



Cite this: RSC Adv., 2017, 7, 23709

Received 21st March 2017

Accepted 13th April 2017

DOI: 10.1039/c7ra03304d

rsc.li/rsc-advances

Novel $[NF_2O]^+$ and $[N_3NFO]^+$ -based energetic oxidizers for solid propellants with super high specific impulse

 Yi Yu, ^a Jifeng Chen, ^a Rubo Zhang, ^b Yuchuan Li^{*a} and Siping Pang^{*a}

Novel $[NF_2O]^+$ and $[N_3NFO]^+$ -based energetic oxidizers were designed, and their structures, thermal stabilities, and energetic properties were investigated via density functional theory (DFT). The analysis of the bond dissociation energies (from 93.4 to 120.8 kcal mol⁻¹) for the screened salts suggests that they possess better thermal stabilities than the reported $[NF_2O]^+SbF_6^-$ (89.8 kcal mol⁻¹), and compound 5 was the most stable energetic salt. All the screened salts possess a positive oxygen balance ranging from 13% to 50%. Due to a positive oxygen balance, the specific impulses of the compounds 5, 11–14 (>300 s) were superior to those of ammonium perchlorate (AP) and ammonium dinitramide (ADN) when the optimized ratio of oxidizer/aluminium/PBAN (%) was 76 : 10 : 14. Considering their thermal stability and chemical reactivity, compounds 5 and 11 with super high specific impulses can be regarded as excellent candidates for novel potential solid propellants.

Introduction

Energetic oxidizers, such as ammonium perchlorate (AP) and ammonium dinitramide (ADN), are compounds that release excess of oxygen, which react with fuels and produce large amounts of hot gases for propulsion.^{1,2} With high densities and heats of formation, high specific impulse, and high thermal stabilities, energetic salts have been a research hotspot of explosive compounds, which are useful as propellants, explosives, and pyrotechnics.^{3–7} As two parts of energetic salts, the cations and anions take the role of fuels and oxidizers, respectively. The introduction of non-oxygen containing cations into the energetic salts will decrease the oxygen balance of the salts, restricting the promotion of the energetic performance of the energetic salts; thus, it is often sufficient to oxidize the carbon content only to CO and not to CO₂ to achieve near maximum performance.⁸ However, oxygen-containing cations remain elusive because once nitrated or oxidized, the ability of neutral Lewis bases to form cationic species is greatly reduced.^{9,10} Therefore, novel oxidizers with oxygen-containing cations will be helpful to improve the oxygen balance and energetic performance of the energetic salts; for example, $[NF_2O]^+$ and $[N_3NFO]^+$, which have been reported decades ago, but the research on these cations moved slowly due to their strong oxidizability and reactivity.^{11,12}

The study history of $[NF_2O]^+$ originated from the chemistry of trifluoramine oxide (NF_3O) in the 1960s,^{13–16} and it was found in

1969 that it exists in an ionic form in its Lewis acid adducts.¹⁷ In 1999, $[NF_2O]^+$ was used as an intermediate to synthesize the amazing N_5^+ cation.¹⁸ To date, only several $[NF_2O]^+$ -based salts with paired anions such as SbF_6^- , $Sb_2F_{11}^-$, AsF_6^- , BF_4^- , and $B_2F_7^-$ have been synthesized, among which $[NF_2O]^+SbF_6^-$ was the most stable salt.¹⁹ As another fluoride oxide cation, $[N_3NFO]^+$ was reported in the form of $[N_3NFO]^+SbF_6^-$ in 2007,²⁰ which exists in the form of two different configurations: *trans*- $[N_3NFO]^+$ and *cis*- $[N_3NFO]^+$. Since then, no novel $[N_3NFO]^+$ -based salts have been experimentally or theoretically reported. These two cations compose of nitrogen, oxygen, and fluorine in their structures. On the one hand, contrary to the commonly used fuel cations, oxygen in their composition will help to improve the oxygen balance. On the other hand, fluorine can also act as a strong oxidizer to further improve the oxygen balance of the energetic salts. Moreover, hydrogen is not present in the cations, which helps to improve the entire density of the energetic salts. These characteristics make them excellent candidates as strong energetic oxidizers to be used as potential propellants when combined with appropriate anions. Fig. 1 demonstrates the superior specific impulse performance of oxygen-rich $[NF_2O]^+$ and $[N_3NFO]^+$ -based oxidizers in contrast to that of traditional fuel cation-based oxidizers, based on theoretical calculations.

Our previous study indicated that the combination of AlF_4^- with a strong oxidizing N_5^+ cation may stabilize the salt,²¹ which prompted us to pair AlF_4^- with the two cations to evaluate their stabilities. Additionally, the possibility of synthesizing novel $[NF_2O]^+$ and $[N_3NFO]^+$ salts with good stability lead us to carry out further investigations. To evaluate the role of the central atom in SbF_6^- for the stability of $[NF_2O]^+[SbF_6]^-$, the Sb atom

^aSchool of Materials Science & Engineering, Beijing Institute of Technology, Beijing 100081, PR China. E-mail: liyuchuan@bit.edu.cn; pangsp@bit.edu.cn

^bSchool of Chemistry, Beijing Institute of Technology, Beijing 100081, PR China



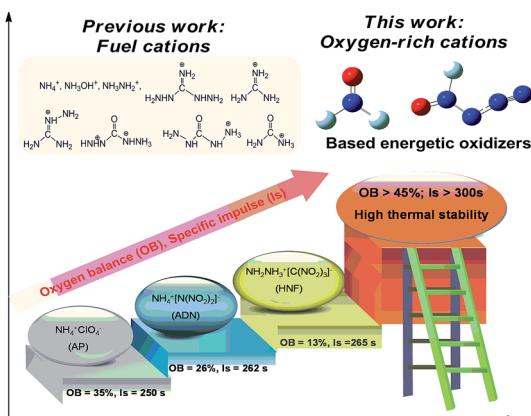


Fig. 1 The superior performance of oxygen-rich $[\text{NF}_2\text{O}]^+$ and $[\text{N}_3\text{NFO}]^+$ -based oxidizers in contrast to that of traditional fuel cation-based oxidizers.

was substituted by Al to estimate its stability. Herein, a series of novel $[\text{NF}_2\text{O}]^+$ and $[\text{N}_3\text{NFO}]^+$ salts with paired anions (ClO_4^- , NO_3^- , $\text{N}(\text{NO}_2)_2^-$, SO_3CF_3^- , SO_3NF_2^- , AlF_4^- , and AlF_6^-) were designed to study their structures, stabilities, and energetic performances. *Via* employing density functional theory, we screened out compounds with good stability and energetic performance, expanding the novel application of these cations.

Computational details

All the calculations were performed using the Gaussian 09 package. Geometry optimizations of the starting structures were carried out at the M06-2X level with the 6-311+G(d) basis set for the C, N, F, O, Cl, S, and Al atoms, and SDD for other heavy metal atoms.²² Note that the M06-2X functional is very suitable to multi-nitrogen or pure nitrogen molecules according to previous studies.²³ Each optimized structure was confirmed as a local energy minima on the potential-energy surface without imaginary frequencies.

Based on the Born–Haber energy cycle, as shown in Scheme 1, the heat of formation of a salt can be simplified by eqn (1), where ΔH_L is the lattice energy of the ionic salts, which can be predicted by the formula suggested by Jenkins *et al.*, as shown in eqn (2),²⁴ where n_M and n_X depend on the nature of the ions M_p^+ and X_q^- , respectively, and are equal to 3 for monatomic ions, 5 for linear polyatomic ions, and 6 for nonlinear polyatomic ions. The equation for the lattice potential energy U_{POT} (kJ mol^{-1}) has the form of (3), where ρ (g cm^{-3}) is the density, M

(g mol^{-1}) is the chemical formula mass of the ionic compound, and the coefficients γ ($\text{kJ mol}^{-1} \text{cm}$) and δ (kJ mol^{-1}) were obtained from the literature.²⁴

For all the related compounds, the theoretical density was obtained from the molecular weight divided by the average molecular volume, which has been successfully applied to high-nitrogen compounds. For an ionic crystal with the formula unit M_pX_q , where M denotes the cation and X denotes the anion, its volume is simply the sum of the volumes of the ions contained in the formula unit. The volume of each ion is defined as inside a contour of 0.001 electrons bohr⁻³ density that was evaluated using a Monte Carlo integration. We performed 100 single-point calculations for the optimized structure of each ion to obtain an average volume. The detonation properties of these compounds were calculated using EXPLO 5 (v6.01).

$$\begin{aligned} \Delta H_f^\circ (\text{salt}, 298 \text{ K}) = & \sum \Delta H_f^\circ (\text{anion}, 298 \text{ K}) \\ & + \sum \Delta H_f^\circ (\text{cation}, 298 \text{ K}) - \Delta H_L \end{aligned} \quad (1)$$

$$\Delta H_L = U_{\text{POT}} + [p(n_M/2 - 2) + q(n_X/2 - 2)]RT \quad (2)$$

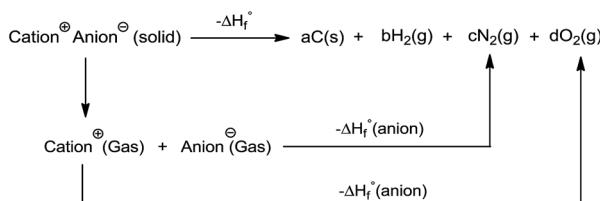
$$U_{\text{POT}} (\text{kJ mol}^{-1}) = \gamma(\rho_m/M_m)^{1/3} + \delta \quad (3)$$

Results and discussion

Structures

Our previous study showed that the M06-2X method was reliable for the optimization of polynitrogen compounds,²³ based on which we obtained the optimized structures of all the compounds, as depicted in Fig. 2. The structure of the separated $[\text{N}_3\text{NFO}]^+$ is planar, and it can exist in the form of two stereoisomers depending on whether the azido group and the fluorine ligand are *cis* (*Z* isomer) or *trans* (*E* isomer) with respect to each other. At our employed calculation level, the bond lengths of the N–F bonds in the two isomers significantly differ, with the *Z* isomer exhibiting a longer bond by 0.042 Å (reported 0.046 Å).⁸ The energy calculation also shows that the *Z* isomer is favoured by 0.78 kcal mol⁻¹, which is slightly different from 0.6 kcal mol⁻¹ reported in the literature.¹³ Thus, the stable *Z* isomer was selected for our further investigations.

There are two aspects worth noting in compound 2 (Fig. 2). One interesting case is that the three fluorine ions in AlF_6^- are obviously separated from the central Al ion. The partially dissociated Al–F bond suggests that the three F ions are not strongly bound to the central Al ion such that F ion tends to shift to $[\text{NF}_2\text{O}]^+$, promoting the decomposition of $[\text{NF}_2\text{O}]^+$. However, the shorter Al–F bond (1.7012 Å) in AlF_4^- indicates a stronger interaction between Al and F than that of AlF_6^- (1.8071 Å), which is relatively more difficult to be broken in the decomposition reaction. The bond strength of the central atom (Al) with the ligand (F⁻) in the counter-ion may be significant for the stability of $[\text{NF}_2\text{O}]^+$. A similar case exists in 9 where the two F ions are separated from the central Al ion. Another noteworthy case is the structural destruction of the anion after



Scheme 1 The Born–Haber energy cycle.



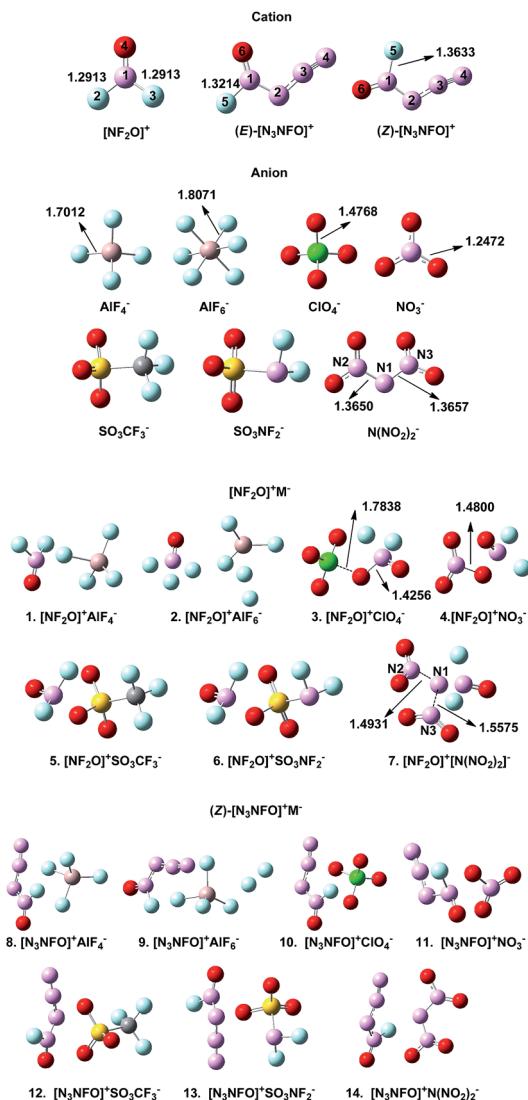


Fig. 2 The optimized structures of the separated $[\text{NF}_2\text{O}]^+$, $(\text{Z})\text{-}[\text{N}_3\text{NFO}]^+$, $(\text{E})\text{-}[\text{N}_3\text{NFO}]^+$, and $(\text{Z})\text{-}[\text{NF}_2\text{O}]^+\text{M}^-$ ($\text{M}^- = \text{ClO}_4^-$, NO_3^- , $\text{N}(\text{NO}_2)_2^-$, SO_3CF_3^- , SO_3NF_2^- , AlF_4^- , and AlF_6^-).

its combination with $[\text{NF}_2\text{O}]^+$, which may be caused by their intermolecular electrostatic interactions. For compounds 3, 4, and 7, when ClO_4^- , NO_3^- , and $\text{N}(\text{NO}_2)_2^-$ are combined with $[\text{NF}_2\text{O}]^+$, the stable structures of the anions are interrupted. The Cl–O bond noted in 3 is lengthened (1.7838 Å) when compared with the Cl–O bond (1.4768 Å) in separated ClO_4^- , resulting in Cl–O bond breakage and a newly formed N–O bond. A similar bond breakage occurs in 4 in which the N–O bond of NO_3^- is broken with much larger distance of N–O (1.4800 Å) than that of the separated NO_3^- (1.2472 Å). The structure of 7 obviously demonstrates its structural instability, where the distance of N1–N2 and N2–N3 in $\text{N}(\text{NO}_2)_2^-$ is 1.4931 Å and 1.5575 Å, respectively, completely decomposing to two separated NO_2^- molecules in the gas phase.

Based on the inspection of the optimized structures, we preliminarily screened out the possible stable compounds 1, 5, 6, 8, 10–14 to carry out further investigations.

Stabilities

The energies released upon the formation of salts by the isolated ions are reflected in the binding energy (ΔE) values, which can be expressed by the following equation.

$$\Delta E = E_0(\text{cation}) + E_0(\text{anion}) - E_0(\text{salt})$$

A larger binding energy means greater stability after the formation of the salts. For comparison, the calculated binding energy of $[\text{NF}_2\text{O}]^+\text{SbF}_6^-$ and the screened compounds are listed in Table 1. Obviously, each compound has a larger binding energy ranging from 93.4 to 120.8 kcal mol⁻¹ than that of $[\text{NF}_2\text{O}]^+\text{SbF}_6^-$ ($\Delta E = 89.8$ kcal mol⁻¹), indicating they have relatively better stabilities than $[\text{NF}_2\text{O}]^+\text{SbF}_6^-$. For compounds 1 and 8 with the same paired anion AlF_4^- , 8 gives a larger ΔE , indicating that $(\text{Z})\text{-}[\text{N}_3\text{NFO}]^+$ can better stabilize AlF_4^- -based salts when compared with $[\text{NF}_2\text{O}]^+$. In addition, our calculations show that for the same paired anions, $(\text{Z})\text{-}[\text{N}_3\text{NFO}]^+$ and $(\text{E})\text{-}[\text{N}_3\text{NFO}]^+$ -based salts have close ΔE values with a difference of 2 kcal mol⁻¹, which is due to the close free energies of separated $(\text{Z})\text{-}[\text{N}_3\text{NFO}]^+$ and $(\text{E})\text{-}[\text{N}_3\text{NFO}]^+$.²⁵ The data in Table 1 shows compound 5 has the largest binding energy ($\Delta E = 120.8$ kcal mol⁻¹), indicating that it is the most stable salt, whereas 1 is the least stable ($\Delta E = 93.4$ kcal mol⁻¹). To sum up, the thermal stability order is 5 > 6 > 11 > 12 > 8 ≈ 14 > 10 > 13 ≈ 1 > $[\text{NF}_2\text{O}]^+\text{SbF}_6^-$, based on which it can be inferred that the screened compounds have great possibilities to be synthesized.

The energy gap ($\Delta E_{\text{LUMO-HOMO}}$) of the highest occupied molecular orbital (HOMO) and the lowest unoccupied molecular orbital (LUMO) is essential for kinetic stability and chemical reactivity during the chemical processes with electron transfer or leap.^{26,27} The HOMOs and LUMOs of the selected salts are depicted in Fig. 3. For the series of $(\text{E})\text{-}[\text{N}_3\text{NFO}]^+$ salts, the distributions of orbitals and $\Delta E_{\text{LUMO-HOMO}}$ values are similar to that of the Z isomers, which are not shown. An inspection of Fig. 3 demonstrates that the LUMO orbitals mainly locate on the part of the cations, and the HOMO orbitals mainly locate on the ions. It's well known that a higher energy gap implies lower chemical reactivity and a lower gap implies higher chemical reactivity. Therefore, the $\Delta E_{\text{LUMO-HOMO}}$ values

Table 1 The binding energy (ΔE) and the energy gaps ($\Delta E_{\text{LUMO-HOMO}}$) of the screened compounds and $[\text{NF}_2\text{O}]^+\text{SbF}_6^-$

Compd.	ΔE [kcal mol ⁻¹]	$\Delta E_{\text{LUMO-HOMO}}$ [eV]
$[\text{NF}_2\text{O}]^+\text{SbF}_6^-$	89.8	7.85
1	93.4	8.17
5	120.8	11.02
6	114.1	9.92
8	96.4	7.49
10	95.6	5.12
11	99.0	8.72
12	97.6	5.14
13	93.8	5.11
14	96.3	4.77

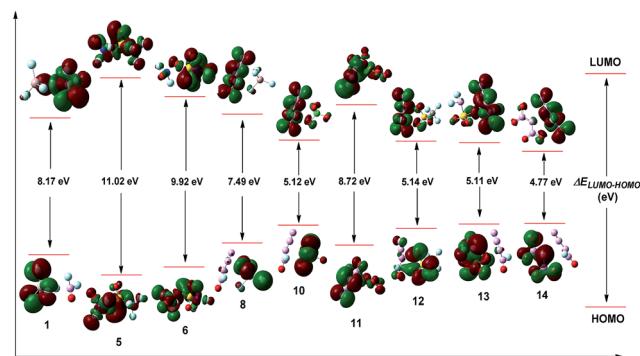


Fig. 3 The depicted HOMOs, LUMOs, and energy gaps ($\Delta E_{\text{LUMO-HOMO}}$) of the selected compounds.

in Table 1 indicate that compounds **1**, **5**, **6**, and **11** have lower chemical reactivity than $[\text{NF}_2\text{O}]^+\text{SbF}_6^-$, with $\Delta E_{\text{LUMO-HOMO}}$ values ranging from 8.17 eV to 11.02 eV, and the compound with the lowest chemical reactivity should be **5** (11.02 eV). In summary, the screened salts have a chemical reactivity order of $5 < 6 < 11 < 1 < [\text{NF}_2\text{O}]^+\text{SbF}_6^- < 8 < 12 \approx 13 \approx 10 < 14$.

Energetic properties

The heat of formation (HOF) is an important parameter to obtain the detonation properties of a compound, and the accurate heats of formation of the cations and anions of the screened compounds were calculated at the G2 level of theory. The lattice energy (ΔH_L) was mainly determined by the volume and type of the salt. For all the salts, the ratio of the number of cations to anions was 1 : 1 and the calculated ΔH_L values in Table 2 range from 507 to 569 kJ mol⁻¹. Due to the low HOF of AlF_4^- , compounds **1** and **8** with AlF_4^- as the anion show lower HOFs when compared with other compounds, whose values are -1517 and -1160 kJ mol⁻¹, respectively, among which **14** has the highest positive heat of formation up to 637 kJ mol⁻¹. Moreover, all the related salts exhibit certain detonation properties such as detonation velocities (*D*) and detonation

pressures (*P*), among which **14**, composed of $\text{N}(\text{NO}_2)_2^-$ and $[\text{N}_3\text{NFO}]^+$, shows a comparable energetic performance with AP and ADN, better than that of other compounds with NO_3^- , ClO_4^- , AlF_4^- , SO_3CF_3^- , and SO_3NF_2^- as paired anions.

The oxygen balance (OB) indicates the degree to which an explosive can be oxidized. As an important index for identifying the potential of oxidants, compounds with positive OBs indicate that there is enough oxygen in their composition to convert all the carbon and hydrogen atoms to carbon monoxide and water, respectively, whereas negative OBs indicate an insufficient oxygen content for complete oxidation. All the screened salts possess positive OBs ranging from 13% to 50%. Compounds **1** and **13** have positive OBs of 23% and 25%, respectively, which are comparable to those of AP and ADN. Particularly, the OBs of **10** (50%), **11** (47%), and **14** (45%) are much higher than those of AP and ADN (26%).

The value of excellent oxidizers can be reflected in the application as propellants. As an important parameter, the specific impulse (*I_s*) is used to determine the performance of solid propellants or binders. The *I_s* values of the relevant compounds were calculated under isobaric conditions at 7 MPa with an initial temperature of 3300 K using EXPLO5 (v6.01). Compounds **10**, **11**, and **14** show impulse values of 236 s, 264 s, and 260 s, which are much higher than those of AP (*I_s* = 157 s) and ADN (*I_s* = 202 s) as neat compounds. The impulse calculations for neat compounds **5** and **6** were not performed due to the problem of temperature limits during the calculations. The addition of aluminium and PBAN binder as a fuel drastically increases the *I_s* for all the compounds. We carried out *I_s* calculations with different combustion parameters: oxidizer/aluminium/PBAN [%], and the calculations indicate that the optimized ratio of the mixtures with aluminium and PBAN binder was composed of 76% oxidizer, 10% aluminium, and 14% binder. Surprisingly the calculations show compound **12** achieves the highest *I_s* value at 464 s, far exceeding that of any other already known solid propellant. This unexpected effect may be caused by the introduction of carbon in **12** when compared with the other salts. Moreover, compounds **5**, **11**, **13**,

Table 2 Properties of screened $[\text{NF}_2\text{O}]^+$ and (Z)- $[\text{N}_3\text{NFO}]^+$ -based oxidizers, AP, and ADN

Compd.	OB ^a [%]	ρ^b [g cm ⁻³]	ΔH_L^c [kJ mol ⁻¹]	ΔH_f^d [kJ mol ⁻¹]	V_D^e [km s ⁻¹]	P^f [GPa]	I_s^g [s]	I_s^h [s]	I_s^i [s]	I_s^j [s]
1	23	2.18	569	-1517	2.69	2.3	138	216	227	239
5	18	1.82	507	-947	4.51	12.0	—	324	263	244
6	32	1.76	532	-344	2.72	2.1	—	279	224	257
8	17	2.03	539	-1160	5.14	11.1	187	220	217	205
10	50	1.97	538	389	6.61	19.0	236	297	289	271
11	47	1.80	557	413	7.03	20.0	264	309	296	284
12	13	2.06	512	-624	5.34	13.8	191	464	124	247
13	25	2.01	519	-3	6.17	15.4	206	361	80	263
14	45	1.85	524	637	7.56	23.9	260	309	294	286
ADN ²⁸	26	1.81	—	-150	7.86	23.6	202	264	269	202
AP ²⁹	26	1.95	—	-296	6.37	15.8	157	256	261	244

^a Oxygen balance. ^b Density. ^c Lattice energy. ^d Calculated enthalpy of formation (kJ mol⁻¹). ^e Detonation velocity. ^f Detonation pressure. ^g Specific impulse(s) of neat compounds. ^h Specific impulse(s) of mixtures with aluminium and/or PBAN binder as 76 : 10 : 14. Binder composition 6% polybutadiene acrylic acid, 6% polybutadiene acrylonitrile and 2% bisphenol A ether. ⁱ Original NASA composition (70 : 16 : 14) for the space shuttle solid rocket boosters. ^j Specific impulse(s) of mixtures with aluminium and/or PBAN binder as 80 : 20 : 0.

and **14** are quite promising oxidizers, with superior I_s values (>300 s) relative to those of AP ($I_s = 256$ s) and ADN ($I_s = 264$ s) with similar formulations. Our calculations show that the I_s values of all the screened oxidizers were superior to those of AP and ADN, except for **1** and **8**. Considering the relative thermal stability and chemical reactivity discussed above, the calculation results suggest that compounds **5** and **11** may be used as a new generation of superenergetic oxidizers with potential applications as solid rocket propellants.

Conclusions

A series of $[\text{NF}_2\text{O}]^+$ and $[\text{N}_3\text{NFO}]^+$ -based energetic oxidizers were designed employing density functional theory. The calculated binding energies indicate that compounds **1**, **5**, **6**, **8**, **10–14** ($\Delta E > 93$ kcal mol $^{-1}$) have better thermal stabilities than previously reported NF_2OSbF_6 ($\Delta E = 89.8$ kcal mol $^{-1}$). The energy gap ($\Delta E_{\text{LUMO-HOMO}}$) analysis predicted that **1**, **5**, **6**, and **11** exhibit a lower chemical reactivity than NF_2OSbF_6 , in which **5** has the lowest chemical reactivity. Amazingly with good thermal stabilities, compounds **5**, **11–14** show superior I_s values (>300 s) to those of ammonium perchlorate (AP) and ammonium dinitramide (ADN) with an optimized ratio of oxidizer/aluminium/PBAN (%) of 76 : 10 : 14. While considering the kinetic stability, compounds **5** and **11** may be used as potential solid propellants. Moreover, our exploration using experiments is currently underway. We hope our work may supply new insight into the expansion of novel energetic oxidizers and candidates for future solid propellants.

Conflict of interest

The authors declare no competing financial interest.

Acknowledgements

The authors acknowledge the financial support received from the National Natural Science Foundation of China (21576026 and U1530262).

Notes and references

- X. X. Zhao, S. H. Li, Y. Wang, Y. C. Li, F. Q. Zhao and S. P. Pang, *J. Mater. Chem. A*, 2016, **4**, 5495–5504.
- Q. J. Axthammer, B. Krumm and T. M. Klapötke, *J. Org. Chem.*, 2015, **80**, 6329–6335.
- R. P. Singh, R. D. Verma, D. T. Meshri and J. n. M. Shreeve, *Angew. Chem., Int. Ed.*, 2006, **45**, 3584–3601.
- D. M. Badgujar, M. B. Talawar, S. N. Asthana and P. P. Mahulikar, *J. Hazard. Mater.*, 2008, **151**, 289–305.
- C. Darwich, T. M. Klapötke and C. M. Sabaté, *Chem.-Eur. J.*, 2008, **14**, 5756–5771.
- R. Wang, H. Xu, Y. Guo, R. Sa and J. n. M. Shreeve, *J. Am. Chem. Soc.*, 2010, **132**, 11904–11905.
- T. Fendt, N. Fischer, T. M. Klapötke and J. Stierstorfer, *Inorg. Chem.*, 2011, **50**, 1447–1458.
- C. B. Jones, R. Haiges, T. Schroer and K. O. Christe, *Angew. Chem., Int. Ed.*, 2006, **45**, 4981–4984.
- W. Liu, Q.-H. Lin, Y.-Z. Yang, X.-J. Zhang, Y.-C. Li, Z.-H. Lin and S.-P. Pang, *Chem.-Asian J.*, 2014, **9**, 479–486.
- G. W. Drake, S. Bolden, J. Dailey, M. J. McQuaid and D. Parrish, *Propellants, Explos., Pyrotech.*, 2012, **37**, 40–51.
- C. A. Wamser, W. B. Fox, B. Sukornick, J. R. Holmes, B. B. Stewart, R. Juurik, N. Vanderkooi and D. Gould, *Inorg. Chem.*, 1969, **8**, 1249–1253.
- K. O. Christe and D. A. Dixon, *J. Am. Chem. Soc.*, 1992, **114**, 2978–2985.
- V. Plato, W. D. Hartford and K. Hedberg, *J. Chem. Phys.*, 1970, **53**, 3488–3494.
- W. B. Fox, J. S. MacKenzie and R. Vitek, *Inorg. Nucl. Chem. Lett.*, 1970, **6**, 177–179.
- W. H. Kirchhoff and D. R. Lide Jr, *J. Chem. Phys.*, 1969, **51**, 467–468.
- R. R. Smardzewski, *J. Chem. Phys.*, 1974, **60**, 2193.
- S. Abramowitz and I. W. Levin, *J. Chem. Phys.*, 1969, **51**, 463–464.
- K. O. Christe, W. W. Wilson, J. A. Sheehy and J. A. Boatz, *Angew. Chem., Int. Ed.*, 1999, **38**, 2004–2009.
- K. O. Christe and W. Maya, *Inorg. Chem.*, 1969, **8**, 1253–1257.
- W. W. Wilson, R. Haiges, J. A. Boatz and K. O. Christe, *Angew. Chem., Int. Ed.*, 2007, **46**, 3023–3027.
- Y. Yu, Y.-c. Li, J.-f. Chen, C.-h. Sun, J.-s. Li, G.-j. Fan, S.-p. Pang and R.-b. Zhang, *RSC Adv.*, 2015, **5**, 104841–104845.
- M. T. Nguyen and T.-K. Ha, *Chem. Phys. Lett.*, 2000, **317**, 135–141.
- C. Qi, R.-B. Zhang and S.-P. Pang, *RSC Adv.*, 2013, **3**, 17741–17748.
- H. D. B. Jenkins, D. Tudela and L. Glasser, *Inorg. Chem.*, 2002, **41**, 2364–2367.
- T. Yu, Y.-Z. Liu, R. Haiges, K. O. Christe, W.-P. Lai and B. Wu, *RSC Adv.*, 2014, **4**, 28377–28389.
- Z. Zhou and R. G. Parr, *J. Am. Chem. Soc.*, 1989, **111**, 7371–7379.
- Y.-F. Li, X.-W. Fan, Z.-Y. Wang and X.-H. Ju, *J. Mol. Struct.: THEOCHEM*, 2009, **896**, 96–102.
- M. r. Y. Nagamachi, J. I. S. Oliveira, A. M. Kawamoto and R. d. C. s. L. Dutra, *J. Aerosp. Technol. Manage.*, 2009, **1**, 153–160.
- E. Gökcinar and T. M. Klapötke, *Turk. J. Chem.*, 2010, **34**, 953–967.

