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The influence of charge compensation defects on the spectroscopic properties of europium doped $\text{Ca}_9\text{Y}(\text{PO}_4)_7$

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In this study, a series of $\text{Ca}_9\text{Y}(\text{PO}_4)_7$ compounds doped with 5% of Eu ions and with different [Y]/[Ca] ratios was synthesized using the Pechini method. Due to the presence of two different cations in the studied matrix (Ca^{2+} and Y^{3+}) that are available for europium substitution, the synthesis parameters were selected to obtain the incorporation of the europium ions (5% mol) into the calcium sites only. The phase composition as well as elemental analysis and the spectroscopic measurements, carried out before and/or after the reduction of Eu^{3+} were performed to characterize the obtained phosphors. The XRD patterns show that in all cases the obtained materials consist of pure phase of $\text{Ca}_9\text{Y}(\text{PO}_4)_7$. PLE and PL spectra measured before the reduction indicate that for the materials with the excess of yttrium, an additional Eu^{3+} site is observed which was not observed for materials with a deficiency of yttrium. For samples obtained after the reduction, the luminescence of Eu^{2+} depends on the changes of the Y/Ca ratio. The observed differences in Eu^{2+} luminescence are discussed taking into account the creation of different compensation defects depending on change of the Y/Ca ratio.

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Introduction

Recently, white light emitting diodes (w-LED) have become a very popular source of light. They are characterized by high efficiency, long lifetime and environmental friendliness. There are several methods that are used to obtain white LEDs. One of the methods consists of a combination of a near-UV LED with red, green and blue phosphors. This method is very appealing because of the great colour rendering index of obtained materials and lower production cost. Thus, the investigations on tricolour phosphors which are suitable for excitation at the near UV-region receive great attention because of their important applications in lighting. Selection of a host that can be activated by different ions emitting in different regions of the spectrum is a very important part of LED design. The widely studied hosts of phosphors are based on borates, phosphates, aluminates, silicates, vanadates, *etc.*^{1–6} Among them, phosphates have been paid intense attention due to their excellent properties.⁷ One of the promising materials in phosphates family is $\text{Ca}_9\text{Y}(\text{PO}_4)_7$, which has two different cations (Ca^{2+} and Y^{3+}) that are available for lanthanides substitution.^{8–11} This material doped with Eu^{2+} and Eu^{3+} ions simultaneously consists of two components that

are essential to achieve the emission of white light. $\text{Ca}_9\text{Y}(\text{PO}_4)_7$ belongs to the whitlockite family and can be described by the general formula: $\text{Ca}_9\text{R}(\text{PO}_4)_7$, where R = Ln, Pr, Eu, Tb, Dy, Ho, Er, Y, Bi, Al, Lu.^{12–18}

The structural model of $\text{Ca}_9\text{Y}(\text{PO}_4)_7$ was proposed by Golubev and Lazoryak¹⁹ and the representation of the unit cell and coordination of the cation sites in $\text{Ca}_9\text{Y}(\text{PO}_4)_7$ are shown in Fig. 1 and 2, respectively.^{13,14,17,20}

Because the crystallographic structure of $\text{Ca}_9\text{Y}(\text{PO}_4)_7$ is isostructural to the $\beta\text{-Ca}_3(\text{PO}_4)_2$, the cation sites in $\text{Ca}_9\text{Y}(\text{PO}_4)_7$ are usually defined in basis of structure of $\text{Ca}_3(\text{PO}_4)_2$.^{13,15} The crystal lattice of $\beta\text{-Ca}_3(\text{PO}_4)_2$ has five different cation sites: seven coordinated Ca(1) positions, eight coordinated Ca(2) and Ca(3) positions, three coordinated Ca(4) positions, which are statistically half occupied by the Ca^{2+} ions and half empty (cation vacancies), and the six coordinated Ca(5) positions, distorted

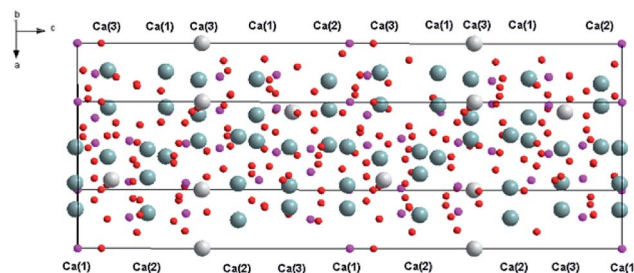


Fig. 1 Representation of the $\text{Ca}_9\text{Y}(\text{PO}_4)_7$ supercell.

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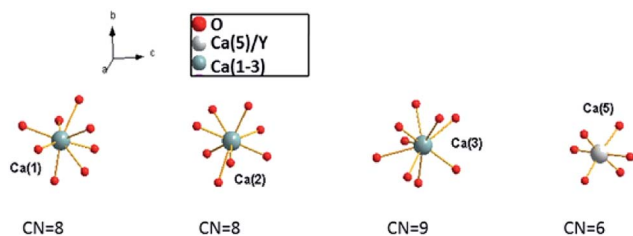


Fig. 2 Coordination of Ca(1), Ca(2), Ca(3) and Ca(5)/Y sites.

octahedral sites, fully occupied by Ca^{2+} ions.^{21,22} Similar situation can be found in $\text{Ca}_9\text{Y}(\text{PO}_4)_7$ with exception that Ca(3) site is nine coordinated, Ca(5) sites are occupied by the Y^{3+} ions and all Ca(4) sites are empty (due to the charge imbalance of Ca^{2+} – Y^{3+} substitution in Ca(5)). The substitution of Ca^{2+} by the Y^{3+} ions in Ca(5) sites does not change the space group of the crystallographic structure and $\text{Ca}_9\text{Y}(\text{PO}_4)_7$ crystallizes in the trigonal $R3c$ (161) space group as in the case of $\text{Ca}_3(\text{PO}_4)_2$. Independently of the coordination number, the Ca(1), Ca(2), Ca(3) sites are characterized by the C_1 point symmetry, whereas the point symmetry of Ca(5) site is C_3 . The values of $\text{Ca}_9\text{Y}(\text{PO}_4)_7$ lattice parameters are as follows: $a = 10.4442 \text{ \AA}$, $c = 37.324 \text{ \AA}$, $V = 3525.89 \text{ \AA}^3$, $Z = 6$.¹²

This paper describes the synthesis, structural properties and the effect of substitution of calcium by europium ions. The main objectives are focused on the investigation of the relations between the material content (the Y/Ca ratio), compensation of the Eu^{3+} ions in the Ca^{2+} sites and the efficiency of the reduction process of the Eu^{3+} . To our best knowledge, although the $\text{Ca}_9\text{Y}(\text{PO}_4)_7:\text{Eu}^{3+}$ material is known, there are no published papers regarding the spectral properties of $\text{Ca}_9\text{Y}(\text{PO}_4)_7:\text{Eu}^{3+}$, Eu^{2+} .

Experimental

Polycrystalline samples of europium doped $\text{Ca}_9\text{Y}(\text{PO}_4)_7$ with different Y/Ca concentration ratio were synthesized by Pechini method, which is described in details elsewhere.^{13,23–26} The 5 mol% of europium was substituted into the calcium sites in all cases. To obtain the desired phosphate, the following reagents were used: $\text{Ca}(\text{NO}_3)_2 \cdot 4\text{H}_2\text{O}$ ($\geq 99\%$ Sigma-Aldrich), $\text{Y}(\text{NO}_3)_3 \cdot 4\text{H}_2\text{O}$ (99.999% Sigma-Aldrich), $(\text{NH}_4)_2\text{HPO}_4$ ($\geq 99.99\%$ Sigma-Aldrich), and $\text{Eu}(\text{NO}_3)_3 \cdot 5\text{H}_2\text{O}$ (99.9% Sigma-Aldrich), citric acid (Avantor Performance Materials Poland S.A.) and ethylene glycol (Chempur).

At first, the citric acid was dissolved in distilled water and the ethylene glycol in the appropriate amount was added to the acid solution (the fixed molar ratio of citric acid and ethylene glycol to total chelate metal cations was 5 : 1).²³ Then the calcium, yttrium and europium nitrates dissolved in distilled water were added to the obtained acid–glycol solution. During the addition of the reagents, the mixture was subsequently stirred. Finally the aqueous solution of $(\text{NH}_4)_2\text{HPO}_4$ was added to the mixture of all substrates. The resulting solution was heated under reflux for about 1 hour and then the excess of solvent was evaporated at $100 \text{ }^\circ\text{C}$. Next, the prepared polyester was dried in the temp. of $80 \text{ }^\circ\text{C}$ for 12 h. After drying, the obtained product was calcined

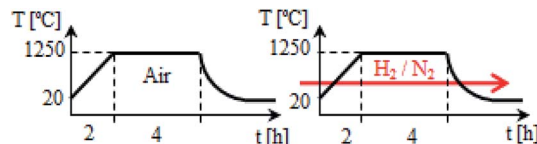


Fig. 3 Scheme of the two-step synthesis.

at $1250 \text{ }^\circ\text{C}$ for 4 h under inert gas atmosphere in the muffle furnace (LHT 04-17, NABERTHERM). Then, the product was cooled, grounded and again calcined at $1250 \text{ }^\circ\text{C}$ for 4 h in a reducing atmosphere, using a mixture of hydrogen (5 vol%) and nitrogen (95 vol%) in an electrical tubular furnace (PRS 75WM, CZYLOK). The scheme of calcination procedure is shown in Fig. 3. In this way materials: $\text{Ca}_{9.45}\text{Eu}_{0.05}\text{Y}_{0.5}(\text{PO}_4)_7$ (denoted as Y0.5), $\text{Ca}_{9.2}\text{Eu}_{0.05}\text{Y}_{0.75}(\text{PO}_4)_7$ (Y0.75), $\text{Ca}_{8.95}\text{Eu}_{0.05}\text{Y}_1(\text{PO}_4)_7$ (Y1), $\text{Ca}_{8.45}\text{Eu}_{0.05}\text{Y}_{1.5}(\text{PO}_4)_7$ (Y1.5) and $\text{Ca}_{7.95}\text{Eu}_{0.05}\text{Y}_2(\text{PO}_4)_7$ (Y2) were obtained.

The phase purity of all samples was precisely confirmed using powder X-ray diffraction (XRD) measurements (BRUKER D2PHASER) at the rate of 0.02° and counting time 0.4 s per step in the 2θ range from 20° to 70° , with Cu $K\alpha$ ($K\alpha$: 1.54 \AA) operating at 30 kV/10 mA. Luminescence excitation spectra were measured using the Horiba spectrofluorometer (FluoroMax-4P TCSPC) equipped with 150 W xenon lamp and R928 Hamamatsu PMT. Photoluminescence spectra were acquired using Shamrock SR750 D1 grating spectrometer followed by iDus 420 CCD detector (Andor Technology) and the CW He–Cd laser was used as an excitation source. All spectra were carefully corrected taking into account the instrumental response function. Time resolved emission spectroscopy of the samples was performed on an experimental setup consisting of tunable parametric light source PG401/SH pumped by 30 ps pulse YAG:Nd PL laser 2143 A/SS (Ekspla) and 2501S grating spectrometer (Brucker Optics) coupled to C4334-01 streak camera (Hamamatsu) which acquires streak images with the resolution of 640×480 points (spectral \times temporal). The time ranges available for a measurement span from 10 ms down to 1 ns.²⁷

Elemental analysis was carried out by a Hitachi S-3400N scanning electron microscope (SEM) operated at 25 kV, equipped with an energy-dispersive X-ray Thermo Noran spectrometer (EDS). To determine the average composition of the specimen, the powders were applied on a graphite conductive tape suitable for high vacuum to obtain a flat layer. The analysis was performed in $250 \times 250 \text{ }\mu\text{m}$ area three times in various parts of the samples. Presented results were calculated as an average value from the measurements.

Results and discussion

The XRD patterns of europium doped samples obtained before and after the reduction process are shown in Fig. 4. As it can be seen in Fig. 4, regardless of the difference in Y/Ca concentration ratio, the obtained materials are characterized by the pure phase of $\text{Ca}_9\text{Y}(\text{PO}_4)_7$ (PDF card no. 00-046-0402). The diffraction patterns do not possess signals from additional, impurity



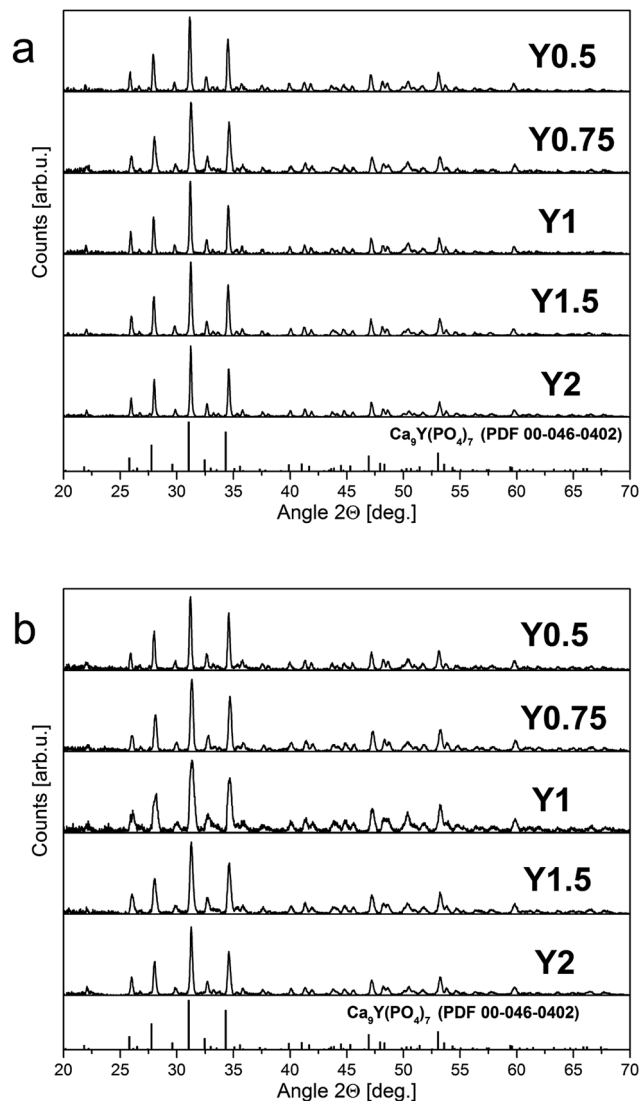


Fig. 4 XRD patterns of samples obtained before (a) and after (b) reduction.

phases. Moreover, the introduction of europium ions or the changes of Y/Ca ratio do not cause any significant alteration to the host structure. After the reduction process the phase composition or the structure of the materials is not changed.

Fig. 5 shows the XRD pattern of Y1 sample with the two PDF standards: PDF 00-046-0402 for $\text{Ca}_9\text{Y}(\text{PO}_4)_7$ and PDF 04-014-2292 for $\text{Ca}_3(\text{PO}_4)_2$. As mentioned previously, the crystallographic structure of $\text{Ca}_9\text{Y}(\text{PO}_4)_7$ is isostructural to the $\text{Ca}_3(\text{PO}_4)_2$ and the XRD signals of both phases are identical. Thus, in this case, the XRD method is not sufficient to distinguish these two members of the whitlockite family. To confirm the incorporation of yttrium into the material lattice, the elemental analysis was performed. Table 1 presents the EDX results for the series of materials obtained after the reduction process. As it can be seen, the materials contain the yttrium ions and the concentration of Y^{3+} increases in order from Y0.5 to Y2. However, in all cases the real concentration is approximately two times lower than assumed. The presence of yttrium ions and the absence of

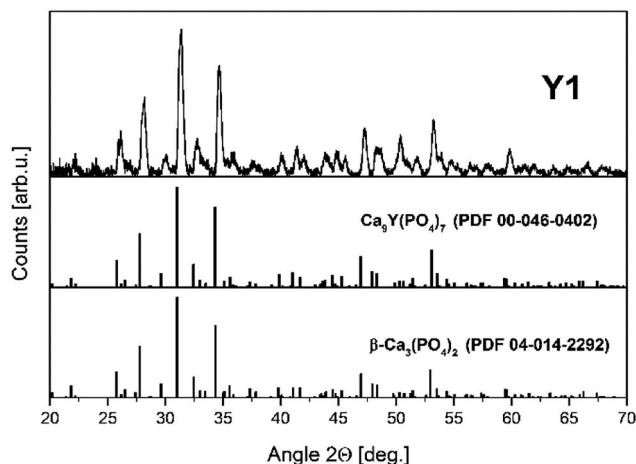


Fig. 5 XRD pattern of Y1 sample with $\text{Ca}_9\text{Y}(\text{PO}_4)_7$ and $\text{Ca}_3(\text{PO}_4)_2$ PDF standards.

Table 1 Nominal and achieved concentration of Y^{3+}

Sample	Nominal atomic%		Experimental atomic%	
	Ca	Y	Ca	Y
Y0.5	95.0	5.0	97.3	2.7
Y0.75	92.5	7.5	94.8	5.2
Y1	89.9	10.1	94.3	5.7
Y1.5	84.9	15.1	86.7	13.4
Y2	79.9	20.1	82.1	17.9

any additional crystal phases of yttrium in the obtained materials confirm that the yttrium ions are incorporated into the crystal lattice.

Optical properties

The normalized, partial photoluminescence excitation (PLE) and photoluminescence (PL) spectra obtained at room temperature for the $\text{Ca}_9\text{Y}(\text{PO}_4)_7:\text{Eu}^{3+}$ system with different Y/Ca ratios (Y0.5; Y0.75; Y1; Y1.5; Y2) are presented in Fig. 6. PLE spectra are shown in narrow spectral region between 390 and 405 nm to reveal the differences between the individual spectra. PL spectra of materials obtained after the first step of the synthesis under excitation at 393 nm and 394 nm consist of narrow lines, which correspond to the $^5\text{D}_0 \rightarrow ^7\text{F}_J$ ($J = 1, 2, 3, 4$) transitions in Eu^{3+} ion. In the case of Y0.5 and Y0.75 materials, the emission spectra do not depend on the excitation wavelength (Fig. 6a and b). For the stoichiometric material (Y1) and materials with the excess of yttrium (Y1.5 and Y2), it is observed that PL spectra depend on the excitation wavelength. Under 393 nm excitation, it is observed that the emission spectra are identical as the ones obtained for Y0.5 and Y0.75 (Fig. 6c), whereas under 394 nm excitation, the emission spectral shapes are clearly different from the ones obtained with 393 nm excitation. The PL spectra under excitation at 394 nm are characterized by higher intensity of magnetic dipole transition $^5\text{D}_0 \rightarrow$



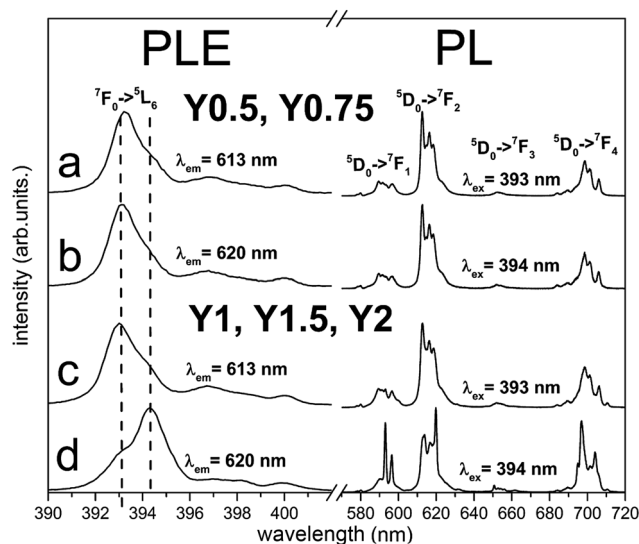


Fig. 6 PL and partial PLE spectra of $\text{Ca}_9\text{Y}(\text{PO}_4)_7:\text{Eu}^{3+}$ measured at RT before reduction. PLE spectra are recorded monitoring at 613 nm and 620 nm wavelength for all samples. The emission spectra were collected under excitation at 393 nm and 394 nm for all samples.

${}^7\text{F}_1$ and electric dipole transition ${}^5\text{D}_0 \rightarrow {}^7\text{F}_4$ than the respective ones obtained under 393 nm excitation (Fig. 6d). It is worth noting that emission related to the Eu^{2+} was not observed in the investigated samples.

The PLE spectra for all samples were recorded monitoring luminescence at 613 nm (Fig. 6a and c) and 620 nm (Fig. 6b and d) and the excitation line shown on Fig. 6 corresponds to ${}^7\text{F}_0 \rightarrow {}^5\text{L}_6$ transition. The most intensive lines seen in PLE spectra have peaks at 393 nm and 394 nm. It can be noticed that the relative intensities of these lines depend on the monitored emission wavelength and Y/Ca concentration. For Y0.5 and Y0.75, the PLE spectra do not depend on the observation wavelength and the ${}^7\text{F}_0 \rightarrow {}^5\text{L}_6$ line is seen at 393 nm. In the case of Y1, Y1.5 and Y2 different excitation spectra were obtained by monitoring 613 and 620 nm luminescence. With observation at 613 nm, the excitation line attributed to the ${}^7\text{F}_0 \rightarrow {}^5\text{L}_6$ transition is seen at 393 nm, while by monitoring at 620 nm, the line is observed at 394 nm. These spectroscopic features provide evidence that in the case of Y1, Y1.5 and Y2 samples, there is an additional center of Eu^{3+} which is not observed for samples with deficiency of yttrium (Y0.5 and Y0.75). The ratio of emission intensities attributed to the ${}^5\text{D}_0 \rightarrow {}^7\text{F}_2$ and the ${}^5\text{D}_0 \rightarrow {}^7\text{F}_1$ transitions, denoted as asymmetric ratio, is used as a measure for the asymmetry of Eu^{3+} centers in phosphors. For the Eu^{3+} center excited with 393 nm, the asymmetric ratio is equal to 6, while for Eu^{3+} center excited with 394 nm this ratio diminishes to about 1.5. This leads to the conclusion that Eu^{3+} ions excited with 394 nm occupy the sites in $\text{Ca}_9\text{Y}(\text{PO}_4)_7$ with a higher inversion symmetry than the ones excited with 393 nm. More distorted sites are attributed to Eu^{3+} replacing divalent Ca^{2+} sites with a triclinic C_1 point group. The sites with higher overall symmetry were attributed to Eu^{3+} replacing Y^{3+} sites with a trigonal C_3 point group.

According to the Tanner work²⁸ where the theoretically allowed transition lines of ${}^5\text{D}_0 \rightarrow {}^7\text{F}_j$ of Eu^{3+} ions at certain

point symmetry groups are present, in our case the 3 peaks of ${}^5\text{D}_0 \rightarrow {}^7\text{F}_1$, 5 peaks of ${}^5\text{D}_0 \rightarrow {}^7\text{F}_2$ and 9 peaks of ${}^5\text{D}_0 \rightarrow {}^7\text{F}_4$ transitions should be observed in the spectra of Eu^{3+} occupying the sites with C_1 symmetry, while in the sites of C_3 symmetry only 2 peaks of ${}^5\text{D}_0 \rightarrow {}^7\text{F}_1$, 3 peaks of ${}^5\text{D}_0 \rightarrow {}^7\text{F}_2$ and 6 peaks of ${}^5\text{D}_0 \rightarrow {}^7\text{F}_4$ transitions should be present. As it can be seen in Fig. 6, the aforementioned splitting is observed only in the emission spectra of Eu^{3+} excited at 394 nm for the Y1, Y1.5, Y2 materials. In this case, the 394 nm excites mainly Eu^{3+} ions that occupy sites with C_3 symmetry. In the case of Eu^{3+} incorporated into the sites with C_1 symmetry, the situation is much more complicated due to the presence of three different C_1 sites that are available for Eu^{3+} substitution. These sites are present in all materials and the Eu^{3+} luminescence cannot be easily distinguished among the sites.

Therefore considering the spectra presented in Fig. 6, one can conclude that when there is a deficiency of yttrium (the case of Y0.5 and Y0.75), Eu^{3+} ions replace almost only divalent Ca^{2+} sites, while for the stoichiometric material (Y1) and materials with the excess amount of yttrium (Y1.5 and Y2), Eu^{3+} ions occupy, besides the Ca^{2+} sites, also Y^{3+} sites.

Another approach to study the number and distortions symmetry of Eu^{3+} centers is to analyze Eu^{3+} time-resolved luminescence spectra. This technique was used for the samples obtained before the reduction and with both: deficiency (Y0.5) and excess (Y2) of yttrium. The spectra were measured under 290 nm excitation wavelength. This excitation wavelength corresponds to the charge transfer transition from the oxygen anion to the empty state of $4f^7$ of Eu^{3+} and excites Eu^{3+} incorporated in the both sites simultaneously.¹⁴ This excitation was used because there was no possibility to use 393 nm and 394 nm excitation wavelength which can be used to excite Eu^{3+} centers independently.

The time-resolved measurement technique is very useful, because one can observe changes of the emission spectrum shape in time.²² The time-resolved luminescence spectra obtained under 290 nm excitation and for acquisition times: 0–3 ms (solid curves) and 7–10 ms (dotted curves) for Y0.5 and Y2 samples are presented in Fig. 7a and b, respectively. In the case of Y0.5, the shape of the emission profile and the intensity ratio of the individual lines do not change significantly in time. Decay curves of Y0.5 monitoring luminescence at different wavelengths with the spectral width of 2 nm ($\lambda_{\text{obs}} = 593$ nm, 612 nm and 696 nm) are plotted in Fig. 8a. All curves are single exponential and, to obtain decay times, single exponential functions were fitted. The decay profiles are plotted along with the calculated decay times. Taking into account the fact that the emission profile does not change in time significantly and that the decay times with a measurement error of monitoring luminescence at different wavelengths are equal, we can conclude and confirm the existence of a single Eu^{3+} site in Y0.5 sample. A different situation is seen for Y2 sample. In the case of Y2, the intensity ratio of the individual lines changes in time, for 0–3 ms time acquisition the collected emission spectra are similar to the ones collected for Y0.5, while in 7–10 ms time acquisition the higher intensity magnetic dipole transition ${}^5\text{D}_0 \rightarrow {}^7\text{F}_1$ and electric dipole transition ${}^5\text{D}_0 \rightarrow {}^7\text{F}_4$ are observed



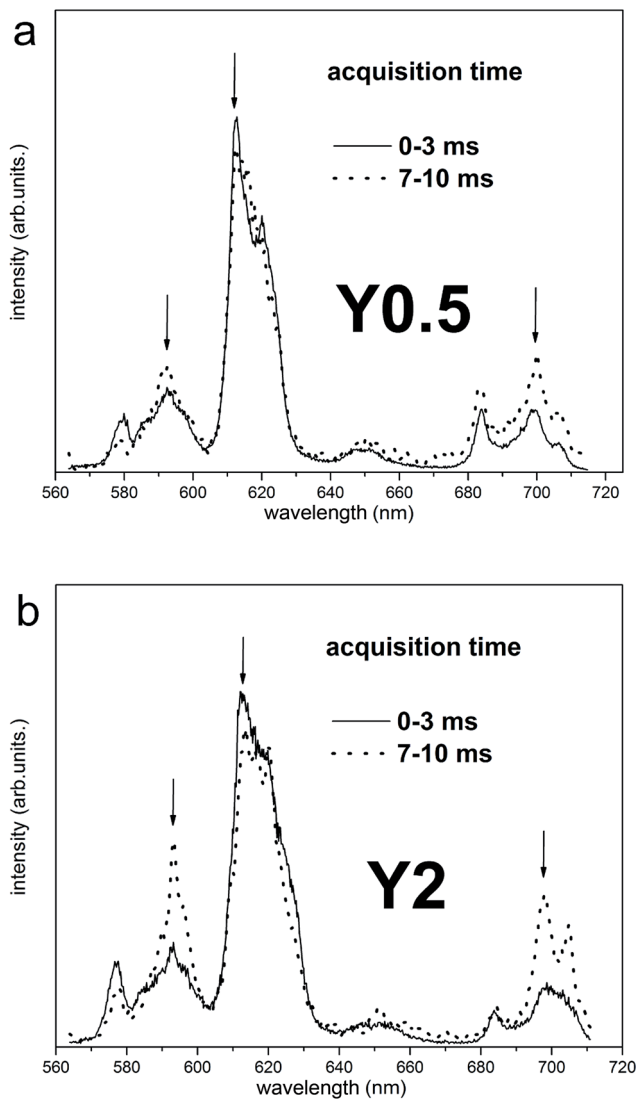


Fig. 7 Time-resolved luminescence spectra for Y0.5:0.05 Eu³⁺ (a) and Y2:0.05 Eu³⁺ (b) excited with wavelength 290 nm for two different acquisition time.

(Fig. 7b). A similar spectrum to the ones collected in 7–10 ms time acquisition was obtained for Y1, Y1.5 and Y2 samples at 394 nm steady state excitation (see Fig. 6d). The calculated decay times for Y2 ranges from 1.50 to 1.98 ms with an error of 0.2 ms monitoring ⁵D₀ → ⁷F_{1,2,4} transitions at 593, 612 and 696 nm wavelengths. Taking into account the fact that the emission profile changes in time significantly and that the decay times with a measurement error of monitoring luminescence at different wavelengths differs, we can conclude the existence of more than one Eu³⁺ site in Y2 sample. As it was discussed above, the more distorted sites are attributed to Eu³⁺ replacing divalent Ca²⁺ sites with a C₁ point group and these sites are expected to have a shorter decay time than the site of Eu³⁺ replacing Y³⁺, which is a more symmetrical site.

The normalized emission spectra obtained for the samples after the reduction using excitation wavelength equal to 270 nm are presented in Fig. 9. The chosen wavelength is justified since

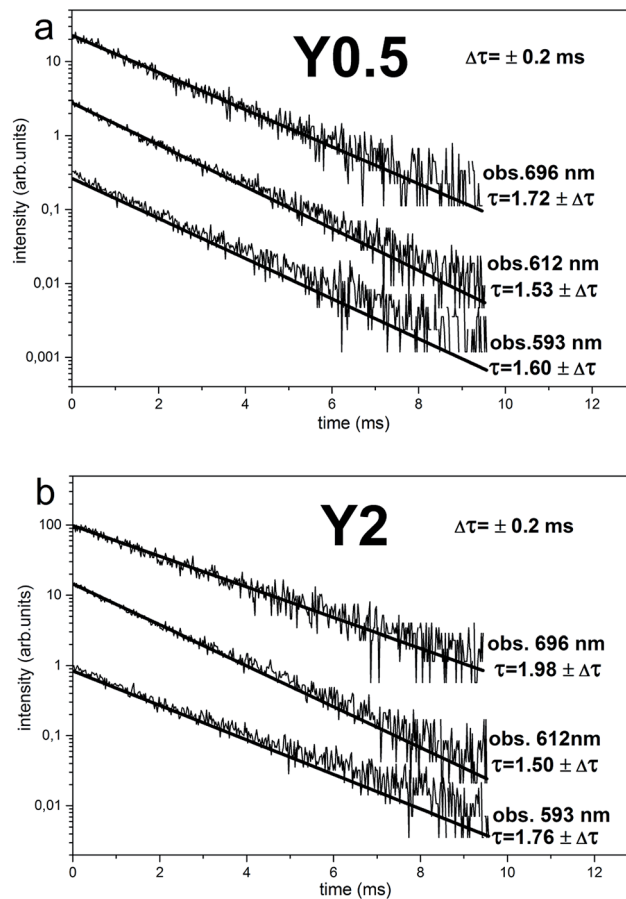


Fig. 8 Decay curves of Eu³⁺ emission for Y0.5:0.05 Eu³⁺ (a) and Y2:0.05 Eu³⁺ (b). The luminescence observed at 593, 612 and 698 nm.

it excites the europium in both oxidation states and corresponds to the intersection of both Eu²⁺ excitation spectra (Fig. 9) as well as ligand–metal charge transfer band of Eu³⁺.¹⁴ The emission spectra for all samples consist of broad bands related to the 4f⁶5d¹ → 4f⁷ transitions occurring in the Eu²⁺ ions and the narrow emission lines attributed to the ⁵D₀ → ⁷F_J transitions in Eu³⁺. The presence of the latter in all cases indicates that the complete reduction did not occur. However, the intensity of the Eu³⁺ in comparison to intensity of Eu²⁺ luminescence diminishes with increasing of the nominal concentration of yttrium in the series of samples. The luminescence of Eu²⁺ consists of two separated bands: first one with a maximum at 420 nm and the second with a maximum at 488 nm. The contribution of these two bands to the overall emission spectra changes, depending on the Y/Ca ratio in the studied materials. In PL spectrum of the Y0.5 sample, the band with a maximum at 420 nm dominates. For higher concentration of Y in the material series (Y0.75, Y1), the contribution of the band at 488 nm suddenly emerges, and for the materials with an excess of yttrium (Y1.5, Y2), the band at 488 nm can solely be seen. The Fig. 9 also shows the excitation spectra observed for 430 and 510 nm wavelengths. As it can be seen the PLE spectra, the line shape depends on the observation wavelength. The excitation spectrum collected for emission, monitored at 430 nm, consists of a single band with a maximum at 268 nm, whereas in the PLE



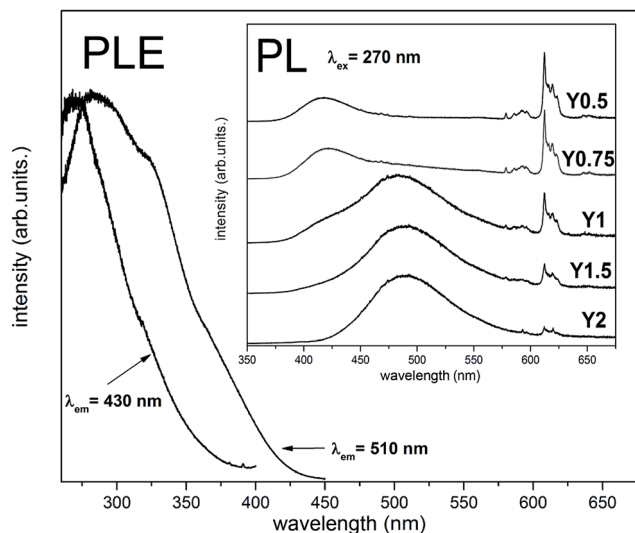


Fig. 9 Excitation spectra monitored at two wavelength corresponding to maxima of luminescence (for materials Y0.5 and Y0.75—obs. 430 nm and for materials Y1, Y1.5 and Y2—obs. 510 nm). Inset: the emission spectra of $\text{Ca}_9\text{Y}(\text{PO}_4)_7$ after reduction, collected at RT under excitation at 270 nm for all samples.

spectrum, monitored at 510 nm, two clear broad bands with maximums at 285 nm and 325 nm can be seen.

To analyze the Eu^{3+} sites remaining after the reduction process, emission spectra under 393 and 394 nm excitation were also measured and the results are presented in Fig. 10. The emission and excitation spectra of Eu^{3+} collected for all materials after the reduction have the same profile as respective ones before the reduction, but with lower intensities. Here, as previously, it can be observed that Eu^{3+} ions incorporated into

the Y0.5 and Y0.75 samples occupy only one site, as opposed to the two sites in the Y1 and Y2 samples.

Taking into account the criteria for short and long distance charge compensation described in detail in ref. 29 the temperature dependence of $\text{Eu}^{2+}/\text{Eu}^{3+}$ emission intensity ratio was also checked for Y0.5 and Y2 materials. The spectra are presented in Fig. 10a and b. The excitation wavelength was 270 nm since it allows to excite the Eu^{3+} through the charge transfer transition and the Eu^{2+} through the $8\text{S}_{7/2} \rightarrow 4\text{f}^65\text{d}$ transition. The mutual relationship of the emission intensity of Eu^{3+} and Eu^{2+} depends on the given sample, as discussed above. The emission spectra were scaled to Eu^{3+} intensity (multiplied or divided by a specific value) to maintain a constant intensity of Eu^{3+} emission for all temperatures for a given sample. Such a normalization is justified assuming that there are no non-radiative processes that cause Eu^{3+} emission decrease in measured temperature range²⁴ and allows to observe changes in Eu^{2+} intensity. It is seen in Fig. 11 that for both Y0.5 and Y2 samples, the intensity of Eu^{2+} decreases in respect to Eu^{3+} when temperature is increased from 10 K to 400 K. Since the intensity of Eu^{2+} does not increase when temperature increases (and in fact decreases due to non-radiative processes), we postulate that the compensation of Eu^{3+} replacing Ca^{2+} ions takes place at a short distance for both samples.

Mechanisms proposed

The observed differences in the emission spectra collected for samples obtained after the reduction can be qualitatively explained considering differences in charge compensation mechanisms. The differences arise from the changes in Y/Ca concentration ratio in materials series. In the case of materials with deficiency of Y^{3+} (Y0.5 and Y0.75), the additional positive charge created during the incorporation of Eu^{3+} ions into the Ca^{2+} sites (in those materials only Ca^{2+} sites are occupied by the Eu^{3+} ; see discussion of Fig. 5 and 8) is compensated by the chemically induced Ca' defects. Moreover, the concentration of Ca' is higher than Eu_{Ca} and the excess of the negative charge is compensated by the e.g. oxygen vacancies ($\text{V}_{\text{O}}^{\bullet\bullet}$). Generally, during a reduction process, the gaseous hydrogen, considered as a reducing agent, creates additional oxygen vacancies ($\text{V}_{\text{O}}^{\bullet\bullet}$) on the surface of the material particles.³⁰ Because the charge compensation in the materials has a local character (short distance compensation), the reduction of Eu^{3+} to Eu^{2+} will occur, when the compensation defect of Eu_{Ca} is locally eliminated. The $\text{V}_{\text{O}}^{\bullet\bullet}$ created on the surface during the reduction cannot cause the elimination of Ca' from the surroundings of Eu^{3+} and for the reduction of the europium ion, the created oxygen vacancy has to be moved from the surface to the vicinity of the Eu^{3+} and Ca' . The situation described above is presented in Fig. 12. This kind of compensation mechanism changes the local environment of Eu^{2+} center, because one oxygen anion is removed from the vicinity of a reduced europium ion. The emission occurring from this luminescent center is strongly shifted to the blue region in comparison to the other Eu^{2+} site observed in $\text{Ca}_9\text{Y}(\text{PO}_4)_7$. The maximum of Eu^{2+} appears at 420 nm. Such an effect has been observed by Fang *et al.*³¹ in

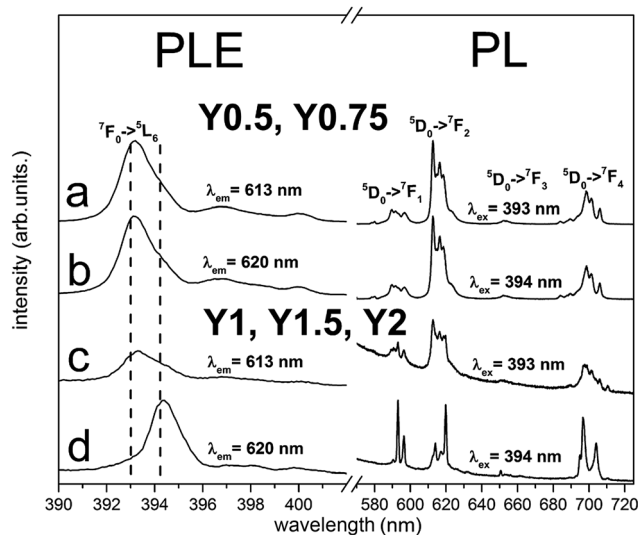


Fig. 10 PL and partial PLE spectra of $\text{Ca}_9\text{Y}(\text{PO}_4)_7:\text{Eu}$ after reduction measured at room temperature. PLE spectra of the emission were recorded at (613 nm) for all samples and at 620 nm for Y1–Y2. The emission spectra was collected under excitation at 393 nm and 394 nm for all samples.



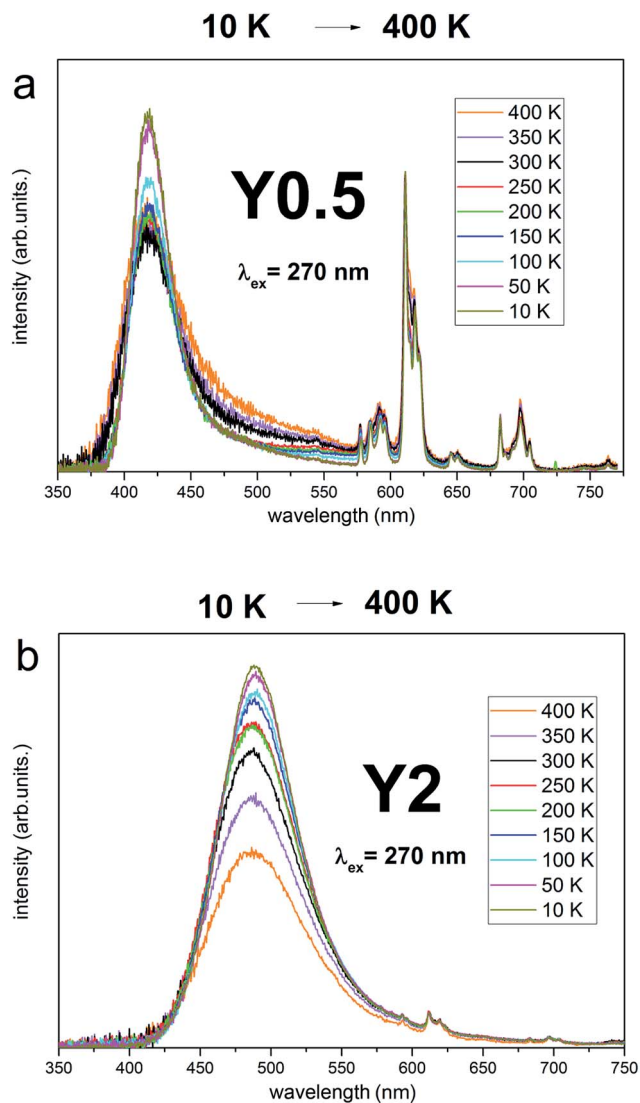


Fig. 11 Luminescence spectra of Y0.5:Eu²⁺,Eu³⁺ (a) and Y2:Eu²⁺,Eu³⁺ (b) at different temperature under pulsed excitation with 270 nm.

BaMgSiO₄:Eu and has been attributed to the creation of oxygen vacancy in the first coordination sphere of the europium ion. It is worth noting that the emission of Eu²⁺ with a maximum at

420 nm is typical for europium doped Ca₃(PO₄)₂ in rhombohedral form.^{32,33} Thus, the explanation can be the presence of a local crystal domain: *e.g.* similar symmetry and coordination number of Eu²⁺ as in the case of Ca₃(PO₄)₂. This situation is achieved after the removal of the oxygen anion from the vicinity of europium in the samples with an excess of calcium.

In the studied phosphates, the oxygen ions are a part of phosphorus tetrahedrons, in which elements are bonded with partially covalent bonds. For this reason, the energy necessary for migration of oxygen vacancies, V_O^{••}, is relatively high (compared to *e.g.* calcium vacancies, V_{Ca}^{••}). It means that *via* the proposed compensation mechanism, only a small part of Eu³⁺ ions (mostly located on the surface) are reduced during the process and intense emission lines of Eu³⁺ are observed in the spectra (Fig. 9).

A different situation is observed in the case of materials with an excess of Y³⁺ (Y1.5 and Y2). Here, the Eu³⁺ ions are incorporated in Ca²⁺ and Y³⁺ sites. The Eu_{Ca}^{••} impurity does not need compensation due to the same oxidation state of Eu³⁺ and Y³⁺, while the additional charge of Eu_{Ca}^{••} defects is compensated by the naturally created cation vacancies (*e.g.* V_{Ca}^{••}). Moreover, in the Y1.5 and Y2 materials, chemically induced, positively charged Y_{Ca}[•] defects are also compensated by the cation vacancies. During the reduction process, the V_O^{••} defects are created in the same way as described previously. Due to the short distance charge compensation, the reduction of Eu³⁺ will occur if Eu_{Ca}^{••} compensating agent (V_{Ca}^{••}) is eliminated. Due to the difference in the type of calcium and oxygen bonding with the nearest neighbors, it is expected that the energy necessary for migration of V_{Ca}^{••} is much smaller than in the case of the V_O^{••} defects. From the energy difference, it stems that the migration of calcium vacancies to the surface is much more probable than the migration of oxygen vacancies to the bulk crystal to compensate each other.³⁰ Fig. 13 shows the scheme of the described situation. This mechanism does not change the local environment of the reduced europium ions and the luminescence of Eu²⁺ occurs with a maximum of the emission band located at 488 nm.²⁴ Due to the high migration rate of V_{Ca}^{••}, all Eu³⁺ ions compensated with the calcium vacancies are reduced and the Eu²⁺ luminescence is relatively intense. However, a part of the europium ions is incorporated into the yttrium sites without compensation defects. It suggests that the luminescence of

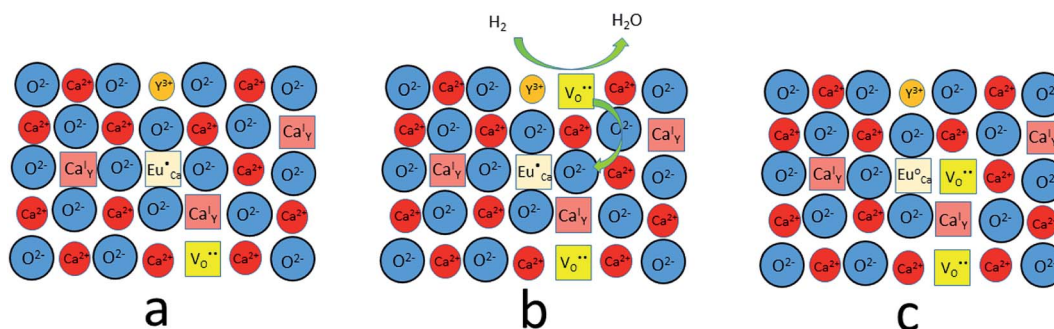


Fig. 12 A scheme of the compensation mechanism occurred in deficiency of Y³⁺; (a) defects created after synthesis in air; (b) reduction process; (c) distribution of defects after reduction process.



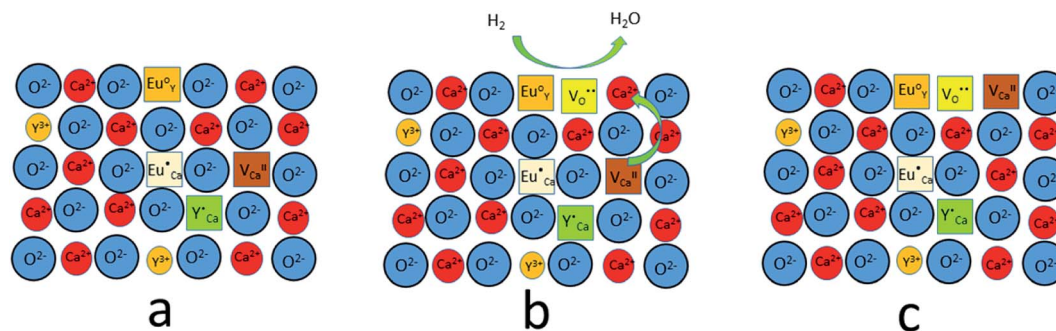


Fig. 13 A scheme of the compensation mechanism occurred in excess of Y^{3+} ; (a) defects created after synthesis in air; (b) reduction process; (c) distribution of defects after reduction process.

unreduced Eu^{3+} ions observed in the emission spectra (Fig. 10) mostly originates from the Eu^{2+} luminescent centers. However, it is assumed that both compensation mechanisms are present in all materials, depending on the randomly occurring distribution of Eu^{3+} compensating defects. It should be emphasized that the probability of encountering a given compensation mechanism depends on the Y/Ca concentration ratio. For the materials with deficiency of Y^{3+} , the mechanism that includes V_O'' migration is more probable and in the case of the materials with the excess of Y^{3+} , a more probable mechanism involves migration of V_{Ca}'' defects. For the stoichiometric material (Y1), both of them are almost equally probable.

Conclusions

The series of luminescent materials based on $Ca_3Y(PO_4)_7$ matrix with different Y/Ca concentration ratios were obtained in order to study the influence of charge compensation defects on the luminescence of europium ions incorporated into the matrix. Depending on the value of the ratio, different defects are created to compensate the additional positive charge of Eu_{Ca} . For the materials with the deficiency of Y^{3+} , Ca_Y' are mainly created, while for the materials with the excess of Y^{3+} , the main defects are V_{Ca}'' . It was assumed that both of the defects have different migration rates. The differences in the luminescence of Eu^{2+} observed in the series of materials were attributed to the differences in the defects migration rates. Moreover, it has been shown that the presence of Eu^{2+} emission band with a maximum at 420 nm observed for the materials with the deficiency of Y^{3+} is a result of changes in the local environment of Eu^{2+} center, because one oxygen anion is removed from the vicinity of the reduced europium ion.

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