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# Mechanistic study on ligand-controlled copper-catalyzed regiodivergent silacarboxylation of allenes with carbon dioxide and silylborane†

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Recently, Tsuji and co-workers reported a ligand controlled regiodivergent silacarboxylation of allenes with CO<sub>2</sub> and a silylborane in the presence of a copper catalyst. Product **P**<sub>1</sub> was generated in the *rac*-Me-DuPhos (**L**<sub>A</sub>) assisted system while product **P**<sub>2</sub> was obtained in the PCy<sub>3</sub> (**L**<sub>B</sub>) assisted system. To account for the observed regioselectivity, density functional theory (DFT) calculations have been carried out in the present study. We calculated the detailed mechanisms and analyzed electronic and steric factors in the selectivity-determining states **TS**<sub>3A</sub>, **TS**<sub>9A</sub>, **TS**<sub>2B</sub>, and **TS**<sub>5B</sub> to gain insight into the origin of the observed regioselectivity. For the **L**<sub>A</sub>-assisted system, the regioselectivity is predicted to be determined by the great steric effects of **TS**<sub>9A</sub>. The more stable geometry of **TS**<sub>3A</sub> causes **P**<sub>1</sub> to be the major product. For the **L**<sub>B</sub>-assisted system, electronic effects can account for the regioselectivity. The stronger attractive interaction in **TS**<sub>5B</sub> compared with that in **TS**<sub>2B</sub> leads to the generation of **P**<sub>2</sub>.

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## 1. Introduction

As useful building blocks in organic synthesis, organosilanes can undergo a good deal of silicon-assisted transformations.<sup>1</sup> In particular, silylcupration of C–C multiple bonds (especially the allene derivatives) has become one of the most reliable and powerful procedures that forms both vinyl- and allyl-silanes.<sup>2</sup> However, how to control the selectivity has become the serious problem since the addition reactions across unsaturated substrates often result in many regio- and stereo-isomers.<sup>2g,3</sup> The first-row transition metals have often been used as catalysts to solve the problem,<sup>4</sup> not only due to their considerably low price, but the differing electronic structures of 3d transition metals provide promise for developing significant molecular structures and uncovering new organometallic reactivities at a fundamental level.<sup>5</sup> In addition, as a greenhouse gas providing serious harm to the environment, carbon dioxide (CO<sub>2</sub>) capture, utilization, and conversion have become hot topics in research.<sup>6</sup>

Recently, Tsuji and co-workers reported a ligand controlled regiodivergent silacarboxylation of allenes with CO<sub>2</sub> and a silylborane in the presence of a copper catalyst (Scheme 1).<sup>7</sup> As shown in Scheme 1, the ratio of two different products **P**<sub>1</sub> and **P**<sub>2</sub> can be highly controlled by the proper choice of ligand. The major product is **P**<sub>1</sub> in reaction condition A (when Cu(OAc) is

used as the catalyst precursor, hexane as the solvent, and the *rac*-Me-DuPhos (**L**<sub>A</sub>) as the ligand). In contrast, the major product is changed to **P**<sub>2</sub> when a mixture of CuCl/NaOAc was used as the catalyst precursor, THF as the solvent, and PCy<sub>3</sub> (**L**<sub>B</sub>) as the ligand (reaction condition B, Scheme 1).

To account for the observed regioselectivity of this reaction, a plausible mechanism has been proposed by Tsuji *et al.*<sup>7</sup> As shown in Scheme 2, a Cu catalyst precursor first reacts with one reactant **R**<sub>1</sub> to form a silylcopper complex **1**. Next, another reactant **R**<sub>2</sub> coordinates to the Cu center of **1** via the terminal double bond and generates two isomers **2** and **5** according to the orientation of the unsymmetrical allene **R**<sub>2</sub>, which is followed by the subsequent alkene insertion to form the allylcopper complex **3** and vinylcopper complex **6**, respectively. With CO<sub>2</sub> inserting at the  $\gamma$ -position of **3**, copper carboxylate complex **4** is afforded. Finally, through the  $\sigma$ -bond metathesis between **4** and **R**<sub>1</sub>, boron carboxylate **P**<sub>0-1</sub> is produced, accompanied with the **1** regenerating. Through the conversion with H<sub>3</sub>O<sup>+</sup> and Me<sub>3</sub>SiCHN<sub>2</sub> further, **P**<sub>0-1</sub> could afford the final product **P**<sub>1</sub>. In contrast, the insertion of CO<sub>2</sub> into **6** generates intermediate **7**, followed by the  $\sigma$ -bond metathesis with **R**<sub>1</sub> to provide boron carboxylate **P**<sub>0-2</sub> and regenerate **1**. Finally, final product **P**<sub>2</sub> could be obtained through the further conversion of **P**<sub>0-2</sub> with H<sub>3</sub>O<sup>+</sup> and Me<sub>3</sub>SiCHN<sub>2</sub>.

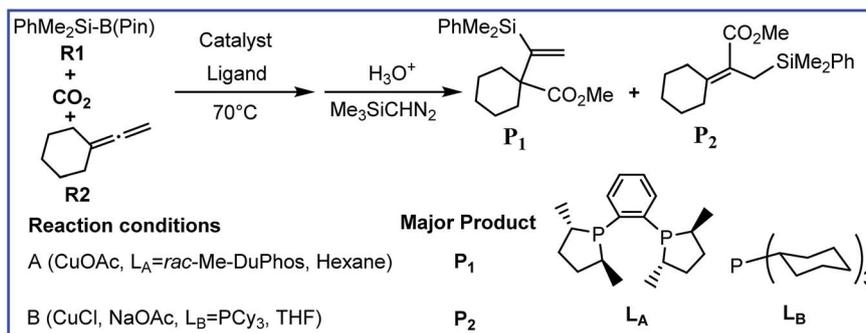
Although a plausible mechanistic pathway has been proposed by Tsuji group, some key issues still need to be further discussed. (1) What is the detailed reaction mechanism? (2) How is the origin of this high ligand-controlled regioselectivity? Here we would like to report our detailed density functional theory (DFT) calculations on the reaction mechanisms in order to gain insight into the interesting experimental observations

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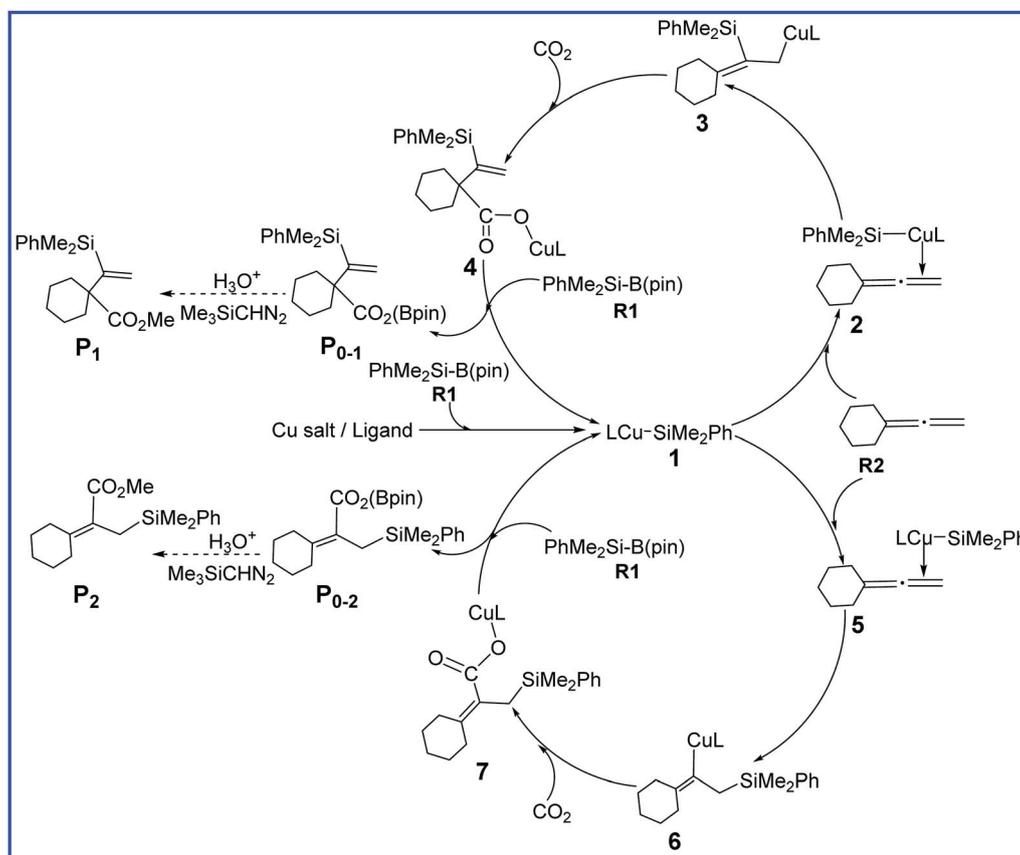
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Scheme 1 Cu-catalyzed silacarboxylation of allenes with CO<sub>2</sub> and silylborane reported by Tsuji group.



Scheme 2 Mechanisms of Cu-catalyzed silacarboxylation of allenes with CO<sub>2</sub> and silylborane proposed by Tsuji group.

and distinct selectivity. We expect this work would help understand the detailed mechanisms and design new related reactions.

## 2. Computational details

In our calculations, the geometries of all intermediates and transition states were optimized at the B3LYP<sup>8</sup> level of DFT in gas phase. A mixed basis set of 6-311+G(d) for Cu and 6-31G(d,p) for other atoms (BSI) was used, which has been proved to be appropriate for the copper-catalyzed reactions.<sup>9</sup> Frequency analysis were carried out to ensure all stationary points as

minima (zero imaginary frequencies) or transition states (one imaginary frequency) at the B3LYP/BSI level. When necessary, intrinsic reaction coordinates (IRC)<sup>10</sup> calculations were applied for transition state to confirm it actually connecting the corresponding two minima. To consider solvent effects, we performed single-point energy calculations for all the gas-phase optimized structures with SMD<sup>11</sup> solvent model at the M06 (ref. 12)/BSII (6-311+G(d) for Cu and 6-311++G(d,p) for other atoms) level. Natural bond orbital (NBO)<sup>13</sup> charges were calculated for selected structures at the M06/BSII level. Free energies (kcal mol<sup>-1</sup>) obtained at the M06/BSII/B3LYP/BSI level were used in



the following discussion. All the calculations were carried out with the GAUSSIAN 09 packages.<sup>14</sup>

### 3. Results and discussion

In this work, we first investigate the  $\text{Cu}/\text{L}_\text{A}$ -catalyzed mechanisms leading to  $\text{P}_{0-1}$  and  $\text{P}_{0-2}$ , respectively (reaction A). Then we study the reaction mechanisms catalyzed by  $\text{Cu}/\text{L}_\text{B}$  (reaction B) to probe the regioselectivity observed experimentally.

#### 3.1 $\text{Cu}/\text{L}_\text{A}$ -catalyzed reaction mechanisms

The free energy diagrams for the reaction to lead to product  $\text{P}_{0-1}$  catalyzed by  $\text{Cu}/\text{L}_\text{A}$  are calculated and shown in Fig. 1. Firstly, the precatalyst  $\text{CuOAc}$  (**1A**) coordinates with the ligand  $\text{L}_\text{A}$  to form the active catalyst **2A** by releasing the free energy of  $57.0 \text{ kcal mol}^{-1}$ . Subsequently, the transmetalation of **2A** with **R1** occurs to generate a three-coordinate  $\text{Cu}(\text{i})$  intermediate **3A** and dissociate the complex  $\text{B}(\text{pin})\text{-OAc}$ , through a concerted four-centered transition state **TS1A** related to  $\sigma$ -bond metathesis. This step is thermodynamically and kinetically feasible, with free energy decreasing of  $6.8 \text{ kcal mol}^{-1}$  as well as an attainable activation free

energy ( $8.6 \text{ kcal mol}^{-1}$ ). The terminal double bond of reactant **R2** then coordinates to the Cu center of silylcopper intermediate **3A** along with the dissociation of one arm of  $\text{L}_\text{A}^{15-17}$  via the transition state **TS2A** and generates intermediate **4A** with a barrier of  $8.6 \text{ kcal mol}^{-1}$ . The subsequent alkene insertion step requires a low activation energy barrier of  $7.5 \text{ kcal mol}^{-1}$  via transition state **TS3A** to form the allylcopper intermediate **5A**. The energy barrier is  $14.4 \text{ kcal mol}^{-1}$  for the migratory insertion step (**3A**  $\rightarrow$  **TS3A**). In the next step,  $\text{CO}_2$  inserts at the  $\gamma$ -position of **5A** to provide copper carboxylate complex **7A** via a six-membered-ring transition state **TS4A** with the activation energy barrier of  $12.2 \text{ kcal mol}^{-1}$ . The stepwise pathway for the  $\text{CO}_2$  insertion step ( $\text{CO}_2$  first inserting into the C–Cu bond of **5A** to give carboxylate intermediate **6A**, which then undergoing a [1,3]-sigmatropic rearrangement to afford **7A**) was also considered (see the red line in Fig. 1), but the calculated results indicates that such a proposal is inaccessible due to the forbidden high barrier ( $87.8 \text{ kcal mol}^{-1}$ ). Finally, **7A** combines with **R1** to form intermediate **8A**, which then produces boron carboxylate  $\text{P}_{0-1}$  and regenerates **3A** via a  $\sigma$ -bond metathesis transition state **TS7A** with an attainable free energy barrier  $13.3 \text{ kcal mol}^{-1}$ . Through the conversion with  $\text{H}_3\text{O}^+$  and  $\text{Me}_3\text{SiCHN}_2$  further,  $\text{P}_{0-1}$  could afford the final product **P1**. As shown in Fig. 1,

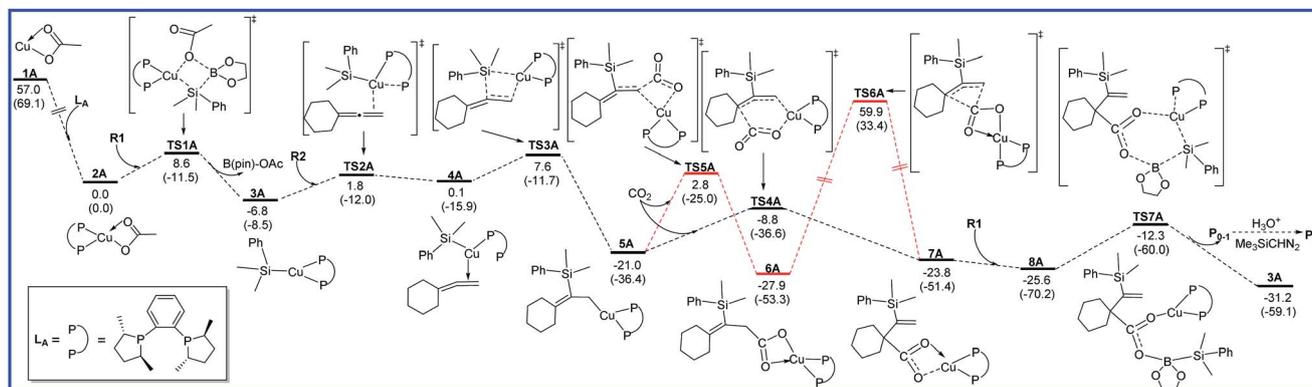


Fig. 1 Free energy diagrams to lead to  $\text{P}_{0-1}$  catalyzed by  $\text{Cu}/\text{L}_\text{A}$ . The relative free energies and relative enthalpic energies (in parentheses) are given in  $\text{kcal mol}^{-1}$ .

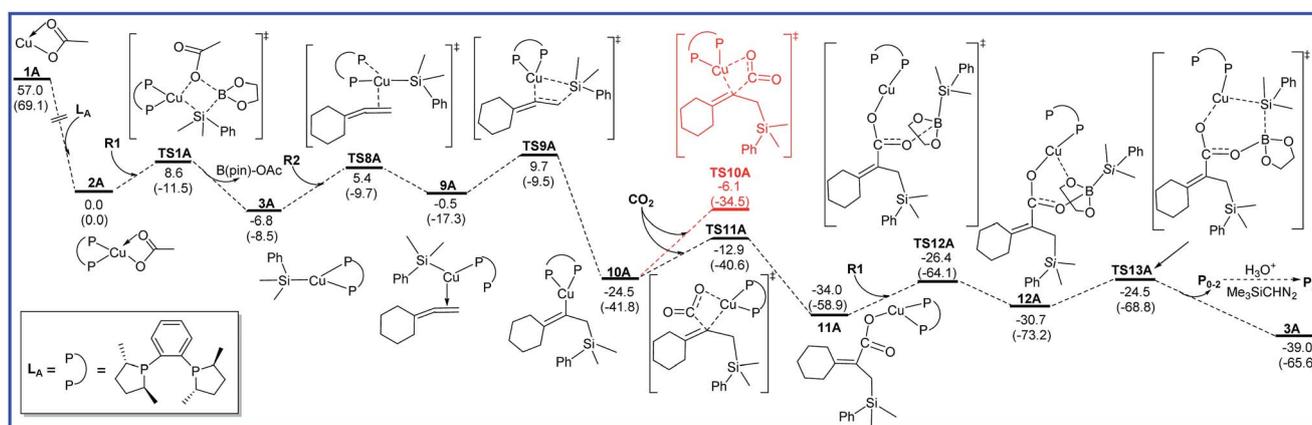


Fig. 2 Free energy diagrams to lead to  $\text{P}_{0-2}$  catalyzed by  $\text{Cu}/\text{L}_\text{A}$ . The relative free energies and relative enthalpic energies (in parentheses) are given in  $\text{kcal mol}^{-1}$ .



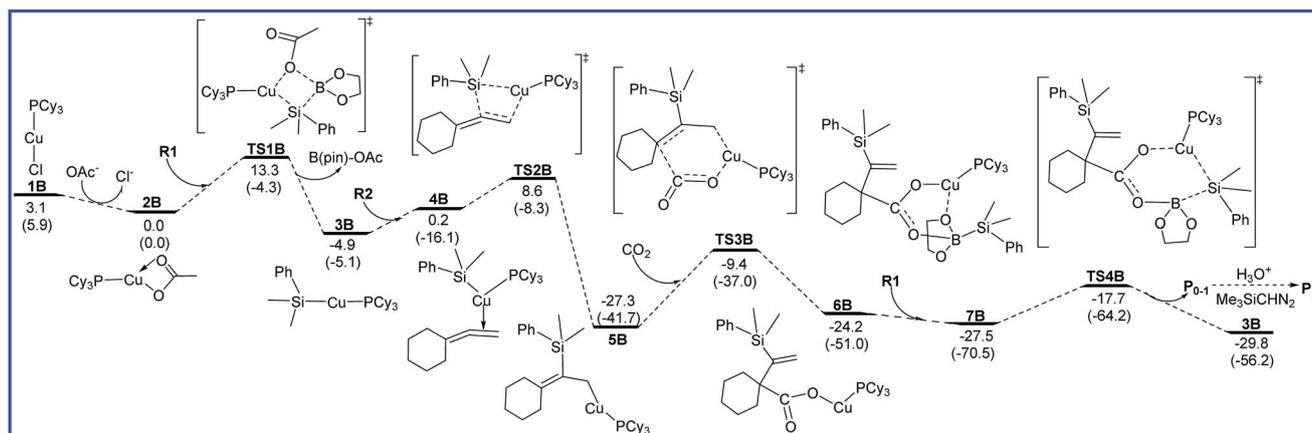


Fig. 3 Free energy diagrams to lead to  $P_{0-1}$  catalyzed by  $Cu/L_B$ . The relative free energies and relative enthalpic energies (in parentheses) are given in kcal mol<sup>-1</sup>.

the rate-determining step in reaction A to produce  $P_{0-1}$  is the alkene migratory insertion step ( $3A \rightarrow TS3A$ ) and the overall activation barrier is 14.4 kcal mol<sup>-1</sup>.

As shown in Fig. 2, R2 can also coordinate with **3A** via another conformation to produce intermediate **9A** with a barrier of 12.2 kcal mol<sup>-1</sup> (**TS8A**). An intramolecular migratory insertion then occurs to form the vinylcopper intermediate **10A** through the transition state **TS9A**, which requires the activation energy barrier of 10.2 kcal mol<sup>-1</sup>. In the following step, CO<sub>2</sub> inserts into the C–Cu bond of **10A** to afford a copper carboxylate intermediate **11A** through a four-membered-ring transition state **TS11A** with a free energy barrier of 11.6 kcal mol<sup>-1</sup>. Herein, another CO<sub>2</sub> insertion transition state **TS10A** was excluded due to the higher free energy compared with **TS11A** (–6.1 vs. –12.9 kcal mol<sup>-1</sup>). Then, **11A** binds up with R1 to form intermediate **12A**, followed by the  $\sigma$ -bond metathesis (**TS13A**) to afford boron carboxylate  $P_{0-2}$  and regenerate **3A** with an attainable free energy barrier 9.5 kcal mol<sup>-1</sup>. Finally, final product  $P_2$  could be obtained through the further conversion of  $P_{0-2}$  with H<sub>3</sub>O<sup>+</sup> and Me<sub>3</sub>SiCHN<sub>2</sub>. It can be found from Fig. 2 that the alkene

migratory insertion process ( $3A \rightarrow TS9A$ ) is also the rate-determining step with the overall activation barrier of 16.5 kcal mol<sup>-1</sup> in reaction A leading to product  $P_{0-2}$ .

Taking the full process shown in Fig. 1 and 2 into account, the alkene migratory insertion step is the regioselectivity-determining step, and the competition between **TS3A** and **TS9A** determines the regioselectivity ( $P_{0-1}$  or  $P_{0-2}$ ). The calculated free energy difference (2.1 kcal mol<sup>-1</sup>) between **TS3A** and **TS9A** is reasonable agreement with the experimental major product  $P_1$ .

### 3.2 Cu/L<sub>B</sub>-catalyzed reaction mechanism

The free energy diagrams for the reaction leading to  $P_{0-1}$  catalyzed by  $Cu/L_B$  are shown in Fig. 3. At the entrance of the reaction, precatalyst PCy<sub>3</sub>CuCl (**1B**) transforms into the active catalyst PCy<sub>3</sub>CuOAc (**2B**) via the ligand substitution with the additive NaOAc. This process is exoergic by 3.1 kcal mol<sup>-1</sup>. **2B** sequentially undergoes  $\sigma$ -bond metathesis with R1 to give the intermediate **3B** with a barrier of 13.3 kcal mol<sup>-1</sup>. The transmetalation pathway of **1B** with R1 was also considered, but the calculated high barrier rules out the possibility (see Fig. S1 in

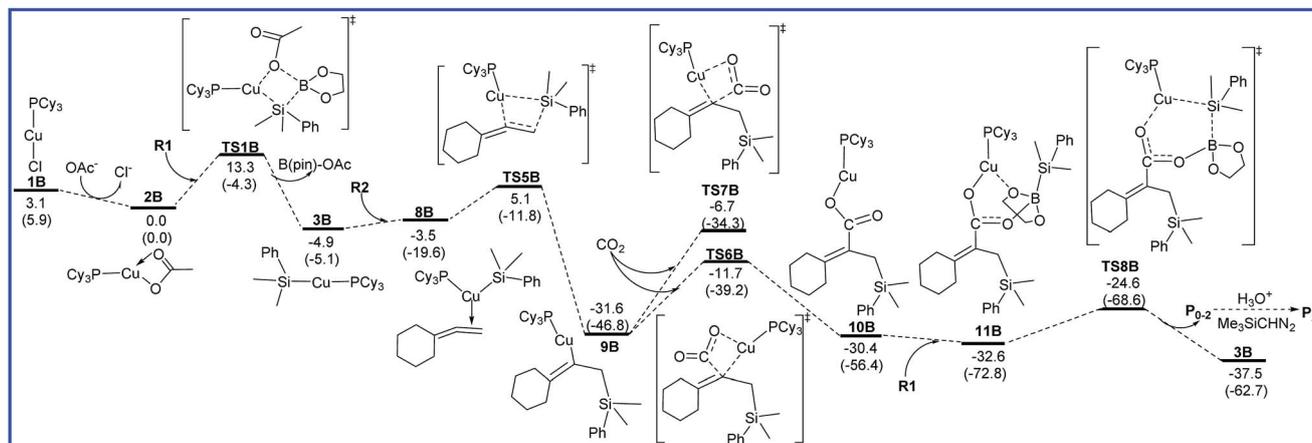


Fig. 4 Free energy diagrams to lead to  $P_{0-2}$  catalyzed by  $Cu/L_B$ . The relative free energies and relative enthalpic energies (in parentheses) are given in kcal mol<sup>-1</sup>.



Table 1 Regioselectivity-determining and rate-determining steps in reactions A and B, along with related activation energies (kcal mol<sup>-1</sup>)

Reactions	Major product	Regioselectivity-determining step <sup>a</sup> ( $\Delta G_1^\ddagger$ )		Rate-determining step <sup>b</sup> ( $\Delta G_2^\ddagger$ )			
Reaction A (Cu/L <sub>A</sub> )	P <sub>1</sub>	Alkene insertion	3A → TS3A	(14.4)	Alkene insertion	3A → TS3A	(14.4)
			3A → TS9A	(16.5)		3A → TS9A	(16.5)
Reaction B (Cu/L <sub>B</sub> )	P <sub>2</sub>	Alkene insertion	3B → TS2B	(13.5)	CO <sub>2</sub> insertion	5B → TS3B	(17.9)
			3B → TS5B	(10.0)		9B → TS6B	(19.9)

<sup>a</sup>  $\Delta G_1^\ddagger$ , Gibbs free activation energy of rate-determining step. <sup>b</sup>  $\Delta G_2^\ddagger$ , Gibbs free activation energy of rate-determining step.

ESI<sup>†</sup>). Then, R<sub>2</sub> inserts into the Cu–Si bond *via* transition state TS2B to form an allylcopper intermediate 5B. The energy barrier of insertion process is 13.5 kcal mol<sup>-1</sup>. The following CO<sub>2</sub> insertion step requires the energy barrier of 17.9 kcal mol<sup>-1</sup> to generate a copper carboxylate intermediate 6B *via* a six-membered-ring transition state TS3B. Finally,  $\sigma$ -bond metathesis of 6B with R<sub>1</sub> occurs to afford boron carboxylate P<sub>0-1</sub> and regenerate 3B with a free energy barrier 9.8 kcal mol<sup>-1</sup>. As shown in Fig. 3, the rate-determining step for the reaction leading to P<sub>0-1</sub> catalyzed by Cu/L<sub>B</sub> is CO<sub>2</sub> insertion (5B → TS3B), and the overall energy barrier is 17.9 kcal mol<sup>-1</sup>.

The free energy diagrams for the reaction leading to P<sub>0-2</sub> catalyzed by Cu/L<sub>B</sub> are shown in Fig. 4. From intermediate 3B, the regioselectivity occurs. R<sub>2</sub> inserts into the Cu–Si bond of 3B towards another orientation *via* the transition state TS5B to form the vinylcopper intermediate 9B with an energy barrier of 10.0 kcal mol<sup>-1</sup>. Then, CO<sub>2</sub> inserts into the C–Cu bond of 9B through a four-membered-ring transition state TS6B to form a copper carboxylate intermediate 10B with a free energy barrier of 19.9 kcal mol<sup>-1</sup>. Another transition state TS7B for the CO<sub>2</sub> insertion step was also calculated, but the higher relative free energy (6.1 kcal mol<sup>-1</sup>) indicates that this pathway is infeasible. Finally, 10B binds up with R<sub>1</sub> to generate intermediate 11B, which is followed by the  $\sigma$ -bond metathesis (TS8B) to afford boron carboxylate P<sub>0-2</sub> and regenerate 3B with an attainable free energy barrier 8.0 kcal mol<sup>-1</sup>. As shown in Fig. 4, for the reaction leading to P<sub>0-2</sub> catalyzed by Cu/L<sub>B</sub>, the CO<sub>2</sub> insertion process is found to be the rate-determining step with the overall activation barrier of 19.9 kcal mol<sup>-1</sup> (9B → TS6B).

As shown in Fig. 3 and 4, the regioselectivity-determining steps to lead to P<sub>0-1</sub> and P<sub>0-2</sub> catalyzed by Cu/L<sub>B</sub> are 3B → TS2B (13.5 kcal mol<sup>-1</sup>) and 3B → TS5B (10.0 kcal mol<sup>-1</sup>), respectively. The relative free energy difference (3.5 kcal mol<sup>-1</sup>) between TS2B and TS5B is reasonable agreement with the experimental major product P<sub>2</sub>.

Other possible pathways were also considered. For example, in the transmetalation process of 1B and R<sub>1</sub>, the pathway that PCy<sub>3</sub>Cu moiety attacking toward B(pin) moiety, accompanied with the R<sub>1</sub> moiety attacking toward Cl moiety was also considered, but the calculated results indicated that such a proposal is inaccessible (see Fig. S2 in ESI<sup>†</sup>). The inaccessible pathways that CO<sub>2</sub> insertion prior to the alkene insertion step were also calculated (see Fig. S3 in ESI<sup>†</sup>). The impact of additive NaOAc on the regioselectivity was also calculated and ruled out (see Fig. S4 in ESI<sup>†</sup>).

In order to give a clear picture, the mechanistic scenarios for reactions A and B are summarized in Table 1. For reactions A

and B, the regioselectivity-determining and rate-determining steps together with activation energy barriers are presented. It can be seen from the table that the regioselectivity-determining step is the reaction of 3A (3B) and R<sub>2</sub> process, which is the divergence of P<sub>0-1</sub> and P<sub>0-2</sub>. In addition, we can see all the reverse reactions of 3A (3B) and R<sub>2</sub> are unachievable kinetically with too high activation energies from the preceding energy diagrams. Consequently, it is predictive that reactions of 3A (3B) and R<sub>2</sub> are irreversible and thus this divergent step determines the major product (P<sub>1</sub> or P<sub>2</sub>). In other words, the major product is determined by the relative stability of the two divergent transition states from reaction of 3A (3B) and R<sub>2</sub>.

For Cu/L<sub>A</sub>-catalyzed reaction A, generation of P<sub>1</sub> is calculated to be kinetically more favored than that of P<sub>2</sub>. The activation energy barrier difference between the two reactions is 2.1 kcal mol<sup>-1</sup> (TS3A *vs.* TS9A). The rate-determining steps are same to the regioselectivity-determining steps which predicted to be the alkene insertion for the two reactions. For Cu/L<sub>B</sub> reaction B, P<sub>2</sub> becomes the major product, and the activation energy difference between the two reactions of 3B + R<sub>2</sub> is enlarged to 3.5 kcal

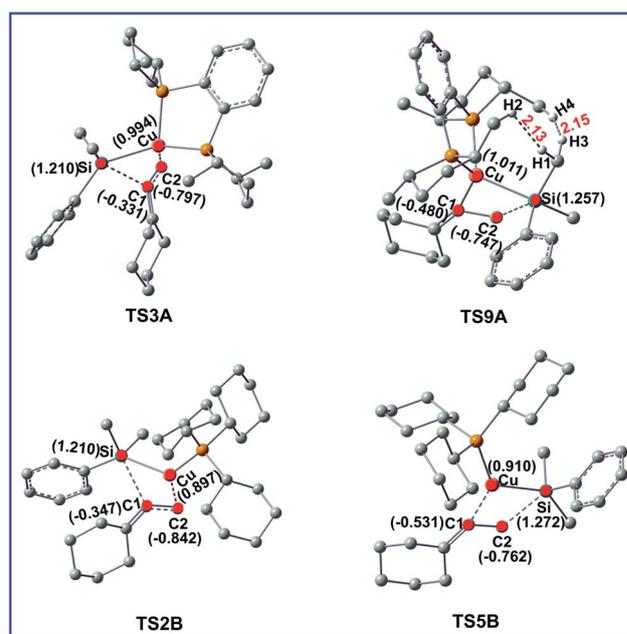


Fig. 5 Optimized structures for the transition states TS3A, TS9A, TS2B, and TS5B. Some important bond distances are given in Å. The hydrogen atoms not participating in the reaction have been omitted for clarity. The values in parentheses are the related NBO charges (in e).



mol<sup>-1</sup>. The rate-determining step is CO<sub>2</sub> insertion process for the two reactions.

### 3.3 Origin of the regioselectivity

To gain insight into the origin of the observed regioselectivity, we checked both steric and electronic effects in the selectivity-determining transition states **TS3A**, **TS9A**, **TS2B**, and **TS5B**. The important bond distances and NBO charges for the four transition states are shown in the Fig. 5.

For Cu/L<sub>A</sub>-catalyzed reaction A, the steric effects are the major factors to account for the regioselectivity. As shown in Fig. 5, the H1...H2 and H3...H4 distances in **TS9A** are 2.13 and 2.15 Å, respectively, which are less than the sums of the van der Waals radii of hydrogen (1.20 Å). Thus, the great steric effects destabilize **TS9A** and rule out the possibility to lead to **P**<sub>2</sub> product in this reaction. In contrast, there are no obvious electronic effects difference in Cu/L<sub>A</sub>-catalyzed reaction (**TS3A** and **TS9A**).

For Cu/L<sub>B</sub>-catalyzed reaction B, there are no obvious steric bias both in **TS2B** and **TS5B**. Compared the corresponding NBO charges in Fig. 5, it can be found that the Cu atom become more negatively charged when the attached ligand L<sub>A</sub> is replaced by L<sub>B</sub>. As shown in Fig. 5, more negative C2 atom combines with the less positive Cu atom in **TS2B**, while combines with the more positive atom Si in **TS5B**. Obviously, the stronger attractive interaction caused by electronic factors in **TS5B** leads to the major product **P**<sub>2</sub>.

## 4. Conclusions

We have presented a detailed and plausible mechanism for the copper-catalyzed ligand controlled regiodivergent silacarboxylation of allenes with CO<sub>2</sub> and silylborane by using DFT calculations. Two reactions are included in this study, reactions A (CuOAc as the catalyst and product **P**<sub>1</sub> is favored with the bidentate L<sub>A</sub> ligand), and B (CuCl/NaOAc as the catalyst and product **P**<sub>2</sub> is favored with the monodentate L<sub>B</sub> ligand). The rate-determining steps for the reactions to form products **P**<sub>1</sub> and **P**<sub>2</sub> catalyzed by Cu/L<sub>A</sub> both are the alkene insertion step, while rate-determining steps for the reactions catalyzed by Cu/L<sub>B</sub> both are the CO<sub>2</sub> insertion step.

The origin of the experimentally observed selectivity on different products has been analyzed based on the above mechanistic details. It is found that regioselectivity depends on the relative feasibility of alkene migratory insertion step. For Cu/L<sub>A</sub>-catalyzed system, the generation of product **P**<sub>1</sub> is preferred mainly due to the great steric effects of **TS9A** that rule out the possibility to lead to **P**<sub>2</sub>. For Cu/L<sub>B</sub>-catalyzed system, electronic effects become the major factors to produce the major product **P**<sub>2</sub>. The in-depth understanding for the divergent reaction mechanisms and distinct regiochemistry could be beneficial in designing new related reactions.

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