

Cite this: *RSC Adv.*, 2017, 7, 34574Received 8th May 2017  
Accepted 30th June 2017

DOI: 10.1039/c7ra05196d

rsc.li/rsc-advances

## Versatile magnetic resonance singlet tags compatible with biological conditions†

Stefan Glöggler,<sup>ID</sup>\* Stuart J. Elliott, Gabriele Stevanato, Richard C. D. Brown and Malcolm H. Levitt<sup>ID</sup>

The long lifetime of nuclear singlet states holds promise for the development of molecular tracers to study motional processes in proteins with increased precision or to act as imaging contrast agents. We introduce a singlet tag (STAG) based on bromoacrylate that is readily attached to a variety of biomolecules, compatible with biological conditions, and is relatively insensitive to the presence of molecular oxygen. The generality of this approach is demonstrated by tagging building blocks for biomolecules using a simple labelling route based on commercially available starting materials. Additionally, we demonstrate a singlet NMR experiment on the endogenous molecule phosphoenolpyruvate.

### Introduction

Nuclear singlet states are NMR “silent” states that do not give rise directly to NMR signals but which may be populated and detected indirectly.<sup>1–30</sup> Such a state is created *e.g.* if the spins of two  $I = 1/2$  nuclei couple antiparallel to each other to form a total spin of 0. The singlet spin state is immune to the motional modulation of the in pair dipole–dipole interaction, with other relaxation mechanisms strongly attenuated. The time constant ( $T_S$ ) for the equilibration of the singlet state population with the triplet populations can exceed the longitudinal relaxation time ( $T_1$ ) by more than one order of magnitude, with a  $T_S$  value exceeding 1 hour being observed in one case.<sup>13</sup> Various applications of singlet states have been proposed, including the utilization of long  $T_S$  times to probe slow dynamic processes in biomolecules and the binding of drugs,<sup>6,10,15–17</sup> and as a means to support nuclear hyperpolarization in the context of molecular imaging.<sup>18,21,23,28,31–43</sup> In some cases, the protons of biomolecules support nuclear singlet states. Examples include molecules relevant to neuroscience such as aspartate and glutamine, peptides, proteins and adenosinediphosphate (ADP).<sup>15–17,28,29</sup> Proton singlet states with lifetimes  $T_S$  up to 4.5 s have been reported under biologically compatible conditions (pH 7.0–7.4 and 25–37 °C) even in the presence of dissolved molecular oxygen and without deuteration of the solvent.<sup>28</sup>

A more versatile approach is offered by artificial prosthetic groups supporting long-lived nuclear singlet states. In this paper we call such entities singlet tags (STAGs). Previously described examples include bromothiophene and fluorinated phenylacetyl

moieties linked to tripeptides.<sup>28,29</sup> The measured  $T_S$  values were 11.7 s in D<sub>2</sub>O for the bromothiophene tag, and 2.6 s for the fluorinated phenylacetyl tag.<sup>28,29</sup> However, in order to be applicable to *in vitro* and *in vivo* experiments, the singlet NMR phenomena should be insensitive to oxygen and functional in non-deuterated solvents (especially H<sub>2</sub>O rather than D<sub>2</sub>O, which is toxic for cell experiments and may change protein conformations). Ideally, the singlet states should be sustainable at high magnetic fields, and their <sup>1</sup>H NMR resonances should be in a spectral region that does not overlap with the signal from water (4.7 ppm in a proton spectrum), signals from amino acids in proteins (aliphatic and aromatic functionalities, ranging from 0–5 ppm and 7–9 ppm in a proton spectrum) and typical signals from metabolites (0–5.3 ppm and 6.8–8.1 ppm in a proton spectrum).<sup>44,45</sup>

In this article we explore the use of substituted acrylates as STAGs. The acrylate moiety ( $\text{CH}_2=\text{C}(\text{CO}_2^-)$ ) contains a pair of protons with similar chemical shifts and which are very close in space (internuclear distance  $\sim 171.2$  pm).<sup>46</sup> The acrylate proton signals appear in a favourable chemical shift region around 6–7 ppm which does not typically overlap with other common proton signals. Acrylates have found a wide range of applications and are mostly used as monomers for polymerization reactions.<sup>47</sup> Other important fields are their utilization as Michael acceptors, with particular use in click chemistry reactions<sup>48,49</sup> and as precursors for imaging contrast agents.<sup>31–38,50–52</sup> In the present case, the close proximity of the acrylate protons reduces the influence of relaxation agents such as paramagnetic species on nuclear singlet order, except at very short range. The relatively small shift difference of about 0.6 ppm between acrylate protons makes it feasible to spin-lock singlet states at high magnetic field by applying on resonant radiofrequency irradiation.<sup>4</sup> We demonstrate a general synthetic route for bromoacrylate STAGs which utilizes inexpensive and commercially available reactants. The route allowed us to label several amino acids as well as

School of Chemistry, University of Southampton, Southampton SO171BJ, UK. E-mail: stefan.gloeggler@mpibpc.mpg.de

† Electronic supplementary information (ESI) available: Chemical synthesis, relaxation analysis, details of the NMR sequence timings. See DOI: 10.1039/c7ra05196d



a glucose derivative. We show that the bromoacrylate singlet lifetime in all of these compounds significantly exceeds  $T_1$  under conditions suitable for biological applications, *i.e.* in an aqueous phosphate buffer in the presence of molecular oxygen and without solvent deuteration. Furthermore, we have identified a metabolite, phosphoenolpyruvate, which contains an acrylate functionality and which supports a nuclear singlet state with a decay time constant  $T_S$  more than four times longer than  $T_1$ . Phosphoenolpyruvate becomes converted into pyruvate during glycolysis which converts into lactate in cancer cells.<sup>53</sup>

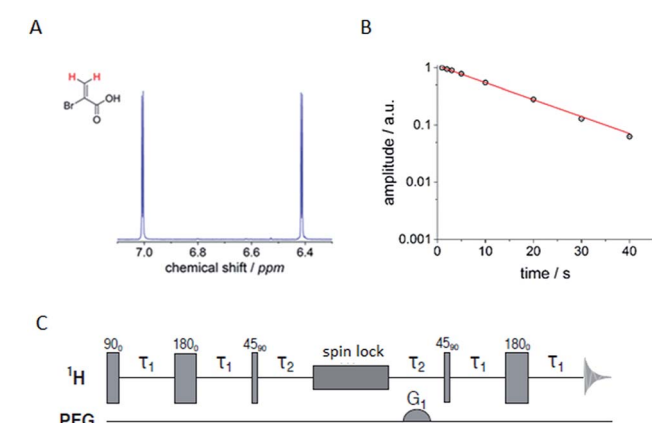
## Results and discussion

The  $^1\text{H}$  NMR spectrum of bromoacrylic acid (BrAc, Fig. 1A), shows a distinct AX pattern in the range of 6–7 ppm. The shift

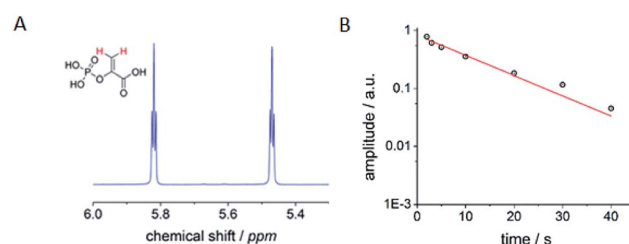
difference between the acrylate protons is 0.6 ppm in  $\text{H}_2\text{O}$ . The measured values of  $T_1$  and  $T_S$  for BrAc solutions at two different magnetic fields and a selection of solvent preparations are shown in Table 1. The singlet relaxation times  $T_S$  were estimated by using the pulse sequence of Sarkar *et al.* [ref. 6] as shown in Fig. 1C (see Table 1 of the ESI† for pulse sequence timings). Experiments exploited singlet spin-locking by applying an on resonant radiofrequency field with a nutation frequency of 3 kHz (at magnetic fields of 11.7 T and 22.2 T). Note that the values of  $T_S$  are approximately four times larger than the values of  $T_1$  under the same conditions. The weak dependence of the relaxation times on the magnetic field strength indicate that chemical shift anisotropy is not a highly significant relaxation mechanism for this system.

Phosphoenolpyruvate (PEP, Fig. 2), is a metabolite of the glycolysis pathway which contains an acrylic moiety. Since glycolysis is up-regulated in a variety of cancer cells, PEP (as well as glucose derivatised STAGs) might serve as molecular imaging agents for cancer targeting. The  $^1\text{H}$  NMR spectrum of PEP is displayed in Fig. 2A. The shift difference between the acrylate protons is 0.4 ppm in  $\text{H}_2\text{O}$ . The spectrum shows the characteristic  $J$ -splitting of the acrylic protons which is further split by the  $^4J$ -coupling of the protons to the  $^{31}\text{P}$  phosphorus nucleus (100% natural abundance, spin  $I = 1/2$ ).

The measured values of  $T_1$  and  $T_S$  for PEP solutions at two different magnetic fields in PBS (phosphate buffered saline, 360 mg  $\text{NaH}_2\text{PO}_4$  and 266 mg  $\text{Na}_2\text{HPO}_4$  in 300 mL ultrapure  $\text{H}_2\text{O}$ ) buffer are shown in Table 1. The nutation frequencies of



**Fig. 1** Investigation of the BrAc-STAG. (A)  $^1\text{H}$  NMR spectrum of bromoacrylic acid (BrAc) in  $\text{H}_2\text{O}$  (PBS) at 11.7 T and 25 °C. The spectrum displays two doublets (AX-pattern) of the indicated protons in red, which are used to support singlet order. The NMR signal is in a region downfield of typical aliphatic protons and upfield of most aromatic protons found in metabolites. (B) Decay of proton singlet order in BrAc, observed using the pulse sequence in (C) in the presence of a spin-locking field (nutation frequency = 3 kHz). The observed singlet decay time constant  $T_S$  is 5.7 times longer than  $T_1$  ( $T_S = 14.7 \pm 0.2$  s,  $T_1 = 2.8 \pm 0.1$  s). (C) Pulse sequence to generate singlet order, spin-lock it, and back-convert it into observable magnetization with subsequent acquisition of the signal [ref. 6 and 16]. A pulse field gradient is utilized as a filter to remove coherences that are not originating from the singlet state after back-conversion into observable magnetization. The timings  $\tau_1$  and  $\tau_2$  depend on the  $J$ -couplings and chemical shift difference of the protons, a table of which is displayed in the ESI Table 1†.

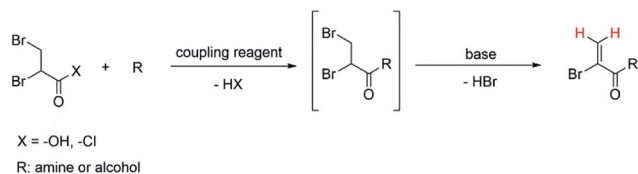


**Fig. 2** Investigation of PEP. (A)  $^1\text{H}$  NMR spectrum of the two indicated protons in phosphoenol pyruvate in  $\text{H}_2\text{O}$  (PBS) at a field of 11.7 T. The triplet-like multiplet structure is due to a superposition of the  $^2J$  coupling between the protons with an additional  $^4J$  coupling to the  $^{31}\text{P}$  nucleus ( $I = 1/2$ , natural abundance of close to 100%). (B) Decay of singlet order in PEP (the data points are within the standard error of 12%). The observed value of  $T_S$  is 3.6 longer than  $T_1$ .

**Table 1** Experimental summary of bromoacrylic acid (BrAc) and phosphoenol pyruvate (PEP) over a wide range of magnetic fields, solvents and oxygenation conditions. A spin-locking field with 3 kHz nutation frequency was applied during  $T_S$ -measurements

Molecule	Magnetic field	Solvent	Oxygen condition	$T_1/\text{s}$	$T_S/\text{s}$
BrAc	11.7 T	$\text{D}_2\text{O}$	Degassed	$3.0 \pm 0.1$	$44.6 \pm 1.0$
BrAc	11.7 T	$\text{D}_2\text{O}$	Non-degassed	$3.3 \pm 0.1$ s	$22.8 \pm 2.0$
BrAc	11.7 T	$\text{H}_2\text{O}$ (PBS)	Non-degassed	$2.8 \pm 0.1$	$14.7 \pm 0.2$
BrAc	22.2 T	$\text{H}_2\text{O}$ (PBS)	Non-degassed	$2.9 \pm 0.1$	$14.1 \pm 0.5$
PEP	11.7 T	$\text{H}_2\text{O}$ (PBS)	Non-degassed	$1.4 \pm 0.1$	$12.5 \pm 1.5$
PEP	22.2 T	$\text{H}_2\text{O}$ (PBS)	Non-degassed	$1.6 \pm 0.1$	$12.1 \pm 0.1$





**Fig. 3** General synthetic route for labelling molecules with BrAc-STAGs. For the labelling of alcohols or amines, 2,3-dibromopropionic acid or its acid chloride are utilized as a reagent. During the reaction a dibromo-intermediate is formed from which hydrobromic acid is eliminated upon reaction with a base. The intermediate is typically not isolated since the reactions are performed in the presence of bases (4-dimethylaminopyridine, triethylamine and diisopropylethylamine) and thus yield the final acrylate in a one-pot procedure.

**Table 2** Tagged molecules

Compound	Structure	Residue
1		Free acid
2		Ethyl
3		Propargyl
4		Serine
5		Alanine
6		Glucose
7		Phosphoenol pyruvate

the spin-locking fields for the  $T_S$  measurements were 3 kHz at magnetic fields of 11.7 T and 22.2 T. The results are similar to those obtained for BrAc. The dependence of the relaxation times on magnetic field strength is again found to be weak, indicating that chemical shift anisotropy is not a strong relaxation mechanism for this system. We developed a general and robust route for coupling the bromoacrylate singlet tag to a variety of molecules. Since the hydrophilicity of bromoacrylic acid itself causes handling difficulties and promotes unwanted side reactions, we developed an alternative procedure (Fig. 3) starting from 2,3-dibromopropionic acid (or its acid chloride), which is subsequently converted into BrAc upon the addition of bases.<sup>51,54</sup> 2,3-dibromopropionic acid or its acid chloride underwent reactions with amines or alcohols, which are functional moieties found in a variety of metabolites, amino acids or peptides. In the presence of coupling reagents and a base we postulate that an intermediate product is formed based on the 2,3-dibromopropionic acid. However, the base slowly eliminates HBr from the intermediate structure so that in most cases the acrylic acid derivative is yielded in a one-pot process. Specific bases that we have utilized are 4-dimethylaminopyridine (DMAP), triethylamine (TEA) and diisopropylethylamine (DIPEA). The tagged molecules that we have synthesized and investigated are esters or amides, based on the following compounds: ethyl alcohol, propargyl alcohol, serine, alanine and acetylglucosylamine. To verify that different sites in amino acids can be tagged, alanine was labelled at the *N*-terminus and serine *via* its alcohol moiety. The labelling of propargylalcohol was chosen since the BrAc-STAG carrying an alkyne moiety from propargylalcohol may serve as a precursor for click reactions.<sup>55</sup> The tagged structures are presented in Table 2. Singlet lifetimes in non-degassed PBS (unless indicated otherwise) are shown in Table 3. All of the demonstrated molecules show a significantly longer  $T_S$  than  $T_1$ , proving that singlet states with a significant larger  $T_S/T_1$  ratio can be maintained in conditions that are compatible with biological investigations. Moreover, the presented STAGs may be utilized to label a variety of biomolecules or derivatives thereof and maintain their longevity. The STAGs described here require the application of a relatively strong radiofrequency field to quench singlet-triplet transitions. Applications which exploit the long singlet lifetime of the spin-locked singlet order may therefore be restricted to *in vitro* applications. To further improve the STAGs we intend to reduce the chemical shift

**Table 3** Relaxation times for tagged systems under non-degassed conditions at 25 °C. Singlet lifetimes were measured by using the pulse sequence described in Fig. 1C. Longitudinal relaxation was measured by using the inversion recovery experiment. A spin-locking field of 3 kHz was applied for singlet relaxation experiments

Compound	$T_1/s$	$T_S/s$	$T_S/T_1$	$B_0/T$	$c/mM$	Spin-lock/kHz	Solvent
1	$2.8 \pm 0.1$	$14.7 \pm 0.2$	$5.3 \pm 0.2$	11.7	10	3	PBS
2	$2.4 \pm 0.1$	$9.9 \pm 0.3$	$4.1 \pm 0.3$	11.7	10	3	PBS
3	$2.4 \pm 0.1$	$13 \pm 2$	$5 \pm 2$	11.7	10	3	PBS
4	$1.4 \pm 0.1$	$13.5 \pm 0.5$	$9.6 \pm 0.5$	11.7	10	3	PBS
5	$1.5 \pm 0.1$	$11.9 \pm 0.2$	$7.9 \pm 0.2$	11.7	10	3	PBS
6 <sup>a</sup>	$1.3 \pm 0.1$	$8.3 \pm 0.5^a$	$6.4 \pm 0.5$	11.7	10	3	MeOH
7	$1.4 \pm 0.1$	$13 \pm 2$	$9 \pm 2$	11.7	10	3	PBS

<sup>a</sup> Measured in MeOH due to limited solubility in PBS buffer.



difference of the protons, allowing operation with a very small or absent spin-locking field, leading to applicability *in vivo* without heating effects.

## Conclusions

We have identified haloacrylate STAGs that can be used as molecular labels for singlet NMR experiments in a variety of biological relevant molecules and which can be used under conditions relevant for biological investigations. The bromoacrylate STAG may be introduced by an efficient one-pot synthetic procedure that is readily performed with commercially available products. The metabolite phosphoenolpyruvate (PEP) contains an acrylate entity which also supports long-lived nuclear singlet order. This is a step towards metabolic imaging of PEP in living systems. Acrylate tags may also have applications in areas outside singlet NMR. Possible applications include their role as precursors for imaging contrast agents, including PET-tracers and hyperpolarized MRI tracers upon the hydrogenation with tritium or para-hydrogen respectively.<sup>31–38,50,51</sup> The BrAc tag molecules may serve as new building blocks for Michael additions or thio-bromo reactions in click chemistry.<sup>49,56</sup> Furthermore, we envision utilization of the haloacrylates as building blocks for polymerization reactions.

## Experimental details

NMR experiments were conducted on Bruker Avance III 500 or 950 MHz spectrometers with a Bruker TBO probe or a Bruker TCI cryo probe running TopSpin 3.2. software. Substrates 1–5 and 7 were dissolved in PBS buffer (360 mg NaH<sub>2</sub>PO<sub>4</sub> and 266 mg Na<sub>2</sub>HPO<sub>4</sub> in 300 mL ultrapure H<sub>2</sub>O) at pH = 7.0 at a concentration of 10 mM. 6 was dissolved in methanol (CH<sub>3</sub>OH). The pulse sequence displayed in Fig. 1C was used to perform singlet experiments for which the coupling parameters and respective timings are listed in the ESI.† The pulsed field gradient in the sequence had a sine shape with 5 Gcm<sup>−1</sup> maximal amplitude and a duration of 3 ms. During the singlet sustaining period a continuous wave decoupling pulse with 3 kHz amplitude was applied, centred in between the two protons resonances of interest. A recycle delay of 80 s was used in protonated solvents and 300 s in D<sub>2</sub>O after sample degassing. 4 averages were acquired for each data point and experiments were repeated 6 times. Longitudinal relaxation times were estimated by the inversion recovery pulse sequence with 4 averages whereby each experiment was repeated 6 times. The recycle delay was set to 20 s.

## Acknowledgements

This research was supported by the Engineering and Physical Sciences Research Council (UK), grant codes EP/N002482 and EP/L505067/1, the Marie-Sklodowska-Curie program of the European Community (655864-SingMet), the Wolfson Foundation and Bruker Biospin (UK). We would like to acknowledge helpful support and discussion from L. J. Brown and G. Pileio. We would also like to acknowledge C. Griesinger and M.

Zweckstetter (Max-Planck-Institute for Biophysical Chemistry and Deutsches Zentrum für Neurodegenerative Erkrankungen) for 950 MHz spectrometer access.

## Notes and references

- 1 M. Carravetta and M. H. Levitt, *J. Am. Chem. Soc.*, 2004, **126**, 6228–6229.
- 2 M. Carravetta, O. G. Johannessen and M. H. Levitt, *Phys. Rev. Lett.*, 2004, **92**, 153003.
- 3 M. Carravetta and M. H. Levitt, *J. Chem. Phys.*, 2005, **122**, 214505.
- 4 G. Pileio and M. H. Levitt, *J. Chem. Phys.*, 2009, **130**, 214501.
- 5 G. Pileio, M. Carravetta and M. H. Levitt, *Proc. Natl. Acad. Sci. U. S. A.*, 2011, **107**, 17135–17139.
- 6 R. Sarkar, P. R. Vasos and G. Bodenhausen, *J. Am. Chem. Soc.*, 2007, **129**, 328–334.
- 7 A. S. Kiryutin, A. N. Pravdivtsev, A. V. Yurkovskaya, H.-M. Vieth and K. L. Ivanov, *J. Phys. Chem. B*, 2016, **120**, 11978–11986.
- 8 G. Pileio, M. Carravetta, E. Hughes and M. H. Levitt, *J. Am. Chem. Soc.*, 2008, **130**, 12582–12583.
- 9 M. C. D. Tayler and M. H. Levitt, *Phys. Chem. Chem. Phys.*, 2011, **13**, 5556–5560.
- 10 P. Ahuja, R. Sarkar, P. R. Vasos and G. Bodenhausen, *J. Chem. Phys.*, 2007, **127**, 134112.
- 11 G. Pileio, M. Concistre, M. Carravetta and M. H. Levitt, *J. Magn. Reson.*, 2006, **182**, 353–357.
- 12 T. Theis, Y. Feng, T. Wu and W. S. Warren, *J. Chem. Phys.*, 2014, **140**, 014201.
- 13 G. Stevanato, J. T. Hill-Cousins, P. Håkansson, S. S. Roy, L. J. Brown, R. C. D. Brown, G. Pileio and M. H. Levitt, *Angew. Chem., Int. Ed.*, 2015, **54**, 3740–3743.
- 14 S. J. Elliott, L. J. Brown, J.-N. Dumez and M. H. Levitt, *Phys. Chem. Chem. Phys.*, 2016, **18**, 17965–17972.
- 15 R. Buratto, D. Mammoli, E. Chiarparin, G. Williams and G. Bodenhausen, *Angew. Chem., Int. Ed.*, 2014, **53**, 11376–11380.
- 16 N. Salvi, R. Buratto, A. Bornet, S. Ulzega, I. R. Rebollo, A. Angelini, C. Heinis and G. Bodenhausen, *J. Am. Chem. Soc.*, 2012, **134**, 11076–11079.
- 17 R. Buratto, A. Bornet, J. Milani, D. Mammoli, B. Vuichoud, N. Salvi, M. Singh, A. Laguerre, S. Passemard, S. Gerber-Lemaire, S. Jannin and G. Bodenhausen, *ChemMedChem*, 2014, **9**, 2509–2515.
- 18 Y. Feng, R. M. Davis and W. S. Warren, *Nat. Phys.*, 2012, **8**, 831–837.
- 19 A. K. Grant and E. Vinogradov, *J. Magn. Reson.*, 2008, **194**, 46–57.
- 20 W. S. Warren, E. Jensita, R. T. Branca and X. Chen, *Science*, 2009, **323**, 1711–1714.
- 21 G. Pileio, S. Bowen, C. Laustsen, M. C. D. Tayler, J. T. Hill-Cousins, L. J. Brown, R. C. D. Brown, J. H. Ardenkjaer-Larsen and M. H. Levitt, *J. Am. Chem. Soc.*, 2013, **135**, 5084–5088.
- 22 Y. Feng, T. Theis, X. Liang, Q. Wang and W. S. Warren, *J. Am. Chem. Soc.*, 2013, **135**, 9632–9635.





- 23 T. Theis, G. X. Ortiz Jr, A. W. J. Logan, K. E. Claytor, Y. Feng, W. P. Huhn, V. Blum, S. J. Malcolmson, E. Y. Chekmenev, Q. Wang and W. S. Warren, *Sci. Adv.*, 2016, **2**, e1501438.
- 24 Y. Zhang, X. Duan, P. C. Soon, V. Sychrovsky, J. W. Canary and A. Jerschow, *ChemPhysChem*, 2016, **17**, 2967–2971.
- 25 Y. Zhang, P. C. Soon, A. Jerschow and J. W. Canary, *Angew. Chem., Int. Ed.*, 2014, **53**, 3396–3399.
- 26 S. Cavadini, J. Dittmer, S. Antonijevi and G. Bodenhausen, *J. Am. Chem. Soc.*, 2005, **127**, 15744–15748.
- 27 P. Ahuja, R. Sarkar, P. R. Vasos and G. Bodenhausen, *J. Am. Chem. Soc.*, 2009, **131**, 7498–7499.
- 28 S. J. DeVience, R. Walsworth and M. S. Rosen, *NMR Biomed.*, 2013, **26**, 1204–1212.
- 29 R. Buratto, D. Mammoli, E. Canet and G. Bodenhausen, *J. Med. Chem.*, 2016, **59**, 1960–1966.
- 30 M. C. D. Tayler and M. H. Levitt, *Phys. Chem. Chem. Phys.*, 2011, **13**, 9128–9130.
- 31 M. Goldman, H. Johannesson, O. Axelsson and M. Karlsson, *C. R. Chim.*, 2006, **9**, 357–363.
- 32 J.-B. Hövener, E. Y. Chekmenev, K. C. Harris, W. H. Perman, T. T. Tran, B. D. Ross and P. Bhattacharya, *Magn. Reson. Mater. Phys., Biol. Med.*, 2009, **22**, 123–134.
- 33 K. W. Waddell, A. M. Coffey and E. Y. Chekmenev, *J. Am. Chem. Soc.*, 2011, **133**, 97–101.
- 34 R. Reineri, T. Boi and S. Aime, *Nat. Commun.*, 2015, **6**, 5858.
- 35 K. Münnemann, M. Kölzer, I. Blakey, A. K. Whittaker and K. J. Thurecht, *Chem. Commun.*, 2012, **46**, 1583–1585.
- 36 J. A. Tang, F. Gruppi, R. Fleysheer, D. K. Sodickson, J. W. Canary and A. Jerschow, *Chem. Commun.*, 2011, **47**, 958–960.
- 37 S. Glöggler, S. Wagner and L.-S. Bouchard, *Chem. Sci.*, 2015, **6**, 4261–4266.
- 38 T. Gutmann, M. Sellin, H. Breitzke, A. Stark and G. Buntkowsky, *Phys. Chem. Chem. Phys.*, 2009, **11**, 9170–9175.
- 39 J. H. Ardenkjaer-Larsen, B. Fridlund, A. Gram, G. Hansson, L. Hansson, M. H. Lerche, R. Servin, M. Thaning and K. Golman, *Proc. Natl. Acad. Sci. U. S. A.*, 2003, **100**, 10158–10163.
- 40 K. Golman, R. in 't Zandt and M. Thaning, *Proc. Natl. Acad. Sci. U. S. A.*, 2006, **103**, 11270–11275.
- 41 S. J. Nelson, J. Kurhanewicz, D. B. Vigneron, P. E. Z. Larson, A. L. Harzstark, M. Ferrone, M. can Crieckinge, J. W. Chang, R. Bok, I. Park, G. Reed, L. Carvajal, E. J. Small, P. Munster, V. K. Weinberg, J. H. Ardenkjaer-Larsen, A. P. Chen, R. E. Hurd, L.-I. Odegardstuen, F. J. Robb, J. Tropp and J. A. Murray, *Sci. Transl. Med.*, 2013, **5**, 198ra108.
- 42 J. Kurhanewicz, D. B. Vigneron, K. Brindle, E. Y. Chekmenev, A. Comment, C. H. Cunningham, R. J. DeBerardinis, G. G. Green, M. O. Leach, S. S. Rajan, R. R. Rizi, B. D. Ross, W. S. Warren and C. R. Malloy, *Neoplasia*, 2011, **13**, 81–97.
- 43 A. Comment, *J. Magn. Reson.*, 2016, **264**, 39–48.
- 44 D. S. Wishart, C. G. Bigam, A. Holm, R. S. Hodges and B. D. Sykes, *J. Biomol. NMR*, 1995, **5**, 67–81.
- 45 V. Govindaraju, K. Young and A. A. Maudsley, *NMR Biomed.*, 2000, **13**, 129–153.
- 46 M. A. Higgs and R. L. Sass, *Acta Crystallogr.*, 1963, **16**, 657–661.
- 47 J. M. G. Cowie and V. Arrighi, *Polymers: Chemistry and Physics of Modern Materials*, 3rd edn, CRC Press, Taylor and Francis Group, Boca Raton, FL, 2007.
- 48 D. P. Nair, M. Podgorski, S. Chatani, T. Gong, W. Xi, C. R. Fenoli and C. N. Bowman, *Chem. Mater.*, 2014, **26**, 724–744.
- 49 C. E. Hoyle, A. B. Lowe and C. N. Bowman, *Chem. Soc. Rev.*, 2010, **39**, 1355–1387.
- 50 M. M. Herth, M. Barz, D. Moderegger, M. Allmeroth, M. Jahn, O. Thews, R. Zentel and F. Rösch, *Biomacromolecules*, 2009, **10**, 1697–1703.
- 51 S. W. Rhee and M. J. Tanga, *J. Labelled Compd. Radiopharm.*, 2000, **43**, 925–932.
- 52 J. L. Turner, D. Pan, R. Plummer, Z. Chen, A. K. Whittaker and K. L. Wooley, *Adv. Funct. Mater.*, 2005, **15**, 1248–1254.
- 53 R. A. Cairns, I. S. Harris and T. W. Mak, *Nat. Rev. Cancer*, 2011, **11**, 85–95.
- 54 V. Pace, L. Castoldi, A. R. Alcantara and W. Holzer, *Green Chem.*, 2012, **14**, 1859–1863.
- 55 G. C. Tron, T. Pirali, R. A. Billington, P. L. Canonico, G. Sorba and A. A. Genazzani, *Med. Res. Rev.*, 2008, **28**, 278–308.
- 56 B. M. Rosen, G. Lligadas, C. Hahn and V. Percec, *J. Polym. Sci., Part A: Polym. Chem.*, 2009, **47**, 3921–3939.

