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Construction of SnO₂/graphene-like g-C₃N₄ with enhanced visible light photocatalytic activity†

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In this work, SnO₂/graphene-like g-C₃N₄ (SnO₂/GL-C₃N₄) composite photocatalysts with large surface areas and abundant coupling heterointerfaces were synthesized using a hydrothermal method. The as-prepared photocatalysts exhibited distinctly manifested efficient visible light activities toward organic pollutant degradation, demonstrating remarkable synergistic effects between SnO₂ and graphene-like g-C₃N₄ (GL-C₃N₄). The 25 wt% SnO₂/GL-C₃N₄ composite showed optimal photocatalytic activity under visible light irradiation, which was almost 9 and 2.5 times as high as that of SnO₂ and GL-C₃N₄, respectively. In addition, the possible photocatalytic mechanism of rhodamine B (RhB) degradation by SnO₂/GL-C₃N₄ under visible light was also discussed in detail. Moreover, this work would provide a facile way for the fabrication of novel two dimension/two dimension (2D/2D) GL-C₃N₄-based photocatalysts with high and stable performance for pollutant degradation.

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1. Introduction

Currently, with the rapid development of global industrialization, industrial wastewater treatment has become a big problem. Many measures have been used to alleviate the environmental issues. Photocatalysis with clean and economical energy has revealed wide prospects in application for the treatment of organic wastewater. Therefore, using semiconductor-based photocatalysts for the degradation of organic pollutants is considered as a sustainable and prospective solution for wastewater purification.¹

Graphitic carbon nitride (g-C₃N₄) is a promising heterogeneous metal-free, non-toxic semiconductor with narrow band gap, good thermal stability, high chemical stability and interesting electronic properties,^{2,3} has been widely researched for catalysis-driven application,^{4,5} electrochemical application,⁶ bio-chemical application,⁷ biosensor application⁸ and so on. And yet, it is found that the layered stacking structure and small specific surface area of g-C₃N₄ hinder the effective separation of photoinduced electron-hole pairs, thus made g-C₃N₄ suffering from rapid recombination of photogenerated electron-hole pairs. Therefore, researchers have made great efforts to solve these problems. It is well known that two dimension (2D) materials have advantages: (1) atomic thickness which could shorten the routine that photogenerated electrons and holes

had to pass in order to participate in the photocatalytic reaction quickly, (2) large surface areas that could expose more active sites.^{9–12} Indeed, it has been reported that g-C₃N₄ could be easily exfoliated into 2D graphene-like g-C₃N₄ nanosheets (GL-C₃N₄)¹³ with larger surface area and high photocatalytic activity.^{14,15} The obtained GL-C₃N₄ had a typical 2D structure with thickness of 2–3 nm and band gap of 2.85 eV, exhibited excellent performance in catalysis.¹⁰ However, the visible light photocatalytic activity of GL-C₃N₄ doesn't meet expectations,¹⁶ therefore, continue to decrease the high recombination rate of photo-generated electrons and holes of GL-C₃N₄ is necessary. Lately, 2D/2D materials have attracted extensive attention due to (1) enlarged contact area in interface region that could accelerate the charge carrier separation and inhibit its recombination,^{17,18} (2) forming a low Schottky barrier height junction which could reduce contact resistance and facilitate carrier transport.¹⁹ It is obvious that combining GL-C₃N₄ with other 2D materials forming 2D/2D structures so as to obtain more surface area and high coupling heterointerfaces extent, that could effectively facilitate separation/transfer of photoinduced electron-hole pairs, and thus improve the catalytic performance of GL-C₃N₄.¹⁸ Several studies have proven that combining g-C₃N₄ nanosheet with other 2D materials to form 2D/2D structures can indeed efficiently suppressed the recombination of photoinduced electron-hole pairs, endues the composites with enhanced properties as a result of the synergistic effects see in carbonized poly(furfural alcohol)/g-C₃N₄,¹⁷ rGO/g-C₃N₄,²⁰ SnNb₂O₆/g-C₃N₄.¹⁸

SnO₂, a n-type semiconductor, has become one of the hot spots in the field of photocatalysis research attributed to its high photosensitivity, good transparency, environmental

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friendliness and low cost.²¹ SnO₂ based semiconductors such as Fe₂O₃/SnO₂,²² CuO–SnO₂,²³ SnO₂/TiO₂,²⁴ SnO/SnO₂,²⁵ SnO₂/Fe₂O₃,²⁶ SnO₂/ZnS²¹ showed stable and outstanding performances in photocatalysis. Especially, 2D SnO₂ nanosheets can provide many promising properties for improving photocatalytic performance, which to a large extent benefit from their confined thickness and large specific surface. The existing literature proved that compounding SnO₂ nanosheets with other suitable materials forming composites with intimate connection and continuity in lattice, like SnO₂ nanosheets–CdSe nanocrystallites,²⁷ porous ZnO–SnO₂ nanosheets,²⁸ could help for facilitating the separations of photoinduced electron–hole pairs more effectively. However, there are few reports coupling GL–C₃N₄ with SnO₂ nanosheet together as to construct 2D/2D SnO₂ nanosheets/graphene-like g–C₃N₄ (SnO₂/GL–C₃N₄) in order to enhance photocatalytic performance for both GL–C₃N₄ and SnO₂ in the visible light region.

In this work, we report novel hybrid architecture of SnO₂/GL–C₃N₄ for photocatalytic environmental application. The SnO₂/GL–C₃N₄ photocatalysts were synthesized by an environmentally friendly hydrothermal method *via* the *in situ* growth of GL–C₃N₄ in the presence of SnO₂ nanosheets. The visible-light-driven photocatalytic activities of SnO₂/GL–C₃N₄ composites were evaluated by the degradation of RhB. It has been found that the optimal photocatalytic activity over 25 wt% SnO₂/GL–C₃N₄ was almost 9 times as high as that of SnO₂ precursor, and 2.5 times higher than that of GL–C₃N₄ precursor, respectively. The crystal structure, chemical state of elements, chemical composition, structural information, microstructure and grain morphology of the synthesized photocatalysts were analyzed by X-ray diffraction (XRD), X-ray photoelectron spectra (XPS), Fourier transform infrared spectra (FT-IR), scanning electron microscopy (SEM) and transmission electron microscopy (TEM), respectively. In addition, photoluminescence (PL) spectra, transient photocurrent response and electrochemical impedance spectroscopy (EIS) analysis also demonstrated that the introduction of SnO₂ could decrease the high recombination rate of photogenerated electron–hole pairs, which would consequently improve the visible-light photo-catalytic activity of GL–C₃N₄.

2. Experimental section

2.1. Materials

Stannous chloride (SnCl₂·2H₂O, A.R. grade), ammonia solution (NH₃·H₂O, A.R. grade), dicyandiamide (C₂H₄N₄, C.P. grade), ethanol (C₂H₅OH, A.R. grade), ammonium chloride (NH₄Cl, A.R. grade), all the starting materials were used directly without purification.

2.2. Material synthesis

SnO₂ nanosheets,²⁹ g–C₃N₄ and GL–C₃N₄ (ref. 13) were synthesized according to the reported procedure. The SnO₂/GL–C₃N₄ photocatalysts were synthesized by hydrothermal method. In this process, 0.015 g SnO₂, 0.1 g g–C₃N₄ and 0.5 g NH₄Cl were added into 40 mL solvent composed of 10 mL alcohol and 30

mL deionized water under vigorous stirring and ultrasonication for 1 h. Afterwards, the mixture was transferred into 50 mL Teflon-lined autoclave, and putted into a 180 °C drying oven for 12 h. At last, the photocatalyst was collected by centrifugation, washing with ethanol and distilled water repeatedly, and frozen drying for 12 h. The obtained product was 15 wt% SnO₂/GL–C₃N₄. Different weight percent (15 wt%, 25 wt%, 35 wt%) of SnO₂/GL–C₃N₄ photocatalysts were synthesized with the similar procedure.

2.3. Characterization

The compositions and crystal structure of the as-prepared samples were analyzed by XRD obtained on a Bruker D8 diffractometer using Cu K α radiation ($\lambda = 1.5418 \text{ \AA}$) in 2θ range of 10–80° with the scanning rate of 7° min^{−1}. The nitrogen adsorption–desorption isotherms were recorded on a TriStar II 3020 surface area and porosity analyzer. XPS was used to further determine the composition and electronic state of the SnO₂ nanosheets, GL–C₃N₄ and SnO₂/GL–C₃N₄ using an ESCA Lab MKII X-ray photo-electron spectrometer using Mg K α radiation. FT-IR was performed on a FT-IR Nexus 470 spectrometer, taking with KBr as standard material at room temperature to reveal the structural information of obtained samples. The compositions and morphologies of the samples were observed by SEM, TEM, high resolution transmission electron microscope (HRTEM), elemental mapping and energy-dispersive X-ray spectroscopy (EDS). And SEM, elemental mapping, EDS were taken with a JEOL-JSM-7001F while TEM, HRTEM were operated at 200 kV on a JEOL-JEM-2010 made in Japan. Ultraviolet-visible (UV-vis) diffuse reflection spectra (DRS) of the samples were recorded on an UV-2425 UV-vis spectrophotometer made in Shimadzu, Japan from the range of 200 nm to 800 nm, and BaSO₄ was employed as a reflectance standard. Photocurrents and EIS were performed at a CHI 660B electro-chemical workstation made in Chenhua Instrument Company (Shanghai, China). PL experiments were obtained on QuantaMaster & TimeMaster Spectrofluorometer using excitation wavelength at 387 nm and 275 nm, respectively. X-band equivalent series resistance (ESR) spectra were measured by a JES FA200 spectrometer at ambient temperature, operating at 9.5 GHz.

2.4. Photocatalytic experiments

Photocatalytic activities of the as-synthesized catalysts were assessed by photocatalytic degradation of RhB irradiated under the visible light in a custom-made photochemical reactor with air bubbling. 50 mg of the photocatalyst powder was directly dispersed in 50 mL of RhB aqueous solution (10 mg L^{−1}) with continuous stirring at 30 °C. At first, all the suspensions were continuously stirred for half an hour in dark condition, in order to establish an adsorption/desorption balance. Then open the spherical Xe lamp (300 W) with a cut-off light filter ($\lambda > 420 \text{ nm}$) which was used as visible light source. At every 30 min, about 4.0 mL solution suspension would be taken out and detected by a UV-vis spectrophotometer (Cary, 8584; NYSE:A) at 553 nm which is the characteristic maximum absorption



wavelength of RhB. The following formula (eqn (1)) was used to obtain the photocatalytic degradation efficiency (E) of RhB:

$$E = \left(1 - \frac{C}{C_0}\right) \times 100\% = \left(1 - \frac{A}{A_0}\right) \times 100\% \quad (1)$$

where C_0 , A_0 is the concentration of solution supernatant with an adsorption/desorption equilibrium at the initial reaction time, and the corresponding values of absorbance, respectively; C , A is the concentration of solution supernatant, and the corresponding values of absorbance at time t , respectively. In this work, we use the values of absorbance between 400–600 nm to obtain the photocatalytic degradation efficiency of RhB.

2.5. Photoelectrochemical measurements

Photocurrents were measured in a standard three-electrode experimental system, employing a Pt wire as the counter electrode and the sample modified indium-tin oxide glass (ITO) as the working electrode while a saturated Ag/AgCl electrode as the reference electrode. Phosphate buffered saline (PBS, pH = 7.0) was used as the supporting electrolyte solution of photocurrent measurements and the light source was a 500 W Xe arc lamp. GL-C₃N₄, SnO₂ and SnO₂/GL-C₃N₄ electrodes were prepared with the procedure: At first, dispersing 5.0 mg sample powder in 1.0 mL ethylene glycol by ultrasound. Afterward, 20 μ L as-prepared dispersion (5 mg mL⁻¹) was one-off drop-cast onto a piece of ITO (0.5 \times 1 cm²) with the infrared lamp irradiation, as to form the sample modified ITO electrode. EIS of different electrodes were measured in 0.1 M KCl solution containing 5 mM K₄[Fe(CN)₆]/K₃[Fe(CN)₆] under sunless conditions. And every Nyquist plot was recorded in frequency range of 100 MHz to 100 kHz.

3. Results and discussion

3.1. XRD, FT-IR analyses

The XRD patterns of the GL-C₃N₄, SnO₂ and SnO₂/GL-C₃N₄ samples with different weight percent of SnO₂ were shown in Fig. 1(a). It could be seen that the peaks of SnO₂/GL-C₃N₄ were distinctly similar to that of SnO₂, indicating that SnO₂ has no obvious change during the material preparation. In addition, the SnO₂/GL-C₃N₄ composites showed the sole SnO₂ peaks without the GL-C₃N₄ peak owing to the reason that, the peak of

(002) crystal plane of GL-C₃N₄ maybe overlapped with the peak of (110) crystal plane of SnO₂ since the two were very close. The pure SnO₂ exhibited several peaks at $2\theta = 26.7^\circ$, 33.7° , 38.1° , 51.8° , 54.9° , 61.7° , 65.4° , 71.1° and 78.4° corresponding to (110), (101), (200), (211), (220), (310), (301), (202) and (321) plane of a tetragonal rutile-like SnO₂.^{29,30} The strong and sharp peak at $2\theta = 27.8^\circ$ could be corresponded to the (002) diffraction plane of GL-C₃N₄, this is consistent with the reported literature.¹ No impurity peaks in GL-C₃N₄, SnO₂ and SnO₂/GL-C₃N₄, indicating that those photocatalysts were successfully prepared in this work.

The FT-IR spectrum enables us to distinguish the molecular structure of SnO₂, GL-C₃N₄ and SnO₂/GL-C₃N₄ composites (Fig. 1(b)). With regard to the SnO₂/GL-C₃N₄ composites, both GL-C₃N₄ and SnO₂ could be detected confirming the SnO₂/GL-C₃N₄ composites were successful synthesized. The peaks at 807, 1100–1650, 2400 cm⁻¹ in the GL-C₃N₄ spectrum corresponded to the unique breathing mode of triazine units,^{31–33} characteristic stretching modes of the C=N and C–N heterocycles,³¹ adsorption of CO₂ in the environment, respectively. And the broad band at 3000–3600 cm⁻¹ was attributed to the residual N–H group.³⁴ SnO₂ showed only a broad strong peak at 400–700 cm⁻¹, which could be ascribed to the anti-symmetric stretching vibration of Sn–O–Sn.^{25,35,36} Compared with pure SnO₂, the band at 597 cm⁻¹ was red-shifted in SnO₂/GL-C₃N₄ composites, suggesting that there were interactions between SnO₂ and GL-C₃N₄.

3.2. XPS analysis

XPS analyses were conducted to determine the composition and surface electronic state of SnO₂, GL-C₃N₄ and 25 wt% SnO₂/GL-C₃N₄ composite. As shown in Fig. 2(a), SnO₂ was composed of Sn, O elements and for the pure GL-C₃N₄, only C and N elements were detected, while the 25 wt% SnO₂/GL-C₃N₄ was composed of C, N, Sn and O elements. High-resolution of C 1s, N 1s, Sn 3d and O 1s spectra of the material were plotted in Fig. 2(b)–(e) in details. It could be observed in Fig. 2(b) that C 1s for GL-C₃N₄ had two peaks at 288.2 and 284.7 eV. The peak at 284.7 eV due to carbon contamination which could be observed on the XPS characterization for all samples in Fig. 2(b), while the C 1s at 288.2 eV was associated with sp²-bonded carbon (N–C=N) which had been shift to 288.4 eV in 25 wt% SnO₂/GL-C₃N₄.^{31,37,38} The shift of C 1s at 288.4 eV in 25 wt% SnO₂/GL-C₃N₄ could be

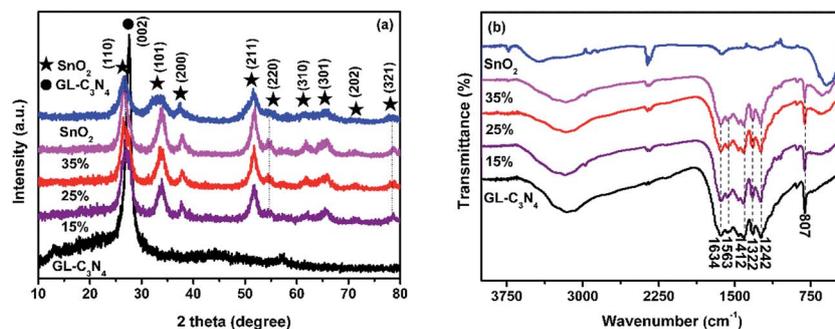


Fig. 1 (a) XRD patterns, (b) FT-IR spectra of SnO₂, GL-C₃N₄, SnO₂/GL-C₃N₄ composites.



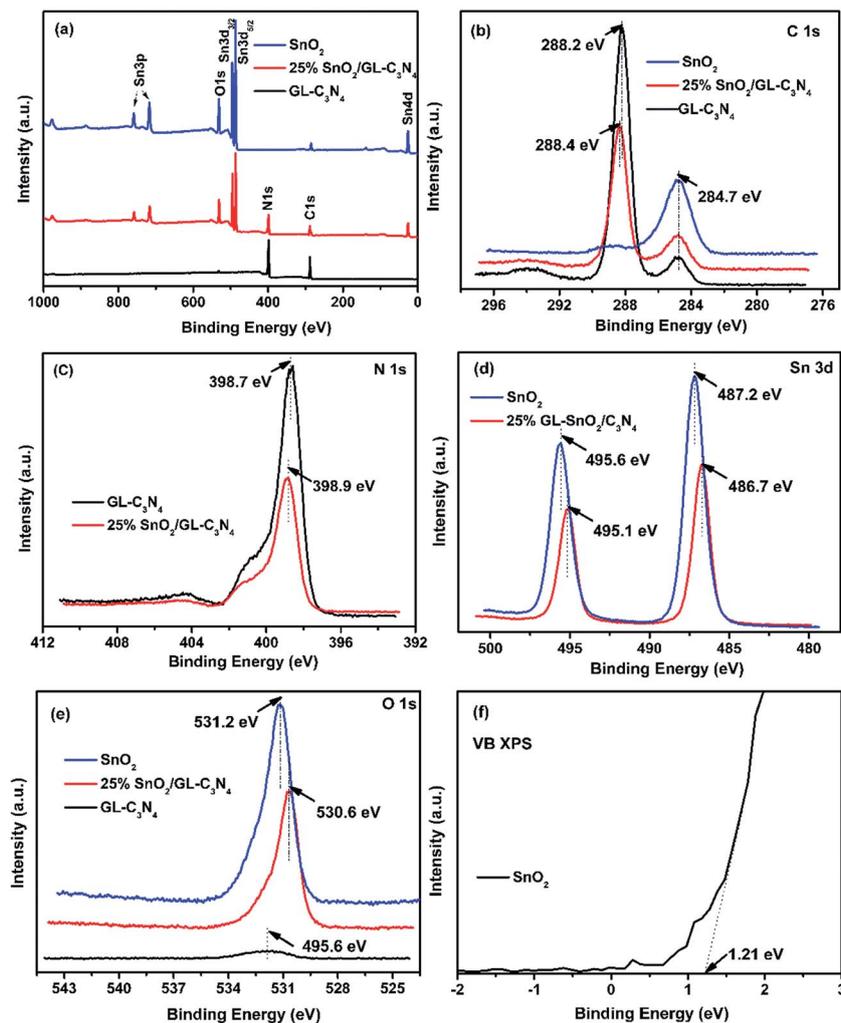


Fig. 2 XPS spectra of SnO_2 , $\text{GL-C}_3\text{N}_4$ and 25 wt% $\text{SnO}_2/\text{GL-C}_3\text{N}_4$ composite. (a) survey spectra, (b) C 1s, (c) N 1s, (d) Sn 3d, (e) O 1s, (f) VB XPS of SnO_2 .

attributed to the strong interactions between SnO_2 and $\text{GL-C}_3\text{N}_4$ which had influenced the chemical environments of N-C=N . Fig. 2(c) showed that the N 1s peak located at 398.7 eV in $\text{GL-C}_3\text{N}_4$ originated from C=N-C coordination³⁷ was shift to 398.9 eV in 25 wt% $\text{SnO}_2/\text{GL-C}_3\text{N}_4$ composite proving that there was existing an interaction between SnO_2 and $\text{GL-C}_3\text{N}_4$. The binding energy of Sn 3d peaks in SnO_2 located at 495.6 and 487.2 eV were ascribed to $\text{Sn } 3d_{5/2}$ and $\text{Sn } 3d_{3/2}$ of Sn^{4+} oxidation state, respectively³⁹ (Fig. 2(d)). Obviously, the corresponding Sn 3d peaks of 25 wt% $\text{SnO}_2/\text{GL-C}_3\text{N}_4$ were shift to 495.1 and 486.7 eV, respectively. The shift providing evidence of the interactions between SnO_2 and $\text{GL-C}_3\text{N}_4$. The O 1s peak in SnO_2 at 531.2 eV should be ascribed to Sn-O-Sn (lattice O),²¹ which changed to 530.6 eV in 25 wt% $\text{SnO}_2/\text{GL-C}_3\text{N}_4$ also demonstrating that there were some interactions between SnO_2 and $\text{GL-C}_3\text{N}_4$ (Fig. 2(e)). Fig. 2(f) showed the VB X-ray photoelectron spectra of SnO_2 , which was 1.21 eV. The valence band value of as-prepared SnO_2 was different from that of normal SnO_2 in previous works. The cause of this phenomenon was discussed in detail in DRS section. Together with FT-IR and XRD results

we knew that, $\text{GL-C}_3\text{N}_4$, SnO_2 and $\text{SnO}_2/\text{GL-C}_3\text{N}_4$ were successfully prepared in this work, and there were interactions between SnO_2 and $\text{GL-C}_3\text{N}_4$ in $\text{SnO}_2/\text{GL-C}_3\text{N}_4$ composites.

3.3. SEM and TEM analyses

Fig. 3 displayed typical SEM and TEM images of SnO_2 , $\text{GL-C}_3\text{N}_4$, 25 wt% $\text{SnO}_2/\text{GL-C}_3\text{N}_4$ composite as well as the elemental mapping, EDS of 25 wt% $\text{SnO}_2/\text{GL-C}_3\text{N}_4$ composite, HRTEM images of pure SnO_2 and 25 wt% $\text{SnO}_2/\text{GL-C}_3\text{N}_4$. As shown in Fig. 3(a) and (f), $\text{GL-C}_3\text{N}_4$ was a few layered structure similar to the structure of graphene, which was consistent with previous reports.³³ Fig. 3(b), (g) and (i) showed SEM image, TEM image, HRTEM image recorded from frame marked region in (g) of bulk SnO_2 , respectively. While it could be seen in Fig. 3(b) that, SnO_2 exhibited layer structure formed by the tiny pieces with size smaller than that of $\text{GL-C}_3\text{N}_4$. Fig. 3(g) and (i) showed that the synthesized SnO_2 lamellar nanosheets had irregular shape and size, which were unlike normal 2D SnO_2 (ref. 29 and 40) nanosheets (which had complete crystallization, thin and



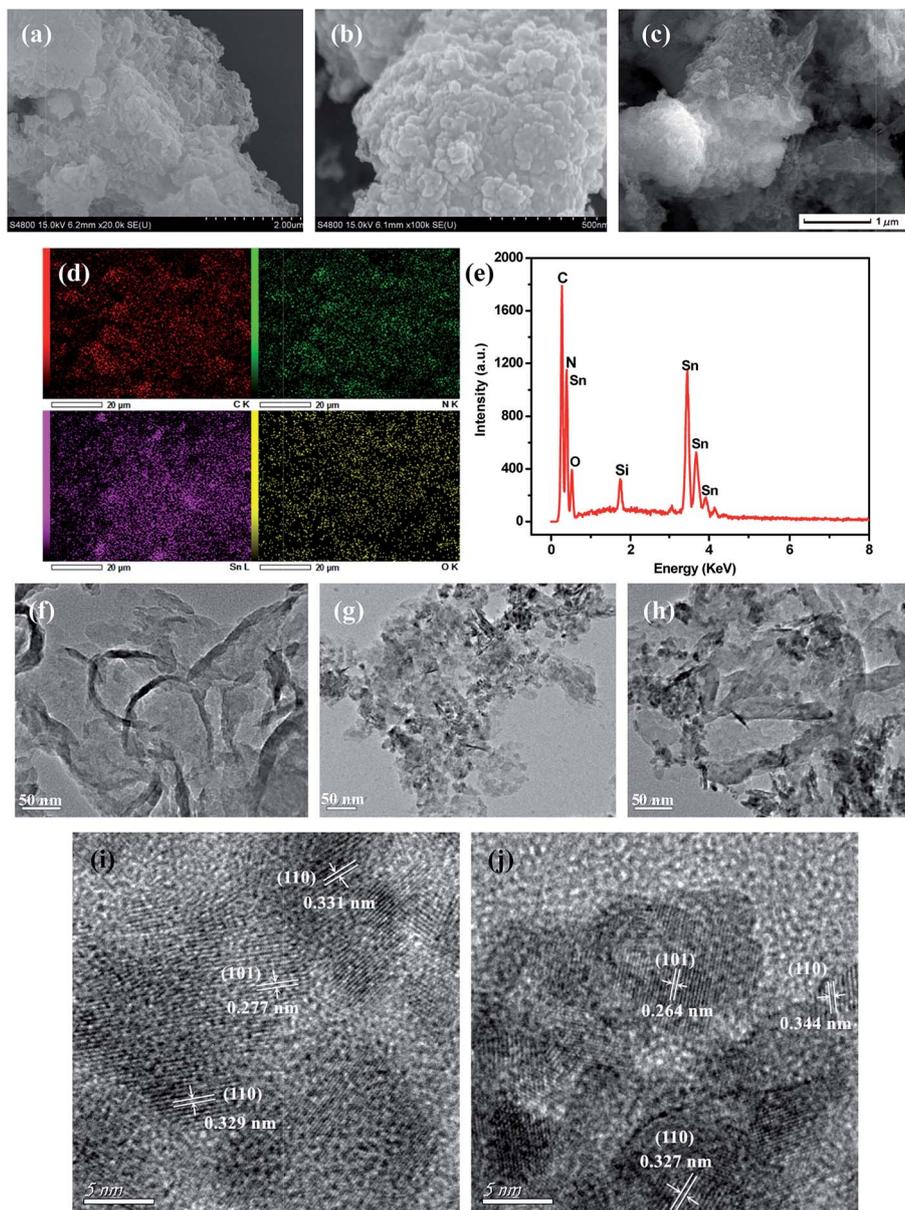


Fig. 3 SEM images: (a) GL-C₃N₄, (b) SnO₂, (c) 25 wt% SnO₂/GL-C₃N₄; (d) Elemental mapping of 25 wt% SnO₂/GL-C₃N₄; (e) EDS of 25 wt% SnO₂/GL-C₃N₄; TEM images: (f) GL-C₃N₄, (g) SnO₂, (h) 25 wt% SnO₂/GL-C₃N₄; HRTEM images of (i) pure SnO₂, (j) 25 wt% SnO₂/GL-C₃N₄.

structured morphology). The lattice spacing of SnO₂ nanosheet were estimated to be near 0.27 nm, 0.33 nm, and corresponding to the (101), (110) plane of SnO₂, respectively (Fig. 3(i)). So the irregular shape of the synthesized SnO₂ nanosheets may be caused by the incomplete crystallization which could due to the uneven cooling in Teflon-lined autoclave. Fig. 3(c), (h) and (j) were SEM image, TEM image, HRTEM image recorded from frame marked region in (h) of 25 wt% SnO₂/GL-C₃N₄ composite, respectively. As expected, the 25 wt% SnO₂/GL-C₃N₄ composite exhibited nanosheets structure and in which SnO₂ was wrapped around GL-C₃N₄, indicating that both SnO₂ and GL-C₃N₄ were exfoliated by this hydrothermal method. While the presence of SnO₂ may have influenced the exfoliating of g-C₃N₄ into 2D GL-C₃N₄, this caused GL-C₃N₄ to have loose structure and smaller

size.⁴¹ As shown in Fig. 3(j), GL-C₃N₄ displayed large area without crystal lattice and the interplanar spacing were estimated to be near 0.27 nm, 0.33 nm, and corresponding to the (101), (110) plane of SnO₂, verifying the existence of GL-C₃N₄ and SnO₂ nanosheets in 25 wt% SnO₂/GL-C₃N₄. Moreover, the interaction of SnO₂ and GL-C₃N₄ brought them a close integration, and the intimate contact between SnO₂ and GL-C₃N₄ would lead to a close interface between the two semiconductors. And this close interface allowed for the charge transfer between SnO₂ and GL-C₃N₄, promoting the separation of photoinduced electrons and holes, and increased the visible light photocatalytic activity of composites consequently. The elemental mapping and EDS (Fig. 3(d) and (e)) indicated that the 25 wt% SnO₂/GL-C₃N₄ composite containing C, N, Sn and O elements,



verifying the existence of GL-C₃N₄ and SnO₂ nanosheets in 25 wt% SnO₂/GL-C₃N₄. And the elemental mapping of elements C, N, Sn and O were uniformly distributed throughout the whole region of 25 wt% SnO₂/GL-C₃N₄ composite shown in Fig. 3(d). Those results gave reliable evidence that SnO₂ nanosheets were successfully combined with GL-C₃N₄, forming the typical 2D/2D structure which was presumed to play an important part in promoting the photo-induced electrons and holes transfer across the SnO₂ and GL-C₃N₄ contact interface, hence reducing the recombination of photoinduced electron-hole carriers.

3.4. N₂ adsorption-desorption and UV-vis analysis

N₂ adsorption-desorption isotherms in Fig. 4 were used for the analysis of BET specific surface areas of SnO₂/GL-C₃N₄ composites. It could be seen that the BET specific surface area of g-C₃N₄, GL-C₃N₄, SnO₂ and 25 wt% SnO₂/GL-C₃N₄ is 3.51 m² g⁻¹, 71.08 m² g⁻¹, 90.37 m² g⁻¹, 124.12 m² g⁻¹, respectively. The BET specific surface area of GL-C₃N₄ is much larger than that of g-C₃N₄, suggesting that the layered stacking g-C₃N₄ was exfoliated into 2D GL-C₃N₄ nanosheets.^{15,42} SnO₂ has the large BET

specific surface area, which could arise from its smaller lamellar structure. And automatically, the BET specific surface area of 25 wt% SnO₂/GL-C₃N₄ composite is much larger than that of pure GL-C₃N₄ and SnO₂, indicating the formation of 2D/2D structures which could effectively facilitate separation/transfer of photoinduced electron-hole pairs, and thus improve the catalytic performance.

Optical absorption property is an important parameter for the photocatalyst. UV-vis absorption spectra of SnO₂, GL-C₃N₄ and SnO₂/GL-C₃N₄ composites are shown in Fig. 5(a). It was clearly that the optical absorption edges of SnO₂ and GL-C₃N₄ were located approximately at 506 and 460 nm, respectively, which was estimated from ones absorption onset. The basal absorption edges of the SnO₂/GL-C₃N₄ composites were as similar with GL-C₃N₄, indicating that both Sn and O were not doped in the GL-C₃N₄ structure. Compared to the SnO₂ semicrystals,⁴³ SnO₂ nanoparticles,⁴⁴ SnO₂ nanotubes,⁴⁵ and other SnO₂ samples with the absorption edge of 305–355 nm originated from its wide band gap in previous work, the as-synthesized SnO₂ nanosheets in this paper showed strong visible light absorption. This result revealed that the as-prepared SnO₂ nanosheets had a narrow band gap and could be excited by visible light. DRS spectra are related to many factors, such as the molecular structure, color, grain size, the degree of dry of the catalyst, particularly evident influence of the color.⁴⁶ The absorption intensity in visible light region was enhanced with the increasing SnO₂ content in the SnO₂/GL-C₃N₄ composites which could be due to the brown color of the SnO₂. The optical absorption data and SEM/TEM results worked together implying that SnO₂ nanosheets were uniformly recombined with GL-C₃N₄ in this work.

The energy level and band gap of the semiconductors play crucial roles in determining these physical properties.¹⁸ The band gap energies (E_g) of SnO₂, GL-C₃N₄ and SnO₂/GL-C₃N₄ composites can be estimated respectively according to the following formula eqn (2).

$$(Ah\nu)^2 = h\nu - E_g \quad (2)$$

where A , h , ν and E_g are absorbance, Planck constant, light frequency and band gap energy, respectively. From Fig. 5(b), it could be seen that the band gap of the SnO₂, GL-C₃N₄ and 15

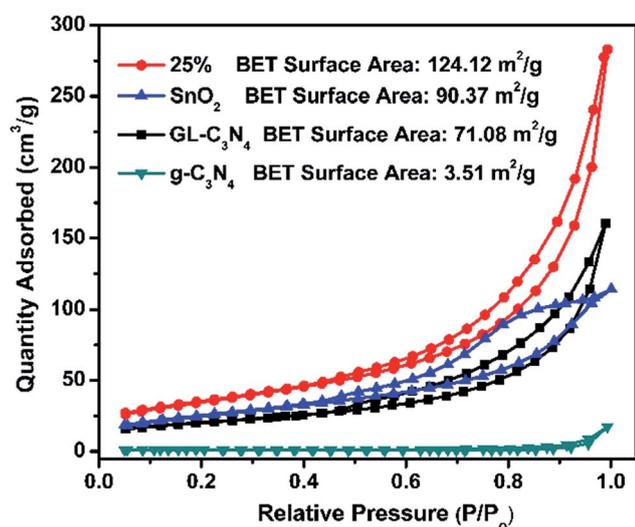


Fig. 4 N₂ adsorption-desorption isotherms of g-C₃N₄, GL-C₃N₄, SnO₂ and 25% SnO₂/GL-C₃N₄ composite.

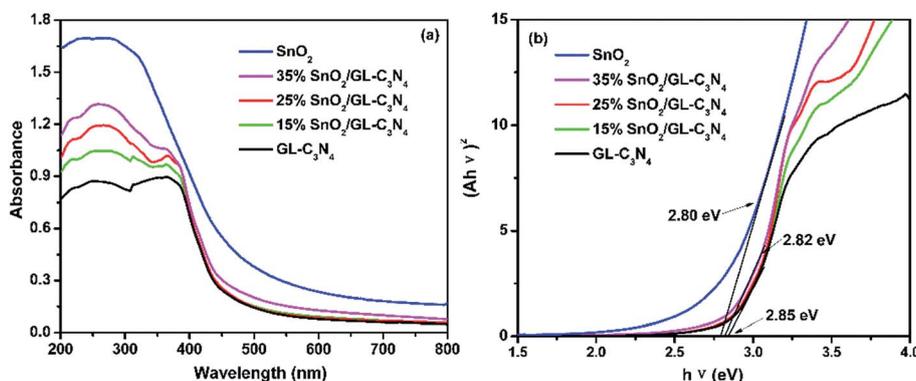


Fig. 5 (a) UV-vis diffuse reflectance spectra of the samples; (b) estimated band gap of SnO₂, GL-C₃N₄ and SnO₂/GL-C₃N₄ composites.



wt% SnO₂/GL-C₃N₄, 25 wt% SnO₂/GL-C₃N₄, 35 wt% SnO₂/GL-C₃N₄ were calculated to be ~2.80 eV, 2.85 eV, 2.85 eV and 2.82 eV, respectively. Compare to GL-C₃N₄, the 35 wt% SnO₂/GL-C₃N₄ composite (at high SnO₂ content) showed narrower band gap, which indicates that the existence of SnO₂ could significantly affect the energy gap. The experimental result for the E_g of GL-C₃N₄ was in agreement with the reported literature,¹³ but for SnO₂, the E_g was significantly smaller than the literature value of 3.8 eV.⁴⁷ The reason for this phenomenon could probably be attributed to the imperfect crystallization of SnO₂.⁴⁸

3.5. PL analysis

The PL spectra of GL-C₃N₄, 25 wt% SnO₂/GL-C₃N₄ in Fig. 6(a) were obtained by monitoring with an excitation wavelength of 387 nm; and SnO₂, 25 wt% SnO₂/GL-C₃N₄ in Fig. 6(b) were obtained by monitoring with an excitation wavelength of 275 nm. As shown in Fig. 6(a), the GL-C₃N₄ and 25 wt% SnO₂/GL-C₃N₄ composite presented strong PL emission at 469 nm which was corresponded to the band gap for the recombination of photogenerated electron-hole pairs.³⁴ Clearly, an apparent fluorescence quenching was observed for 25 wt% SnO₂/GL-C₃N₄ compared with pure GL-C₃N₄, indicating the suppressed recombination among photogenerated charge carriers in this photocatalyst.⁴⁹ SnO₂ exhibited higher emission intensity of PL spectra at 358 nm compared with 25 wt% SnO₂/GL-C₃N₄ with a commonly excitation wavelength of 275 nm.^{50,51} The decreased photoluminescence intensity of 25 wt% SnO₂/GL-C₃N₄

indicating that it had lower recombination rate of photoinduced electron-hole pairs than GL-C₃N₄ and SnO₂, providing evidence of a direct effect in charge recombination between GL-C₃N₄ and SnO₂.

3.6. Transient photocurrents and EIS analyses

As is well known that, semiconductor with unique band gap structure may have photoelectric properties which could convert can sun light energy into electricity. And the photoelectric conversion efficiency of the SnO₂/GL-C₃N₄ could be analyzed by the photocurrent intensity of the sample. In order to further confirm the composites playing an important role in the photocatalytic reaction, four switch source loop transient photocurrent response experiment of the samples were made. In Fig. 7(a), the 25 wt% SnO₂/GL-C₃N₄ composite showed an obvious enhancement of photocurrent with respect to the bulk sample under visible light irradiation, which was almost 10 and 4 times as high as that of SnO₂ and GL-C₃N₄, respectively, indicating an improvement of charge separation. Photocurrent results confirmed that the composites had a more efficient separation of photogenerated electrons and holes. It was also confirmed that SnO₂ and GL-C₃N₄ administered together effectively at the interface and there were interactions (which playing an important role in the effective separation of electron-hole pairs of GL-C₃N₄) between SnO₂ and GL-C₃N₄ which was also reflected in the XPS analyses and FT-IR spectra. And 25 wt% SnO₂/GL-C₃N₄ had the highest separation efficiency of the

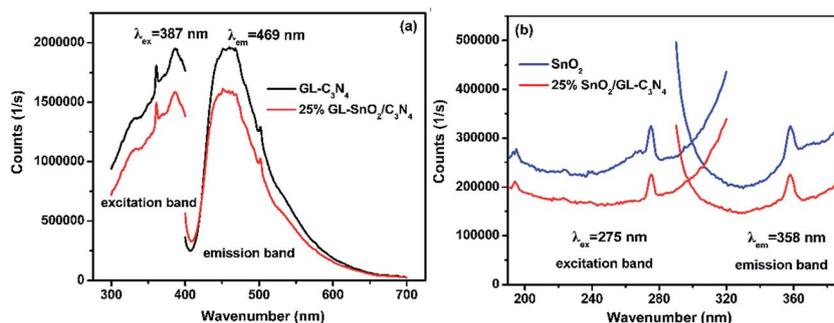


Fig. 6 PL spectra of (a) GL-C₃N₄, 25 wt% SnO₂/GL-C₃N₄ and (b) SnO₂, 25 wt% SnO₂/GL-C₃N₄.

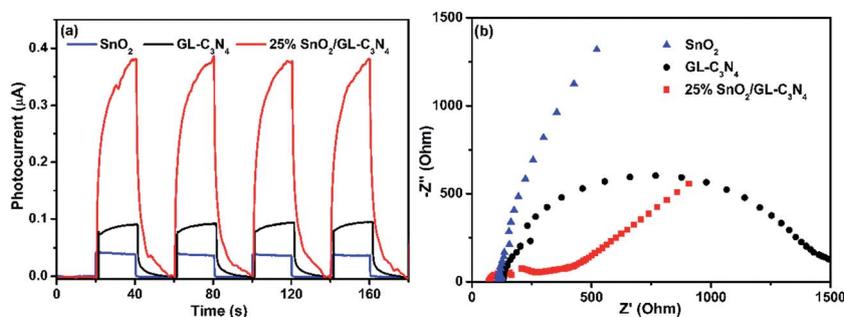


Fig. 7 (a) Transient photocurrent response for GL-C₃N₄, SnO₂ and 25 wt% SnO₂/GL-C₃N₄ composite in PBS (pH = 7.0) aqueous solution under visible light irradiation. (b) EIS of GL-C₃N₄, SnO₂ and 25 wt% SnO₂/GL-C₃N₄ composite in a 0.1 M KCl solution containing 5.0 mM Fe(CN)₆³⁻/Fe(CN)₆⁴⁻.



samples, which was in good consistent with the photocatalytic activities. The EIS measurements were conducted to gain deeper insight into the charge transport behavior of SnO₂/GL-C₃N₄ in Fig. 7(b). It has been proved that the radius of every circle in the Nyquist diagram is based on the charge transfer process at the interface of corresponding electrode/electrolyte, and the smallest radius corresponding with a lowest charge-transfer resistance.⁵² It could be seen from Fig. 7(b) that there was a decline in the radius when introduced SnO₂ into the system, indicating that the introduction of SnO₂ decreased charge-transfer resistance and raised the interfacial charge transfer. Therefore, the separation of photogenerated electron-holes could be improved by combining SnO₂ with GL-C₃N₄. 25 wt% SnO₂/GL-C₃N₄ showed the smallest radius demonstrating that it had a more effective separation of photoinduced electrons and holes. EIS results were in good consistent with the four switch source loop transient photocurrent response experiment and PL spectra of the composites.

3.7. Photocatalytic activity test

From the characterization we mentioned above, we knew that the composites showed enhanced adsorption capability, low recombination rate and high separation efficiency. Those as-prepared composites were supposed to be efficient visible-light driven photocatalysts. The photocatalytic activities of the samples were evaluated from RhB degradation under visible light irradiation. As shown in Fig. 8(a), GL-C₃N₄ and SnO₂ showed weak photocatalytic activities towards RhB degradation, of which the photocatalytic degradation efficiencies were 36% and 10%, respectively. While the degradation rate for SnO₂/GL-C₃N₄ composites (15 wt%, 25 wt%, 25 wt% physically blending, 35 wt%) under visible-light irradiation for 3 h were 88%, 92%, 78% and 71%, respectively (Fig. 8(a)). Obviously, an appropriate SnO₂ ratio in composite was critical significant for the improvement of the resultant photocatalytic performance. The increased photocatalytic activity of SnO₂/GL-C₃N₄ composites could be attributed to the enhanced separation efficiency of the photoexcited electron-hole pairs. For comparison, a mixture of SnO₂ and GL-C₃N₄ with an appropriate ratio of 25 wt% was also used as one of the photocatalysts. As can be seen from Fig. 8(a), the 25 wt% physically blending of SnO₂ and GL-C₃N₄ showed

much lower photocatalytic activity in RhB degradation than that of 25 wt% SnO₂/GL-C₃N₄ composite. The result also indicated that there was an interaction leading to a close interface between the two semiconductors in SnO₂/GL-C₃N₄ composites. And this interaction allowed for the charge transferring between SnO₂ and GL-C₃N₄, thus promoting the separation of photoexcited electron-hole pairs, leading to a high photocatalytic activity of composites. Fig. 8(b) was the absorption spectra of 25 wt% SnO₂/GL-C₃N₄ composite in photodegradation of RhB. Under visible light irradiation for 3 h, the color of solution suspension was changed from red gradually to colorless. Also the absorption peak near 553 nm decreased gradually, which suggested that RhB was oxidation degraded completely.

3.8. Cycling test

The stability of the synthesized 25 wt% SnO₂/GL-C₃N₄ composite was studied by a 4-run cycling test. The photocatalytic performance of 25 wt% SnO₂/GL-C₃N₄ composite was investigated by the RhB degradation under visible light irradiation for 3 h. It could be seen from Fig. 9 that the degradation efficiency of RhB over 25 wt% SnO₂/GL-C₃N₄ after 4 recycles was very high. The stability of 25 wt% SnO₂/GL-C₃N₄ composite was

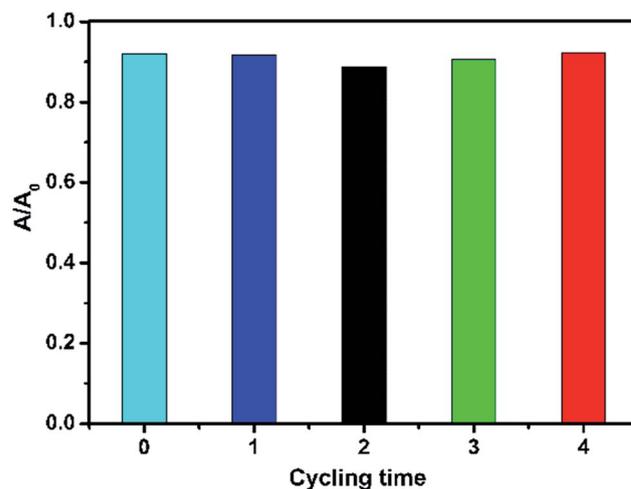


Fig. 9 The recycle test for RhB degradation over 25 wt% SnO₂/GL-C₃N₄ under visible-light irradiation.

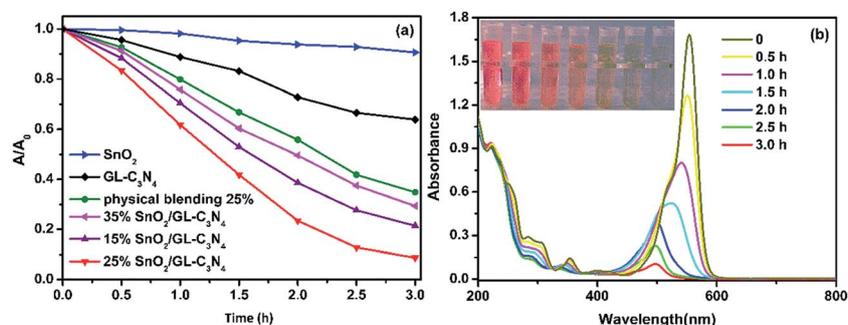


Fig. 8 (a) Photocatalytic activities of the samples for the degradation of RhB under visible-light irradiation; (b) absorption spectral changes of RhB under visible light irradiation.



also investigated by XRD (Fig. S1†) and FT-IR analyses (Fig. S2†). It could be seen from Fig. S1 and S2,† the recycled 25 wt% SnO₂/GL-C₃N₄ composite named “25 wt% after” had no obvious change in location compared with 25 wt% SnO₂/GL-C₃N₄ composite. These results made it clearly that 25 wt% SnO₂/GL-C₃N₄ remained effective and recyclable for RhB degradation.

3.9. ESR, trapping experiments and mechanism of RhB degradation under visible light

For purpose of investigate the active species produced in the photodegradation process under visible light irradiation, ESR spin-trap technique with DMPO (5,5-dimethyl-1-pyrroline-*N*-oxide) was used. As everyone knows, H₂O₂ was formed from photogenerated electron-hole pairs in the reactions with adsorbed water or oxygen, where $\cdot\text{OH}$ and O₂^{•-} were possibly involved.¹³ As can be shown in Fig. 10(a), SnO₂/GL-C₃N₄ treated without visible light irradiation exhibited no ESR signal. While under visible light irradiation, signals attributed to a DMPO-O₂^{•-} species were detected successfully in SnO₂/GL-C₃N₄ dispersions in methanolic media, indicating that O₂^{•-} reactive species were generated during the reaction. In addition, $\cdot\text{OH}$ signals were also detected in visible light irradiation during the reaction (Fig. 10(b)). The results confirmed that both O₂^{•-} and $\cdot\text{OH}$ were main active species.

The trapping experiments were also employed in Fig. 11, as to further study the mechanism. Usually, *t*-BuOH was employed as the part of hydroxyl radical scavenger, and 2Na-EDTA was used as holes scavenger, while N₂ was inlet in order to prevent O₂^{•-} forming.^{37,53} As shown in Fig. 11, the addition of N₂ significantly suppressed the degradation of RhB over 25 wt% SnO₂/GL-C₃N₄ photocatalyst, while *t*-BuOH and 2Na-EDTA had suppressed the RhB degradation in varying degrees. Those results confirmed that O₂^{•-} was main active specie during RhB photocatalytic degradation process, which was consist with the ESR result. While $\cdot\text{OH}$ and holes were also contributed part of the role in RhB photocatalytic degradation.

Fig. 12 describes the possible photocatalytic mechanism of RhB degradation over SnO₂/GL-C₃N₄ samples under visible light. The band gaps of SnO₂ and GL-C₃N₄ are 2.80 eV and 2.85 eV, respectively. Since the conduction band (CB) edge potential of SnO₂ is estimated as -1.59 eV and valence band

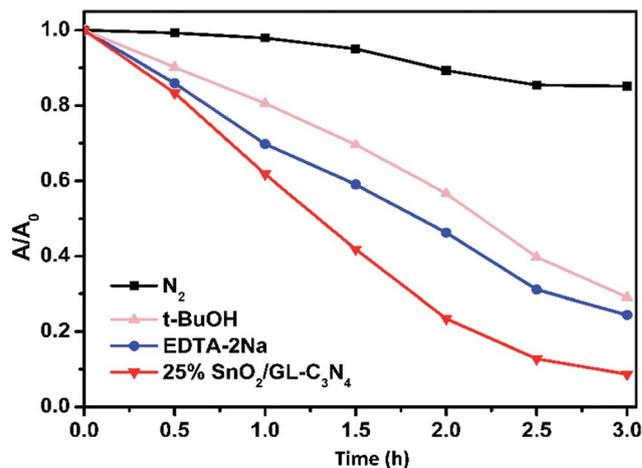


Fig. 11 Photodegradation of RhB over 25 wt% SnO₂/GL-C₃N₄ photocatalyst with different quenchers.

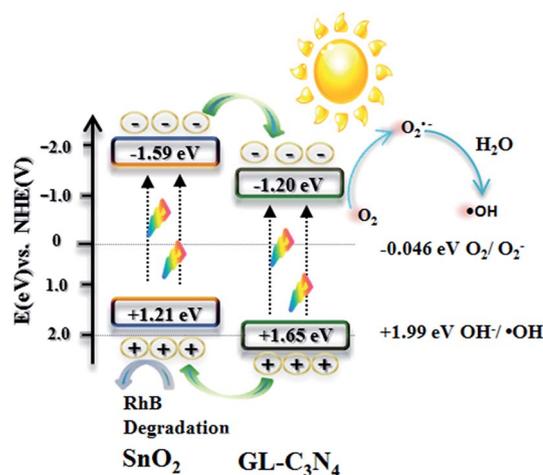


Fig. 12 Photocatalytic mechanism in the presence of SnO₂/GL-C₃N₄ under visible light.

(VB) is 1.21 eV. Compared with the common SnO₂ with wide band gap, the prepared SnO₂ can be excited and produces the photogenerated electron/hole pairs under visible light

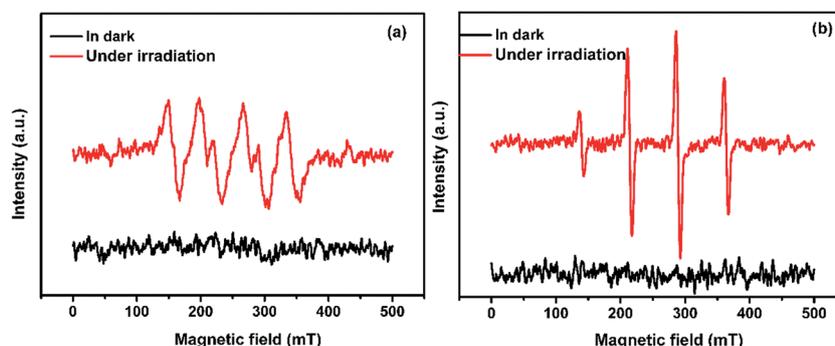


Fig. 10 ESR spectra of radical adducts trapped by DMPO: (a) DMPO-O₂^{•-} radical species detected for the 25 wt% SnO₂/GL-C₃N₄ dispersion in methanol, (b) DMPO- $\cdot\text{OH}$ radical species detected for the 25 wt% SnO₂/GL-C₃N₄ dispersion in water.



irradiation due to its narrow band gap. While the CB and VB edge potential of GL-C₃N₄ are estimated to be -1.20 eV and 1.65 eV.¹³ From the information mentioned above, it is clearly to know that both the CB and VB edges of GL-C₃N₄ are more positive than these of SnO₂. Therefore, when SnO₂/GL-C₃N₄ is excited under visible light irradiation, the corresponding photo-generated electrons in SnO₂ could easily transfer to the CB edge of GL-C₃N₄, while the corresponding photoinduced holes in GL-C₃N₄ could be easily accumulated in the VB edge of SnO₂. Meanwhile, the photogenerated electrons on the CB edge of GL-C₃N₄ can be trapped by oxygen molecules in the aqueous solution to form O₂^{•-} reactive oxygen species, which plays an important role in the degradation of RhB. Afterwards, part of the O₂^{•-} reactive oxygen species continue turning into •OH radical and play its part in the photodegradation of RhB.^{54,55} At the same time, the photogenerated holes on the VB edge of SnO₂ can directly oxidize RhB. Corresponds to the trapping experiments and ESR results, the photocatalytic mechanism confirmed that O₂^{•-}, •OH and holes worked together in combination during the RhB photodegradation process.

4. Conclusion

In summary, the SnO₂/GL-C₃N₄ composites were synthesized in suit by hydrothermal method using GL-C₃N₄ and SnO₂ as precursors. And 25 wt% SnO₂/GL-C₃N₄ composite showed the highest photocatalytic activity, which could degrade about 92% of RhB under visible-light irradiation within 3 h, which was almost 9 and 2.5 times higher than that of SnO₂ and GL-C₃N₄, respectively. The high photodegradation rate of RhB was ascribed to the suitable band potentials of the two components and efficient electron-hole separations *via* interface of SnO₂ and GL-C₃N₄. Moreover, the trapping experiments and ESR spectra all indicated that O₂^{•-}, •OH and holes worked together in combination during the photocatalytic oxidation of RhB.

Conflicts of interest

There are no conflicts of interest to declare.

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References

- 1 Y. P. Zhu, T. Z. Ren and Z. Y. Yuan, *ACS Appl. Mater. Interfaces*, 2015, 7, 16850–16856.
- 2 Z. W. Zhao, Y. J. Sun and F. Dong, *Nanoscale*, 2015, 7, 15–37.
- 3 F. F. Shi, L. L. Chen, C. S. Xing, D. L. Jiang, D. Li and M. Chen, *RSC Adv.*, 2014, 4, 62223–62229.
- 4 X. C. Hao, X. L. Ji and Q. Zhang, *Mater. Lett.*, 2016, 185, 29–31.
- 5 Y. C. Zhang, Q. Zhang, Q. W. Shi, Z. Y. Cai and Z. J. Yang, *Sep. Purif. Technol.*, 2015, 142, 251–257.
- 6 Y. B. Zhao, M. Shalom and M. Antonietti, *Appl. Catal., B*, 2017, 207, 311–315.
- 7 X. H. Pang, H. J. Bian, W. J. Wang, C. Liu, M. S. Khan, Q. Wang, J. N. Qi, Q. Wei and B. Du, *Biosens. Bioelectron.*, 2017, 91, 456–464.
- 8 S. S. Hu, W. J. Ouyang, L. H. Guo, Z. Y. Lin, X. H. Jiang, B. Qiu and G. N. Chen, *Biosens. Bioelectron.*, 2017, 92, 718–723.
- 9 J. L. Lv, K. Dai, J. F. Zhang, Q. Liu, C. H. Liang and G. P. Zhu, *Sep. Purif. Technol.*, 2017, 178, 6–17.
- 10 A. K. Singh, K. Mathew, H. L. Zhuang and R. G. Hennig, *J. Phys. Chem. Lett.*, 2015, 6, 1087–1098.
- 11 S. Jeong, D. Yoo, J. T. Jang, M. Kim and J. Cheon, *J. Am. Chem. Soc.*, 2012, 134, 18233–18236.
- 12 Y. Y. Liu, H. Xiao and W. A. Goddard, *J. Am. Chem. Soc.*, 2016, 138, 15853–15856.
- 13 J. Yan, Z. G. Chen, H. Y. Ji, Z. Liu, X. Wang, Y. G. Xu, X. J. She, L. Y. Huang, L. Xu, H. Xu and H. M. Li, *Chem.–Eur. J.*, 2016, 22, 4764–4773.
- 14 X. D. Zhang, X. Xie, H. Wang, J. J. Zhang, B. C. Pan and Y. Xie, *J. Am. Chem. Soc.*, 2013, 135, 18–21.
- 15 X. P. Dong and F. X. Cheng, *J. Mater. Chem. A*, 2015, 3, 23642–23652.
- 16 Q. J. Xiang, J. G. Yu and M. Jaroniec, *J. Phys. Chem. C*, 2011, 115, 7355–7363.
- 17 W. N. Xing, C. M. Li, Y. Wang, Z. H. Han, Y. D. Hu, D. H. Chen, Q. Q. Meng and G. Chen, *Carbon*, 2017, 115, 486–492.
- 18 Z. Y. Zhang, D. L. Jiang, D. Li, M. Q. He and M. Chen, *Appl. Catal., B*, 2016, 183, 113–123.
- 19 Y. Y. Liu, P. Stradins and S. H. Wei, *Sci. Adv.*, 2016, 2, e1600069.
- 20 J. H. Feng, Y. Y. Li, M. D. Li, F. Y. Li, J. Han, Y. H. Dong, Z. W. Chen, P. Wang, H. Liu and Q. Wei, *Biosens. Bioelectron.*, 2017, 91, 441–448.
- 21 L. X. Hu, F. Y. Chen, P. F. Hu, L. P. Zou and X. Hu, *J. Mol. Catal. A: Chem.*, 2016, 411, 203–213.
- 22 H. L. Xia, H. S. Zhuang, T. Zhang and D. C. Xiao, *Mater. Lett.*, 2008, 62, 1126–1128.
- 23 H. L. Xia, H. S. Zhuang, T. Zhang and D. C. Xiao, *J. Environ. Sci.*, 2007, 19, 1141–1145.
- 24 S. H. Hwang, C. Kim and J. Jang, *Catal. Commun.*, 2011, 12, 1037–1041.
- 25 K. Santhi, C. Rani and S. Karuppuchamy, *J. Alloys Compd.*, 2016, 662, 102–107.
- 26 C. Q. Zhu, Y. Li, Q. Su, B. G. Lu, J. Q. Pan, J. W. Zhang, E. Q. Xie and W. Lan, *J. Alloys Compd.*, 2013, 575, 333–338.
- 27 S. S. Bhande, E. K. Kim, D. V. Shinde, S. Patil, R. S. Mane and S. H. Han, *Int. J. Electrochem. Sci.*, 2013, 8, 11596–11605.
- 28 R. Lamba, A. Umar, S. K. Mehta and S. K. Kansal, *Talanta*, 2015, 131, 490–498.
- 29 C. Wang, Y. Zhou, M. Y. Ge, X. B. Xu, Z. L. Zhang and J. Z. Jiang, *J. Am. Chem. Soc.*, 2010, 132, 46–47.
- 30 T. Tao, L. J. He, J. Li and Y. H. Zhang, *Mater. Lett.*, 2015, 138, 45–47.



- 31 T. T. Li, L. H. Zhao, Y. M. He, J. Cai, M. F. Luo and J. J. Lin, *Appl. Catal., B*, 2013, **129**, 255–263.
- 32 S. C. Yan, Z. S. Li and Z. G. Zou, *Langmuir*, 2010, **26**, 3894–3901.
- 33 H. Xu, J. Yan, Y. G. Xu, Y. H. Song, H. M. Li, J. X. Xia, C. J. Huang and H. L. Wan, *Appl. Catal., B*, 2013, **129**, 182–193.
- 34 H. Zhao, Y. M. Dong, P. P. Jiang, H. Y. Miao, G. L. Wang and J. J. Zhang, *J. Mater. Chem. A*, 2015, **3**, 7375–7381.
- 35 H. D. Liu, *J. Mater. Sci.: Mater. Electron.*, 2014, **25**, 3353–3357.
- 36 A. Akhundi and A. Habibi-Yangjeh, *Mater. Express*, 2015, **5**, 309–318.
- 37 H. T. Ren, S. Y. Jia, Y. Wu, S. H. Wu, T. H. Zhang and X. Han, *Ind. Eng. Chem. Res.*, 2014, **53**, 17645–17653.
- 38 L. Q. Yea, J. Y. Liu, Z. Jiang, T. Y. Peng and L. Zan, *Appl. Catal., B*, 2013, **142**, 1–7.
- 39 Y. P. Zang, L. P. Li, X. G. Li, R. Lin and G. S. Li, *Chem. Eng. J. (Amsterdam, Neth.)*, 2014, **246**, 277–286.
- 40 T. J. Jiang, Z. Guo, J. H. Liu and X. J. Huang, *Electrochim. Acta*, 2016, **191**, 142–148.
- 41 Y. B. Li, H. M. Zhang, P. Liu, D. Wang, Y. Li and H. J. Zhao, *Small*, 2013, **9**, 3336–3344.
- 42 X. J. She, H. Xu, Y. G. Xu, J. Yan, J. X. Xia, L. Xu, Y. H. Song, Y. Jiang, Q. Zhang and H. M. Li, *J. Mater. Chem. A*, 2014, **2**, 2563–2570.
- 43 L. R. Zheng, Y. H. Zheng, C. Q. Chen, Y. Y. Zhan, X. Y. Lin, Q. Zheng, K. M. Wei and J. F. Zhu, *Inorg. Chem.*, 2009, **48**, 1819–1825.
- 44 T. Uddin Md, Y. Nicolas, C. Olivier, T. Toupance, L. Servant, M. M. Müller, H.-J. Kleebe, J. Ziegler and W. Jaegermann, *Inorg. Chem.*, 2012, **51**, 7764–7773.
- 45 C. Q. Zhu, Y. R. Li, Q. Su, B. G. Lu, J. Q. Pan, J. W. Zhang, E. Q. Xie and W. Lan, *J. Alloys Compd.*, 2013, **575**, 333–338.
- 46 H. Y. Ji, X. C. Jing, Y. G. Xu, J. Yan, H. P. Li, Y. P. Li, L. Y. Huang, Q. Zhang, H. Xu and H. M. Li, *RSC Adv.*, 2015, **5**, 57960–57967.
- 47 M. T. Niu, F. Huang, L. F. Cui, P. Huang, Y. L. Yu and Y. S. Wang, *ACS Nano*, 2010, 681–688.
- 48 C. Wang, J. C. Zhao, X. M. Wang, B. X. Mai, G. Y. Sheng, P. A. Peng and J. M. Fu, *Appl. Catal., B*, 2002, **39**, 269–279.
- 49 Y. J. Zhou, L. X. Zhang, W. M. Huang, Q. L. Kong, X. Q. Fan, M. Wang and J. L. Shi, *Carbon*, 2016, **99**, 111–117.
- 50 Zulfiqar, R. Khan, Y. L. Yuan, Z. Iqbal, J. Yang, W. C. Wang, Z. Z. Ye and J. G. Lu, *J. Mater. Sci.: Mater. Electron.*, 2017, **28**, 4625–4636.
- 51 H. Shen, X. R. Zhao, L. B. Duan, R. D. Liu, H. J. Wu, T. Hou, X. W. Jiang and H. D. Gao, *Appl. Surf. Sci.*, 2017, **391**, 627–634.
- 52 I. Shakir, M. Shahid and D. J. Kang, *Chem. Eng. J. (Amsterdam, Neth.)*, 2013, **225**, 650–655.
- 53 H. Xu, H. Z. Zhao, Y. G. Xu, Z. G. Chen, L. Y. Huang, Y. P. Li, Y. H. Song, Q. Zhang and H. M. Li, *Ceram. Int.*, 2016, **42**, 1392–1398.
- 54 B. Yuan, J. X. Wei, T. J. Hu, H. B. Yao, Z. H. Jiang, Z. W. Fang and Z. Y. Chu, *Chin. J. Catal.*, 2015, **36**, 1009–1016.
- 55 Q. Li, N. Zhang, Y. Yang, G. Z. Wang and D. H. L. Ng, *Langmuir*, 2014, **30**, 8965–8972.

