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Highly permeable ionic liquid 1-butyl-3-methylimidazolium tetrafluoroborate (BMIMBF₄)/CuO composite membrane for CO₂ separation

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An ionic liquid (IL) 1-butyl-3-methylimidazolium tetrafluoroborate (BMIM⁺BF₄⁻)/CuO composite was prepared for CO₂ transport membranes. CuO nanoparticles (NPs) were successfully generated in BMIM⁺BF₄⁻. The formation of CuO NPs was ascribed to the favorable interaction between the ions of BMIM⁺BF₄⁻ and the surface of the CuO particles. To confirm the coordination behavior of the BMIM⁺BF₄⁻/CuO membrane, TGA and FT-Raman spectroscopy were employed. The ideal CO₂/N₂ selectivity in the case of the composite material was 21.0, with a CO₂ permeance of 52.4 GPU, while the neat BMIM⁺BF₄⁻ membrane showed a selectivity of 5.0 and a CO₂ permeance of 17.0 GPU. The enhanced separation performance was attributed to (1) the synergistic effect of well-dispersed CuO nanoparticles, resulting in enhanced CO₂ solubility due to the increased surface of the oxide particle layer and (2) the abundant free ions in BMIM⁺BF₄⁻.

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1. Introduction

Presently, one of the greatest challenges for researchers is finding a thermodynamically efficient process for separating and capturing CO₂.¹ It has been commonly accepted that global warming is related directly to the increased levels of CO₂.²⁻⁴ CO₂ is mainly emitted as a result of the combustion of fossil fuels, certain chemical plant processes, and production of synthesis gas.⁵ Currently, the principal CO₂ separation technologies for post-combustion capture include physical adsorption, physical or chemical absorption, cryogenic distillation, and membrane separation.^{6,7} Among the various types of CO₂ separation technologies, membrane technology promises excellent separation performance while being relatively less energy- and capital intensive.⁸⁻¹¹ However, current commercial polymeric membranes, such as polyimide and cellulose acetate membranes, have low CO₂ selectivity and permeability, since they are based on a solution-diffusion mechanism and rely chiefly on the subtle size differences of the penetrants to accomplish separation.^{12,13} To overcome the low separation performance of polymeric membranes, facilitated transport was introduced so that higher gas permeability and outstanding selectivity could be achieved.¹⁴⁻¹⁶ The facilitated transport membrane has attracted attention due to the reversible reaction with a specific gas component.¹⁷ Hence, the reaction in the membrane can generate another transport mechanism alongside the simple solution-diffusion mechanism.^{18,19}

Given the above, ionic liquids (ILs) have been utilized to prepare facilitated CO₂ transport membranes.^{20,21} ILs have attracted great attention because of their highly tunable properties such as low flammability, high chemical and thermal stability, and extremely low volatility.²² Moreover, ILs based on imidazolium cations present outstanding CO₂ separation performance since the amine group in the imidazolium ring acts a CO₂ carrier in CO₂/CH₄ and CO₂/N₂ separation.²³ For example, CO₂ separation performance utilizing 1-butyl-3-methylimidazolium tetrafluoroborate (BMIM⁺BF₄⁻) showed that CO₂/N₂ and CO₂/CH₄ ideal selectivity were 5 and 4.8, respectively.²⁴ Surprisingly, the CO₂ permeance was 17 GPU.²⁴ Furthermore, in the case of CO₂ separation utilizing pure 1-butyl-3-methylimidazolium nitrate (BMIM⁺NO₃⁻), the ideal CO₂/N₂ selectivity and CO₂ permeance were 4.7 and 5.2 GPU, respectively.²⁵

To enhance the CO₂ separation performance of neat ILs, our group investigated selective CO₂ separation utilizing ILs with lithium tetrafluoroborate (LiBF₄).²⁶ It was established that the facilitated transport membrane consisting of lithium tetrafluoroborate as a CO₂ carrier and BMIM⁺BF₄⁻ exhibits an ideal selectivity of 8.40 for CO₂/N₂ with a CO₂ permeance of 13.36 GPU.²⁶ Lee *et al.* reported that a 1-methyl-3-octylimidazolium tetrafluoroborate (MOIM⁺BF₄⁻)/positively polarized Cu nanoparticle membrane shows CO₂/N₂ and CO₂/CH₄ selectivities of 25.2 and 24.3, respectively, with a CO₂ permeance of 24.2 GPU.²⁷ In addition, our group reported that BMIM⁺NO₃⁻/Cu nanoparticle membranes demonstrate a CO₂ permeance of 9.8 GPU with an ideal selectivity of 12.25 for CO₂/N₂ and CO₂/CH₄.²⁵ These results indicated that Cu NPs could be an excellent carrier for CO₂ separation in ILs.

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Very recently, our group reported that metal oxide nanoparticles could be utilized for a CO₂ separation membrane. For example, an MOIM⁺BF₄⁻/AgO composite membrane showed a high ideal selectivity of 23.4 for CO₂/N₂, with a CO₂ permeance of 18.7 GPU; on the other hand, the neat MOIM⁺BF₄⁻ membrane showed a selectivity of 15.9, with a CO₂ permeance of 12.7 GPU.²⁸ Furthermore, when AgO particles were added to BMIM⁺BF₄⁻, a CO₂/N₂ selectivity of 28.2 and CO₂ permeance of 14.1 GPU were reported.²⁹ These results indicated that the dissociated AgO nanoparticles have strong affinity for CO₂ due to the oxide layer.²⁹

In this study, in order to clarify the effect of the metal oxide in ILs, experiments were conducted with other oxides. Copper(II) oxide was selected because its properties are similar to those of silver oxide. For the BMIM⁺BF₄⁻/CuO composite membrane, the dissociated CuO particles were expected to enhance the solubility of the CO₂ molecules due to the nano-oxide particle layer. Moreover, when CuO was incorporated into BMIM⁺BF₄⁻, the free ions could also enhance the solubility of CO₂ molecules since the quadrupole moment of CO₂ was relatively higher than N₂, resulting in the increased separation performance.

II. Experimental

2.1 Materials

BMIM⁺BF₄⁻ was purchased from Merck KGaA (Darmstadt, Germany). CuO was purchased from Sigma-Aldrich Chemical Co. Ethyl alcohol (>94.0% purity) was purchased from Daejung Chemicals & Metals. All chemicals were used as received.

2.2 Characterization

A sonifier (Branson 450, Branson Ultrasonics Corporation, Danbury CT, USA) with a standard tip was used. Scanning electron microscopy (SEM) images were obtained using a JEOL JSM-5600LV system. Transmission electron microscopy (TEM) images were obtained using a JEM-2010 system operating at 200 kV. The weight loss was measured by thermogravimetric analysis (TGA, TGA Q50, TA Instrument) of the complex membrane in flowing N₂. FT-Raman measurements were performed for the BMIM⁺BF₄⁻/CuO solution at room temperature using a Horiba Jobin Yvon LabRam Aramis apparatus with a diode laser beam at an excitation wavelength of 785 nm.

2.3 Membrane preparation

The membranes were prepared utilizing BMIM⁺BF₄⁻, copper(II) oxide, and ethanol. First, copper(II) oxide was sonicated to disperse in ethanol for 5 min, and then, BMIM⁺BF₄⁻ was added into the mixture. The dispersion was heated at 85 °C for 24 h to evaporate ethanol. The final solution was coated onto a polysulfone support (Toray Chemical Korea Inc.) and then cast using an RK control coater (Model 202, Control Coater RK Print-Coat Instruments Ltd., UK). The best performance of BMIM⁺BF₄⁻/CuO was observed at a weight ratio of 1/0.007.

2.4 Gas separation performance

Gas flow rates were measured with a bubble flow meter at an upstream pressure of 2 kgf·cm⁻² and atmospheric downstream

pressure. The unit of gas permeance is GPU, with 1 GPU = 1 × 10⁻⁶ cm³ (STP)/(cm² s cmHg). Selectivity was defined as ideal selectivity (CO₂ permeance/N₂ permeance).

III. Results and discussion

3.1 Transmission electron microscope (TEM) image of BMIM⁺BF₄⁻/CuO

TEM images were recorded to investigate the size and dispersity of the dissociated copper oxide particles in BMIM⁺BF₄⁻. As shown in Fig. 1, no significant copper(II) oxide agglomerates were present, and CuO particles were mostly observed as small nanoparticles with an average diameter of 10 nm. These results indicated that the dissociated CuO nanoparticles are advantageous for increasing the solubility of CO₂ molecules.

3.2 Separation performance

Fig. 2 presents the cross-section structures of the BMIM⁺BF₄⁻/CuO composite membrane. The utilized polysulfone support had a finger-like structure. The thickness of the coated selective layer was approximately 4.1 μm. Based on the SEM image, it was suggested that the selective layer was not penetrated into the polysulfone microporous support.

To investigate the effect of CuO concentration in the BMIM⁺BF₄⁻/CuO membrane on the CO₂ separation performance, the pure gas permeance of CO₂ and N₂ through the BMIM⁺BF₄⁻/CuO composite membrane was measured at increasing weight ratios of CuO, as shown in Fig. 3. As the CuO weight ratio increased, the CO₂ selectivity also sharply increased. As shown in Table 1, the highest CO₂ permeance (52.4 GPU) and CO₂/N₂ selectivity (21.0) were observed at a BMIM⁺BF₄⁻/CuO weight ratio of 1/0.007. At CuO weight ratios higher than 0.007, both CO₂ permeance and selectivity remained unchanged. However, above a CuO weight ratio of 0.04, the copper oxide was not well dispersed in BMIM⁺BF₄⁻, and therefore, significant copper oxide agglomerates were present. A previous study revealed that for neat BMIM⁺BF₄⁻, the CO₂ permeance and CO₂/N₂ selectivity were 17.0 GPU and 5.0, respectively, as shown Table 1.²⁴ On the other hand, when CuO was added to BMIM⁺BF₄⁻, the CO₂ permeance increased from 17 to 52.4 GPU and the CO₂/N₂ selectivity increased from 5.0 to

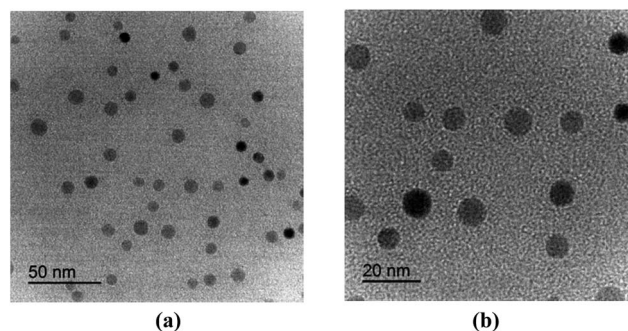


Fig. 1 (a) TEM image of the dissociated CuO nanoparticles in BMIM⁺BF₄⁻ and (b) enlarged image.



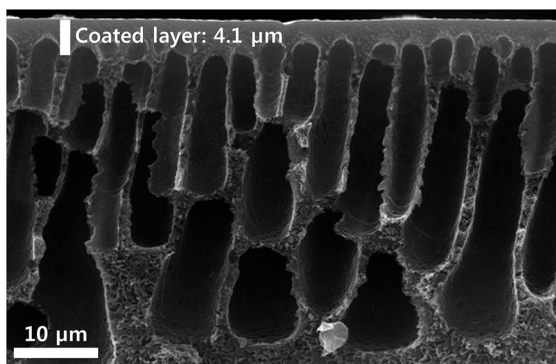


Fig. 2 SEM image of BMIM⁺BF₄⁻/CuO composite membrane.

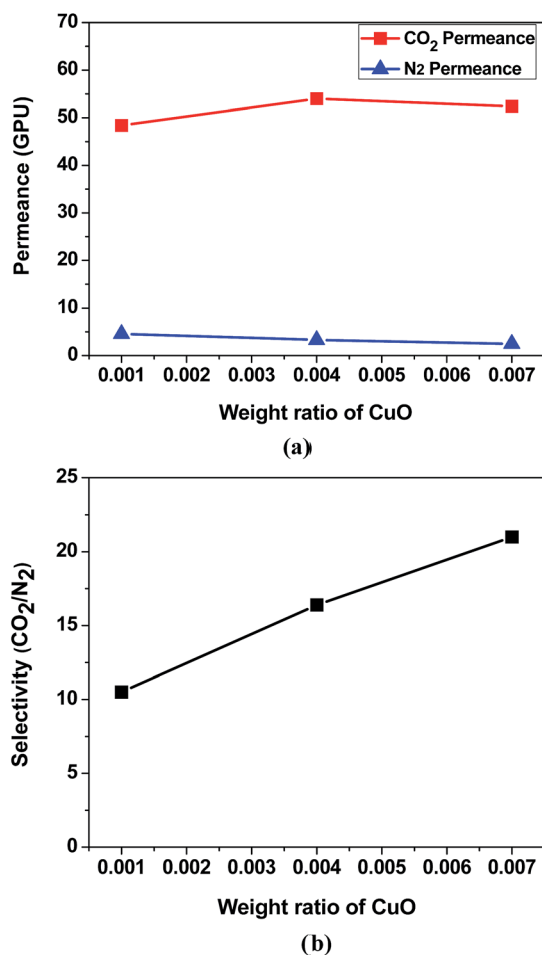


Fig. 3 Gas separation performance of BMIM⁺BF₄⁻/CuO composite membranes: (a) pure gas permeance and (b) ideal CO₂/N₂ selectivity.

21.0. Furthermore, it was confirmed that CO₂ permeance was much higher than the value reported in previous study consisting of BMIMBF₄/AgO composite. (CO₂ permeance: 14.1 GPU)²⁹ These results were ascribed to two factors: (1) well-dispersed CuO nanoparticles, which could lead to enhanced CO₂ solubility due to the increased oxide layer and (2) abundant free ions in BMIM⁺BF₄⁻, enabling the enhanced CO₂ transport.

Table 1 Pure gas permeance and selectivity of neat BMIM⁺BF₄⁻ and BMIM⁺BF₄⁻/CuO membranes with a weight ratio of 1/0.007

	CO ₂ permeance (GPU)	N ₂ permeance (GPU)	CO ₂ /N ₂ selectivity
Neat BMIM ⁺ BF ₄ ⁻ (ref. 24)	17.0	3.4	5.0
BMIM ⁺ BF ₄ ⁻ /AgO ²⁹	14.1	0.5	28.2
BMIM ⁺ BF ₄ ⁻ /CuO	52.4	2.5	21.0

3.3 Thermogravimetric analysis

The thermal properties of the neat BMIM⁺BF₄⁻ and BMIM⁺BF₄⁻/CuO composite membranes were analyzed by TGA. Fig. 4 presents weight loss at temperatures from 350 to 500 °C. The major weight loss was ascribed to the structural decomposition of the IL BMIM⁺BF₄⁻. The BMIM⁺BF₄⁻/CuO composite showed slower decomposition with a smaller weight loss of 70.6% at 450 °C, as compared to that (53.2%) of neat BMIM⁺BF₄⁻ at the same temperature. If the agglomeration of CuO particles in BMIM⁺BF₄⁻ occurred, the decomposition temperature would remain constant. However, this result indicates that CuO particles were well dispersed in the IL BMIM⁺BF₄⁻, because of the coordinative interaction between the CuO surface and BMIM⁺BF₄⁻. Thus, the thermal stability of the BMIM⁺BF₄⁻/CuO composite membrane increased.

3.4 FT-Raman spectroscopy

Interactions between BMIM⁺BF₄⁻ and CuO in the BMIM⁺BF₄⁻/CuO composite membranes were investigated by FT-Raman spectroscopy. The Raman spectrum in the region of the BF₄⁻ stretching bands is shown in Fig. 5. Three different ionic species were observed in the FT-Raman spectra of the ILs: free ions, ion pairs, and ionic aggregates. The BF₄⁻ stretching bands at 765, 770, and 774 cm⁻¹ were assigned to the free ions, ion pairs, and ion aggregates, respectively.³⁰ BMIM⁺ forms ion aggregates and ion pairs with BF₄⁻ in BMIM⁺BF₄⁻ through hydrogen bonding.²⁴ A previous report showed that the Raman spectrum of neat BMIM⁺BF₄⁻ included signals attributed to ion

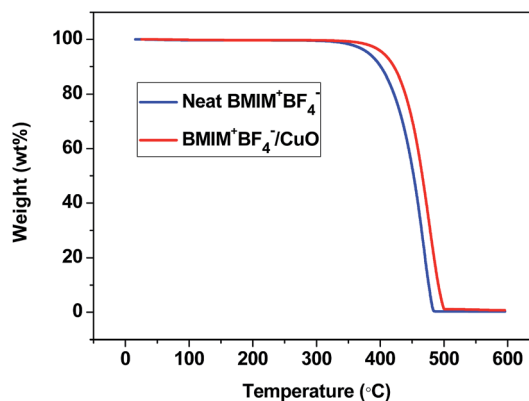


Fig. 4 TGA results for neat BMIM⁺BF₄⁻ (blue) and BMIM⁺BF₄⁻/CuO composite (red).



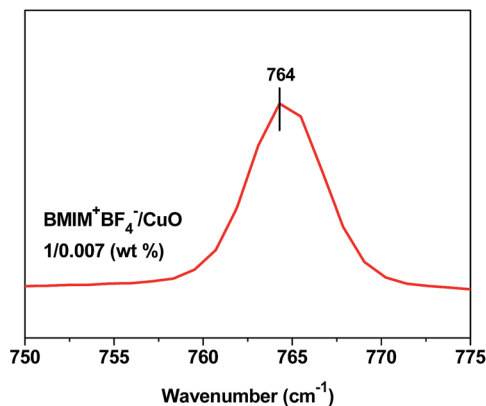


Fig. 5 FT-Raman spectrum of $\text{BMIM}^+\text{BF}_4^-/\text{CuO}$ nanoparticles.

pairs and ion aggregates.²⁴ As the copper oxide is well dispersed in $\text{BMIM}^+\text{BF}_4^-$, free BMIM^+ and BF_4^- ions are observed. These results can be explained by the coordinative interaction between $\text{BMIM}^+\text{BF}_4^-$ and the surface of the CuO nanoparticles, which results in weakened interactions between BMIM^+ and BF_4^- . Therefore, CO_2 separation performance could be enhanced by the increase of solubility due to abundant free ions.

IV. Conclusions

We have successfully stabilized CuO nanoparticles by employing an ionic liquid $\text{BMIM}^+\text{BF}_4^-$. The coordinative interactions of the $\text{BMIM}^+\text{BF}_4^-/\text{CuO}$ composite membrane were characterized by TEM, TGA, and FT-Raman analyses. The separation performance of the $\text{BMIM}^+\text{BF}_4^-/\text{CuO}$ composite membrane was considerably enhanced compared to that of the neat $\text{BMIM}^+\text{BF}_4^-$. When CuO was incorporated into $\text{BMIM}^+\text{BF}_4^-$, the CO_2/N_2 selectivity increased to 21.0, with a CO_2 permeance of 52.4 GPU. These results were attributed to (1) the synergistic effect of well-dispersed CuO nanoparticles, which result in enhanced CO_2 solubility due to the increased surface of the oxide particle layer and (2) the abundant free ions in $\text{BMIM}^+\text{BF}_4^-$.

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Notes and references

1 N. V. Blinova and F. Svec, *J. Mater. Chem. A*, 2014, **2**, 600–604.

- N. Du, H. B. Park, M. M. Dal-Cin and M. D. Guiver, *Energy Environ. Sci.*, 2012, **5**, 7306–7322.
- D. Aaron and C. Tsouris, *Sep. Sci. Technol.*, 2005, **40**, 321–348.
- H. Q. Yang, Z. H. Xu, M. H. Fan, R. Gupta, R. B. Slimane, A. E. Bland and I. Wright, *J. Environ. Sci.*, 2008, **20**, 14–27.
- J. Huang, J. Zou and W. S. Winston Ho, *Ind. Eng. Chem. Res.*, 2008, **47**, 1261–1267.
- C. H. Yu, C. H. Huang and C. S. Tan, *Aerosol Air Qual. Res.*, 2012, **12**, 745–769.
- O. Davidson and B. Metz, *International Panel on Climate Change*, 2005.
- P. Jungho, K. Deukju and N. Sangyong, *Membr. J.*, 2015, **25**, 547–557.
- P. arume and K. Sungjoong, *Membr. J.*, 2016, **26**, 126–134.
- R. W. Baker and K. Lokhandwala, *Ind. Eng. Chem. Res.*, 2008, **47**, 2109–2121.
- W. Yave, A. Car, J. Wind and K. V. Peinemann, *Nanotechnology*, 2010, **21**, 395301.
- A. Meisen and X. Shuai, *Energy Convers. Manage.*, 1997, **38**, S37–S42.
- W. J. Koros and R. Mahajan, *J. Membr. Sci.*, 2000, **175**, 181–196.
- L. Deng, T. J. Kim and M. B. Hagg, *J. Membr. Sci.*, 2009, **340**, 154–163.
- W. J. Ward and W. L. Robb, *Science*, 1957, **156**, 1481–1484.
- R. Yegani, H. Hirozawa, M. Teramoto, H. Himei, O. Okada, T. Takigawa, N. Ohmura, N. Matsumiya and H. Matsuyama, *J. Membr. Sci.*, 2007, **291**, 157–164.
- S. Kasahara, E. Kamio, T. Ishigami and H. Matsuyama, *Chem. Commun.*, 2012, **48**, 6903–6905.
- S. U. Hong, J. H. Jin, J. Won and Y. S. Kang, *Adv. Mater.*, 2000, **12**, 968–971.
- Y. Suzuki, H. Nishide and E. Tsuchida, *Macromol. Res.*, 2000, **33**, 2530–2534.
- C. Yu, T. Zhu, X. Zhang and H. Liu, *Sci. Adv. Mater.*, 2015, **7**, 1213–1220.
- J. Chang, G. H. Hong and S. W. Kang, *RSC Adv.*, 2015, **5**, 69698–69701.
- P. T. Nguyen, B. A. Voss, E. F. Wiesenauer, D. L. Gin and R. D. Noble, *Ind. Eng. Chem. Res.*, 2013, **52**, 8812–8821.
- K. E. Gutowski and E. J. Maginn, *J. Am. Chem. Soc.*, 2008, **130**, 14690–14704.
- J. H. Lee, J. Hong, J. H. Kim, Y. S. Kang and S. W. Kang, *Chem. Commun.*, 2012, **48**, 5298–5300.
- G. H. Hong, J. H. Oh, D. Ji and S. W. Kang, *Chem. Eng. J.*, 2014, **252**, 263–266.
- Y. Choi, G. H. Hong and S. W. Kang, *J. Nanosci. Nanotechnol.*, 2016, **16**, 2832–2835.
- J. H. Lee, J. Hong, J. H. Kim, D. Song, Y. S. Kang and S. W. Kang, *Chem. Eng. J.*, 2014, **235**, 252–256.
- D. Ji and S. W. Kang, *Korean J. Chem. Eng.*, 2016, **33**(2), 666–668.
- D. Ji, Y. S. Kang and S. W. Kang, *Sci. Rep.*, 2015, **5**, 16362.
- S. W. Kang, D. H. Lee, J. H. Park, K. Char, J. H. Kim, J. Won and Y. S. Kang, *J. Membr. Sci.*, 2008, **322**, 281–285.

