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Crystallinity depends on choice of iron salt precursor in the continuous hydrothermal synthesis of Fe–Co oxide nanoparticles

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A series of Fe–Co oxide nanoparticles (NPs) were prepared by a continuous hydrothermal method using iron nitrate and ammonium iron citrate as alternative iron precursors. The crystallinity, Fe/Co composition and element spatial distribution in the synthesised NPs were investigated using X-ray diffraction (XRD), aberration-corrected scanning transmission electron microscopy (ac-STEM) imaging, energy-dispersive X-ray spectroscopy (EDX) and electron energy loss spectroscopy (EELS). Strong dependence on the choice of iron salt was observed. We demonstrate that the presence of ammonium citrate markedly improves the crystallinity of the NPs; an ordered cobalt ferrite alloy is formed. We suggest this is due to the formation of a homogenous reaction environment during the thermal decomposition of ammonium citrate, and the formation of complexes among citrate, Fe and Co ions.

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Introduction

Magnetic nanoparticles (NPs) have shown great potential for a number of important technological applications, such as information storage, medical diagnosis, drug delivery, *etc.*^{1–3} For example, in an alternating magnetic field, magnetic NPs are able to convert electromagnetic energy into thermal energy, which is useful for treatment of cancer by magnetic hyperthermia.^{4,5} Fe–Co oxide NPs have demonstrated great potential linked to their distinctive properties. They show large magnetic anisotropy, moderate saturation magnetization, chemical stability and mechanical hardness.^{6–8} The magnetic properties of Fe–Co oxide NPs are believed to be related to their morphology, composition, and cation arrangements.^{9–11} Joshi *et al.*⁹ reported that the saturation magnetization of cobalt ferrite NPs increases with an increase in size, and that spherical NPs exhibit higher saturation magnetization than faceted irregular ones. The effect of stoichiometry, *i.e.* metal ratio, on the magnetic properties of cobalt ferrite NPs was studied by Sathya *et al.*¹⁰ They reported that the specific absorption rate (relevant for magnetic hyperthermia) and the relaxivity value (significant for magnetic resonance imaging, MRI) of the cobalt ferrite (Co_xFe_{3–x}O₄) NPs could be tuned not only by the NP's

size, but also by the Fe/Co ratio. They concluded that particles of 20 ± 2 nm in size and Co stoichiometry (x) in the $x = 0.5–0.7$ range are promising candidates as heat mediators for both hyperthermia and MRI applications. Fe–Co oxide NPs usually have a spinel structure with unit formula AB₂O₄, where A and B stand for tetrahedral and octahedral cation sites within the close packed array of oxygen atoms.¹² The metal ion arrangement at different sites also has a strong influence on magnetic properties. Previous studies^{13–15} have reported the valence states of the Fe and Co cations and their corresponding distributions in the A and B sites of the cobalt ferrite structure with Mössbauer spectra. Liu *et al.*¹¹ reported that the migration of Co²⁺ ions from A to B sites could result in an increase in the saturation magnetization and coercivity.

The hydrothermal method is one of the most widely used synthesis protocols for the preparation of metal oxide NPs.^{16,17} By adjusting the synthesis conditions, such as pH, reaction temperature, Co to Fe ratio and surfactants, the particle morphology and Co to Fe ratio in the NPs can be adjusted.^{16,18,19} The cation distribution in A or B sites can be tuned by changing the reaction time¹¹ or the quenching speed after post-annealing of the Fe–Co oxide NPs.^{12,20} However, the effect of synthesis conditions on the element spatial distribution of Co and Fe in the NPs is rarely discussed. The degree of inhomogeneous alloying in the NPs will undoubtedly affect their magnetic properties; Antoniak *et al.*²¹ found that the magnetic moment of FePt NPs was reduced by 20–30% with respect to the bulk material due to compositional inhomogeneity in the particles. Furthermore, López-Ortega *et al.*²² synthesized magnetic core/shell NPs of Mn_xFe_{3–x}O₄/Fe_xMn_{3–x}O₄ using a seeded growth procedure; the heterostructured NPs showed a strong exchange

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coupling between the core and shell resulting in a coercivity increase.

Ac-STEM is a powerful tool for characterising binary complex NPs using Z-contrast imaging^{23,24} and STEM-EELS analysis.^{25–27} Here we studied the morphology, Fe/Co ratio and Fe and Co distribution of the hydrothermally prepared Fe–Co oxide NPs by ac-STEM. We found that by changing the iron salt, from iron nitrate to ammonium iron citrate, the internal order (crystallinity, elemental distribution, *etc.*) in the Fe–Co oxide NPs can be changed.

Synthesis procedure and experimental method

The synthesis of a series of Fe–Co oxide samples at Promethean Particles Ltd (Nottingham, UK) was conducted in water using a continuous-flow hydrothermal reactor, which has previously been described in the literature;^{28–32} a schematic diagram is shown in Fig. 1. Briefly, deionised water was pumped into an electric preheater at a rate of 20 ml min⁻¹. The water was heated to 400 °C and flowed into the nozzle reactor section of the system, where it met a flow of aqueous metal salt solution at ambient temperature, pumped in at a rate of 10 ml min⁻¹.

For each sample, the metal salt solution was composed of the specified concentration ratio of iron and cobalt ions, with total concentration 0.05 M. The cobalt salt used was cobalt(II) acetate tetrahydrate (Co(C₂H₃O₂)₂·4H₂O, Sigma-Aldrich). Two different iron salts, iron(III) nitrate nonahydrate (Fe(NO₃)₃·9H₂O, Sigma Aldrich) and ammonium iron(III) citrate (C₆H₈O₇·xFe³⁺·yNH₃, Sigma-Aldrich), were used to synthesise sample series INN (iron nitrate nonahydrate) and AIC (ammonium iron citrate), respectively. A summary of the precursor concentrations used is given in Table 1.

For the AIC-1 sample, the water flow feed also contained 0.25% v/v hydrogen peroxide (H₂O₂, Fisher Scientific, UK). Hydrogen peroxide was used here to drive oxidation of the cobalt precursor. A reaction temperature of 400 °C was chosen to keep a high rate of conversion and the whole system pressure was maintained at 240 bar by a backpressure regulator. The

Table 1 Summary of precursor concentrations used for each sample

Sample code	Fe/Co ratio	Cobalt acetate (M)	Ammonium iron citrate (M)	Iron nitrate (M)
INN-1	1 : 3	0.0375	—	0.0125
INN-2	1 : 1	0.0250	—	0.0250
INN-3	3 : 1	0.0125	—	0.0375
AIC-1	1 : 3	0.0375	0.0125	—
AIC-2	1 : 1	0.0250	0.0250	—
AIC-3	3 : 1	0.0125	0.0375	—

collected samples were washed by centrifugation or filtration and then sent to the University of Birmingham for characterisation.

The samples were imaged in a 200 kV JEM2100F (JEOL) STEM with a spherical aberration probe corrector (CEOS). High-angle annular dark field (HAADF) images were acquired with inner and outer detector semi-angles of 62 and 164 mrad (camera length 10 cm), and probe convergence semi-angle of 19 mrad. The elemental composition was determined by EDX spectra, which were obtained with a Bruker XFlash 4030 detector attached to the microscope. The EELS spectra were recorded with a Gatan Enfina detector at a camera length of 2 cm and a collection semi-angle of 57.8 mrad. Zeta potential measurements were made with a Malvern Zetasizer ZS system. Measurements were performed at 21 °C and repeated at least 5 times per sample using a low volume zeta potential cell, which was washed with ultra-high purity (UHP) water in between each sample. The X-ray diffraction (XRD) patterns were collected using a Bruker D8 powder diffractometer using Cu K α radiation ($\lambda = 1.5406 \text{ \AA}$).

Results and discussion

Fig. 2 shows the HAADF images and corresponding histograms of NP diameter for the samples from the INN and AIC batches. In both cases, the NPs were aggregated. However, the level of agglomeration of the NPs in the AIC sample series was higher than that in the INN series, which agrees with the zeta potential measurements. The average zeta potential (shown in Table 2) of the INN samples is larger than the AIC series, indicating a higher stability and lower level of agglomeration. From the size distribution histograms, it can be seen that the diameter ranged from around 5 to 30 nm, and that the AIC sample series had more large particles (>15 nm). Overall, the average diameters of the two batches of samples were close to each other. Changing the iron salt in the solution seems to have an effect on the dispersion of the hydrothermally synthesised Fe–Co NPs.

The XRD results from the INN and AIC sample series are shown in Fig. 3. The XRD spectra of the INN samples (Fig. 3a) suggested that the Fe–Co oxide NPs from INN-2 and -3 are not well crystallised, since the diffraction peaks were weak and broad. The crystallite sizes of Fe–Co nanoparticles in the INN series are calculated using the Scherrer equation,³³

$$D = K\lambda/\beta \cos \theta,$$

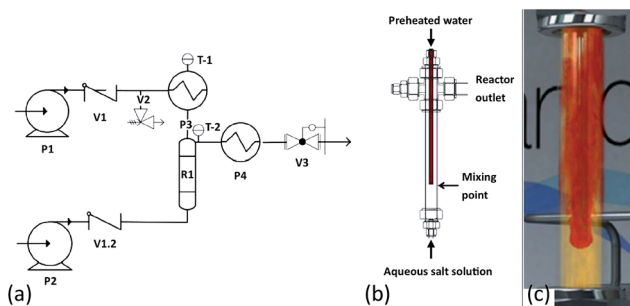


Fig. 1 (a) Flow diagram of continuous flow supercritical water reactor system: P1 and P2 – Gilson HPLC pumps, P3 – water preheater, R1 – nozzle reactor, P4 – cooler, V1, V1.2 – check valves, V2 – pressure relief valve, V3 – Tescom backpressure regulator, T-1,2 – thermocouples. (b) Nozzle reactor design.²⁹ (c) An animated depiction of the nozzle reactor.³⁰



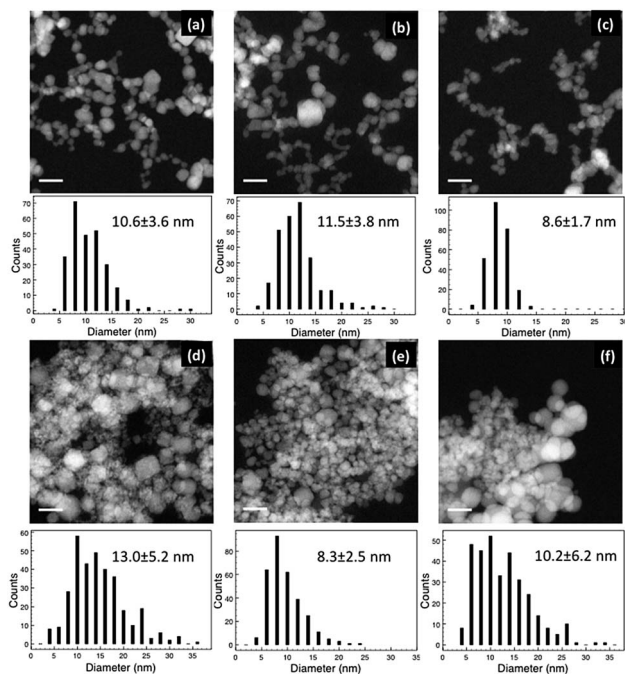


Fig. 2 HAADF images of sample (a) INN-1, (b) INN-2, (c) INN-3, (d) AIC-1, (e) AIC-2 and (f) AIC-3. The corresponding particle diameter distribution histograms are shown underneath. The scale bar is 20 nm.

Table 2 Zeta potential measurements

Sample code	Zeta potential (mV)
INN-1	17.8 ± 0.6
INN-2	20.5 ± 0.7
INN-3	19.7 ± 0.5
AIC-1	18.8 ± 0.4
AIC-2	8.5 ± 0.3
AIC-3	-7.3 ± 0.2

where D is the mean size of the crystalline domains, K is the shape factor and takes the value 0.9 in the calculation, λ is the X-ray wavelength, β is the full width half maximum (FWHM) of the (311) peak obtained from their Gaussian fittings, and θ is the Bragg angle. The average crystallite sizes for sample INN-1, -2 and -3 are 5.8 nm, 3.6 nm and 3.8 nm, respectively. These values are much smaller than those measured from the STEM-HAADF images, which suggests that the Fe-Co NPs probably have segregated Fe-rich and Co-rich regions since these regions have different lattice parameters when the Fe/Co ratio varies.^{15,34} This was also supported by the EELS mapping analysis shown below. Fig. 3b shows clear XRD patterns from the AIC samples, which correspond to the AB_2O_4 structure.^{35,36} The average FWHM value of the (311) peaks for the AIC samples is 0.9° , which is much smaller than that of the INN series (about 2.0°). This indicates that these particles have bigger crystalline domains than those in the INN sample series. The average crystallite sizes calculated from XRD patterns of AIC-1, -2 and -3 are around 11.4 nm, 7.9 nm and 9.1 nm, respectively, which are very close to the

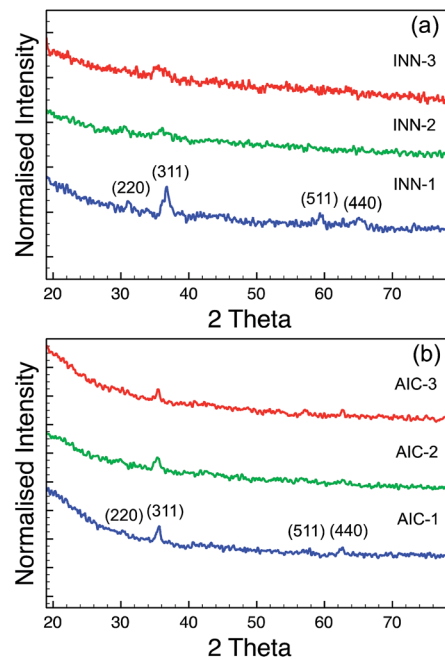


Fig. 3 XRD patterns of the sample series (a) INN and (b) AIC.

microscope measurements, indicating a high level of crystallinity in the particles.

Table 3 shows the elemental compositions of each sample in the INN and AIC series. It is important to note that the Fe/Co ratios of the synthesised INN samples are close to the ratios in the metal salt solutions. In the case of the AIC samples, these two ratios are very different, and the Fe/Co ratios detected in the particles are always higher than in the solutions. This may be caused by the ammonium citrate in the solutions. In the hydrothermal synthesis of iron oxide particles from ammonium iron citrate, it has been reported that the Fe^{3+} can be reduced to Fe^{2+} by the carbon monoxide formed from the thermal decomposition of ammonium citrate in the supercritical water.^{37,38} In a reducing environment, the oxidation of Co^{2+} to Co^{3+} will also be highly suppressed and the precipitation of Co ions from the solution will become more difficult.¹⁶ This could explain why the Co content detected in the AIC samples was always lower than expected. For the samples in the INN series, where iron nitrate was used as the precursor, the nitrate ions might act as an oxidation agent to drive the Co^{2+} to Co^{3+} .

Table 3 Elemental compositions of the samples from INN and AIC batches determined by STEM-EDX

Sample code	Fe/Co ratio	Element (at%)		
		O	Fe	Co
INN-1	0.37 ± 0.02	62.2 ± 2.4	10.2 ± 0.4	27.6 ± 1.0
INN-2	0.91 ± 0.05	61.0 ± 2.7	18.6 ± 0.8	20.5 ± 0.8
INN-3	3.17 ± 0.15	66.4 ± 2.4	25.6 ± 0.8	8.1 ± 0.3
AIC-1	1.87 ± 0.08	58.4 ± 1.9	27.1 ± 0.8	14.5 ± 0.5
AIC-2	2.56 ± 0.11	62.6 ± 1.9	26.9 ± 0.8	10.5 ± 0.3
AIC-3	3.22 ± 0.12	51.5 ± 1.4	37.0 ± 1.0	11.5 ± 0.3



The element spatial distribution of the metals in the samples was studied by STEM-EELS and some mapping results are shown in Fig. 4. We found that the element spatial distributions in the NPs synthesised with iron nitrate (INN batch) were not uniform, as is shown in Fig. 4a and b. A Co-rich region tended to lie at the surface of the particles and an Fe-rich area in the core. However, a complete core-shell structure was not observed, and in some cases, even a “dumbbell” distribution of Fe and Co was also observed. The non-uniform elemental distribution in the NPs of INN samples may be caused by the nitrate anions in the iron salt and the solubility of the Fe and Co ions in the solutions. Previous studies^{16,39} implied that Co oxide NPs could not be precipitated from cobalt nitrate solutions in the absence of a base, or that only a few large particles could be produced when Co nitrate was in contact with supercritical water.⁴⁰ In contrast, the solubility of Fe ions in the iron nitrate solution is several orders of magnitude lower than that of Co in its nitrate solution.¹⁶ Therefore, the nucleation of Fe oxide or hydroxide particles is easier than that of Co oxide or hydroxide during hydrothermal processing. Consequently, the Co cations may then proceed to precipitate on the surface of Fe oxide particles due to ion diffusion in the solution, forming Co and Fe-rich areas. This also explains why the Fe-rich regions tend to be in the core and Co-rich regions on the surface.

For the AIC sample series, we found that the Fe and Co spatial distributions in these NPs were very uniform (shown in Fig. 4c and d). In this case, cobalt ferrite is the likely product as the XRD results indicate. There are two possible reasons for the uniform elemental distributions. First, the CO molecules produced by the

thermal decomposition of citrate^{37,41} are expected to be miscible with supercritical water and provide a homogeneous reaction environment promoting the oxidizing and reducing capability.³⁸ Second, citrate anions can form complexes with Fe and Co cations, respectively. Citrate complexes of Fe or Co have been reported elsewhere in the literatures;^{42,43} the citrate ligand binds to Fe or Co atoms through the O atom of the hydroxyl group and the carboxyl group. A formation of these complexes (among Fe ions, Co ions and citrate) in the cobalt iron salts solution will promote the simultaneous precipitation of Fe and Co ions during the thermal decomposition process,^{41,44} resulting in a uniform elemental distribution.

Controlling the elemental arrangement in magnetic alloy NPs is of great importance, not only because it can change the magnetic properties of the particles²¹ but also because it might influence their toxicities and biocompatibilities. Iron oxide NPs are widely used in *in vivo* applications because of their low toxicity and biocompatibility.^{45,46} For the Fe-Co oxide NPs, the non-uniform distribution of Co oxide (especially when the Co rich region is on the surface of the NPs) may increase the toxicity of the particles, since it has been reported that Co NPs can induce cytotoxicity and genotoxicity in cells.^{47–49} A toxicity investigation of the Fe-Co oxide NPs reported in this work is in progress, and may shed light on the effect of the spatial distributions of the metals in the NPs on the toxicity.

Conclusions

In summary, we have explored a continuous hydrothermal method to prepare Fe-Co oxide NPs using two different iron salts as precursors. We found that the iron salt anions have a significant effect on the degree of crystallinity and aggregation, composition of the NPs, and spatial distribution of the metals in the NPs. The Fe-Co oxide NPs prepared with ammonium iron citrate were found to be well crystallised and had a more uniform Fe and Co distribution than the particles synthesised with an iron nitrate precursor. We suggest that ammonium citrate provides homogeneous reaction environment and promotes the simultaneous precipitation of Fe and Co ions by forming complexes. Overall, our investigation into the effect of precursor salt on the continuous hydrothermal synthesis is helpful to understanding the reaction process, allowing for better control of the NPs produced.

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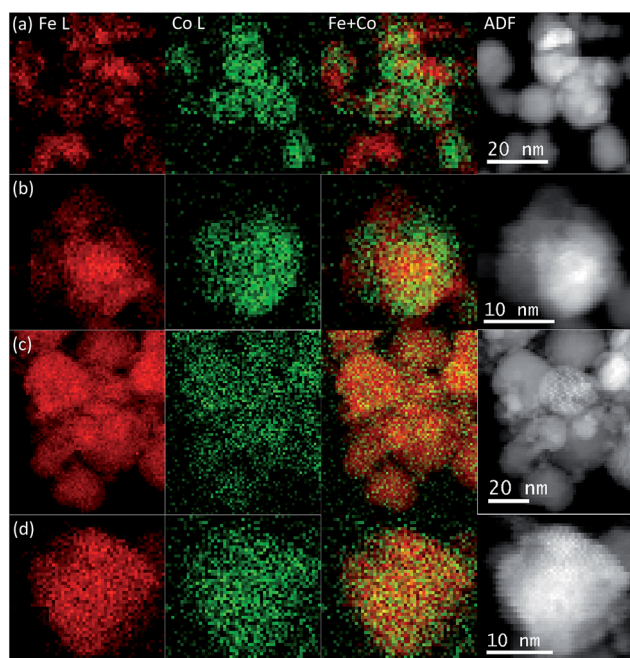


Fig. 4 Examples of EELS mapping of samples INN-2 (a) and AIC-2 (c). (b and d) Are single nanoparticle EELS maps from INN-2 and AIC-2, respectively. From left to the right we show: the Fe (red) and Co (green) L edge maps, a mix of the Fe and Co maps, and the corresponding ADF images.



Notes and references

- 1 A. K. Nikumbh, R. A. Pawar, D. V. Nighot, G. S. Gugale, M. D. Sangale, M. B. Khanvilkar and A. V. Nagawade, *J. Magn. Magn. Mater.*, 2014, **355**, 201–209.
- 2 S. Mornet, S. Vasseur, F. Grasset and E. Duguet, *J. Mater. Chem.*, 2004, **14**, 2161–2175.
- 3 P. K. Jain, X. Huang, I. H. El-Sayed and M. A. El-Sayed, *Acc. Chem. Res.*, 2008, **41**, 1578–1586.
- 4 J. Lee, J. Jang, J. Choi, S. H. Moon, S. Noh, J. Kim, J. G. Kim, I. Kim, K. I. Park and J. Cheon, *Nat. Nanotechnol.*, 2011, **6**, 418–422.
- 5 E. Fantechi, C. Innocenti, M. Zanardelli, M. Fittipaldi, E. Falvo, M. Carbo, V. Shullani, L. Di Cesare Mannelli, C. Ghelardini, A. M. Ferretti, A. Ponti, C. Sangregorio and P. Ceci, *ACS Nano*, 2014, **8**, 4705–4719.
- 6 F. Tihay, G. Pourroy, M. Richard-Plouet, A. C. Roger and A. Kiennemann, *Appl. Catal., A*, 2001, **206**, 29–42.
- 7 D. S. Mathew and R. S. Juang, *Chem. Eng. J.*, 2007, **129**, 51–65.
- 8 B. G. Toksha, S. E. Shirsath, S. M. Patange and K. M. Jadhav, *Solid State Commun.*, 2008, **147**, 479–483.
- 9 H. M. Joshi, Y. P. Lin, M. Aslam, P. V. Prasad, E. A. Schultz-Sikma, R. Edelman, T. Meade and V. P. Dravid, *J. Phys. Chem. C*, 2009, **113**, 17761–17767.
- 10 A. Sathya, P. Guardia, R. Brescia, N. Silvestri, G. Pugliese, S. Nitti, L. Manna and T. Pellegrino, *Chem. Mater.*, 2016, **28**, 1769–1780.
- 11 M. Liu, M. Lu, L. Wang, S. Xu, J. Zhao and H. Li, *J. Mater. Sci.*, 2016, **51**, 5487–5492.
- 12 G. A. Sawatzky, *J. Appl. Phys.*, 1968, **39**, 1204.
- 13 S. Kawano, N. Yamamoto and N. Achiwa, *Jpn. J. Appl. Phys.*, 1974, **13**, 1891–1892.
- 14 P. J. Murray and J. W. Linnett, *J. Phys. Chem. Solids*, 1976, **37**, 1041–1042.
- 15 P. A. Smith, C. D. Spencer and R. P. Stillwell, *J. Phys. Chem. Solids*, 1978, **39**, 107–111.
- 16 L. J. Cote, A. S. Teja, A. P. Wilkinson and Z. J. Zhang, *J. Mater. Res.*, 2003, **210**, 307–317.
- 17 T. J. Daou, G. Pourroy, S. Begin-Colin, J. M. Greneche, C. Ulhaq-Bouillet, P. Legare, P. Bernhardt, C. Leuvrey and G. Rogez, *Chem. Mater.*, 2006, **18**, 4399–4404.
- 18 D. Zhao, X. Wu, H. Guan and E. Han, *J. Supercrit. Fluids*, 2007, **42**, 226–233.
- 19 L. Zhao, H. Zhang, Y. Xing, S. Song, S. Yu, W. Shi, X. Guo, J. Yang, Y. Lei and F. Cao, *J. Solid State Chem.*, 2008, **181**, 245–252.
- 20 G. A. Sawatzky, F. Van Der Woude and A. H. Morrish, *Phys. Rev.*, 1969, **187**, 747–757.
- 21 C. Antoniak, M. Spasova, A. Trunova, K. Fauth, F. Wilhelm, A. Rogalev, J. Minár, H. Ebert, M. Farle and H. Wende, *J. Phys.: Condens. Matter*, 2009, **21**, 336002.
- 22 A. López-Ortega, M. Estrader, G. Salazar-Alvarez, S. Estradé, I. V. Golosovsky, R. K. Dumas, D. J. Keavney, M. Vasilakaki, K. N. Trohidou, J. Sort, F. Peiró, S. Suriñach, M. D. Baró and J. Nogués, *Nanoscale*, 2012, **4**, 5138–5147.
- 23 N. Jian and R. E. Palmer, *J. Phys. Chem. C*, 2015, **119**, 11114–11119.
- 24 F. Yin, Z. W. Wang and R. E. Palmer, *J. Am. Chem. Soc.*, 2011, **133**, 10325–10327.
- 25 I. Romer, Z. Wang, R. C. Merrifield, R. E. Palmer and J. R. Lead, *Environ. Sci. Technol.*, 2016, **50**, 2183–2190.
- 26 M. Baalousha, K. P. Arkill, I. Romer, R. E. Palmer and J. R. Lead, *Sci. Total Environ.*, 2015, **502**, 344–353.
- 27 Y. Li, Z. W. Wang, C.-Y. Chiu, L. Ruan, W. Yang, Y. Yang, R. E. Palmer and Y. Huang, *Nanoscale*, 2012, **4**, 845–851.
- 28 H. Hobbs, S. Briddon and E. Lester, *Green Chem.*, 2009, **11**, 484–491.
- 29 E. Lester, G. Aksomaityte, J. Li, S. Gomez, J. Gonzalez-Gonzalez and M. Poliakoff, *Prog. Cryst. Growth Charact. Mater.*, 2012, **58**, 3–13.
- 30 F. Moro, S. V. Y. Tang, F. Tuna and E. Lester, *J. Magn. Magn. Mater.*, 2013, **348**, 1–7.
- 31 E. Lester, S. V. Y. Tang, A. Khlobystov, V. L. Rose, L. Buttery and C. J. Roberts, *CrystEngComm*, 2013, **15**, 3256.
- 32 Q. Wang, S. V. Y. Tang, E. Lester and D. O'Hare, *Nanoscale*, 2013, **5**, 114–117.
- 33 P. Scherrer, *Nachr. Ges. Wiss. Goettingen, Math.-Phys. Kl., Fachgruppe 2*, 1918, 96.
- 34 T. A. S. Ferreira, J. C. Waerenborgh, M. H. R. M. Mendonça, M. R. Nunes and F. M. Costa, *Solid State Sci.*, 2003, **5**, 383–392.
- 35 Y. I. Kim, D. Kim and C. S. Lee, *Phys. B*, 2003, **337**, 42–51.
- 36 Z. Ma, L. Ren, S. Xing, Y. Wu and Y. Gao, *J. Phys. Chem. C*, 2015, **119**, 23068–23074.
- 37 T. Adschiri, K. Kanazawa and K. Arai, *J. Am. Ceram. Soc.*, 1992, **75**, 1019–1022.
- 38 T. Adschiri, Y. Hakuta and K. Arai, *Ind. Eng. Chem. Res.*, 2000, **39**, 4901–4907.
- 39 L. J. Cote, A. S. Teja, A. P. Wilkinson and Z. J. Zhang, *J. Mater. Res.*, 2002, **19**, 2410–2416.
- 40 Y. Hao and A. Teja, *J. Mater. Res.*, 2003, **18**, 415–422.
- 41 N. S. Gajbhiye and S. Prasad, *Thermochim. Acta*, 1996, **285**, 325–336.
- 42 T. Birsa Čelič, Z. Jagličič, K. Lazar and N. Zabukovec Logar, *Acta Crystallogr.*, 2013, **69**, 490–495.
- 43 Z. Zhou, Y. Deng and H. Wan, *Cryst. Growth Des.*, 2005, **5**, 1109–1117.
- 44 S. Prasad, A. Vijayalakshmi and N. S. Gajbhiye, *J. Therm. Anal. Calorim.*, 1998, **52**, 595–607.
- 45 R. Weissleder, D. D. Stark, B. L. Engelstad, B. R. Bacon, C. C. Compton, D. L. White, P. Jacobs and J. Lewis, *Am. J. Roentgenol.*, 1989, **152**, 167–173.
- 46 Y. Wu, Y. Wang, G. Luo and Y. Dai, *Bioresour. Technol.*, 2009, **100**, 3459–3464.
- 47 E. Papis, F. Rossi, M. Raspanti, I. Dalle-Donne, G. Colombo, A. Milzani, G. Bernardini and R. Gornati, *Toxicol. Lett.*, 2009, **189**, 253–259.
- 48 J. Ponti, E. Sabbioni, B. Munaro, F. Broggi, P. Marmorato, F. Franchini, R. Colognato and F. Rossi, *Mutagenesis*, 2009, **24**, 439–445.
- 49 S. Chattopadhyay, S. K. Dash, S. Tripathy, B. Das, D. Mandal, P. Pramanik and S. Roy, *Chem.-Biol. Interact.*, 2015, **226**, 58–71.

