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# Selective C–H bond hydroxylation of cyclohexanes in water by supramolecular control†

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A new approach for selective hydroxylation of non-activated cyclohexanes using dioxirane generated *in situ* in water through supramolecular control has been developed. Using  $\beta$ -CD and  $\gamma$ -CD as the supramolecular hosts, selective hydroxylation of cyclohexane substrates, including *trans/cis*-1,4-, 1,3- and 1,2-dimethylcyclohexanes and *trans/cis*-decahydronaphthalene, was achieved in up to 54% yield in water. Furthermore, site-selective C–H bond hydroxylation of (+)-menthol was achieved by obstructing the approach of dioxirane to the C–H bond with higher steric hindrance through inclusion complexation with  $\beta$ -CD and  $\gamma$ -CD in water.

## Introduction

Over the decades, significant advancements in the research area of C–H bond oxidation<sup>1–23</sup> have been accomplished using stoichiometric oxidizing reagents,<sup>7,9–11</sup> catalytic transition metal catalysts<sup>12–18</sup> and biomimetic supramolecular complexes.<sup>2,19–23</sup> Notably, selective oxidation of non-activated  $sp^3$  C–H bonds of alkanes is a very useful reaction because it offers a strategy with improved atom- and step-economy to transform inert C–H bonds into versatile functionalities for further synthetic elaborations.<sup>6</sup> However, it is regarded as one of the most challenging tasks in organic chemistry as the functionalization of non-activated C–H bonds generally requires harsh reaction conditions, and the highly reactive oxidizing agents employed exhibit low functional group tolerance.<sup>3,19,24</sup> Furthermore, the presence of multiple C–H bonds with similar reactivity in a single molecule renders site-selective oxidation difficult.<sup>24</sup> In the literature, synthetic methods have been developed for selective oxidation of non-activated  $sp^3$  C–H bonds on the basis of electronic, steric, and substrate-based directing effects.<sup>1–8</sup> However, these approaches are limited by the inherent structural properties of the substrates. For example, less activated aliphatic C–H bonds could be selectively targeted for oxidation using “directing” groups strategically installed in the substrates.<sup>17,25</sup> In view of the significance of C–H bond functionalization, there remains great interest in developing new strategies for selective  $sp^3$  C–H bond oxidation to target a diverse variety of substrates.

Oxidized cyclohexanes are commonly found in the structural skeletons of natural products such as artemisinin,<sup>19</sup> taxanes,<sup>26</sup> terpenes,<sup>27,28</sup> and steroids.<sup>29</sup> Thus, selective oxidation of cycloalkanes would be an appealing reaction for efficient synthesis of these structurally diverse oxidized cyclohexane-containing organic molecules.

Supramolecular host–guest chemistry provides promising approaches for achieving selectivity enhancement in a wide range of organic reactions.<sup>30–33</sup> Generally, supramolecular hosts are able to preferentially position the target sites of the substrates in proximity to the reaction centre, leading to selective organic transformations. On the other hand, access to the non-target sites of the substrates for reaction could be obstructed through inclusion complex formation.

Cyclodextrins (CDs) are amphipathic molecules bearing a hydrophilic exterior surface and a hydrophobic interior cavity (Fig. 1). Due to the hydrophobic cavity of cyclodextrins and their derivatives, “host–guest” inclusion complex formation with a wide diversity of guest molecules through hydrophobic interactions has been extensively employed in supramolecular catalysis.<sup>34–40</sup> Breslow reported a pioneering study on selective C–H bond oxidations of steroids using cyclodextrin-attached metalloporphyrins as catalysts.<sup>39,40</sup> In 2003, we developed a cyclodextrin-ketoester as a supramolecular catalyst for stereoselective alkene epoxidation.<sup>37</sup> Then, Bols and co-workers developed cyclodextrin ketones for organic oxidations.<sup>41–45</sup> In 2012, we studied the site-selective

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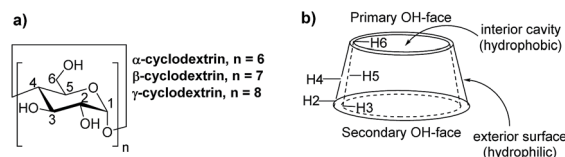


Fig. 1 General structure of cyclodextrins.



oxidation of tertiary C–H bonds of 3,7-dimethyloctyl esters using cyclodextrins as the supramolecular host.<sup>38</sup>

Because selective C–H bond hydroxylation of cyclohexanes using supramolecular control remains largely unexplored and cyclodextrin-promoted reactions are environmentally friendly using water as the solvent, we aim to realize site-selective oxidation of C–H bonds on cyclohexanes in water using cyclodextrins as supramolecular hosts, which would lead to interesting selectivity.<sup>46</sup> Based on our previous studies on cyclodextrins,<sup>37,38,47</sup> we developed a new approach for site-selective hydroxylation of cyclohexane derivatives through supramolecular control using cyclodextrins as the hosts. The inclusion models of *trans*-1,4-dimethylhexane (**1**) and (+)-menthol (**9**) with  $\beta$ -CD and  $\gamma$ -CD as well as the relationship between inclusion models and site-selective C–H bond hydroxylation of **9** have also been investigated. It is noteworthy that  $\beta$ -CD and  $\gamma$ -CD could act as reaction vessels such that the C–H bond hydroxylation could be performed in water without using organic solvents.

## Results and discussion

### The effect of different CDs on C–H bond hydroxylation of cyclohexanes

To study the efficiency of different CDs on C–H bond hydroxylation of cyclohexanes, *trans*-1,4-dimethylcyclohexane (**1**) was employed as the substrate (Table 1). Dioxirane generated *in situ* from 1,1,1-trifluoroacetone and Oxone was used as the oxidizing agent.<sup>48–59</sup> Initially, the reaction was carried out at room temperature by stirring *trans*-1,4-dimethylcyclohexane (**1**) (1.0 mmol) with  $\beta$ -CD (1.1 mmol) and 1,1,1-trifluoroacetone (1.0 mmol) in water (50 mL) for 1 h, and then 8 portions of a mixture of Oxone (2.5 mmol  $\times$  8) and NaHCO<sub>3</sub> (7.75 mmol  $\times$  8) were added within 7 h (one portion was added per hour). After stirring for an additional 16 h, the hydroxylation product **1a** was isolated in 12% yield by flash column chromatography (Table 1, entry 1). Using  $\gamma$ -CD instead of  $\beta$ -CD as the supramolecular host, **1a** was obtained in

32% yield (Table 1, entry 2). On using  $\alpha$ -CD, only a trace amount of **1a** was observed by TLC analysis (Table 1, entry 3). A control experiment was performed in water without cyclodextrin, and no hydroxylation product **1a** was obtained, probably due to the low solubility of **1** in water (Table 1, entry 4). **1a** was obtained in 16% yield when the reaction was conducted in 1 : 1 CH<sub>3</sub>CN/H<sub>2</sub>O without cyclodextrins as the supramolecular host (Table 1, entry 5). These results suggest that  $\beta$ -CD and  $\gamma$ -CD could enhance the hydroxylation yield by acting as a reaction vessel and maximizing the dispersion of **1** through inclusion complex formation.<sup>60</sup> Furthermore,  $\beta$ -CD and  $\gamma$ -CD could bind to **1** and dioxirane so as to place the two reactants in close proximity, stabilizing the transition state of the reaction and reducing the activation energy, and thus accelerating the reaction.<sup>61–65</sup>

Subsequently, *cis*-1,4-dimethylcyclohexane (**2**), *trans/cis*-1,3- and 1,2-dimethylcyclohexane (**3–6**), and *trans/cis*-decahydronaphthalene (**7** and **8**) were employed as substrates. As shown in Table 2, dimethylcyclohexanes (**1–6**) can be oxidized by dioxirane generated *in situ* through complexation with  $\beta$ -CD or  $\gamma$ -CD to give the corresponding hydroxylation products. Notably, *cis*-1,4-dimethylcyclohexane (**2**) and *trans/cis*-1,3-dimethylcyclohexane (**3** and **4**) gave monohydroxylation and dihydroxylation products (Table 2, entries 2–4), while *trans/cis*-1,2-dimethylcyclohexane (**5** and **6**) and *trans/cis*-decahydronaphthalene (**7** and **8**) gave only monohydroxylation products (Table 2, entries 5–8). In the presence of  $\beta$ -CD, *trans/cis*-decahydronaphthalene (**7** and **8**) did not give monohydroxylation or dihydroxylation products. Good yields (54% and 40%) were achieved when  $\gamma$ -CD was used as a supramolecular host. These findings indicate that  $\beta$ -CD and  $\gamma$ -CD could act as reaction vessels through inclusion complexation of hydrophobic cyclohexanes in water. Moreover, these results showed that using  $\gamma$ -CD as the supramolecular host gave higher yields than using  $\beta$ -CD as the supramolecular host.

### <sup>1</sup>H NMR titration for binding of **1** to $\beta$ -CD/ $\gamma$ -CD

To investigate the effect of the cavity sizes of  $\beta$ -CD and  $\gamma$ -CD on C–H bond oxidation of cycloalkanes, we studied the inclusion complex formation between *trans*-1,4-dimethylcyclohexane (**1**) and  $\beta$ -CD or  $\gamma$ -CD by <sup>1</sup>H NMR titration (Fig. 2 and 3).<sup>38,66–70</sup> The signals of H3 and H5 of  $\beta$ -CD or  $\gamma$ -CD were shifted significantly upfield when the amount of **1** increased in the aqueous solution of  $\beta$ -CD or  $\gamma$ -CD, while the chemical shifts of H2 and H4 remained unchanged. These results indicate that **1** interacts with the hydrogens in the hydrophobic cavities of  $\beta$ -CD and  $\gamma$ -CD.

<sup>1</sup>H NMR titration curves were obtained by plotting the change in the chemical shift of H3 against the ratio of **1** to  $\beta$ -CD (Fig. 4) and the change in the chemical shift of H3 against the ratio of **1** to  $\gamma$ -CD (Fig. 5). The stoichiometry for the inclusion complex formation between  $\beta$ -CD and **1** was 1 : 1, while that of the inclusion complex formation between  $\gamma$ -CD and **1** was 1 : 2, as determined by extrapolating the curve.

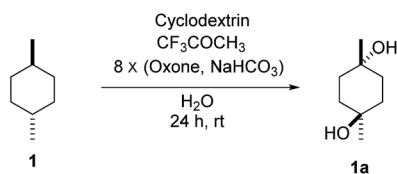
### The effect of CDs on selective C–H bond hydroxylation of mixtures of *trans/cis*-cycloalkane isomers

As both  $\beta$ -CD and  $\gamma$ -CD enhanced the selective C–H bond hydroxylation reactions of *trans/cis*-1,4-dimethylcyclohexanes

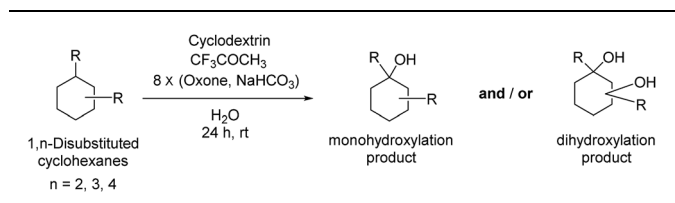
Table 1 Studies on the effect of different cyclodextrins towards C–H bond hydroxylation of *trans*-1,4-dimethylcyclohexane (**1**)<sup>a</sup>

Entry	Cyclodextrins	Yield of <b>1a</b> <sup>b</sup> (%)
1	$\beta$ -CD	12
2	$\gamma$ -CD	32
3	$\alpha$ -CD	Trace <sup>c</sup>
4	Without CD (in H <sub>2</sub> O)	0 <sup>c</sup>
5 <sup>d</sup>	Without CD (in CH <sub>3</sub> CN/H <sub>2</sub> O)	16

<sup>a</sup> Reactions were conducted with **1** (1.0 mmol), 1,1,1-trifluoroacetone (1.0 mmol) and cyclodextrin (1.1 mmol) in H<sub>2</sub>O (50 mL) at room temperature with Oxone (2.5 mmol  $\times$  8) and NaHCO<sub>3</sub> (7.75 mmol  $\times$  8). <sup>b</sup> Yield of isolated product. <sup>c</sup> Observed by TLC analysis. <sup>d</sup> CH<sub>3</sub>CN (25 mL) and H<sub>2</sub>O (25 mL).

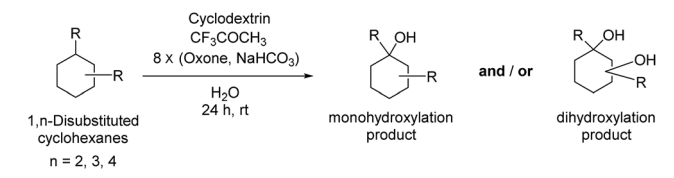


**Table 2** Comparison of C–H bond hydroxylation of various disubstituted cyclohexanes with and without CDs<sup>a</sup>



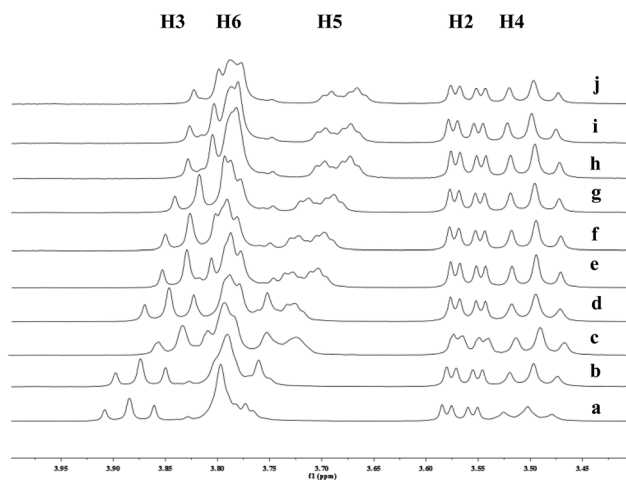
Entry	Substrate	Product	Yield <sup>b</sup> (%)	CDs
1			12	β-CD
			32	γ-CD
			16	Without CD
				(in CH <sub>3</sub> CN/H <sub>2</sub> O)
2			7	β-CD
			11	γ-CD
			15	Without CD
				(in CH <sub>3</sub> CN/H <sub>2</sub> O)
3 <sup>c</sup>			5	β-CD
			9	γ-CD
			7	Without CD
				(in CH <sub>3</sub> CN/H <sub>2</sub> O)
4			11	β-CD
			30	γ-CD
			40	Without CD
				(in CH <sub>3</sub> CN/H <sub>2</sub> O)
5 <sup>c</sup>			3	β-CD
			7	γ-CD
			4	Without CD
				(in CH <sub>3</sub> CN/H <sub>2</sub> O)
6			10	β-CD
			27	γ-CD
			16	Without CD
7			3	β-CD
			7	γ-CD
			5	Without CD
				(in CH <sub>3</sub> CN/H <sub>2</sub> O)

**Table 2** (Contd.)



Entry	Substrate	Product	Yield <sup>b</sup> (%)	CDs
8			— <sup>d</sup>	β-CD
			54	γ-CD
			65	Without CD
				(in CH <sub>3</sub> CN/H <sub>2</sub> O)

<sup>a</sup> Reactions were conducted with substrate (1.0 mmol), and 1,1,1-trifluoroacetone (1.0 mmol) in H<sub>2</sub>O (50 mL) at room temperature with Oxone (2.5 mmol × 8) and NaHCO<sub>3</sub> (7.75 mmol × 8). <sup>b</sup> Yield of isolated product. <sup>c</sup> The substrates 3 and 5 used in the experiments were racemic mixtures. <sup>d</sup> Not detected.



**Fig. 2** Partial <sup>1</sup>H NMR spectra of a mixture of **1** and β-CD in D<sub>2</sub>O (signals of β-CD). Ratios of **1**/β-CD: (a) 0 : 10, (b) 1 : 10, (c) 2 : 10, (d) 4 : 10, (e) 6 : 10, (f) 8 : 10, (g) 1 : 1, (h) 12 : 10, (i) 15 : 10, (j) 2 : 1.

in water, it is of interest to study the C–H bond hydroxylation selectivity in a mixture of *trans* and *cis* isomers. In this regard, a 1 : 1 mixture of *trans*- and *cis*-1,3-dimethylcyclohexane (**3** and **4**) was chosen to study the effect of CDs on the selectivity of C–H bond hydroxylation of the *trans* and *cis* isomers.

As shown in Scheme 1, when a 1 : 1 mixture of *trans*- and *cis*-1,3-dimethylcyclohexane (**3** and **4**) was subjected to the oxidation conditions with β-CD as the supramolecular host, the ratio of the corresponding products **3a** and **4a** was found to be 1 : 1. This result is consistent with the outcomes of the individual oxidations of *trans*- and *cis*-1,3-dimethylcyclohexane (**3** and **4**) under supramolecular control by β-CD. Using γ-CD as the supramolecular host, a 2 : 1 ratio of **3a** and **4a** was obtained. These findings indicate that γ-CD could provide enhanced isomer-selective C–H bond oxidation for 1,3-dimethylcyclohexane (**3** and **4**).



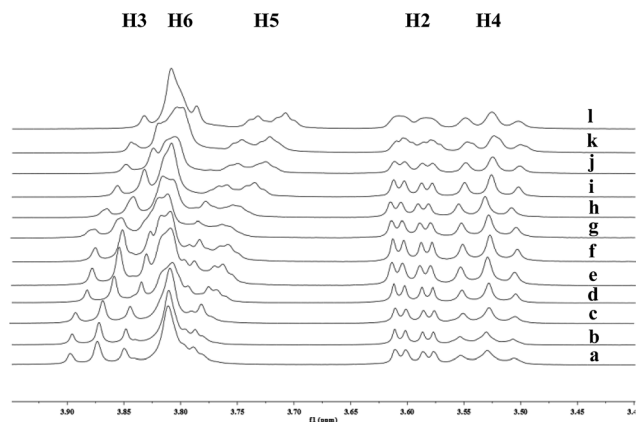


Fig. 3 Partial  $^1\text{H}$  NMR spectra of a mixture of **1** and  $\gamma$ -CD in  $\text{D}_2\text{O}$  (signals of  $\gamma$ -CD). Ratios of **1**/ $\gamma$ -CD: (a) 0 : 10, (b) 1 : 10, (c) 4 : 10, (d) 7 : 10, (e) 1 : 1, (f) 12 : 10, (g) 15 : 10, (h) 18 : 10, (i) 2 : 1, (j) 3 : 1, (k) 3.5 : 1, (l) 4 : 1.

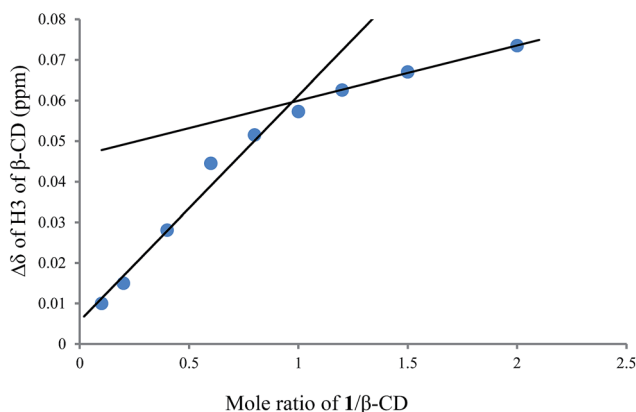


Fig. 4  $^1\text{H}$  NMR titration curve for **1** and  $\beta$ -CD.

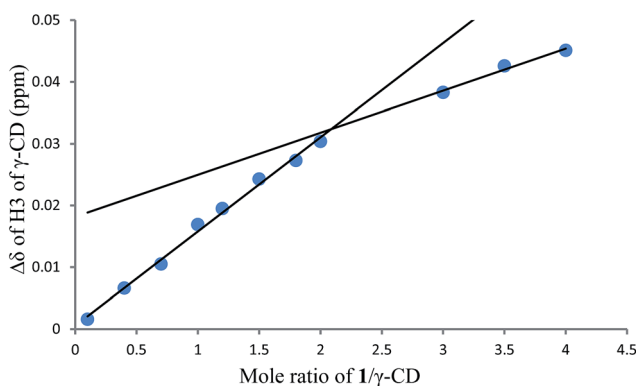
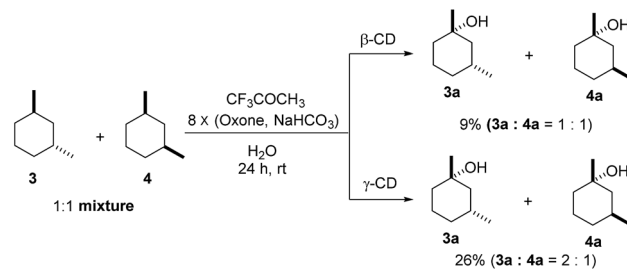
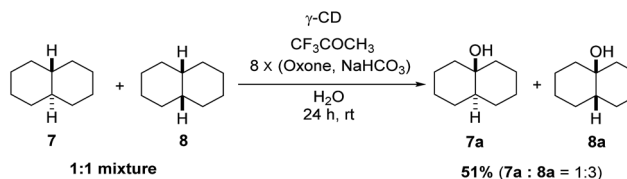


Fig. 5  $^1\text{H}$  NMR titration curve for **1** and  $\gamma$ -CD.

Using  $\gamma$ -CD as the supramolecular host, a 1 : 1 mixture of *trans*- and *cis*-decahydronaphthalene (**7** and **8**) was oxidized giving the corresponding products **7a** and **8a** in a 1 : 3 ratio (Scheme 2). In contrast, the ratio of **7a** and **8a** was 1 : 1.4, when **7** and **8** were individually oxidized in water with  $\gamma$ -CD as the supramolecular host.



Scheme 1 Hydroxylation of a 1 : 1 mixture of **3** and **4** with  $\beta$ -CD and  $\gamma$ -CD as the supramolecular hosts. Reactions were conducted with a 1 : 1 mixture of **3** (0.5 mmol) and **4** (0.5 mmol), 1,1,1-trifluoroacetone (1.0 mmol), and  $\beta$ -CD or  $\gamma$ -CD (1.1 mmol) in water (50 mL) with Oxone (2.5 mmol  $\times$  8) and  $\text{NaHCO}_3$  (7.75 mmol  $\times$  8).



Scheme 2 Hydroxylation of a 1 : 1 mixture of **7** and **8** with  $\gamma$ -CD as the supramolecular host. Reactions were conducted with a 1 : 1 mixture of **7** (0.5 mmol) and **8** (0.5 mmol), 1,1,1-trifluoroacetone (1.0 mmol), and  $\gamma$ -CD (1.1 mmol) in water (50 mL) with Oxone (2.5 mmol  $\times$  8) and  $\text{NaHCO}_3$  (7.75 mmol  $\times$  8).

### The effect of CDs on site-selective C–H bond hydroxylation of (1*S*,2*R*,5*S*)-(+)-menthol (**9**)

Menthol is an easily accessible chiral monoterpene possessing a cyclohexane skeleton with different types of C–H bonds that is an ideal substrate for studying C–H bond hydroxylation. To investigate the effects of  $\beta$ -CD and  $\gamma$ -CD on site-selective C–H bond hydroxylation, (1*R*,2*R*,5*S*)-(+)-menthol (**9**), which bears three tertiary C–H bonds at C2', C5' and C8', was chosen as the substrate.

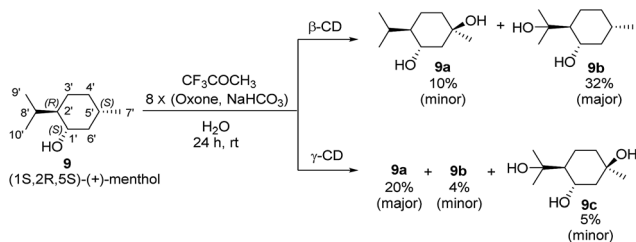
The reaction was carried out at room temperature by stirring **9** (1.0 mmol) with  $\beta$ -CD or  $\gamma$ -CD (1.1 mmol) and 1,1,1-trifluoroacetone (1.0 mmol) in water (50 mL) for 1 h, and then adding 8 portions of a mixture of Oxone (2.5 mmol  $\times$  8) and  $\text{NaHCO}_3$  (7.75 mmol  $\times$  8) within 7 h (one portion was added per hour). After stirring for an additional 16 h, products **9a** and **9b** were isolated by flash column chromatography (Scheme 3).

As shown in Scheme 3, using  $\beta$ -CD as the supramolecular host, the C5'-hydroxylated product **9a** (10% yield; minor) and the C8'-hydroxylated product **9b** (32% yield; major) were obtained with 20% conversion. Using  $\gamma$ -CD as the supramolecular host, significantly increased conversion (50%) was achieved, and the product distribution was reversed with respect to **9a** (20% yield; major) and **9b** (4% yield; minor). In addition, a dihydroxylation product **9c** (5% yield) was obtained when  $\gamma$ -CD was used.

### $^1\text{H}$ NMR titration experiments for binding of **9** with $\beta$ -CD and $\gamma$ -CD

To understand the different product distributions in the hydroxylation of **9** using  $\beta$ -CD and  $\gamma$ -CD as the supramolecular





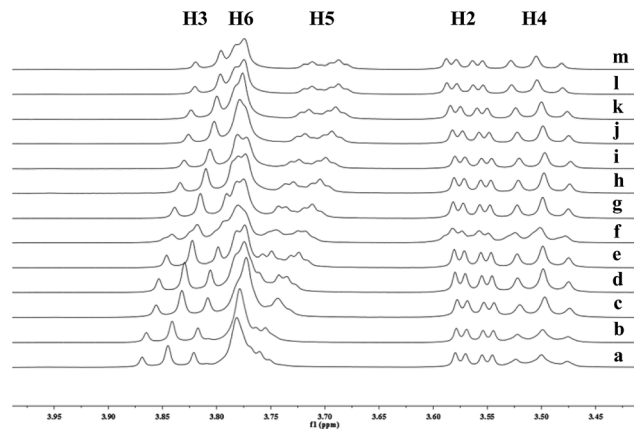
**Scheme 3** Hydroxylation of (+)-menthol (**9**) with  $\beta$ -CD and  $\gamma$ -CD as supramolecular hosts. Reactions were conducted with **9** (1.0 mmol), 1,1,1-trifluoroacetone (1.0 mmol), and  $\beta$ -CD or  $\gamma$ -CD (1.1 mmol) in water (50 mL) with Oxone (2.5 mmol  $\times$  8) and  $\text{NaHCO}_3$  (7.75 mmol  $\times$  8). The yields were calculated based on the conversion of **9**.

hosts, we studied inclusion complexation between **9** and  $\beta$ -CD or  $\gamma$ -CD by  $^1\text{H}$  NMR titration. The peaks of H3 and H5 of the CDs showed significant upfield shifts when the amount of **9** increased in the aqueous solution of CDs (Fig. 6 and 7). Two  $^1\text{H}$  NMR titration curves were plotted based on the change in the chemical shift of H3 against the ratio of **9**/ $\beta$ -CD (Fig. 8) and **9**/ $\gamma$ -CD (Fig. 9). The stoichiometry of the inclusion complex between  $\beta$ -CD and **9** was found to be 1 : 1 and that between  $\gamma$ -CD and **9** was 1 : 2, as determined by extrapolating the curves.

### Studies on the binding geometry of (+)-menthol (**9**) with $\beta$ -CD and $\gamma$ -CD through 2D ROESY experiments

To further investigate the effect of  $\beta$ -CD and  $\gamma$ -CD on site-selective C–H bond oxidation of (+)-menthol (**9**), 2D ROESY experiments were conducted to determine the binding geometry of **9** with  $\beta$ -CD and  $\gamma$ -CD.

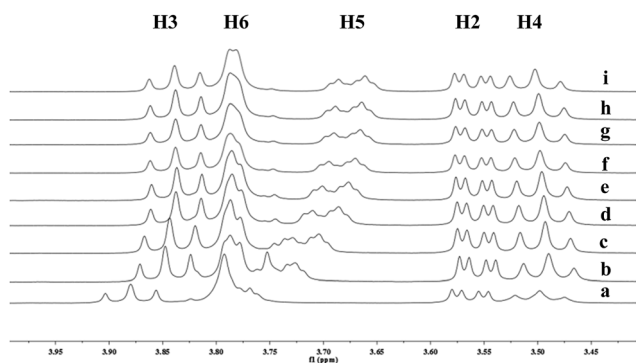
The 2D ROESY spectrum for the binding between **9** and  $\beta$ -CD showed that **9** has NOE correlation signals with hydrogens H3, H5 and H6 of  $\beta$ -CD (Fig. 10), which suggested that **9** was located inside the cavity of  $\beta$ -CD. As shown in Fig. 10, the 2D ROESY spectrum regarding the binding between **9** and  $\beta$ -CD showed that H9' and H10' of **9** have strong NOE correlation signals with H3 and H5 of  $\beta$ -CD and weak NOE correlation signals with H6 of  $\beta$ -CD. Moreover, strong NOE correlation signals of H7' of **9** with H5 of  $\beta$ -CD and weak NOE correlation signals of H7' of **9** with



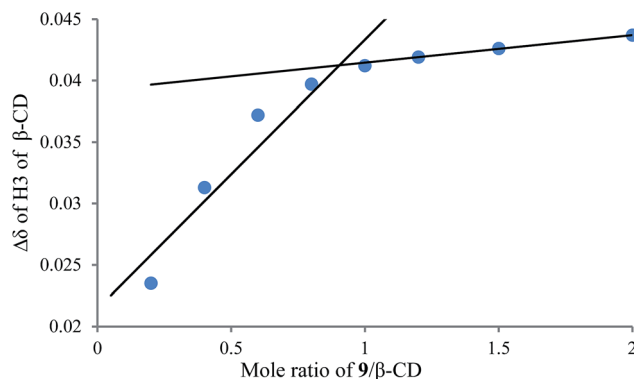
**Fig. 7** Partial  $^1\text{H}$  NMR spectra of a mixture of **9** and  $\gamma$ -CD in  $\text{D}_2\text{O}$  (signals of  $\gamma$ -CD). Ratios of **9**/ $\gamma$ -CD: (a) 0 : 10, (b) 1 : 10, (c) 4 : 10, (d) 7 : 10, (e) 1 : 1, (f) 12 : 10, (g) 15 : 10, (h) 18 : 10, (i) 2 : 1, (j) 2.5 : 1, (k) 3 : 1, (l) 3.5 : 1, (m) 4 : 1.

H6 of  $\beta$ -CD also appeared on the ROESY spectrum. These NOE correlation signals and the 1 : 1 stoichiometry of the complexation between **9** and  $\beta$ -CD suggested that  $\beta$ -CD most likely binds to **9** in two possible binding models: (1) the terminal methyl (H7') of **9** located deep in the  $\beta$ -CD cavity with the terminal isopropyl (H9', H10') of **9** outside of the cavity and close to the secondary OH-face (Fig. 11, binding mode A); and (2) the terminal isopropyl (H9', H10') of **9** located deep in the  $\beta$ -CD cavity with the terminal methyl (H7') of **9** outside of the cavity and close to the primary OH-face (Fig. 11, binding mode B). Further analysis of the strength of the NOE correlation signals indicated that  $\beta$ -CD prefers to bind to **9** in binding mode A rather than binding mode B.

For **9** and  $\gamma$ -CD, the 2D ROESY spectrum (Fig. 12) showed that H7' of **9** has NOE correlation signals with H1', H8' of **9** and H5' of **9** have NOE correlation signals with H9', H10' of **9**. The 2D ROESY spectrum also showed that H3, H5, and H6 of  $\gamma$ -CD have strong NOE correlation signals with H7', H9', and H10' of **9**; H6 of  $\gamma$ -CD has NOE correlation signals with H2', H3', H5', H7', H9', and H10' of **9**; and H5 of  $\gamma$ -CD has NOE correlation signals with H4', H5', H6', and H8' of **9**. Given the 2 : 1 stoichiometry of the complexation between **9** and  $\gamma$ -CD, the



**Fig. 6** Partial  $^1\text{H}$  NMR spectra of a mixture of **9** and  $\beta$ -CD in  $\text{D}_2\text{O}$  (signals of  $\beta$ -CD). Ratios of **9**/ $\beta$ -CD: (a) 0 : 10, (b) 2 : 10, (c) 4 : 10, (d) 6 : 10, (e) 8 : 10, (f) 1 : 1, (g) 12 : 10, (h) 15 : 10, (i) 2 : 1.



**Fig. 8**  $^1\text{H}$  NMR titration curve for **9** and  $\beta$ -CD.



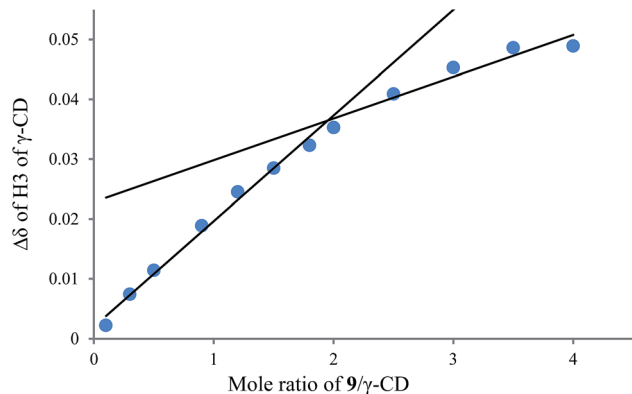


Fig. 9  $^1\text{H}$  NMR titration curve for **9** and  $\gamma$ -CD.

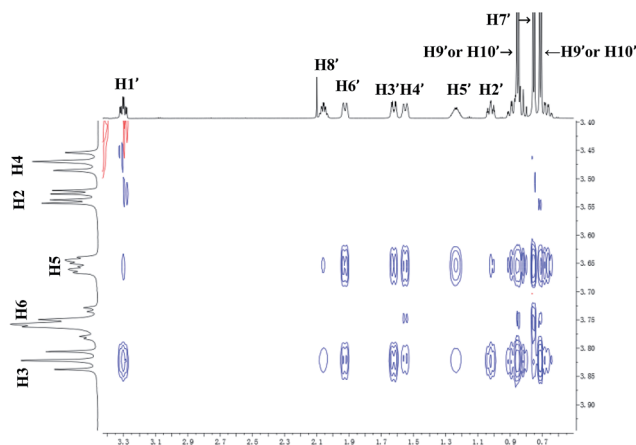


Fig. 10 Partial contour plot of the 600 MHz 2D ROESY spectrum for binding of (+)-menthol (**9**) to  $\beta$ -CD in  $\text{D}_2\text{O}$ .

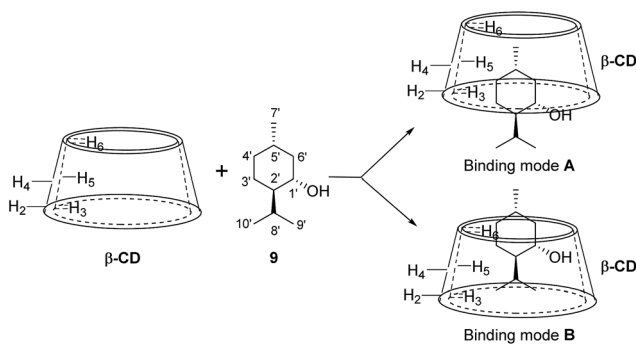


Fig. 11 Proposed binding geometry for the inclusion of (+)-menthol (**9**) in  $\beta$ -CD.

possible binding modes involve two molecules of **9** encased in the cavity of  $\gamma$ -CD. Therefore, two possible binding modes between  $\gamma$ -CD and **9** were proposed: (1) the terminal isopropyl ( $\text{H}9'$ ,  $\text{H}10'$ ) of **9** inserted deep into the  $\gamma$ -CD cavity with the terminal methyl ( $\text{H}7'$ ) of **9** outside of the cavity and close to the secondary OH-face (Fig. 13, binding mode C); and (2) the terminal methyl ( $\text{H}7'$ ) of **9** inserted deep into the  $\gamma$ -CD cavity

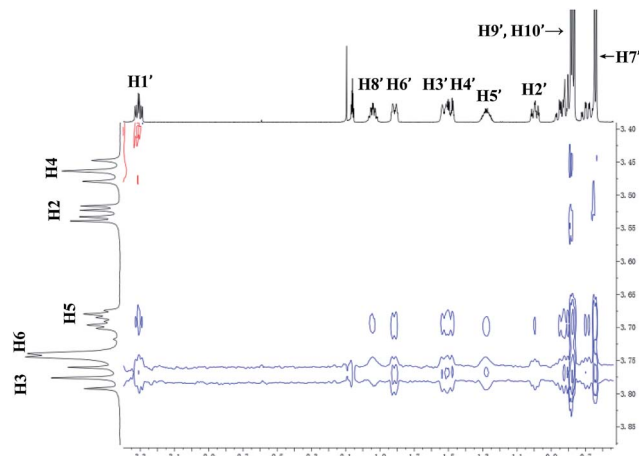


Fig. 12 Partial contour plot of the 600 MHz 2D ROESY spectrum for binding of (+)-menthol (**9**) to  $\gamma$ -CD in  $\text{D}_2\text{O}$ .

with the terminal isopropyl ( $\text{H}9'$ ,  $\text{H}10'$ ) of **9** outside of the cavity and close to the primary OH-face (Fig. 13, binding mode D). The experimental results indicated that binding mode C is the favored binding geometry.

#### Schematic diagrams for site-selective C–H bond oxidation of **9** with $\beta$ -CD and $\gamma$ -CD as supramolecular hosts

Based on the binding geometries of **9** to  $\beta$ -CD (Fig. 11), schematic diagrams for how  $\beta$ -CD affects the site-selectivity of C–H bond hydroxylation were proposed in Fig. 14. The preferred binding geometry of **9** to  $\beta$ -CD, in which the terminal methyl of **9** is located deep in the cavity of  $\beta$ -CD and the terminal isopropyl of **9** is exposed outside of the cavity, renders the C–H bond at position  $\text{C}8'$  more easily subject to hydroxylation by dioxirane than that at position  $\text{C}5'$ . This leads to site-selectivity in the C–H bond hydroxylation. The proposed schematic diagrams are reasonably consistent with the experimental results (see Scheme 3, **9a/9b** = 1 : 3) on the site-selectivity of C–H bond hydroxylation of **9** when using  $\beta$ -CD as the supramolecular host.

In contrast, for the binding of **9** to  $\gamma$ -CD, the preferred binding geometry is such that the C–H bond at position  $\text{C}5'$  is

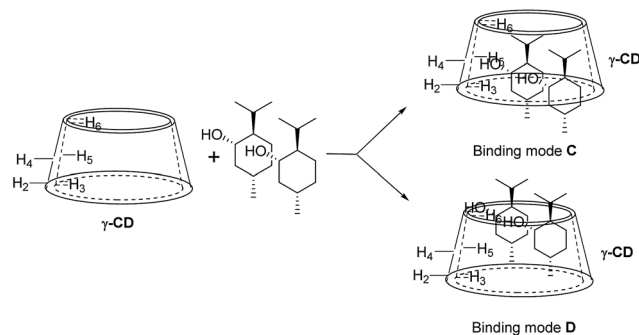


Fig. 13 Proposed binding geometry for the inclusion of (+)-menthol (**9**) in  $\gamma$ -CD.



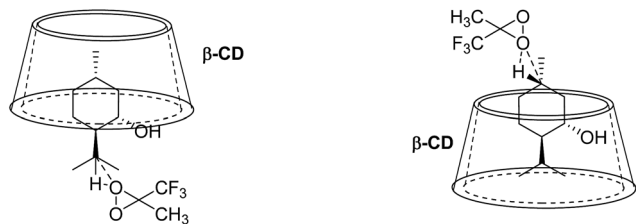


Fig. 14 Schematic diagrams for site-selective C–H bond oxidation of **9** in  $\beta$ -CD.

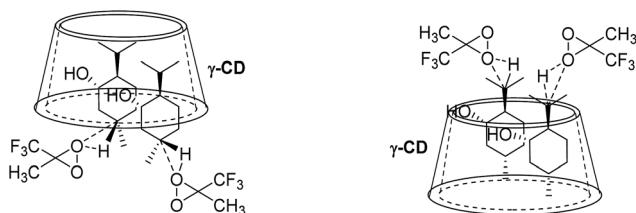


Fig. 15 Schematic diagrams for site-selective C–H bond oxidation of **9** in  $\gamma$ -CD.

more easily approached by dioxirane than that at position C8', which gives **9a** as the major product (Fig. 15). Thus, the binding geometries of **9** to  $\gamma$ -CD and steric hindrance lead to high site-selectivity in the C–H bond hydroxylation of **9** (**9a** : **9b** = 5 : 1).

## Conclusions

In summary, we have investigated the hydroxylation of various cyclohexanes with dioxirane generated *in situ* through supramolecular control using  $\beta$ -CD and  $\gamma$ -CD as the hosts. The results indicated that CDs could act as reaction vessels for hydroxylation reactions performed in water without the use of organic solvents. The stoichiometries of inclusion complexation between  $\beta$ -CD or  $\gamma$ -CD and cyclohexanes were determined by  $^1\text{H}$  NMR titration. The effect of  $\beta$ -CD and  $\gamma$ -CD as supramolecular hosts on chemoselective C–H bond hydroxylation of *trans* and *cis* isomers of disubstituted cyclohexanes was investigated. Site-selective C–H bond hydroxylation of (+)-menthol (**9**) was achieved through inclusion complex formation with  $\beta$ -CD and  $\gamma$ -CD in water.

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## Notes and references

- J. F. Hartwig, *Acc. Chem. Res.*, 2017, **50**, 549–555.
- V. C. C. Wang, S. Maji, P. P. Y. Chen, H. K. Lee, S. S. F. Yu and S. I. Chan, *Chem. Rev.*, 2017, DOI: 10.1021/acs.chemrev.6b00624.

- T. Newhouse and P. S. Baran, *Angew. Chem., Int. Ed.*, 2011, **50**, 3362–3374.
- S. R. Neufeldt and M. S. Sanford, *Acc. Chem. Res.*, 2012, **45**, 936–946.
- T. Brückl, R. D. Baxter, Y. Ishihara and P. S. Baran, *Acc. Chem. Res.*, 2012, **45**, 826–839.
- W. R. Gutekunst and P. S. Baran, *Chem. Soc. Rev.*, 2011, **40**, 1976–1991.
- Y. Qin, L. Zhu and S. Luo, *Chem. Rev.*, 2017, DOI: 10.1021/acs.chemrev.6b00657.
- M. C. Feiters, A. E. Rowan and R. J. M. Nolte, *Chem. Soc. Rev.*, 2000, **29**, 375–384.
- B. H. Brodsky and J. Du Bois, *J. Am. Chem. Soc.*, 2005, **127**, 15391–15393.
- M. Milan, M. Bietti and M. Costas, *ACS Cent. Sci.*, 2017, **3**, 196–204.
- X. Li, X. Che, G.-H. Chen, J. Zhang, J.-L. Yan, Y.-F. Zhang, L.-S. Zhang, C.-P. Hsu, Y. Q. Gao and Z.-J. Shi, *Org. Lett.*, 2016, **18**, 1234–1237.
- F. Kakiuchi and N. Chatani, *Adv. Synth. Catal.*, 2003, **345**, 1077–1101.
- H. M. Davies, *Angew. Chem., Int. Ed.*, 2006, **45**, 6422–6425.
- J. A. Labinger and J. E. Bercaw, *Nature*, 2002, **417**, 507–514.
- M. C. White, *Science*, 2012, **335**, 807–809.
- M. S. Chen and M. C. White, *Science*, 2010, **327**, 566–571.
- M. A. Bigi, S. A. Reed and M. C. White, *Nat. Chem.*, 2011, **3**, 216–222.
- E. McNeill and J. Du Bois, *J. Am. Chem. Soc.*, 2010, **132**, 10202–10204.
- K. Zhang, B. M. Shafer, M. D. Demars, H. A. Stern and R. Fasan, *J. Am. Chem. Soc.*, 2012, **134**, 18695–18704.
- S. Das, C. D. Incarvito, R. H. Crabtree and G. W. Brudvig, *Science*, 2006, **312**, 1941–1943.
- J. Yang, B. Gabriele, S. Belvedere, Y. Huang and R. Breslow, *J. Org. Chem.*, 2002, **67**, 5057–5067.
- B. R. Cook, T. J. Reinert and K. S. Suslick, *J. Am. Chem. Soc.*, 1986, **108**, 7281–7286.
- L. Que Jr and W. B. Tolman, *Nature*, 2008, **455**, 333–340.
- T. Gensch, M. N. Hopkinson, F. Glorius and J. Wencel-Delord, *Chem. Soc. Rev.*, 2016, **45**, 2900–2936.
- B. D. Dangel, J. A. Johnson and D. Sames, *J. Am. Chem. Soc.*, 2001, **123**, 8149–8150.
- A. Mendoza, Y. Ishihara and P. S. Baran, *Nat. Chem.*, 2011, **4**, 21–25.
- T. J. Maimone and P. S. Baran, *Nat. Chem. Biol.*, 2007, **3**, 396–407.
- K. Chen and P. S. Baran, *Nature*, 2009, **459**, 824–828.
- J. R. Hanson, *Nat. Prod. Rep.*, 2010, **27**, 887–899.
- J.-N. Rebilly, B. Colasson, O. Bistri, D. Over and O. Reinaud, *Chem. Soc. Rev.*, 2015, **44**, 467–489.
- Y. Murakami, J.-I. Kikuchi, Y. Hisaeda and O. Hayashida, *Chem. Rev.*, 1996, **96**, 721–758.
- D. H. Leung, D. Fiedler, R. G. Bergman and K. N. Raymond, *Angew. Chem., Int. Ed.*, 2004, **43**, 963–966.
- M. Yoshizawa, J. K. Klosterman and M. Fujita, *Angew. Chem., Int. Ed.*, 2009, **48**, 3418–3438.
- R. Breslow and S. D. Dong, *Chem. Rev.*, 1998, **98**, 1997–2012.



- 35 K. Takahashi, *Chem. Rev.*, 1998, **98**, 2013–2034.
- 36 L. Marinescu and M. Bols, *Curr. Org. Synth.*, 2010, **14**, 1380–1398.
- 37 W.-K. Chan, W.-Y. Yu, C.-M. Che and M.-K. Wong, *J. Org. Chem.*, 2003, **68**, 6576–6582.
- 38 Y. S. Fung, S. C. Yan and M. K. Wong, *Org. Biomol. Chem.*, 2012, **10**, 3122–3130.
- 39 R. Breslow, X. Zhang, R. Xu, M. Maletic and R. Merger, *J. Am. Chem. Soc.*, 1996, **118**, 11678–11679.
- 40 R. Breslow, X. Zhang and Y. Huang, *J. Am. Chem. Soc.*, 1997, **119**, 4535–4536.
- 41 C. Rousseau, B. Christensen, T. E. Petersen and M. Bols, *Org. Biomol. Chem.*, 2004, **2**, 3476–3482.
- 42 L. Marinescu, M. Mølbach, C. Rousseau and M. Bols, *J. Am. Chem. Soc.*, 2005, **127**, 17578–17579.
- 43 T. H. Fenger, L. G. Marinescu and M. Bols, *Eur. J. Org. Chem.*, 2011, **2011**, 2339–2345.
- 44 T. H. Fenger, L. G. Marinescu and M. Bols, *Org. Biomol. Chem.*, 2009, **7**, 933–943.
- 45 J. Bjerre, T. Hauch Fenger, L. G. Marinescu and M. Bols, *Eur. J. Org. Chem.*, 2007, **2007**, 704–710.
- 46 C. I. Herreras, X. Yao, Z. Li and C.-J. Li, *Chem. Rev.*, 2007, **107**, 2546–2562.
- 47 T.-W. Hui, J.-F. Cui and M.-K. Wong, *RSC Adv.*, 2017, **7**, 14477–14480.
- 48 W. Adam, R. Curci and J. O. Edwards, *Acc. Chem. Res.*, 1989, **22**, 205–211.
- 49 R. W. Murray, *Chem. Rev.*, 1989, **89**, 1187–1201.
- 50 S. E. Denmark and Z. Wu, *Synlett*, 1999, **1999**, 847–859.
- 51 Y. Shi, *Acc. Chem. Res.*, 2004, **37**, 488–496.
- 52 D. Yang, *Acc. Chem. Res.*, 2004, **37**, 497–505.
- 53 O. A. Wong and Y. Shi, *Chem. Rev.*, 2008, **108**, 3958–3987.
- 54 R. Curci, L. D'Accolti and C. Fusco, *Acc. Chem. Res.*, 2006, **39**, 1–9.
- 55 R. W. Murray, R. Jeyaraman and L. Mohan, *J. Am. Chem. Soc.*, 1986, **108**, 2470–2472.
- 56 G. Asensio, M. E. Gonzalez-Nunez, C. B. Bernardini, R. Mello and W. Adam, *J. Am. Chem. Soc.*, 1993, **115**, 7250–7253.
- 57 D. Yang, M.-K. Wong, X.-C. Wang and Y.-C. Tang, *J. Am. Chem. Soc.*, 1998, **120**, 6611–6612.
- 58 M.-K. Wong, N.-W. Chung, L. He and D. Yang, *J. Am. Chem. Soc.*, 2003, **125**, 158–162.
- 59 M.-K. Wong, N.-W. Chung, L. He, X.-C. Wang, Z. Yan, Y.-C. Tang and D. Yang, *J. Org. Chem.*, 2003, **68**, 6321–6328.
- 60 C. J. Abelt and J. M. Pleier, *J. Org. Chem.*, 1988, **53**, 2159–2162.
- 61 P. W. N. M. van Leeuwen, *Supramolecular Catalysis*, Wiley-VCH Verlag GmbH & Co. KGaA, 2008.
- 62 J. Kang and J. Rebek Jr, *Nature*, 1997, **385**, 50–52.
- 63 D. M. Vriezema, M. Comellas Aragonés, J. A. A. W. Elemans, J. J. L. M. Cornelissen, A. E. Rowan and R. J. M. Nolte, *Chem. Rev.*, 2005, **105**, 1445–1489.
- 64 J. Lehn and C. Sirlin, *J. Chem. Soc., Chem. Commun.*, 1978, 949–951.
- 65 L. Xu, G. Fang and S. Li, *RSC Adv.*, 2017, **7**, 14046–14052.
- 66 M. Zubiaur and C. Jaime, *J. Org. Chem.*, 2000, **65**, 8139–8145.
- 67 H.-J. Schneider, F. Hacket, V. Rüdiger and H. Ikeda, *Chem. Rev.*, 1998, **98**, 1755–1786.
- 68 D. Salvatierra, C. Jaime, A. Virgili and F. Sánchez-Ferrando, *J. Org. Chem.*, 1996, **61**, 9578–9581.
- 69 A. Botsi, K. Yannakopoulou, B. Perly and E. Hadjoudis, *J. Org. Chem.*, 1995, **60**, 4017–4023.
- 70 D. Salvatierra, X. Sánchez-Ruiz, R. Garduño, E. Cervelló, C. Jaime, A. Virgili and F. Sánchez-Ferrando, *Tetrahedron*, 2000, **56**, 3035–3041.

