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# Synthesis, solution and solid-state fluorescence of 2-diethylaminocinchomeric dinitrile derivatives†

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A facile approach to the synthesis of novel 2-diethylaminocinchomeric dinitriles, which are found to be fluorescent both in the solution and in a solid states, was developed. Absorption, fluorescence and solvatochromic properties as well as a crystal structure of the synthesized compounds were investigated. It was found that the 2-diethylaminocinchomeric dinitrile derivatives with methoxy groups in the aryl moiety possess the most intensive emission in nonpolar solvents with fluorescence quantum yields up to 0.59.

## Introduction

Luminescent dyes are widely used in various fields of industry and science. There are many examples of optical devices, chemo- and biosensors, and nanomaterials which are based on luminescent substances.<sup>1</sup> Because of the theoretical and practical interest, compounds of the cyano pyridine series should be mentioned separately among the diverse fluorescent organic molecules.<sup>2–8</sup> They are commonly used in light-emitting diodes (OLED),<sup>2</sup> photovoltaics,<sup>3</sup> nonlinear optical (NLO) materials,<sup>4</sup> liquid crystalline,<sup>5</sup> dyes,<sup>6</sup> sensors for the detection of metal ions<sup>7</sup> and fluorescent security markers.<sup>8</sup> Such a diverse practical use of fluorophores causes an interest for the targeted synthesis of new fluorescent structures with desired properties.

It was reported earlier, that the cyano group is able to significantly increase fluorescence intensity.<sup>9</sup> Because of its structural simplicity, compactness, strong electron-withdrawing properties and high polarizability, the cyano group has been frequently utilized as a functional unit in the design of advanced optical materials.<sup>10</sup> Various «push–pull» chromophores wherein cyano group or a polynitrile moiety is an acceptor are widely presented in the modern scientific literature. A number of reports describes that including of cyano group in the molecule could increase such characteristics of the substances as solubility, thermal stability, sublimability, optical and electrochemical properties, and so on, making them interesting as advanced functional materials for novel optoelectronic devices.<sup>11</sup> Amino (or alkylamino) group, in such a kind donor–acceptor structures, is often used as a donor

moiety. It was noted that amino substituted pyridines possess a higher fluorescence quantum yield in compare to unsubstituted pyridines, and 2-aminopyridines has the most one.<sup>12</sup> Therefore, the combination of a cyano group and an amino group in one molecule is a promising approach to structures with promising optical properties.

The known compounds with such a donor–acceptor moiety are 2-aminopyridine-3-carbonitriles possessing fluorescent and other practically useful properties.<sup>7,8,13</sup> It was also mentioned that diethylamino group leads to the highest fluorescence quantum yield.<sup>14</sup>

We have reported previously, that pyridines containing two cyano groups as acceptors – 2-oxo-1,2-dihydropyridine-3,4-dicarbonitriles and 2-chlorocinchomeric dinitriles possess fluorescence properties in solutions (an intensive blue emission) and in solid state (an emission in the blue, green and yellow regions of the spectrum).<sup>15</sup> Moreover, the other authors reported that various porphyrazines could be synthesized from these cinchomeric dinitriles.<sup>16</sup> Taking into account the importance of diethylamino group as well as the described prospects of using the cinchomeric dinitrile moiety we decided to construct structures with improved fluorescent properties. Therefore, we had synthesized a series of 2-diethylaminocinchomeric dinitriles combining key fragments in the structure, and studied fluorescence properties of these compounds both in solutions and in a solid state.

## Results and discussions

2-Diethylaminocinchomeric dinitrile derivatives **3** were synthesized by the method shown in Table 1. Starting 4-oxoalkane-1,1,2,2-tetracarbonitriles **1a–c** were prepared by interaction of TCNE and carbonyl compounds in 1,4-dioxane in the presence of hydrochloric acid<sup>17</sup> as well as compounds **1d–h** were obtained according to the «solvent-free» protocol.<sup>18</sup>

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Table 1 Synthesis of compounds 3

| 2  | R <sup>1</sup>                           | R <sup>2</sup> | <sup>a</sup> Yield% |
|----|--|----------------|---------------------|
| 3a | Me                                       | Me             | 82                  |
| 3b | (CH <sub>2</sub> ) <sub>5</sub>          |                | 71                  |
| 3c | (CH <sub>2</sub> ) <sub>6</sub>          |                | 78                  |
| 3d | Ph                                       | H              | 88                  |
| 3e | 4-MeO-C <sub>6</sub> H <sub>4</sub>      | H              | 94                  |
| 3f | 3,4-Di-MeO-C <sub>6</sub> H <sub>3</sub> | H              | 92                  |
| 3g | Ph                                       | Me             | 91                  |
| 3h | 4-MeO-C <sub>6</sub> H <sub>4</sub>      | Me             | 96                  |

<sup>a</sup> Yield has been reported for isolated crude product.

Compounds **1** were further transformed to corresponding 2-chlorocinchomeronic dinitriles **2** by an interaction with hydrochloric acid for the synthesis of compounds **2a–d,g** or with generated *in situ* dry hydrogen chloride to obtain compounds **2e,f,h**.<sup>15b</sup>

Targeted 2-diethylaminocinchomeronic dinitrile derivatives **3** were synthesized as a result of an interaction between 2-chlorocinchomeronic dinitriles **2** and diethylamine. The reactions were carried out in an excess of diethylamine. It was observed, that the interaction carried out in the organic solvents (alcohols, dioxane, THF were tested) took much more time. For example, the reaction of 2-chloro-5,6-dimethylpyridine-3,4-dicarbonitrile **3a** with diethylamine in THF takes 30 h.<sup>16a</sup> Moreover, we have found that final products are not required a further purification after the reaction carried out in the excess of diethylamine without heating. An addition of metal carbonates as a catalyst did not increase a yield of compounds **3** but led to the formation of byproducts obtaining as a result of the addition of diethylamine to cyano group.<sup>19</sup> According to the developed facile protocol a series of 2-diethylaminocinchomeronic dinitrile derivatives **3** were synthesized containing alkyl, cycloalkyl and aryl substituents, including the moieties with electron-donating methoxy groups (Table 1). Compounds **3b–h** are new and were not described in the literature.

2-Diethylaminocinchomeronic dinitrile derivatives **3** are crystalline powders from green to yellow-green color, they were found to be fluorescent both in the solution and in the solid state. There are very few dyes that exhibit intensive fluorescence both in the solid state and in solutions because generally the molecular aggregation in the solid state causes fluorescence quenching (ACQ). It causes an increased interest to study the compounds **3**.

### Spectroscopic properties in solution

The structure of the synthesized pyridines **3** containing conjugated donor and acceptor moieties prompted us to explore their photophysical properties. The absorption and emission spectra

of compound **3f** in various polar and nonpolar solvents are shown in Fig. 1, the obtained spectral characteristics are summarized in Table 2. A position of absorption maxima (342–346 nm and 409–415 nm) does not change significantly in various solvents.

The compound **3f** exhibits fluorescence in various solvents, the most intensive emission was noted in benzene, an emission band undergoes a slight red-shift and significantly loses its intensity with an increase of the polarity of the solvent. The relative fluorescence quantum yield ( $\Phi_F$ ) was also calculated (see Table 2). It decreased when the solvent polarity increased, probably, because of the charge transfer phenomena. The intermolecular charge transfer (ICT)<sup>4c</sup> phenomena occurs, when an electron is transferred from the amino group to the acceptor moiety upon light absorption, producing an (ICT) excited state, which is more sensitive to the solvent polarity. A change in the excited state leads to the change in the lifetime and hence the results are a decrease of radiative emission.

The other compounds **3** possess similar spectral characteristics which are presented in Table 3. Fig. 2 shows absorption and emission spectra of 2-diethylaminocinchomeronic dinitriles **3** in benzene. Solutions of all the studied compounds **3** are characterized with an absorption maximum in the violet region of the spectrum (399–416 nm), corresponding, probably, to the intramolecular charge transfer (ICT) according to TD-DFT calculations. The intensity of this band increases with the growth of the number of donor fragments in the molecule. Moreover, in the spectra of the compounds **3e,f,h** containing MeO– groups a new shortwave absorption band is appeared which corresponds to  $n-\pi^*$  electronic transition in the aryl substituent.

It was observed that the structure of aminopyridines **3** has a little influence on the positions of the absorption and emission maxima. Compounds **3d–h** with aryl substituent are characterized with a slight (up to 15 nm) bathochromic shift of the absorption band and a more intensive fluorescence.

A relative fluorescence quantum yield of the dimethoxy derivative **3f** is the highest ( $\Phi_F$  0.59) and comparable to the used

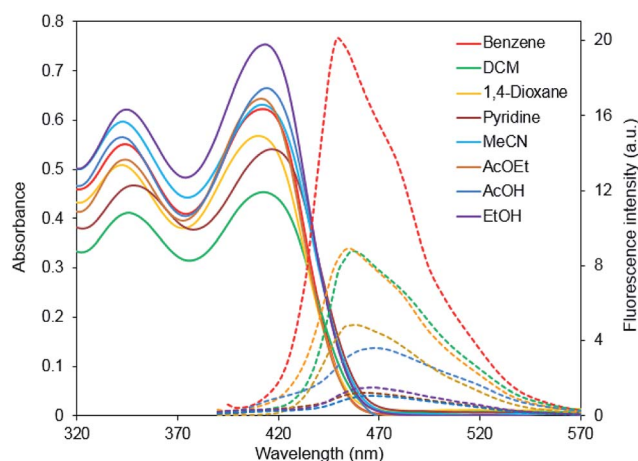


Fig. 1 Absorption (solid) and emission (dotted) spectra of **3f** in various solvents, excitation wavelength is 375 nm.

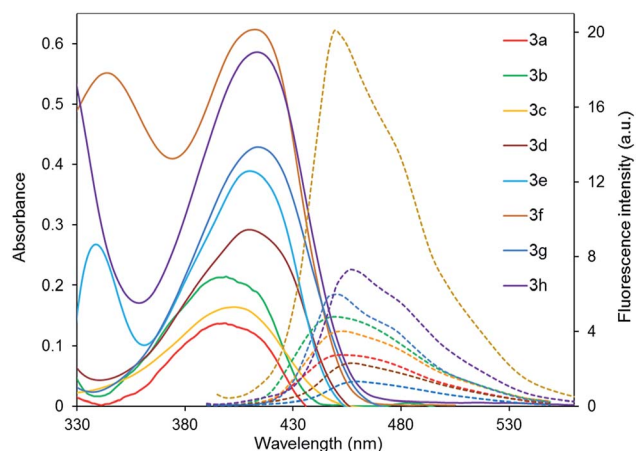


**Table 2** Spectral characteristics of the compounds **3** in various solvents

| Solvent  | $\lambda_{\text{abs;max}}$ , nm | $\log \epsilon_{\text{max}}$ | $\lambda_{\text{flu;max}}$ , nm | Stokes shift, $\text{cm}^{-1}$ (nm) | $\Phi_{\text{F}}$ |
|----------|---------------------------------|------------------------------|---------------------------------|-------------------------------------|-------------------|
| Benzene  | 343                             | 4.04                         | 449                             | 2000 (37)                           | 0.59              |
|          | 412                             | 4.10                         |                                 |                                     |                   |
| DCM      | 346                             | 3.92                         | 456                             | 2342 (44)                           | 0.33              |
|          | 412                             | 3.96                         |                                 |                                     |                   |
| Dioxane  | 343                             | 4.01                         | 456                             | 2520 (47)                           | 0.29              |
|          | 409                             | 4.06                         |                                 |                                     |                   |
| Pyridine | 349                             | 3.97                         | 465                             | 2475 (48)                           | 0.06              |
|          | 417                             | 4.04                         |                                 |                                     |                   |
| MeCN     | 343                             | 4.08                         | 468                             | 2845 (55)                           | 0.11              |
|          | 413                             | 4.09                         |                                 |                                     |                   |
| AcOEt    | 342                             | 4.02                         | 459                             | 2544 (48)                           | 0.14              |
|          | 411                             | 4.11                         |                                 |                                     |                   |
| AcOH     | 343                             | 4.05                         | 468                             | 2729 (53)                           | 0.04              |
|          | 415                             | 4.12                         |                                 |                                     |                   |
| EtOH     | 346                             | 4.09                         | 466                             | 2637 (51)                           | 0.05              |
|          | 415                             | 4.18                         |                                 |                                     |                   |

**Table 3** Spectral properties of the compounds **3** in benzene

| Compound  | $\lambda_{\text{abs;max}}$ , nm | $\log \epsilon_{\text{max}}$ | $\lambda_{\text{flu;max}}$ , nm | Stokes shift, $\text{cm}^{-1}$ (nm) | $\Phi_{\text{F}}$ |
|-----------|---------------------------------|------------------------------|---------------------------------|-------------------------------------|-------------------|
| <b>3a</b> | 399                             | 3.44                         | 453                             | 2988 (54)                           | 0.46              |
| <b>3b</b> | 399                             | 3.63                         | 450                             | 2840 (51)                           | 0.45              |
| <b>3c</b> | 403                             | 3.52                         | 453                             | 2739 (50)                           | 0.53              |
| <b>3d</b> | 410                             | 3.77                         | 457                             | 2508 (47)                           | 0.20              |
| <b>3e</b> | 339, 410                        | 3.73, 3.89                   | 449                             | 2119 (39)                           | 0.41              |
| <b>3f</b> | 344, 412                        | 4.04, 4.10                   | 449                             | 2000 (37)                           | 0.59              |
| <b>3g</b> | 416                             | 3.93                         | 460                             | 2299 (44)                           | 0.11              |
| <b>3h</b> | 319, 416                        | 4.12, 4.07                   | 457                             | 2157 (41)                           | 0.35              |

**Fig. 2** Absorption (solid) and emission (dotted) spectra of compounds **3** in benzene, excitation wavelength is 375 nm.

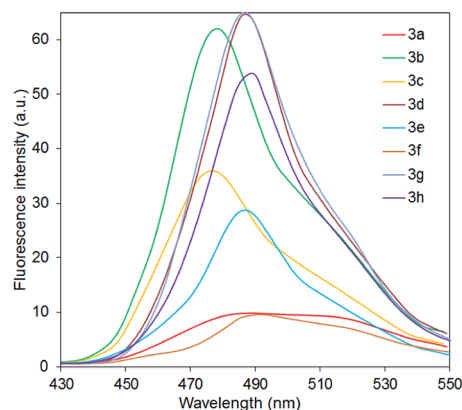
standard (7-hydroxy-4-methylcoumarine in 0.1 M phosphate buffer with pH = 10,  $\Phi_{\text{F}}$  0.63, excitation wavelength is 375 nm). Compounds **3d** and **3g** possess the lowest emission intensity probably due to the absence of the donating moieties in the aryl

substituent. The full characteristics of the luminescence are given in the Table 3.

### Solid-state fluorescence of synthesized compounds

Microcrystalline compounds **3a-f** are bright green or yellow-green. Upon illumination by an UV lamp, compounds **3b-e,g,h** are brightly emissive in the blue-green region of the spectrum while compounds **3a,f** are less emissive. The solid-state fluorescence spectra were recorded at room temperature and showed a main band around 480 nm and a shoulder around 510 nm. The emission bands appeared in a narrow region (478–491 nm) (see Fig. 3, Table 4). Nevertheless, a substituent in the aryl moiety possesses a significant influence on the emission intensity. Methoxy group insertion led to the decreasing of solid-state fluorescence intensity for compound **3h** by about 20% in compare to unsubstituted compound **3g**. In compare to unsubstituted compound **3d** methoxy derivative **3e** has a 55% less intensive emission, and dimethoxy derivative **3f** has an 85% less one. It should be noted that an observed regularity is opposite to the fluorescence in solution.

Also, an interesting phenomenon is the wide region of the excitation of solid-state fluorescence which is located in area of 320–470 nm. The highest emission intensity is observed for compounds **3** upon excitation by irradiation of the blue region

**Fig. 3** Solid-state emission spectra of compounds **3**, excited by 351 nm.**Table 4** Photophysical parameters of the crystalline compounds **3**

| <b>3</b>  | $\lambda_{\text{ex; max}}$ , nm | $\lambda_{\text{em; max}}$ , nm | R.I. <sup>a</sup> | HOMO <sup>b</sup> , eV | LUMO <sup>b</sup> , eV | GAP <sup>b</sup> , eV |
|-----------|---------------------------------|---------------------------------|-------------------|------------------------|------------------------|-----------------------|
| <b>3a</b> | 444                             | 489                             | 0.18              | −6.2                   | −2.29                  | 3.91                  |
| <b>3b</b> | 446                             | 478                             | 1.16              | −6.15                  | −2.24                  | 3.91                  |
| <b>3c</b> | 451                             | 481                             | 0.71              | −6.14                  | −2.29                  | 3.85                  |
| <b>3d</b> | 455                             | 486                             | 1.21              | −6.36                  | −2.57                  | 3.79                  |
| <b>3e</b> | 450                             | 486                             | 0.54              | −6.18                  | −2.42                  | 3.76                  |
| <b>3f</b> | 460                             | 491                             | 0.18              | −6.03                  | −2.41                  | 3.62                  |
| <b>3g</b> | 453                             | 487                             | 1.23              | −6.20                  | −2.44                  | 3.76                  |
| <b>3h</b> | 459                             | 488                             | 1.00              | −6.09                  | −2.36                  | 3.73                  |

<sup>a</sup> Excited by 351 nm, R.I. – relative intensity of emission is given in compare to **3h**. <sup>b</sup> Calculated in the basis TDDFT(B3LYP)/6-31+G(d,p).



of the spectrum with wavelengths of 450–460 nm, which is clearly shown in Fig. 4.

The quantum-chemical calculations in the basis TDDFT(B3LYP)/6-31+G(d,p) (see Fig. 5) showed that electron density of the highest occupied molecular orbital (HOMO) is localized on the diethylamino moiety and the lowest unoccupied molecular orbital (LUMO) involved cyano group. It indirectly confirms that absorption of light could lead to the intramolecular charge transfer from the donor moiety to acceptor. The energy levels of Frontier orbitals and gap values are also given in Table 4.

Continuing our earlier studies,<sup>15b</sup> it is important to note that the fluorescence intensity increases significantly, and its band shifts to the long-wavelength region upon transformation of 2-chloropyridines **2** to 2-diethylaminopyridine derivatives **3** (Fig. 6). It demonstrates the promise of this work, as well as the subsequent studies of spectral properties of other 3,4-dicyano-substituted aminopyridines.

### Crystal structures characterization

The obtained single crystal X-ray crystallographic data of the synthesized compounds can be used to gain a better understanding of the relationship between the optical properties and the molecular conformation and packing mode. Single crystals of **3c**, **3d**, **3f** and **3g** suitable for X-ray structural analysis were prepared by slow evaporation of acetonitrile solution, at room temperature (Fig. 7).

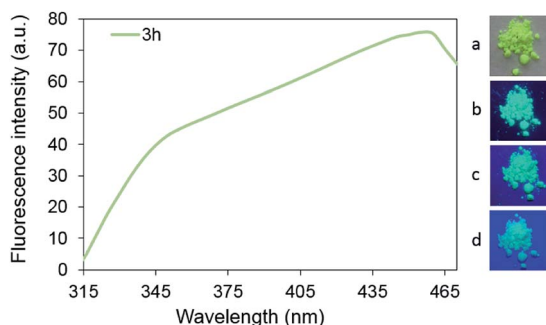


Fig. 4 Solid-state excitation spectrum of compound **3h** and photos under daylight (a), UV (b: 365 nm) and visible (c: 400 nm, d: 455 nm) irradiation.

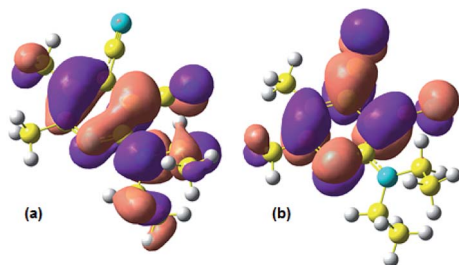


Fig. 5 (a) HOMO and (b) LUMO for compound **3a**, computed with TDDFT(B3LYP)/6-31+G(d,p). The pink (purple) lobes indicate a positive (negative) isocontour value.

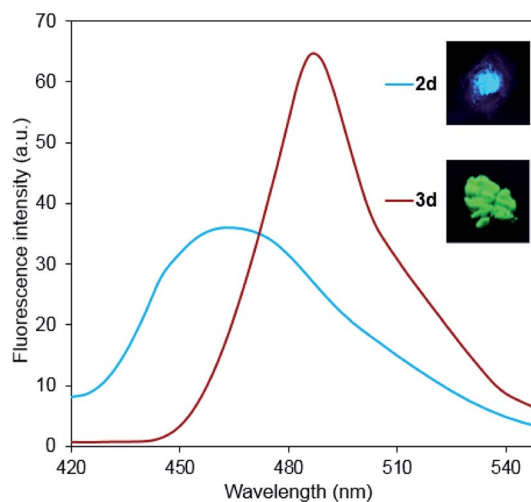


Fig. 6 Solid-state emission spectra of compounds **2d**, **3d** and photos under UV (365 nm) irradiation.

The pyridine ring of the **3g** (Fig. 7d) is overloaded with substituents, which leads to a substantial deformation of the heterocycle. The deviation of the atoms C3, C2, N1 from the plane through C4...C5...C6 – are 0.05, –0.20, –0.12 Å, respectively. The angle along the C6...C3...N3 line of pyridine cycle and cyano group at the third position is 171.1, while for another cyano group at the fourth position, the angle along the N1...C4...N4 line is 177.8. The heterocycle forms a dihedral angle –130.5 with a phenyl ring. Despite the significant steric interactions, the amino group (nitrogen atom N2) actively interacts with the  $\pi$ -heterocycle system. The bond length C2–N2 (1.351 Å) coincides with the value of the delocalized bond N1 C2 (1.354 Å). The nitrogen atom N2 is located outside (0.1 Å) of plane running through atoms C2, C17, C18. This means, that the nitrogen atom, despite its conjugation with the  $\pi$ -system, retains, to some extent, the tetrahedral ( $sp^3$ ) configuration.

Interestingly, there are two independent molecules in the structure of **3f** (Fig. 7c). The two molecules possess the different conformations with the different torsion angles. The greatest difference is observed in twisted conformation in which the torsion angles between pyridine ring and ethyl moieties. Torsion angles C3...C2...N2...C17 and C3...C2...N2...C18 are –34.7 and 167.5 respectively, while angles C3A...C2A...N2A...C17A and C3A...C2A...N2A...C17A are –25.9 and 171.6.

Pyridine ring in the molecule **3f** as well as in the compound **3g** is significantly deformed. The deviations from the plane running through the atoms C4...C5...C6 (C4A...C5A...C6A) have C2 –0.12 and C2A –0.07 Å. The angle C6...C3...N3 (C6A...C3A...N3) of the line along the pyridine cycle and the cyano group at the third position is 171.7 (173.7), while for the cyano group at the fourth position the angle N1...C4...N4 (N1A...C4A...N4A) is 178.9 (177.6).

Amino groups of N2, N2A are conjugated to pyridine rings with parameters close to **3g** N2...C2 is 1.356 Å and N2A...C2A is 1.354 Å. In the crystal, molecules form dimers, as a result of the  $\pi$ – $\pi$  interaction. The shortest distance C12...C6A is 3.364 Å.



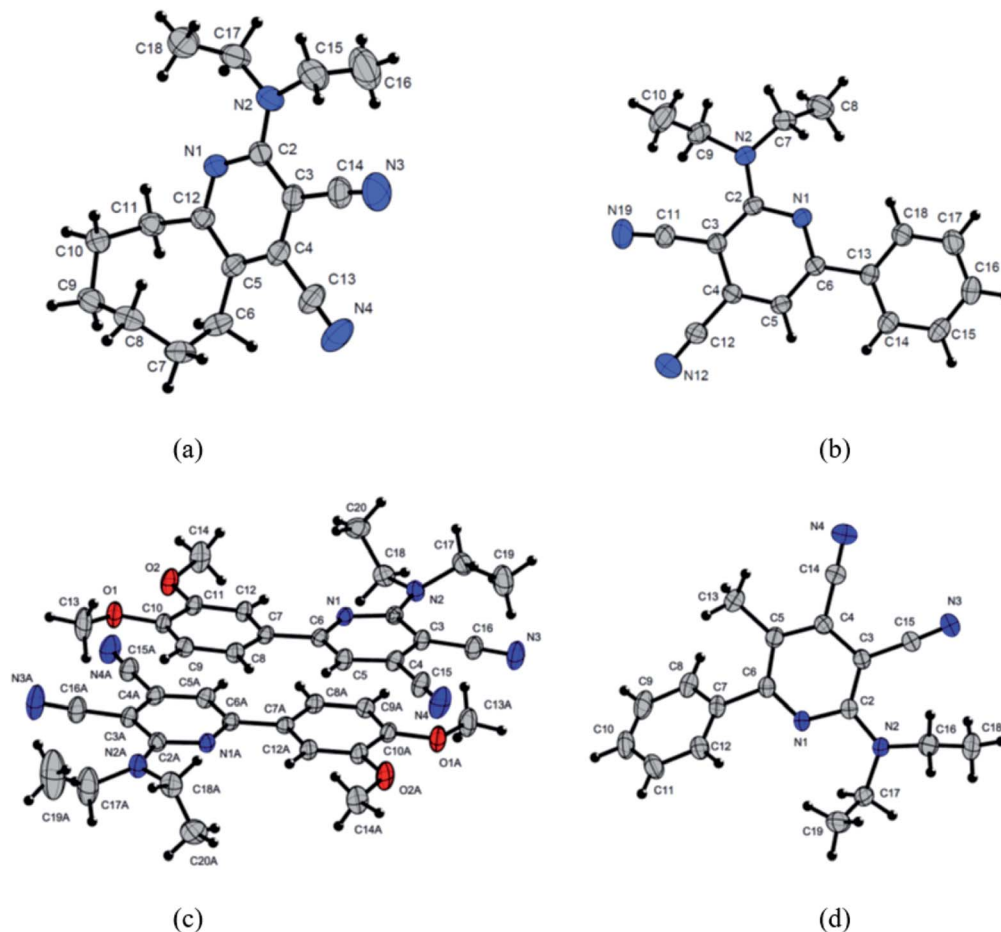


Fig. 7 Single-crystal structures of **3c** (a), **3d** (b), **3f** (c) and **3g** (d).

Phenyl ring in structure of **3f** lie in the same plane with the pyridine rings in contrast to the molecule **3g**. This difference can be explained by the influence of the methyl group at the fifth position of the pyridine ring, which creates steric hindrances for the planar arrangement of the rings in the case of **3g**. In the crystal of compound **3f**, wherein the hydrogen atom in the fifth position, the repulsion between the atoms H5...H8 (H5...H8a) is compensated by the interaction of the lone electron pair of the nitrogen atom, N1 (N1A), and the hydrogen atoms, H12 (H12A), respectively, for each of the independent molecules.

The structure of the molecule **3d** (Fig. 7b) in the crystal is similar to the structure of the **3f** molecule. Small differences are observed in the packing of molecules in a crystal. Molecules form stacks in which the values of the distances between neighboring molecules are substantially similar, but greater than 3.5 Å, that is, more than the sum of van der Waals radii of carbon-carbon.

Pyridine cycles in molecule **3d** are planar, in contrast to molecules **3f** and **3g**. Nevertheless, the cyano group in the third position deviates from the plane of the pyridine ring, and the angle along the line C6...C3...N19 is 173.1, while the cyano group in the fourth position does not deviate, and the angle N1...C4...N12 is 178.3.

Conformational features of the 8-membered cycle of the molecule **3c** can be seen in Fig. 7a. As in the structures presented earlier, the amino group (ethyl moieties) prevents from the  $\pi$ - $\pi$  interacting adjacent molecules in the stack. The angle C12...C3...N3 along the pyridine ring and the cyano group at the third position is 173.0, while with the cyano group at the fourth position the angle N1...C4...N4 is to 179.4.

To summarize the obtained data for the compounds studied, the deviation of angles for cyano groups in the third position of pyridine cycle for all the crystals obtained is observed in crystalline form and is in the range of 171–173°. It is probably the result of the influence of the adjacent diethylamine fragment, but on the other hand it also can indicate a significant contribution of the intermolecular charge-transfer (ICT) structure. The latter is supported by the fact that exactly the acceptor cyano group in the third position is in effective conjugation with the donor diethylamine fragment.

Moreover, the obtained data for crystal **3f** shows a difference in the torsion angles for the diethylamino group in two independent molecules of the dimer (Fig. 7c). A different conjugation degree of the amino group with the  $\pi$ -heterocycle system is the consequence of the above. This is probably the reason for the appearance of a “shoulder” in the emission spectra in the solid state and in solutions for all compounds. It is especially



noticeable for the solid state emission of compounds **3a** and **3f** (Fig. 3).

## Conclusions

In conclusion, we have developed the facile method for the synthesis of 2-diethylaminocinchomeric dinitrile derivatives and have studied fluorescence properties thereof in the solutions and in the solid state. High fluorescence quantum yield was observed for the solutions of pyridine derivatives with electron donating groups in nonpolar solvents. Moreover, an intensive solid-state emission in the blue-green region of the spectrum was observed for alkyl substituted 2-diethylaminocinchomeric dinitriles, as well as compounds with unsubstituted aryl moieties. This work is a continuation of the series of papers aimed to the study of the correlation of the structural framing of cyano-substituted pyridines and their derivatives with their photoluminescent properties.<sup>15</sup>

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