form an oxide film on the surface during the corrosion process. If the formed oxide film can prevent the further corrosion, it is called the passive process. But when the voltage continues to increase, the local passive oxide film will break, so that the obvious corrosion phenomenon occurs in the local area, this phenomenon is called pitting, and the potential of the occurrence is pitting potential (E_{pit}). When the corrosion voltage exceeded the pitting potential, the current density increased rapidly, which indicated that the oxide film on the surface of the sample was broken, and the new corrosion happened. As shown in Fig. 5, for the polarization curve, all the cathode sections contain two sections, but in the anode polarization curve CAW showed a significant passive behavior, with



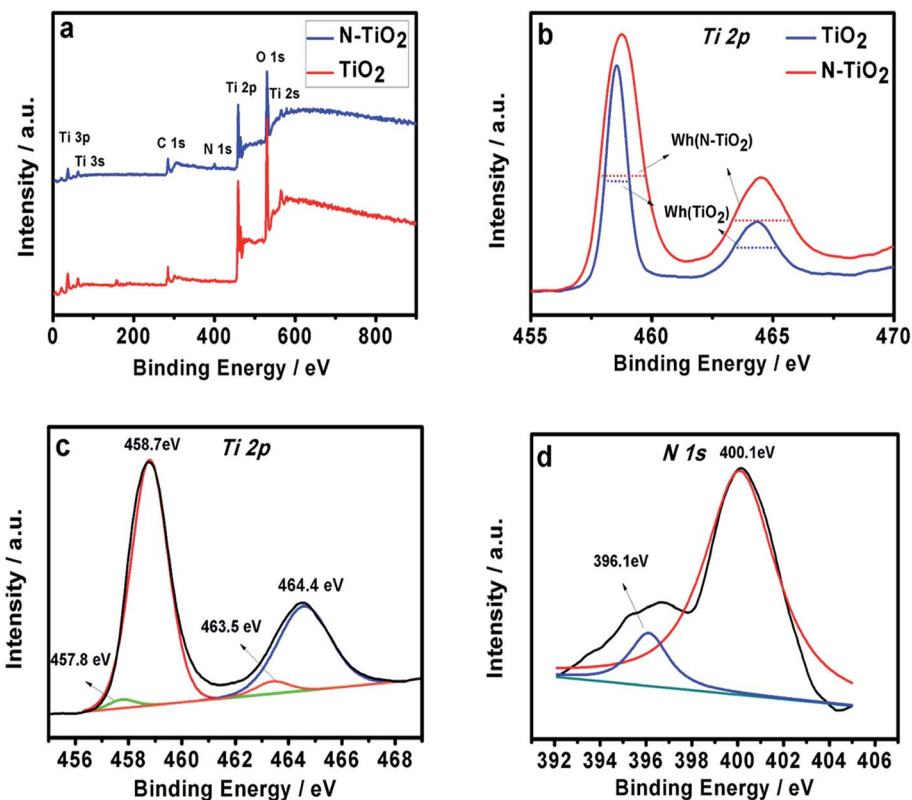


Fig. 3 (a) XPS survey spectra of TiO_2 and N-doped TiO_2 . High-resolution XPS spectra of Ti 2p peaks of TiO_2 (b) and N-doped TiO_2 as well as Ti 2p (c) and N 1s (d) peaks of N- TiO_2 .

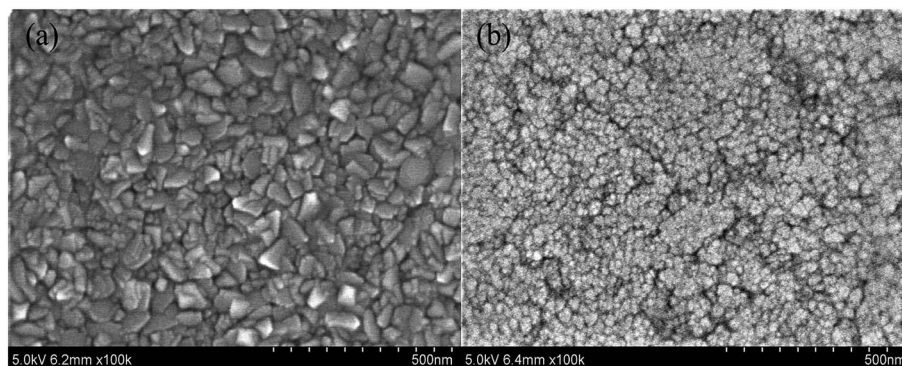


Fig. 4 SEM surface images of CAWs coated with TiO_2 thin film (a) and N- TiO_2 thin film (b) after annealing at 450 °C.

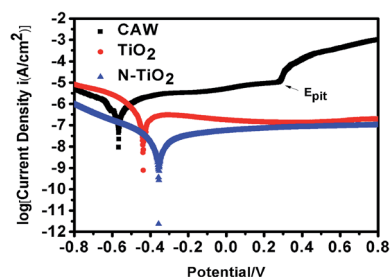


Fig. 5 Potentiodynamic polarization curves of pristine CAW, TiO_2 coat-CAW and N- TiO_2 coated CAW.

a pitting potential (E_{pit}) of 0.284 V. When the corrosion voltage exceeded 0.284 V, the current density of CAW increased rapidly, which indicated that the oxide film on the surface of CAW was broken, and the new corrosion occurred. But TiO_2 and N- TiO_2 coated CAWs have no obvious anode area passive behavior and local pitting corrosion behavior was not observed. It indicated that TiO_2 and N- TiO_2 coated CAWs have no oxide film rupture during electrochemical corrosion.

The corrosion process can be also evaluated by the corrosion potential (E_{corr}) and corrosion current density (I_{corr}) from the polarization curve. Generally, the higher corrosion potential (E_{corr}) and lower corrosion current density (I_{corr}) indicate that



the sample has a low corrosion rate and good corrosion resistance.^{37,38} The corrosion behaviors of unmodified CAW as well as TiO₂ and N-TiO₂ CAW in artificial saliva were studied in detail by polarization and EIS techniques. There is a linear relationship in the strong polarization region between the current and voltage, the slope of the curve is called the Tafel slope. Corrosion potential and corrosion density can be obtained by the Tafel extrapolation method. The intersection of cathode and anode slope is the corrosion current value.

From Table 2, the E_{corr} values of uncoated CAWs and N-TiO₂ coated CAWs were increased from -0.5659 to -0.3557 V, and the positive shifting of the E_{corr} clearly shows an increase in corrosion resistance for TiO₂ and N-TiO₂ coated samples. Another indicator of corrosion behavior is the corrosion current density (I_{corr}), which is proportional to the rate of corrosion. The corrosion current density of CAW was measured as $5.620 \mu\text{A cm}^{-2}$, while the corrosion current density of TiO₂ and N-TiO₂ coatings were measured as $1.722 \mu\text{A cm}^{-2}$ and $0.8161 \mu\text{A cm}^{-2}$, respectively. The high corrosion potential means better corrosion resistance. Since corrosion current density is proportional to corrosion rate, lower corrosion current density also shows better corrosion resistance. Therefore, we can conclude that the corrosion rate of N-TiO₂ coated CAW is lower than those of pristine CAW and TiO₂ coated CAW. This is due to its more dense surface coverage, as shown in Fig. 4; thus N-TiO₂ coating is more effective to prevent the contact between the electrolyte and the base metal material, and then hinder the occurrence of corrosion reaction. This result is also consistent with that of in Table 2. The protective rates of TiO₂ and N-TiO₂ coatings were 69.36%, 85.47% respectively. The protection rate (η) in Table 2 is calculated by the follow equation:³⁹

$$\eta = (I_{\text{corr}}^0 - I_{\text{corr}}) / I_{\text{corr}}^0 \times 100\% \quad (1)$$

where I_{corr}^0 represents the corrosion current density of CAW, and I_{corr} represents the current density of coating samples. According to the results of polarization curves, the coating CAWs have more positive corrosion potential than unmodified CAW. Furthermore, N-doped TiO₂ coating is better than pristine TiO₂ coating.

TiO₂ thin films exhibit ceramic based structures, which have low electrical conductivity and lower charge transport, thus improving the electrochemical barrier properties. Therefore, lowering the electron conductivity can delay the electrochemical process, and this is considered to be the electrochemical mechanism of corrosion resistance of TiO₂ coated samples. In general, for a protective coating, in addition to strong corrosion resistance, the coating materials should be also separated completely the protecting base material and corrosive medium.

The electrochemical reaction does not occur at the interface between the coating and the substrate.

TiO₂ nanoparticle coating thin films were prepared by magnetron sputtering. The coating surface is flat and compact with good sealing performance, which can effectively prevent corrosive medium to contact substrate material. At the same time from observation of the anodic passivation region of polarization curves, the formation of a dense passivation layer on the substrate surface limits the occurrence of the anodic reaction. N-TiO₂ coating can more effectively enhance the corrosion resistance than pure TiO₂ coating.

The corrosion current density obtained by Tafel extrapolation is not very accurate, and high speed scanning error or calculation of slope may lead to cathodic or anodic curve distortion, which will cause some errors to slope extrapolation corrosion current value. It can only roughly estimate the corrosion rate of samples, as well as the relative ranking of corrosion resistance. Thus the corrosion properties of different samples are further analyzed by the electrochemical impedance spectroscopy (EIS) diagrams. Fig. 6 shows the Nyquist, Bode-phase and Bode-magnitude plots for unmodified, TiO₂ coated and N-doped TiO₂ coated CAWs after immersing in artificial saliva for 30 min at open circuit potential. The high frequency capacitance behavior of the coating sample indicates the corrosion resistance, and the low frequency capacitance behavior indicates the capacitance process at the coating material/metal interface. According to the Nyquist plots, all of the plots are simple semicircular, indicating that each sample has a single time constant. The curves appear to have a similar shape, but they are very different in their sizes, which represents the number of charge transfer in the corrosion process. It also may explain the similar process occurring in all the coatings, but each case is carried out under different material surfaces. Compared with the unmodified CAW, the TiO₂ and N-doped TiO₂ coated samples have a larger capacitance loop. It can be observed from the Nyquist plots the capacitance loop size of N-doped TiO₂ sample is about 2 orders of magnitude relative to CAW, is about 1 order of magnitude relative to TiO₂ group. D. Y. Yu *et al.* pointed out that the larger the capacitance loop, the better the corrosion resistance of the layer.⁴⁰ Therefore, TiO₂ and N-doped TiO₂ coating can significantly improve the corrosion resistance of CAW, especially for the N-TiO₂ coating. The improvement of the corrosion resistance of the material is due to the formation of a stable structure, effectively preventing the penetration of the electrolyte. The TiO₂ nanoparticle coating produced by magnetron sputtering can effectively reduce the porosity of the material surface and form a sealed barrier, which can prevent the electrolyte to etch the substrate material. Thus, the permeation and corrosion of the electrolyte can be delayed, and the charge transfer resistance can be further increased which can enhance the barrier property for the corrosion protection of the substrate material. Therefore, low corrosion rates and high impedances obtained by electrochemical measurements confirmed that the TiO₂ nanoparticle coating offered good barrier properties.

Besides Bode plots, phase angle plots are also provided to analyze the corrosion process. The high impedance modulus in

Table 2 Comparison of E_{corr} , I_{corr} and η on different surfaces

Sample	E_{corr} (mV)	I_{corr} ($\mu\text{A cm}^{-2}$)	η (%)
CAW	-565.9	5.620	—
TiO ₂	-434.1	1.722	69.36
N-TiO ₂	-355.7	8.161×10^{-1}	85.47



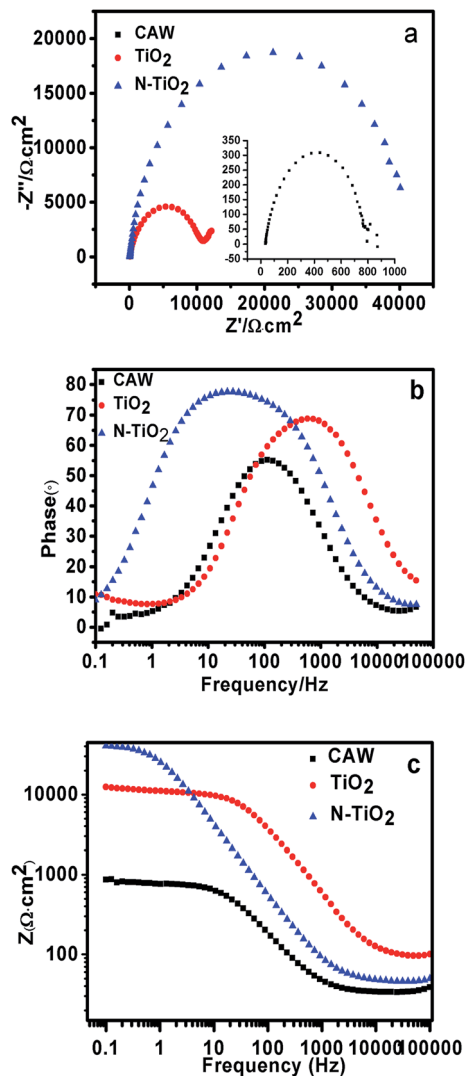


Fig. 6 EIS data of CAW and coated samples in AS after immersion for 30 min (a) Nyquist plots, (b) Bode phase plots and (c) Bode magnitude plots.

the low frequency region of the Bode diagram represents the barrier performance of the system.⁴¹ Low frequency high impedance modulus and one time constant revealed that the film has effective barrier property.^{42–44} In the Bode-phase plots, the higher phase angle and the low frequency impedance indicate the better corrosion resistance and passivity of the stable structure. At the same time it is believed that inert and homogenous film deposition can lead to higher phase angle and corrosion resistance. In Fig. 6b and c, the coated samples have higher phase angle and low frequency impedance than CAW, especially for N-TiO₂.

Simple equivalent circuit of EIS electrochemical model was shown in Fig. 7. It was able to analyse impedance data accurately and understand corrosion mechanism further. By using the software Nova1.8 to fit equivalent circuit model, we found that it is the most suitable one for our results. The corresponding equivalent circuit principle diagram is presented in Fig. 7b. In this circuit, R_{ct} is the charge transfer resistance and

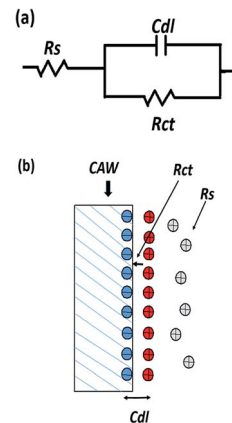


Fig. 7 (a) Equivalent circuit used for fitting the experimental data of TiO₂ and N-TiO₂ coatings, (b) corresponding equivalent circuit principle diagram.

C_{dl} is the double layer capacitance that the corrosion reaction occurred at the coating material/substrate interface on the surface of CAW. In this case the most areas are well covered without corrosion, while R_s is the resistance of the solution. The data associated with the model are listed in Table 3.

As known, a good protective coating should be close to the base material as much as possible, so that the base material can be separated from corrosion medium completely. However, even a very dense coating, there also may exist some micro/nano-porous structures. These small holes will form the channels from each other which allow small molecules of corrosive medium to pass and to corrode base materials. This kind of corrosion also occurred in our case. Although the nanoparticle coatings produced by magnetron sputtering are uniform and compact, artificial saliva corrosion medium will also corrode the base materials to some extent. Therefore, the coating coverage rate is very important to the corrosion resistance of CAW.

The definition of conductor resistance shows that R_{ct} value is inversely proportional to the surface area of the sample. We assume that the base material is the size of the unit and θ is the coverage rate, while the corroded area of samples by artificial saliva corrosion medium through micropore would be $(1 - \theta)$. Then the R_{ct} of bare CAW and coated samples can be written as:

$$R_{ct}^0 = A^0/I \quad (2)$$

$$R_{ct} = A/(1 - \theta) \quad (3)$$

where R_{ct}^0 is the charge transfer resistance of CAW and R_{ct} is the one of coated samples. A and A^0 are factors. Assuming A and A^0 are the same, then the coverage rate θ can be derived as:

Table 3 Comparison of R_s , R_{ct} , C_{dl} and θ on different surfaces

Sample	R_s (Ω cm ²)	R_{ct} (k Ω cm ²)	C_{dl} (μ F cm ⁻²)	θ (%)
CAW	33.5	758	6.34	—
TiO ₂	40.5	10 400	2.713	92.7
N-TiO ₂	46.1	43 800	2.159	98.26



$$1 - \theta = R_{ct}^0/R_{ct} \quad (4)$$

Yuan *et al.* have already used eqn (4) to calculate the coverage rate of the coated samples.⁴⁵ From Table 3 the coverage rate of TiO₂ and N-TiO₂ coated samples is 92.70% and 98.26%, respectively. In this case the charge transfer resistance (R_{ct}) of N-TiO₂ will increase, as shown in Table 3. R_{ct} of N-TiO₂ was 43 800 kΩ cm², which is much larger than those of CAW and TiO₂-coated samples which were 10 400 kΩ cm² and 758 kΩ cm², respectively.

According to comprehensive analysis of EIS results and Tafel curves, we can conclude that the coating TiO₂ can improve the corrosion resistance of the wire. It is due to two points. One is that titanium dioxide is a kind of ceramic material, and its low conductivity can reduce the charge transport, thus improving the electrochemical barrier performance.⁴⁶ The other is that the TiO₂ film can effectively block the contact reaction between the substrate and the electrolyte.⁴⁶ Meanwhile, we found the N-doped TiO₂ has a better property. Lindgren *et al.* reported that N-doping can diminish the grain size of TiO₂ and decrease the roughness of TiO₂ thin film.⁴⁷ As shown in Fig. 4a and b, N-doped TiO₂ thin film is more dense and compact, and its surface is smoother than that of un-doped TiO₂ film because of its smaller grain size. The smaller the particle can decrease inter-particle distance and pore size of the films. The coverage rates of TiO₂ and N-doped TiO₂ films showed in Table 3 also prove it. Thus the more compact structure and higher coverage rate of N-doped TiO₂ film can reduce micropore and seal

channel better, which are more effective to prevent the contact between the electrolyte and the base material, and then hinder the occurrence of corrosion reaction. R_{ct} of N-doped TiO₂ is much larger than that of TiO₂, also proves this point.

After electrochemical experiments, the color of the welding point of CAW becomes black and rough which can be seen by the naked eye. Fig. 8 shows corrosion image of the samples after electrochemical corrosion. Fig. 8a and b are the corrosion images of unmodified CAW after electrochemical etching, and a large area of corrosion can be found on the surface. It was observed from Fig. 8b that not only the corrosion area is large but also the corrosion depth is deep and welding points are severely damaged. Some black areas and corrosion products on the un-doped TiO₂ coated surface can be seen in Fig. 8c. This indicated that the un-doped TiO₂ coated surface has also undergone mild corrosion behavior. As expected, for N-doped TiO₂ coated CAW, its surface is still flat without obvious corrosion pits and corrosion product after electrochemical process, as demonstrated in Fig. 8d. The SEM results are in agreement with the electrochemical polarization curves and EIS results which indicate good corrosion resistance of N-doped TiO₂ coating.

3.3 Antibacterial test

Fig. 9 and 10 show the antimicrobial properties of each sample. It can be clearly seen that the group tested with N-doped TiO₂ thin film shows the most effective antimicrobial ability with

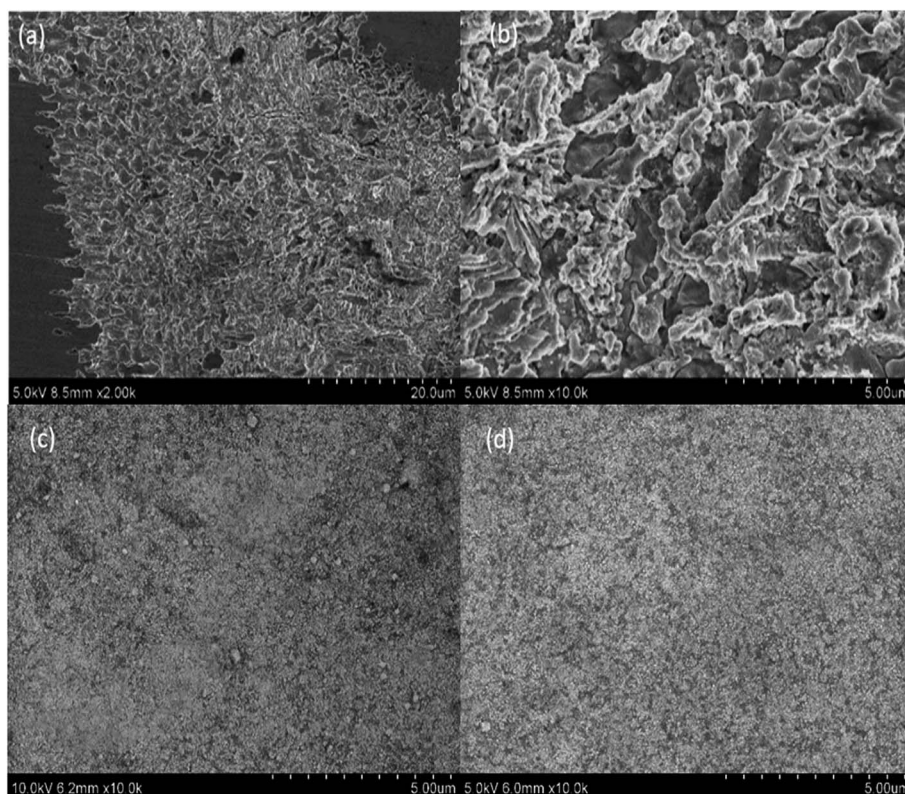


Fig. 8 SEM surface images of (a, b) CAW, (c) TiO₂-coated CAW and (d) N-TiO₂-coated CAW after potentiodynamic polarization tests.



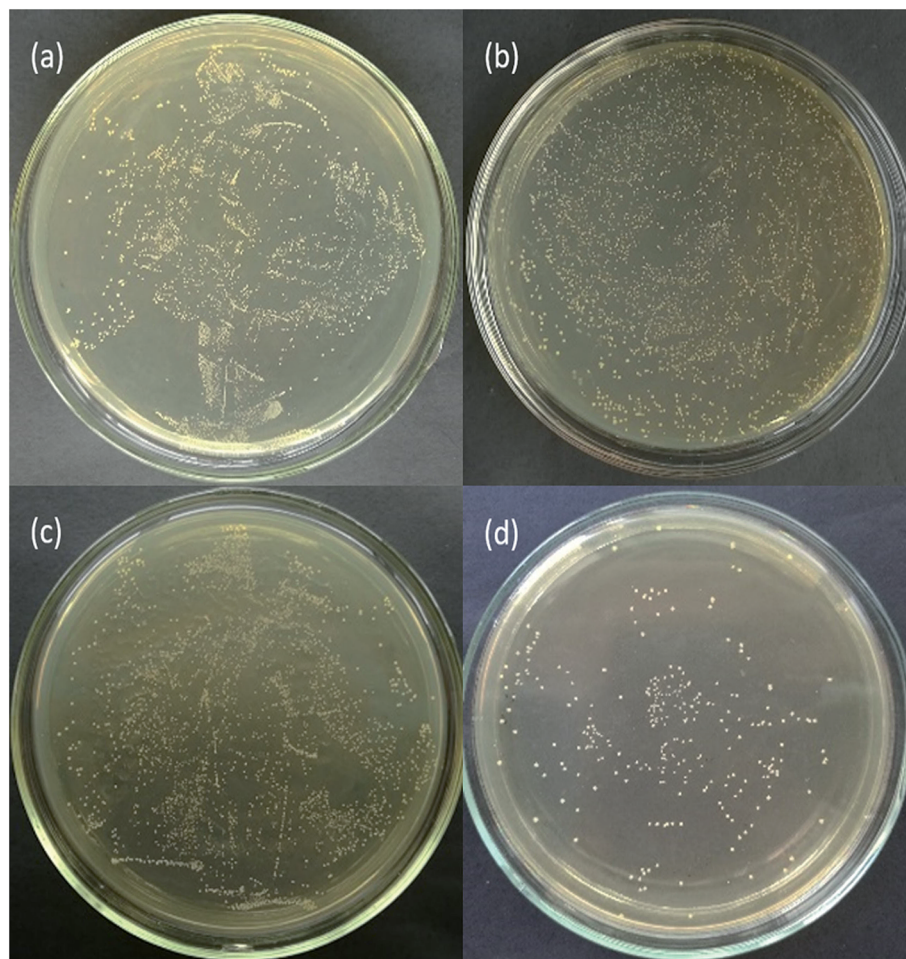


Fig. 9 Photographs of antimicrobial resistance of *Streptococcus mutans*: (a) control group, (b) CAW, (c) TiO₂ and (d) N-TiO₂.

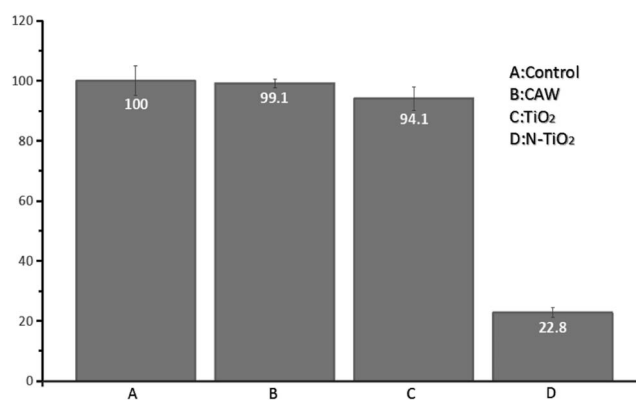


Fig. 10 Antibacterial effects of the samples against *Streptococcus mutans*.

a rate of 87.2%. The group filmed only with TiO₂ shows a rate of 5.9% which indicates no bactericidal effect according to standard reduction of bacteria criteria (standard reduction of bacteria criteria: less than 0–20% reduction indicates no bactericidal effect; 20–50% reduction indicates a low bactericidal effect; between 50% and 70% reduction indicates an

expressive bactericide; and >70% reduction is considered a powerful bactericidal effect³⁵). Shuai Cao *et al.* have found that the photo-catalytic activities of N-doped TiO₂ substrates were much higher than that of un-doped TiO₂ substrates.³⁵ This finding coincides with our result. The main bactericidal mechanism of TiO₂ was hydroxyl radical attack and the lipid peroxidation reaction.⁴⁸ Huang *et al.* pointed out that when TiO₂ photo-catalytic surface first contact with bacteria, the wall of the cell and the cytoplasmic membrane would be damaged, which would lead to the free efflux of intracellular contents.⁴⁹ However, TiO₂ only produces hydroxyl radical under irradiation with UV light. It has been demonstrated that N-doped TiO₂ is a visible-light-sensitive photo-catalyst. In our experiments, the UV light has been filtered, thus the bactericidal rate of TiO₂ is rather low whereas the bactericidal rate of N-doped TiO₂ filmed substrates is obvious under visible light irradiation.

Many patients might not be able to maintain oral hygiene during orthodontic treatment. This may cause the increase of the plaque accumulation around orthodontic arch wire, which leads to an increase of the possibility of caries. *Streptococcus mutans* is one of the main compositions of dental plaque. It can make use of sucrose for substrate to synthesize extracellular glucan, fructan and intracellular polysaccharide. Glucan which



Table 4 The growth rate and the safety level of hamster fibroblasts in different time^a

Group	A value			RGR (%)			Safety level		
	24 h	72 h	120 h	24 h	72 h	120 h	24 h	72 h	120 h
Negative control	0.449 ± 0.035	0.673 ± 0.028	1.147 ± 0.035	100	100	100	0	0	0
Positive control	0.053 ± 0.032**	0.094 ± 0.030**	0.229 ± 0.028**	12	14	20	4	4	4
CAW	0.431 ± 0.026	0.609 ± 0.018**	1.044 ± 0.035*	96	90	91	1	1	1
TiO ₂	0.452 ± 0.030	0.66 ± 0.019	1.149 ± 0.030	101	98	101	0	1	0
N-TiO ₂	0.474 ± 0.034	0.679 ± 0.025	1.169 ± 0.080	105	101	102	0	0	0

^a *: $p < 0.05$ vs. NC, **: $p < 0.01$ vs. NC.

is the important cause of caries can mediate bacteria adhesion and promote the formation of plaque, causing more bacteria to accumulate. *Streptococcus mutans* is also the main reason responsible for caries. It can ferment a variety of carbohydrates rapidly to produce a large quantity of acid which will make the local pH less than 5.5. Due to the protection of dental plaque, the buffer action of saliva could be easily avoided. Thus the decreased local pH can maintain for quite a long time which would cause local hard tissue demineralization. The caries will appear when the demineralization is severe enough. In our research the use of N-doped TiO₂ coated arch wire with high antibacterial activity under visible light irradiation would kill the most of *Streptococcus mutans* which is the main reason responsible for caries as stated above. This result indicated that N-doped TiO₂ modified compound arch wire has a high potential in the clinical practice.

3.4 Cytotoxicity

In order to realize the clinical application, the cytotoxicity of CAW, TiO₂ and N-doped TiO₂ coated CAWs was respectively tested. The results are shown in the Table 4. The detailed procedures of the cytotoxicity tests are provided in ESI.† The relative growth rate (RGR) equals to the absorbance values of experimental group/the absorbance values of control group × 100%. According to the calculated RGR shown in Table 4, the cytotoxicity levels of the three types of the materials are evaluated. Level 0: RGR ≥ 100%; level 1: 75% ≤ RGR < 100%; level 2: 0% ≤ RGR < 75%; level 3: 25% ≤ RGR < 50%; level 4: 1% ≤ RGR < 25%; level 5: RGR < 1%. Level 0 and level 1 are safe levels, while level 3–5 are unsafe levels. When the cytotoxicity is level 2, it should be further evaluated according to the status of the cell culture.

According to the literature report,⁶ the copper content should be less than 25% in the oral materials, although a small amount of copper can promote cell activity. The concentration of copper ions is greater than 10.85 mg L⁻¹ for more than 24 h, which will show the cytotoxicity. According to the results shown in Table 4, the safety level of CAW modified with N-doped TiO₂ was 0 even though it was soaked in the solution for 120 h. This result suggested that N-doped TiO₂ nanocrystal thin film can effectively prevent the release of Cu and N-doped TiO₂ modified CAW has good biological safety.

4. Conclusions

In summary, TiO₂ and N-doped TiO₂ nanocrystal thin films were deposited on the surface of CAW by RF magnetron sputtering method. The morphologies and structures of the coated materials were systematically characterized and their corrosion resistance in artificial saliva was electrochemically assessed. As expected, the TiO₂ and N-TiO₂ coated CAWs can greatly improve the corrosion resistance of pristine CAW in artificial saliva. It was demonstrated that N-doped TiO₂ film showed much better anticorrosion ability than TiO₂ film. Additionally, the antibacterial experiments confirmed that N-doped TiO₂ group exhibited higher antibacterial activity against *Streptococcus mutans* with a rate of 87.2% under visible light irradiation. Therefore, the N-doped TiO₂ coating can improve corrosion resistance and antibacterial properties of CAW obviously. The results of the cytotoxicity test demonstrated that N-doped TiO₂ coated CAW has a high potential in the practical dental treatments.

Conflicts of interest

There are no conflicts to declare.

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