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Significance of the electron-density of molecular fragments on the properties of manganese(III) β -diketonato complexes: an XPS and DFT study

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DFT and XPS studies were conducted on a series of nine manganese(III) complexes of the general formula $[\text{Mn}(\beta\text{-diketonato})_3]$, with the ligand β -diketonato = dipivaloylmethanato (1), acetylacetonato (2), benzoylacetonato (3), dibenzoylmethanato (4), trifluoroacetylacetonato (5), trifluorothienoylacetonato (6), trifluorofuroylacetonato (7), trifluorobenzoylacetonato (8) and hexafluoroacetylacetonato (9). The binding energy position of the main and satellite structures of the $\text{Mn} 2p_{3/2}$ photoelectron line, as well as the spin-orbit splitting, gave insight into the electronic structure of these manganese(III) complexes. DFT calculations showed that an experimental sample of the d^4 $[\text{Mn}(\beta\text{-diketonato})_3]$ complex can contain a mixture of different bond stretch isomers and different electronic states, in dynamic equilibrium with one other. The presence of more than one isomer in the experimental sample, as well as interaction between an unpaired $2p$ electron (originating after photoemission) and an unpaired $3d$ electron, which aligned anti-parallel to the unpaired $2p$ electron, caused broadening of the $\text{Mn} 2p$ photoelectron lines. Multiplet splitting simulations of these photoelectron lines, similar to those calculated by Gupta and Sen for the free $\text{Mn}^{(III)}$ ion, gave good fits with the observed $\text{Mn} 2p_{3/2}$ photoelectron lines. The XPS spectra of complexes with unsymmetrical β -diketonato ligands were simulated with two sets of multiplet splitting peaks, representing both the *mer* and *fac* isomers. The satellite structures obtained in both the $\text{Mn} 2p_{3/2}$ photoelectron line (shake-up peaks) and the ligand $F 1s$ photoelectron line (shake-down peaks), are representative of the ligand-to-metal charge transfer during photoionisation. The binding energies of the $\text{Mn} 2p$, $F 1s$ and $S 2p$ electrons, as well as the amount of charge transfer from ligand-to-metal, are both dependent on the electronegativity of the different groups attached to the β -diketonato ligand.

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1 Introduction

Knowledge of the electronic structure and properties of transition metal complexes is important, due to their role in a variety of chemical and biological processes. High spin manganese(III) complexes have various applications, for example as building blocks in molecular magnets,^{1–8} catalysis,^{9–14} redox mediators in dye-sensitised solar cells,^{15,16} and in biochemical reaction cycles.^{17–21} Due to the four unpaired d electrons of high spin $\text{Mn}^{(III)}$ complexes, their geometry displays the Jahn-Teller effect and also the X-ray photoelectron spectroscopy (XPS) lines of these complexes split into many peaks, due to a coupling interaction between the unpaired core p -electrons (originating after photoemission) and the unpaired valence d -electrons.²² Recently an X-ray photoelectron spectroscopy study on $[\text{Mn}^{(III)}(\text{ferrocenyl-}\beta\text{-diketonato})_3]$ complexes, where the β -diketonato ligand contained at least one ferrocenyl group,²³ showed that a charge

transfer process from the β -diketonato ligand to the $\text{Mn}^{(III)}$ metal centre is responsible for prominent shake-up satellite peaks of the main $\text{Mn} 2p$ photoelectron lines. These results further showed that the binding energies of the $\text{Mn} 2p$ photoelectron lines are linearly influenced by the sum of the Gordy scale group electronegativities²⁴ of the different R groups attached to the β -diketonato ligands (RCOCHCOFc^- (Fc = ferrocenyl group)), varying over a binding energy range of 0.55 eV. By adding the XPS data of $[\text{Mn}^{(III)}(\beta\text{-diketonato})_3]$ complexes, where the β -diketonato ligand does not contain an electron donating ferrocenyl group, we hereby enlarge the range of binding energies of the $\text{Mn} 2p$ photoelectron lines by nearly 100%, to a range of 1.07 eV. Thus, in continuation of our interest in the geometrical^{25,26} and electronic structure²⁷ of $[\text{Mn}(\beta\text{-diketonato})_3]$ complexes containing β -diketonato ligands with different electondonating and -withdrawing properties, we hereby present a combined study of theoretical quantum chemistry calculations (using density functional theory, DFT) and experimental measurements (using XPS) on the series of nine non-ferrocene-containing $[\text{Mn}(\beta\text{-diketonato})_3]$ complexes shown in Fig. 1 (with β -diketonato = dipivaloylmethanato (1), acetylacetonato (2), benzoylacetonato

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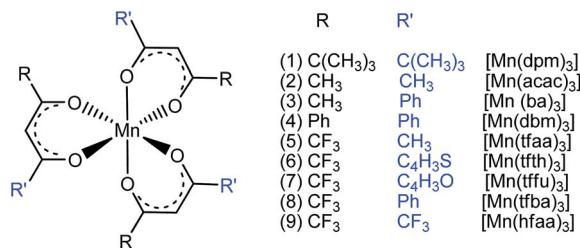


Fig. 1 Structure of the different $[\text{Mn}^{\text{III}}(\beta\text{-diketonato})_3]$ complexes under investigation. The β -diketonato ligands of complexes 1, 2, 4 and 9 are symmetrical, while complexes 3, 5–8 have unsymmetrically substituted β -diketonato ligands.

(3), dibenzoylmethanato (4), trifluoroacetylacetonato (5), trifluorothenoylacetonato (6), trifluorofuroylacetonato (7), trifluorobenzoylacetonato (8) and hexafluoroacetylacetonato (9). Results are also compared with previously published results²³ on $[\text{Mn}^{\text{III}}(\text{ferrocenyl-}\beta\text{-diketonato})_3]$ complexes.

2 Results and discussion

The spectra obtained by X-ray photoelectron spectroscopy (XPS) of 3d transition metal complexes containing unpaired d electrons, frequently contain a variety of several features which arise from the electronic structure of the material. $[\text{Mn}(\beta\text{-diketonato})_3]$ complexes are high-spin paramagnetic octahedral complexes, with a d-electron occupation of $(t_{2g})^3(e_g)^1$.²⁸ Since the electronic configurations $(t_{2g})^3d_{z^2}^1$ and $(t_{2g})^3d_{x^2-y^2}^1$ are degenerate, Jahn-Teller geometric distortion of $[\text{Mn}(\beta\text{-diketonato})_3]$ complexes occurs in order to remove the degeneracy, with a consequent lowering in symmetry and energy.²⁹ Elongation Jahn-Teller distortion in $[\text{Mn}(\beta\text{-diketonato})_3]$ complexes is characterised by two long bonds along the z-direction, as well as four short bonds in the xy-plane, and a d-electron occupation of $d_{xy}^1d_{xz}^1d_{yz}^1d_{z^2}^1$, with the highest occupied molecular orbital (HOMO) of mainly d_{z^2} character. The elongation along the z-direction stabilises the energy of the d_{z^2} HOMO. Compression Jahn-Teller distortion in $[\text{Mn}(\beta\text{-diketonato})_3]$ complexes on the other hand, is

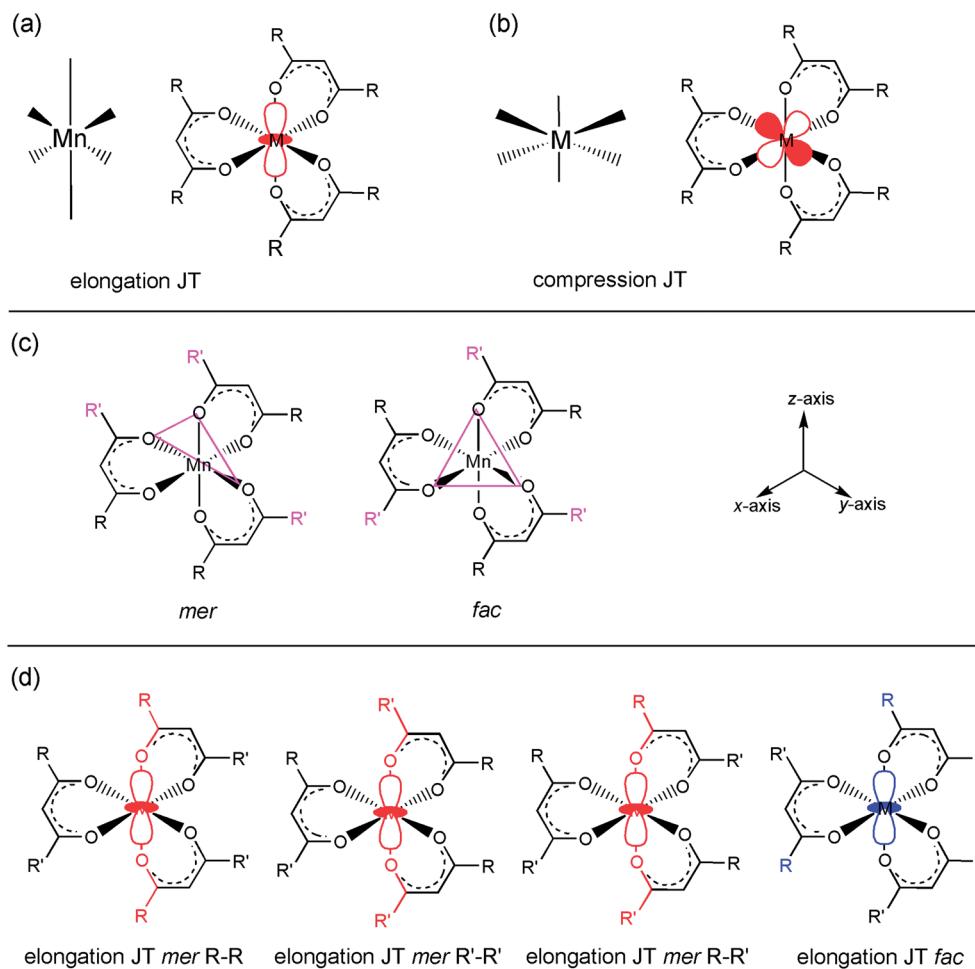


Fig. 2 Illustration of the two electronic states of $[\text{Mn}(\beta\text{-diketonato})_3]$ complexes, containing a symmetrical β -diketonato ligand ($\text{RCOCHCOR}'^-$), leading to either (a) elongation ($d_{xy}^1d_{xz}^1d_{yz}^1d_{z^2}^1$) or (b) compression ($d_{xy}^1d_{xz}^1d_{yz}^1d_{x^2-y^2}^1$) Jahn-Teller distortion. (c) The fac and mer isomers of $[\text{Mn}(\beta\text{-diketonato})_3]$ complexes, containing an unsymmetrical β -diketonato ligand ($\text{RCOCHCOR}'^-$) with $\text{R} \neq \text{R}'$. (d) Four possible elongation ($d_{xy}^1d_{xz}^1d_{yz}^1d_{z^2}^1$) Jahn-Teller bond stretch isomers for $[\text{Mn}(\beta\text{-diketonato})_3]$ complexes, containing an unsymmetrical β -diketonato ligand ($\text{RCOCHCOR}'^-$) with $\text{R} \neq \text{R}'$. The d_{z^2} HOMO or $d_{x^2-y^2}$ HOMO has been superimposed on the structures.

characterised by two short bonds along the z -direction, as well as four long bonds in the xy -plane, and a d-electron occupation of $d_{xy}^1 d_{xz}^1 d_{yz}^1 d_{z^2-y^2}^1$ instead, with the HOMO of mainly $d_{x^2-y^2}$ character. Elongation along the x and y -directions in this case, stabilises the energy of the $d_{x^2-y^2}$ HOMO.^{25,27} The two electronic states $d_{xy}^1 d_{xz}^1 d_{yz}^1 d_{z^2}^1$ and $d_{xy}^1 d_{xz}^1 d_{yz}^1 d_{x^2-y^2}^1$, respectively leading to elongation and compression Jahn–Teller distortion in $[\text{Mn}(\beta\text{-diketonato})_3]$ complexes, are illustrated in Fig. 2(a) and (b), for $[\text{Mn}(\beta\text{-diketonato})_3]$ complexes containing a symmetrically substituted β -diketonato ligand ($\text{RCOCHCOR}'^-$), with groups $R = R'$.

2.1 DFT study

The influence of the ligands upon the geometry, energy and electronic structure of $[\text{Mn}(\beta\text{-diketonato})_3]$ complexes, can conveniently be studied by quantum chemical calculations. DFT calculations conducted on both the two electronic states $d_{xy}^1 d_{xz}^1 d_{yz}^1 d_{z^2}^1$ and $d_{xy}^1 d_{xz}^1 d_{yz}^1 d_{x^2-y^2}^1$ of complexes 1–9, show that for 1–9, the elongation Jahn–Teller distortion geometry is generally more stable than the compression Jahn–Teller distortion geometry, see Table 1.^{26,30} Both elongation³¹ as well as compression Jahn–Teller distortion^{32,33} has experimentally been observed in crystal structures obtained for $[\text{Mn}(\beta\text{-diketonato})_3]$ complexes. However, compression Jahn–Teller distortion as observed for $[\text{Mn}(\text{acac})_3]$,³³ has actually been attributed to an

unusual consequence of unknown crystal packing effects,³⁴ stating that typically the tetragonal elongation is the “natural” Jahn–Teller effect of $[\text{Mn}(\text{acac})_3]$, in agreement with the DFT calculations obtained in this study (Hacac = acetylacetone).

When the $[\text{Mn}(\beta\text{-diketonato})_3]$ complex contains an unsymmetrically substituted β -diketonato ligand (with groups $R \neq R'$, as for complexes 3, 5–8), then two geometrical isomers are possible, namely both the *fac* and *mer* isomers, depending on the arrangement of the unsymmetrical β -diketonato ligands around the metal; see Fig. 2(c). In the case of unsymmetrical complexes, the elongation Jahn–Teller distortion of the *mer* isomer can occur along any of the three different $\text{O}_{\beta\text{-diketonato}}\text{–Mn–O}_{\beta\text{-diketonato}}$ bonds, leading to as much as three different bond stretch Jahn–Teller elongation *mer* isomers for the electronic state $d_{xy}^1 d_{xz}^1 d_{yz}^1 d_{z^2}^1$, as illustrated in Fig. 2(d).^{26,30} The OLYP/TZP energies for these four (including the one *fac*) different elongation $[\text{Mn}^{\text{III}}(\beta\text{-diketonato})_3]$ isomers of unsymmetrical complexes 3, 5–8, are given in Table 1. The DFT calculations show that for these unsymmetrical complexes 3, 5–8, (i) the most stable isomer is a *mer* isomer, (ii) the *fac* isomer is generally slightly less stable than the different *mer* isomers, and (iii) the energy differences between the different bond stretch isomers of $[\text{Mn}(\beta\text{-diketonato})_3]$ complexes are small. This result indicates that an experimental sample of $[\text{Mn}(\beta\text{-diketonato})_3]$ can contain a mixture of all four different bond stretch isomers (and possibly also different electronic states), in dynamic equilibrium with one other. This result agrees with the fact that experimental crystal structures of both the *fac*^{15,35} and *mer* isomers²⁶ of $[\text{Mn}(\beta\text{-diketonato})_3]$ complexes have been characterised.

While the DFT study focused on the energy and character of the different $[\text{Mn}(\beta\text{-diketonato})_3]$ isomers, which are mainly determined by the valence Mn-d electrons, an X-ray photoelectron spectroscopy (XPS) study of a series of these $[\text{Mn}(\beta\text{-diketonato})_3]$ complexes, gave insight into the energies and properties of the core electrons of Mn and other elements present in $[\text{Mn}(\beta\text{-diketonato})_3]$.

2.2 XPS study

X-ray photoelectron spectroscopy (XPS) is a convenient characterisation technique to study the chemical and oxidation states of the elements present in a sample.³⁶ The binding energy position, shape and intensity of a photoelectron line can additionally provide a wealth of information regarding the magnetic character,³⁷ electronic properties^{38,39} and metal-ligand bond of the elements under investigation.^{40–42}

The XPS of complexes 1–9 in this study all displayed the Mn 2p (641–652 eV), C 1s (*ca.* 284.5 eV) and O 1s (*ca.* 528.8 eV) photoelectron lines, while the XPS of the five fluorine-containing complexes 5–9 also showed the F 1s photoelectron line (main line at *ca.* 688.5 eV; satellite line at *ca.* 684.9 eV) and complex 6 displayed the S 2p photoelectron lines (main line at *ca.* 164.7 eV; satellite at *ca.* 161.2 eV). The Mn 2p region of the XPS spectra of these non-ferrocene containing $[\text{Mn}^{\text{III}}(\beta\text{-diketonato})_3]$ complexes 1–9 of this study, is shown in Fig. 3, with selected data tabulated in Table 2. The Mn 2p_{3/2} and Mn

Table 1 DFT calculated energies of the indicated electronic states of $[\text{Mn}(\beta\text{-diketonato})_3]$ complexes 1–9

No	Relative energy in eV	
	$d_{xy}^1 d_{xz}^1 d_{yz}^1 d_{z^2}^1$	$d_{xy}^1 d_{xz}^1 d_{yz}^1 d_{x^2-y^2}^1$
1	$[\text{Mn}(\text{dpm})_3]^a$	0.000
2	$[\text{Mn}(\text{acac})_3]^a$	0.000
3	$[\text{Mn}(\text{ba})_3]^a$	0.000 (<i>mer</i> Ph–Ph) ^a 0.015 (<i>mer</i> Ph–Me) ^a 0.025 (<i>mer</i> Me–Me) ^a 0.020 (<i>fac</i>) ^a
4	$[\text{Mn}(\text{dbm})_3]^a$	0.000
5	$[\text{Mn}(\text{tfaa})_3]$	0.000 (<i>mer</i> CF ₃ –CF ₃) 0.030 (<i>mer</i> Me–CF ₃) 0.064 (<i>mer</i> Me–Me) 0.044 (<i>fac</i>)
6	$[\text{Mn}(\text{tfth})_3]^b$	0.000 (<i>mer</i> Th–Th) 0.016 (<i>mer</i> CF ₃ –CF ₃) 0.025 (<i>mer</i> CF ₃ –Th) 0.090 (<i>fac</i>)
7	$[\text{Mn}(\text{tfvu})_3]$	0.000 (<i>mer</i> CF ₃ –CF ₃) 0.019 (<i>mer</i> CF ₃ –Fu) 0.055 (<i>mer</i> Fu–Fu) 0.085 (<i>fac</i>)
8	$[\text{Mn}(\text{tfba})_3]$	0.000 (<i>mer</i> CF ₃ –CF ₃) 0.023 (<i>mer</i> CF ₃ –Ph) 0.041 (<i>mer</i> Ph–Ph) 0.071 (<i>fac</i>)
9	$[\text{Mn}(\text{hfaa})_3]^a$	0.000

^a From ref. 30. ^b From ref. 26. ^c Compression Jahn–Teller input geometry spontaneously converged to elongation Jahn–Teller geometry.



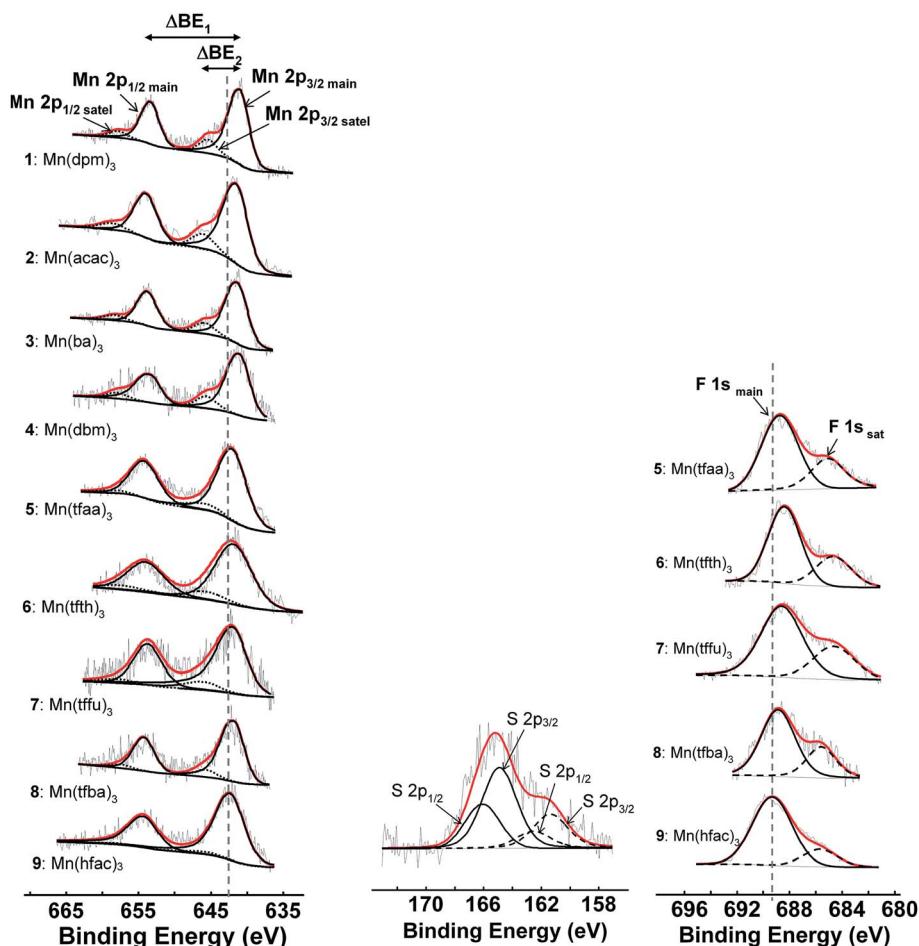


Fig. 3 Comparative XPS spectra, showing one peak fitted for the main and satellite peaks of (left) the Mn 2p_{3/2} and 2p_{1/2} photoelectron lines, (middle) the S 2p_{3/2} and 2p_{1/2} photoelectron lines, and (right) the F 1s photoelectron lines of complexes 1–9. The photoelectron lines are charge corrected against C 1s, at 284.5 eV. The grey dashed line at 642.4 eV (left) and 689.3 eV (right), illustrates the shift of the binding energy of complexes 1 to 8 relative to complex 9.

2p_{1/2} photoelectron lines are located respectively between 641.67–642.38 eV and 652.82–653.56 eV, with the presence of a satellite peak (due to the shake-up mechanism) at a binding energy of *ca.* 4 eV higher than the main Mn 2p photoelectron lines. The maximum binding energy of the main Mn 2p_{3/2} peak of all the [Mn(β-diketonato)₃] complexes 1–9, is found at *ca.* 3 eV higher than the binding energy of the Mn 2p_{3/2} of the free manganese metal Mn⁰, which is located at *ca.* 639 eV.⁴³ This 3 eV shift to a higher binding energy indicates that Mn in fact exists in oxidation state III in the complex, which is further confirmed by the *ca.* 641 eV binding energy reported for other Mn^{III} species,⁴⁴ as well as by the 641.2 eV previously reported in literature for [Mn(acac)₃], complex 2, in ref. 45.

Both the Mn 2p_{3/2} and Mn 2p_{1/2} photoelectron lines exhibit broad peaks, with a fairly high value of 4–6 eV measured for the full width at half maximum (FWHM). The FWHM of the obtained peaks is rather large in comparison with the FWHM of 2.28 eV, reported for the Mn 2p_{1/2} photoelectron line of the cubic alloy Al₆₀Pd₂₅Mn₁₅, which was measured at a pass energy of 29.5 eV,⁴⁶ and in comparison to the Mn 2p_{1/2} photoelectron line of elemental Mn⁰, which gave a FWHM of 2.52 eV

(measured at a pass energy of 35.75 eV).⁴⁷ The unexpectedly high FWHM value observed for these Mn^{III} photoelectron lines might be due to various reasons:

(i) Mn^{III} is known to exist in a high spin paramagnetic state,^{48,49} and the FWHM of high spin Mn^{III} is generally higher than that of a low spin Mn complex. For example, the FWHM of high spin Mn^{III} in Mn₂O₃ was measured as 1.75 eV, while the FWHM of low spin Mn^{VI} in MnO₂ was found to be only 0.91 eV (both measured at 20 eV pass energy).⁵⁰

(ii) A higher pass energy allows for a larger number of electrons to reach the detector, but the energy determination becomes less precise, causing the FWHM to increase. In this study a large pass energy had to be used in order to obtain well defined peaks (the same pass energy of 93.90 eV was used for all the complexes in the Mn 2p region), which contributed to the higher FWHM. For example, the FWHM of elemental Mn⁰ is 1.14 eV at 25 eV pass energy, 0.79 eV at 20 eV pass energy and 0.74 eV at 10 eV pass energy,⁵⁰ therefore obtaining smaller FWHM values with decreasing pass energy.

(iii) The larger than expected FWHM value is also an indication of more than one species most probably being present in

Table 2 Sum of the Gordy scale group electronegativities of all the R-groups, $\Sigma\chi_R = 3(\chi_R + \chi_{R'})$ of $[\text{Mn}(\text{RCOCHCOR}')_3]$. The Binding energies at the maximum binding energy (BE) of the Mn 2p_{3/2} photoelectron lines, the spin orbit splitting of maximum binding energies of the Mn 2p photoelectron lines (ΔBE_1), the binding energy of the satellite Mn 2p_{3/2} photoelectron line (BE Mn 2p_{3/2sat}), as well as the peak separations between main Mn 2p_{3/2} and the satellite Mn 2p_{3/2} photoelectron lines (ΔBE_2), the ratio area % between main Mn 2p_{3/2} and the satellite Mn 2p_{3/2} photoelectron lines (I_{ratio}). The binding energies of the maximum binding energy (BE) of the main and satellite F 1s photoelectron lines, as well as the ratio area % between main F 1s and the satellite F 1s, I_{ratio}

Complex	$\Sigma\chi_R = 3(\chi_R + \chi_{R'})$	ν_{CO} (cm ⁻¹)	BE Mn 2p _{3/2main} (eV)	ΔBE_1^b (eV)	BE Mn 2p _{3/2sat} (eV)	ΔBE_2^c (eV)	BE Mn 2p _{3/2main} <i>mer</i> (eV)	BE F 1s _{main} (eV)	BE F 1s _{sat} (eV)	BE F 1s _{sat} (eV)	I_{ratio}^g
$[\text{Mn}(\text{dpm})_3]$, 1	13.62	1591	641.67	11.88	645.87	4.20	0.16	641.49	— ^f	— ^f	— ^f
$[\text{Mn}(\text{acac})_3]$, 2	14.04	1590	641.73	11.91	646.09	4.36	0.17	— ^e	— ^f	— ^f	— ^f
$[\text{Mn}(\text{ba})_3]$, 3	13.65	1591	641.69	11.89	645.98	4.29	0.16	641.51	— ^f	— ^f	— ^f
$[\text{Mn}(\text{dbm})_3]$, 4	13.26	1592	641.67	11.80	645.99	4.32	0.17	— ^e	— ^f	— ^f	— ^f
$[\text{Mn}(\text{tfaa})_3]$, 5	16.05	1598	642.05	12.21	645.81	3.76	0.11	641.87	688.49	684.88	0.41
$[\text{Mn}(\text{tfth})_3]$, 6	15.33	1602	641.93	12.02	645.83	3.90	0.13	641.75	688.20	684.60	0.45
$[\text{Mn}(\text{tfuu})_3]$, 7	15.39	1597	641.97	12.04	645.84	3.87	0.13	641.79	688.33	684.48	0.44
$[\text{Mn}(\text{tfba})_3]$, 8	15.66	1601	641.99	12.12	645.83	3.84	0.11	641.81	688.48	684.63	0.42
$[\text{Mn}(\text{hfac})_3]$, 9	18.06	1639	642.38	12.38	645.73	3.35	0.05	— ^e	689.32	685.70	0.23

^a Sum of the Gordy scale group electronegativities, see ref. 24. ^b $\Delta\text{BE}_1 = \text{BE}_{\text{Mn } 2p_{3/2\text{main}}} - \text{BE}_{\text{Mn } 2p_{3/2\text{sat}}}$. ^c $\Delta\text{BE}_2 = \text{BE}_{\text{Mn } 2p_{3/2\text{sat}}} - \text{BE}_{\text{Mn } 2p_{3/2\text{main}}}$. ^d $I_{\text{ratio}} =$ ratio between the intensities of the satellite and main Mn 2p_{3/2} photoelectron lines = $(I_{\text{Mn } 2p_{3/2\text{sat}}})/(I_{\text{Mn } 2p_{3/2\text{main}}})$. ^e The R-groups of the ligands of the complex are equivalent, hence the complex does not show *mer* and *fac* isomers. ^f The R-groups of the ligands of the complexes do not contain F.

^g $I_{\text{ratio}} =$ ratio between the intensities of the satellite and main F 1s photoelectron lines = $(I_{\text{F } 1s_{\text{sat}}})/(I_{\text{F } 1s_{\text{main}}})$.

the sample, namely the *fac* as well as different bond stretch *mer* isomers, as discussed in the previous section.

(iv) The large FWHM value could also be attributed to final state effects, such as the coupling of the angular momenta of the unpaired electrons of the core level after photoionisation, with one of the four unpaired 3d electrons of high-spin Mn^{III} of the outer orbitals (causing multiplet splitting), as well as crystal field- and electrostatic interactions.^{51,52} Multiplet splitting is discussed in a separate section below.

2.2.1 Binding energies. The spectral deconvolution of the Mn 2p_{3/2} photoelectron lines indicates that complexes **1–9** all display different binding energies (BE_{Mn 2p_{3/2}}, see Fig. 3 and Table 2), which are influenced by their chemical environment, which in this case are the different groups attached to the β -diketonato ligands in $[\text{Mn}^{\text{III}}(\beta\text{-diketonato})_3]$ complexes **1–9**. It is known that for a fixed oxidation state, the binding energy of a metal-electron increases as the electronegativity of the ligand attached to the metal increases.⁵³ For β -diketones and metal- β -diketonato complexes, the electronegativity of the β -diketonato ligands can conveniently be expressed in terms of $\Sigma\chi_R$, which is the sum of the Gordy scale group electronegativities,²⁴ of the different R groups attached to the β -diketonato ligands ($\text{RCOCHCOR}'^-$). It has been shown that the relative electron-withdrawing or electron-donating ability of different R groups in β -diketone ligands as well as metal- β -diketonato complexes, influences chemical properties like electrochemistry,^{54,55} spectroscopic response,^{56,57} kinetics,^{58,59} turn-over frequency of catalysis,⁶⁰ and NMR peak positions⁶¹ of the β -diketone ligands and metal- β -diketonato compounds. For $[\text{Mn}^{\text{III}}(\text{RCOCHCOR}')_3]$ complexes, the sum of the Gordy scale group electronegativities is $\Sigma\chi_R = 3(\chi_R + \chi_{R'})$. Evaluation of the different binding energies of the Mn 2p_{3/2} photoelectron lines of complexes **1–9**, shows that the binding energy of the Mn 2p_{3/2} electrons increases as $\Sigma\chi_R$ in $[\text{Mn}^{\text{III}}(\text{RCOCHCOR}')_3]$ increases, see Fig. 4A. An increase in the electron withdrawing character of the β -diketonato

ligands, expressed as increasing $\Sigma\chi_R$, causes electron density to move away from the manganese metal centre, resulting in a more positive manganese metal centre. The manganese centre therefore binds more strongly to its core electrons, leading to an increase in the binding energy of the Mn 2p_{3/2} electrons.⁶⁰ The linear relationship obtained between the binding energy (BE) of the Mn 2p_{3/2} core electrons of complexes **1–9** and $\Sigma\chi_R$, the sum of the electronegativity of the different R groups in $[\text{Mn}^{\text{III}}(\text{RCOCHCOR}')_3]$, can be fitted by the following equation (Fig. 4A):

$$\text{BE}_{\text{Mn } 2p_{3/2}} = 0.1532 \Sigma\chi_R + 639.6; R^2 = 0.993 \text{ (for complexes 1–9)}$$

For comparative reasons, data of previously published $[\text{Mn}^{\text{III}}(\text{ferrocenyl-}\beta\text{-diketonato})_3]$ complexes, containing an additional Fe metal,²³ is also considered at this point. Considering published data²³ from a previous study of the $[\text{Mn}^{\text{III}}(\text{ferrocenyl-}\beta\text{-diketonato})_3]$ complexes **10–14**, containing at least one ferrocenyl group on the β -diketonato ligand (with ligands ferrocenyl- β -diketonato = $(\text{FcCOCHCOF}_3)^-$ (**10**), $(\text{FcCOCHCOCH}_3)^-$ (**11**), $(\text{FcCOCHCOPh})^-$ (**12**), $(\text{FcCOCHCOFc})^-$ (**13**) as well as (ferrocenyl- β -diketonato)₃ = $((\text{FcCOCHCOF}_3)^-)_2(\text{FcCOCHCOCH}_3)^-$ (**14**), where Fc = ferrocenyl), and including this data in the linear fit of $\text{BE}_{\text{Mn } 2p_{3/2}}$ with $\Sigma\chi_R$, it was found that the above equation stays basically the same, within experimental accuracy (see Fig. 4A), namely:

$$\text{BE}_{\text{Mn } 2p_{3/2}} = 0.1503 \Sigma\chi_R + 639.6; R^2 = 0.995 \text{ (for complexes 1–14)}$$

This result illustrates that the above equation can be used to a high degree of accuracy, to also determine the binding energy of related $[\text{Mn}(\beta\text{-diketonato})_3]$ complexes, if the electronegativity χ_R of the groups attached to the β -diketonato ligands ($\text{RCOCHCOR}'^-$), is known.



2.2.2 Spin-orbit splitting. Another important feature of the Mn 2p XPS peaks, is the difference between the maximum binding energy of the main $2p_{3/2}$ and $2p_{1/2}$ peaks, known as the spin-orbit splitting. For complexes **1–9**, a spin orbit splitting ($\Delta BE_1 = BE_{\text{Mn } 2p_{1/2\text{main}}} - BE_{\text{Mn } 2p_{3/2\text{main}}}$) was observed in the range of $\Delta BE_1 = 11.8\text{--}12.4$ eV. This is in the same order as reported data of $\Delta BE_1 = ca. 11.6$ eV, for Mn^{III} in Mn_2O_3 .^{62–64}

Spin-orbit splitting is an initial state effect, and the magnitude of the separation is dependent on the degree of spin density delocalisation of the electrons in the valence orbitals.^{65–67} Thus it is expected that a more electron withdrawing group (with a higher degree of spin density delocalisation) would be associated with a larger spin-orbit splitting. The correlation between the sum of the Gordy group electronegativities ($\Sigma \chi_R$) and the spin-orbit splitting (ΔBE_1) between the Mn $2p_{3/2}$ and Mn $2p_{1/2}$ photoelectron lines (see Fig. 4B), is in agreement with the above.

It was found that the spin-orbit splitting ΔBE_1 of the main Mn $2p_{3/2}$ and $2p_{1/2}$ peaks for complexes **1–9**, increases linearly as $\Sigma \chi_R$ increases (see Fig. 4B):

$$\Delta BE_1 = 0.1185 \Sigma \chi_R + 10.25; R^2 = 0.970 \text{ (for complexes 1–9)}$$

$$\Delta BE_1 = 0.1110 \Sigma \chi_R + 10.36; R^2 = 0.972 \text{ (for complexes 1–14)}$$

Here again we observe that enlarging the $\Sigma \chi_R$ range, by including Fe-containing ligands from the previous study²³ of $[\text{Mn}^{\text{III}}(\text{ferrocenyl-}\beta\text{-diketonato})_3]$ complexes **10–14** into the linear fit of ΔBE_1 with $\Sigma \chi_R$, does not have any significant influence on the accuracy of the fit or the linear relationship. The presence or absence of the second metal Fe in the ligand of these $[\text{Mn}^{\text{III}}(\text{diketonato})_3]$ complexes, does not influence the binding energies of the Mn 2p XPS peaks any more than merely the normal effect expected from the electronegativity of the ferrocenyl group.

2.2.3 Shake-up Mn 2p satellite peaks. In addition to the main photoelectron lines displayed by the XPS of complexes containing high-spin metals, satellite structures are often

observed. These structures are due to final state effects. They are observed either as shake-up peaks at higher binding energies than the main photoelectron line, or as shake-down peaks at lower binding energies than the main photoelectron line. The position and intensity of these satellite structures both are dependent on the ability of the electron hole that has been created due to photoemission, to be stabilised by the electronic structure of the surrounding environment of the complex. This ability is related to the intrinsic properties of the complex. Published results show that there is a link between the intensity of these shake-up or shake-down satellite peaks and the magnetic moment, μ ,⁶⁸ as well as the spin density of the corresponding complex.⁶⁹

As mentioned earlier and shown in Fig. 3, the Mn $2p_{3/2}$ shake-up satellite peak is located at *ca.* 4 eV higher binding energy than the main Mn $2p_{3/2}$ peak (see also columns 4 and 6 in Table 2). It previously has been shown for paramagnetic Cu(II) complexes, that the intensity of the satellite peak in comparison with the main metal M $2p_{3/2}$ peak, namely $I_{\text{ratio}} = (I_{\text{Mn } 2p_{3/2\text{sat}}}) / (I_{\text{Mn } 2p_{3/2\text{main}}})$, increases (non-linearly) as the spin density of the ligands attached to the Cu metal increases.⁶⁹ In this study it has been found however, that the ratio of the intensity of the satellite peak towards the main Mn $2p_{3/2}$ peak, $I_{\text{ratio}} = (I_{\text{Mn } 2p_{3/2\text{sat}}}) / (I_{\text{Mn } 2p_{3/2\text{main}}})$, is inversely proportional to $\Sigma \chi_R$, which is a measure of the electronegativity of the different ligands. This “opposite” trend will be explained below.

Satellite peaks are generally caused by one of two phenomena. In the first instance, shake-up peaks commonly observed for paramagnetic transition metals occur, when the kinetic energy of the emitted photoelectron is being reduced due to energy transfer to a non-photoelectron of the same element (causing the non-photoelectron to be in an excited state), which leads to a “shake-up” peak at a higher binding energy than the main line.⁷⁰ On the other hand, shake-up peaks also occur due to a strong configuration interaction in the final state, involving significant ligand-to-metal charge transfer instead (for example in Ni(II) and Co(II)), which results in an “extra” 3d electron, compared with the initial state.⁷¹ The excited energy transfer satellite mechanism in the first instance dominates over the latter charge-transfer satellite mechanism, when firstly the difference between the maximum binding energies of the satellite and the main metal peaks is large (namely the charge transfer energy, ΔBE_2), when secondly the core-hole-d-electron Coulomb attraction energy is small, and thirdly when there are many empty 3d orbitals.⁷² However, for the d⁴ complexes **1–9** in this study, with as many as 6 empty 3d orbitals each, the observed shake-up peaks can best be interpreted in terms of the latter instance, namely a ligand-to-metal charge transfer during photoionisation, since ΔBE_2 is only a small value (less than 5 eV). Also, in the F-containing $[\text{Mn}^{\text{III}}(\beta\text{-diketonato})_3]$ complexes **5–9**, the corresponding shake-down peaks are observed for the main F 1s photoelectron line, as well as for the main S 2p line of the S-containing complex **6** (see Section 2.2.4 below). Comparable shake-down satellites were reported in a previous study, for the photoelectron lines associated with the Fe 2p electrons of the Fe-containing ligand, for related $[\text{Mn}^{\text{III}}(\text{ferrocenyl-}\beta\text{-diketonato})_3]$ complexes.²³ In addition, the latter charge transfer model also accounts quantitatively for the observed intensities of the

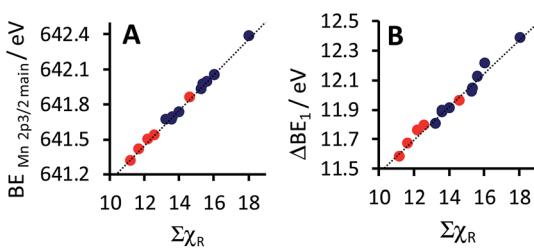


Fig. 4 Relationship between $\Sigma \chi_R$, the sum of the Gordy scale group electronegativities of the R-groups on the β -diketonato ligand of complexes **1–14**, with (A) the maximum binding energy ($BE_{\text{Mn } 2p_{3/2\text{main}}}$) of the main Mn $2p_{3/2}$ photoelectron line, as well as with (B) the spin orbit splitting of the maximum binding energy (ΔBE_1) of the main Mn $2p_{3/2}$ and Mn $2p_{1/2}$ photoelectron lines ($\Delta BE_1 = BE_{\text{Mn } 2p_{1/2\text{main}}} - BE_{\text{Mn } 2p_{3/2\text{main}}}$). Data of complexes **1–9** from this study is shown in blue (from Table 2) and data of complexes **10–14** from a previous study in red (from ref. 23).



satellites to the main peaks (I_{ratio}),⁷² as was found in this study for complexes **1–9**.

The electron configuration of Mn(III) in complexes **1–9** is $1s^2 2s^2 2p^6 3s^2 3p^6 3d^4$. However, after ligand-to-metal charge transfer during photoionisation of a 2p electron, the final state is $1s^2 2s^2 2p^5 3s^2 3p^6 3d^{4+n}$, where $n \approx 1$ denotes the amount of charge transferred from the ligands to the Mn metal.⁴⁵ Ligand-to-metal charge transfer during photoionisation, therefore leads to a change in spin density. As more charge is transferred from the ligand to the metal (leading to the “extra” 3d electron during photoionisation), the larger the satellite charge transfer peak (shake-up peak) of the metal becomes. The inversely proportional relationship observed between $\Sigma\chi_R$ and I_{ratio} (Fig. 5B) affirms that β -diketonato ligands with higher $\Sigma\chi_R$ values (for example $\Sigma\chi_R = 18.06$ for complex **9**), indicating stronger electron withdrawing properties than others, would obviously transfer less charge to the Mn metal centre during photoionisation of a 2p electron, than the ligands of, for example complex **4**, with a smaller value of $\Sigma\chi_R = 13.26$, therefore withdrawing much less electrondensity. In literature, this result also agrees with the linear increase observed for the intensity ratio, I_{ratio} , with the decrease of the Pauling electronegativity of the halides,^{73,74} for three series of metal-dihalides, namely CoX_2 , FeX_2 and MnX_2 , where $\text{X} = \text{F, Cl or Br}$. The same increase in I_{ratio} with decreasing $\Sigma\chi_R$ can be seen in the example of the Mn-dihalides series from literature: MnF_2 ($I_{\text{ratio}} = 0.06$, $\chi_F = 3.98$), MnCl_2 ($I_{\text{ratio}} = 0.28$, $\chi_F = 3.16$) and MnBr_2 ($I_{\text{ratio}} = 0.38$, $\chi_F = 2.96$).⁷² However, metal-to-ligand charge transfer, as observed for Cu(II) in literature,⁶⁹ results in the opposite trend between I_{ratio} and ligand electronegativity (than ligand-to-metal charge transfer) during photoexcitation. The linear relationship obtained between the I_{ratio} of the Mn 2p_{3/2} peaks and $\Sigma\chi_R$, the sum of the Gordy scale group electronegativities of the R-groups on the ligands, fits the following linear equation (see Fig. 5B):

$$I_{\text{ratio}} = -0.0256 \Sigma\chi_R + 0.5172; R^2 = 0.970 \text{ (for complexes 1–9)}$$

$$I_{\text{ratio}} = -0.0248 \Sigma\chi_R + 0.5040; R^2 = 0.971 \text{ (for complexes 1–14)}$$

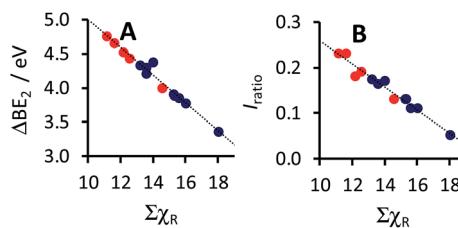


Fig. 5 Relationship between $\Sigma\chi_R$, the sum of the Gordy scale group electronegativities of the R-groups on the β -diketonato ligand, of complexes **1–14**, with (A) the charge transfer energy (ΔBE_2), namely the difference between the maximum binding energy of the main Mn 2p_{3/2} photoelectron line and the satellite Mn 2p_{3/2} photoelectron line ($\Delta BE_2 = BE_{\text{Mn} 2p_{3/2\text{sat}}} - BE_{\text{Mn} 2p_{3/2\text{main}}}$), as well as with (B) the intensity ratio (I_{ratio}) of the satellite to the main Mn 2p_{3/2} photoelectron lines, $I_{\text{ratio}} = (I_{\text{Mn} 2p_{3/2\text{sat}}}) / (I_{\text{Mn} 2p_{3/2\text{main}}})$. Data of complexes **1–9** from this study is shown in blue (from Table 2) and data of complexes **10–14** from literature in red (from ref. 23).

2.2.4 Shake-down F 1s and S 2p satellite peaks. Detail of the spectral deconvolution of the F 1s photoelectron line, present in the XPS of the five CF_3 -containing complexes **5–9**, is shown in Fig. 3. Similarly to the Mn 2p photoelectron lines, the binding energy position of the F 1s photoelectron line, is also influenced by the sum of the Gordy scale group electronegativities, $\Sigma\chi_R$, of the R-groups attached to the β -diketonato ligand. Again, (as in Fig. 4A for the main Mn 2p_{3/2} photoelectron line), as the electronegativity of the ligands, $\Sigma\chi_R$, increases, the binding energy ($BE_{\text{F} 1s}$) of the F 1s photoelectron line also increases, see Fig. 6A. The linear relationship obtained between the binding energy of the main F 1s photoelectron line ($BE_{\text{F} 1s\text{main}}$), and the sum of the electronegativity of the different R groups, $\Sigma\chi_R$, of the five fluorine-containing complexes in this study, **5–9**, can be fitted by the following linear equation:

$$BE_{\text{F} 1s} = 0.3834 \Sigma\chi_R + 682.39; R^2 = 0.979 \text{ (for the F-containing complexes 5–9)}$$

The substructure of the F 1s photoelectron lines displays a strong shake-down satellite peak, *ca.* 3.7 eV lower than the main F 1s photoelectron line. This shake-down peak of the main F 1s photoelectron line is associated with the corresponding shake-up peak of the Mn 2p photoelectron, which depicts the ligand-to-metal charge transfer during photoionisation. The shake-down peak of the F 1s photoelectron line can be considered representative of some of the charge transfer from the F-containing ligand to the Mn metal.

The I_{ratio} of the F 1s peaks ($I_{\text{ratio}} = (I_{\text{F} 1s_{\text{sat}}} / (I_{\text{F} 1s_{\text{main}}}))$) is a measure of the amount of charge transferred from the CF_3 -containing ligand to the Mn metal of these five complexes **5–9**. The intensity ratio, and thus the amount of charge transferred from the fluorine-containing ligand to the Mn metal decreases, as the electronegativity of the CF_3 -containing ligand ($\Sigma\chi_R$) increases (see Fig. 6B). This inversely proportional linear relationship obtained between the I_{ratio} of the satellite to the main F 1s peaks with electronegativity $\Sigma\chi_R$, fits the following equation (see Fig. 5B):

$$I_{\text{ratio}} = -0.0798 \Sigma\chi_R + 1.674; R^2 = 0.989 \text{ (for the F-containing complexes 5–9)}$$

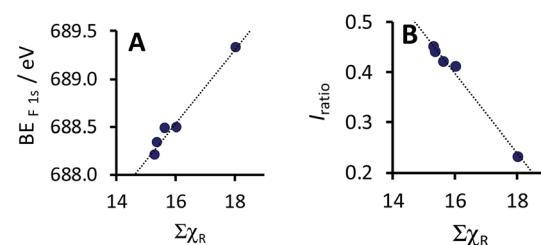


Fig. 6 Relationship between $\Sigma\chi_R$, the sum of the Gordy scale group electronegativities of the R-groups on the β -diketonato ligand, of the five CF_3 -containing complexes **5–9**, with (A) the maximum binding energy ($BE_{\text{F} 1s\text{main}}$) of the main F 1s photoelectron line, as well as with (B) the intensity ratio (I_{ratio}) of the satellite and main F 1s photoelectron lines, $I_{\text{ratio}} = (I_{\text{F} 1s_{\text{sat}}}) / (I_{\text{F} 1s_{\text{main}}})$. Data from Table 2.



The above result implies that the symmetrically substituted ligands in complex **9**, containing as much as six strongly electron withdrawing CF_3 groups, will subsequently transfer even less charge to the Mn metal in complex **9**, than in complexes **5–8**, which each contain only three electron withdrawing CF_3 groups. This observed result agrees with the lowest amount of charge “received” by Mn in complex **9** (Fig. 5B), as given by $I_{\text{ratio}} = (I_{\text{Mn } 2\text{p}_{3/2\text{sat}}} / (I_{\text{Mn } 2\text{p}_{3/2\text{main}}}) = 0.05$ (see the point farthest to the right in Fig. 5B).

The observed ligand-to-metal charge transfer during photoionisation, occurs from the β -diketonato ligands to Mn. Therefore charge is also transferred from the other atoms on the β -diketonato ligand to the Mn metal. For example, in complex **6**, containing sulphur in the groups substituted on the ligand, the main S 2 $\text{p}_{3/2}$ peak at 164.8 eV also exhibits a shake-down charge transfer satellite peak at 161.5 eV, with a spin-orbit splitting (ΔBE_1) of 3.3 eV, as seen in Fig. 3. It however was not possible to also identify the satellite peaks associated with the C and O lines in complexes **1–9**, due to the presence of adventitious C and O.

2.2.5 Fitting of *fac* and *mer* isomers. As mentioned previously, the large FWHM (full width at half maximum) obtained for the Mn 2p spectral lines, could be an indication that more than one species is present in the sample. The FWHM obtained for the unsymmetrical substituted complexes **3, 5–8**, for which both *fac* and *mer* isomers are possible, is slightly larger (namely 4.9, 5.3, 6.2, 5.1 and 4.7 eV respectively) than the FWHM obtained for complexes **1, 2, 4** and **9** (3.9, 4.4, 4.5 and 5.1 eV respectively), containing symmetrically substituted β -diketonato ligands. Two peaks, one for the three *mer* bond stretch isomers and one for the *fac* isomer, can successfully be fitted into the original Mn 2p peak. The fitting of these simulated peaks into the original Mn 2p peak, results in a much smaller FWHM for the unsymmetrical complexes **3, 5–8** (of 3.8, 4.1, 4.7, 3.9 and 3.6 eV respectively). These fitted values are then comparable to the original FWHM obtained for the symmetrical complexes **1, 2, 4** and **9** (in the range of 3.9–5.1 eV). See Fig. 7, which depicts the fitting of complex **3**, as an example of the unsymmetrical complexes. Additionally, the CHI squared value (indicating the goodness of fit) decreased, when two main peaks instead of one were fitted for the unsymmetrical complexes **3, 5–8**. For example, for complex **3**, CHI squared decreased from 1.6 to 1.5 when fitting two main peaks, thereby accommodating both the *fac* and the *mer* isomers. The DFT calculations in Section 2.1.1 above, showed that the *fac* isomer is generally less stable than the *mer* isomer, and also that three *mer* bond stretch isomers are possible for unsymmetrical complexes **3, 5–8**. Since the statistical distribution between the three different *mer* bond stretch isomers and the one *fac* isomer is 3 : 1, the two Gaussian peaks were fitted in a ratio of 3 : 1 into the original Mn 2p peaks for the unsymmetrical complexes **3, 5–8**. The Mn(III) β -diketonato complexes containing unsymmetrically substituted β -diketonato ligands (**3, 5–8**), were further simulated to have two Mn main $2\text{p}_{3/2}$ peaks, as well as two satellite peaks; one satellite structure is associated with the three more stable *mer* bond stretch isomers (at lower energy) and the other satellite peak with the less stable *fac* isomer (at higher energy). The difference in binding energy between the *mer* and the *fac* isomers is *ca.* 0.01 eV.

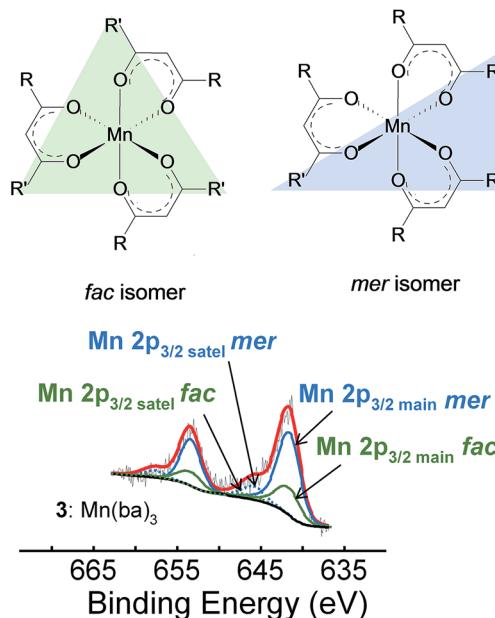


Fig. 7 XPS spectra of the unsymmetrical complex **3** (as a representative), showing the two simulated peaks, as an example for the complexes with unsymmetrical β -diketonato ligands, representing both the *mer* and *fac* isomers (a ratio of 3 : 1 was simulated), as well as the satellite shake-up peaks of the Mn 2p area at a higher binding energy.

The *mer* and *fac* simulation of complex **3** is shown in Fig. 7, as an example for the other unsymmetrical complexes.

However, the FWHM obtained above is still larger than expected. A more accurate fitting of the peaks due to additional multiplet splitting, commonly observed in high spin paramagnetic complexes, is therefore needed.

2.2.6 Multiplet splitting. Interaction between an unpaired 2p electron (originating after photoemission) and an unpaired 3d electron, which aligns anti-parallel to the unpaired 2p electron, causes additional splitting of the original Mn 2p photoelectron lines (also contributing to the increased FWHM) of complexes **1–9**.⁵⁰ The photoelectron lines of the Mn 2p presented in this study, are therefore even more complex than described thus far, because of the additional effect of multiplet splitting.^{75,76} Gupta and Sen^{51,52} theoretically calculated the splitting of the 2p peaks of a high-spin free Mn^{3+} ion, to be in a ratio area percentage of 1 : 1 : 1.35 : 0.7 : 0.3 for the five splitted 2p peaks. Fig. 8 shows a simulation of the theoretical multiplet splitting ratio as calculated by Gupta and Sen,^{51,52} simulated for both complex **2** (representative of complexes **1, 2, 4** and **9** with symmetric ligands), as well as for complex **3** (representative of complexes **3, 5–8** with unsymmetric ligands, in this case fitting both the *fac* and *mer* isomers). These multiplet splitting peaks consist of five components (labelled peaks 1–5 for symmetrical complex **2**, as well as f1–f5 (*fac*) and m1–m5 (*mer*) for the unsymmetrical complex **3**). For complex **2**, the CHI squared value decreased from 1.8 to 1.3 when multiplet peaks instead of one single peak were fitted for the Mn 2p_{3/2} photoelectron line, indicating a good fit. The CHI squared value for unsymmetrical complex **3**, also decreased further from 1.5 (where merely two main peaks for the *fac* and *mer* isomers



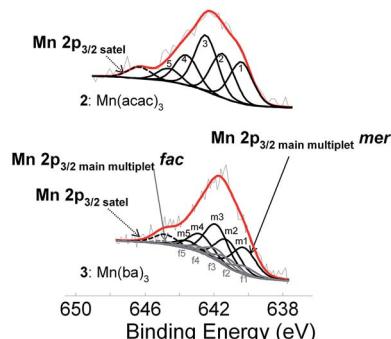


Fig. 8 Detailed Mn 2p_{3/2} area of complex 2 (top, showing the simulated multiplet splitting of the symmetrical complex 2), depicting the five components, labelled peaks 1–5; as well as the detailed Mn 2p_{3/2} area of complex 3 (bottom, showing the simulated multiplet splitting of the *fac* (grey line) and the *mer* (black line) isomers of the unsymmetrical complex 3). In the latter complex 3, the peaks labelled m1–m5 above the spectra belong to the *mer* isomer, while the five components of the peaks labelled f1–f5 below the spectra, belong to the *fac* isomer. No multiplet splitting, or separate *fac* or *mer* peaks, were fitted to the satellite peak.

were fitted for complex 3) to a CHI squared value of 1.3, when also applying the multiplet splitting peak fitting. The overall FWHM of the Mn 2p_{3/2} photoelectron line when using multiplet splitting, decreased (from the original range of 3.9–6.2 eV, before applying any simulation, see Section 2.2.5) to a range of 1.4–1.8 eV, which is in the same order as was found for related free Mn ions: For example, the calculated Gupta and Sen multiplet splitting distribution for Mn²⁺ and Mn³⁺ free ions, when applied to Mn²⁺ in MnO, and to Mn³⁺ in manganite, resulted in FWHM values of the multiplet peaks of 1.7 for Mn²⁺ and 1.25 eV for Mn³⁺,⁷⁷ which are similar to the above improved FWHM values for the Mn 2p_{3/2} peaks (range of 1.4–1.8 eV), after fitting the multiplet splitting.

3 Experimental section

3.1 Synthesis

A series of nine β -diketonato manganese(III) complexes of formula [Mn(RCOCHCOR')₃], with symmetrically substituted groups R = C(CH₃)₃ [1]; (CH₃) [2]; (C₆H₅) [4] and (CF₃) [9], or of formula [Mn(RCOCHCOR')₃], with unsymmetrically substituted groups R = (CH₃) and R' = (C₆H₅) [3]; and also R = (CF₃) and R' = (CH₃) [5], (OC₄H₉S) [6], (OC₄H₉S) [7] or (C₆H₅) [8], were prepared and characterised according to published methods.⁷⁸

3.2 XPS measurements

The X-ray photoelectronic spectroscopic (XPS) analysis was carried out with a PHI 5000 Versaprobe, using monochromatic Al K α X-ray radiation, with $h\nu = 1486.6$ eV. The hemispherical analyser pass energy, used to record the high resolution spectra, was maintained at 29.35 eV for C 1s and O 1s, and at 93.90 eV for Mn 2p, S 2p, and F 1s, in order to obtain peaks with good resolution, using 1 eV per step. In order to acquire a neutral charge on the surface of the sample, a low energy electron beam was utilised. All the binding energies of the photoelectron

Table 3 DFT calculated energies of the indicated spin states of [Mn(acac)₃] optimized with the indicated functionals

Spin	BLYP	OLYP	B3LYP	CAM-B3LYP ^a	M06-D3 ^a	M065X ^a
1	0.000	0.33	0.27	0.27	0.92	0.91
2	0.232	0.00	0.00	0.00	0.00	0.00

^a Optimized using the Gaussian 09 program package⁹⁰ and the 6-311G(d,p) basis set.

spectra were referenced against the lowest binding energy of the C–C simulated adventitious C 1s photoelectron line, which was set at 284.8 eV. The XPS data was analysed, utilising the Multipak version 8.2c computer software.⁷⁹

3.3 DFT calculations

Density functional theory (DFT) calculations were conducted with the ADF (Amsterdam Density Functional) 2013 programme.⁸⁰ The GGA (Generalised Gradient Approximation) functional OLYP (Handy-Cohen and Lee-Yang-Parr),^{81–84} a all-electron TZP (Triple ζ polarised) basis set, and spin-unrestricted open shell calculations were used for geometry optimisations. The OLYP functional proved to be a good choice for correctly calculating the ground state of high spin 3d metal-complexes.^{26,85–89} To further support the reliability of the computational method used in this work, the $S = 1$ (two unpaired electrons) and $S = 2$ (four unpaired electrons) spin states of [Mn(acac)₃] as representative example of [Mn(β -diketonato)₃] complexes, were optimized using a selection of functionals, see Table 3. These results indicate that the OLYP, B3LYP, CAM-B3LYP (B3LYP with long-range correction), M06-D3 (M06 with the Grimme empirical dispersion correction) and M065X functionals correctly predicted the experimental high state of $S = 2$ of [Mn(acac)₃], in agreement with experimental EPR studies.²⁸

4 Conclusions

A DFT study of the energy and character of the Frontier orbitals of [Mn(β -diketonato)₃] complexes shows that a highest occupied molecular orbital (HOMO) of mainly d_{z^2} character is responsible for the elongation of two Mn-ligand bonds along the z -direction. [Mn(β -diketonato)₃] complexes containing an unsymmetrically substituted β -diketonato ligand, can have one *fac* and three *mer* bond stretch isomers. An XPS study of the Mn 2p core electron binding energies, provides insight into the influence of the electronic environment caused by the electron donating/withdrawing properties of the β -diketonato ligands, on the binding energy strength. The main and satellite peaks of the Mn 2p XPS photoelectron lines were fitted with multiplet splitting peaks, as calculated for the free Mn(III) ion by Gupta and Sen, as well as with satellite shake-up peaks, due to charge transfer. Two sets of multiplet splitting peaks, representing the *mer* and *fac* isomers, were simulated for the Mn 2p_{3/2} photoelectron lines of those [Mn(β -diketonato)₃] complexes containing unsymmetrically substituted β -diketonato ligands. The relative intensity of the

satellite peak of both the Mn 2p signal of the metal and the F 1s signal in the five fluorine-containing ligands, is related to the amount of charge transferred between the β -diketonato ligands and the Mn^{III} metal centre.

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References

- (a) M. Viciano-Chumillas, S. Tanase, I. Mutikainen, U. Turpeinen, L. J. de Jongh and J. Reedijk, *Inorg. Chem.*, 2008, **47**, 5919; (b) P. Artus, C. Boskovic, J. Yoo, W. E. Streib, L.-C. Brunel, D. N. Hendrickson and G. Christou, *Inorg. Chem.*, 2001, **40**, 4199.
- J. Yoo, A. Yamaguchi, M. Nakano, J. Krzystek, W. E. Streib, L.-C. Brunel, H. Ishimoto, G. Christou and D. N. Hendrickson, *Inorg. Chem.*, 2001, **40**, 4604.
- G. E. Granroth, M. W. Meisel, M. Chaparala, T. Jolicœur, B. H. Ward and D. R. Talham, *Phys. Rev. Lett.*, 1996, **77**, 1616.
- J. S. Miller, C. Vazquez, N. L. Jones, R. S. McLean and A. J. Epstein, *J. Mater. Chem.*, 1995, **5**, 707.
- A. L. Barra, A. Caneschi, D. Gatteschi and R. Sessoli, *J. Am. Chem. Soc.*, 1995, **117**, 8855.
- J. S. Miller, C. Vazquez, J. C. Calabrese, R. S. McLean and A. J. Epstein, *Adv. Mater.*, 1994, **6**, 217.
- J. S. Miller, J. C. Calabrese, R. S. McLean and A. J. Epstein, *Adv. Mater.*, 1992, **4**, 498.
- A. Caneschi, D. Gatteschi, R. Sessoli, A. L. Barra, L. C. Brunel and M. Guillot, *J. Am. Chem. Soc.*, 1991, **113**, 5873.
- (a) R. K. Sodhi and S. Paul, *Catal. Lett.*, 2011, **141**, 608; (b) Y.-F. Wang and S. Chiba, *J. Am. Chem. Soc.*, 2009, **131**, 12570.
- R. van Gorkum, E. Bouwman and J. Reedijk, *Inorg. Chem.*, 2004, **43**, 2456.
- R. I. Khusnutdinov, N. A. Shchadneva, A. R. Baiguzina, Y. Y. Lavrentieva and U. M. Dzhemilev, *Russ. Chem. Bull.*, 2002, **51**, 2074.
- (a) J. R. Bryant, J. E. Taves and J. M. Mayer, *Inorg. Chem.*, 2002, **41**, 2769; (b) K. A. Campbell, M. R. Lashley, J. K. Wyatt, M. H. Nantz and R. D. Britt, *J. Am. Chem. Soc.*, 2001, **123**, 5710.
- P. Magnus, A. H. Payne, M. J. Waring, D. A. Scott and V. M. Lynch, *Tetrahedron Lett.*, 2000, **41**, 9725.
- M. J. S. Dewar and T. Nakaya, *J. Am. Chem. Soc.*, 1968, **90**, 7134.
- S. Carli, E. Benazzi, L. Casarin, T. Bernardi, V. Bertolasi, R. Argazzi, S. Caramori and C. A. Bignozzi, *Phys. Chem. Chem. Phys.*, 2016, **18**, 5949.
- I. R. Perera, A. Gupta, W. Xiang, T. Daeneke, U. Bach, R. A. Evans, C. A. Ohlin and L. Spiccia, *Phys. Chem. Chem. Phys.*, 2014, **16**, 12021.
- K. A. Campbell, D. A. Force, P. J. Nixon, F. Dole, B. A. Diner and R. D. Britt, *J. Am. Chem. Soc.*, 2000, **122**, 3754.
- R. D. Britt, J. M. Peloquin and K. A. Campbell, *Annu. Rev. Biophys. Biomol. Struct.*, 2000, **29**, 463.
- K. A. Campbell, E. Yikilmaz, C. V. Grant, W. Gregor, A.-F. Miller and R. D. Britt, *J. Am. Chem. Soc.*, 1999, **121**, 4714.
- I. Fridovich, *Annu. Rev. Biochem.*, 1995, **64**, 97.
- K. M. Faulkner, S. I. Liochev and I. Fridovich, *J. Biol. Chem.*, 1994, **269**, 23471.
- E. Paparazzo, *J. Electron Spectrosc. Relat. Phenom.*, 2006, **154**, 38.
- B. E. Buitendach, E. Erasmus, J. W. Niemantsverdriet and J. C. Swarts, *Molecules*, 2016, **21**, 1427.
- For a comprehensive list of Gordy scale group electronegativities related to β -diketonato R-groups, see J. Conradie, *J. Chem. Soc., Dalton Trans.*, 2015, **44**, 1503.
- R. Freitag, T. J. Muller and J. Conradie, *J. Chem. Crystallogr.*, 2014, **44**, 352.
- R. Gostynski, P. H. van Rooyen and J. Conradie, *J. Mol. Struct.*, 2016, **1119**, 48.
- R. Freitag and J. Conradie, *J. Chem. Educ.*, 2013, **90**, 1692.
- S. L. Dexheimer, J. W. Gohdes, M. K. Chan, K. S. Hagen, W. H. Armstrong and M. P. Klein, *J. Am. Chem. Soc.*, 1989, **111**, 8923.
- S. Geremia and N. Demitri, *J. Chem. Educ.*, 2005, **82**, 460.
- J. Conradie, *Comput. Theor. Chem.*, 2016, **1087**, 1.
- Cambridge Structural Database (CSD), Version 5.36, May 2015 update. CSD, reference codes: ACACMN21, ACACMN22, ACACMN22, ACACMN23, ACACMN24, JINPIF, JINPIF01, NOSHUA, QAYYEU, XESPIW.
- P. Magnus, A. H. Payne, M. J. Waring, D. A. Scott and V. Lynch, *Tetrahedron Lett.*, 2000, **41**, 9725.
- J. P. Fackler Jr and A. Avdeef, *Inorg. Chem.*, 1974, **13**, 1864.
- J. Krzystek, G. J. Yeagle, J.-H. Park, M. W. Meisel, R. D. Britt, L.-C. Brunel and J. Telser, *Inorg. Chem.*, 2003, **42**, 4610.
- B. E. Buitendach, E. Erasmus, M. Landman, J. W. Niemantsverdriet and J. C. Swarts, *Inorg. Chem.*, 2016, **55**, 1992.
- J. W. Niemantsverdriet, Microscopy and Imaging, in *Spectroscopy in Catalysis: An Introduction*, Wiley-VCH Verlag GmbH & Co. KGaA, 3rd edn, Weinheim, Germany, 2007.
- I. L. Eremenko, S. E. Nefedov, A. A. Sidorov, M. A. Golubnichaya, P. V. Danilov, V. I. Ikorskii, Y. G. Shvedenkov, V. M. Novotortsev and I. I. Mioseev, *Inorg. Chem.*, 1999, **38**, 3764.
- P. Pfluger and G. B. Street, *J. Chem. Phys.*, 1984, **80**, 544.
- C. Battistoni, J. L. Dormann, D. Fiorani, E. Paparazzo and S. Viticoli, *Solid State Commun.*, 1981, **39**, 581.
- A. E. Bocquet, *Phys. Rev. B: Condens. Matter Mater. Phys.*, 1992, **46**, 3771.
- A. V. Naumkin, T. M. Ivanova, A. V. Shchukarev, A. A. Sidorov, M. A. Kiskin, V. M. Novotortsev and I. L. Eremenko, *Russ. J. Inorg. Chem.*, 2008, **53**, 1614.
- A. E. Bocquet, *J. Electron Spectrosc. Relat. Phenom.*, 1996, **82**, 87.
- V. DiCastro and G. Polzonetti, *J. Electron Spectrosc. Relat. Phenom.*, 1989, **48**, 117.



44 L.-T. Chang, C.-Y. Wang, J. Tang, T. Nie, W. Jiang, C.-P. Chu, S. Aranfin, L. He, M. Afsal, L.-J. Chen and K. L. Wang, *Nano Lett.*, 2014, **14**, 1823.

45 A. J. Nelson, J. G. Reynolds and J. W. Roos, *J. Vac. Sci. Technol. A*, 2000, **18**, 1072.

46 C. J. Jenks, S.-L. Chang, J. W. Anderegg, P. A. Thiel and D. W. Lynch, *Phys. Rev. B: Condens. Matter Mater. Phys.*, 1996, **54**, 6301.

47 A. R. Chourasia and D. R. Chopra, *Surf. Sci. Spectra*, 1995, **3**, 74.

48 S. Wang, W.-R. He, M. Ferbinteanu, Y.-H. Li and W. Huang, *Polyhedron*, 2013, **52**, 1199.

49 B. J. Kennedy and K. S. Murry, *Inorg. Chem.*, 1985, **24**, 1552.

50 M. C. Biesinger, B. P. Payne, A. P. Grosvenor, L. W. M. Lau, A. R. Gerson and R. S. C. Smart, *Appl. Surf. Sci.*, 2011, **257**, 2717.

51 R. P. Gupta and S. K. Sen, *Phys. Rev. B: Solid State*, 1974, **10**, 71.

52 R. P. Gupta and S. K. Sen, *Phys. Rev. B: Solid State*, 1975, **12**, 15.

53 J. W. Niemantsverdriet, *Spectroscopy in Catalysis*, John Wiley & Sons, 3rd edn, Weinheim, 2007.

54 E. Erasmus, J. Conradie, A. Muller and J. C. Swarts, *Inorg. Chim. Acta*, 2007, **360**, 2277.

55 E. Erasmus and J. C. Swarts, *New J. Chem.*, 2013, **37**, 2862.

56 R. Liu, J. Conradie and E. Erasmus, *J. Electron Spectrosc. Relat. Phenom.*, 2016, **206**, 46.

57 J. Conradie and E. Erasmus, *Polyhedron*, 2016, **119**, 142.

58 J. J. C. Erasmus and J. Conradie, *Dalton Trans.*, 2013, **42**, 8655.

59 M. M. Conradie, J. J. C. Erasmus and J. Conradie, *Polyhedron*, 2011, **30**, 2345.

60 M. M. Conradie, J. Conradie and E. Erasmus, *Polyhedron*, 2014, **79**, 52.

61 A. Kuhn and J. Conradie, *Electrochim. Acta*, 2010, **56**, 257.

62 H. Cao, X. Wu, G. Wang, J. Yin, G. Yin, F. Zhang and J. Liu, *J. Phys. Chem. C*, 2012, **116**, 21109.

63 M. Salavati-Niasari, F. Mohandes, F. Davar and K. Saberyan, *Appl. Surf. Sci.*, 2009, **256**, 1476.

64 M. Liu, G.-J. Zhang, Z.-R. Shen, P.-C. Sun, D.-T. Ding and T.-H. Chen, *Solid State Sci.*, 2009, **11**, 118.

65 J. C. Carver, G. K. Schweitzer and T. A. Carlson, *J. Chem. Phys.*, 1972, **57**, 973.

66 T. M. Ivanova, A. V. Shchukarev, A. V. Naumkin, A. A. Sidorov, M. A. Kiskin, V. M. Novotortsev and I. L. Eremenko, *Russ. J. Inorg. Chem.*, 2008, **53**, 1929.

67 E. Agostinelli, C. Battistoni, D. Fiorani, G. Mattogno and M. Nogues, *J. Phys. Chem. Solids*, 1989, **50**, 269.

68 Y. G. Borod'ko, S. I. Vetchinkin, S. L. Zimont, I. N. Ivleva and Y. M. Shul'ga, *Chem. Phys. Lett.*, 1976, **42**, 264.

69 M. S. Ioffe and Y. G. Borod'ko, *J. Electron Spectrosc. Relat. Phenom.*, 1977, **11**, 235.

70 A. P. Grosvenor, B. A. Kobe, M. C. Biesinger and N. S. McIntyre, *Surf. Interface Anal.*, 2004, **36**, 1564.

71 J. S. H. Q. Perera, D. C. Frost and C. A. McDowell, *J. Chem. Phys.*, 1980, **72**, 5151.

72 J. Park, S. Ryu, M.-S. Han and S.-J. Oh, *Phys. Rev. B: Condens. Matter Mater. Phys.*, 1988, **37**, 10867.

73 L. Pauling, *J. Am. Chem. Soc.*, 1932, **54**, 3570.

74 L. Pauling, *Nature of the Chemical Bond*, Cornell University Press, 1960.

75 E. Ünveren, E. Kemnitz, S. Hutton, A. Lippitz and W. E. S. Unger, *Surf. Interface Anal.*, 2004, **36**, 92.

76 A. P. Grosvenor, M. C. Biesinger, R. S. C. Smart and N. S. McIntyre, *Surf. Sci.*, 2006, **600**, 1771.

77 H. W. Nesbitt and D. Banerjee, *Am. Mineral.*, 1998, **83**, 305.

78 R. Freitag and J. Conradie, *Electrochim. Acta*, 2015, **158**, 418.

79 F. Moulder, W. F. Stickle, P. E. Sobol and K. D. Bomben, *Handbook of X-ray Photoelectron Spectroscopy*, ULVAC-PHI, Inc., Enzo, Chigasaki, Japan, 1995.

80 G. te Velde, F. M. Bickelhaupt, E. J. Baerends, C. F. Guerra, S. J. A. van Gisbergen, J. G. Snijders and T. Ziegler, *J. Comput. Chem.*, 2001, **22**, 931.

81 N. C. Handy and A. J. Cohen, *Mol. Phys.*, 2001, **99**, 403.

82 C. Lee, W. Yang and R. G. Parr, *Phys. Rev. B: Condens. Matter Mater. Phys.*, 1998, **37**, 785.

83 B. G. Johnson, P. M. W. Gill and J. A. Pople, *J. Chem. Phys.*, 1993, **98**, 5612.

84 T. V. Russo, R. L. Martin and P. J. Hay, *J. Chem. Phys.*, 1994, **101**, 7729.

85 M. Swart, A. R. Groenhof, A. W. Ehlers and K. Lammertsma, *J. Phys. Chem. A*, 2004, **108**, 5479.

86 J. Conradie and A. Ghosh, *J. Phys. Chem. B*, 2007, **111**, 12621.

87 J. Conradie and A. Ghosh, *J. Chem. Theory Comput.*, 2007, **3**, 689.

88 M. M. Conradie and J. Conradie, *Electrochim. Acta*, 2015, **152**, 512.

89 J. Conradie, T. Wondimagegn and A. Ghosh, *J. Phys. Chem. B*, 2008, **112**, 1053.

90 M. J. Frisch, G. W. Trucks, H. B. Schlegel, G. E. Scuseria, M. A. Robb, J. R. Cheeseman, G. Scalmani, V. Barone, B. Mennucci, G. A. Petersson, H. Nakatsuji, M. Caricato, X. Li, H. P. Hratchian, A. F. Izmaylov, J. Bloino, G. Zheng, J. L. Sonnenberg, M. Hada, M. Ehara, K. Toyota, R. Fukuda, J. Hasegawa, M. Ishida, T. Nakajima, Y. Honda, O. Kitao, H. Nakai, T. Vreven, J. A. Montgomery jr, J. E. Peralta, F. Ogliaro, M. Bearpark, J. J. Heyd, E. Brothers, K. N. Kudin, V. N. Staroverov, T. Keith, R. Kobayashi, J. Normand, K. Raghavachari, A. Rendell, J. C. Burant, S. S. Iyengar, J. Tomasi, M. Cossi, N. Rega, J. M. Millam, M. Klene, J. E. Knox, J. B. Cross, V. Bakken, C. Adamo, J. Jaramillo, R. Gomperts, R. E. Stratmann, O. Yazyev, A. J. Austin, R. Cammi, C. Pomelli, J. W. Ochterski, R. L. Martin, K. Morokuma, V. G. Zakrzewski, G. A. Voth, P. Salvador, J. J. Dannenberg, S. Dapprich, A. D. Daniels, O. Farkas, J. B. Foresman, J. V. Ortiz, J. Cioslowski and D. J. Fox, *Gaussian 09, Revision D.01*, Gaussian Inc., Wallingford CT, 2010.

