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Naphthalimide-based fluorescent gelator for construction of both organogels and stimuli-responsive metallogels†

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A new kind of naphthalimide-based gelator **NP** containing terpyridine group was designed and characterized. NPs can form both organogels and metallogels in 2-methoxyethanol via different gelation approaches. It was also found that the **NP**/Cd(OAc)₂ metallogel showed multiple stimuli-responsive properties such as heat, Na₂S and shear stress with reversible phase changes.

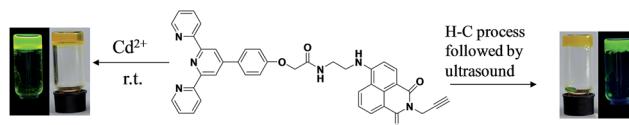
Introduction

In the past two decades, design and synthesis of low molecular weight gelators is of increasing interest due to their potential applications in catalysis, biochemistry, tissue engineering, pollutant removal and sensing materials.^{1–9} Organic compounds are found efficient gelators in organic solvents or water due to the delicate and synergistic intermolecular interactions including hydrogen bonding, hydrophobic and π – π stacking interaction. For example, Ajayaghosh *et al.* report many kinds of π organogelators with distinct optical properties;¹⁰ D. K. Smith demonstrates a series of 1,3:2,4-dibenzylidene-D-sorbitol (DBS) organogelator with varies applications such as pollutant removal and gel electrolytes;¹¹ Lu and Xue synthesize the carbazole-based π -conjugated gelator for vapor sensing.¹² Fan designs a series of stimuli-responsive dendritic organogels.¹³

Recently, using metal complexes to control molecular assembly is also a well-recognized strategy for constructing gels, in which the facile and dynamic metal-ligand coordination interaction endows the gel with unique and new properties such as enhanced catalytic performance, luminescence, as well as selective and sensitive sensing.^{14–18} For example, Bhattacharjee *et al.* synthesize a novel copper-based hydrogel for chemical fixation of CO₂;¹⁹ Zhang develops metallogel-base arrays to sense ions and anions with high sensitivity and selectivity;²⁰ Liu presents L-histidine based metallogel with dual self-healing and shrinkage properties.²¹ However, to date, it is still difficult to predict the gelation or crystallization process based on molecular structures, and the gelators that

can both form organogels and metallogels have been scarcely reported.

In previous works, we also designed many naphthalimide-based organogelator that can form functional organogels and hydrogels containing cholesterol and sugar groups respectively.^{22–24} In consideration of the synthesis of fluorescent gelator that can form both organogels and metallogels, terpyridine group was introduced in to the naphthalimide backbone, a novel gelator **NP** was designed and characterized, which contained amide group as hydrogen bonding sites, naphthalimide unit as π segment, and terpyridine group for binding with metal ions. We anticipated that the binding of terpyridine with metal ions could change the gelation pathway *via* coordination interaction and other cooperatively non-covalent interactions, for example, easy gel formation just at room temperature rather than heating–cooling process or ultrasound treatment, leading to novel functions such as multiple stimuli-responsive properties. As expected, **NP** can form opaque gel in organic solvents such as 1,4-dioxane, 2-methoxyethanol, and THF *via* the heating–cooling process followed by sonication. In the presence of Cd(OAc)₂ or Cu(OAc)₂, room temperature gelation of **NP** was observed just by staying for minutes or shaking–rest (Scheme 1). Further study showed that **NP**/Cd(OAc)₂ metallogel exhibited thixotropic properties and could be responsive to heat and Na₂S reversibly. This work provide a new strategy to the design of gelators with different gelation pathways tuned by multiple stimuli such as heat, ultrasound and ions.



Scheme 1 The chemical structures of **NP**, photos of **NP** gel and **NP**/Cd(OAc)₂ metallogel.

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Experimental

Materials

All starting materials were obtained from commercial supplies and used without further purification. 4-Vinyl pyridine, 4-bromobenzaldehyde, Cu(OAc)₂, 4-Br-1,8-naphthalic anhydride, phenyl isothiocyanate and other reagents were supplied from Shanghai Darui Fine Chemical Co. Ltd.

Techniques

IR spectra were recorded by using an IRPRESTIGE-21 spectrometer (Shimadzu). SEM images of the xerogels were carried out by using SSX-550 (Shimadzu) and FE-SEM S-4800 (Hitachi) instruments. Samples were prepared by spinning the gels on glass slides and coating them with Au. NMR spectra were performed on a Bruker Advance DRX 400 spectrometer operating at 500 and 125 MHz for ¹H NMR and ¹³C NMR spectroscopy, respectively. Fluorescence spectra were collected on an Edinburgh instrument FLS-920 spectrometer with a Xe lamp as an excitation source. The X-ray diffraction pattern (XRD) was generated by using a Bruker AXS D8 instrument (Cu target; $\lambda = 0.1542$ nm) with a power of 40 kV and 50 mA. UV-vis absorption and fluorescent spectra were recorded on a UV-vis 2550 spectroscope (Shimadzu). Sonication treatment of a sol was performed in a KQ-500DB ultrasonic cleaner (maximum power, 100 W, 40 kHz, Kunshan Ultrasound Instrument Co., Ltd., China).

Results and discussion

The synthesis and characterization of **NP** could be seen from ESI.† The gelation test was carried out in a tube by tube inversion method.²⁴ From Table S1,† by heating–cooling process, no gels were observed. Whereas, stable and yellow emissive gels could be formed when the heated solution was subjected to ultrasound at room temperature in the solvents of butanol, 1,4-dioxane, THF and 2-methoxyethanol. Due to the coordination nature of terpyridine group with metal ions, we also tested the assembly properties of **NP** in the presence of metal ions (Table S2†). The formed metallogeles in 2-methoxyethanol by heating–cooling process showed tunable colors (Fig. S1†). For example, the metallogeles of **NP** with Cu(OAc)₂, FeCl₂, and Cd(OAc)₂ exhibited green, red and yellow color respectively. The counter anions also had obvious impact on the gel colors. For example, in contrast to that of **NP**/Cu(OAc)₂ metallogele, the **NP**/CuSO₄ metallogele showed brown earth color. Such color tunable gels might be found potential application in special ion salts sensing.

Notably, **NP**/Cd(OAc)₂ metallogele (with molecular ratio of 1 : 1) can form stable and thixotropic gels just at room temperature by staying or shaking–rest method, indicating the spontaneous coordination interaction between **NP** and Cd(OAc)₂, and such coordination interaction was certified by ¹H NMR experiment (Fig. S2†). By heating the gel, it became to a sol, which reverted to a gel again after cooling. The above results suggested the tunable gelation approaches of **NP** in 2-

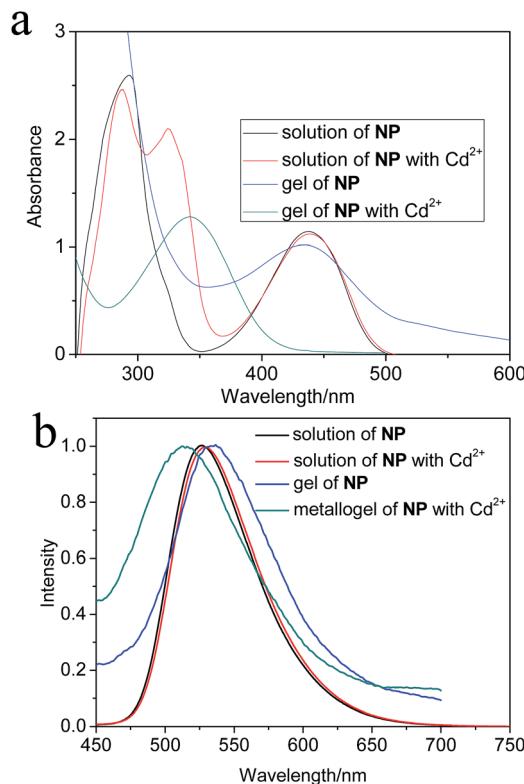


Fig. 1 (a) UV-vis spectra of **NP** solutions (10^{-4} M), gel (25 mg mL^{-1}), and **NP**/Cd(OAc)₂ metallogele (25 mg mL^{-1} , with molecular ratio of **NP** and Cd(OAc)₂ = 1 : 1); (b) corresponding fluorescence spectra of **NP** solutions, gel, and **NP**/Cd(OAc)₂ metallogele.

methoxyethanol triggered by multiple stimuli such as heating, shaking and ultrasound.

UV-vis and fluorescent studies of both solutions and gels were performed to examine the aggregation changes. From Fig. 1a, the organogel **NP** and the solutions all showed broad peak at 440 nm and 293 nm, which was attributed to the ICT process of 4-naphthalimide and pyridyl segments respectively. The peak of **NP**/Cd(OAc)₂ metallogele blue shifted from 440 nm to 342 nm in comparison to that of the diluted solution, indicating the H aggregation mode of fluorophores. In contrast to the **NP**/Cu(OAc)₂ metallogele, there was no obvious spectral changes in UV-vis spectra (Fig. S3†).

In diluted solution state, both the solution of **NP** and **NP**/Cd(OAc)₂ aggregates diluted from the metallogele showed maximum emission peak at 527 nm; while, in contrast to the emission peak of the solution, the organogel displayed 9 nm red shift, and 14 nm blue shift were observed in the metallogele (Fig. 1b). The **NP**/Cu(OAc)₂ metallogele also had the same maximum emission peak banded at 541 nm as that of **NP**/Cd(OAc)₂ metallogele. As an typical example, the temperature dependent fluorescent changes of **NP**/Cd(OAc)₂ metallogele were studied. In Fig. 2, by heating the gel from 0 to 90 °C, gel-to-sol transition happened, together with obvious fluorescent quenching with a factor of 3.3. The result revealed that π – π stacking of **NP** played an important role for the aggregates in the metallogeles. The fluorescent titrations of **NP** (10^{-5} M) with



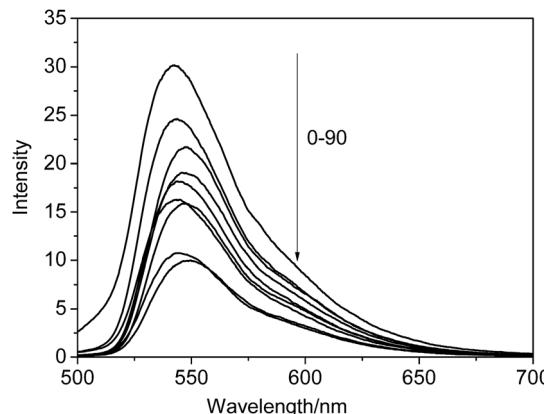


Fig. 2 Temperature dependent fluorescent changes of **NP**/ $\text{Cu}(\text{OAc})_2$ metalloge. Unit for 0–90: $^{\circ}\text{C}$.

metal ions were also studied. Addition of $\text{Cd}(\text{OAc})_2$ caused remarkable fluorescence enhancement of **NP** due to the efficient charge transfer from terpyridine group to Cd^{2+} (Fig. 3a). While, in the presence of Cu^{2+} or Fe^{2+} ion, fluorescent quenching was observed (Fig. 3b and S4†). The method of continuous variation (job plot experiments of **NP** with Cd^{2+}) in fluorescent intensity changes revealed a 1 : 1 stoichiometry of **NP** with Cd^{2+} (Fig. S5†). The corresponding UV-vis titration of **NP** with ions could be also seen from Fig. S6,† it was observed that **NP** showed no selectivity toward ions. The above results indicated that the coordination interaction between **NP** and $\text{Cd}(\text{OAc})_2$ had highly impact on both the ICT process and aggregate mode of **NP** molecules.

SEM experiments were also examined to check the macro-changes of **NP** aggregates. The **NP** gel from 2-methoxyethanol and 1,4-dioxane was comprised of densely entangled nanofibers (Fig. 4a, b and S7†). Porous structure were observed in the **NP** gel in butanol (Fig. S8†). In the presence of $\text{Cd}(\text{OAc})_2$, the as prepared gel showed sphere structure in which spheres were linked by short nanofibers (Fig. 4c and d). The **NP**/ FeCl_2 and **NP**/ CaCl_2 metallogels exhibited ribbon and ellipse structure respectively (Fig. S9 and S10†). The **NP**/ $\text{Cu}(\text{OAc})_2$ metalloge was dominated by very thin nanofibers with diameter of 20 nm (Fig. S11†). Interestingly, ion coordination induced supramolecular chirality from achiral components in gel system was observed. Seen from Fig. 5, the organogel and **NP**/ $\text{Cu}(\text{OAc})_2$ metalloge showed no CD signals. Whereas, the **NP**/ $\text{Cu}(\text{OAc})_2$

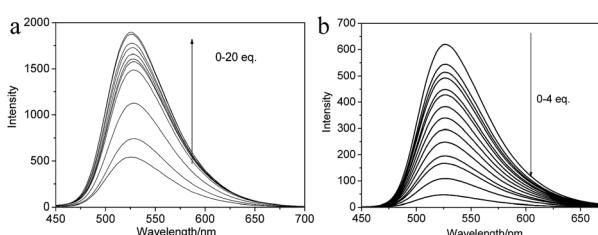


Fig. 3 Fluorescent titrations of **NP** (10^{-5} M) solution upon the addition of metal salts; (a) upon the addition of $\text{Cd}(\text{OAc})_2$; (b) upon the addition of $\text{Cu}(\text{OAc})_2$.

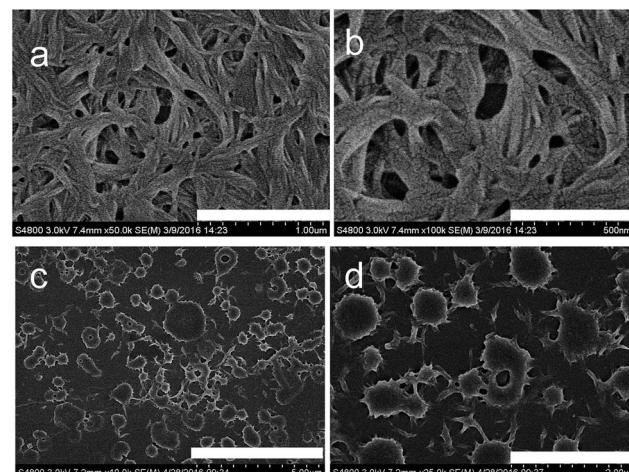


Fig. 4 SEM images of the xerogels. (a) **NP** xerogel from 2-methoxyethanol; (b) was the magnification picture of (a); (c) **NP**/ $\text{Cd}(\text{OAc})_2$ metalloge; (d) was the magnification spectra of (c). Scale bar: 1 μm ; (b) 500 nm; (c) 5 μm ; (d) 2 μm .

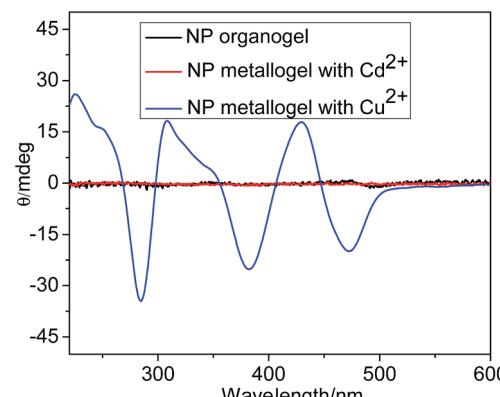


Fig. 5 CD spectra of **NP** organogel and metallogels.

metalloge of exhibited three positive peaks at 225, 307, 429 nm and three negative peaks at 284, 382, 473 nm, which might be resulted from the symmetry breaking of chiral assembly of terpyridine and naphthalimide units. XRD experiments were also studied to examine the different aggregation mode of organogel and metalloge. Seen from Fig. S12,† the peak of **NP** organogel positioned at 3.3 nm was closed to twice the length of a single molecule, indicating the dimer structure. The *d* value of 2.1 nm from the **NP**/ $\text{Cd}(\text{OAc})_2$ was rationally ascribed to the length of **NP** with $\text{Cd}(\text{OAc})_2$.

The **NP**/ $\text{Cd}(\text{OAc})_2$ metalloge could also response to both chemical and physical stimuli such as Na_2S and shaking with reversible phase changes. When coating Na_2S on the gel surface, the gel became red gradually and finally collapsed (Fig. S13†). The gel phase could be reversed by further addition of AgNO_3 . Notably, the color changes were irreversible, and the mechanism was unclear. Moreover, by shaking the gels, the gel transformed to a sol, and it could be reformed by staying for minutes, indicating the thixotropic properties of the metalloge. To further check the thixotropic properties of the **NP**/ $\text{Cd}(\text{OAc})_2$



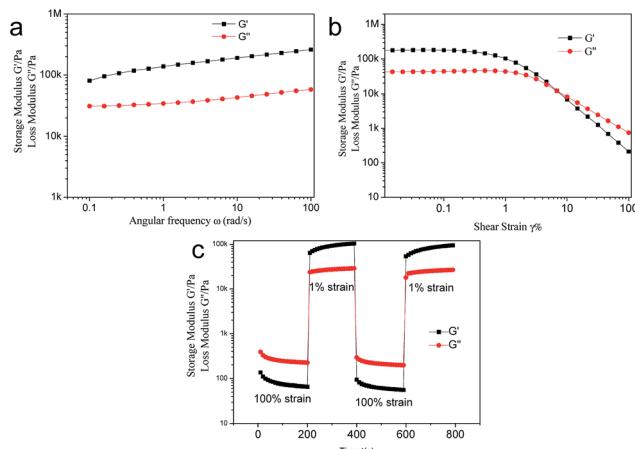


Fig. 6 (a) Frequency measurements of G' and G'' for NP/Cd(OAc)₂ metalloge with strain at 0.1%; (b) strain sweep experiment of NP/Cd(OAc)₂ metalloge with angular frequency of 10 rad s⁻¹; (c) recovery test for NP/Cd(OAc)₂ metalloge, alternating strain amplitudes of 100% and 1%.

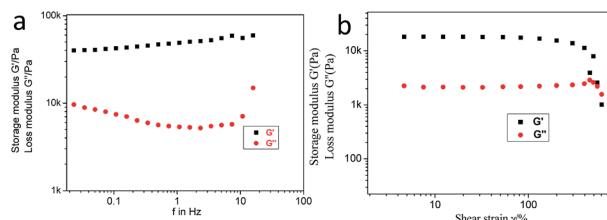


Fig. 7 (a) Frequency measurements of G' and G'' for NP organogel with strain at 0.1%; (b) strain sweep experiment of NP organogel with angular frequency of 10 rad s⁻¹.

metalloge, rheological experiments were studied. From the static frequency measurements, the storage modulus of metalloge (G') was much higher than that of the loss modulus (G''), which was in accordance with the gel state (Fig. 6a). Fig. 6b showed that the gel to sol transition point was at 6.7% strain. The recovery experiments could be shown in Fig. 6c. The gel became to a sol when exposed to 100% strain for 200 s ($G' < G''$); when the strain was decreased to 1%, the gel reformed again within 200 s. The modulus values of NP organogel was much lower than that of the NP/Cd(OAc)₂ metalloge, which revealed that the metal salt might also behave as linkers to cross-linked the assembly, leading to the higher mechanical strength (Fig. 7).

Conclusion

In conclusion, a novel gelator NP containing terpyridine and naphthalimide fluorophores was designed and characterized, which can form both organogels and metalloge in 2-methoxyethanol. Ultrasound was necessary in the organogel formation process by heating-cooling process. The coordination of NP with ion salts such as FeCl₂, Cu(OAc)₂ and CaCl₂ can also promote the gelation without ultrasound. Notably, transparent and fluorescent metalloge could be obtained by staying or

shaking-rest just at room temperature in the presence of Cd(OAc)₂. The NP/Cd(OAc)₂ metalloge exhibited multiple stimuli responsive properties such as heat, Na₂S, and shearing. This work provided a good paradigm for understanding how the gelation process was controlled by multiple stimuli. Although we and others report many naphthalimide-based gelators, the naphthalimide-based metalloge were still rare.²⁵

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Notes and references

- 1 C. D. Jones and J. W. Steed, *Chem. Soc. Rev.*, 2016, **45**, 6546–6596.
- 2 H. Yang, S. Zhang, K. Liu and Y. Fang, *RSC Adv.*, 2016, **6**, 109969–109977.
- 3 D. Mandal, S. Dinda, P. Choudhury and P. K. Das, *Langmuir*, 2016, **32**, 9780–9789.
- 4 X. H. Cao, N. Zhao, R. H. Li, H. T. Lv, Z. W. Zhang, A. P. Ao and T. Yi, *Chem.-Asian J.*, 2016, **11**, 3196–3204.
- 5 Y. Gao, F. Zhao, Q. Wang, Y. Zhang and B. Xu, *Chem. Soc. Rev.*, 2010, **39**, 3425–3433.
- 6 W. Zheng, L. J. Chen, G. Yang, B. Sun, X. Wang, B. Jiang, G. Q. Yin, L. Zhang, X. P. Li, M. H. Liu, G. S. Chen and H. B. Yang, *J. Am. Chem. Soc.*, 2016, **138**, 4927–4937.
- 7 G. Cravotto and P. Cintas, *Chem. Soc. Rev.*, 2009, **38**, 2684–2697.
- 8 Q. Pei, Q. Tang, Z. Tan, Z. Lu, L. He and B. Gong, *RSC Adv.*, 2017, **7**, 22248–22255.
- 9 X. Cao, N. Zhao, G. Zou, A. Gao, Q. Ding, G. Zeng and Y. Wu, *Soft Matter*, 2017, DOI: 10.1039/c7sm00714k.
- 10 S. S. Babu, V. K. Praveen and A. Ajayaghosh, *Chem. Rev.*, 2014, **114**, 1973–2129.
- 11 B. O. Okesola, V. M. P. Vieira, D. J. Cornwell, N. K. Whitelaw and D. K. Smith, *Soft Matter*, 2015, **11**, 4768–4787.
- 12 L. Zhai, M. Liu, P. Xue, J. Sun, P. Gong, Z. Zhang, J. Sun and R. Lu, *J. Mater. Chem. C*, 2016, **4**, 7939–7947.
- 13 Y. Feng, Y. He and Q. Fan, *Chem.-Asian J.*, 2014, **9**, 1724–1750.
- 14 A. Y.-Y. Tam and V. W.-W. Yam, *Chem. Soc. Rev.*, 2013, **42**, 1540–1567.
- 15 T. Feldner, M. Häring, S. Saha, J. Esquena, R. Banerjee and D. D. Diaz, *Chem. Mater.*, 2016, **28**, 3210–3217.
- 16 B. Jiang, L. Chen, G. Yin, Y. Wang, W. Zheng, L. Xu and H. Yang, *Chem. Commun.*, 2017, **53**, 172–175.
- 17 H. Yao, H. Wu, J. Chang, Q. Lin, T. Wei and Y. Zhang, *New J. Chem.*, 2016, **40**, 4940–4944.
- 18 S. Sarkar, S. Dutta, S. Chakrabarti, P. Bairi and T. Pal, *ACS Appl. Mater. Interfaces*, 2014, **6**, 6308–6316.

- 19 C. K. Karan, M. C. Sau and M. Bhattacharjee, *Chem. Commun.*, 2017, **53**, 1526–1529.
- 20 Q. Lin, T. Lu, X. Zhu, T. Wei, H. Li and Y. Zhang, *Chem. Sci.*, 2016, **7**, 5341–5346.
- 21 M. Liu, L. Zhang and T. Wang, *Chem. Rev.*, 2015, **115**, 7304–7397.
- 22 T. Wang, X. D. Yu, Y. Li, J. Ren and X. Zhen, *ACS Appl. Mater. Interfaces*, 2017, **9**, 13666–13675.
- 23 X. D. Yu, X. Ge, H. Lan, Y. Li, L. Geng, X. Zhen and T. Yi, *ACS Appl. Mater. Interfaces*, 2015, **7**, 24312–24327.
- 24 Z. Ma, P. Zhang, X. Yu, H. Lan, Y. Li, D. Xie, J. Li and T. Yi, *J. Mater. Chem. B*, 2015, **3**, 7366–7371.
- 25 J. I. Lovitt, C. S. Hawes, A. D. Lynes, B. Haffner, M. E. Möbiusc and T. Gunnlaugsson, *Inorg. Chem. Front.*, 2017, **4**, 296–308.

