

## PAPER

 View Article Online  
 View Journal | View Issue
Cite this: *RSC Adv.*, 2017, 7, 25066

Received 16th March 2017

Accepted 24th April 2017

DOI: 10.1039/c7ra03134c

rsc.li/rsc-advances

# Facile synthesis of mesoporous silica by CO<sub>2</sub>/N<sub>2</sub> switchable templates using a convenient compound†

Jianzhong Jiang,<sup>ID</sup>\*<sup>a</sup> Jinchao Yin,<sup>‡</sup><sup>a</sup> Li Dai,<sup>‡</sup><sup>a</sup> Rukuan Liu<sup>b</sup> and Zhihong Xiao<sup>b</sup>

A facile synthesis method for the preparation of mesoporous silica with CO<sub>2</sub>/N<sub>2</sub> switchable templates using a commercially available compound was described. Dodecyl dimethyl ammonium bicarbonate, which can be *in situ* formed by *N,N*-dimethyldodecylamine and CO<sub>2</sub>, was used as a soft template, and 98% of the templates could be easily removed.

Mesoporous silica nanoparticles (MSNs) have attracted extensive attention in recent years and have been applied in various fields such as catalysis, adsorption and storage, controlled drug release, targeted drug delivery, and simultaneous diagnosis, owing to their low density, large void space, large specific surface area, and good biocompatibility.<sup>1</sup>

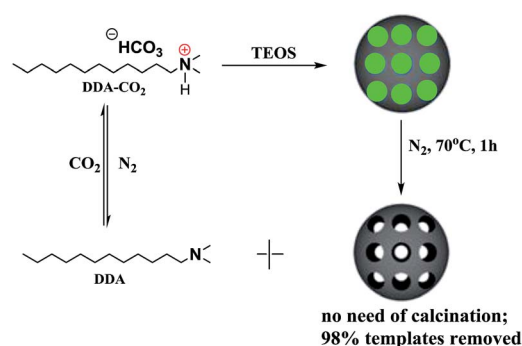
Soft templates are most commonly used in the preparation of various mesoporous silica materials. MSNs with spherical pores can always be obtained after the removal of the templates. The use of different removal methods certainly influence the characteristics of mesoporous materials. The most common method to remove the templates is calcination. The calcination process causes structure shrinkage or collapse, no recovery of the surfactants, and sacrifice of the surface hydroxyl groups.<sup>2</sup> Pinnavaia and co-workers have extensively studied alkylamines as silica templates;<sup>3</sup> however, solvent extraction was needed to remove the templates. Although solvent extraction allows the recovery and reuse of expensive templates,<sup>4</sup> a great amount of solvent is needed. Supercritical carbon dioxide extraction,<sup>5</sup> ultrasonic extraction,<sup>6</sup> microwave-assisted methods,<sup>7</sup> and oxidative treatment, such as UV/H<sub>2</sub>O<sub>2</sub> treatment<sup>8</sup> and oxidation using perchlorates,<sup>9</sup> have also been employed to remove the templates. Recently, Zhao *et al.* reported UV irradiation and solvent extraction to remove an azobenzene-derived surfactant template that was used to prepare MSNs, and 80% of the surfactant could be recycled.<sup>10</sup> However, new strategies are still needed for the facile and non-destructive preparation of MSNs.

Jessop *et al.* have developed CO<sub>2</sub>-switchable surfactants.<sup>11</sup> Compared to other triggers, such as light or pH, CO<sub>2</sub> is an

inexpensive, environmentally benign, and effective trigger for switchable surfactants. Hydrophobic long-chain alkyl amidines could be protonated and become water-soluble surfactants upon the addition of CO<sub>2</sub>. Amidinium ion could then be deprotonated by removing CO<sub>2</sub> using N<sub>2</sub>. However, most of the switchable surfactants need complex synthesis methods and are commercially unavailable.

Herein, we described a CO<sub>2</sub>/N<sub>2</sub> switchable template for the synthesis of MSNs using a commercially available compound, *N,N*-dimethyldodecylamine (abbreviated as DDA), which can be *in situ* converted to amphiphilic dodecyl dimethyl ammonium bicarbonate (abbreviated as DDA-CO<sub>2</sub>) by bubbling CO<sub>2</sub>. The templates collapsed after the preparation of the MSNs and could be easily removed from the mesopores by bubbling N<sub>2</sub> at 70 °C and washing with water/acetone. Most of the template (98%) could be removed without calcination (Scheme 1).

In an initial experiment, DDA was used as a template for the preparation of MSNs. As shown in Fig. 1a, there was no peak in the X-ray diffraction (XRD) curve, which indicated that DDA was not a suitable template for MSNs. Pinnavaia *et al.* have reported that primary amines can be used as a silica template. However,



Scheme 1 The synthesis of MSNs using a CO<sub>2</sub>/N<sub>2</sub> switchable template.

<sup>a</sup>The Key Laboratory of Food Colloids and Biotechnology, Ministry of Education, School of Chemical and Material Engineering, Jiangnan University, 1800 Lihu Road, Wuxi, Jiangsu 214122, P. R. China. E-mail: jzjiang@jiangnan.edu.cn

<sup>b</sup>Institute of Bioenergy, Hunan Academy of Forestry, Changsha 410004, China

† Electronic supplementary information (ESI) available. See DOI: 10.1039/c7ra03134c

‡ The two authors contributed equally to this work.



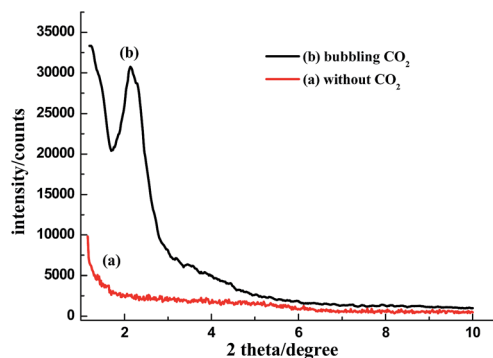


Fig. 1 The X-ray diffraction (XRD) curves obtained for the MSNs fabricated using *N,N*-dimethyldodecylamine aqueous solution: (a) without bubbling  $\text{CO}_2$  and (b) after bubbling  $\text{CO}_2$  (0.1 MPa).

contrary to primary amines, *N,N*-dimethyldodecylamine is too hydrophobic to form micelles in water.

DDA can be protonated to amphiphilic ammonium bicarbonate ( $\text{DDA-CO}_2$ , surface-active form) by bubbling  $\text{CO}_2$ , and the opaque aqueous solution of DDA quickly turns transparent. Bubbling the latter solution with  $\text{N}_2$  or air converts the surface-active form to its uncharged surface-inactive form (Scheme 1). The interconversion can be monitored by measuring the conductivity of the aqueous solution (Fig. S1†). Therefore, the *in situ* formed amphiphilic  $\text{DDA-CO}_2$  was used as a soft template for the preparation of MSNs.

As shown in Scheme 1, tetraethyl orthosilicate (TEOS) and  $\text{Na}_2\text{EDTA}$  were then added into the DDA aqueous solution in a bottle under stirring. The bottle was connected to a  $\text{CO}_2$  cylinder (0.1 MPa). The reaction was carried out under static conditions at a certain temperature for 6 days.  $\text{N}_2$  was then bubbled into the suspension at  $70^\circ\text{C}$  for an hour when the reaction was completed. The product was ultrasonically washed 6 times using distilled water and acetone (5 minutes per washing step) at room temperature.

Silica was investigated using XRD. There was one strong and narrow peak at 2 degree in the XRD curve, as shown in Fig. 1b, which indicated the disordered mesopores in the silica product. The result is similar to that previously reported in the literature.<sup>2a,12</sup> At neutral pH  $\sim 7$ , only disordered amorphous silica was obtained.

Thermogravimetric analysis (TGA) was performed to determine the template residue in the mesopores (Fig. 2). The weight loss can be divided into the following stages: the first stage is associated with the removal of physically adsorbed water (temperature below  $150^\circ\text{C}$ ), the second stage is associated with the decomposition of organic compounds ( $150$  to  $350^\circ\text{C}$ ), and the third stage is associated with the dehydroxylation of the silicon hydroxyl groups (temperature higher than  $350^\circ\text{C}$ ).<sup>13</sup> It was calculated that the template residue in the as-made silica without any treatment was 30% (Fig. 2a). Although the template residue was decreased to 22% after washing with acetone and water, the washing process did not efficiently remove the template from the mesopores (Fig. 2b). The template residue in silica was only 2% after bubbling  $\text{N}_2$  and washing (Fig. 2d), which was close to the residue value obtained after calcination ( $\sim 1\%$ , Fig. 2e).

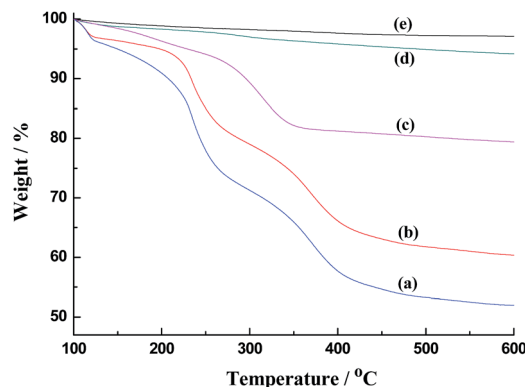


Fig. 2 Thermogravimetric analysis of the MSNs fabricated using the  $\text{CO}_2/\text{N}_2$  switchable template ( $\text{DDA-CO}_2$ ,  $1.1 \times 10^{-4}$  M). (a) The as-made sample without any treatment; (b) with washing 20 times with distilled water and acetone; (c) the as-made MSNs without  $\text{Na}_2\text{EDTA}$  and bubbling  $\text{N}_2$  to the remove template, (d) with bubbling  $\text{N}_2$  and washing with water/acetone, and (e) calcination.

The IR transmission spectra show a slight peak at  $2900\text{ cm}^{-1}$  corresponding to the stretching vibration band of  $-\text{CH}_3$  and  $-\text{CH}_2$ , which indicates the efficient removal of the organic templates by calcination and bubbling  $\text{N}_2$  (Fig. S2†). The relative intensity of the Si-OH bending bands around  $960\text{ cm}^{-1}$  was nearly the same as that observed for the as-made and bubbling  $\text{N}_2$  samples, whereas a much weaker peak intensity was observed for the calcined sample. This result indicated that calcination would inevitably reduce the level of silanol groups.

Compare with the time-consuming process of calcination or solvent extraction, the entire process of bubbling  $\text{N}_2$  and washing only required 1.5 h. There are three reasons that  $\text{DDA-CO}_2$  template could be efficiently removed just by switching off the template upon bubbling  $\text{N}_2$ . The first reason is that  $\text{DDA-CO}_2$  can be easily deprotonated into its surface-inactive compound (DDA). The hydrophilic DDA cannot form micelles as amphiphilic  $\text{DDA-CO}_2$ . Therefore, the micelles collapse after the  $\text{DDA-CO}_2$  is deprotonated by bubbling  $\text{N}_2$ .<sup>11b</sup> The second reason is that the negative charges on the wall of the mesopores in a neutral solution were much lower than those under basic conditions. The electrostatic attraction between the micelles and the wall of the mesopores is weaker under neutral conditions. The third reason is the addition of  $\text{Na}_2\text{EDTA}$  to the surfactant solution. Wong and Xiao have reported that  $\text{EDTA}^{2-}$  can co-assemble with the micelles of the cationic surfactant *via* electrostatic interactions to afford a low incorporation of the surfactant on the mesoporous,<sup>14</sup> which facilitates the removal of the template. The TGA results show that the addition of  $\text{Na}_2\text{EDTA}$  decreases the template residue from 15% to 2% (Fig. 2c and d).

The recyclability of the template was also studied. The solution after bubbling  $\text{N}_2$  and washing was obtained, evaporated, and then extracted with diethyl ether. The diethyl ether layer was evaporated and DDA was then obtained. Most of the DDA could be obtained ( $\sim 97\%$ ) and could be reused at least 10 times although there was a small weight loss during the entire process (Fig. S3†).

Scanning electron microscopy (SEM) (Fig. 3a) and transmission electron microscopy (TEM) (Fig. 3b) were also



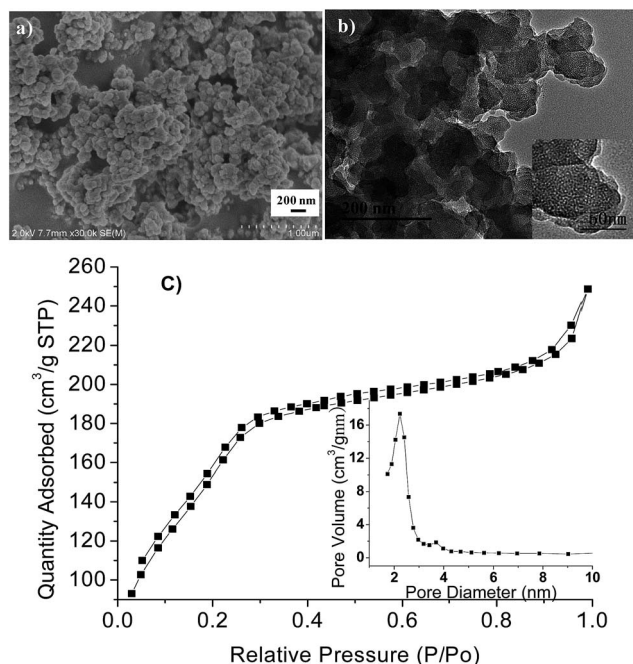


Fig. 3 (a) SEM image of the MSNs after removal of the template by bubbling  $N_2$ , (b) TEM image, and (c) the  $N_2$  adsorption isotherm and pore size distribution.

performed to observe the morphology and structure of the silica product. It was seen that the silica product comprised aggregated particles with sizes of 150 nm and had a clear mesoporous structure. Moreover, the  $N_2$  adsorption isotherm shown in Fig. 3c was a type IV isotherm according to IUPAC nomenclature.<sup>15</sup> The steep increase at a relative pressure of  $0.9 < p/p_0 < 1.0$  indicated

that the pores were formed by stacked particles. The pore size distribution curve indicated a uniform pore size of 2.3 nm. The Brunauer–Emmett–Teller (BET) surface area  $S_{BET}$  was  $565.8 \text{ m}^2 \text{ g}^{-1}$  and pore volume at  $P/P_0 = 0.99$  was  $0.4 \text{ cm}^3 \text{ g}^{-1}$ .

As the hydrocarbon chain length is an important factor that can affect the surface activity of a surfactant and the size of the micelles, tertiary amines with a carbon number of 12, 14, and 16 in the alkyl chain were used under the same conditions (Fig. 4a–c). The particle size decreased upon increasing the alkyl chain (Fig. S4†). In general, the longer the cation alkyl chain, the stronger the cation amphiphilicity. The tertiary amines with a longer alkyl chain have a higher surface activity, which can facilitate the surfactant assembly with the silica monomer, resulting in smaller particles.<sup>16,17</sup> The 2-theta of peak in the XRD pattern decreased as the alkyl chain increased (Fig. 4a–c), which indicated larger cell spacing.

As shown in Fig. 4d–f, the intensity at higher temperature was stronger, which indicated better long-range order. In addition, the morphology at higher temperature was more uniform (Fig. S5†). DDA- $CO_2$  with a much lower surface activity at lower temperature could not effectively assemble with the silica monomer; thus, it was harder to form the uniform mesoporous particles.<sup>1b</sup>

In conclusion, a new method to synthesize mesoporous silica using an inexpensive  $N_2/CO_2$  switchable surfactant as the template was introduced, in which the template could be easily removed *via* bubbling  $N_2$ .

## Acknowledgements

This work was supported by the National Natural Science Foundation of China (21473080, 20901032) and the Fundamental Research Funds for the Central Universities (JUSRP51405A).

## Notes and references

- (a) Y. H. Deng, J. Wei, Z. K. Sun and D. Y. Zhao, *Chem. Soc. Rev.*, 2013, **40**, 4254; (b) H. P. Lin and C. Y. Mou, *Acc. Chem. Res.*, 2002, **35**, 927.
- (a) D. Zhao, J. Feng, Q. Huo, N. Melosh, G. Fredrickson, B. Chmelka and G. Stucky, *Science*, 1998, **279**, 548; (b) F. Berube and S. Kaliaguine, *Microporous Mesoporous Mater.*, 2008, **115**, 469; (c) R. Grieken, G. Calleja, G. Stucky, J. Melero, R. Garcia and J. Iglesias, *Langmuir*, 2003, **19**, 3966.
- (a) Y. Mori and T. J. Pinnavaia, *Chem. Mater.*, 2001, **13**, 2173; (b) R. W. Hicks and T. J. Pinnavaia, *Chem. Mater.*, 2003, **15**, 78; (c) I. Park, Z. Wang and T. J. Pinnavaia, *Chem. Mater.*, 2005, **17**, 383.
- (a) H. Ji, Y. Fan, W. Jin, C. Chen and N. Xu, *J. Non-Cryst. Solids*, 2008, **354**, 2010; (b) C. Knofel, M. Lutecki, C. Martin, M. Mentens, V. Hornebecq and P. Llewellyn, *Microporous Mesoporous Mater.*, 2010, **128**, 26.
- (a) S. Kawi and M. Lai, *AIChE J.*, 2002, **48**, 1572; (b) Z. Huang, L. Xu, J. Li, S. Kawi and A. Goh, *Sep. Purif. Technol.*, 2011, **77**, 112.
- J. Jin, L. Cao, G. X. Su, C. Y. Xu, Z. Zhang, X. Gao, H. Liu and H. T. Liu, *Ultrason. Sonochem.*, 2014, **21**, 1688.

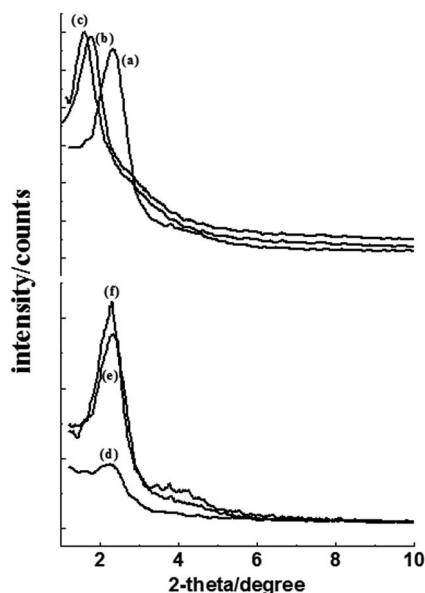


Fig. 4 (a)–(c): The XRD patterns of the silica products synthesized using tertiary amines with different carbon number in the alkyl chain (a: 12, b: 14, and c: 16). (d)–(f): The XRD patterns of the silica products synthesized using DDA- $CO_2$  at different temperatures (d: 10 °C, e: 20 °C, and f: 30 °C).



- 7 B. Tian, X. Liu, C. Yu, F. Gao, Q. Luo, S. Xie, B. Tu and D. Zhao, *Chem. Commun.*, 2002, **11**, 1186.
- 8 L. Xiao, J. Li, H. Jin and R. Xu, *Microporous Mesoporous Mater.*, 2006, **96**, 413.
- 9 H. Cai and D. Zhao, *Microporous Mesoporous Mater.*, 2009, **118**, 513.
- 10 J. Wei, Y. Liu, J. Chen, Y. Li, Q. Yue, G. Pan, Y. Yu, Y. Deng and D. Zhao, *Adv. Mater.*, 2014, **26**, 1782.
- 11 (a) Y. Liu, P. G. Jessop, M. Cunningham, C. A. Eckert and C. L. Liotta, *Science*, 2006, **313**, 958; (b) P. G. Jessop, S. M. Mercera and D. J. Heldebrant, *Energy Environ. Sci.*, 2012, **5**, 7240.
- 12 (a) E. Prouzet and T. J. Pinnavaia, *Angew. Chem., Int. Ed. Engl.*, 1997, **36**, 516; (b) C. G. Goltner and M. Antonietti, *Adv. Mater.*, 1997, **9**, 431; (c) M. Antonietti and C. Goltner, *Angew. Chem., Int. Ed. Engl.*, 1997, **36**, 910.
- 13 (a) S. Udayakumara, A. Pandurangan and P. K. Sinha, *Appl. Catal., A*, 2004, **272**, 267; (b) L. K. C. de Souza, J. J. R. Pardaul and J. R. Zamian, *Powder Technol.*, 2012, **229**, 1.
- 14 (a) R. K. Rana, V. V. Murthy, J. Yu and M. S. Wong, *Adv. Mater.*, 2005, **17**, 1145; (b) Q. Xiao, Y. Zhong, W. Zhu, T. Chen and L. Wang, *Microporous Mesoporous Mater.*, 2008, **116**, 339.
- 15 L. B. McCusker, F. Liebau and G. Engelhardt, *Pure Appl. Chem.*, 2001, **73**, 381.
- 16 K. Yano and Y. Fukushima, *J. Mater. Chem.*, 2004, **14**, 1579.
- 17 B. G. Trewyn, C. M. Whitman and V. S.-Y. Lin, *Nano Lett.*, 2004, **4**, 2139.

