Ir/C and Brønsted acid functionalized ionic liquids: an efficient catalytic system for hydrogenation of nitrobenzene to p-aminophenol

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Introduction

p-Aminophenol (PAP) is an important commercial chemical raw material, which is widely used in the field of pharmaceuticals, dyestuffs, polyurethanes and photographic chemicals. PAP is usually synthesized by the reduction of p-nitrochlorobenzene or p-nitrophenol, which has several drawbacks including multistep processes, scarcity of raw materials and environmental problems. Compared to the above synthetic methods, the one-pot synthesis of PAP from nitrobenzene (NB) is regarded as the most convenient, economical and challenging method at present. This process involves two steps: the first step is the hydrogenation of NB to an intermediate phenylhydroxylamine to aniline. On the other hand, the Brønsted acid functionalized ionic liquids with sulfonic acid and bisulfate anions were acidic enough to catalyze the Bamberger rearrangement to form p-aminophenol from phenylhydroxylamine. On this basis, a new catalytic system constructed by Ir/C and Brønsted acid functionalized ionic liquid was applied, for the first time, to the one-pot hydrogenation of nitrobenzene to p-aminophenol. Our results indicate that the PAP selectivity of Ir/C and Brønsted acid functionalized ionic liquid was far more than that of the traditional Pt/C and sulfuric acid catalyst system. Furthermore, the dually functionalized ionic liquid ([HSO3-b-N-Bu][SO3H]) can be used simultaneously as an acid catalyst and also as a surfactant, due to its higher lipophilicity. Therefore, our new catalytic system has unique advantages in the hydrogenation of nitrobenzene to p-aminophenol.

Scheme 1 Hydrogenation of nitrobenzene to p-aminophenol.
BFILs also have unique properties like water-stability and structure-adjustability, which make them widely used in the field of esterification, hydrolysis, alkylation, the Biginelli reaction, hydrodeoxygenation, and Beckmann rearrangement instead of traditional mineral acids. Recently, Cui et al. reported their work on the one-pot hydrogenation of NB to PAP using BFIL [HSO₃₋-B-N-Me₃][HSO₄] as the acid catalyst and Pt/SiO₂ as the hydrogenation catalyst.55

Based on our previous research results and the performance of BFILs in the hydrogenation of NB to PAP, we constructed a new catalytic system by the combination of Ir/C and BFILs to catalyze the one-pot hydrogenation of NB to PAP. The catalytic performance of this system was then studied and compared with that of the traditional Pt/C and sulfuric acid system. In addition, considering the structure-adjustability of BFILs, we also wanted to modify the structure of the BFILs to improve the lipophilicity, and found real dual functionalized ionic liquids with the functions of acid catalyst and surfactant. The dual functionalization of acidic ionic liquids might afford an opportunity for us to carry out the hydrogenation of NB to PAP without adding surfactant.

**Experimental**

**Materials**

N-Methylimidazole, triethylamine, tripropylamine, tributylamine, pyridine, 1,4-butasulphonate and hexadecyl trimethyl ammonium bromide (CTAB) were purchased from Aladdin Reagent Co. Ltd. Sulfuric acid, nitrobenzene, aniline ethyl acetate, diethyl ether, acetonitrile, ethanol and isopropanol were purchased from Shanghai Lingfeng Chemical Reagent Co. Ltd. All reagents and raw materials were commercially available and used without any further purification.

**Preparation of catalysts**

**Hydrogenation catalysts.** The commercial coconut shell activated carbon with Brunauer–Emmett–Teller (BET) surface area of 1662 m² g⁻¹ was provided by Fujian Xinsen Carbon Co. Ltd. (China). The activated carbon was outgassed in a vacuum at 383 K overnight. A desired volume of H₂IrCl₆ or H₂PtCl₆ (Sino-Platinum Metals Co., Ltd.) aqueous solution (0.035 g L⁻¹), with nominal Ir or Pt loading of 1 wt%, was added to an aqueous suspension of the activated carbon. The aqueous suspension was subsequently dried at 353 K for 5 h. The formed catalyst was then reduced by hydrogen with 3 MPa pressure in distilled water at 363 K for 9 h. After the reduction was complete, the product was washed with ethyl acetate (20 mL × 3) and diethyl ether (20 mL × 3), respectively, then dried for 6 hours in a vacuum at 353 K. The BFIL SF-A was obtained. SF-B ([Et₃-NH][HSO₄]) and SF-C ([HPY][HSO₄]) were prepared according to the procedure described above for SF-A with the reaction of triethylamine or pyridine with sulfuric acid (Scheme 2).

The preparation of type B BFILs DF-A ([SO₃H-bmim][HSO₄]) was similar to that used in the literature⁴⁶ and the detailed procedure is as follows: N-methylimidazole (16.42 g, 0.2 mol) and equimolar 1,4-butasulphonate (27.23 g, 0.2 mol) were mixed in a 100 mL single-necked flask. Then, the mixture was stirred at 333 K under nitrogen until the white solid was completely generated. The solid was washed with ethyl acetate (20 mL × 3) and dried in vacuum at 353 K for 6 hours. A stoichiometric amount of sulphuric acid was added dropwise to the zwitterion and the mixture was stirred at 353 K for 6 hours, resulting in the formation of DF-A. The product was washed with ethyl acetate (20 mL × 3) and dried under vacuum at 383 K for 6 hours. DF-B1 ([HSO₃₋-b-N-Et₃][HSO₄]) and DF-C ([HSO₃₋-bPy][HSO₄]) were prepared according to the procedure described above for DF-A, with triethylamine and pyridine, respectively, instead of N-methylimidazole.

Type C BFILs DF-B2 ([HSO₃₋-b-N-Pr₃][HSO₄]) and DF-B3 ([HSO₃₋-b-N-Bu₃][HSO₄]) were prepared using the same procedure described above, replacing N-methylimidazole with tripropylamine and tributylamine.⁶⁸⁹

**Characterization of catalysts**

**TEM and XPS.** X-ray Photoelectron Spectroscopy (XPS) was conducted on a Thermo ESCALAB 250 Axis Ultra using monochromatic Al Kα radiation (hv = 1486.6 eV). The binding energies were regulated according to the C 1s peak at 284.8 eV from the samples.

The particle size of Ir or Pt on the active carbon surface was determined by transmission electron microscopy (TEM) using a Tecnai G2 F30 S-Twin microscope (Philips-FEI Co.). At least 200 individual Ir particles were counted for each catalyst. The Ir or Pt particle size of the catalysts, dₛ, was calculated using the following equation: 

\[
dₛ = \frac{\sum n_i d_i^3}{\sum n_i d_i^2},
\]

where the visible particle size dᵢ on the micrographs was measured by a computerized system.

![Scheme 2 Structures of the BFILs.](Image)
**1H-NMR and IR spectroscopy of ionic liquids.** The structures of BFILs were characterized by 1H-NMR and IR spectroscopy. Some of the BFILs were further characterized by 13C-NMR. 1H-NMR and 13C-NMR spectra were recorded on a Bruker AVANCE III spectrometer in D2O. FT-IR measurements were performed using KBr disc on a Nicolet iS50 FT-IR infrared spectrometer. The spectral data were obtained as follows and were all in agreement with that reported in the corresponding ref. 27 and 40–42.

Spectral data for SF-A ([Hmim][HSO4]): 1H-NMR (500 MHz, D2O) δ 9.06 (s, 1H), 7.70 (t, 1H), 7.65 (t, 1H), 3.88 (s, 3H). 13C-NMR (125MHz, D2O) δ 35.86, 120.41, 122.95, 135.86. IR (cm\(^{-1}\)): 3149, 2967, 1639, 1586, 1546, 1408, 761.

Spectral data for SF-B ([Et3N][HSO4]): 1H-NMR (500 MHz, D2O) δ 3.09 (q, 6H), 1.18 (t, 9H). IR (cm\(^{-1}\)): 2991, 2951, 1476, 1399, 1165, 1036.

Spectral data for SF-C ([Hpy][HSO4]): 1H-NMR (500 MHz, D2O) δ 8.66 (d, 2H), 8.50 (m, 1H), 8.08–7.84 (m, 2H). IR (cm\(^{-1}\)): 3071, 1639, 1545, 1488.

Spectral data for DF-A ([SOH2-b-nim][HSO4]): 1H-NMR (500 MHz, D2O) δ 8.59 (s, 1H), 7.32 (m, 2H), 4.10 (t, 2H), 3.80–3.72 (m, 3H), 2.83–2.62 (m, 2H), 1.88 (m, 2H), 1.76–1.41 (m, 2H). 13C-NMR (125 MHz, D2O) δ 21.50, 28.43, 35.87, 49.28, 51.07, 72.82, 123.64, 123.55, 136.56. IR (cm\(^{-1}\)): 3156, 3115, 2960, 1575, 1169, 1031, 794, 747, 579.

Spectral data for DF-B1 ([HSO3-b-N-Et3][HSO4]): 1H-NMR (500 MHz, D2O) δ 3.08 (q, 6H), 2.99 (dd, 2H), 2.76 (dd, 2H), 1.76–1.52 (m, 4H), 1.06 (t, 9H). 13C-NMR (125 MHz, D2O) δ 7.36, 19.98, 21.90, 50.26, 52.28, 56.10. IR (cm\(^{-1}\)): 2991, 1488, 1396, 1189, 730, 587.

Spectral data for DF-B2 ([HSO3-b-N-Pr3][HSO4]): 1H-NMR (500 MHz, D2O) δ 3.19–3.08 (m, 2H), 3.07–2.98 (m, 6H), 2.88–2.74 (m, 2H), 1.77–1.67 (m, 2H), 1.64 (dd, 2H), 1.61–1.46 (m, 6H), 0.80 (t, 9H). IR (cm\(^{-1}\)): 2976, 2884, 1487, 1473, 1169, 1054, 720, 579.

Spectral data for DF-B3 ([HSO3-b-N-Bu3][HSO4]): 1H NMR (500 MHz, D2O) δ 3.01–2.89 (m, 6H), 2.89–2.80 (m, 2H), 2.71–2.60 (m, 2H), 1.60–1.45 (m, 4H), 1.43–1.32 (m, 6H), 1.15–1.02 (m, 6H), 0.66 (q, 9H). IR (cm\(^{-1}\)): 2963, 2876, 1472, 1168, 1054, 726, 579.

**Acidity of ionic liquids.**

In order to explore the relationship between the catalytic performance and the acidity of BFILs, the acidity of BFILs was measured by an acid–base titration method. Phenolphthalein was used as an indicator, and the concentration of NaOH aqueous solution was 2 mol L\(^{-1}\).

\[
\text{Acidity} = \frac{n(\text{NaOH})}{n(\text{BFILs})}
\]

**Solubility of NB in the reaction solution**

The solubility of NB in the reaction solution was checked by HPLC, which was similar to that used in the literature. The method is as follows: the slight excess of 6 NB was sonicated in an ultrasonic washer for 30 minutes and then stirred at 353 K for 5 hours. After that, 1 mL of the upper homogeneous was transferred to the ethanol solution and diluted to a fixed volume, then determined by HPLC.

**Evaluation of catalyst**

This reaction was carried out in a 75 mL stainless steel autoclave (Parr MRS5000). In a typical experiment, NB, catalyst, BFILs or 10% sulphuric acid solution and CTAB were introduced into the autoclave. After the air in the reactor was replaced with nitrogen and hydrogen, three times, respectively, the reactions were performed at 0.8 MPa H\(_2\), 353 K, the optimal reaction conditions that were previously determined. The final products were analyzed by HPLC on an Agilent 1260 (column Agilent TC-C18, 5 \(\mu\)m) system with UV detector (\(\lambda = 254\) nm). The column temperature was 298 K. CH\(_3\)CN/H\(_2\)O (70/30, v/v) was used as the mobile phase and the flow rate was 1 mL min\(^{-1}\).

**Recycling of Ir/C and ionic liquids**

The Ir/C catalyst in the system was separated by vacuum filtration, and the aniline by-product was separated by ordinary extraction. After most of the water was removed by distillation under reduced pressure, PAP precipitated in the ice-water bath, due to its low solubility in aqueous solution. The remaining BFILs could then be reused.

**Results and discussion**

**Hydrogenation catalyst**

In our previous study of the hydrogenation of halogenated nitrobenzene, we had compared the catalytic performance of the Ir/C and the Pd/C catalysts. DFT calculations showed that adsorption properties of halogenated nitrobenzene on the Ir surface were different from that on the Pd surface. The distance between oxygen and Ir was much shorter than that between oxygen and Pd. Moreover, in our study of the hydrogenation of NB to AN, we also found that the amount of PHA intermediate detected in the solution of the Ir/C catalyst was much larger than that of the Pd/C catalyst. The adsorption properties of halogenated nitrobenzene on Ir aroused our interest. As we all know, for the hydrogenation of NB to PAP, NB was first hydrogenated to the PHA intermediate and then the PHA intermediate was desorbed from the surface of the hydrogenation catalyst and converted to PAP through the acid-catalysed Bamberger rearrangement. Therefore, the desorption properties of the PHA intermediate from the hydrogenation catalyst directly determined the amount of PHA involved in the next Bamberger rearrangement. Considering the above results, we conceived of Ir as the hydrogenation catalyst for the one-pot hydrogenation of NB to PAP. Thus, we first compared the catalytic performance of Ir with the traditional Pt catalyst.

The XPS spectra of Ir/C and Pt/C are shown in Fig. 1. The Ir 4f\(_{7/2}\) could be deconvoluted into two peaks at about 61.1 eV and...
62.2 eV, which were assigned to metallic Ir(0) and Ir(4+) respectively.\textsuperscript{45,46} The Pt 4f\textsubscript{7/2} could be deconvoluted into two peaks at about 71.5 eV and 72.3 eV, which were assigned to metallic Pt(0) and Pt(2+) respectively.\textsuperscript{47,48} The presence of Ir(4+) and Pt(2+) might be related to the oxidation of Ir and Pt particles in air.\textsuperscript{46}

High resolution transmission electron microscopy (HR-TEM) images (Fig. 2a, b, e and f) revealed that Ir(Pt) particles were found on the surface of the activated carbon support. Moreover, the hydrogenation of NB to PAP did not cause any appreciable changes in the catalysts. At the same time, transmission electron microscopy (TEM) images and particle size distribution of the Ir(Pt) particles (Fig. 2c and d) demonstrated that the surface of metal particles were in a small and uniform state and the average particle size of the Ir/C catalyst was very close to that of the Pt/C catalyst. All these results indicated that the Ir and Pt should have similar amounts of active sites under the same metal load.

The hydrogen consumption rates for the hydrogenation of NB to AN were then studied and the results are shown in Fig. 3, where we can easily find that the hydrogen consumption rate of the Pt/C catalyst was much higher than that of the Ir/C catalyst. Obviously, this significant difference in the hydrogenation rate was not caused by the number of metal active sites. In order to find out the reason, the product distribution of the Pt/C catalyst for hydrogenation of NB to AN was then investigated and the results were demonstrated in Fig. 4. Compared to our previous study results,\textsuperscript{13} we found that the amount of the PHA intermediate distribution of Pt/C was much lower than that of Ir/C (about 21%). This indicated that the PHA intermediate desorbed more easily from the Ir surface than from the Pt surface. The more easily PHA was desorbed from the metal surface, the less PHA was left on the metal surface to be further hydrogenated to AN, which should be the main reason why the hydrogen consumption rate of Ir/C was much lower than that of Pt/C.

As an intermediate of the hydrogenation of NB to PAP, the distinct adsorption properties of PHA on different metal catalysts could not only affect the hydrogen consumption rate, but also influence the PAP selectivity. Fig. 3 showed the hydrogen consumption rate of the hydrogenation of NB to PAP. As we had speculated, the hydrogen consumption rate of Ir and Pt catalysts all decreased sharply, compared to the hydrogenation of NB to AN, and the hydrogen consumption rate of the Ir/C catalyst was still much lower than that of the Pt/C. On the other hand, we further studied the catalytic performance of Ir/C and Pt/C for the hydrogenation of NB to PAP (Table 1). Our results indicated that the PAP selectivity of the Ir/C catalyst was significantly higher than that of the Pt/C. We thought the reason was that the easier desorption of PHA from Ir allowed more PHA to be involved in the following Bamberger rearrangement, which led to the lower hydrogen consumption rate and higher selectivity of PAP. Thus, we concluded that Ir was more suitable than Pt for the hydrogenation of NB to PAP.

**Bromsted acid functionalized ionic liquids**

When the hydrogenation catalyst was determined, we selected BFILs as acid catalysts and constructed a new catalytic system...
with Ir/C for the hydrogenation of NB to PAP. In this regard, three types of BFILs were synthesized. The type A BFILs (SF-A, SF-B and SF-C) were single functionalized ionic liquids, which were characterized by the bisulfate ion as the anion and N-methylimidazole, triethylamine or pyridine as the cations, respectively. Based on the preparation of type A BFILs, type B BFILs (DF-A, DF-B1 and DF-C) were designed with butyl sulfonic acid functionalized to the N atom. The purpose of this reform was to increase the acid strength of the corresponding BFILs. In order to enhance the lipophilicity of DF-B1, type C BFILs (DF-B2 and DF-B3) were then prepared by replacing the ethyl group in DF-B1 with \( n \)-propyl (DF-B2) or \( n \)-butyl (DF-B3) respectively. Type C BFILs were more targeted to functionalization, compared to DF-B1. They were a type of dually functionalized ionic liquid and were expected to be used as acid catalyst and surfactant at the same time.

The hydrogen consumption rates in the BFILs were firstly investigated in the presence of surfactant. As shown in Fig. 5, the hydrogen consumption rates in type A and type B BFILs were different. We could easily see that the hydrogen consumption rate in type A BFILs was much higher than that in type B BFILs. The hydrogen consumption rates in DF-A, DF-B1 and DF-C were basically within the same range. The above experimental results could be explained through the mechanism of the hydrogenation of NB to PAP. This reaction involved two competitive parallel reactions, and the acidity of BFILs affects the trend of the Bamberger rearrangement reaction. When the acidity of BFILs was stronger, the Bamberger rearrangement was dominant in the parallel reaction and the over hydrogenation side reaction was inhibited, which led to the lower hydrogen consumption rate. The acidity of type B BFILs stemmed not only from the bisulfate anion, but also the sulfonic acid, which made them more acidic than type A BFILs and resulted in the lower hydrogen consumption rate.

![Fig. 3](image-url) Hydrogen consumption rate of the Ir/C and Pt/C catalysts for the hydrogenation of NB to AN \(^a\) and hydrogenation of NB to PAP \(^b\). Reaction conditions: (a) 0.02 mol of NB, 0.1 g of catalyst, 25 mL of ethanol, \( P_{\text{H}_2} = 1.0 \) MPa, \( T = 353 \) K, stirring rate = 1200 rpm. (b) 0.02 mol of NB, 0.1 g of catalyst, 25 mL of 10% sulphuric acid solution, 0.025 g of CTAB, \( P_{\text{H}_2} = 0.8 \) MPa, \( T = 353 \) K, stirring rate = 1000 rpm.

![Fig. 4](image-url) Product distribution of catalytic hydrogenation of NB to AN on the Pt/C catalyst. Reaction conditions: 0.05 g of catalyst, 25 mL of ethanol, 0.02 mol of NB, \( P_{\text{H}_2} = 1.0 \) MPa; \( T = 353 \) K, stirring rate = 1200 rpm.

![Fig. 5](image-url) The hydrogen consumption rate of the reaction in different BFILs in the presence of CTAB. Reaction conditions: 0.02 mol of NB, 25 mL of H\(_2\)O, 0.1 g of Ir/C, 0.025 g of CTAB, 353 K, 0.8 MPa, 1000 rpm, 0.05 mol BFILs of type A, 0.025 mol BFILs of type B.

<table>
<thead>
<tr>
<th>Catalyst</th>
<th>Conversion (%)</th>
<th>Selectivity (%)</th>
</tr>
</thead>
<tbody>
<tr>
<td></td>
<td></td>
<td>PAP</td>
</tr>
<tr>
<td>1% Pt/C</td>
<td>100</td>
<td>40.5</td>
</tr>
<tr>
<td>1% Ir/C</td>
<td>100</td>
<td>60.4</td>
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</table>

\(^a\) Reaction conditions: 0.1 g of catalyst, 25 mL of 10% sulphuric acid solution, 0.02 mol of NB, 0.025 g of CTAB, \( P_{\text{H}_2} = 0.8 \) MPa, \( T = 353 \) K, stirring rate = 1000 rpm.
Considering the influence of the acidity on the Bamberger rearrangement reaction, we thought that it was necessary to further compare the acidity of the different BFILs quantitatively. Therefore, the acidity of the BFILs was quantitatively tested by sodium hydroxide titration. The relationship between the amount of acid and the catalytic performance of different kinds of BFILs was studied. The results are presented in Table 2, where it can be seen that the selectivity of PAP was proportional to the acidity of BFILs. The strong acidity of BFILs benefited the Bamberger rearrangement to form PAP and thus increased the selectivity of PAP; the relative acidity of the SO_3H-functionalized ILs was stronger than that of the anion-functionalized ILs and thus, the type B BFILs should possess stronger acidity than that of the type A BFILs. Due to the acidity of type A BFILs being much weaker than that of type B, there was an obvious gap in the PAP selectivity between type A and type B BFILs. We could also see that the PAP selectivity in type B BFILs were very close to that in sulfuric acid. Among all the BFILs of type B, DF-A showed the best catalytic performance and the PAP selectivity of DF-A was even higher than that of sulfuric acid. We thought that this reason could be attributed to its weak steric hindrance and strong acidity. Our results indicate that BFILs could already replace sulfuric acid for the hydrogenation of NB to PAP in the presence of surfactant.

Due to the lower solubility of NB in the reaction solution, surfactant must be added for the hydrogenation of NB to PAP. The addition of surfactant led to some problems such as a complicated system and troublesome operation. Considering the properties of the BFILs, we further explored the catalytic performance of different BFILs without the existence of surfactant. The result is shown in Table 3. As we predicted, the PAP selectivity greatly decreased both the type B BFILs and sulfuric acid, implying the non-negligible importance of surfactant for such a complex reaction system. However, the catalytic performance of DF-B1 caught our attention. Although DF-B1 was the weakest acid among all the type B BFILs, it showed the highest PAP selectivity without the existence of surfactant. We speculated that the catalytic performance of the BFILs was also related to lipophilicity in addition to acidity. The lipophilicity of the BFILs increased the solubility of the reactants, which would make the reaction occur more easily. This provided us with a good idea for modifying the DF-B1. In order to further enhance the lipophilicity of DF-B1, we increased the length of the carbon chain connected to the N atom by replacing the ethyl group in DF-B1 with the n-propyl or n-butyl carbon chain. As expected, the solubility of NB in type C BFILs improved significantly, and the selectivity of PAP increased with the elongation of the carbon chain. DF-B3 showed the best catalytic performance with 50.8% selectivity of PAP. Our results indicated that the DF-B3 dually functionalized ionic liquid could simultaneously be used as an acid catalyst and surfactant for the hydrogenation of NB to PAP.

Recyclability is one of the important properties of catalysts. Herein, we chose the catalytic system composed of Ir/C and DF-B3 to conduct the recycling experiments. As shown the Fig. 6, the conversion of NB remained stable and the selectivity of PAP decreased slightly after being recycled three times. This indicated that the new catalytic system of Ir/C and DF-B3 had good stability in the hydrogenation of NB to PAP.

### Conclusions

The desorption properties of the PHA intermediate from the hydrogenation metal catalyst and the acidity of the acid catalyst...
were the key factors in the one-pot hydrogenation of NB to PAP. Our work included the following two aspects: the PHA intermediate was more easily desorbed from the surface of the Ir catalyst, compared to the traditional Pt catalyst; the BFILs with bisulfate ion and sulfonic acid exhibited similar acidity to sulfuric acid. A new catalyst system was constructed through the combination of Ir/C and BFILs and applied to the catalysis of the one-pot hydrogenation of NB to PAP. Our results indicate that the PAP selectivity of Ir/C and [SO3H-bmim][HSO4] was far greater than that of Pt/C and sulfuric acid. In addition, we found that the dually functionalized ionic liquid ([HSO3-b-n-Bu3][HSO4]) could be used simultaneously as acid catalyst and surfactant, due to its higher lipophilicity. Therefore, the great catalytic performance of the Ir/C and BFILs system could open new ground for the study of the catalyzed hydrogenation of NB to PAP.

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Notes and references