

Cite this: *RSC Adv.*, 2017, 7, 32120

Photocatalytic degradation of imidazolium ionic liquids using dye sensitized TiO₂/SiO₂ composites†

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Ionic liquids (ILs) are widely applied in diverse fields, however, ILs bring considerable challenges to the ecosystem when exposed to the environment due to their cytotoxicity and high chemical stability. It is thus increasingly important to investigate measures for the degradation of IL wastes in industrial processes. This paper presents the preparation of dye-sensitized photocatalysts (DCQ-TiO₂/SiO₂) and their applications in the degradation of 4 imidazolium ILs (1-butyl-3-methylimidazolium bromide, [BMIM]Br; 1-butyl-3-methylimidazolium tetrafluoroborate, [BMIM]BF₄; 1-butyl-3-methylimidazolium hexafluorophosphate, [BMIM]PF₆; 1-butyl-3-methylimidazolium bis(trifluoromethylsulfonyl)imide, [BMIM]NTf₂). The photocatalysts are prepared through *in situ* incorporation of TiO₂ into silica matrices and sensitization with 2,9-dichloroquinacridone (DCQ). The photocatalysts are then characterized with N₂ adsorption-desorption isotherm measurements, transmission electron microscopy (TEM), X-ray diffraction (XRD), Raman spectroscopy, Fourier transform infrared spectroscopy (FT-IR), elemental analysis and ultraviolet-visible diffuse reflectance spectroscopy (UV-vis DRS). The results show that these photocatalysts exhibit high catalytic activities when they are applied in the degradation of imidazolium ILs. The degradation efficiency for [BMIM]Br can reach up to 95% under simulated sunlight irradiation in 180 minutes. The photodegradation intermediates of [BMIM]⁺ are identified as harmless and easily biodegradable moieties.

Received 2nd May 2017
Accepted 12th June 2017

DOI: 10.1039/c7ra04939k

rsc.li/rsc-advances

1. Introduction

In the past two decades, ionic liquids (ILs) have attracted considerable interest in the diverse fields of synthesis, catalysis, materials science, separation technology,¹ and CO₂ capture.^{2,3} The increasing research interest is due to the following unique physicochemical characteristics of ILs: tuneable structures and properties, negligible vapor pressure, excellent chemical and thermal stability⁴ as well as high polarity and ionic conductivity.⁵ However, the high thermal and chemical stability lead to the persistence and accumulation of ILs in the environment in the case of any industrial leakage.^{6,7} Meanwhile, there have been several reports discussing the toxicity of ILs to algae, bacteria, fungi, invertebrates, and vertebrates.^{8–12} Exposed ILs will thus be a challenge to terrestrial ecosystems.¹³ It is essential to investigate efficient and economic measures for the degradation of ILs.

A series of studies reported the mineralization of ILs with Fenton reagents,^{14–16} UV radiation,^{17–19} sonication,²⁰ and gamma irradiation^{21,22} as aids. The use of UV radiation showed good performance in the degradation of ILs. It is important to further

evaluate the efficiency of UV usage, the catalytic activity and recyclability of TiO₂.²⁰ Due to the high photocatalytic activity, nontoxicity and good chemical stability, TiO₂ is regarded as an ideal and powerful photocatalyst.²³ However, its band gap is as high as 3.2 eV and it can only be excited by the UV light that makes up only 3–5% of the entire solar spectrum.^{24–26} In addition, there are also difficulties in practical applications of TiO₂. Although the common form of ultrafine powder increases the surface area and the number of active sites,²⁷ it is easy to be agglomerated into large particles in aqueous solutions. The photocatalytic performance is thus reduced.²⁸

To solve the aforementioned problem, several studies have been performed to support TiO₂ onto a wide variety of materials such as zeolites,^{27,28} glass beads²⁹ and glass capillaries³⁰ for the photocatalytic degradation of methylene blue, methyl orange and agri-food industry effluents. In this study, a series of TiO₂/SiO₂ are *in situ* synthesized and then sensitized with 2,9-dichloroquinacridone (DCQ) in order to show better photocatalytic abilities upon full solar spectrum irradiation. The dye-sensitized photocatalysts (DCQ-TiO₂/SiO₂) are characterized through BET, TEM, XRD, Raman, FT-IR, elemental analysis and UV-vis DRS, and are then used in the degradation of some imidazolium ionic liquids. Experimental results show that the dye sensitization of titania-silica materials with DCQ can significantly increase the photodegradation efficiency. More than 95% of [BMIM]Br can be degraded within 180 min under simulated sunlight irradiation. The photodegradation

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† Electronic supplementary information (ESI) available. See DOI: 10.1039/c7ra04939k



intermediates of $[\text{BMIM}]^+$ are identified as 2-oxo-2-ureidoacetic acid, butyric acid and others. These intermediates are harmless and can be easily biodegraded in the ecosystem.

2. Experimental

2.1 Synthesis of the photocatalysts

In this study, $\text{TiO}_2/\text{SiO}_2$ composite materials were prepared according to the method reported in literature.^{31,32} Triblock copolymer P123 ($\text{EO}_{20}\text{PO}_{70}\text{EO}_{20}$, $M_w = 5800$, Sigma-Aldrich) was used as the structure-directing agent, tetramethoxysilane (TMOS, Acros) and tetrabutyl orthotitanate (TBOT, Acros) were used as the silica source and titania source, respectively. All reagents were used as received. The catalysts obtained are denoted as $X\% \text{TiO}_2/\text{SiO}_2$ ($X = 0, 10, 20, 30, 40$), where X stands for the mass ratio of Ti in the whole synthesis gel ($X = 0$, the catalyst obtained is SBA-15). In a typical synthesis of $30\% \text{TiO}_2/\text{SiO}_2$, 8.0 g of P123 was dissolved in 160 mL of 1 mol L^{-1} HCl solution under magnetic stirring at room temperature. Subsequently, 9.2 g of TMOS and 6.6 g of TBOT were added to the solution. The resulting mixture was stirred for 24 h at 40 °C. Hydrothermal crystallization was then performed in a Teflon-lined autoclave for 24 h at 110 °C. The white product was separated using vacuum filtration and was then washed with 1 mol L^{-1} HCl (25 mL \times 3). The solid obtained was dried for 10 h at 100 °C and subsequently calcined for 5 h at 550 °C to remove the template. The dye-sensitized photocatalyst DCQ-30% $\text{TiO}_2/\text{SiO}_2$ was prepared in the following way: 2.0 g of 30% $\text{TiO}_2/\text{SiO}_2$ was dispersed in 100 mL of acidic methanol solution (pH = 4.7) under vigorous stirring for 15 min. 40 mL methanol solution of 0.01 mol L^{-1} DCQ (Gaoyou Auxiliary Factory) was then added. The mixture was stirred for 24 h at room temperature. After the reaction, the precipitate was separated using vacuum filtration and was washed with methanol (60 mL \times 3). The solid obtained was then dried for 10 h at 100 °C in vacuum.

2.2 Characterization of the photocatalysts

The textural properties of the photocatalysts were determined by N_2 adsorption-desorption isotherms at -196°C using a Micromeritics ASAP-2010. Prior to adsorption, the samples were thoroughly degassed at 150 °C for 8 h under vacuum to remove water and other physically adsorbed species. The specific surface area (S_{BET}) was calculated through the standard Brunauer-Emmett-Teller (BET) method. TEM images were captured on a JEM-2100 microscope (JEOL) operated at 200 kV. XRD patterns of the samples were obtained on a D/max-3A diffractometer (Rigaku) equipped with a CuK_α X-ray source (40 kV, 30 mA, $\lambda = 0.154 \text{ nm}$) and Ni filter. Diffractograms were recorded in the 2θ ranges of $0.6\text{--}10^\circ$ (small angle) and $10\text{--}80^\circ$ (wide angle) at a scanning step of 0.02° . Raman spectroscopy was carried out on a LabRamHR spectrometer (Horiba Jobin Yvon Inc.). All experiments were performed on the electrode surface by using a 632.8 nm excitation line from an internal He-Ne laser with a power of 6 mW. FT-IR spectra were recorded on an Equinox-55 spectrophotometer (Bruker) by KBr wafer technique. Elemental analysis was performed on a Vario EL III

(Elementar). UV-vis DRS were obtained in the wavelength of 200–800 nm by a UV-vis spectrophotometer (Shimadzu U-3010) equipped with an integrating sphere assembly and BaSO_4 was used as a reflectance standard.

2.3 Test of photocatalytic activity

The photocatalytic degradation of ionic liquids was conducted using $X\% \text{TiO}_2/\text{SiO}_2$ and DCQ- $X\% \text{TiO}_2/\text{SiO}_2$ as photocatalysts under simulated sunlight irradiation provided by a 300 W Xe lamp (CEL-HXUV300, Beijing Au light company). A quartz reactor with cooling jacket was used to keep the temperature controlled at 25 °C. In each experiment, 0.2 g of photocatalyst was ultrasonically dispersed in a 100 mL of solution, where the concentrations of ionic liquids and H_2O_2 were 45 mg L^{-1} and 14 mg L^{-1} respectively, and pH was adjusted to 3.9. The suspension was magnetically stirred at 200 rpm and then stored in dark for 30 min to reach adsorption-desorption equilibrium. 1 mL of sample was then taken out as a blank control to eliminate the effects of adsorption. During the degradation process, 1 mL of sample was collected by a pipette at 30 min intervals and then filtered through a $0.45 \mu\text{m}$ cellulose membrane (Millex, Millipore). The degradation degree of ionic liquids was measured by high-performance liquid chromatography (HPLC, Agilent 1200) equipped with a reverse-phase C18 analytical column (Zorbax SB-C18). A water-acetonitrile mixture (70 : 30, v/v) was used as the mobile phase. The flow rate was set to be 1 mL min^{-1} . The degradation efficiency ($D\%$) was then calculated according to eqn (1):

$$D\% = (C_0 - C_t)/C_0 \times 100 \quad (1)$$

where C_0 is the initial concentration of ionic liquids, and C_t is the concentration of ionic liquids after irradiation for a selected time interval. To identify intermediates in the photocatalytic degradation of ionic liquids (mainly the cation of $[\text{BMIM}]^+$), liquid chromatography-mass spectrometry (LC/MS) was performed on an Agilent 1100/Esquire HCT PLUS equipped with an electrospray ionization source.

3. Results and discussion

3.1 N_2 adsorption-desorption isotherms

The textural properties of SBA-15, $X\% \text{TiO}_2/\text{SiO}_2$, and DCQ- $X\% \text{TiO}_2/\text{SiO}_2$ samples were investigated and shown in Fig. 1. All samples display type IV isotherms and hysteresis loops in the relative pressure range of $0.5 < P/P_0 < 0.9$. These features are associated with capillary condensation of N_2 in primary mesopores,²⁸ which therefore indicate that the samples are mesoporous materials. Hysteresis loops of $X\% \text{TiO}_2/\text{SiO}_2$ and DCQ- $X\% \text{TiO}_2/\text{SiO}_2$ show reduced sharpness and are shifted to lower relative pressures (P/P_0) over a larger range of P/P_0 with the incorporation of TiO_2 and DCQ in the matrix of SiO_2 , indicating that pores in these samples are less uniform and smaller.^{32–34} Detailed information on S_{BET} , average pore diameter (D_{BJH}), and pore volume (V_{pore}) is given in Table S1 (ESI†). It shows that higher TiO_2 loadings in the samples result in the decrease of S_{BET} . Further, S_{BET} of DCQ- $X\% \text{TiO}_2/\text{SiO}_2$ are smaller than that



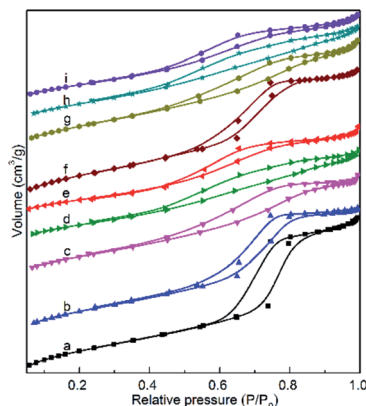


Fig. 1 N_2 adsorption-desorption isotherms of the samples: (a) SBA-15; (b)–(e) 10, 20, 30, 40% TiO_2/SiO_2 , (f)–(i) DCQ-10, 20, 30, 40% TiO_2/SiO_2 .

of their counterpart ($X\%$ TiO_2/SiO_2) due to the introduction of DCQ. Nonetheless, the S_{BET} values of all materials remain high (836.9–496.6 $m^2 g^{-1}$).

3.2 XRD analysis

Small-angle XRD patterns of SBA-15, $X\%$ TiO_2/SiO_2 , and DCQ- $X\%$ TiO_2/SiO_2 samples are shown in Fig. 2A. SBA-15 produces three reflections of (1 0 0), (1 1 0), and (2 0 0), which correspond to a two-dimensional hexagonal meso-structure of the $p6mm$ space group with high structural periodicity.^{35,36} The samples of 10% TiO_2/SiO_2 and DCQ-10% TiO_2/SiO_2 produce XRD patterns similar to that of SBA-15, indicating that they retain the SBA-15 structure. However, the diffraction intensities of the (1 1 0) and (2 0 0) reflections decrease gradually with increasing TiO_2 loadings, implying that the unit cell dimensions and the order of the meso-structures will be affected by the incorporation of more TiO_2 ,^{35,36} as confirmed by the TEM images.

Wide-angle XRD patterns are presented in Fig. 2B. SBA-15 shows diffraction patterns with a broad peak at 22.2° . All

patterns of TiO_2/SiO_2 and DCQ- TiO_2/SiO_2 samples show peaks at $2\theta = 24.5^\circ, 38.0^\circ, 48.2^\circ, 54.3^\circ, 62.9^\circ, 69.9^\circ$, and 75.3° (JCPDS-ICDD no. 21-1272), which are assigned to the anatase reflections of (1 0 1), (0 0 4), (2 0 0), (1 0 5 and 2 1 1), (2 0 4), (1 1 6 and 2 2 0), and (2 1 5), respectively.²⁸ Intensities of these peaks increase markedly with increasing TiO_2 loadings. No peak for the (1 1 0) rutile reflection at 27.4° has been detected in the patterns of all samples. These results indicate that TiO_2 in the samples is pure anatase with no secondary crystal phases of either brookite or rutile. It was reported that anatase in pure TiO_2 would transform to rutile when the calcination temperature was higher than $400^\circ C$.³⁷ In this study, no rutile phase is observed although all samples are calcined at $550^\circ C$. It could be attributed to the stabilizing effect of SBA-15 that enhanced the thermal stability of the anatase crystalline phase.

3.3 Raman spectroscopy

Raman spectra of all samples are shown in Fig. 3. Peaks at 149, 401, 519, and 645 cm^{-1} can be found in the spectra of all samples, except for DCQ. These peaks are assigned to the $E_{g(1)}$, $B_{1g(1)}$, A_{1g} (or $B_{1g(2)}$), and $E_{g(3)}$ modes, respectively. Meanwhile, there are no bands at 268 and 601 cm^{-1} , which correspond to the TiO_2 rutile phase.^{28,38,39} This observation is consistent with the above XRD results, indicating clearly that only the anatase phase of TiO_2 is present in these samples. Furthermore, the intensities of these four peaks increase with increasing TiO_2 loadings. Other peaks at 291, 1199, 1229, 1311, 1571, 1595, and 1650 cm^{-1} in Fig. 3f–j is regarded as characteristic bands of DCQ.⁴⁰ In particular, bands at 1311, 1571, and 1650 cm^{-1} can be assigned to the coupled ring vibration, aromatic vibration, and CO stretching vibration of the quinacridone groups in DCQ.

3.4 TEM

Some TEM images of SBA-15 and DCQ- $X\%$ TiO_2/SiO_2 samples are selected and shown in Fig. 4. Other images are presented in the ESI (Fig. S1–S9†). The long-range ordering and hexagonally arranged pores of SBA-15 can be well recognized in Fig. 4A. The regular morphology is maintained in the samples of DCQ-10%

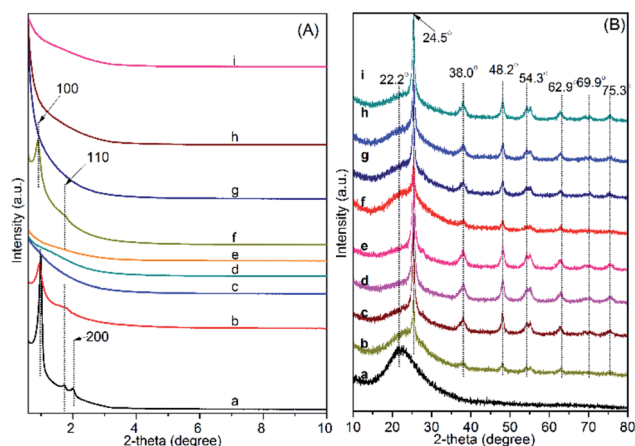


Fig. 2 Small-angle (A) and wide-angle (B) XRD spectra of the samples: (a) SBA-15; (b)–(e) 10, 20, 30, 40% TiO_2/SiO_2 , (f)–(i) DCQ-10, 20, 30, 40% TiO_2/SiO_2 .

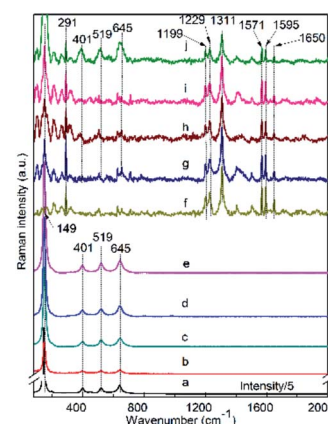


Fig. 3 Raman spectra of the photocatalysts: (a) TiO_2 ; (b)–(e) 10, 20, 30, 40% TiO_2/SiO_2 ; (f) DCQ; (g)–(j) DCQ-10, 20, 30, 40% TiO_2/SiO_2 .



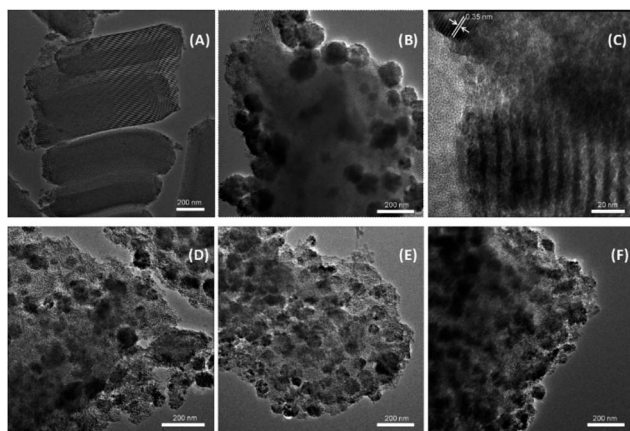


Fig. 4 TEM images of (A) SBA-15 and (B–F) DCQ- $X\%$ $\text{TiO}_2/\text{SiO}_2$ samples: (B, C) $X = 10$, (D) $X = 20$, (E) $X = 30$, and (F) $X = 40$.

$\text{TiO}_2/\text{SiO}_2$ (Fig. 4B). The symmetry of local structure gradually decreases with increasing TiO_2 contents (Fig. 4D–F). Fig. 4C also confirms the anatase crystal structure of TiO_2 particles. The lattice spacing (d) is ~ 0.35 nm between adjacent lattice planes. This matches well the distance between the (1 0 1) crystal planes of anatase TiO_2 .²³ The anatase phase of TiO_2 shows generally strong photocatalytic behaviour in various reactions, for example, the decomposition of methylene blue in aqueous solutions.²⁸ Hence, good photocatalytic performance is expected with the presence of anatase TiO_2 in the samples presented in this work.

3.5 FT-IR

FT-IR spectra of the samples SBA-15, $X\%$ $\text{TiO}_2/\text{SiO}_2$, and DCQ- $X\%$ $\text{TiO}_2/\text{SiO}_2$ and their corresponding assignments are shown in Fig. 5 and Table S2.† The peak intensities of all $X\%$ $\text{TiO}_2/\text{SiO}_2$ and DCQ- $X\%$ $\text{TiO}_2/\text{SiO}_2$ samples at 960 cm^{-1} are obviously higher than that of SBA-15. It was reported that the incorporation of TiO_2 into the framework of SBA-15 usually enhanced the intensity of the silanol shoulder (at 960 cm^{-1}) of the ν_{as} (Si–O–Si) band;³⁶ such enhancement of the peak intensity is also a strong evidence about the formation of Ti–O–Si bonds in all

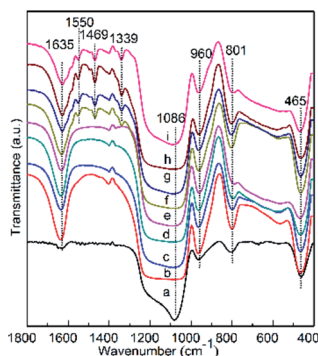


Fig. 5 FT-IR spectra of the samples: (a) SBA-15; (b)–(e) 10, 20, 30, 40% $\text{TiO}_2/\text{SiO}_2$, (f)–(i) DCQ-10, 20, 30, 40% $\text{TiO}_2/\text{SiO}_2$.

Table 1 Elemental analysis of SBA-15 and DCQ- $X\%$ $\text{TiO}_2/\text{SiO}_2$

Samples	C (wt%)	N (wt%)	H (wt%)	DCQ loading (wt%)
SBA-15	0.525	0.150	0.976	—
DCQ-10% $\text{TiO}_2/\text{SiO}_2$	1.327	0.379	2.411	3.12
DCQ-20% $\text{TiO}_2/\text{SiO}_2$	2.225	0.477	1.765	4.45
DCQ-30% $\text{TiO}_2/\text{SiO}_2$	2.426	0.513	2.062	4.94
DCQ-40% $\text{TiO}_2/\text{SiO}_2$	1.893	0.476	2.129	4.44

$X\%$ $\text{TiO}_2/\text{SiO}_2$ and DCQ- $X\%$ $\text{TiO}_2/\text{SiO}_2$ samples. At the same time, bands at 1550, 1469, and 1339 cm^{-1} , which are characteristic of DCQ (Table S2†), can only be found in spectra of DCQ- $X\%$ $\text{TiO}_2/\text{SiO}_2$.

3.6 Elemental analysis and UV-vis DRS

In order to investigate the loadings of DCQ in the samples of DCQ- $X\%$ $\text{TiO}_2/\text{SiO}_2$, elemental analysis is performed and the results are shown in Table 1. The DCQ loadings in the samples range from about 3 to 5 wt%, the sequences are ranked in the following order: DCQ-30% $\text{TiO}_2/\text{SiO}_2 > \text{DCQ-20% } \text{TiO}_2/\text{SiO}_2 > \text{DCQ-40% } \text{TiO}_2/\text{SiO}_2 > \text{DCQ-10% } \text{TiO}_2/\text{SiO}_2$.

Fig. 6 presents the UV-vis diffuse reflectance spectra of SBA-15, TiO_2 , DCQ, and $X\%$ $\text{TiO}_2/\text{SiO}_2$ and DCQ- $X\%$ $\text{TiO}_2/\text{SiO}_2$. As shown in Fig. 6A, all samples show high absorbance in the ultraviolet region (200–400 nm) except for SBA-15. TiO_2 has an adsorption edge at 376 nm. This is consistent with the value of anatase TiO_2 reported in the literature.^{41,42} The adsorption edge of $X\%$ $\text{TiO}_2/\text{SiO}_2$ samples are slightly red shifted from 376 nm to 391 nm with increasing TiO_2 contents. The red shift might be caused by the formation of Ti–O–Si chemical bonds^{41,43} during the *in situ* incorporation of TiO_2 . The absorption intensities of $X\%$ $\text{TiO}_2/\text{SiO}_2$ samples in the ultraviolet region are also found to be increased with increasing amounts of TiO_2 incorporated. It can be observed from Fig. 6B that all DCQ- $X\%$ $\text{TiO}_2/\text{SiO}_2$ samples show broad and prominent absorbance in the visible light region of 450–600 nm. Obviously, this is the result of the sensitization of DCQ, which can further facilitate a more efficient utilization of simulated sunlight and improve the photocatalytic activity under full spectrum irradiation. It can also be seen the absorbance intensity of DCQ- $X\%$ $\text{TiO}_2/\text{SiO}_2$ in the

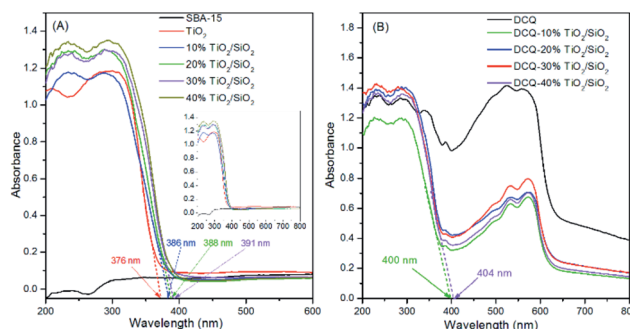


Fig. 6 UV-vis DRS of (A) SBA-15, TiO_2 , $X\%$ $\text{TiO}_2/\text{SiO}_2$ and (B) DCQ, DCQ- $X\%$ $\text{TiO}_2/\text{SiO}_2$ ($X = 10, 20, 30, 40$).



region of 450–600 nm is consistent with the DCQ loadings in all the samples (Table 1). The band gap energies (E_g) of the samples can be estimated according to the empirical eqn (2)^{42,44} and the results are shown in Table S3.†

$$E_g = 1240/\lambda_g \quad (2)$$

where E_g is the energy gap and λ_g is absorption threshold. Table S3† shows that the incorporation of TiO_2 into SBA-15 and the sensitization of the samples with DCQ can get the band gap of TiO_2 narrowed, although the effect is not significant.

3.7 Photocatalytic degradation

The photocatalytic degradation efficiency ($D\%$) of imidazolium ionic liquids were tested and shown in Fig. 7. It can be seen that $D\%$ of [BMIM]Br within 180 min is only 10 with SBA-15 in the blank experiment. $D\%$ of [BMIM]Br can reach about 65 to 75 with un-sensitized $X\%$ $\text{TiO}_2/\text{SiO}_2$ catalysts under simulated sunlight irradiation. This indicates that TiO_2 incorporated into the composites acts as photocatalytic active site and plays an important role on the photodegradation. It is well known that the irradiation (especially the UV irradiation according to DRS results) on the active sites of TiO_2 will generate conduction-band electrons (e^-) and valence band holes (h^+), which react with H_2O_2 to produce $\cdot\text{OH}$ radicals. Both valence band holes and $\cdot\text{OH}$ radicals are highly oxidative and can degrade organic compounds quickly.^{17,45}

It is interesting to notice that the photodegradation performance is considerably improved when DCQ- $X\%$ $\text{TiO}_2/\text{SiO}_2$ is used as catalysts under simulated sunlight irradiation, $D\%$ of [BMIM]Br could reach 90–95 within 180 min. This can be explained by the sensitization of $\text{TiO}_2/\text{SiO}_2$ with DCQ, which can significantly enhance the utilization of solar light and then improve the photodegradation efficiency. According to the results of UV-vis DRS, the enhancement and promotion is

presumably due to the strong absorption of DCQ- $X\%$ $\text{TiO}_2/\text{SiO}_2$ samples in the visible-light region, rather than the band gap narrowing of TiO_2 . The similar result was also reported by Sun *et al.*⁴¹ when N-Ag- TiO_2 catalyst was applied in photocatalytic oxidation of phenol solution under visible light irradiation. It can be seen from Fig. 7 that $D\%$ of [BMIM]Br remains above 80 even after 10 recycling use, which means the photocatalyst has satisfactory stability under simulated sunlight irradiation. Further, three other ILs were adopted to investigate the effect of water solubility on $D\%$. It shows that the worse the water solubility, the lower the degradation efficiency, but $D\%$ of [BMIM]NTF₂ can still reach up to nearly 77 within 180 min.

LC/MS was used to identify the intermediates of [BMIM] cation in the photocatalytic degradation process. Main peaks in the HPLC chromatograms (Fig. S10A†) are located at retention times of 2.6, 2.8, 2.9, and 3.4 min, which can be assigned to [BMIM]⁺ and its photodegradation intermediates (a–c, Fig. S10B–D†). The peak of [BMIM]⁺ disappears rapidly during degradation; meanwhile, the peak intensity of intermediate (a) increases from 30 to 60 min and then decreases from 60 to 180 min. These trends indicate that intermediate (a) is produced first during the degradation reaction and is then degraded continuously in the photocatalytic process. While peak intensities of intermediates (b) and (c) gradually increase from 30 to 180 min. Based on the above analysis, the mechanism of [BMIM]⁺ photocatalytic degradation is proposed in Scheme S1† and can be presented as follows: [BMIM]⁺ is firstly oxidized to 1-butyl-3-methylimidazolidine-2,4,5-trione (a), and then degraded to 2-oxo-2-ureidoacetic acid (b) and butyric acid (c) by ring-opening. And the compound of (b) and (c) will be finally mineralized to CO_2 , H_2O *etc.*

4. Conclusions

In this study, the photocatalysts $\text{TiO}_2/\text{SiO}_2$ and DCQ- $\text{TiO}_2/\text{SiO}_2$ are prepared using *in situ* synthesis and dye sensitization. The results of TEM, XRD analysis, Raman spectroscopy, elemental analysis and UV-vis diffuse reflectance spectra clearly show the formation of TiO_2 - SiO_2 mixed oxides in the $\text{TiO}_2/\text{SiO}_2$ and DCQ- $\text{TiO}_2/\text{SiO}_2$ samples. TiO_2 in the samples is pure anatase without secondary crystal phase of brookite or rutile. Results of Raman, FT-IR and UV-vis DRS also indicate that DCQ is successfully incorporated into DCQ- $\text{TiO}_2/\text{SiO}_2$. The photocatalysts are then used in the degradation of imidazolium ionic liquids. The degradation experiments show that dye sensitization of $\text{TiO}_2/\text{SiO}_2$ with DCQ can enhance the photodegradation efficiency by over 20%. Around 95% of [BMIM]Br can be degraded within 180 min with the use of DCQ- $\text{TiO}_2/\text{SBA-15}$ as photocatalyst under simulated sunlight irradiation. The dye-sensitized photocatalysts can thus be applied efficiently in the degradation of ILs, *e.g.* for treating spillage of ILs. The dye sensitized catalysts are expected to be applicable for the degradation of other contaminants in aquatic environments. Such applications are under investigation.

Acknowledgements

This study is performed with the financial support of the National Natural Science Foundation of China (Grant No.

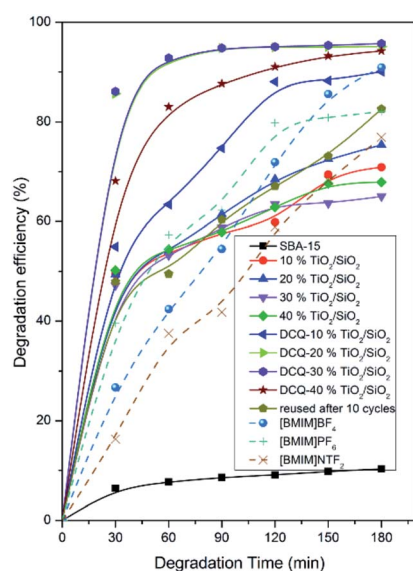


Fig. 7 Degradation efficiency of imidazolium ILs under simulated sunlight irradiation. Solid line: the photodegradation target was [BMIM]Br; dash line: DCQ-30% $\text{TiO}_2/\text{SiO}_2$ was used as the photocatalyst.



21006035 and 21676099), and the Fundamental Research Funds for the Central Universities, South China University of Technology.

Notes and references

- W. Zhu, P. Wu, L. Yang, Y. Chang, Y. Chao, H. Li, Y. Jiang, W. Jiang and S. Xun, *Chem. Eng. J.*, 2013, **229**, 250–256.
- Y. Yu, J. Mai, L. Huang, L. Wang and X. Li, *RSC Adv.*, 2014, **4**, 12756–12762.
- Y. Yu, J. Mai, L. Wang, X. Li, Z. Jiang and F. Wang, *Sci. Rep.*, 2014, **4**, 5997.
- J. Neumann, C. Cho, S. Steudte, J. Koser, M. Uerdingen, J. Thoming and S. Stolte, *Green Chem.*, 2012, **14**, 410–418.
- J. Ma, L. Cai, B. Zhang, L. Hu, X. Li and J. Wang, *Ecotoxicol. Environ. Saf.*, 2010, **73**, 1465–1469.
- J. McFarlane, W. Ridenour, H. Luo, R. Hunt, D. DePaoli and R. Ren, *Sep. Sci. Technol.*, 2005, **40**, 1245–1265.
- T. Thuy Pham, C. Cho and Y. Yun, *Water Res.*, 2010, **44**, 352–372.
- K. Docherty and J. Kulpa, *Green Chem.*, 2005, **7**, 185–189.
- M. Matzke, S. Stolte, K. Thiele, T. Jufferholz, J. Arning, J. Ranke, U. Biermann and B. Jastorff, *Green Chem.*, 2007, **9**, 1198–1207.
- C. Pretti, C. Chiappe, D. Pieraccini, M. Gregori, F. Abramo, G. Monni and L. Intorre, *Green Chem.*, 2006, **8**, 238–240.
- J. Ranke, A. Müller, U. Weber, F. Stock, S. Stolte, J. Arning, R. Störmann and B. Jastorff, *Ecotoxicol. Environ. Saf.*, 2007, **67**, 430–438.
- P. Stepnowski, A. Skladanowski and A. Ludwiczak, *Hum. Exp. Toxicol.*, 2004, **23**, 513–517.
- H. Zhou, Y. Shen, P. Lv, J. Wang and J. Fan, *Sep. Purif. Technol.*, 2013, **104**, 208–213.
- E. Siedlecka, S. Stolte, M. Gołębowski, A. Nienstedt, P. Stepnowski and J. Thöming, *Sep. Purif. Technol.*, 2012, **101**, 26–33.
- E. Siedlecka, M. Gołębowski, Z. Kaczyński, J. Czupryniak, T. Ossowski and P. Stepnowski, *Appl. Catal., B*, 2009, **91**, 573–579.
- E. Siedlecka, W. Mroziński, Z. Kacynski and P. Stepnowski, *J. Hazard. Mater.*, 2008, **154**, 893–900.
- P. Stepnowski and A. Zaleska, *J. Photochem. Photobiol., A*, 2005, **170**, 45–50.
- A. Morawski, M. Janus, I. Goc-Maciejewska, A. Syguda and J. Pernak, *Pol. J. Chem.*, 2005, **79**, 1929–1935.
- R. Katoh and K. Takahashi, *Radiat. Phys. Chem.*, 2009, **78**, 1126–1128.
- X. Li, J. Zhao, Q. Li, L. Wang and S. Tsang, *Dalton Trans.*, 2007, 1875–1880.
- E. Bosse, L. Berthon, N. Zorz, J. Monget, C. Berthon, I. Bisel, S. Legand and P. Moisy, *Dalton Trans.*, 2008, 924–931.
- G. Le Rouzo, C. Lamouroux, V. Dauvois, A. Dannoux, S. Legand, D. Durand, P. Moisy and G. Moutiers, *Dalton Trans.*, 2009, 6175–6184.
- J. Schneider, M. Matsuoka, M. Takeuchi, J. Zhang, Y. Horiuchi and D. Bahnemann, *Chem. Rev.*, 2014, **114**(19), 9919–9986.
- J. Chao, Z. Xie, X. Duan, Y. Dong, Z. Wang, J. Xu, B. Liang, B. Shan, J. Ye, D. Chen and G. Shen, *CrystEngComm*, 2012, **14**, 3163–3168.
- J. Yu, W. Ho, J. Yu, H. Yip, P. Wong and J. Zhao, *Environ. Sci. Technol.*, 2005, **39**, 1175–1179.
- J. Di, S. Li, Z. Zhao, Y. Huang, Y. Jia and H. Zheng, *Chem. Eng. J.*, 2015, **281**, 60–68.
- J. Yang, J. Zhang, L. Zhu, S. Chen, Y. Zhang, Y. Tang, Y. Zhu and Y. Li, *J. Hazard. Mater.*, 2006, **137**, 952–958.
- Y. Acosta-Silva, R. Nava, V. Hernández-Morales, S. Macías-Sánchez, M. Gómez-Herrera and B. Pawelec, *Appl. Catal., B*, 2011, **110**, 108–117.
- M. Tototzintle, I. Oller, A. Ramírez, S. Malato and M. Maldonado, *Chem. Eng. J.*, 2015, **273**, 205–213.
- C. Shen, Y. Wang, J. Xu and G. Luo, *Chem. Eng. J.*, 2015, **277**, 48–55.
- G. Li and X. Zhao, *Ind. Eng. Chem. Res.*, 2006, **45**(10), 3569–3573.
- G. Wittmann, K. Demeestere, A. Dombi, J. Dewulf and H. Van Langenhove, *Appl. Catal., B*, 2005, **61**, 47–57.
- W. Yue, C. Randorn, P. S. Attidekou, Z. Su, J. T. S. Irvine and W. Zhou, *Adv. Funct. Mater.*, 2009, **19**, 2826–2833.
- W. Jung, S. Baek, J. Yang, K. Lim, M. Lee, G. Lee, S. Park and S. Hong, *Catal. Today*, 2008, **131**, 437–443.
- M. Mazaj, W. Stevens, N. Logar, A. Ristić, N. Tušar, I. Arčon, N. Daneu, V. Meynen, P. Cool, E. Vansant and V. Kaučič, *Microporous Mesoporous Mater.*, 2009, **117**, 458–465.
- S. Das, M. Bhunia and A. Bhaumik, *J. Solid State Chem.*, 2010, **183**, 1326–1333.
- H. Yang, H. Lin, Y. Chien, J. Wu and H. Wu, *Catal. Lett.*, 2009, **131**, 381–387.
- R. Zukerman, L. Vladman, L. Titelman, L. Zeiri, N. Perkas, A. Gedanken, M. Landau and M. Herskowitz, *Mater. Chem. Phys.*, 2010, **122**, 53–59.
- M. Alvaro, C. Aprile, M. Benitez, E. Carbonell and H. García, *J. Phys. Chem. B*, 2006, **110**, 6661–6665.
- F. Schulte, K. Brzezinka, K. Lutzenberger, H. Stege and U. Panne, *J. Raman Spectrosc.*, 2008, **39**, 1455–1463.
- H. Sun, G. Zhou, S. Liu, H. Ang, M. Tadé and S. Wang, *Chem. Eng. J.*, 2013, **231**, 18–25.
- A. Police, S. Basavaraju, D. Valluri, V. Muthukonda, S. Machiraju and J. Lee, *Chem. Eng. J.*, 2014, **247**, 152–160.
- J. Li, S. Zhou, G. Hong and C. Chang, *Chem. Eng. J.*, 2013, **219**, 486–491.
- J. Xu, K. Li, W. Shi, R. Li and T. Peng, *J. Power Sources*, 2014, **260**, 233–242.
- P. Wang, T. Zhou, R. Wang and T. Lim, *Water Res.*, 2011, **45**, 5015–5026.

