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Low-temperature molten-salt synthesis and upconversion of novel hexagonal NaBiF₄:Er³⁺/Yb³⁺ micro-/nanocrystals

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A series of NaBiF₄:Er³⁺/Yb³⁺ micro-/nanocrystals were synthesized *via* the low-temperature molten-salt method in NH₄NO₃ flux. The influences of the amount of NaF and NH₄NO₃ flux used, the reaction time and reaction temperature on the crystal phase structure, morphology and size of the resultant samples were systematically investigated. The morphology, size, doped concentrations of Er³⁺ and Yb³⁺, pumping power, and temperature dependences of UC emissions were studied in detail under 980 nm irradiation. The NaBiF₄:3 mol% Er³⁺/20 mol% Yb³⁺ crystal exhibits the strongest green UC emission, which was observed under several mW 980 nm excitation at room temperature. The energy transfer between Bi³⁺ and Er³⁺ during UC emission process was also discussed at length.

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1. Introduction

In the past decades, lanthanide-doped upconversion (UC) inorganic materials have evoked a tremendous amount of attention to the ongoing research. UC refers to a system sequentially absorbing two or more long-wavelength (low-energy) photons, combining their energies, and emitting one short-wavelength (high-energy) photon. Compared to semiconductor quantum dots, these UC materials exhibit some advantages of low toxicity, large Stokes shift, sharp emission band, high quantum yield, long lifetime, weak background luminescence, and good resistance to photo-bleaching, blinking and photochemical degradation.^{1–3} Owing to their outstanding properties, they have been proposed for diverse potential applications in photovoltaic,^{3–5} photocatalysis,^{4,5} sensor and biological labels.^{6–9}

Er³⁺, as being an excellent activated ion, can emit different wavelength photons ranging from ultraviolet (UV) through visible (red, blue, green) to near infrared (NIR) light. Unfortunately, Er³⁺ suffers from inferior absorption near 980 nm (compared with Yb³⁺), and low Er³⁺ doping concentration (in the range 0.2–2%) to avoid significant concentration-quenching.^{5–9} To enhance UC efficiency, a popular approach is adopted where a sensitizer with a reasonable cross section in NIR region is co-doped due to an efficient energy transfer between the two. Yb³⁺ has a relatively large absorption cross section (1.2×10^{-20} cm²) compared with that of Er³⁺ (1.7×10^{-21} cm²) and broad absorption band matching with the emission wavelength of powerful commercial laser diodes, and its ²F_{5/2} excited state resonates well

with the ⁴I_{11/2} excited state of Er³⁺, thus Yb³⁺ can significantly improve UC efficiency of Er³⁺.^{10–13}

It is generally accepted that Bi³⁺ ions possess some inherent features of non-toxicity, low cost, and large rare-earth ion admittance. From the point of optical physics, Bi³⁺ ions have 6s² electronic configuration (¹S₀ corresponds to the ground state, whereas ³P₀, ³P₁ and ³P₂ are due to three excited states) and a broad absorption and emission band,^{14–16} which is usually used in downconversion and UC emission system to enhance luminescence.^{16–21} For example, the Bi³⁺ ion in hexagonal NaYF₄ possesses a broad absorption near 300 nm and a emission band ranging from 400 to 500 nm.²² The introduction of the optimal Bi³⁺ can obviously enhance UC emission of hexagonal NaYF₄:Er³⁺/Yb³⁺ materials.²¹ In addition, fluoride compounds possess a host lattice of low phonon energies (less than 350 cm⁻¹) causing to decrease nonradiative relaxation and in turn to yield relatively strong UC emission, they have served as very efficient UC matrixes. For instance, NaREF₄ (RE = Y, Gd and Lu) is one kind of the efficient NIR-vis UC host materials.^{23–27} Therefore, Bi³⁺-based fluorides can combine the advantages of both fluorides and Bi³⁺ ions, they are suggested to be novel UC materials with efficient luminescence and have been attracted more and more attention.^{28–31} As one member of Bi³⁺-based fluorides, NaBiF₄ belongs to the hexagonal system with the space group $P\bar{3}$ and the unit cell parameters: $a = 6.144$ Å, $c = 3.721$ Å (ref. 32) and has the same structure as NaREF₄,^{1–3} thereby, the addition of Er³⁺ and Yb³⁺ can not generate the crystal lattice distortion and exhibits the capacity of heavily rare-earth ion doping. The Yb³⁺/RE³⁺ (RE = Er, Tm, and Ho) doped hexagonal NaBiF₄ nanocrystals has demonstrated to be one potential excellent UC material.³¹

In past several decades, rare-earth doped NaREF₄ materials with high dispersability were synthesized in organic solvents by

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hydro(solvo)thermal methods.^{1–11,33–35} In contrast with the above methods, the most advantage of low-temperature molten-salt method is an economical mass, convenient, effective, facile and environmental friendly approach. And it is very easy to obtain the materials with clean surface, chemical purification, and few residual impurities below the melting point of the as-grown crystals. Ammonium nitrate (NH₄NO₃) exhibits some inherent advantages, including abundance in nature, low melting point, and easy deliquescence in water/alcohol causing to easily separate from the flux and the target compound. Thus, this synthetic route has been adopted to fabricate hexagonal NaYF₄:Er³⁺/Yb³⁺,³⁶ tetragonal LiYF₄:Er³⁺/Yb³⁺,³⁷ cubic BaGdF₅:Ce³⁺/Er³⁺/Yb³⁺ (ref. 38) and monazite LaPO₄:Eu³⁺.³⁹ Recently, the Yb³⁺/RE³⁺ (RE = Er, Tm, and Ho) doped hexagonal NaBiF₄ nanocrystals have been fabricated in ethylene glycol at room temperature.³¹ Unfortunately, the degree of the crystalline and their UC emissions need to be further improved. Therefore, here, hexagonal NaBiF₄:Er³⁺/Yb³⁺ (h-NaBiF₄:Er³⁺/Yb³⁺) micro-/nanocrystals will be prepared by the low-temperature molten-salt synthesis in NH₄NO₃ flux. In this synthetic route, the phase structure, shape and size of the samples could be highly determined by the intrinsic structure of target compound and the growth surroundings (such as the reaction temperature, the reaction time, and the usage amounts of raw materials and flux). To obtain high-quality NaBiF₄:Er³⁺/Yb³⁺ micro-/nanocrystals, it is very necessary to fastidiously control these parameters. Herein, we will mainly concentrate upon the influence of the usage amount of reaction materials NaF and flux NH₄NO₃, the reaction temperature and the reaction time on the chemical composition, the morphology and size of the resultant products. On this basis, the corresponding UC properties will be investigated and discussed in detail. Finally, the

UC emission dynamics of h-NaBiF₄:Er³⁺/Yb³⁺ will be studied by the pumping power, and temperature dependences of UC emissions and fluorescence decays.

2. Experimental

2.1 Material preparation

NaBiF₄:Er³⁺/Yb³⁺ samples were synthesized by the low temperature molten-salt method in NH₄NO₃ flux. Reagent-grade NaF, NH₄NO₃, Bi(NO₃)₃·5H₂O, Yb(NO₃)₃·5H₂O and Er(NO₃)₃·5H₂O powders were used to synthesize hexagonal Er³⁺/Yb³⁺-codoped NaBiF₄ crystals. All the reagent-grade powders and NH₄NO₃ with different usage amounts were mixed well. The mixtures were put into 15 ml capacity crucibles. After the lids were tightly closed, the crucibles were placed in an oven. The sealed tank was heated to different reaction temperature (160, 175, 210 and 250 °C), and held for different hours (0.5, 1, 2, 4, 8, 12 and 24 h) in an oven, and then cooled to room temperature naturally. The detailed experimental parameters including the reaction temperature, reaction time, the doping concentration of Er³⁺ and Yb³⁺, the NaF usage amount and the usage amount of NH₄NO₃ flux were listed in Table 1. After being washed several times with deionized water and ethanol, the precipitates were dried at 70 °C for 12 hours in vacuum.

2.2 Material characterization

The samples were examined by X-ray diffraction (XRD), transmission electron microscopy (TEM), energy-dispersive X-ray spectroscopy (EDS) and photoluminescence (PL). XRD analyses were carried out on a Bruker D8-Advance diffractometer with graphite-monochromatized Cu K α radiation (40 kV/60 mA, graphite monochromator, $\lambda = 0.1541$ nm). The size,

Table 1 The detailed experimental parameters and crystal phase of the as-obtained samples

Sample	NaF/Ln molar ratio	NH ₄ NO ₃ /Ln molar ratio	Reaction temperature (°C)	Reaction time (h)	Crystal phase
S1	3	5	250	24	c-BiF ₃ + t-BiOF
S2	4	5	250	24	
S3	5	5	250	24	h-NaBiF ₄
S4	6	5	250	24	h-NaBiF ₄
S5	8	5	250	24	h-NaBiF ₄
S6	5	8	250	24	h-NaBiF ₄
S7	5	10	250	24	h-NaBiF ₄
S9	6	8	250	24	h-NaBiF ₄
S10	6	10	250	24	h-NaBiF ₄
S11	8	8	250	24	h-NaBiF ₄
S12	8	10	250	24	h-NaBiF ₄
S13	5	5	210	24	h-NaBiF ₄
S14	5	5	180	24	h-NaBiF ₄ + c-BiF ₃ + t-BiOF
S15	5	5	160	24	m-NH ₄ BiF ₄ + c-BiF ₃ + t-BiOF
S16	3	5	250	12	h-NaBiF ₄
S17	3	5	250	8	h-NaBiF ₄
S18	3	5	250	4	h-NaBiF ₄
S19	3	5	250	2	h-NaBiF ₄
S20	3	5	250	1	h-NaBiF ₄ t-BiOF
S21	3	5	250	0.5	h-NaBiF ₄ t-BiOF



morphology and chemical compositions of products were determined by a transmission electron microscopy (TEM, JEOL2010) operating at 200 kV and a JSM 6700F scanning electron microscope (SEM) equipped with the energy dispersive X-ray spectrum (EDS). Structural information of the nanocrystals was measured by a high-resolution transmission electron microscopy (HRTEM).

Downconversion (DC) luminescence spectra were performed using an Edinburgh Instruments FLS920 fluorescence spectrometer equipped with both continuous (450 W) and pulsed xenon lamps. UC emission spectra were carried out upon 976 nm excitation with a pump power of ~ 200 mW (power density ~ 20 W cm $^{-2}$), provided by a mode-locked picosecond Ti:sapphire laser (700–1000 nm, pulse width ≤ 1.5 ps, Tsunami, Spectra-Physics). The measurements were repeated in an integrating sphere in order to reduce the effect of the scattering on the UC properties. Power-dependence of UC emissions of Er $^{3+}$ was studied upon excitation by a power-controllable 976 nm diode laser (DPL-II, Module HTL98M10) with two emission monochromators to record the emission spectra in the wavelength range of 200–850 nm (with a Hamamatsu R928 photomultiplier tube). Its maximum power output is 5 W. To eliminate the noise from tunable Ti:sapphire laser, 980 nm LD, and both continuous (450 W) and pulsed xenon lamps, different filters were kept in front of the detector. The emission monochromator's slits were set as small as possible to maximize the instrumental resolution. The best wavelength resolution was 0.05 nm. The line intensities and positions of the measured spectra were calibrated according to the FLS920 correction curve and standard mercury lamp. All the spectra were measured at room temperature and under identical conditions for each serial test. Luminescence decays were recorded using a customized UV to mid-infrared steady-state and phosphorescence lifetime spectrometer (FSP920-C, Edinburgh) equipped with a digital oscilloscope (TDS3052B, Tektronix) and a tunable mid-band OPO pulse laser as the excitation source (410–2400 nm, 10 Hz, pulse width ≤ 5 ns, Vibrant 355II, OPOTEK). For the low-temperature measurements, the samples were mounted on a closed-cycle cryostat (10–350 K, DE202, Advanced Research Systems).

3. Results and discussion

3.1 Phase structure

When the molar ratio of NaF to Ln(NO $_3$) $_3$ (hereinafter to be referred as NaF/Ln) was 3, the resultant product consists of cubic BiF $_3$ (c-BiF $_3$) (JCPDS no. 74-0144) and tetragonal BiOF (t-BiOF) (JCPDS no. 86-1478) (Fig. 1a) and the XRD peaks due to hexagonal NaBiF $_4$ (h-NaBiF $_4$) (JCPDS no. 041-0796) were not observed. When the NaF/Ln molar ratio increased to 4, h-NaBiF $_4$ phase appeared and played a predominant role (Fig. 1a), and yet a small quantity of c-BiF $_3$ could be still detected. As the NaF/Ln molar ratio increased to 5, the diffraction peaks due to c-BiF $_3$ disappeared and those due to t-BiOF continued to weaken. The further increasing NaF/Ln molar ratio to 8, the diffraction peaks due to the t-BiOF vanished, all the diffraction peaks were in

good agreement with those of pure h-NaBiF $_4$ and no trace of characteristic peaks for other impurity phases was observed, implying that the samples obtained have the same crystallographic structure as pure h-NaBiF $_4$ crystal. Thus, the h-NaBiF $_4$ can be synthesized at NaF/Ln molar ratio above 5, and the NH $_4$ NO $_3$ molten solution system can be developed to prepare the h-NaBiF $_4$ materials.

From Fig. 1b, it can be seen that the variance of the reaction temperature also highly affects the phase structure of the products. For the sample obtained at 210 and 250 °C for 12 hours, only h-NaBiF $_4$ (JCPDS no. 041-0796) was obtained. When the reaction temperature decreased to 175 °C, the t-BiOF and monoclinic NH $_4$ BiF $_4$ (m-NH $_4$ BiF $_4$) (JCPDS no. 041-0796) were formed. With the reaction temperature further decreasing to 160 °C, the m-NH $_4$ BiF $_4$ became more and more and played a predominant role while the amount of h-NaBiF $_4$ significantly decreased. As is known, the thermodynamic growth regime is highly dependent upon a sufficient supply of thermal energy, and the crystal with the most stable structure preferential grows.⁴⁰ Since h-NaBiF $_4$ is thermodynamically more stable than m-NH $_4$ BiF $_4$ and t-BiOF, m-NH $_4$ BiF $_4$ and t-BiOF are susceptible to transform to h-NaBiF $_4$. The energy barrier hinders the formation of the h-NaBiF $_4$, a sufficient supply of thermal energy must be required to surmount this energy barrier to adjust the environment of Bi $^{3+}$ (Er $^{3+}$ and Yb $^{3+}$) and Na $^+$ occupation sites.⁴¹ Therefore in this thermal system m-NH $_4$ BiF $_4$ and t-BiOF can be formed at low temperature, relatively high reaction temperature is more favorable for this transformation (m-NH $_4$ BiF $_4$ \rightarrow NaBiF $_4$ and t-BiOF \rightarrow h-NaBiF $_4$) in the molten salt process and the pure h-NaBiF $_4$ crystals can be synthesized at the reaction temperature of 210 °C and above.

To study the impact of the reaction time on the phase structure and the growth mechanism, the XRD patterns of the products prepared for different reaction times of 0.5, 1, 2, 4, 8, 12 and 24 h, as shown in Fig. 1c, indicates that the crystal phase of the samples is strongly dependent on the reaction time. When the reaction time was 0.5 h, a great amount of h-NaBiF $_4$ and a small amount of t-BiOF were also observed and m-NH $_4$ BiF $_4$ (JCPDS no. 041-0796) was hardly detected. The t-BiOF became less and less with the elongation of the reaction time from 0.5 hour to 4 hours and disappeared as the reaction time extended to 8 and 12 hours. It is demonstrated that the longer reaction time is also beneficial to transform from t-BiOF to h-NaBiF $_4$ and the pure h-NaBiF $_4$ crystals can be also obtained when the reaction time exceeds 8 h. Moreover, in comparison to (101) and (201) XRD peaks owing to the hexagonal NaBiF $_4$, the (110) and (110) XRD peaks becomes stronger with the extension of the reaction time, implying that the [0001] direction is the preferred direction of crystal growth.

XRD patterns of the products with different concentrations of Er $^{3+}$ or Yb $^{3+}$ for 24 hours are shown in Fig. 1d, e and f with h-NaBiF $_4$ standard card (no. 41-0796). All the XRD peaks are in good agreement with those of the h-NaBiF $_4$, showing that all the products have the same crystal structure as pure h-NaBiF $_4$. Because of the replacement of Bi $^{3+}$ with smaller Er $^{3+}$ and Yb $^{3+}$ ions in the host lattice, the diffraction peaks shift slightly to higher angle side as a result of the shrinkage of the unit-cell volume.



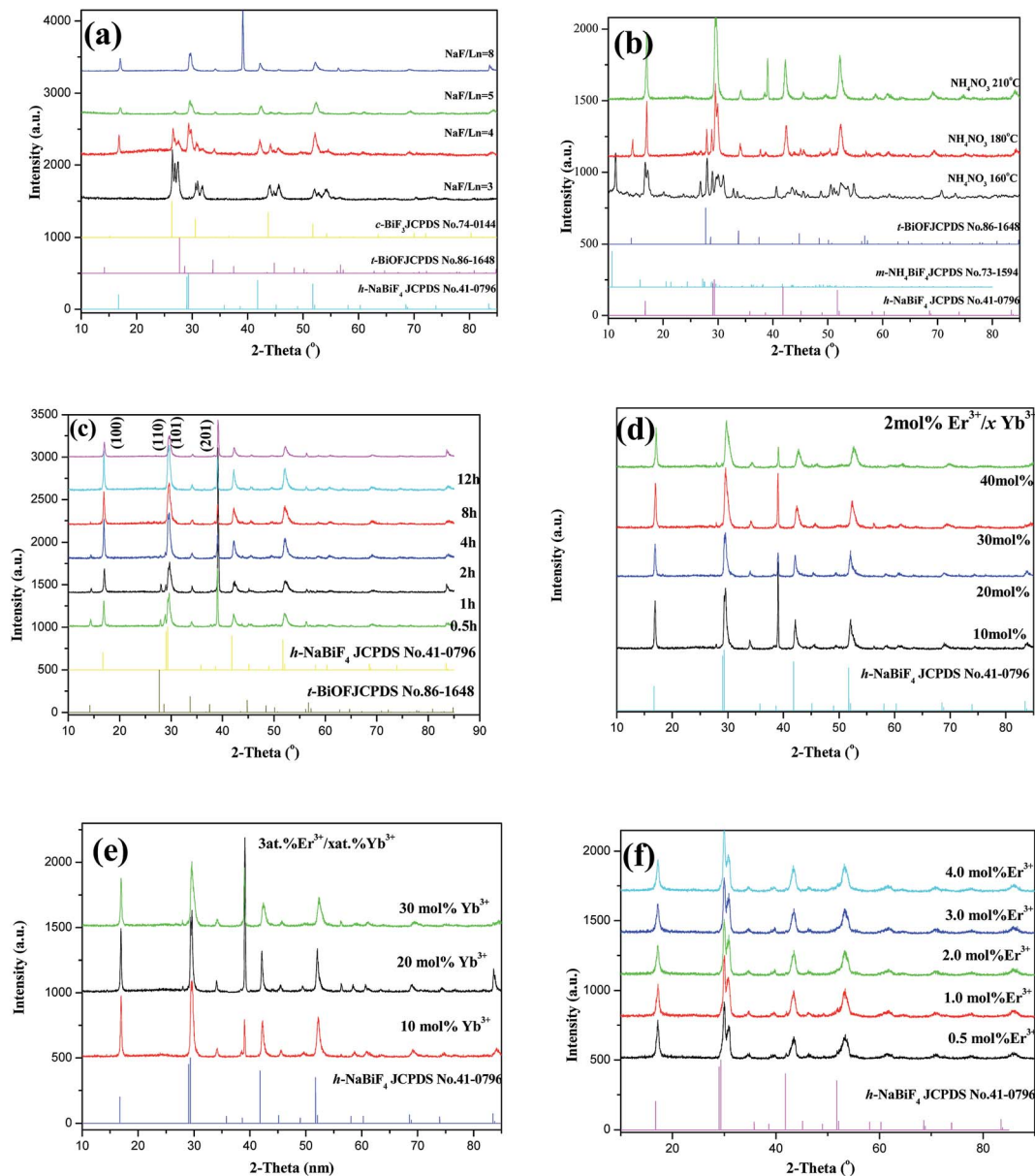


Fig. 1 The XRD patterns of the samples prepared under different conditions. (a) Different molar ratios of NaF/Ln; (b) different temperatures; (c) different times; (d) doped with 2 mol% Er^{3+} and different Yb^{3+} concentrations; (e) doped with 3 mol% Er^{3+} and different Yb^{3+} concentrations; (f) doped with 20 mol% Yb^{3+} and different Er^{3+} concentrations.

3.2 Morphology

Fig. 2–5 illustrate some representative SEM images of $\text{h-NaBiF}_4:\text{Er}^{3+}/\text{Yb}^{3+}$ prepared under different reaction conditions, as listed in Table 1. When the molar ratio of NaF/Ln was 3.0, the corresponding sample is composed of a large number of tiny nanoparticles attached to the surface of large polyhedral particles with size of *ca.* 0.6 μm . The product prepared at the NaF/Ln molar ratio of 4 mainly consists of many prismatic microrods with 4–5 μm in length and *ca.* 0.8 μm in diameter. Their lateral surfaces (Fig. 2a and 3a) are covered with a layer of minute irregular nanoparticles while their end planes are coarse, concave and irregular. When the molar ratio of NaF/Ln was increased to 5.0, the product maintains rod-like morphology

and becomes even uniform. Their sizes and end planes keep nearly unchanged, whereas their lateral surfaces become relatively smooth and have some fine slit along the rods. With the increasing NaF/Ln molar ratio to 6.0, the morphology of the product still exhibits prismatic microrods, whereas their length and diameter slightly decrease to 3–4 μm and *ca.* 0.8 μm , respectively. Some small holes appear on their surfaces. Further increasing NaF/Ln molar ratio to 8.0 deduces the prisms incomplete and their length and diameter are 3–5 μm and 0.6 μm , respectively. There are a great number of relatively large holes on their side surfaces of these prisms.

To study the impact of the variation of the amount of flux NH_4NO_3 upon the morphology and size of the products, the



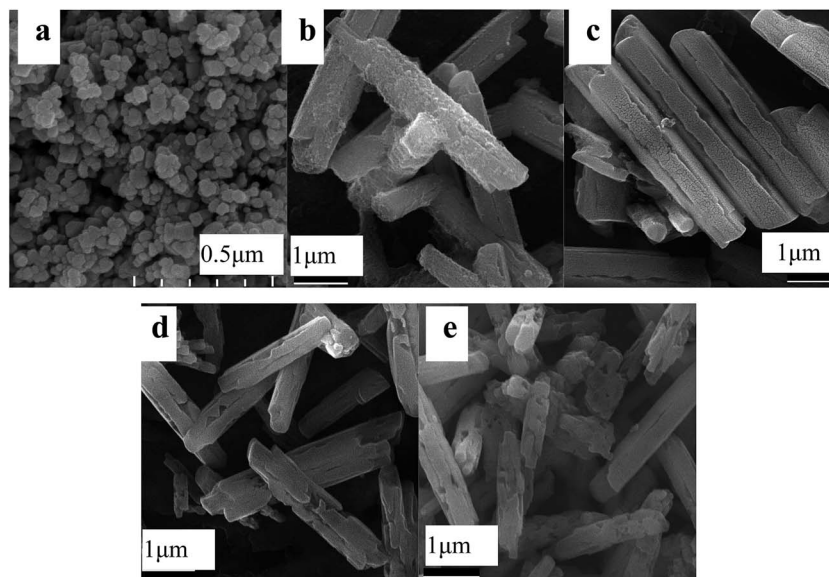


Fig. 2 Representative SEM images of the as-obtained products at 250 °C for 24 hours with the $\text{NH}_4\text{NO}_3/\text{Ln}$ molar ratio of 5, and different NaF/Ln molar ratios of 3 (a), 4 (b), 5 (c), 6 (d) and 8 (e).

SEM images of the samples with different $\text{NH}_4\text{NO}_3/\text{Ln}$ molar ratios were shown in Fig. 3. The NH_4NO_3 usage amount produces a great influence upon their sizes. When the NaF/Ln molar ratio was fixed to be 5 and 6, the molar ratio of $\text{NH}_4\text{NO}_3/\text{Ln}$ increased to 8 and 10, the resultant products still maintain prism-like morphology. Meanwhile, their diameters also nearly stay at the same level (*ca.* 0.8 μm), and the mean

lengths decrease to *ca.* 3.0 μm . For the NaF/Ln molar ratio of 8, the $\text{NH}_4\text{NO}_3/\text{Ln}$ molar ratio enhances from 5 through 8 to 10, the diameters of these rods also keep almost invariable (*ca.* 0.5 μm), the lengths experience a significant downward from 3.0 μm through 2.0 μm to 1.0 μm . That is to say, the larger $\text{NH}_4\text{NO}_3/\text{Ln}$ molar ratio is, the smaller the aspect ratio length/diameter becomes.

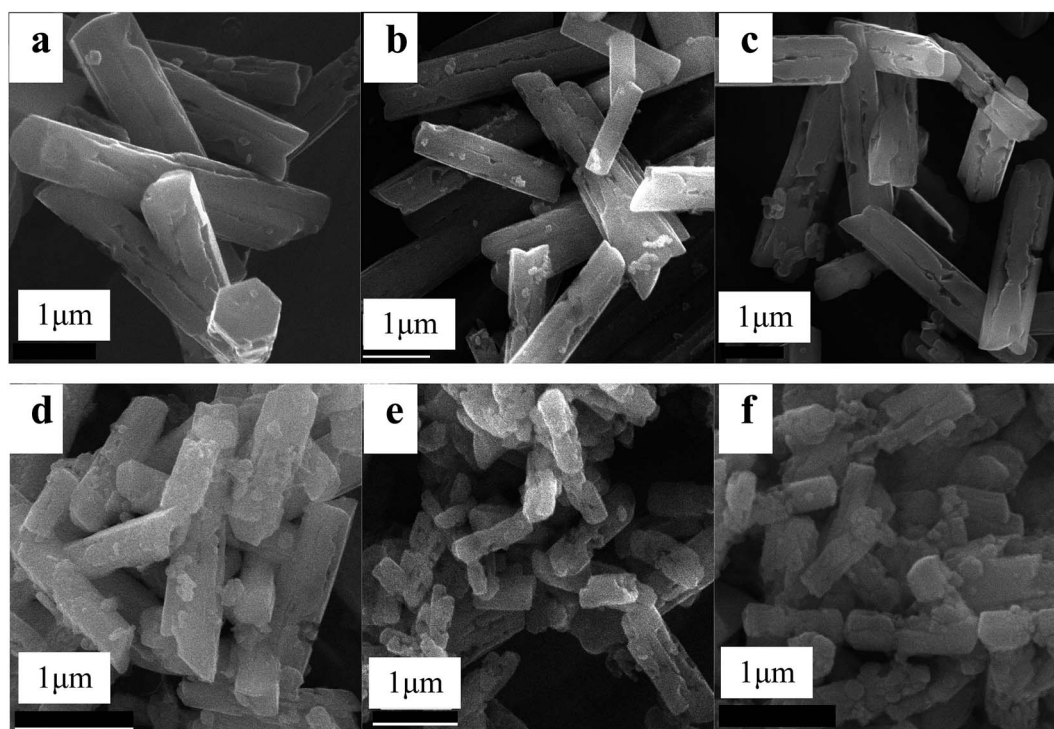


Fig. 3 Representative SEM images of the as-obtained products at 210 °C for 24 hours with the $\text{NaF}/\text{NH}_4\text{NO}_3/\text{Ln}$ molar ratio of 5/8/1 (a), 5/10/1 (b), 6/8/1 (c), 6/10/1 (d), 8/8/1 (e) and 8/10/1 (f).



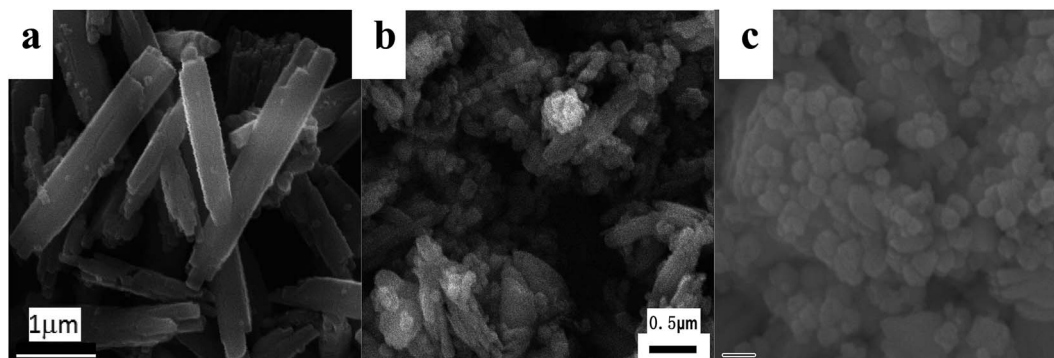


Fig. 4 Representative SEM images of the as-obtained products under for 24 hours with the $\text{NH}_4\text{F}/\text{Ln}$ molar ratio of 5 and $\text{NH}_4\text{NO}_3/\text{Ln}$ molar ratios of 8 at reaction temperature of 210 °C (a), 180 °C (b) and 160 °C (c).

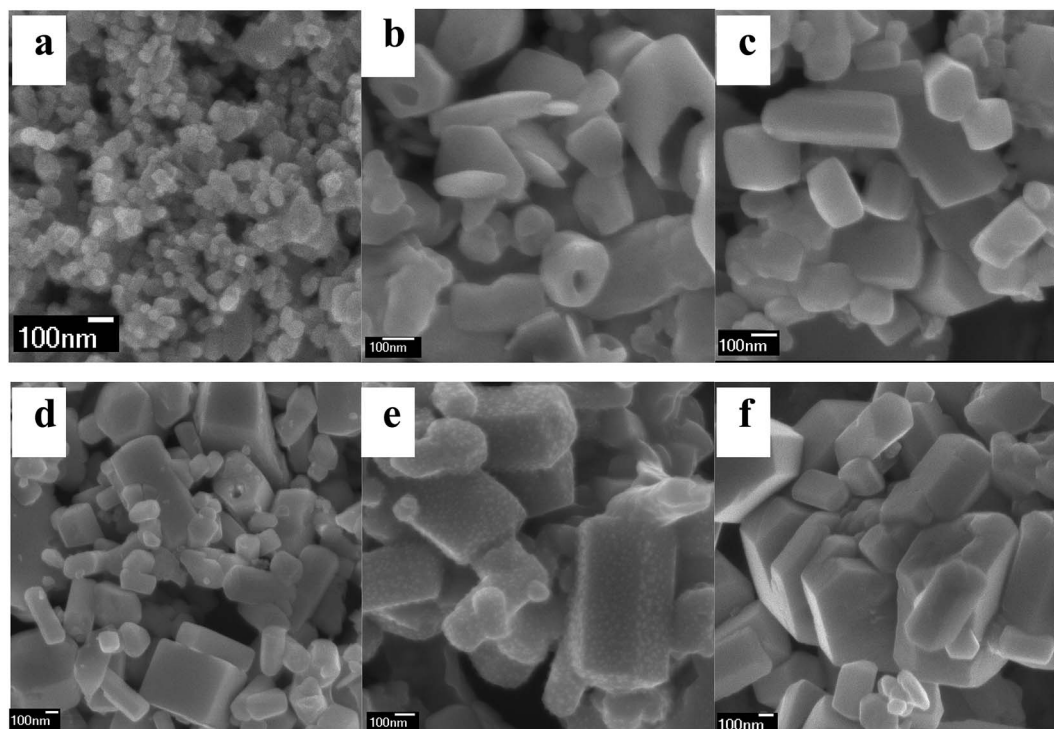


Fig. 5 Representative SEM images of the as-obtained products under at 250 °C for 24 h with the $\text{NH}_4\text{F}/\text{Ln}$ molar ratio of 5 and $\text{NH}_4\text{NO}_3/\text{Ln}$ molar ratios of 8 for the reaction time of 0.5 h (a), 1 h (b), 2 h (c), 4 h (d), 8 h (e) and 12 h (f).

To elucidate the effect of the reaction temperature on the morphology the detailed temperature-dependent experiments were carried out under the similar reaction conditions. When the reaction temperature decreases from 250 °C to 210 °C, the corresponding sample takes on prismatic shape with the length of *ca.* 2.0–3.0 μm, the diameter of *ca.* 0.3–0.5 μm and irregular end planes (Fig. 4a). In comparison with S3, the length and diameter simultaneously have a downward trend. Further decreasing reaction temperature to 180 °C deduces the prismatic rods disappear, and some tiny nanoparticles aggregate together into the prismatic embryonic form (Fig. 4b). When the reaction temperature is 160 °C, the prismatic embryonic form vanish, a great number of tiny nanoparticles have come into being (Fig. 4c).

To shed light on the morphology evolution of h- NaBiF_4 crystal, taking S3 as an example, the morphologies of the intermediates at different growth stages are shown in Fig. 5 and 6. The intermediates exhibit different morphologies at different reaction times. The reaction treatment for 0.5 hour leads to the formation of small hexagonal prisms with average sizes of 10–50 nm in length and 20–50 nm in diameter (Fig. 5a and 6a). The high-magnification TEM image indicates that the fringe distance (*d*-spacing) is 0.311 nm (Fig. 6b), corresponding to the distance between neighboring (110) planes of h- NaBiF_4 (JCPDS no. 041-0796). The selected area electron diffraction (SAED) pattern, as shown in Fig. 6c, shows the diffraction rings corresponding to (110), (111), (210), (102) and (310) planes of h-



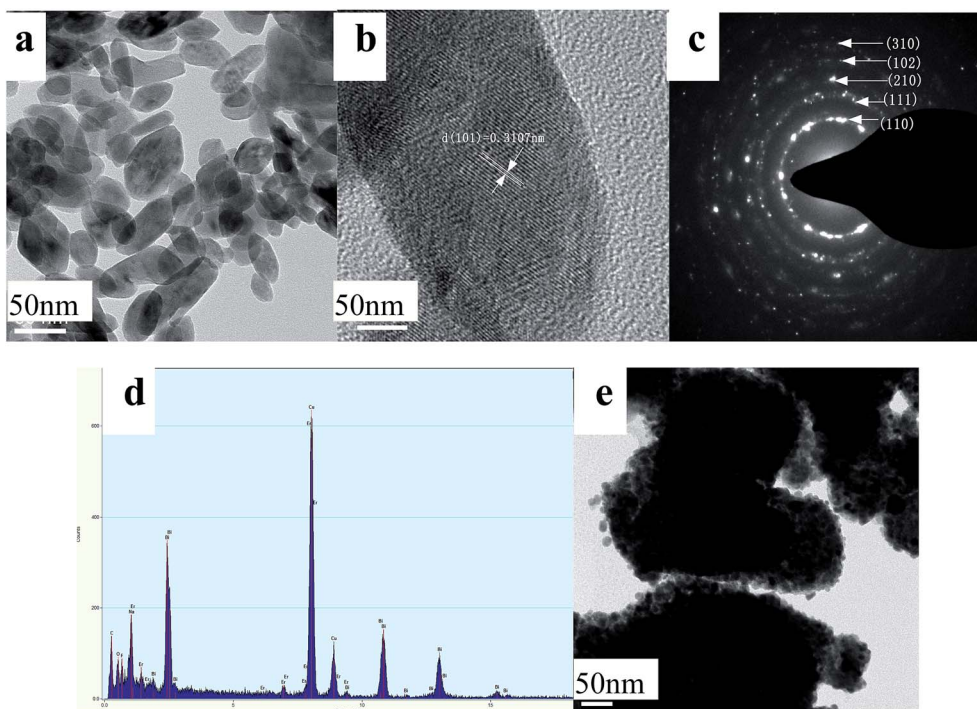


Fig. 6 High-magnification TEM image (a), high-resolution TEM image (b), SAED pattern (c); and EDS spectra (d) of as-prepared $\text{NaBiF}_4:3 \text{ mol\% Er}^{3+}/20 \text{ mol\% Yb}^{3+}$ sample S21 for 1.0 h, and high-magnification TEM image (e) of as-prepared $\text{NaBiF}_4:3 \text{ mol\% Er}^{3+}/20 \text{ mol\% Yb}^{3+}$ sample S20 for 1.0 h.

NaBiF_4 . The compositional analysis by EDS pattern spectrum (as shown in Fig. 6d) reveals that the sample is composed of Er, Yb, Bi, Na and F elements. When the reaction time was increased to 1.0 h, some short and imperfect hexagonal prismatic particles are formed (Fig. 5b). TEM image reveals that these prismatic products are actually polycrystalline, constructed by small particles (Fig. 6e). The holding time of 2.0 h reduced these short and imperfect hexagonal prisms to evolve into the regular ones with the length of 200–500 nm and the diameter of 150–200 nm and convex end planes (Fig. 5c). When the reaction time was extended to 4 and 8 h, the hexagonal prisms continue to grow large while the small particles become less and less (Fig. 5d and e). As the reaction time elongated to 12 h, most of small nanoparticles disappeared and transformed to the hexagonal prismatic rods or plates (Fig. 5f). When the reaction time further extended to 24 h, these irregular nanoparticles depleted and converted to the hexagonal prismatic rods or plates (Fig. 2c). The above results reveal that the rod-like particles undergo dissolution–recrystallization transformation in order to minimize the surface energy of the system, and the h- NaBiF_4 prisms grow along *c*-axis direction at the consumption of smaller nanoparticles by Ostwald ripening process.⁴²

3.3 Formation mechanism

Herein we have experimentally and independently study the impacts of (1) the reaction time, (2) the reaction temperature, (3) the usage amount of NaF (4) the usage amount of NH_4NO_3 flux upon the resultant morphology and chemical composition.

Based on the above results, a possible phase and morphology evolution mechanism is described as follows.

At the beginning, $\text{Ln}(\text{NO}_3)_3 \cdot 5\text{H}_2\text{O}$ ($\text{Ln} = \text{Bi}, \text{Yb}$ and Er) and NaF are mixed with the desired amount of NH_4NO_3 salts to form a precursor, which then is reacted above the melting point to provide a fluid phase. Since the melting point of NH_4NO_3 , $\text{Bi}(\text{NO}_3)_3 \cdot 5\text{H}_2\text{O}$, $\text{Yb}(\text{NO}_3)_3 \cdot 5\text{H}_2\text{O}$ and $\text{Er}(\text{NO}_3)_3 \cdot 5\text{H}_2\text{O}$ is less than 100 °C, they quickly melt, dissociate and form Bi^{3+} , Yb^{3+} , Er^{3+} , NH_4^+ and NO_3^- ions during the heat-up process, whereas too high melting point of NaF (*ca.* 993 °C) cannot melt in such low experimental temperature, howbeit NaF can be partly dissolved to form Na^+ and F^- ions due to the existence of crystalliferous water. In addition, some Bi^{3+} ions tend to experience a hydrolysis process and generate BiO^+ ions during warming-up process.

When the usage amount of F^- is beyond the stoichiometric ratio of h- NaBiF_4 , Na^+ , Ln^{3+} ($\text{Ln} = \text{Bi}, \text{Yb}, \text{Er}$) and NH_4^+ ions in the molten salt interact with F^- to generate m- NH_4BiF_4 and h- NaBiF_4 small nuclei. If the number of F^- ions can not exceed this ratio, the m- NH_4BiF_4 and c- BiF_3 take precedence to produce in this molten-salt system. Additionally, those BiO^+ ions react with F^- ions and produce t- BiOF small nuclei. In thermodynamics, m- NH_4BiF_4 and t- BiOF are lower than h- NaBiF_4 and c- BiF_3 , and susceptible to evolve inevitably to h- NaBiF_4 and c- BiF_3 seed through dissolution–renucleation process. These nuclei quickly aggregate together and grown into t- BiOF and h- NaBiF_4 crystals. As the reaction temperature enhances and reaction time extends, the rate of transformation from m- NH_4BiF_4 and t- BiOF to h- NaBiF_4 becomes larger, more m- NH_4BiF_4 and t- BiOF convert to h- NaBiF_4 by dissolution–recrystallization process.



On the basis of the general theory of crystal growth, the growth crystal is relevant to the relative growth rate of different crystal faces. The different crystal rates give rise to diverse appearance of the crystalline. Since h-NaBiF₄ micro-/nano-crystals have a hexagonal shape, their surfaces are typical (001) top/bottom planes and six prismatic side planes of the energetically equivalent (100) family. In this molten salt system, the growth rate on the *c*-axis in the crystal growth process is quicker than that on the directions perpendicular to *c*-axis, thereby the sample takes on the hexagonal prism. At the same time, the selective adhesion of some simple ions such as NH₄⁺ ions and Na⁺ ions on some specific crystalline planes can change the epitaxial growth rate of different crystallographic direction.³⁸ In our experimental surrounding, NH₄⁺ or Na⁺ ions preferentially adsorb the (001) crystal plane and thereby remarkably prohibit the *c*-axis (*i.e.* [001]) directed growth, there is a tendency to shorten the length of prism. The observation of the concave structure at top/bottom facets suggests that the growth rate of the prismatic side facets is little faster than that of the top/bottom facets. These prism-like crystals exhibit the presence of some defects (such as holes and linear slits) on their side surfaces of these prisms, which may be reflected on the preferential 2D-nucleation at the border of the growing face. This behavior can be resulted from the morphological instability generated by the competition between the supersaturation (destabilizing factor) and the surface tension (stabilizing factor).^{43,44}

3.4 Upconversion emission

To study the effects of Er³⁺ and Yb³⁺ concentrations on the UC properties of h-NaBiF₄:Er³⁺/Yb³⁺ products, Fig. 7 depicts the room temperature UC emission spectra of h-NaBiF₄:Er³⁺/Yb³⁺ samples with different doping concentrations upon 980 nm excitation. The strong green UC emission band is observed in the region of 512–573 nm, attributed to the ²H_{11/2}/⁴S_{3/2} → ⁴I_{15/2} transitions of Er³⁺. The red UC emission between 630–700 nm belongs to ⁴F_{9/2} → ⁴I_{15/2} transition of Er³⁺. It is striking that the predominant green emission is much stronger than the red emission. Thus, it can emit intense green light with high purity. Besides green and red UC emission, other weak UC emissions at 370–390, 400–428, and 445–460 nm, stem from the ⁴G_{11/2}, ²H_{9/2}, ⁴F_{3/2} and ⁴F_{5/2} states to the ⁴I_{15/2} ground state of Er³⁺,^{9–12} respectively. The UC emission in the range between 485 and 505 nm is attributed to the overlap of the cooperative UC of Yb³⁺,^{45–47} ⁴F_{7/2}–⁴I_{15/2} transition of Er³⁺ (ref. 9–12) and the ³P₁–¹S₀ transition of Bi³⁺ ion^{14,15} (see Fig. 7a). It is clearly seen from Fig. 7a and b that the UC emission is strongly dependent upon the concentration of Er³⁺ and Yb³⁺ ions. In the samples with the constant Er³⁺ concentration (2 mol% and 3 mol%) the UC emission intensities except ⁴F_{7/2}–⁴I_{15/2} UC emission increase as the Yb³⁺ concentration increases from 10 mol% to 20 mol%, and the maximum UC emission intensity appears at 20 mol% Yb³⁺ doping concentration. Then further increasing Yb³⁺ doping concentration (such as 30 mol%, 40 mol%) deduces the

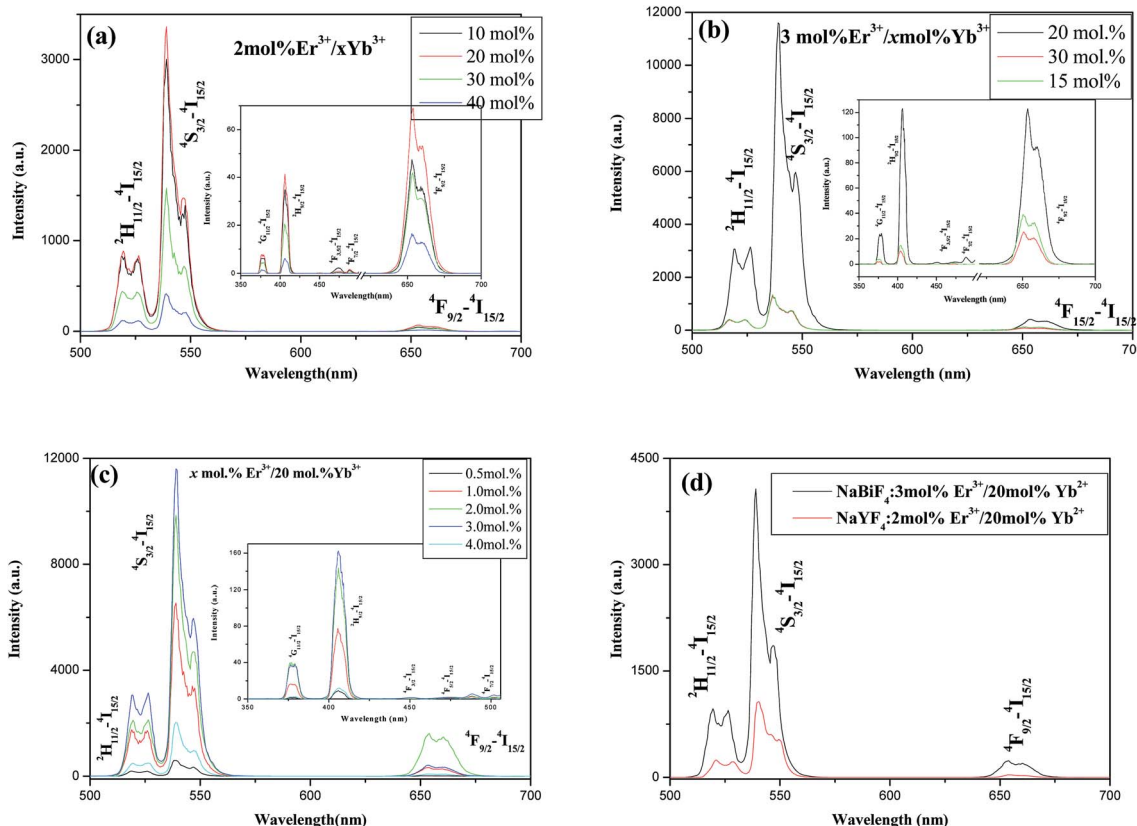


Fig. 7 The UC emission spectra of the h-NaBiF₄:Er³⁺/Yb³⁺ samples with different concentrations of Er³⁺ and Yb³⁺: (a) 2 mol% Er³⁺/xYb³⁺ (x = 10, 20, 30 and 40 mol%); (b) 3 mol% Er³⁺/xYb³⁺ (x = 10, 20 and 30 mol%); (c) xEr³⁺/20 mol% Yb³⁺ (x = 0.5, 1, 2, 3 and 4 mol%).



UC emission to fall (Fig. 7a and b). In similar fashion in the samples with the constant Yb^{3+} doping (20 mol%), the UC emission starts an upward trend as the Er^{3+} doping concentration increases from 0.5 mol% to 2 mol%, arrives at a peak at 3.0 mol% Er^{3+} doping concentration, then further increasing Er^{3+} concentration deduces to fall (see Fig. 7c). Thus, NaBiF_4 :3 mol% $\text{Er}^{3+}/20$ mol% Yb^{3+} sample exhibits the strongest UC emission, and possesses stronger green UC emission about 4 times than hexagonal NaYF_4 :2 mol% $\text{Er}^{3+}/20$ mol% Yb^{3+} nanocrystals under experimental condition (Fig. 7d). The size and morphology of the latter exhibit small conical hexagonal prisms with *ca.* 50 nm in diameter and 80–100 nm in length (see Fig. 7b and e in ref. 36).

Fig. 8 shows the UC emission spectra of NaBiF_4 :3 mol% $\text{Er}^{3+}/20$ mol% Yb^{3+} prepared under different conditions (including the reaction time, the reaction temperature, NaF/Ln and NH_4NO_3 /Ln molar ratio). It can be seen that the UC emissions are similar in the line, and the only difference between these bands are their intensities. From the figure plugged in Fig. 8a, the as-obtained sample at the NaF/Ln molar ratio of 5 has the highest UC emission intensity, and the strong green emission can be observed under several mW 980 nm excitation. The samples prepared at the NaF/Ln molar ratio of 3 including c-BiF₃ and t-BiOF possess the lowest UC intensity, whereas the NaF/Ln molar ratio of 4 leads to the formation of h- NaBiF_4 , thus substantially improving the intensity of UC emission. The resulting samples at the NaF/Ln molar ratio of 5, 6, and 8 exhibit prism with the almost large diameter (*ca.* 0.8 μm), their mean lengths become short. Similar to the hexagonal NaYF_4 :2 mol%

$\text{Er}^{3+}/20$ mol% Yb^{3+} , the green UC emission gradually decreased as the NaF/Ln molar ratio increased from 5 through 6 to 8, that is to say, the larger the length of the prism is, the stronger the UC emission is (see Fig. 3 and Table 1).

In addition, the change of the NH_4NO_3 usage amount can alter UC emission intensities of the samples (Fig. 8b). For the fixed NaF/Ln molar ratio (5, 6 and 8, respectively), as NH_4NO_3 /Ln molar ratio gradually increases, the UC emission intensity of the corresponding sample also gradually reduces (plugged in Fig. 8b). This can be also related to the morphology and size of the sample. The increase of NH_4NO_3 /Ln molar ratio deduces the length of the rod to shorten, thereby the UC emission of the corresponding sample also decreases.

With the enhancing reaction temperature and elongating reaction time, the UC emission became stronger and stronger (Fig. 8c and d). As mentioned above, the higher reaction temperature and longer reaction time are favorable for the transition from m- NH_4BiF_4 , t-BiOF to h- NaBiF_4 in the molten salt procedure and the proportion of h- NaBiF_4 increasing. Thus, the UC emission intensity of the samples is dramatically improved.

To determine the nature of the UC emission mechanisms responsible, we study the evolution of the intensities as a function of the pumping power for NaBiF_4 :3 mol% $\text{Er}^{3+}/20$ mol% Yb^{3+} sample. The UC emission intensity I should scale as P^n at low power limit (far from the saturation), where n stands for the number of pumping phonons absorbed per UC phonon emitted and P represents the pump power.^{48–51} A plot of $\ln I$ vs. $\ln P$ yields a straight line with the slope n . Fig. 9 shows the plots for

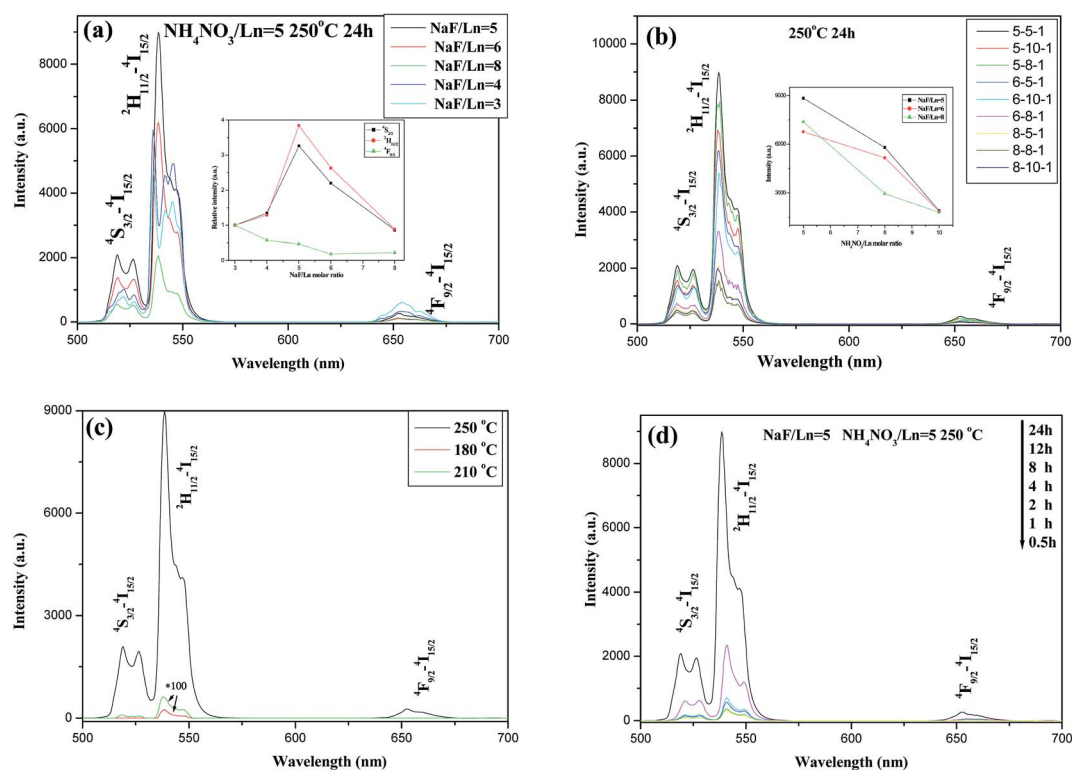


Fig. 8 The UC emission spectra of the h- NaBiF_4 : $\text{Er}^{3+}/\text{Yb}^{3+}$ samples with different conditions. (a) Different NaF/Ln molar ratios; (b) different NH_4NO_3 /Ln molar ratios; (c) different reaction temperatures; (d) different reaction times.



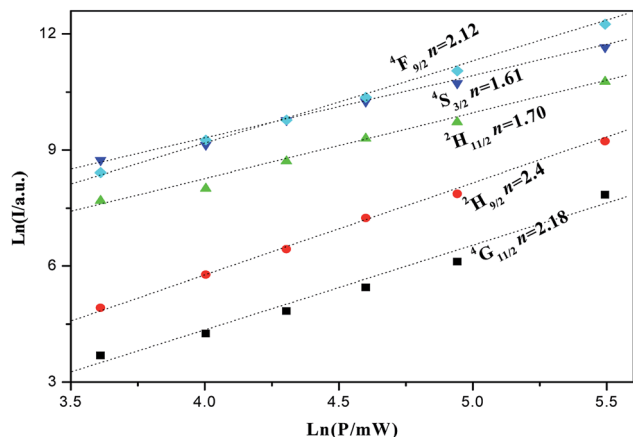


Fig. 9 Logarithmic plot of the integrated UC emission intensity of h-NaBiF₄:3 mol% Er³⁺/20 mol% Yb³⁺ sample at 295 K after 980 nm excitation.

$^4G_{11/2}$, $^2H_{9/2}$, $^4F_{9/2}$, $^2H_{11/2}$, $^4S_{3/2}$, $^4F_{7/2}$ and emission. The slopes for $^2H_{11/2}$ and $^4S_{3/2}$ are 1.70 and 1.61, respectively, which are close to 2, indicating that two-phonon process deal with the UC mechanism responsible for populating these emitting levels, whereas the slope for $^4F_{9/2}$, $^4G_{11/2}$ and $^2H_{9/2}$ is 2.12, 2.4 and 2.18, respectively, implying that it is necessary to feed $^4G_{11/2}$ and $^2H_{9/2}$ level *via* three-phonon process. In addition, the slope for $^2H_{11/2}$ and $^4G_{11/2}$ is relatively larger than that for $^4S_{3/2}$ and $^2H_{9/2}$,

indicating that a faster growth of the former than the latter. This can be resulted from the thermal effect. As the pump power increases to high enough, the strong absorption brings forth the enhancement of the sample temperature at the irradiated region, generating the thermal effect and the fluorescence quenching. The small energy gap of $^2H_{11/2}/^4S_{3/2}$ and $^4G_{11/2}/^2H_{9/2}$ pairs deduces more Er³⁺ ions to feed the $^2H_{11/2}$ and $^4G_{11/2}$ upper states *via* the thermal distribution as the pump power rise gradually. Therefore, the $^2H_{11/2}$ and $^4G_{11/2}$ state has a large slope *n*. In these samples, the predominant green emission in UC emission fluorescence indicates a larger population of the $^2H_{11/2}/^4S_{3/2}$ level and a smaller population of the $^4F_{9/2}$ level. In this case, the $^2H_{11/2}(\text{Er}^{3+1}) + ^4I_{15/2}(\text{Er}^{3+2}) \rightarrow ^4I_{9/2}(\text{Er}^{3+1}) + ^4I_{13/2}(\text{Er}^{3+2})$ cross relaxation between Er³⁺ and the $^2F_{7/2}(\text{Yb}^{3+}) + S_{3/2}(\text{Er}^{3+}) \rightarrow ^2F_{5/2}(\text{Yb}^{3+}) + ^4I_{13/2}(\text{Er}^{3+})$ ^{9-12,48-51} energy back transfer becomes important. As a result, two NIR excitation photons are used to feed the $^4I_{13/2}$ state of Er³⁺. Since $^4F_{9/2}$ level is populated from $^4I_{13/2}$ through energy transfer between Er³⁺ and Yb³⁺. Consequently, the $^4F_{9/2}$ population, and thus red emission becomes a three-photons process.

To investigate UC emission mechanism and study the temperature dependence UC characteristics, the UC emission spectra of NaBiF₄:3 mol% Er³⁺/20 mol% Yb³⁺ at different temperatures under the excitation wavelength of 980 nm were depicted in Fig. 11a and b. The intensity of the UC emission band attributed to the $^4S_{3/2} \rightarrow ^4I_{15/2}$ and $^2H_{9/2} \rightarrow ^4I_{15/2}$ transition of Er³⁺ exhibits the same temperature dependent

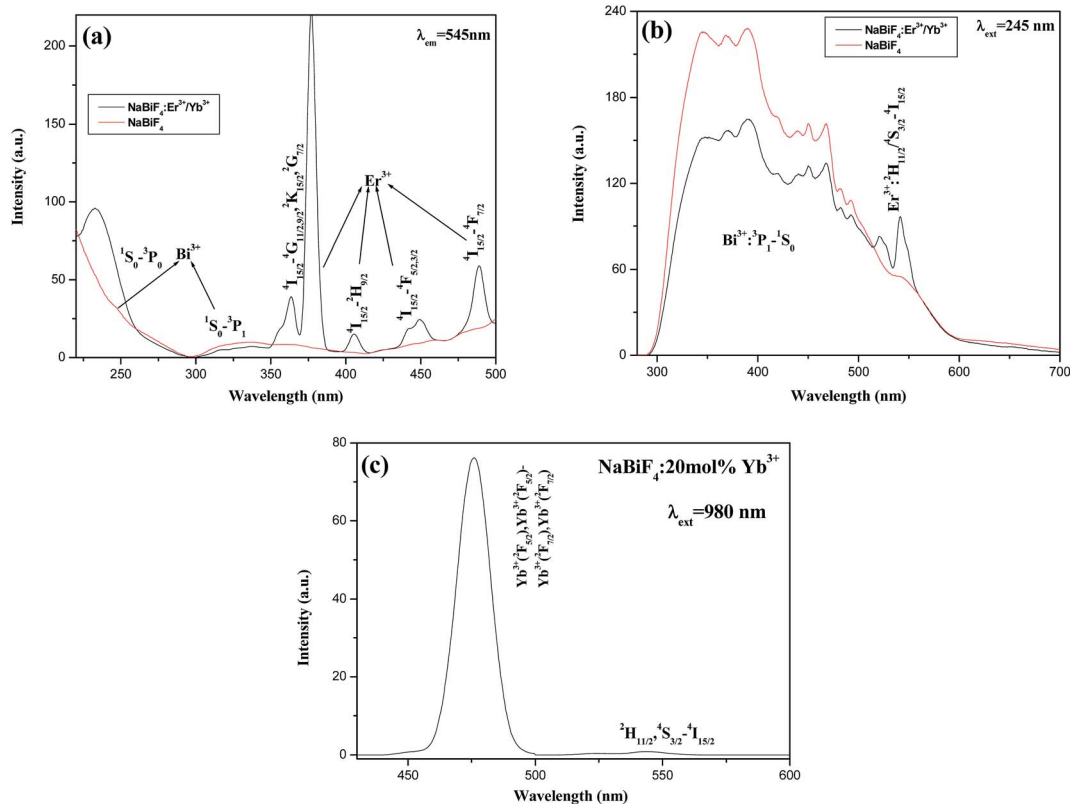


Fig. 10 The excitation spectra (a) and emission spectra (b) of h-NaBiF₄:3 mol% Er³⁺/20 mol% Yb³⁺ and pure h-NaBiF₄; (c) the cooperative UC spectra of h-NaBiF₄:20 mol% Yb³⁺.



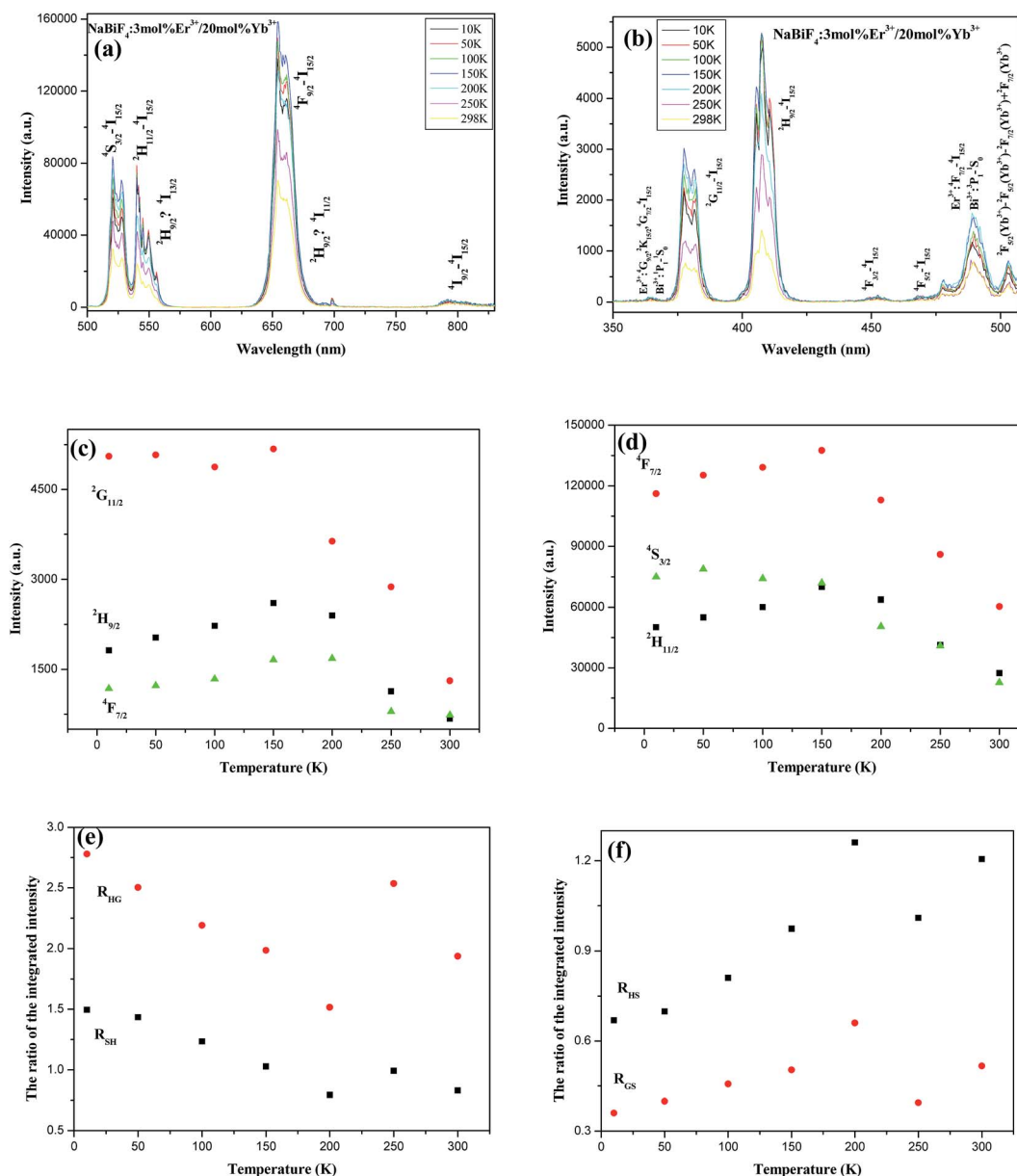


Fig. 11 (a and b) The UC emission spectra of the $\text{NaBiF}_4:3 \text{ mol}\% \text{Er}^{3+}/20 \text{ mol}\% \text{Yb}^{3+}$ sample at different temperatures (the excitation wavelength is 980 nm). (c and d) The UC emission intensity of $\text{NaBiF}_4:3 \text{ mol}\% \text{Er}^{3+}/20 \text{ mol}\% \text{Yb}^{3+}$ sample as a function of the temperature; (e and f) the R_{HG} and R_{GS} of $\text{NaBiF}_4:3 \text{ mol}\% \text{Er}^{3+}/20 \text{ mol}\% \text{Yb}^{3+}$ samples a function of the temperature.

behaviors, which may be resulted from that the population on the ${}^2\text{H}_{9/2}$ state of Er^{3+} can be excited from $\text{S}_{3/2}/{}^2\text{H}_{11/2}$ state *via* the cross relaxation process of ${}^4\text{S}_{3/2}(\text{Er}^{3+1}) + {}^4\text{I}_{13/2}(\text{Er}^{3+2}) \rightarrow {}^2\text{H}_{9/2}(\text{Er}^{3+1}) + {}^4\text{I}_{15/2}(\text{Er}^{3+2})$.^{9–12} They stay at the same level with increasing temperature from 10 to 150 K, then experience a significant downward trend. The other UC emissions exhibit another similar temperature dependent behavior. With decreasing temperature the emission intensities increase initially, reach a maximum at 150 K, and then decrease (Fig. 11c). The reduction in the emission can be resulted from the decrease in resonant transfer because of the thermal depopulation of the higher-energy ${}^4\text{I}_{11/2}$ Er^{3+} levels, which has a resonant match with the ${}^4\text{F}_{7/2}$ level. The decrease of emission intensity can be originated from the temperature dependence of

the multi-phonon de-excitation of the ${}^4\text{S}_{3/2}({}^2\text{H}_{11/2})$ and ${}^4\text{F}_{9/2}$ level. Since it is most likely for a relatively tightly coupled pair to transfer to a single ion is and such pairs have the probability to relax back to ${}^4\text{I}_j$ levels again, the UC fluorescence quenching takes place. The fluorescence quenching becomes more effective than there occurs for the excitation into ${}^4\text{F}_{7/2}$ of Er^{3+} aggravating the fluorescence quenching with increasing temperature. For ${}^4\text{F}_{7/2}$, the emissions originating mainly from relatively isolated ions cause to reduce rate relaxation rate of ion-pair and to decrease the UC emission.^{49–54}

In order to better understand the temperature-dependent behavior, R_{HS} (the ratio of ${}^2\text{H}_{11/2} \rightarrow {}^4\text{I}_{15/2}$ to ${}^4\text{S}_{3/2} \rightarrow {}^4\text{I}_{15/2}$) and R_{GH} (the ratio of ${}^4\text{G}_{11/2} \rightarrow {}^4\text{I}_{15/2}$ to ${}^2\text{H}_{9/2} \rightarrow {}^4\text{I}_{15/2}$) as a function of temperature are shown as Fig. 11d, respectively.



R_{HS} and R_{GH} possess the same temperature dependence characteristics. R_{HS} and R_{GH} firstly increase as the temperature enhances from 10 K to 200 K, then have a downward trend and hit rock bottom when the temperature increases to 250 K, finally recover to have an upward trend. Such phenomenon that the decreasing trend at 200–250 K indicates that there must be some additional UC channel energy transfers are involved with Yb^{3+} cooperative UC followed by the energy transfer to Bi^{3+} , and the nonradiative energy transfer from the excited levels of Er^{3+} to the $^3\text{P}_1$ manifold of Bi^{3+} , followed by energy back energy to another manifolds of Er^{3+} .

To authenticate the existence of energy transfer process between Bi^{3+} and Er^{3+} , the excitation and emission spectra of $\text{NaBiF}_4:3 \text{ mol}\% \text{ Er}^{3+}/20 \text{ mol}\% \text{ Yb}^{3+}$ and NaBiF_4 have been measured and are depicted in Fig. 10. The excitation and emission wavelengths are 245 and 545 nm, respectively. In the excitation spectra (Fig. 10a), there exists broad absorption bands owing to Bi^{3+} ions,^{15–19} while some emission bands due to the transitions of Er^{3+} ions are also observed in the emission spectra (Fig. 10b),^{1–12} implying that the interaction between Bi^{3+} and Er^{3+} takes place. In addition, the anti-Stokes UC fluorescence spectra of $\text{NaBiF}_4:20 \text{ mol}\% \text{ Yb}^{3+}$ samples under 980 nm excitation at room temperature was also shown in Fig. 10c. The observed anti-Stokes in the range of 470–510 nm is believed to be cooperative UC causing by the simultaneous radiative relaxation of the excited $\text{Yb}^{3+}\text{--Yb}^{3+}$ pairs accompanied by the emission of a visible photon in the expression $^2\text{F}_{5/2}(\text{Yb}^{3+})\text{--}^2\text{F}_{5/2}(\text{Yb}^{3+}) \rightarrow ^2\text{F}_{7/2}(\text{Yb}^{3+}) + ^2\text{F}_{7/2}(\text{Yb}^{3+}) + h\nu^{45–47}$ and the emission from the $^3\text{P}_1$ excited state to the $^1\text{S}_0$ ground state of Bi^{3+} .^{14,15,22}

To further validate the rationality of the energy transfer from Bi^{3+} to Er^{3+} , the UC emission spectra of $\text{NaYF}_4:2 \text{ mol}\% \text{ Er}^{3+}/20 \text{ mol}\% \text{ Yb}^{3+}$ at different temperatures and the emission decay curves of $^4\text{S}_{3/2} \rightarrow ^4\text{I}_{15/2}$ transition in $\text{h-NaBiF}_4:3 \text{ mol}\% \text{ Er}^{3+}/20 \text{ mol}\% \text{ Yb}^{3+}$ (the former) and $\text{h-NaYF}_4:2 \text{ mol}\% \text{ Er}^{3+}/20 \text{ mol}\% \text{ Yb}^{3+}$ (the latter) were measured under the excitation wavelength of 980 nm, as shown in Fig. 12. Both h-NaBiF_4 and h-NaYF_4 belong to the hexagonal system with the same space group $P\bar{3}$, and the crystal cell parameter of h-NaBiF_4 ($a = 6.144 \text{ \AA}$, $c = 3.721 \text{ \AA}$)³² is slightly larger than that of h-NaYF_4 ($a = 5.96 \text{ \AA}$, $c = 3.53 \text{ \AA}$).⁵⁵ In the same doping concentration, the mean distance

between Yb^{3+} ions in the former is larger than that in the latter. The cooperative UC intensity depends upon the shortest distances between Yb^{3+} ions, and the smaller the distance between Yb^{3+} ions, the stronger the cooperative UC emission,^{45–47} thereby, the cooperative UC intensity of the former is weaker than that of the latter. As a matter of fact, the UC emission intensity of the former is stronger than that of the latter (see figure). It is suggested that there exists the energy transfer between Er^{3+} ion or Yb^{3+} ion and Bi^{3+} ion in $\text{h-NaBiF}_4:3 \text{ mol}\% \text{ Er}^{3+}/20 \text{ mol}\% \text{ Yb}^{3+}$ sample. In addition, the decay curves of both the samples could be well fitted into a single exponential function (Fig. 12a). It is obvious that the lifetime of the $^4\text{S}_{3/2}$ state in the former (267 μs) is longer than that in the latter (147 μs). Moreover, the former exhibits longer rise and delay times than the latter. This further affirms the existence of energy transfer between Er^{3+} ion and Bi^{3+} ions in $\text{h-NaBiF}_4:3 \text{ mol}\% \text{ Er}^{3+}/20 \text{ mol}\% \text{ Yb}^{3+}$.

Based on the above analysis, the energy diagram showing the proposed UC emission mechanism and energy transfer process in $\text{NaBiF}_4:\text{Er}^{3+}/\text{Yb}^{3+}$ materials was depicted in Fig. 13. The initially increasing Yb^{3+} doping rapidly promotes the rate of energy transfer from Yb^{3+} to Er^{3+} , enhancing the population on the green-emitting participating $^4\text{F}_{7/2}$ and red-emitting participating $^4\text{I}_{11/2}$ manifolds, which results in the improvement of the UC emission. Too high Yb^{3+} doping concentration induces an efficient energy back-transfer process $^4\text{F}_{7/2}(\text{Er}^{3+}) + ^2\text{F}_{5/2}(\text{Yb}^{3+}) \rightarrow ^4\text{I}_{11/2}(\text{Er}^{3+}) + ^2\text{F}_{5/2}(\text{Yb}^{3+})^{9–12}$ and $^4\text{F}_{9/2}(\text{Er}^{3+}) + ^2\text{F}_{5/2}(\text{Yb}^{3+}) \rightarrow ^4\text{I}_{13/2}(\text{Er}^{3+}) + ^2\text{F}_{5/2}(\text{Yb}^{3+})^{9–12}$ causing the decrease of the population on the $^4\text{F}_{7/2}$, $^4\text{I}_{11/2}$, $^2\text{H}_{11/2}/^4\text{S}_{3/2}$ and $^4\text{F}_{9/2}$ manifold, accordingly the green UC emission and red one both experiences a significant downward trend. At low Er^{3+} doping concentration, there are not enough Er^{3+} ions to be populated or excited. The increase of Er^{3+} concentration makes the average distance between Er^{3+} ions shrink and the interaction between ions intensifies. This induces the population on the $^4\text{F}_{7/2}$, $^4\text{I}_{11/2}$, $^2\text{H}_{11/2}$, $^4\text{S}_{3/2}$ and $^4\text{F}_{9/2}$ states increase, and gives rise to the increase of the population on the green and red-emitting manifolds, thus, the red and green emission experiences an upward trend. When the Er^{3+} doping concentration achieves a certain crisis, the average interaction distance between Er^{3+} and Yb^{3+} is beyond the critical distance of the energy back transfer $^4\text{I}_{11/2}(\text{Er}^{3+}) + ^2\text{F}_{7/2}(\text{Yb}^{3+}) \rightarrow ^4\text{I}_{15/2}(\text{Er}^{3+}) +$

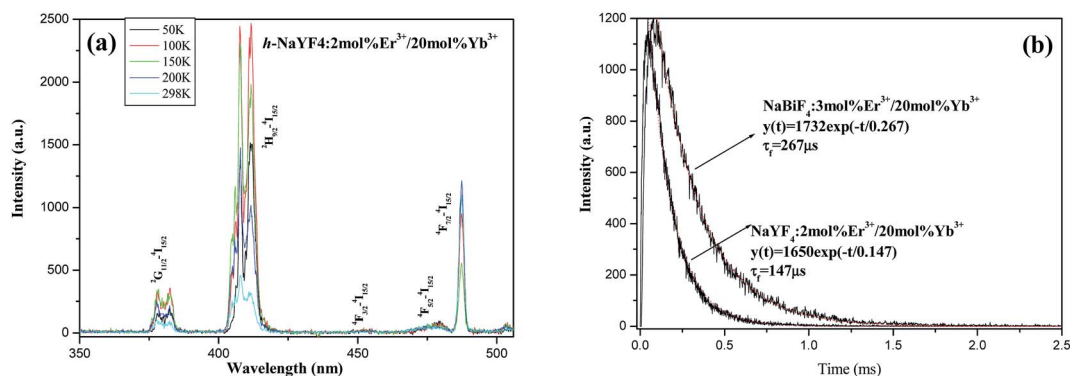


Fig. 12 (a) The UC emission spectra of the $\text{NaYF}_4:2 \text{ mol}\% \text{ Er}^{3+}/20 \text{ mol}\% \text{ Yb}^{3+}$ sample at different temperatures (the excitation wavelength is 980 nm). (b) The decay curves UC emission intensity of 532 nm of $\text{NaBiF}_4:3 \text{ mol}\% \text{ Er}^{3+}/20 \text{ mol}\% \text{ Yb}^{3+}$ and $\text{NaYF}_4:2 \text{ mol}\% \text{ Er}^{3+}/20 \text{ mol}\% \text{ Yb}^{3+}$ samples.



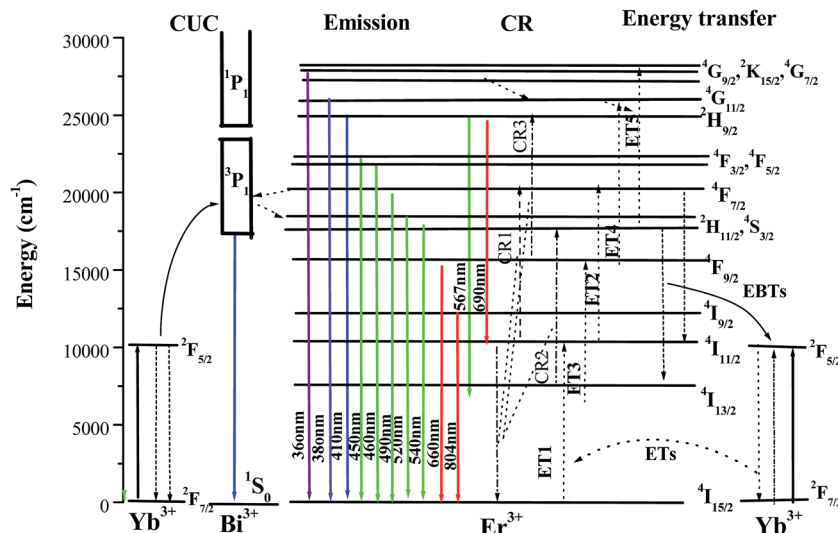


Fig. 13 Schematic energy level diagrams of Bi^{3+} , Er^{3+} and Yb^{3+} ions as well as the proposed energy transfer of $\text{NaBiF}_4:\text{Er}^{3+}/\text{Yb}^{3+}$ under excitation of 980 nm.

$^2\text{F}_{5/2}(\text{Yb}^{3+})$ (EBT),⁹⁻¹² which leads to a considerable drop the population on the red-emitting $^4\text{F}_{9/2}$ manifold and the green-emitting $^2\text{H}_{11/2}/^4\text{S}_{3/2}$ and violet emitting $^4\text{G}_{11/2}$ manifold. As a consequence, the intensity of the visible emission abruptly falls off, while the green one quickly rises. Obviously, 4 mol% exceeds the critical concentration, whereas 3 mol% is below this threshold.

4. Conclusions

The hexagonal $\text{NaBiF}_4:\text{Er}^{3+}/\text{Yb}^{3+}$ micro-/nanocrystals were synthesized *via* low-temperature molten-salt method in NH_4NO_3 flux. The influences of the reaction time, the reaction temperature, the content of NaF and NH_4NO_3 on the crystal phase of the resulting products were studied at length. This research also investigates and discusses in detail the effects of shapes, sizes, doping concentration of Er^{3+} and Yb^{3+} , the pumping power and the temperature on the UC emission of $\text{NaBiF}_4:\text{Er}^{3+}/\text{Yb}^{3+}$ micro-/nanocrystals. The results show that the $\text{NaBiF}_4:3 \text{ mol}\% \text{ Er}^{3+}/20 \text{ mol}\% \text{ Yb}^{3+}$ sample exhibits the strongest UC emission and can be comparable to the most efficient UC hexagonal $\text{NaYF}_4:\text{Er}^{3+}/\text{Yb}^{3+}$ micro-/nanocrystals. Two-phonon process takes responsibility to populate $^4\text{S}_{3/2}/^2\text{H}_{11/2}$ of Er^{3+} , whilst the $^4\text{F}_{9/2}$ is involved in a three-photon process. The energy transfer between the excited levels of Er^{3+} and the $^3\text{P}_1$ manifold of Bi^{3+} and cooperative UC of Yb^{3+} are favorable to improve upconversion emission of Er^{3+} . This work is significant not only to provide one novel synthesis method of Bi-based fluoride UC materials but also for the better understanding of their essential UC emission mechanisms (particularly the energy transfer from Bi^{3+} to Er^{3+}).

Conflicts of interest

There are no conflicts to declare.

Acknowledgements

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