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Modifications of polyethersulfone membrane by doping sulfated-TiO₂ nanoparticles for improving anti-fouling property in wastewater treatment

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Polyethersulfone (PES) composite membranes mixed with sulfated-TiO₂ nanoparticles were fabricated using a non-solvent induced phase inversion method. Sulfating had little effect on particle size and crystal phase of TiO₂ nanoparticles. However, the roles of each nanoparticle as doping materials to reduce membrane fouling for the PES membrane were apparent. The PES membrane embedded with sulfated-TiO₂ nanoparticles showed higher pure water permeability and larger membrane pore size while lowering fouling rate than the PES membrane embedded with bare TiO₂ nanoparticles (without sulfating) for bovine serum albumin solution as a feed. Water uptake capacity and hydrophilicity of the PES membrane was improved significantly by introducing the sulfated-TiO₂ nanoparticles into the membrane. As a dosage of sulfated TiO₂ nanoparticles was above 0.8 wt%, however, membrane porosity was reduced due to aggregation of the nanoparticles causing membrane pore blockage. The sulfated-TiO₂ nanoparticles improved membrane hydrophilicity while reducing fouling rate under strong acidity caused by incorporation of sulfate groups with PES polymer chains.

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1 Introduction

Polyethersulfone (PES) is one of the most common polymers used for the fabrication of ultrafiltration (UF) or nanofiltration (NF) membranes in water and wastewater treatment.^{1,2} The PES membranes are hydrophobic intrinsically with high chemical and thermal stability. Nevertheless, membrane fouling caused by adsorption of organic compounds present in water and wastewater on membrane surface or within membrane pores is a main hurdle to be resolved. The membrane fouling decreases membrane performance by decreasing in permeate flux with time at constant-pressure filtration or increasing in trans-membrane pressure (TMP) with filtration time at constant-flux operation. The membrane fouling requires a frequent chemical cleaning, thereby shortening membrane life-time.

Interests in nanoparticles as doping materials to improve polymeric PES membranes are growing rapidly in water and wastewater treatment applications. Many efforts have been made to embed various types of inorganic nanoparticles into membrane structure as doping materials to enhance membrane properties, *e.g.*, membrane hydrophilicity and antifouling characteristics. Generally, inorganic nanoparticles such as silica (SiO₂),³ alumina (Al₂O₃),⁴ silver (Ag)⁵ and zirconia (ZrO₂)⁶ have been applied as doping materials for the PES membrane. In particular, the TiO₂ nanoparticle is one of the mostly used as

inorganic fillers in fabricating hybrid organic-inorganic composite membrane for water treatment applications.^{7,8} The use of TiO₂ nanoparticles can provide beneficial effect on membrane functionality, *e.g.*, photocatalytic activity, with desired control of membrane fouling. Doping the TiO₂ nanoparticles into membrane structure also increases membrane hydrophilicity without significant reduction in membrane chemical and mechanical stability.

Nevertheless, high surface energy and strong van der Waals attractive force employed by TiO₂ nanoparticles should lead to their aggregation easily within the membrane structure during its fabrication.⁹ Nano-TiO₂ particles can show tendency to aggregate due to their high specific surface area and the hydroxyl group on the TiO₂ surface. The nanoparticle aggregation in casting solution reduces nanoparticle dispersions, thereby leading to less improved membrane performance due to loss of reactivity of the nanomaterials. As a result, novel functionality provided by TiO₂ nanoparticles for improving membrane property can be reduced. To overcome this problem, the TiO₂ nanoparticles are often modified by surface grafting with hydrophilic polymers such as polyacrylic acid (PAA), polyethylene glycol (PEG) or other inorganic materials such as silane.^{7,10,11} Synthesis of binary composite materials has also been considered to improve not only hydrophilicity of the inorganic fillers but also intrinsic properties of the membrane.^{2,8,12,13} Alternatively, non-stoichiometric inorganic metal oxide nanoparticles with super-hydrophilicity has also been used to maintain catalytic functionality of polymeric membrane in water treatment applications.^{14,15}

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Sulfated metal oxides are a useful group of strong acid catalysts. It played an important role particularly in industrial applications due to their high efficacy for some reactions such as hydration/dehydration reactions or polymerization processes.^{13,16} Interaction of sulfated oxides with metals is known as forming novel materials with greater acidity and higher surface area than the bare metal oxides (without sulfating).¹⁷ Because of the strong acidity and inductive property employed by S=O bonds, a large number of hydroxide groups can be formed. Sulfated-metal oxides have been studied extensively as doping materials for the proton-exchange membrane in direct methanol fuel cells (DMFC) and photocatalytic membrane reactors. Ren *et al.* made first attempt to prepare a proton conductive membrane by embedding sulfated-ZrO₂ nanoparticles into N115 membrane for DMFC applications. Proton conductivity of the membrane was improved by high acidity of sulfated-ZrO₂. Sgambetterra *et al.* also synthesized sulfated-TiO₂ nanoparticles for the PVDF-HFP membrane to enhance proton conductivity with DMFC system and photocatalytic membrane reactors.^{18–21} A few works have been published to improve the membrane performance in wastewater treatment by adding sulfated inorganic materials.^{22,23} In most of previous works, the results took the benefits of nonstoichiometric metal oxide or binary metal oxide together with sulfated metal oxide. Zhang, *et al.*,²² synthesized sulfated Y-doped zirconia nanoparticles and embedded them into PSF membrane matrix. For synthesis of this nanomaterial, yttrium oxide and zirconium oxychloride were used as the precursors to prepare Y-doped zirconia. After sulfating with sulfuric acid, the sulfated Y-doped zirconia nanoparticles were obtained. The fouling behavior of composite membranes was tested by filtering oily wastewater. The optimum dosage of this nanomaterial was reported as 15 wt% to achieve lowest fouling rate. In other work,²³ the TiO₂ was firstly deposited on the surface of SiO₂ nanotube structure, and then sulfated with sulfuric acid. The composite PSF/STSNs (sulfated TiO₂ deposited SiO₂ nanotubes) membranes (inorganic material dosage: 10 wt%) were used to filter oily wastewater. Membrane permeate flux during fouling tests was higher than that obtained by bare PSF membrane as well as the membrane embedded with pure SiO₂. Although the inorganic materials were sulfated, membrane performance was improved by several reasons, for example, hydrophilicity of metal oxides (SiO₂, TiO₂), strong activity of nonstoichiometric material (Y-doped SiO₂), binary metal oxide (TiO₂-SiO₂) and the superacidity induced by incorporation of SO₄²⁻. However, influence of bare sulphated TiO₂ nanoparticles as doping materials on membrane characterization and its performance in water treatment has not been studied yet. The objectives in this study were to synthesize sulfated-TiO₂ nanoparticles and embed them into the PES membrane. The PES mixed matrix membrane was tested for water treatment applications. Intrinsic properties and antifouling characteristics of the PES mixed matrix membrane (MMM) embedded with sulfated-TiO₂ nanoparticles at various dosages were observed using model protein compound.

2 Experimental and materials

2.1 Materials

Polyethersulfone (PES) was used as a base polymer (Solvay, Belgium). A 1-methyl-2-pyrrolidone (NMP, 99.5%) was used as a solvent (Sigma Aldrich). Titania (TiO₂) nanoparticles commercially available (Sigma Aldrich, 21 nm) and ammonium sulfate (Sigma Aldrich) were used to synthesize sulfated-TiO₂ nanoparticles. For fouling test, a bovine serum albumin (BSA, BOVOGEN, Australia) was used as a model foulant material. The BSA solution was prepared in 1 g L⁻¹ of phosphate buffer solution (PBS) to maintain solution pH of about 7.6.

2.2 Preparation of sulfated TiO₂ nanoparticles (SNP)

The sulfated TiO₂ nanoparticles were synthesized through one step sulfating method. A 1 g of TiO₂ nanoparticles were added into 10 mL ammonium sulfate solution (2 mol L⁻¹) followed by vigorous stirring for 4 h at room temperature. After that, a paste solution formed was poured into DI water and rinsed for three times. A white powder obtained was dried at 60 °C for 24 h and then calcined at 400 °C for 4 h to obtain the sulfated TiO₂ nanoparticles.

2.3 Preparation of PES mixed matrix membranes

PES mixed matrix membranes were prepared using non-solvent induced phase inversion method by immersing precipitation. The PES cast solution containing 18 wt% of the PES polymer in 1-methyl-2-pyrrolidone (NMP) solvent was applied as polymer matrix. Various amounts of sulfated TiO₂ nanoparticles were dispersed into a NMP solution under ultrasonication for 0.5 h. The PES polymer was added into the nanoparticle suspension and then stirred for 12 h at room temperature followed by degassing for 6 h. After homogenous casting solution is formed, the composite PES films were casted by a Filimograph (K4340 Automatic Film Applicator, Elcometer K4340) with 150 μm film thickness. The prepared films were immersed into DI water as non-solvent at 18 °C. After 5 min precipitation, the membranes were then moved into another water bath and soaked into it for 48 h before testing the membrane. In this study, 0, 0.1, 0.8 and 2.0 by weight percent of the sulfated TiO₂ nanomaterials were tested as additive for PES membrane. This was denoted in this study as SNP0, SNP0.1, SNP0.8 and SNP2.0, respectively. The PES membrane fabricated at 2 wt% of bare TiO₂ nanoparticles without any sulfating (NP2.0) was also tested for comparison.

2.4 Characterization of sulfated TiO₂ nanoparticles

The sulfated TiO₂ nanoparticles synthesized were characterized by using X-ray diffraction (XRD) and transmission electron microscope (TEM). The X-ray diffraction (XRD) patterns of TiO₂ nanoparticles were obtained using a Model DMAX-2500 (Rigaku) with Cu anticathode radiation. The diffractograms were recorded between 10° and 90° in steps of 0.02° with count time of 20 s at each point. Transmission electron microscopy (TEM) microphotographs and energy-dispersive X-ray (EDX) microanalyses were performed by a JEOL JEM2100F electronic



microscope operating at 200 kV. Particle size distributions of synthesized particles were measured by using Zeta Potential/Particle Size Analyzer ELS-Z.

2.5 Membrane characterization

Atomic force microscopy (AFM) images were obtained using a Nanoscope Multimode IVa (Bruker) scanning microscope under the tapping mode in air. The AFM was applied to estimate the surface roughness of the prepared PES membranes by scanning $2\ \mu\text{m} \times 2\ \mu\text{m}$ sample area. Plain and cross-sectional views of the PES membrane were conducted using a high-resolution scanning electron microscopy (HR-SEM) with an accelerating voltage of 15.0 kV.

Membrane porosity (ε) was determined by gravimetric method, as defined in the following eqn (1):

$$\varepsilon = \frac{W_w - W_d}{A \times L \times \rho_w} \quad (1)$$

where W_w is a weight of wet membrane, W_d is a weight of dry membrane, A is a membrane effective area ($4\ \text{cm}^2$), ρ_w is water density ($998\ \text{kg m}^{-3}$) and L is membrane thickness (m). The average membrane pore radius (nm) was calculated by Guerout-Elford-Ferry equation as shown in equation (2).

$$r_m = \sqrt{\frac{(2.9 - 1.75\varepsilon)8\eta LQ}{\varepsilon \times A \times \Delta P}} \quad (2)$$

where η is the water viscosity ($8.9 \times 10^{-4}\ \text{Pa s}$), Q is the volume of permeated pure water per unit time ($\text{m}^3\ \text{s}^{-1}$), and ΔP is the operational pressure which is 1.4 bar.

Membrane water uptake was estimated by measuring the weight difference between wet and dry membrane. Membrane coupon with surface area of $4\ \text{cm}^2$ was dried at $80\ ^\circ\text{C}$ for overnight and then weighed as W_d . The membrane was then soaked into DI water for 24 h. After taking out the membrane, residual water was removed gently. The weight of membrane coupon was measured immediately and recorded as W_w . Water uptake of membrane was calculated by using eqn (3) as shown below. Five membrane samples were tested for each test and the average value was taken.

$$\text{Water uptake (\%)} = \frac{W_w - W_d}{W_d} \times 100\% \quad (3)$$

Contact angle of each PES membrane fabricated was measured by sessile drops with contact angle analyzer (Phoenix-300, SEO Corporation). Five static contact angles were measured for each membrane sample.

2.6 Pure water flux and anti-compaction test

With PES membranes prepared, membrane filtration tests were conducted at dead-end filtration mode. The feed solution was pressurized into a membrane filtration cell (Millipore, effective filtration area is $0.0028\ \text{m}^2$) at constant pressure of 1.4 bar. A permeate which is produced by the membrane was collected on a reservoir located on a digital weight balance. Accumulated permeate volume with filtration time was recorded by data

acquisition system. The pure water flux (PWF) of membrane was calculated by eqn (4) below.

$$J = \frac{V}{A \times \Delta t} \quad (4)$$

where V is the permeate volume collected (L) for sampling time (Δt , h), A is the effective filtration area (m^2) and Δt is time difference. Membrane compaction was also estimated by the relative difference between initial pure water flux and pure water flux stabilized using eqn (5).

$$\text{Compaction rate (\%)} = \frac{(J_0 - J_1) \times 100\%}{J_0} \quad (5)$$

where J_0 is the initial flux of membranes ($\text{L m}^{-2}\ \text{h}^{-1}$) and J_1 is the stable flux of membranes ($\text{L m}^{-2}\ \text{h}^{-1}$).

2.7 Fouling experiment

In order to know anti-fouling properties of the PES mixed matrix membrane synthesized in this study, a BSA solution prepared ($1\ \text{g L}^{-1}$) was added into the dead-end filtration cell. The permeate flux decline with filtration time was measured at 1.4 bar for 60 min operation. The rejection of BSA compound was then also measured using eqn (6) as below.

$$\text{Rejection (\%)} = \frac{(C_f - C_p)}{C_f} \times 100\% \quad (6)$$

where C_f and C_p (mg L^{-1}) are the BSA concentration in feed solution and permeate, respectively. The BSA concentration was measured with UV-VIS spectroscopy at 280 nm as wavelength.

3 Results and discussion

3.1 Characterization of sulfated TiO_2 nanoparticles

Fig. 1 shows microscopic images of the sulfated- TiO_2 nanoparticles taken by Field Emission-Transmission Electron Microscopy (FE-TEM). Significant difference between sulfated- TiO_2 nanoparticles and virgin TiO_2 ones was not observed. The EDX analysis with the sulfated TiO_2 nanoparticles detected S element, indicating that sulfating should be incorporated with TiO_2 nanoparticles successfully (0.79% weight percent). The XRD results of sulfated TiO_2 and virgin TiO_2 nanoparticles are compared in Fig. 2. Both nanoparticles are associated with anatase and rutile crystal structure. The XRD peak appeared at 2 theta of 25, 48 and 55, which belongs to (101), (200) and (211) as anatase phase, respectively. Whereas, the peaks positioned at 2

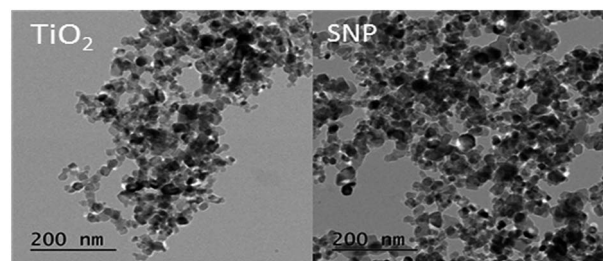


Fig. 1 TEM images of TiO_2 (left) and sulfated TiO_2 nanoparticles (right).



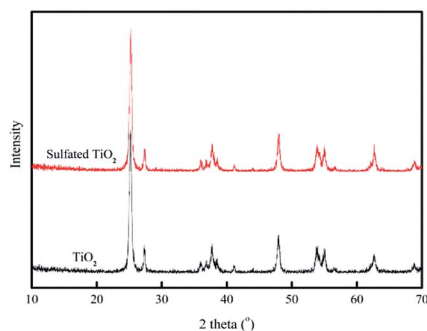


Fig. 2 XRD results with TiO_2 and sulfated TiO_2 nanoparticles.

theta of 27, 35 and 41 are related to the (110), (101) and (111) as rutile phase, respectively. Based upon XRD results, sulfating TiO_2 nanoparticles do not likely affect crystal structure of the TiO_2 particles significantly. The particle size distributions of synthesized TiO_2 nanoparticles are shown in Fig. 3. After sulfating treatment, particle size of the TiO_2 nanoparticle became smaller than the bare TiO_2 particles (128.3 nm vs. 172.5 nm), indicating that the sulphated TiO_2 particles in suspension had better dispersity.

3.2 Characterization of PES mixed matrix membrane

Fig. 4 shows SEM images of PES membrane embedded with sulfated TiO_2 nanoparticles and TiO_2 nanoparticles at different dosages. Results show more deposit of TiO_2 nanoparticles on PES membrane than that of sulfated TiO_2 nanoparticles. During phase separation in membrane fabrication, the hydrophilic TiO_2 nanoparticles can be migrated from the membrane pore matrix which is hydrophobic to the non-solvent (water) to reduce interfacial energy between casting solution and water. As a result, the TiO_2 nanoparticles tend to be released toward the membrane surface with their aggregate forms.

Cross-sectional views of each PES membranes are also shown in Fig. 4. Asymmetric cross-sectional structure of the membrane consisting of a dense skin layer and porous support layer having finger-like macroporous structure was observed.

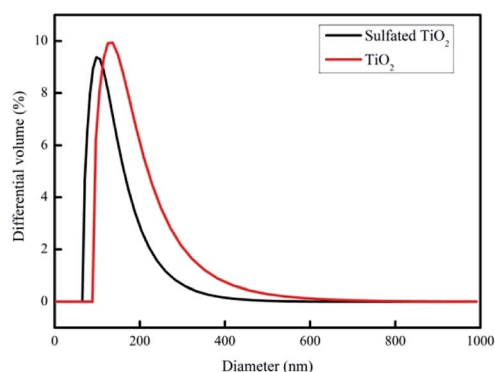


Fig. 3 Size distribution of sulfated TiO_2 nanoparticles and non-sulfated TiO_2 nanoparticles. The nanoparticles were dispersed in DI water for tests.

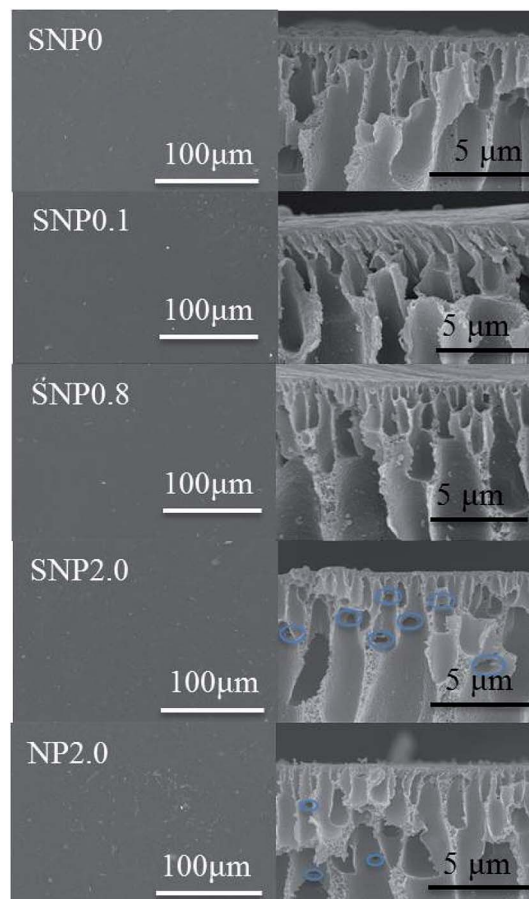


Fig. 4 Surface morphology and cross sectional view images of PES composite membranes.

Thinner skin layer as well as the appearance of lateral pores (blue circles in Fig. 4) with oval shape having horizontal orientation could result in higher membrane permeability as the sulphated TiO_2 nanoparticles were applied as doping materials.²⁴ Porosity and pore size of each membrane fabricated in this study are presented in Table 1. With TiO_2 nanoparticles at 2.0 wt%, membrane porosity was reduced possibly due to membrane pore blockage caused by aggregation of the TiO_2 nanoparticles.

Membrane hydrophilicity was estimated by measuring contact angle and water uptake capacity. Contact angle of the bare PES membrane was 68° , but it was reduced to 57° and 53° for the TiO_2 and sulfated TiO_2 nanoparticles as doping material, respectively. It was found that the sulfated TiO_2 nanoparticles increased hydrophilicity of PES membrane. This result corresponded to higher water uptake capacity of the PES membrane embedded with sulfated- TiO_2 nanoparticles, as shown in Fig. 5. Compared to the bare PES membrane (SNP0), the SNP2.0 membrane provided better anti-compaction against trans-membrane pressure (Fig. 6). Because sulfated TiO_2 nanoparticles usually have strong acidity, particle aggregation should be mitigated by resisting the pull and impact from outside surroundings.



Table 1 Membrane porosity, mean pore size, BSA rejection and surface roughness

Membrane	Porosity (%)	Mean pore radius (nm)	BSA rejection (%)	Surface roughness (nm)
SNP0	80.2 ± 2.6	11.6 ± 0.9	96.5 ± 1.1	95.7
SNP2.0	75.3 ± 1.5	16.6 ± 0.4	99.3 ± 0.2	65.8
NP2.0	73.9 ± 2.0	14.9 ± 0.4	99.0 ± 0.7	91.9

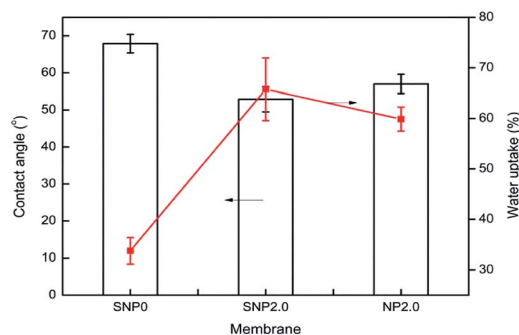


Fig. 5 Contact angle measurement and water uptake result.

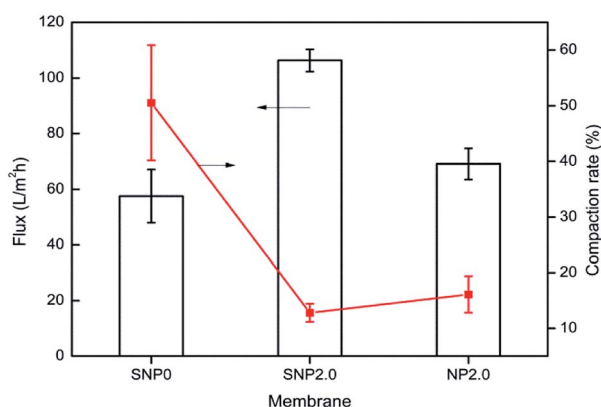


Fig. 6 Membrane pure water flux and anti-compaction properties.

In order to investigate the performance of the PES mixed matrix membrane fabricated in this study, the dead end filtration experiments were performed at constant pressure. The pure water flux was measured for each PES membrane fabricated. As shown in Fig. 6, the pure water flux of SNP2.0 membrane was two times higher than that obtained by bare PES membrane. This high permeate flux is associated with membrane pore size and hydrophilicity improved by doping SNP2.0. Pure water flux of SNP2.0 membrane was much higher than that of NP2.0 membrane (106 LMH vs. 69 LMH) at the same transmembrane pressure applied.

3.3 Membrane fouling tests

Membrane fouling is an inevitable phenomenon during membrane filtration due to the deposit of foulant on membrane surface and/or within membrane pores. At constant pressure filtration, the permeate flux should decrease with time due to

the progress of membrane fouling. The fouling rate was measured by using a dead-end filtration cell with the bovine serum albumin solution (BSA, 1 g L⁻¹). As shown in Fig. 7, the SNP2.0 membrane results in lowest fouling rate, but the rejection with BSA compound is greatest (Table 1) during 60 min of BSA filtration, there was about 40% flux decline with sulphated TiO₂ nanoparticles-embedded PES membrane. Safarpour *et al.*, observed more than 60% in flux decline with 500 mg L⁻¹ BSA filtration with PES membrane embedded with graphene oxide/TiO₂ nanoparticles.²⁵ Fang also found that more than 80% flux decline with 1 g L⁻¹ of BSA solution with iron-tannin-framework modified PES UF membrane.²⁶ Since the sulfated TiO₂ nanoparticles are more acidic and hydrophilic than TiO₂ nanoparticles, the sulfated TiO₂ nanoparticles should be combined with water molecule more easily due to their hydrogen bonds. As result, water molecules are expected to be passed through membrane pores smoothly. With AFM observations, lower surface roughness was associated with higher antifouling performance (Fig. 8). Surface roughness of the bare PES membrane (SNP0), the PES membrane embedded with

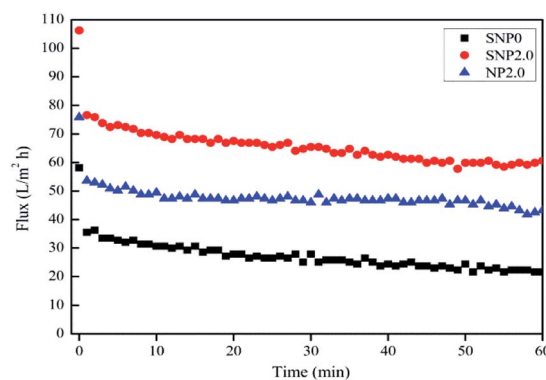
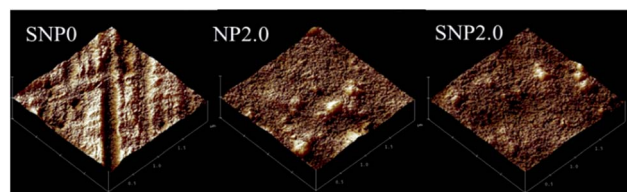


Fig. 7 Flux decline during fouling test with BSA solution.

Fig. 8 AFM images of PES membrane (SNP0), PES membrane embedded with sulfated-TiO₂ nanoparticles (SNP2.0) and PES membrane embedded with TiO₂ nanoparticles (NP2.0) at 2 wt% dosage.

sulfated-TiO₂ nanoparticles (SNP2.0) and bare TiO₂ nanoparticles (NP2.0) is 95.7, 65.8 and 91.9 nm, respectively. Membrane surface roughness reduces a repulsive energy barrier between membrane surface and foulant materials. Higher membrane surface roughness should trap more foulants, resulting in higher fouling rate.

3.4 Effect of sulfated TiO₂ dosage on membrane performance

Effect of sulfated TiO₂ nanoparticles dosage on membrane performance was investigated at different dosages of nanoparticles, 0.1, 0.8 and 2.0 wt%. Overall porosity are similar to bare PES membrane (77% for SNP0.1 and 79.6% for SNP0.8). At lowest dosage of sulfated TiO₂ nanoparticles (0.1 wt%), membrane porosity was similar to that of the bare PES membrane. This may be attributed to the well-dispersivity of the sulfated TiO₂ nanoparticles without forming their aggregate within membrane pore matrix. Highest dosage of sulfated TiO₂ nanoparticles (2.0 wt%), however, decreased membrane porosity by 6% due to aggregation of the sulfated TiO₂ nanoparticles within the membrane. The morphology of PES composite membranes doped with 0.1 wt% and 0.8 wt% of sulfated TiO₂ nanoparticles did not show any morphological difference in cross-sectional view (Fig. 4).

Based upon membrane porosity and pure water flux, mean pore size of the membrane was calculated using Guerout-Elford-Ferry equation. It was estimated that membrane pore size was increased by increasing sulfated-TiO₂ dosage to 0.1 wt%. At this dosage, exchange rate can be faster between solvent and non-solvent (water) during membrane precipitation because the sulfated-TiO₂ nanoparticles can increase hydrophilicity in casting solution. As the dosage was increased to 0.8 wt%, however, membrane pore size was reduced significantly, from 22.2 ± 1.6 nm to 12.1 ± 1.1 nm. Increasing the dosage of sulfated TiO₂ nanoparticles can increase a viscosity of casting solution, leading to delay exchange rate between solvent and non-solvent during membrane precipitation. Interestingly, further increase in the dosage of sulfated TiO₂ nanoparticles to 2.0 wt% increased membrane pore size again (16.6 ± 0.4 nm). Viscosity effect in casting solution should be overwhelmed by membrane hydrophilicity enhanced by high dosage of sulfated-TiO₂ nanoparticles.

The particle size distributions of synthesized TiO₂ nanoparticles are shown in Fig. 3. After sulfating treatment, particle size of the TiO₂ nanoparticle was smaller than the bare TiO₂ particles (128.3 nm vs. 172.5 nm). This indicates the sulphated TiO₂ particles in suspension had better dispersivity. The water

update capacity was 33.8, 46, 50.8 and 68.9% for SNP0, SNP0.1, SNP0.8 and SNP2.0, respectively. This result is consistent with higher membrane hydrophilicity at higher dosage of sulfated TiO₂ nanoparticles as observed in this study. Additionally, anti-compaction property of the PES membrane was improved by increasing sulfated-TiO₂ dosage. Results on all these information are summarized in Table 2. These phenomena can be explained by two main reasons. First, the particles with small size and large curvature can decentralize stress and impact from outside surroundings efficiently. Second, at high sulfated-TiO₂ dosage, free motion of polymer chains is partly restricted by intermolecular force between polymer chains and sulfated TiO₂ nanoparticles. This should help to improve dispersivity of the nanoparticles in membrane pore matrix. With SNP0.1, the pure water flux of the PES membrane was about four times higher than that of bare PES membrane (SNP0) (57.6 ± 9.6 vs. 217.8 ± 11.6 L m⁻² hr). However, the pure water flux was reduced to 65.6 ± 10.9 L m⁻² hr as dosage of sulfated TiO₂ nanoparticles was above 0.8 wt%. These observations indicate strongly that pure water flux of the PES membrane doped with sulfated-TiO₂ nanoparticles should be affected more by membrane hydrophilicity than membrane pore size.

Interaction between sulfated-TiO₂ nanoparticles and PES membrane is illustrated in Fig. 9a. Due to strong acidity and inductive effect by S=O bond, strong hydrogen bonds should be formed between Ti-OH group on sulfated TiO₂ nanoparticles and O element in PES chains. This interaction helps to improve both anti-compaction of the membrane and dispersivity of the nanoparticles within membrane pore matrix.

In addition to the hydroxide group employed by TiO₂ nanoparticles, the inductive effect by S=O of SO₄²⁻ groups can form

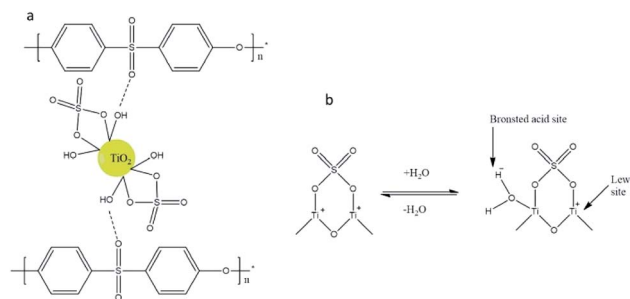


Fig. 9 Conceptual schematics to understand interaction between sulfated TiO₂ nanoparticles and PES polymer chain (a), formation of Brønsted acid site and Lewis acid site with sulfated TiO₂ nanoparticles (b).

Table 2 Membrane properties and permeability of PES composite membrane doped with different filler concentration

Membrane	Water uptake (%)	Contact angle (°)	Compaction rate (%)	Flux (L m ⁻² h ⁻¹)
SNP0	33.8 ± 2.6	67.9 ± 2.5	50.5 ± 10.3	57.6 ± 9.6
SNP0.1	46.0 ± 2.0	57.6 ± 3.1	64.7 ± 7.7	217.8 ± 11.6
SNP0.8	50.8 ± 3.1	55.1 ± 1.4	44.2 ± 10.9	65.6 ± 10.9
SNP2.0	68.9 ± 6.6	52.9 ± 3.4	12.8 ± 1.6	106.3 ± 4.0



Lewis-type acid sites and Brønsted acid sites (Fig. 9b). The Brønsted acid sites are formed by uptake of water molecules while the Lewis acid sites are formed by highly covalent property of adsorbed sulfates. Both acid sites can enhance membrane hydrophilicity with sulfated-TiO₂ nanoparticles.^{22,27}

4 Conclusions

Sulfated-TiO₂ nanoparticles were synthesized by simple sulfating method and applied as doping material for PES membrane successfully. Embedding the sulfated-TiO₂ nanoparticles into PES membrane improved membrane permeability with increasing pure water flux while reducing organic fouling as compared to the PES membrane embedded with bare TiO₂ nanoparticles as doping material. The sulfated TiO₂ nanoparticles resulted in larger membrane pore size, higher membrane hydrophilicity and better anti-compaction against transmembrane pressure compared to the bare TiO₂ nanoparticles under the same dosage. However, optimum dosage of the sulfated-TiO₂ nanoparticle existed with the PES membrane at 0.8 wt% above which no beneficial effect of the doping material was observed on the performance of PES membrane.

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Notes and references

- 1 P. S. Goh, B. C. Ng, W. J. Lau and A. F. Ismail, *Sep. Purif. Rev.*, 2014, **44**, 216–249.
- 2 A. Sotto, J. Kim, J. M. Arsuaga, G. del Rosario, A. Martínez, D. Nam, P. Luis and B. Van der Bruggen, *J. Mater. Chem. A*, 2014, **2**, 7054–7064.
- 3 X. Li, J. Huang, Y. Zhang, Y. Lv, Z. Liu and Z. Shu, *Desalin. Water Treat.*, 2015, **57**, 10980–10987.
- 4 L. Yan, Y. Li, C. Xiang and S. Xianda, *J. Membr. Sci.*, 2006, **276**, 162–167.
- 5 A. Mollahosseini, A. Rahimpour, M. Jahamshahi, M. Peyravi and M. Khavarpour, *Desalination*, 2012, **306**, 41–50.
- 6 Y. Zhang, X. Shan, Z. Jin and Y. Wang, *J. Hazard. Mater.*, 2011, **192**, 559–567.
- 7 S. S. Madaeni, S. Zinadini and V. Vatanpour, *J. Membr. Sci.*, 2011, **380**, 155–162.
- 8 Y. Zhang, F. Liu, Y. Lu, L. Zhao and L. Song, *Desalination*, 2013, **324**, 118–126.
- 9 V. M. Gun'ko, E. F. Voronin, E. M. Pakhlov, V. I. Zarko, V. V. Turov, N. V. Guzenko, R. Lebeda and E. Chibowski, *Colloids Surf., A*, 2000, **166**, 187–201.
- 10 A. Razmjou, J. Mansouri and V. Chen, *J. Membr. Sci.*, 2011, **378**, 73–84.
- 11 H. Song, J. Shao, Y. He, B. Liu and X. Zhong, *J. Membr. Sci.*, 2012, **405–406**, 48–56.
- 12 G. Zeng, Y. He, Z. Yu, Y. Zhan, L. Ma and L. Zhang, *Appl. Surf. Sci.*, 2016, **371**, 624–632.
- 13 Y. Zhang, L. Wang and Y. Xu, *Desalination*, 2015, **358**, 84–93.
- 14 P. K. Doolin, S. Alerasool and D. J. Zalewski, *Catal. Lett.*, 1994, **25**, 209–223.
- 15 K. Guan, B. Lu and Y. Yin, *Surf. Coat. Technol.*, 2003, **173**, 219–223.
- 16 X. Wang, J. C. Yu, Y. Hou and X. Fu, *Adv. Mater.*, 2005, **17**, 99–102.
- 17 R. S. Rodrigo, J. M. H. Enríquez, A. C. Mares, J. A. M. Banda, R. G. Alamilla, M. Picquart and T. L. Goerne, *Catal. Today*, 2005, **107–108**, 838–843.
- 18 J. Wu, Z. Cui, C. Zhao, H. Li, Y. Zhang, T. Fu, H. Na and W. Xing, *Int. J. Hydrogen Energy*, 2009, **34**, 6740–6748.
- 19 M. Sgambetterra, S. Panero, J. Hassoun and M. A. Navarra, *Ionics*, 2013, **19**, 1203–1206.
- 20 D. Liu, L. Geng, Y. Fu, X. Dai and C. Lü, *J. Membr. Sci.*, 2011, **366**, 251–257.
- 21 S. Ren, G. Sun, C. Li, S. Song, Q. Xin and X. Yang, *J. Power Sources*, 2006, **157**, 724–726.
- 22 Y. Zhang, P. Cui, T. Du, L. Shan and Y. Wang, *Sep. Purif. Technol.*, 2009, **70**, 153–159.
- 23 Y. Zhang and P. Liu, *J. Membr. Sci.*, 2015, **493**, 275–284.
- 24 V. Vatanpour, S. S. Madaeni, L. Rajabi, S. Zinadini and A. A. Derakhshan, *J. Membr. Sci.*, 2012, **401–402**, 132–143.
- 25 M. Safarpour, V. Vatanpour and A. Khataee, *Desalination*, 2016, **393**, 65–78.
- 26 X. Fang, J. Li, X. Li, S. Pan, X. Sun, J. Shen, W. Han, L. Wang and B. Van der Bruggen, *J. Colloid Interface Sci.*, 2017, **505**, 642–652.
- 27 G. D. Yadav and J. J. Nair, *Microporous Mesoporous Mater.*, 1999, **33**, 1–48.

