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Mechanofluorochromic properties of fluorescent molecules based on a dicyanomethylene-4H-pyran and indole isomer containing different alkyl chains via an alkene module[†]

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A series of 3-, 4-, and 5-position indole-substituted dicyanomethylene-4H-pyran (I3PM, I4PM, and I5PM) derivatives with different alkyl chains were synthesized and their mechanofluorochromic (MFC) properties were investigated. Compared with the alkyl chains, the isomerization of the indole unit plays a more important role in regulating the MFC properties of these compounds. The increase of the alkyl chain length and the isomerization of the alkyl chain are favorable to the effective mechanofluorochromism of the I3PM derivatives, however, the former exhibits the opposite effect on that of the I4PM derivatives. The opposite alkyl length-dependent MFC behavior is closely related to the degree of distortion of the molecular conformation, namely the more twisted the molecular conformation, the more pronounced the MFC phenomenon. In particular, the introduction of the alkyl chains is completely detrimental to the MFC properties of the I5PM derivatives. X-Ray diffraction experiments indicate that the MFC properties of some of these compounds should be mainly attributed to the transformation from a crystalline state to an amorphous state because the weak intermolecular interactions and loose molecular packing modes in their solid-state samples can be easily damaged under external force stimuli as revealed by the single crystal X-ray diffraction analysis.

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Introduction

Mechanofluorochromic (MFC) organic molecules are a class of “smart” materials whose solid-state fluorescence colors can display obvious changes in response to external force stimuli, and have received considerable interest owing to their potential applications in the fields of fluorescent sensors,¹ data storage,² security inks,³ optoelectronic devices,⁴ etc. In general, to make an organic fluorescence molecule have MFC properties, first the molecule must be capable of emitting fluorescence in the solid state, and its solid morphology must be able to undergo a transition from a crystalline to an amorphous state or from a crystalline state to another crystalline state under pressure stimuli.⁵ Although the development new fluorescent molecular structures and the structural modification of traditional fluorescent molecular frameworks to meet the above two conditions are both crucial for obtaining new MFC materials, the latter

looks more alluring and more easily accepted by researchers. Generally speaking, the introduction of alkyl chains to the conjugated skeleton of organic fluorescence molecules is to mainly tune the solubility of the materials, but some recent studies show that alkyl chains can also influence the intermolecular interactions and molecular packing modes of the materials in the solid state, and thereby play a crucial role in tuning and altering their MFC properties.⁶ Although the change of the length of alkyl chain can greatly affect MFC behaviors of some fluorescent molecules, this effect is very complicated. For instance, longer alkyl chains are favorable to the more remarkable MFC phenomena of 9,10-bis[(9,9-dialkylfluorene-2-yl)vinyl-2]anthracenes,⁷ 9,10-bis(*N*-alkylphenoxythiazin-3-yl-vinyl-2)anthracenes,⁸ and 9,10-bis(*p*-alkoxystyryl)anthracenes,⁹ whereas obviously disadvantageous to those of 9,10-bis[(*N*-alkylcarbazol-3-yl)-vinyl-2]anthracenes¹⁰ and 9,10-bis(*N*-alkylindole-3-yl-vinyl-2)anthracenes,¹¹ and 2,5-dialkylcarbazole-substituted terephthalate derivatives.¹² Moreover, the branched alkyl chains are reported to be advantageous to obtain high-contrast MFC compounds for 9,10-bis(butoxystyryl)anthracenes.¹³ Additionally, some other ways have also proven to be very convenient in regulating the MFC properties of the fluorescent molecules, such as the introduction of appropriate aromatic substituents¹⁴ and heavy atoms¹⁵ as well as changing the position of the

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substituent.¹⁶ But so far the influence of the structure of the molecule on its MFC property is still not very clear and the regular pattern is hard to sum up, which is worthy of further study for the development of MFC-active materials based on the traditional fluorescent molecular frameworks.

4H-Pyran is a traditional and excellent building unit for constructing donor- π -acceptor (D- π -A) fluorescent compounds in the applications of organic electroluminescent diodes, molecular logic gates, and fluorescent sensors.¹⁷ However, MFC materials based on 4H-pyran unit have rarely been reported until now. Very recently, we found that some indene-1,3-dionemethylene-4H-pyran derivatives containing different lengths of alkoxy groups showed obvious alkoxy length-dependent MFC properties and the longer alkoxy chain led to more remarkable MFC phenomena.¹⁸ Moreover, for naphthylethylene-substituted dicyanomethylene-4H-pyran derivatives and cyano(ethoxycarbonyl)methylene-4H-pyran derivatives, the position isomerism of the naphthylethylene end groups were found to significantly affect their solid-state fluorescence color changes under external stimuli.¹⁹ To our knowledge, the combined effects of two different factors, namely different alkyl groups and the isomerization of aromatic substituents on the MFC properties of the fluorescent molecules has not been reported. Herein, a series of indole-substituted dicyanomethylene-4H-pyran (IPM) derivatives with different alkyl chains and varying the linking position of the indole from 3-, 4- to 5-position at 4H-pyran ring, namely **I3PM**, **I4PM**, and **I5PM**, were synthesized. The combined effects of isomerization and length of the alkyl chains and position isomerization of the indole unit on the MFC properties of the target compounds were systematically investigated. The results show the isomerization of indole unit plays a more important role in regulating the MFC properties of these compounds compared with the alkyl chains. The isomerization and length of the alkyl chains exhibit different effects on the MFC properties of these fluorescent molecules when the substitution position of the indole unit is different. The increase of the alkyl chain length and the isomerization of the alkyl chain are favorable to the effective MFC properties of the **I3PM** derivatives, however, the former exhibits opposite effects on those of the **I4PM** derivatives. In particular, the introduction of the alkyl chains is completely detrimental to the MFC properties of the **I5PM** derivatives. The results signify that the choice of the alkyl chains and the isomerization of indole unit can be used to regulate the MFC properties of the IPM derivatives. X-Ray diffraction (XRD) experiments reveal that the reversible MFC properties of the IPM derivatives should be mainly attributed to the transformation between amorphous and crystalline states.

Experimental

Measurements and materials

NMR spectral measurements were performed on a Bruker DRX 500 NMR spectrometer. Mass spectrometry spectra were performed with a Finnigan DECA-30000 LCQ Deca mass spectrometer or a Finnigan MAT95XP mass spectrometer. Elemental analysis was recorded on an Elementar Vario MICRO analyzer.

Fluorescence measurements were carried out with a HITACHI F-7000 fluorometer. X-Ray diffraction (XRD) measurements were carried out with a Bruker D8 Advance X-ray diffractometer. Single-crystal X-ray diffraction measurements were carried out with a Bruker-Nonius Smart Apex CCD diffractometer with graphite monochromated Mo K α radiation. 2,6-Dimethyl-4-pyrone (**1**), malononitrile, 1H-indole-3-carbaldehyde (**3a**), 1H-indole-4-carbaldehyde (**4a**), 1H-indole-5-carbaldehyde (**5a**), and various alkyl bromides were purchased from commercial suppliers and used directly.

Synthesis of *N*-substituted indolecarboxaldehydes **3b–3d/4b–4d/5b–5d**

General procedure: Alkyl bromide (20.0 mmol) was slowly added to a solution of **3a/4a/5a** (2.47 g, 17.0 mmol) and KOH (1.43 g, 25.60 mmol) in DMSO (30 mL). The mixture was stirred overnight at room temperature, and then added water and extracted with dichloromethane. The combined organic layers were washed with water and brine, dried over anhydrous Na₂SO₄, and evaporated to dryness *in vacuo*. The residue was purified by column chromatography through silica gel with petroleum ether/ethyl acetate (v/v = 4/1) to give the pure *N*-substituted indolecarboxaldehyde. The characterization data for the intermediates **3b–3d/4b–4d/5b–5d** are listed as follows:

1-Butyl-1H-indole-3-carbaldehyde (3b). Yellow liquid (3.29 g), 95.9% yield. ¹H NMR (CDCl₃, 500 MHz): δ 9.95 (s, 1H), 8.31 (dd, 3J = 6.5 Hz, 4J = 2.5 Hz, 1H), 7.65 (s, 1H), 7.36–7.27 (m, 3H), 4.09 (t, J = 7.5 Hz, 2H), 1.84–1.79 (m, 2H), 1.35–1.31 (m, 2H), 0.93 (t, J = 7.5 Hz, 3H) ppm. ¹³C NMR (CDCl₃, 125 MHz): δ 184.2, 138.2, 137.1, 125.2, 123.6, 122.6, 121.8, 117.8, 110.0, 46.7, 31.5, 19.8, 13.4 ppm. Anal. calcd for C₁₃H₁₅NO: C, 77.58; H, 7.51; N, 6.96. Found: C, 77.99; H, 7.56; N, 6.91.

1-Isobutyl-1H-indole-3-carbaldehyde (3c). Yellow liquid (3.20 g), 93.3% yield. ¹H NMR (CDCl₃, 500 MHz): δ 9.96 (s, 1H), 8.29 (dd, 3J = 6.5 Hz, 4J = 3.0 Hz, 1H), 7.64 (s, 1H), 7.34–7.26 (m, 3H), 3.91 (d, J = 7.5 Hz, 2H), 2.24–2.16 (m, 1H), 0.92 (d, J = 6.5 Hz, 6H) ppm. ¹³C NMR (CDCl₃, 125 MHz): δ 184.4, 138.8, 137.4, 125.3, 123.7, 122.7, 121.9, 117.8, 110.2, 54.7, 29.0, 20.0 ppm. Anal. calcd for C₁₃H₁₅NO: C, 77.58; H, 7.51; N, 6.96. Found: C, 78.02; H, 7.55; N, 7.02.

1-Dodecyl-1H-indole-3-carbaldehyde (3d). White solid (4.93 g), 92.3% yield. ¹H NMR (CDCl₃, 500 MHz): δ 10.00 (s, 1H), 8.31 (dd, 3J = 6.5 Hz, 4J = 2.0 Hz, 1H), 7.70 (s, 1H), 7.38–7.29 (m, 3H), 4.15 (t, J = 7.0 Hz, 2H), 1.90–1.87 (m, 2H), 1.33–1.25 (m, 18H), 0.88 (t, J = 7.0 Hz, 3H) ppm. ¹³C NMR (CDCl₃, 125 MHz): δ 184.3, 138.1, 137.2, 125.5, 123.8, 122.8, 122.1, 118.0, 110.0, 47.3, 31.9, 29.7, 29.53, 29.46, 29.37, 29.27, 29.1, 26.8, 22.6, 14.0 ppm. Anal. calcd for C₂₁H₃₁NO: C, 80.46; H, 9.97; N, 4.47. Found: C, 80.92; H, 9.93; N, 4.44.

1-Butyl-1H-indole-4-carbaldehyde (4b). Dark yellow liquid (2.56 g), 74.6% yield. ¹H NMR (CDCl₃, 500 MHz): δ 10.25 (s, 1H), 7.62 (d, J = 7.5 Hz, 2H), 7.35–7.29 (m, 3H), 4.17 (t, J = 7.0 Hz, 2H), 1.85–1.79 (m, 2H), 1.35–1.30 (m, 2H), 0.94 (t, J = 7.5 Hz, 3H) ppm. ¹³C NMR (CDCl₃, 125 MHz): δ 192.9, 136.6, 131.2, 128.4, 126.9, 126.3, 120.6, 115.6, 101.3, 46.1, 32.4, 20.0,



13.5 ppm. Anal. calcd for $C_{13}H_{15}NO$: C, 77.58; H, 7.51; N, 6.96. Found: C, 77.29; H, 7.47; N, 6.92.

1-Isobutyl-1*H*-indole-4-carbaldehyde (4c). Dark yellow liquid (2.34 g), 68.2% yield. 1H NMR ($CDCl_3$, 500 MHz): δ 10.24 (s, 1H), 7.59 (d, J = 7.5 Hz, 2H), 7.32–7.28 (m, 3H), 3.94 (d, J = 7.5 Hz, 2H), 2.21–2.13 (m, 1H), 0.90 (d, J = 7.0 Hz, 6H) ppm. ^{13}C NMR ($CDCl_3$, 125 MHz): δ 193.0, 137.0, 131.8, 128.4, 126.9, 126.2, 120.6, 115.8, 101.3, 54.1, 29.7, 20.1 ppm. Anal. calcd for $C_{13}H_{15}NO$: C, 77.58; H, 7.51; N, 6.96. Found: C, 77.18; H, 7.46; N, 6.90.

1-Dodecyl-1*H*-indole-4-carbaldehyde (4d). Yellow liquid (3.86 g), 72.2% yield. 1H NMR ($CDCl_3$, 500 MHz): δ 10.25 (s, 1H), 7.62 (d, J = 7.5 Hz, 2H), 7.35–7.30 (m, 3H), 4.16 (t, J = 7.0 Hz, 2H), 1.85–1.82 (m, 2H), 1.30–1.25 (m, 18H), 0.90 (t, J = 7.5 Hz, 3H) ppm. ^{13}C NMR ($CDCl_3$, 125 MHz): δ 192.9, 136.7, 131.2, 128.4, 126.9, 126.5, 120.6, 115.6, 101.4, 46.5, 31.8, 30.3, 29.52, 29.45, 29.38, 29.25, 29.11, 26.9, 22.6, 14.0 ppm. Anal. calcd for $C_{21}H_{31}NO$: C, 80.46; H, 9.97; N, 4.47. Found: C, 81.05; H, 9.94; N, 4.49.

1-Butyl-1*H*-indole-5-carbaldehyde (5b). Yellow liquid (2.23 g), 65.0% yield. 1H NMR ($CDCl_3$, 500 MHz): δ 10.00 (s, 1H), 8.12 (d, J = 1.0 Hz, 1H), 7.77 (dd, 3J = 8.5 Hz, 4J = 1.0 Hz, 1H), 7.39 (d, J = 8.5 Hz, 1H), 7.17 (d, J = 3.5 Hz, 1H), 6.63 (d, J = 3.0 Hz, 1H), 4.11 (t, J = 7.0 Hz, 2H), 1.83–1.77 (m, 2H), 1.35–1.28 (m, 2H), 0.93 (t, J = 7.5 Hz, 3H) ppm. ^{13}C NMR ($CDCl_3$, 125 MHz): δ 192.2, 139.0, 129.6, 129.1, 128.2, 126.3, 121.4, 109.8, 103.0, 41.2, 32.1, 19.9, 13.4 ppm. Anal. calcd for $C_{13}H_{15}NO$: C, 77.58; H, 7.51; N, 6.96. Found: C, 77.97; H, 7.48; N, 7.01.

1-Isobutyl-1*H*-indole-5-carbaldehyde (5c). Yellow liquid (2.00 g), 58.3% yield. 1H NMR ($CDCl_3$, 500 MHz): δ 10.02 (s, 1H), 8.14 (s, 1H), 7.77 (d, J = 8.5 Hz, 1H), 7.40 (d, J = 8.5 Hz, 1H), 7.16 (d, J = 2.5 Hz, 1H), 6.65 (s, 1H), 3.94 (d, J = 7.5 Hz, 2H), 2.24–2.15 (m, 1H), 0.93 (d, J = 7.0 Hz, 6H) ppm. ^{13}C NMR ($CDCl_3$, 125 MHz): δ 192.4, 139.6, 130.3, 129.2, 128.2, 126.4, 121.6, 110.1, 103.1, 54.3, 29.7, 20.2 ppm. Anal. calcd for $C_{13}H_{15}NO$: C, 77.58; H, 7.51; N, 6.96. Found: C, 77.89; H, 7.55; N, 6.99.

1-Dodecyl-1*H*-indole-5-carbaldehyde (5d). White solid (3.94 g), 73.8% yield. 1H NMR ($CDCl_3$, 500 MHz): δ 10.03 (s, 1H), 8.15 (s, 1H), 7.78 (dd, 3J = 8.5 Hz, 4J = 1.0 Hz, 1H), 7.41 (d, J = 9.0 Hz, 1H), 7.19 (d, J = 3.5 Hz, 1H), 6.65 (d, J = 3.5 Hz, 1H), 4.14 (t, J = 7.0 Hz, 2H), 1.86–1.83 (m, 2H), 1.31–1.24 (m, 18H), 0.88 (t, J = 7.0 Hz, 3H) ppm. ^{13}C NMR ($CDCl_3$, 125 MHz): δ 192.3, 139.3, 129.7, 129.3, 128.3, 126.5, 121.6, 109.9, 103.2, 46.7, 31.9, 30.2, 29.56, 29.49, 29.42, 29.28, 29.15, 26.9, 22.6, 14.1 ppm. Anal. calcd for $C_{21}H_{31}NO$: C, 80.46; H, 9.97; N, 4.47. Found: C, 79.75; H, 10.01; N, 4.50.

Synthesis of IPM derivatives

General procedure: Compound 2 (172.2 mg, 1.0 mmol) and compound 3/4/5 (3.0 mmol) were dissolved in the solution of acetonitrile (10 mL) and piperidine (0.2 mL). The mixture was heated at 80 °C under a nitrogen atmosphere for 24 h. After cooling, the mixture was poured into methanol and suction filtered to give the crude product. The crude product was washed with acetone and methanol three times, respectively,

and then dried to give the pure target compound. The characterization data for various IPM derivatives are listed as follows:

2-(2,6-Bis((E)-2-(1*H*-indol-3-yl)vinyl)-4*H*-pyran-4-ylidene)malononitrile (I3PM). Red solids (362.4 mg), 85.0% yield. 1H NMR ($DMSO-d_6$, 500 MHz): δ 11.87 (s, 2H), 8.25 (d, J = 8.0 Hz, 2H), 8.04–7.98 (m, 4H), 7.51 (d, J = 8.0 Hz, 2H), 7.28–7.19 (m, 6H), 6.92 (s, 2H) ppm. ^{13}C NMR ($DMSO-d_6$, 125 MHz): δ 160.3, 156.3, 137.6, 132.6, 132.5, 124.6, 122.7, 120.9, 120.8, 116.6, 113.3, 112.7, 112.4, 104.3, 52.6 ppm. MS (ESI, m/z): 427.10 ($M^+ + H$). Anal. calcd for $C_{28}H_{18}N_4O$: C, 78.86; H, 4.25; N, 13.14. Found: C, 79.29; H, 4.22; N, 13.22.

2-(2,6-Bis((E)-2-(1-butyl-1*H*-indol-3-yl)vinyl)-4*H*-pyran-4-ylidene)malononitrile (I3PMBn). Red solids (483.2 mg), 89.7% yield. 1H NMR ($CDCl_3$, 500 MHz): δ 7.94 (d, J = 7.5 Hz, 2H), 7.72 (d, J = 15.5 Hz, 2H), 7.51 (s, 2H), 7.23 (d, J = 8.0 Hz, 2H), 7.36–7.30 (m, 4H), 6.67 (d, J = 16.0 Hz, 2H), 6.48 (s, 2H), 4.19 (t, J = 7.0 Hz, 4H), 1.91–1.87 (m, 4H), 1.43–1.37 (m, 4H), 0.99 (t, J = 7.0 Hz, 6H) ppm. ^{13}C NMR ($CDCl_3$, 125 MHz): δ 159.7, 156.2, 137.6, 132.1, 131.2, 125.9, 123.2, 121.6, 120.5, 116.5, 113.5, 112.9, 110.5, 104.8, 55.6, 46.7, 32.0, 20.2, 13.6 ppm. MS (ESI, m/z): 539.25 ($M^+ + H$). Anal. calcd for $C_{36}H_{34}N_4O$: C, 80.27; H, 6.36; N, 10.40. Found: C, 80.81; H, 6.41; N, 10.49.

2-(2,6-Bis((E)-2-(1-isobutyl-1*H*-indol-3-yl)vinyl)-4*H*-pyran-4-ylidene)malononitrile (I3PMBi). Orange-red solids (477.2 mg), 88.6% yield. 1H NMR ($CDCl_3$, 500 MHz): δ 7.96–7.94 (m, 2H), 7.73 (dd, 3J = 16.0 Hz, 4J = 3.5 Hz, 2H), 7.49 (d, J = 3.5 Hz, 2H), 7.42 (dd, 3J = 7.5 Hz, 4J = 3.0 Hz, 2H), 7.36–7.27 (m, 4H), 6.69 (dd, 3J = 16.0 Hz, 4J = 3.5 Hz, 2H), 6.49 (d, J = 3.5 Hz, 2H), 3.39 (dd, 3J = 7.0 Hz, 4J = 3.5 Hz, 4H), 2.29–2.24 (m, 2H), 0.99 (dd, 3J = 6.5 Hz, 4J = 3.5 Hz, 12H) ppm. ^{13}C NMR ($CDCl_3$, 125 MHz): δ 159.7, 156.2, 137.8, 132.8, 131.3, 125.8, 123.2, 121.6, 120.5, 116.5, 113.6, 112.7, 110.7, 104.8, 55.5, 54.5, 29.3, 20.3 ppm. MS (ESI, m/z): 539.25 ($M^+ + H$). Anal. calcd for $C_{36}H_{34}N_4O$: C, 80.27; H, 6.36; N, 10.40. Found: C, 80.92; H, 6.39; N, 10.50.

2-(2,6-Bis((E)-2-(1-dodecyl-1*H*-indol-3-yl)vinyl)-4*H*-pyran-4-ylidene)malononitrile (I3PMD). Orange-red solids (662.3 mg), 86.8% yield. 1H NMR ($CDCl_3$, 500 MHz): δ 7.94 (d, J = 8.0 Hz, 2H), 7.72 (d, J = 15.5 Hz, 2H), 7.50 (s, 2H), 7.42 (d, J = 8.5 Hz, 2H), 7.36–7.29 (m, 4H), 6.67 (d, J = 16.0 Hz, 2H), 6.47 (s, 2H), 4.17 (t, J = 7.0 Hz, 4H), 1.92–1.89 (m, 4H), 1.35–1.25 (m, 36H), 0.87 (t, J = 7.0 Hz, 6H) ppm. ^{13}C NMR ($CDCl_3$, 125 MHz): δ 159.7, 156.2, 137.6, 132.1, 131.2, 125.9, 123.2, 121.6, 120.5, 116.4, 113.5, 112.9, 110.5, 104.7, 55.6, 47.0, 31.9, 30.0, 29.6, 29.5, 29.4, 29.3, 29.2, 27.0, 22.7, 14.1 ppm. MS (ESI, m/z): 763.45 ($M^+ + H$). Anal. calcd for $C_{52}H_{66}N_4O$: C, 81.84; H, 8.72; N, 7.34. Found: C, 82.34; H, 8.67; N, 7.40.

2-(2,6-Bis((E)-2-(1*H*-indol-4-yl)vinyl)-4*H*-pyran-4-ylidene)malononitrile (I4PM). Orange-red solids (300.1 mg), 70.4% yield. 1H NMR ($DMSO-d_6$, 500 MHz): δ 11.50 (s, 2H), 8.06 (d, J = 16.5 Hz, 2H), 7.58–7.46 (m, 8H), 7.21 (d, J = 7.5 Hz, 2H), 7.10 (s, 4H) ppm. ^{13}C NMR ($DMSO-d_6$, 125 MHz): δ 159.3, 156.3, 137.4, 136.5, 127.1, 126.4, 126.0, 121.6, 121.1, 118.6, 115.9, 114.0, 106.9, 100.6, 55.7 ppm. MS (ESI, m/z): 427.10 ($M^+ + H$). Anal. calcd for $C_{28}H_{18}N_4O$: C, 78.86; H, 4.25; N, 13.14. Found: C, 79.25; H, 4.23; N, 13.02.



2-(2,6-Bis((E)-2-(1-butyl-1*H*-indol-4-yl)vinyl)-4*H*-pyran-4-ylidene)malononitrile (I4PMBn). Orange-red solids (430.4 mg), 79.9% yield. ^1H NMR (CDCl_3 , 500 MHz): δ 7.92 (d, J = 16.0 Hz, 2H), 7.42 (d, J = 8.5 Hz, 2H), 7.37 (d, J = 7.0 Hz, 2H), 7.28–7.23 (m, 4H), 6.86 (d, J = 16.0 Hz, 2H), 6.77 (d, J = 2.5 Hz, 2H), 6.55 (s, 2H), 4.16 (t, J = 7.0 Hz, 4H), 1.87–1.81 (m, 4H), 1.40–1.32 (m, 4H), 0.95 (t, J = 7.0 Hz, 6H) ppm. ^{13}C NMR (CDCl_3 , 125 MHz): δ 158.6, 156.0, 136.9, 136.6, 129.6, 127.5, 126.8, 121.6, 119.9, 118.3, 115.6, 111.9, 106.7, 99.3, 58.6, 46.4, 32.4, 20.2, 13.6 ppm. MS (EI, m/z): 538.29 (M^+ , 56%), 495.23 (8%), 394.25 (22%), 226.10 (57%), 186.12 (49%). Anal. calcd for $\text{C}_{36}\text{H}_{34}\text{N}_4\text{O}$: C, 80.27; H, 6.36; N, 10.40. Found: C, 80.83; H, 6.33; N, 10.32.

2-(2,6-Bis((E)-2-(1-isobutyl-1*H*-indol-4-yl)vinyl)-4*H*-pyran-4-ylidene)malononitrile (I4PMBi). Orange-red solids (360.7 mg), 67.0% yield. ^1H NMR (CDCl_3 , 500 MHz): δ 7.99 (d, J = 16.0 Hz, 2H), 7.42 (dd, 3J = 14.0 Hz, 4J = 8.0 Hz, 4H), 7.29–7.27 (m, 4H), 6.93 (d, J = 16.0 Hz, 2H), 6.82 (d, J = 2.5 Hz, 2H), 6.64 (s, 2H), 3.99 (d, J = 7.5 Hz, 4H), 2.28–2.20 (m, 2H), 0.97 (d, J = 6.5 Hz, 12H) ppm. ^{13}C NMR (CDCl_3 , 125 MHz): δ 158.7, 156.0, 137.0, 136.9, 130.2, 127.4, 126.9, 121.6, 120.0, 118.4, 115.6, 112.1, 106.7, 99.2, 58.8, 54.4, 29.7, 20.3 ppm. MS (ESI, m/z): 539.25 (M^+ + H). Anal. calcd for $\text{C}_{36}\text{H}_{34}\text{N}_4\text{O}$: C, 80.27; H, 6.36; N, 10.40. Found: C, 79.86; H, 6.32; N, 10.35.

2-(2,6-Bis((E)-2-(1-dodecyl-1*H*-indol-4-yl)vinyl)-4*H*-pyran-4-ylidene)malononitrile (I4PMD). Orange-red solids (600.3 mg), 78.7% yield. ^1H NMR (CDCl_3 , 500 MHz): δ 7.94 (d, J = 16.0 Hz, 2H), 7.39 (dd, 3J = 19.0 Hz, 4J = 7.0 Hz, 4H), 7.26 (d, J = 3.5 Hz, 2H), 7.24 (d, J = 7.0 Hz, 2H), 6.88 (d, J = 16.0 Hz, 2H), 6.77 (d, J = 2.5 Hz, 2H), 6.59 (s, 2H), 4.14 (t, J = 7.0 Hz, 4H), 1.86–1.83 (m, 4H), 1.31–1.23 (m, 36H), 0.85 (t, J = 6.5 Hz, 4H) ppm. ^{13}C NMR (CDCl_3 , 125 MHz): δ 158.7, 156.0, 137.0, 136.6, 129.6, 127.5, 126.9, 121.6, 120.0, 118.4, 115.6, 111.9, 106.7, 99.3, 58.7, 46.7, 31.9, 30.3, 29.6, 29.5, 29.4, 29.3, 29.2, 27.0, 22.7, 14.1 ppm. MS (ESI, m/z): 763.45 (M^+ + H). Anal. calcd for $\text{C}_{52}\text{H}_{66}\text{N}_4\text{O}$: C, 81.84; H, 8.72; N, 7.34. Found: C, 82.16; H, 8.77; N, 7.38.

2-(2,6-Bis((E)-2-(1*H*-indol-5-yl)vinyl)-4*H*-pyran-4-ylidene)malononitrile (I5PM). Orange-red solids (311.7 mg), 73.1% yield. ^1H NMR ($\text{DMSO}-d_6$, 500 MHz): δ 11.39 (s, 2H), 8.07 (s, 2H), 7.92 (d, J = 16.0 Hz, 2H), 7.63 (d, J = 16.5 Hz, 2H), 7.48 (d, J = 8.5 Hz, 2H), 7.43 (s, 2H), 7.29 (d, J = 16.0 Hz, 2H), 6.81 (s, 2H), 6.54 (s, 2H) ppm. ^{13}C NMR ($\text{DMSO}-d_6$, 125 MHz): δ 159.6, 156.2, 140.0, 137.1, 128.0, 126.7, 126.3, 122.5, 120.6, 116.0, 115.5, 112.1, 105.7, 102.1, 54.7 ppm. MS (ESI, m/z): 427.10 (M^+ + H). Anal. calcd for $\text{C}_{28}\text{H}_{18}\text{N}_4\text{O}$: C, 78.86; H, 4.25; N, 13.14. Found: C, 78.51; H, 4.28; N, 13.06.

2-(2,6-Bis((E)-2-(1-butyl-1*H*-indol-5-yl)vinyl)-4*H*-pyran-4-ylidene)malononitrile (I5PMBn). Orange solids (245.2 mg), 45.5% yield. ^1H NMR (CDCl_3 , 500 MHz): δ 7.84 (s, 2H), 7.67 (d, J = 16.0 Hz, 2H), 7.48 (d, J = 8.5 Hz, 2H), 7.38 (d, J = 8.5 Hz, 2H), 7.15 (s, 2H), 6.69 (d, J = 16.0 Hz, 2H), 6.58–6.56 (m, 4H), 4.14 (t, J = 7.5 Hz, 4H), 1.88–1.82 (m, 4H), 1.40–1.34 (m, 4H), 0.97 (t, J = 7.5 Hz, 6H) ppm. ^{13}C NMR (CDCl_3 , 125 MHz): δ 159.1, 156.2, 139.8, 137.3, 129.2, 129.0, 126.2, 122.6, 120.6, 115.8, 115.3, 110.2, 106.0, 102.2, 57.6, 46.4, 32.4, 20.1, 13.6 ppm. MS (ESI, m/z): 539.25 (M^+ + H). Anal. calcd for $\text{C}_{36}\text{H}_{34}\text{N}_4\text{O}$: C, 80.27; H, 6.36; N, 10.40. Found: C, 79.75; H, 6.31; N, 10.33.

2-(2,6-Bis((E)-2-(1-isobutyl-1*H*-indol-5-yl)vinyl)-4*H*-pyran-4-ylidene)malononitrile (I5PMBi). Brown solids (223.7 mg), 41.5% yield. ^1H NMR (CDCl_3 , 500 MHz): δ 7.85 (s, 2H), 7.68 (d, J = 15.5 Hz, 2H), 7.48 (d, J = 8.5 Hz, 2H), 7.38 (d, J = 8.5 Hz, 2H), 7.13 (s, 2H), 6.70 (d, J = 15.5 Hz, 2H), 6.59 (d, J = 16.0 Hz, 4H), 3.94 (d, J = 7.5 Hz, 4H), 2.25–2.19 (m, 2H), 0.96 (d, J = 6.5 Hz, 12H) ppm. ^{13}C NMR (CDCl_3 , 125 MHz): δ 159.1, 156.2, 139.8, 137.6, 129.8, 129.0, 126.3, 122.6, 120.6, 115.8, 115.4, 110.5, 106.0, 102.1, 57.7, 54.3, 29.7, 20.3 ppm. MS (ESI, m/z): 539.25 (M^+ + H). Anal. calcd for $\text{C}_{36}\text{H}_{34}\text{N}_4\text{O}$: C, 80.27; H, 6.36; N, 10.40. Found: C, 79.69; H, 6.33; N, 10.48.

2-(2,6-Bis((E)-2-(1-dodecyl-1*H*-indol-5-yl)vinyl)-4*H*-pyran-4-ylidene)malononitrile (I5PMD). Brown solids (162.7 mg), 21.3% yield. ^1H NMR (CDCl_3 , 500 MHz): δ 7.85 (s, 2H), 7.70 (d, J = 15.5 Hz, 2H), 7.49 (d, J = 8.5 Hz, 2H), 7.39 (d, J = 8.5 Hz, 2H), 7.15 (s, 2H), 6.72 (d, J = 16.0 Hz, 2H), 6.63 (s, 2H), 6.57 (s, 2H), 4.14 (t, J = 7.0 Hz, 4H), 1.88–1.85 (m, 4H), 1.33–1.26 (m, 36H), 0.86 (t, J = 6.5 Hz, 6H) ppm. ^{13}C NMR (CDCl_3 , 125 MHz): δ 159.1, 156.2, 139.9, 137.3, 129.2, 129.1, 126.3, 122.6, 120.7, 115.8, 115.4, 110.3, 106.0, 102.2, 57.8, 46.7, 31.9, 30.3, 29.6, 29.5, 29.4, 29.3, 29.2, 27.0, 22.7, 14.1 ppm. MS (EI, m/z): 761.24 (M^+ – H, 100), 382.23 (31%), 298.23 (78%), 207.03 (69%), 186.10 (73%), 158.08 (67%). Anal. calcd for $\text{C}_{52}\text{H}_{66}\text{N}_4\text{O}$: C, 81.84; H, 8.72; N, 7.34. Found: C, 81.23; H, 8.69; N, 7.29.

Results and discussion

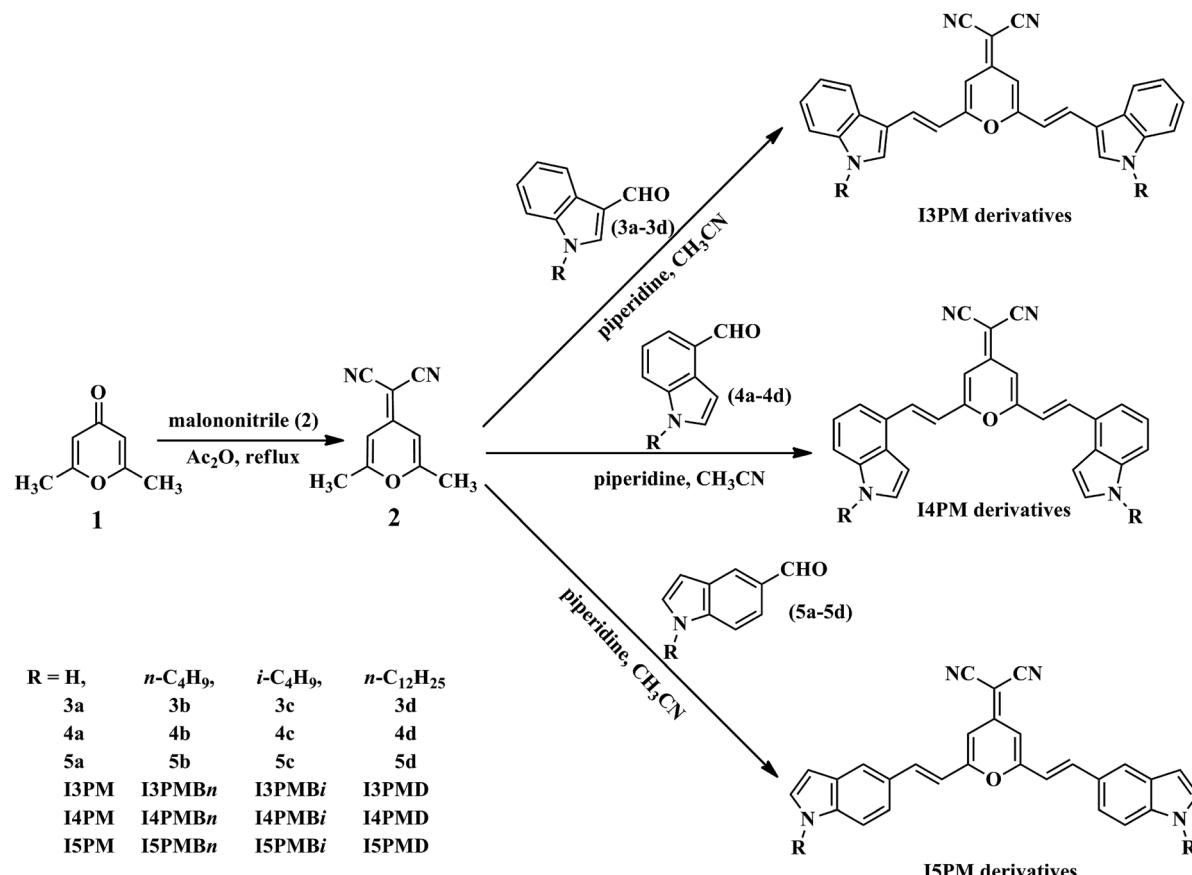
Synthesis and features of IPM derivatives

The chemical structures and synthetic route of IPM derivatives are depicted in Scheme 1. The intermediate 2-(2,6-dimethyl-4*H*-pyran-4-ylidene)malononitrile (2) was synthesized according to the previous literature procedure using 2,6-dimethyl-4-pyrone (1) as the starting material.²⁰ The intermediates 3b–3d/4b–4d/5b–5d were obtained from the alkylation of 3a/4a/5a with the corresponding alkyl bromide in the presence of KOH, respectively, adopting a method similar to the previous literature procedures (Scheme S1, ESI[†]).²¹ And then the IPM derivatives could be facilely prepared from a Knoevenagel reaction of the intermediate 2 with 3a–3d/4a–4d/5a–5d, respectively. The chemical structures of IPM derivatives were confirmed by NMR spectroscopy, mass spectrometry, and elemental analysis. Among these compounds, the structures of I3PM, I3PMBn, I3PMBi, I4PMBn, I4PMBi, and I5PM were further confirmed by single-crystal X-ray diffraction analysis. These target compounds are soluble in common organic solvents, such as chloroform, dichloromethane, toluene, tetrahydrofuran (THF), and ethyl acetate (EA).

Photophysical properties of IPM derivatives in solution and solid state

The absorption and fluorescence spectra of the IPM derivatives in THF solution were measured as shown in Fig. S1 and S2 (ESI[†]), respectively, and the data are presented in Table S1 (ESI[†]). The IPM derivatives in THF solution have similar absorption and emission spectra, and exhibit three or four absorption peaks in the range of 333–487 nm, and one emission





Scheme 1 Chemical structures and synthetic routes of IPM derivatives.

peaks in the range of 529–580 nm. Compared with **I3PM**, **I4PM**, and **I5PM**, the corresponding *N*-alkylated IPM derivatives show a slight redshift in absorption spectra as well as fluorescence spectra because of the introduction of the electron-donating alkyl group. However, the different alkyl groups or isomerized indole units do not show obvious influence on their absorption and fluorescence spectra in THF solution.

Furthermore, the fluorescence spectra of the IPM derivatives in the solid state were investigated. As shown in Fig. S3 (ESI[†]) and Table 1, the original solid samples of **I3PM**, **I4PM**, and **I5PM** emit red or orange fluorescence, and appear emission peaks at 634, 607, and 608 nm, respectively. Compared with **I3PM** and **I4PM**, the corresponding *N*-alkylated IPM derivatives in the solid state exhibit an obvious blue shift in emission spectra, whereas, the emission spectra **I5PM** derivatives are related irregularly to the change of the alkyl chain. It is worth mentioning that the increase of alkyl chain length (from *n*-butyl group to *n*-dodecyl group) leads to the blue shift for the **I3PM** derivatives and the red shift for the **I4PM** derivatives in the emission spectra, showing an interesting correlation with their MFC behaviors which will be mentioned below. Contrary to the situation in the solution, the alkyl groups exhibit an obvious influence on the solid-state fluorescence properties of IPM derivatives.

MFC properties of IPM derivatives

The MFC behaviors of the original samples of IPM derivatives were explored (Fig. 1). When the original solid samples of **I3PM** derivatives were put in a mortar and ground with a pestle, the change of the fluorescence color was not obvious for **I3PM** and **I3PMBn**. Although grinding led to an obvious red shift in the emission spectra of **I3PM** and **I3PMBn** (Fig. 2 and Table 1) and

Table 1 Emission wavelengths of the IPM derivatives in solid state

Compound	$\lambda_{\text{Original}}$ (nm)	λ_{Ground} (nm)	$\Delta\lambda_{\text{MC}}^a$ (nm)	λ_{Fumed} (nm)
I3PM	634	655	21	n.d. ^b
I3PMBn	622	631	9	n.d.
I3PMBi	601	632	31	618
I3PMD	591	604	13	592
I4PM	607	634	27	613
I4PMBn	588	621	33	583
I4PMBi	586	626	40	585
I4PMD	590	599	9	n.d.
I5PM	608	618	10	615
I5PMBn	613	614	1	n.d.
I5PMBi	588, 675 (s) ^c	593, 675 (s)	5	n.d.
I5PMD	603, 657 (s)	603, 663 (s)	6	n.d.

^a $\lambda_{\text{MC}} = \lambda_{\text{Ground}} - \lambda_{\text{Original}}$. ^b n.d. = no detection. ^c s = shoulder peak.



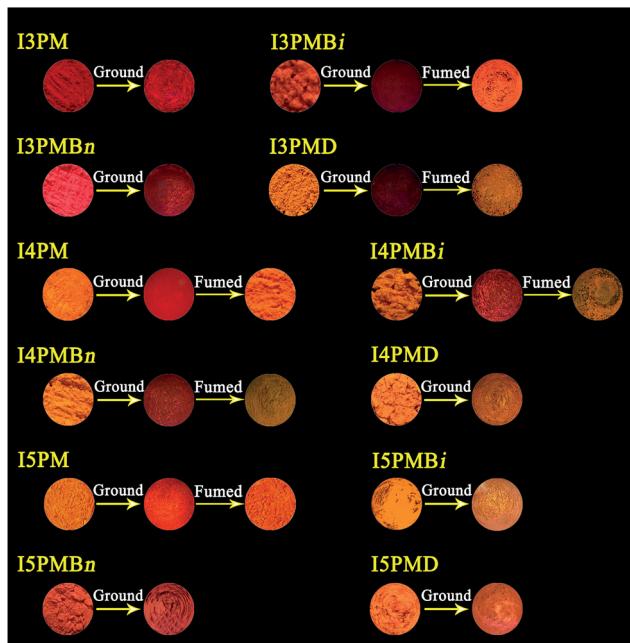


Fig. 1 Photographs of the solid samples of IPM derivatives upon pressing and solvent-fuming (EA vapor) under a 365 nm UV lamp.

the grinding-induced spectral shifts ($\Delta\lambda_{\text{MFC}} = \lambda_{\text{Ground}} - \lambda_{\text{Original}}$) were 21 and 9 nm, respectively, these two compounds did not show obvious MFC phenomena because the fluorescence color change from red to more red could not be well recognized by the

naked eye. However, a color change was clearly observed from orange to red for **I3PMBi** ($\Delta\lambda_{\text{MFC}} = 31$ nm) and **I3PMD** ($\Delta\lambda_{\text{MFC}} = 13$ nm), and the fluorescence colors of their ground samples could be changed into orange by solvent-fuming (exposed to EA vapor), suggesting these two compounds exhibited reversible MFC properties. The results indicated that the isomerization of the alkyl chain and the increase of the alkyl chain length were favorable to the effective mechanofluorochromism of the **I3PM** derivatives.

For **I4PM** derivatives, by grinding the original samples of **I4PM**, **I4PMBn**, and **I4PMBi** with orange fluorescence, the corresponding red-emitting ground samples were obtained and their $\Delta\lambda_{\text{MFC}}$ values were 27, 33, and 31 nm, respectively. Different from **I3PMD**, **I4PMD** containing a dodecyl chain did not show the fluorescence color change in response to press stimuli, indicating that the increase of the alkyl chain length were disadvantageous to the MFC properties of the **I4PM** derivatives. As for **I5PM** derivatives, interestingly, only the original samples of **I5PM** exhibited reversible fluorescence color change upon grinding ($\Delta\lambda_{\text{MFC}} = 10$ nm) and solvent-fuming, the fluorescence spectra and colors of the other homologues had no significant change ($\Delta\lambda_{\text{MFC}} = 5\text{--}6$ nm). Thereby the introduction of the alkyl chain was not advantageous for the MFC properties of the **I5PM** derivatives. As mentioned above, the alkyl chains exhibited different effects on the MFC properties of the IPM derivatives when the substitution position of the indole unit was different.

To understand the different MC behaviors of the IPM derivatives upon external stimuli, XRD measurements were

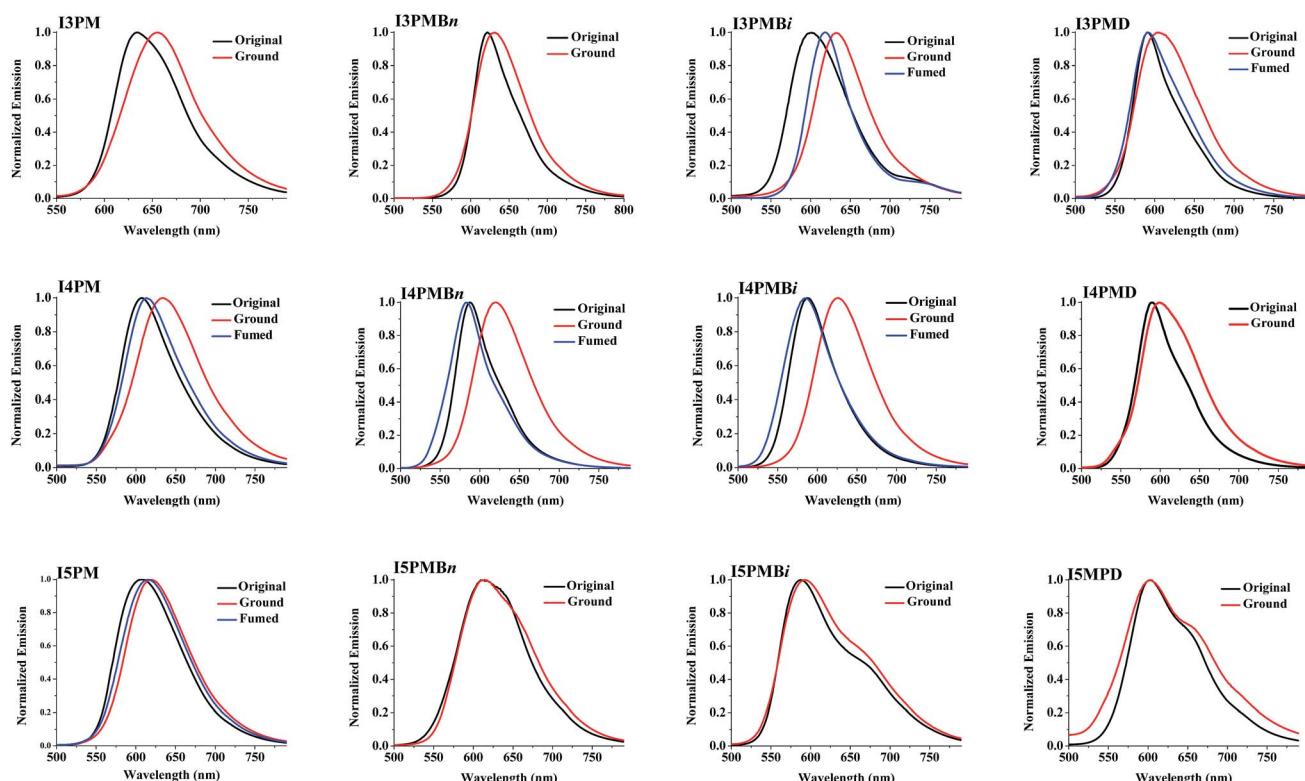


Fig. 2 Normalized fluorescence spectra of the solid samples of IPM derivatives upon external stimuli.



carried out for the samples (Fig. 3, S4, and S5, ESI[†]). As shown in Fig. 3, XRD measurements indicated all the **I3PM** derivatives showed different structures of molecular aggregation before and after grinding. The XRD curves of the original samples of **I3PMBi** and **I3PMD** displayed sharp and intense diffraction peaks, indicating they were crystalline structures. These diffraction peaks disappeared or decreased obviously after grinding, indicating the destruction of crystalline structures and the formation of amorphous structures. Thus, we could conclude that the original samples underwent a morphological change from a crystalline state to an amorphous state upon grinding. The diffraction peaks were recovered when the ground samples were fumed by EA vapor and the recovery led to the blue-shifted fluorescence spectra as well as the recovery of the fluorescence colors. It was noted that the morphology of **I3PM** and **I3PMB** was also significantly changed under the stimulation of pressure, which was consistent with the red-shifted emission wavelengths observed in the fluorescence spectra (Fig. 2), but the fluorescence color change was hardly distinguished. For the **I4PM** and **I5PM** derivatives, the grinding treatment led to the transformation from crystalline state to amorphous state of the MFC-active compounds, such as **I4PM**, **I4PMBn**, **I4PMBi**, and **I5PM**, whereas, the morphology of the MFC-inactive compounds including **I4PMD**, **I5PMBn**, **I5PMBi**, and **I5PMD** still maintained (Fig. S4 and S5, ESI[†]). The results indicated the MFC properties of the IPM derivatives should be mainly attributed to the transformation between amorphous and crystalline states.

Crystal structures of IPM derivatives

To gain deep insight into the underlying reasons for the photophysical behavior of the IPM derivatives in the solid state, single crystals of some of the target compounds were obtained and the crystal structure analysis was performed.²² Single crystals of **I3PM** and **I5PM** were obtained from a THF/1,4-dioxane solution (v/v = 1 : 1), and those of **I3PMBn**, **I3PMBi**, **I4PMBn**, and **I4PMBi** were cultured from a CHCl₃/ethanol solution (v/v = 1 : 1), respectively. The crystal structures of these compounds are shown in Fig. 4 and the selected crystal data are given in Table 2. The unit cells of **I3PM**, **I3PMBn**, **I4PMBi**, and **I5PM** are all triclinic with space group $\bar{P}1$, and those of **I3PMBi** and **I4PMBn** are monoclinic with space group $P12_1/c1$ and $C 2/c$, respectively.

For **I3PM**, the dihedral angles between the 4H-pyran cycle and the two indole units are 3.61° and 9.64°, respectively, indicating its molecule shows a certain degree of distortion. And there are no obvious π – π stacking interactions in the crystal of **I3PM**. One kind of aliphatic C–H···N bonds (2.688 Å) restrict the two adjacent **I3PM** molecules in the same layer (Fig. 5 and Table 3). Because of the unit cell of **I3PM** includes three 1,4-dioxane molecules, these solvent molecules connect the **I3PM** molecules of the upper and lower layers by two kinds of C–H···O bonds (2.645 and 2.687 Å), one kind of N–H···O hydrogen bond (2.899 Å), and C–H··· π bond (2.796 Å). The intermolecular interactions in crystal **I3PM** promote the formation of a stairs-appearance structure. In comparison with **I3PM**, **I5PM** has a more twisted conformation because the

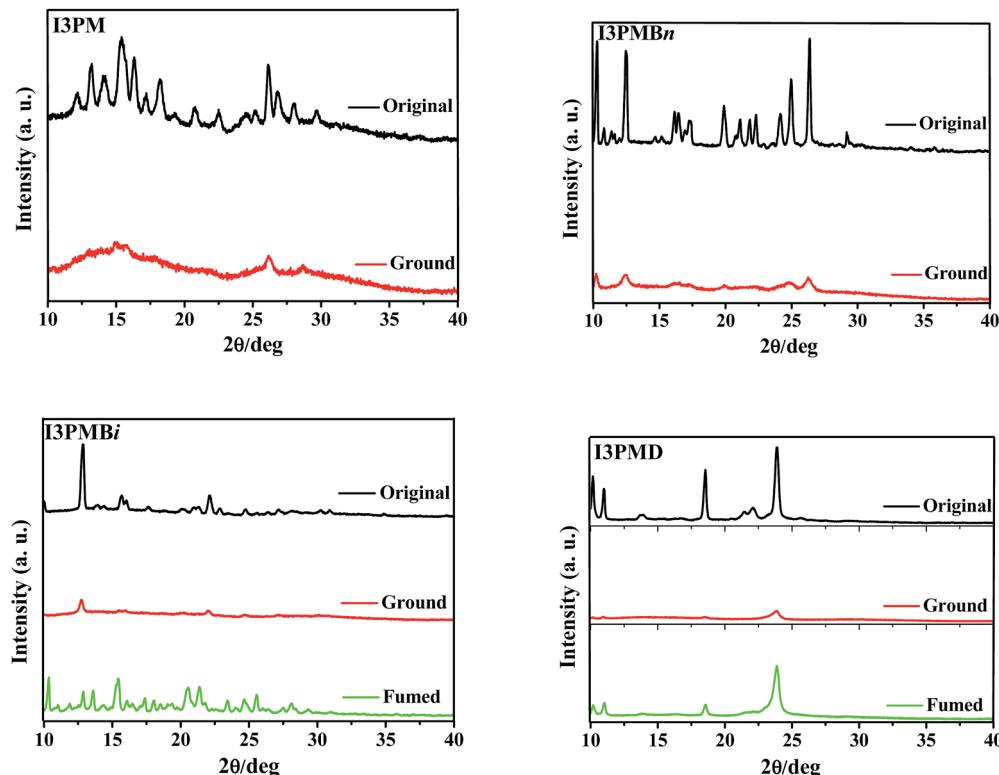


Fig. 3 XRD curves of the solid samples of **I3PM** derivatives upon external stimuli.

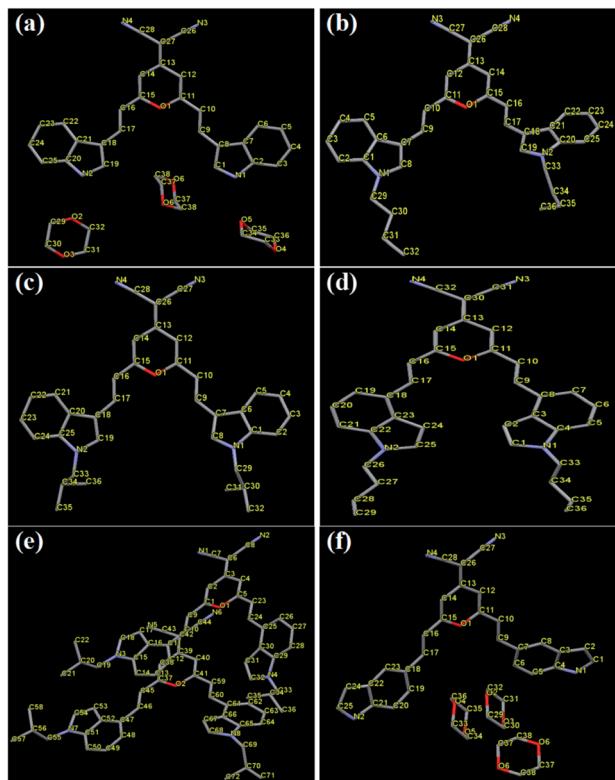


Fig. 4 Single-crystal structures of **I3PM** (a), **I3PMBn** (b), **I3PMBi** (c), **I4PMBn** (d), **I4PMBi** (e), and **I5PM** (f). Hydrogen atoms are omitted for clarity.

dihedral angle between the 4*H*-pyran cycle and one of the indole units reaches 27.82° , which can use as an explanation for the facts that **I5PM** has blue-shifted absorption and emission

Table 2 Crystal data of **I3PM**, **I3PMBn**, **I3PMBi**, **I4PMBn**, **I4PMBi**, and **I5PM**

	I3PM	I3PMBn	I3PMBi	I4PMBn	I4PMBi	I5PM
Empirical formula	$C_{38}H_{38}N_4O_6$	$C_{36}H_{34}N_4O$	$C_{36}H_{34}N_4O$	$C_{36}H_{34}N_4O$	$C_{36}H_{34}N_4O$	$C_{38}H_{38}N_4O_6$
Formula weight	646.72	538.67	538.67	538.67	538.67	646.72
Temperature (K)	298(2)	293(2)	296(2)	296(2)	273(2)	298(2)
Crystal system	Triclinic	Triclinic	Monoclinic	Monoclinic	Triclinic	Triclinic
Space group	$P\bar{1}$	$P\bar{1}$	$P12_1/c1$	$C\ 2/c$	$P\bar{1}$	$P\bar{1}$
<i>Z</i>	2	2	4	8	4	2
<i>D</i> _{calcd} [mg/m ³]	1.231	1.187	0.893	1.159	1.019	1.234
<i>F</i> (000)	684	572	1144	2288	1144	684
θ range [°]	2.20–26.00	1.76–25.00	2.38–66.78	2.31–65.00	3.10–64.99	2.32–24.92
<i>R</i> ₁ [$I > 2\sigma(I)$]	0.0724	0.0832	0.0978	0.0793	0.0917	0.0811
<i>wR</i> ₂ [$I > 2\sigma(I)$]	0.2047	0.1684	0.2622	0.2138	0.2407	0.2091
<i>a</i> [Å]	10.4610(14)	8.3257(12)	19.030(3)	39.856(3)	11.5539(4)	9.885(5)
<i>b</i> [Å]	12.7008(17)	12.1141(17)	21.991(3)	8.1733(4)	17.9477(6)	13.056(6)
<i>c</i> [Å]	14.931(3)	15.672(2)	9.8034(12)	19.7806(10)	17.9799(5)	15.400(8)
α [deg]	101.290(3)	100.531(3)	90	90	105.002(2)	70.424(8)
β [deg]	103.112(3)	90.188(4)	102.551(7)	106.633(2)	96.230(2)	80.392(9)
γ [deg]	109.122(2)	103.883(3)	90	90	99.152(2)	68.547(8)
<i>V</i> [Å ³]	1745.0(5)	1506.8(4)	4004.5(9)	6174.0(6)	3511.4(2)	1740.7(15)
GOF	0.971	1.054	0.950	0.913	0.954	1.037
<i>R</i> _(int)	0.0300	0.0477	0.0837	0.1099	0.0896	0.0654
No. of reflns collected	13 631	8415	22 108	20 863	40 675	14 608
No. of unique reflns	6740	5293	6829	5158	11 555	6358
<i>R</i> ₁ (all data)	0.1338	0.1608	0.1414	0.1746	0.1375	0.1275
<i>wR</i> ₂ (all data)	0.2363	0.2008	0.3153	0.2966	0.2779	0.2355

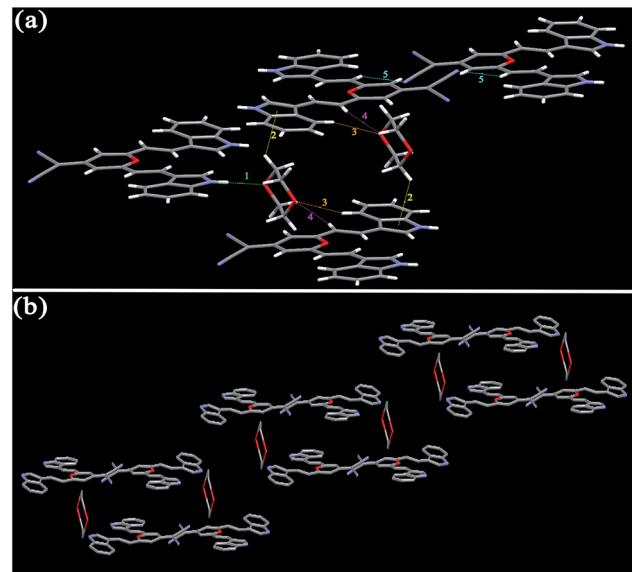


Fig. 5 Intermolecular interactions (a) and molecular packing (b) in **I5PM**.

spectra in THF solution (Table S1†) as well as in solid state (Table 1). There are weak π – π stacking in crystal **I5PM** because the distances between the indole ring and the 4*H*-pyran ring in the upper and lower layers are measured to be 3.642 and 3.698 Å. Furthermore, the crystal structure of **I5PM** is stabilized by one kind of C–H···N bond (2.239 Å) between the two adjacent **I5PM** molecules in the same layer, and one kind of C–H···O bond (2.705 Å) and five kinds of C–H··· π bonds (2.727–3.362 Å) between the 1,4-dioxane solvent molecules and **I5PM** molecules

Table 3 Summarization of the intermolecular interactions in **I3PM**, **I3PMBn**, **I3PMBi**, **I4PMBn**, **I4PMBi**, and **I5PM**

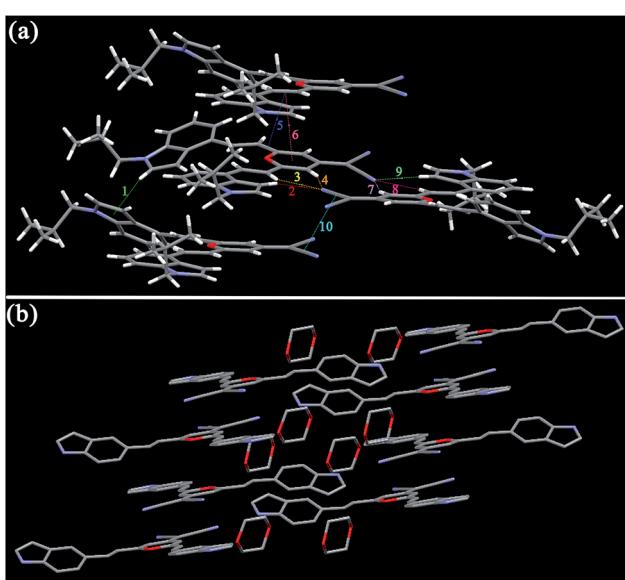
Crystal	No.	II ^a	d/Å	Crystal	No.	II ^a	d/Å
I3PM	1	N1–H1···O5	2.889	I3PMBi	1	H32···N4	2.740
	2	H35B···π(pyrrole)	2.796		2	π(Ph)···π(pyran)	4.530
	3	H22···O4	2.687		3	π(pyrrole)···π(pyran)	4.798
	4	H16···O4	2.645		4	H34A···π(CN)	2.837
	5	H10···N3	2.688		5	π(Ph)···π(pyran)	3.836
I5PM	1	π(pyran)···π(pyrrole)	3.698	I4PMBn	1	H25···N3	2.710
	2	π(pyran)···π(Ph)	3.642		2	H29C···H36B	2.244
	3	H32B···π(C=C)	3.118		3	H29C···H36A	2.351
	4	H16···O2	2.705		4	H28A···H36A	1.943
	5	H30A···π(Ph)	2.844		5	H20···π(Ph)	2.829
	6	H30A···π(pyrrole)	3.062		6	H32···π(pyrrole)	2.943
	7	H1A···N4	2.239		7	H17···N5	2.632
	8	H34B···π(Ph)	3.362		8	H9···N5	2.699
	9	H34B···π(pyrrole)	2.727		9	H2···N5	2.560
I3PMBn	1	H29B···π(pyran)	3.373	I4PMBi	1	π(C=C)···π(C=C)	3.296
	2	π(pyrrole)···π(C=C)	3.654		2	π(C=C)···π(pyran)	3.590
	3	H2···H23	2.340		3	H38···N1	2.634
	4	H30B···π(Ph)	2.827		4	H45···N1	2.747
	5	H29A···π(pyrrole)	3.083		5	H53···N1	2.493
					6	π(CN)···π(CN)	3.374
					7		
					8		
					9		
					10		

^a Intermolecular interaction.

(Fig. 6 and Table 3). As a result, the molecules of **I5PM** pack in a reverse lamellar structure. For the crystals of **I3PM** and **I5PM**, the presence of solvent molecules indicates that there are a large number of defects or holes in their offset lamellar crystal structures.²³ This kind of loose crystal structure is easily destroyed by planarization of the molecular conformation or slip deformation upon grinding, which increases molecular conjugation and promotes a red shift of the emission spectra.^{23a} The destruction of crystalline structures means the transformation from a crystalline state to an amorphous state, which is confirmed by the results of XRD experiments (Fig. 3 and S5†).

In the case of crystals **I3PMBn**, **I3PMBi**, and **I4PMBn**, the two dihedral angles between the two indole units and the 4H-pyran cycle are 7.67 and 14.32° (**I3PMBn**), 5.89 and 19.79° (**I3PMBi**), and 10.01 and 13.04° (**I4PMBn**), respectively. For crystal **I4PMBi**, there were two independent conformational molecules whose corresponding dihedral angles are 5.77 and 23.58°, and 5.23 and 30.22°, respectively. Based on the above mentioned, we can conclude the following results: (i) the introduction of *n*-butyl group and i-butyl group on the nitrogen atom of the indole unit enhances the degree of distortion of the molecules, which can be used to account for the obvious blue shift of the solid-state emission spectra of these four molecules relative to those of the corresponding parent molecules **I3PM** and **I4PM** (Table 1); (ii) compared with *n*-butyl group, i-butyl group is more favorable for the formation of a twisted conformation, and therefore the solid-state emissions of **I3PMBi** and **I3PMBi** show blue shift compared to those of **I3PMBn** and **I3PMBn**, respectively; (iii) **I4PM** derivatives possess a more twisted conformation than the corresponding **I3PM** derivatives, which is consistent with the fact that the former has shorter emission wavelength in the solid state.

Furthermore, we investigated the intermolecular interactions and packing arrangements of these compounds. For crystal **I3PMBn**, the distance between the indole unit and the olefinic double bond in the upper and lower molecules is measured to be 3.654 Å, indicating the existence of weak π–π stacking (Fig. 7 and Table 3). Three kinds of aliphatic C–H···π bonds (2.827, 3.083, and 3.373 Å), and one kind of van der Waals force between aromatic hydrogen atoms in the phenyl units with a distance of 2.340 Å stabilizes the crystal structure. Similar to **I3PM**, **I3PMBn** adopts a reverse lamellar packing pattern. However, there is only one kind of C–H···N bond (2.740 Å) in the crystal of **I3PMBi**, which results in

Fig. 6 Intermolecular interactions (a) and molecular packing (b) in **I5PM**.

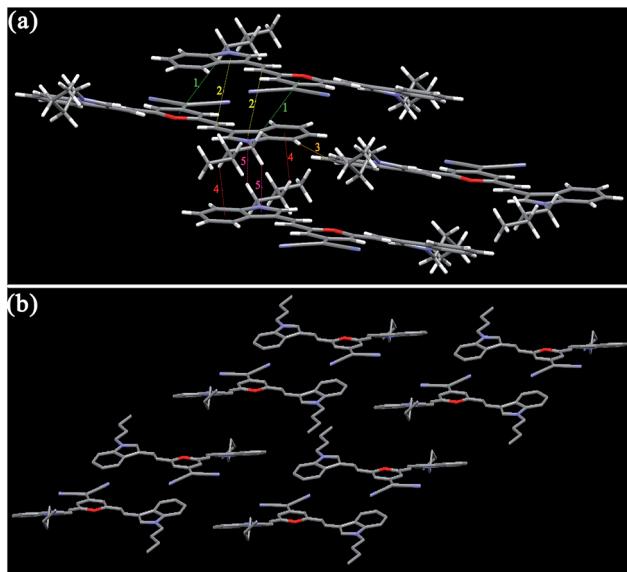


Fig. 7 Intermolecular interactions (a) and molecular packing (b) in **I3PMBn**.

a herringbone arrangement with a herringbone angle of 86.03° (Fig. 8 and Table 3). Herein, the *i*-butyl group obviously decreases the intermolecular interactions between the molecules of **I3PMBi**, which causes that **I3PMBi** has larger $\Delta\lambda_{MFC}$ value and shows more significant solid-state fluorescence change compared with **I3PMBn** containing *n*-butyl group.

As shown in Fig. 9, the molecules of **I4PMBn** in the crystal are stabilized by one kind of π - π stacking interaction (3.836 \AA) between the *4H*-pyran ring and the phenyl unit, one kind of C-H \cdots N bond (2.710 \AA), two kinds of C-H \cdots π bonds (2.829 and 2.837 \AA), and three kinds of van der Waals force between the hydrogen atoms in *n*-butyl groups (1.943 , 2.244 and 2.351 \AA)

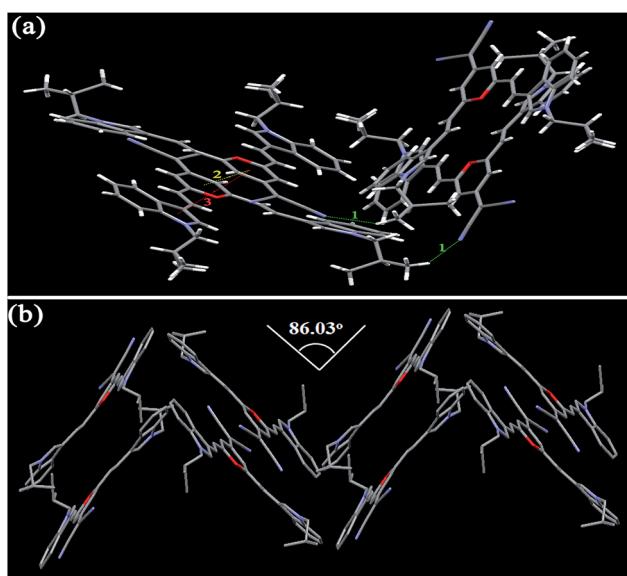


Fig. 8 Intermolecular interactions (a) and molecular packing (b) in **I3PMBi**.

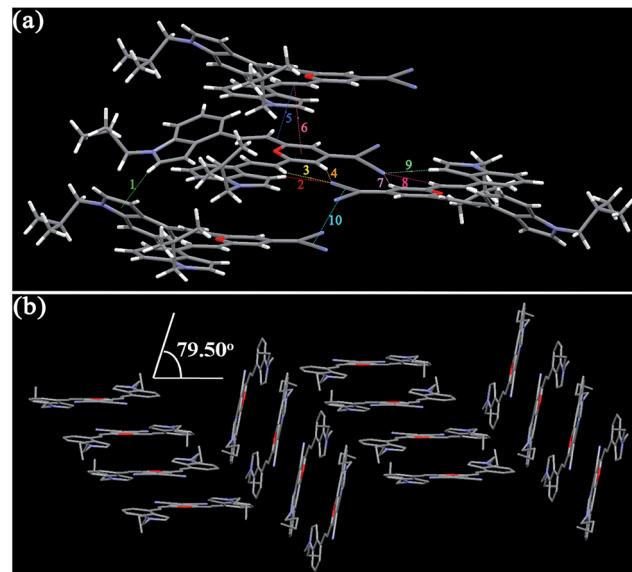


Fig. 9 Intermolecular interactions (a) and molecular packing (b) in **I4PMBi**.

(Table 3). The molecules of **I4PMBn** pack in a herringbone sharp including a lamellar arrangement and the herringbone angle is 79.50° , which is similar to those of **I3PMBi**. For crystal **I4PMBi**, six kinds of C-H \cdots N bonds (2.493 – 2.747 \AA) restrict the molecules in the same layer, three kind of π - π stacking interaction (3.296 , 3.590 and 3.374 \AA) and two kinds of C-H \cdots π bonds (2.874 and 2.943 \AA) stabilize the molecules in the upper and lower layers (Fig. 10 and Table 3). These intermolecular interactions cause the molecules of **I4PMBi** to adopt a face-to-face lamellar stacking mode. Herein, compared with the herringbone packing containing a lamellar arrangement of **I4PMBn**, purely lamellar arrangement of **I4PMBi** makes it easier to break down by external pressure, which leads to its larger $\Delta\lambda_{MFC}$ value.

For the introduction of different alkyl chains, two factors, namely the isomerization of the alkyl chain and the alkyl chain length, are considered. In view of the **I3PM** and **I4PM** derivatives, according to the above crystal structure analysis, *i*-butyl group leads to a weaker intermolecular interaction and/or a looser stacking mode than *n*-butyl group due to a more twisted molecular conformation, which can be used to account for the fact that the isomerization of the alkyl chain is benefit for the remarkable MFC properties. But for the length of alkyl chain, its influence is obviously controlled by the substitution position of the indole unit according to the experimental results. The increase of the alkyl chain length is favorable to the MFC properties of the **I3PM** derivatives, whereas is unfavorable to those of the **I4PM** derivatives. This unusual phenomenon is likely to be related to the change of the molecular conformation before and after grinding. Because there is no strong π - π stacking as well as other intermolecular interactions in the original crystalline samples of the **I3PM** and **I4PM** derivatives, their emissions should mainly depend on the molecular conformations. As shown in Table 1, the original crystalline samples of **I3PMD** and **I4PMD** show a blue shift and red shift



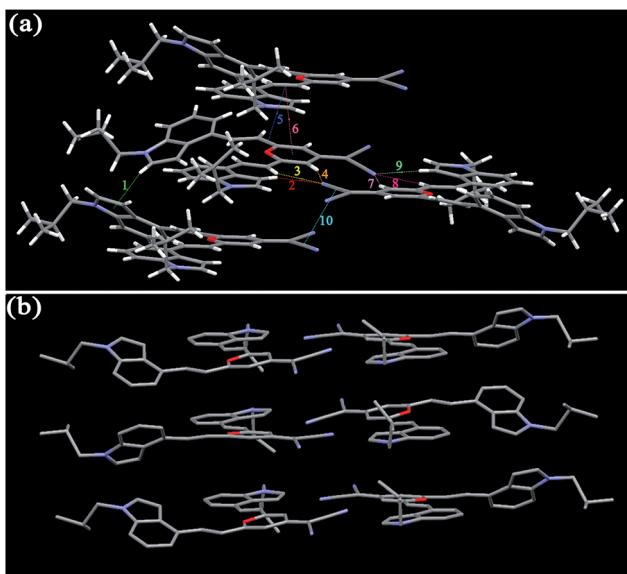


Fig. 10 Intermolecular interactions (a) and molecular packing (b) in I4PMBi.

compared with those of **I3PMBn** and **I4PMBn**, respectively, indicating **I3PMD** and **I4PMBn** might have a more twisted conformation and a poorer π -conjugation. After grinding, a better π -conjugation for the ground samples of **I3PMD** and **I4PMBn** is anticipated, and thus **I3PMD** and **I4PMBn** have a larger $\Delta\lambda_{MFC}$ value. Based on the results in the case, a regular pattern can be concluded as follows: if the emission wavelengths of the original crystalline samples of a series of fluorescent compounds gradually blue-shifted with the increase of alkyl chain length, the homologue with longer alkyl chain will exhibit higher contrast MFC behavior after grinding;^{7-9,24} otherwise shorter alkyl chain is more beneficial to the MFC property,¹⁰⁻¹² which is in good agreement with the previous reports.

Conclusions

A series of new fluorescent molecules based on dicyanomethylene-4H-pyran and indole unit *via* alkene module were synthesized in order to investigate the combined effects of the introduction of different alkyl chains and the isomerization of indole unit on their MFC properties. The results show that the isomerization of indole unit plays a more important role in regulating their MFC activities than the choice of the alkyl chains. For the **I3PM** derivatives, only **I3PMBi** and **I3PMD** show obvious MFC phenomenon, and the increase of the alkyl chain length and the isomerization of the alkyl chain are favorable to their effective mechanofluorochromism. But interestingly, longer alkyl chain is not advantageous for the MFC activities of the **I4PM** derivatives. The opposite alkyl length-dependent mechanofluorochromism is closely related to the degree of distortion of the molecular conformation of the homologues, that is, the greater the degree of distortion of molecular conformation, the larger grinding-induced spectral shifts. More

specially, the introduction of the alkyl chains is completely detrimental to those of the **I5PM** derivatives. Based on the results of crystal structure analysis, the weak intermolecular interactions and loose packing modes in MFC-active molecules result in that the crystalline structures of the molecules are easily destroyed by pressure stimulation and thus the morphological transformation of the molecular aggregates are realized. Although the transformation from a crystalline state to an amorphous state cannot fully determine whether a molecule has MFC properties, for example, **I3PM** and **I3PMBn** do not exhibit obvious solid-state fluorescence color change upon grinding, there is no doubt the transformation is crucial for the formation of MFC properties. This work provides a good design idea for designing and obtaining new MFC-active materials based on the traditional fluorescent molecular frameworks.

Conflicts of interest

There are no conflicts to declare.

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