

Cite this: *RSC Adv.*, 2017, 7, 26120

Ternary NiCoP nanoparticles assembled on graphene for high-performance lithium-ion batteries and supercapacitors†

Chunde Wang,^{abcd} Yinyin Qian,^{abcd} Jing Yang,^{abcd} Shiqi Xing,^{abcd} Xu Ding^{abcd} and Qing Yang^{*abcd}

Transition metal phosphides have received considerable interest for electrochemical energy storage/conversion and catalysis. In this work, we designed a unique hybrid of NiCoP nanoparticles adhered on quasi-planar structured graphene by assembling 8.5 nm ternary NiCoP nanoparticles on graphene through a solution-phase self-assembly strategy. The NiCoP catalyst in the form of small-size particles wrapped in graphene provided more active sites, a buffer for volume alteration and enhanced conductivity for electrochemical reactions. Typically, the hybrid catalyst demonstrates a high specific capacity of around 532 mA h g⁻¹, excellent cycling stability and superior rate performance when the hybrid material is evaluated as an anode material for lithium-ion batteries, and it shows excellent electrochemical properties with a specific capacitance of 646 F g⁻¹ at 4 A g⁻¹, maintaining 91% of this initial value after 2000 cycles functioning as a supercapacitor.

Received 10th March 2017

Accepted 1st May 2017

DOI: 10.1039/c7ra02910a

rsc.li/rsc-advances

Introduction

With the increasing depletion of non-renewable resources and the growing environmental pollution due to the use of fossil fuels, the global energy crisis is worsening. Developing new energy sources with high power density at low cost has attracted tremendous interest.^{1,2} Lithium ion batteries (LIBs) and supercapacitors are promising energy conversion/storage and power output devices for meeting the demands of digital communications, portable electronic devices and hybrid vehicles.^{3–7} Recently, much effort has been devoted to high-capacity materials with superior long-term stability to meet the growing demands for flexible and practical electric products.^{8–12} To the best of our knowledge, the performance of LIBs and supercapacitors is largely dependent on the structure and morphology of the electrode materials.^{13–15} As a consequence, it is of great significance to seek appropriate nanostructured materials. In the past three decades, a series of nanomaterials with different dimensionalities, such as 0D quantum dots,^{16,17}

1D nanowires,^{18,19} 2D nanosheets^{20–22} and 3D network-like structures,^{23–26} have been found and applied in electrode materials.

Transition metal phosphides (TMPs), especially binary TMPs (MP_x, where M = Cu, Fe, Co, *etc.*), have been considered as alternative anode materials for LIBs because of their exceptionally high gravimetric storage capacities (500–1800 mA h g⁻¹) compared to conventional carbon-based anode materials, their low lithiation-delithiation potentials and the low cost of the raw materials.^{27–33} However, the practical application of LIBs is often hindered on account of various critical challenges. First of all, the main challenge is the large volume alteration of TMPs during charge/discharge: the lithium ion-coupled charge transfer reactions in TMPs induce large irreversible volume changes, which can cause pulverization of the initial crystal structure and the aggregation of the active material. This process of large volume alteration further leads to poor capacity retention. The second key problem is related to the insulating nature of TMPs, including their binary compounds and multi-elemental derivatives: the intrinsically low electronic conductivity of such materials tends to limit their practical applications in most cases. Meanwhile, the same problems are encountered in the development of supercapacitors, in addition to some other energy-conversion systems, when TMPs are employed as active electrode materials. To date, to overcome the above limitations of electrode materials with low conductivity and large volume changes, several effective methods have been developed, such as introducing conductive agents (including carbon shells, carbon networks, graphene nanosheets, *etc.*) and constructing various nanostructures, and it is found that these

^aHefei National Laboratory of Physical Sciences at the Microscale, University of Science and Technology of China (USTC), Hefei 230026, Anhui, P. R. China. E-mail: qyoung@ustc.edu.cn; Fax: +86-551-63606266; Tel: +86-551-63600243

^bDepartment of Chemistry, USTC, Hefei 230026, Anhui, P. R. China

^cLaboratory of Nanomaterials for Energy Conversion, USTC, Hefei 230026, Anhui, P. R. China

^dDepartment of Physics, USTC, Hefei 230026, Anhui, P. R. China

† Electronic supplementary information (ESI) available: Materials including methods, TEM images of the NiCoP/G hybrid after the treatment, XRD and EDX of the samples, Raman spectrum of the NiCoP/G hybrid, discharge-charge voltage profiles of the NiCoP/G hybrid. See DOI: 10.1039/c7ra02910a

strategies can effectively improve the cycling performance, and especially the rate capability, of the electrode materials.^{34–38}

As previously reported, graphene is a two-dimensional single-layer sheet of carbon atoms, and is particularly attractive as a substrate material to host nano-materials based on metal compounds due to its mechanical strength, large surface area and high electronic conductivity.^{39–41} In the present work, we demonstrate that monodisperse ternary NiCoP nanoparticles can be self-assembled on graphene (termed NiCoP/G hybrid) at room temperature by a solution-phase method, and our electrode materials exhibit a high capacity and excellent cycling stability for LIBs and supercapacitors. As a proof of concept, the introduction of carbon materials into the nanocrystals gives rise to a high electrochemical performance compared with that of the pure phase crystals. We present electrochemical data demonstrating the good performance of the NiCoP/G hybrid electrode, which delivered a discharge capacity of 532 mA h g^{−1} during the 100th cycle at a charge/discharge rate of 100 mA g^{−1}, and good rate performance. Meanwhile, the NiCoP/G hybrid showed excellent electrochemical properties with a specific capacitance of 646 F g^{−1} at 4 A g^{−1} and maintained 91% of this initial value after 2000 cycles.

Results and discussion

Monodisperse NiCoP nanoparticles (NPs) were grown *via* a reflux process with some modification in high quality and high yield, and they were assembled on a graphene surface through a solution-phase self-assembly method (see the ESI† for

details). Both NiCoP NPs and NiCoP/G hybrid were characterized by transmission electron microscopy (TEM). Fig. 1a is a typical TEM image of the NiCoP NPs deposited on an amorphous carbon-coated copper grid, and they show a narrow size distribution with a diameter of 8.5 ± 0.5 nm. An SEM image of the NiCoP/G hybrid is provided in Fig. S1.† It is clear to observe the well-dispersed NiCoP NPs deposited uniformly on the surface of the graphene throughout the whole morphology. Fig. 1b and d show TEM images of the NiCoP/G hybrid at different magnifications, confirming that no NPs were detached from the graphene substrate. In the hybrid material, the NiCoP NPs selectively and densely grew on the graphene sheets (the exposed graphene is marked by dashed red circles). During synthesis, to ensure the particles adhesion on graphene, the samples of NPs were washed with acetic acid at 70 °C overnight and annealed at 100 °C under a gas mixture of Ar + 5% H₂ for 1 h to remove the surfactant around each NiCoP NP. The NiCoP NPs in the NiCoP/G hybrid did not show any morphological changes after the treatment (Fig. S2†). As seen in Fig. 1c, high-resolution TEM (HRTEM) images reveal interplanar spacings of 0.22 nm, which can be assigned to the (111) plane of the hexagonal NiCoP crystal structure, consistent with the X-ray diffraction (XRD) data of NiCoP (JCPDS 77-1716, Fig. S3†). The inset in Fig. 1c is the magnified lattice fringe of the NiCoP NPs. To further demonstrate the element distribution, high-angle annular dark-field scanning TEM energy dispersive X-ray spectroscopic (HAADF-STEM-EDX) elemental mappings are provided in Fig. 1e–i, verifying the co-existence of Ni, Co, P and C in the NiCoP/G hybrid. In addition, the elemental ratio of Ni : Co : P in the product was determined to be about 1 : 1 : 1, corresponding well with its stoichiometric ratio (Fig. S4†).

The chemical compositions and states of the NiCoP/G hybrid were investigated by X-ray photoelectron spectroscopy (XPS). The corresponding XPS survey scan shows that the sample contains the elements Ni, Co, P and C. The high resolution XPS spectra of the elements Ni, Co, P and C from the NiCoP/G hybrid are shown in Fig. 2a–d. In the Ni 2p (Fig. 2a) and Co 2p (Fig. 2b)

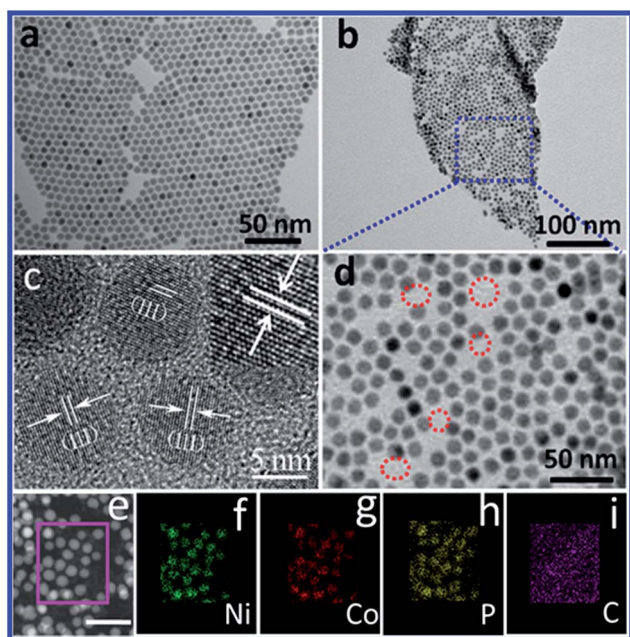


Fig. 1 (a) TEM image of the NiCoP NPs assembled on an amorphous carbon surface, (b) low-magnified, (d) high-magnified TEM images of the NiCoP/G hybrid, (c) HRTEM image of the NiCoP NPs. (e) HAADF-STEM image of the NiCoP/G hybrid, scale bars represent 20 nm. (f–i) Elemental maps of Ni (green), Co (red), P (yellow) and C (purple), respectively.

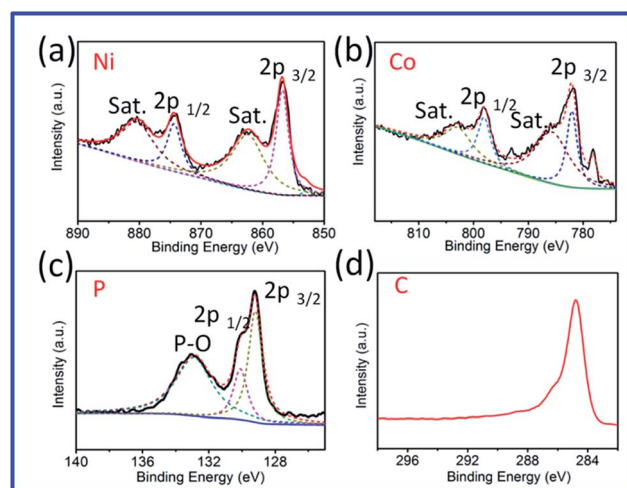


Fig. 2 XPS spectra of (a) Ni 2p, (b) Co 2p, (c) P 2p and (d) C regions for the NiCoP/G hybrid.



spectra, two pairs of spin-orbit doublets and two shake-up (satellites, shortened as sat.) excitations were obtained, indicating the existence of $\text{Ni}^{2+}/\text{Ni}^{3+}$ and $\text{Co}^{2+}/\text{Co}^{3+}$ in the samples, respectively.^{42,43} The peaks at 853.2 and 778.4 eV are close to those of zero valence Ni (852.8 eV) and Co (778.1 eV).^{43,44} It is therefore suggested that the Ni and Co species in NiCoP have a very small positive charge. Accordingly, the high-resolution spectrum of the P 2p region shows two peaks at 130.4 and 129.3 eV, reflecting the binding energy of P 2p_{1/2} and P 2p_{3/2} (Fig. 2c).^{45,46} The peak at 129.3 eV is at lower energy than that of elemental P (130.2 eV), so the P species in NiCoP has a very small negative charge. In addition, the peaks at 133.5 eV show that the P species (like the Ni and Co) is oxidized due to the exposure of the sample to acid solution and air. As shown in Fig. 2d, the peak at 284.8 eV is ascribed to C 1s. Meanwhile, the Raman spectra of the hybrid (Fig. S5†) show two intense peaks located at 1358 and 1588 cm^{-1} , assigned to the D- and G-bands of graphite, respectively, further evidencing the successful modification of graphene in the samples.

Recently, as a new alternative anode material, TMPs have attracted much attention due to their high theoretical capacity. In order to evaluate the NiCoP/G hybrid as a potential anode material for lithium ion batteries, CR2016 type coin-cells with lithium foil as a counter electrode were prepared. Fig. S6† shows the first and 100th galvanostatic charge–discharge curves of the NiCoP/G hybrid from 0.005 to 3.0 V at a current density of 100 mA g^{-1} . The initial discharge and charge specific capacities reached 1034 and 686 mA h g^{-1} , respectively, indicating irreversible losses of about 33.6%, which may be ascribed to the irreversible formation of a solid electrolyte interlayer (SEI) and the partial destruction of the internal structure. The cyclic voltammogram (CV) profiles of the NiCoP/G hybrid for the first three cycles at a scan rate of 0.1 mV s^{-1} are shown in Fig. 3a. The

first discharge for the NiCoP/G hybrid shows a peak at 1.45 V, corresponding to the Li^+ insertion process.²⁹ Another reduction peak located at 0.65 V corresponds to the decomposition of NiCoP into metallic Ni and Co and the formation of amorphous Li_3P and an SEI film. During the charging process, the anodic peaks at 1.01 and 1.42 V are related to the decomposition of the SEI film and Li_3P , and the one at 2.42 V is attributed to the partially reformed crystalline Ni_2P phase.⁴⁷ After the initial cycle, the cathodic peaks shift to near 1.80 and 0.71 V, indicating the effect of the SEI film on the Li^+ kinetic process. A good overlapping of the CV curves from the first and second cycles indicates the reversible Li-storage behavior and high coulombic efficiency.

Stable cyclic performance of the electrode is important for practical application of LIBs. Fig. 3b shows the cyclic performances of the NiCoP/G hybrid electrode at a charge–discharge rate of 100 mA g^{-1} . It can be seen that the reversible capacity of the NiCoP/G hybrid slightly decreases in the first cycle and then keeps a stable value and reaches 532 mA h g^{-1} after 100 cycles, corresponding to 78.2% capacity retention compared with the second cycle, and its coulombic efficiency (the blue dots in Fig. 3b) remains consistently at $\sim 97\%$. As is known, the major drawback that limits the practical application of TMPs is the drastic capacity fading during the cycle process.³¹ Here, the specific capacity retention has been greatly improved for the NiCoP NPs by incorporating graphene sheets. The rate performance of the electrode is also important for the application of LIBs. The tested rate performance (Fig. 3c) shows that the NiCoP/G hybrid electrode exhibits excellent cycling stability. In detail, the reversible capacity of the NiCoP/G hybrid decreases from 680 to 512, 354, 170, 120, and 74 mA h g^{-1} when the current density increases from 0.1 to 0.2, 0.5, 1, 2, and 5 A g^{-1} , respectively. After the current density decreases back to 100 mA g^{-1} , the discharge capacity of the NiCoP/G electrode recovers to 528 mA h g^{-1} . In order to further understand the reason for this outstanding performance of the NiCoP/G hybrid electrode, an AC impedance measurement was conducted. The electrochemical impedance spectra (EIS) of the as-prepared hybrid electrode tested before cycling and after 100 cycles at 0.1 A g^{-1} are presented in Nyquist plots (Fig. 3d). The plots consist of a semicircle in the high frequency region and a sloped line in the low frequency region, which can be attributed to charge transfer processes and the diffusion of lithium ions, respectively. It can be seen that the NiCoP/G hybrid exhibits good electron-transport during the cycling process due to the introduction of graphene. For comparison, the cycling performance of the NiCoP NPs was tested by investigating the charge/discharge process under the same conditions (as seen in Fig. S7a†).

To further evaluate the capacitance performance of the NiCoP/G hybrid electrode material, CV and galvanostatic charge–discharge tests of the NiCoP/G hybrid in a 2.0 M KOH solution were performed. As clearly shown in Fig. 4, a set of rate-dependent CV curves of a supercapacitor based on the NiCoP/G hybrid (Fig. 4a) were acquired at various scan rates from 2 to 20 mV s^{-1} in the potential region of 0.1–0.58 V (vs. Hg/HgO). All of the curves exhibit obvious pseudo-capacitance features with

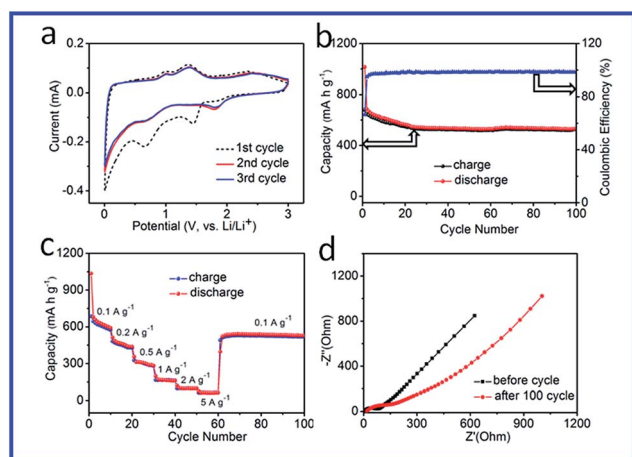


Fig. 3 (a) Cyclic voltammograms of the first three cycles of typical NiCoP/G hybrid obtained between 0.005 and 3.0 V at a scan rate of 0.1 mV s^{-1} . (b) Cycling performance of NiCoP/G hybrid between 0.005 and 3 V (vs. Li^+/Li) at 100 mA g^{-1} . (c) Rate capability of the NiCoP/G hybrid electrode at various current densities (from 0.1 A g^{-1} to 5 A g^{-1}). (d) Electrochemical impedance spectra of the NiCoP/G hybrid electrode before cycling and after 100 cycles at 100 mA g^{-1} .



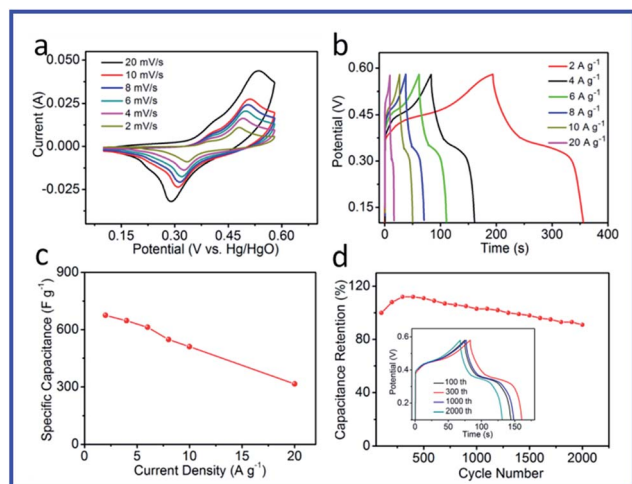


Fig. 4 Electrochemical performances of the NiCoP/G nanostructured electrode: (a) CV results measured at scan rates of 2, 4, 6, 8, 10, 20 mV s^{-1} , (b) galvanostatic charge-discharge curves, (c) specific capacitances at various current densities, (d) evaluation of specific capacitance versus the number of cycles at 4 A g^{-1} . The inset is the galvanostatic charge/discharge curve after the 100th, 300th, 1000th and 2000th cycles.

a pair of well-defined redox peaks. Moreover, the peak currents increased with the increase of the scan rate, indicating that the hybrid architectures are beneficial to fast redox reactions. The detailed electrochemical process of the NiCoP/G hybrid on nickel foam substrate at a scanning rate of 2 mV s^{-1} is illustrated by the CVs in Fig. S8†. Two pairs of redox peaks can be clearly seen, the first redox couple A1/C1 is attributed to the conversion between NiCoP and NiCoPOH and the second redox couple A2/C2 corresponds to the reversible reaction between NiCoPOH and NiCoPO,⁴⁸ represented by the following reaction:

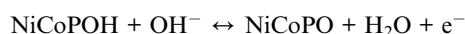


Fig. 4b shows the charge/discharge curves of the NiCoP/G hybrid in the potential range of 0.1–0.58 V (vs. Hg/HgO) at various current densities ranging from 2 to 20 A g^{-1} . The specific capacitance of the electrode can be calculated from the discharge curve according to the following equation:

$$C = \frac{I\Delta t}{m\Delta V}$$

where C (F g^{-1}) is the specific capacitance, I (A) is the charge/discharge current, Δt (s) is the discharge time, m (g) is the mass of the active material and ΔV (V) is the voltage interval of the discharge. The specific capacitances calculated from the discharge curves were 675, 646, 612, 548, 510, and 315 F g^{-1} at current densities of 2, 4, 6, 8, 10 and 20 A g^{-1} , respectively (Fig. 4c). As expected, the NiCoP/G electrode displayed better performance compared with the pure NiCoP electrode (Fig. S7b†). Such superior pseudo-capacitive performance might be attributed to the advantageous structure of the NiCoP NPs assembled on graphene. Small-size particles give rise to high

surface area, providing numerous electro-active sites for the redox reaction, while graphene ensures the conductivity of the hybrid material.

The long-term cycling stability of the supercapacitor was tested through a cyclic charge-discharge process at a current density of 4 A g^{-1} (Fig. 4d). Before activation, the NiCoP/G hybrid exhibited pseudo-capacitances of 577 F g^{-1} . After activation for 300 cycles, the discharge specific capacitance increased greatly to 646 F g^{-1} . After 2000 charge-discharge cycles, the specific capacitances for the NiCoP/G electrode decreased to 525 F g^{-1} . Hence, the specific capacitance of the NiCoP/G hybrid progressively rises to 112% with the increase of the number of cycles up to the 300th cycle. Moreover, the device still remains at 91% of the initial capacitance after 2000 charge-discharge cycles, demonstrating an excellent long-term cycling stability. These data highlight the capability of the NiCoP/G hybrid electrode to meet the requirements of practical energy storage devices.

Conclusions

In summary, a planar hybrid nanostructure of monodisperse NiCoP NPs assembled on graphene has been successfully synthesized. The well-defined NiCoP/G hybrid exhibits excellent lithium storage and supercapacitor properties, which can be attributed to their unique structure. Specifically, the hybrid structure of carbon and nanoparticles effectively improves the structural stability by buffering the volume variation during the cycling process and enhances the conductivity of the electrode material. All these merits would undoubtedly contribute to the electrochemical performance of the electrode materials, leading to high capacity, long cycle life, and superior rate performance for lithium-ion batteries and supercapacitors. This work provides a facile strategy for the synthesis of phosphide particles assembled on graphene for high performance energy storage and conversion.

Acknowledgements

We acknowledge the support from the National Natural Science Foundation of China (21571166 and 51271173), the National Basic Research Program of China (2012CB922001) and the Fundamental Research Funds for the Central Universities (WK6030000019).

Notes and references

- 1 A. J. Bard and M. A. Fox, *Acc. Chem. Res.*, 1995, **28**, 141–145.
- 2 M. W. Kanan and D. G. Nocera, *Science*, 2008, **321**, 1072–1075.
- 3 S. Chen, J. W. Zhu, X. D. Wu, Q. F. Han and X. Wang, *ACS Nano*, 2010, **4**, 2822–2830.
- 4 Z. J. Fan, J. Yan, L. J. Zhi, Q. Zhang, T. Wei, J. Feng, M. L. Zhang, W. Z. Qian and F. Wei, *Adv. Mater.*, 2010, **22**, 3723–3728.
- 5 A. S. Arico, P. Bruce, B. Scrosati, J. M. Tarascon and W. V. Schalkwijk, *Nat. Mater.*, 2005, **4**, 366–377.



- 6 S. I. Park, Y. Xiong, R. H. Kim, P. Elvikis, M. Meitl, D. H. Kim, J. Wu, J. Yoon, C. J. Yu, Z. Liu, Y. Huang, K. C. Hwang, P. Ferreira, X. Li, K. Choquette and J. A. Rogers, *Science*, 2009, **325**, 977–981.
- 7 S. Ju, J. Li, J. Liu, P. C. Chen, Y. G. Ha, F. Ishikawa, H. Chang, C. Zhou, A. Facchetti, D. B. Janes and T. J. Marks, *Nano Lett.*, 2007, **8**, 997–1004.
- 8 Y. Y. Mi, W. Liu, K. R. Yang, J. B. Jiang, Q. Fan, Z. Weng, Y. R. Zhong, Z. S. Wu, G. W. Brudvig, V. S. Batista, H. H. Zhou and H. L. Wang, *Angew. Chem., Int. Ed.*, 2016, **55**, 14818–14822.
- 9 N. Choudhary, C. Li, H. S. Chung, J. L. Moore, J. Y. Thomas and Y. Jung, *ACS Nano*, 2016, **10**, 10726–10735.
- 10 D. H. Youn, S. K. Stauffer, P. H. Xiao, H. M. Park, Y. J. Nam, A. Dolocan, G. Henkelman, A. Heller and C. B. Mullins, *ACS Nano*, 2016, **10**, 10778–10788.
- 11 M. Acerce, D. Voiry and M. Chhowalla, *Nat. Nanotechnol.*, 2015, **40**, 313–318.
- 12 C. K. Zhang, H. Q. Song, C. F. Liu, Y. G. Liu, C. P. Zhang, X. H. Nan and G. Z. Cao, *Adv. Funct. Mater.*, 2015, **25**, 3497–3504.
- 13 X. Y. Feng, Q. Shen, Y. C. Shi and J. X. Zhang, *Electrochim. Acta*, 2016, **220**, 391–397.
- 14 C. Zhou, Y. W. Zhang, Y. Y. Li and J. P. Liu, *Nano Lett.*, 2013, **13**, 2078–2085.
- 15 J. W. Hall, N. Membreno, J. Wu, H. Celio, R. A. Jones and K. J. Stevenson, *J. Am. Chem. Soc.*, 2012, **134**, 5532–5535.
- 16 S. Y. Lim, W. Shen and Z. Q. Gao, *Chem. Soc. Rev.*, 2015, **44**, 362–381.
- 17 A. J. Nozik, M. C. Beard, J. M. Luther, M. Law, R. J. Ellingson and J. C. Johnson, *Chem. Rev.*, 2010, **110**, 6873–6890.
- 18 L. F. Cui, Y. Yang, C. M. Hsu and Y. Cui, *Nano Lett.*, 2009, **9**, 3370–3374.
- 19 L. Q. Mai, F. Yang, Y. L. Zhao, X. Xu, L. Xu and Y. Z. Luo, *Nat. Commun.*, 2011, **2**, 381–385.
- 20 H. Chen, L. F. Hu, M. Chen, Y. Yan and L. M. Wu, *Adv. Funct. Mater.*, 2014, **24**, 934–942.
- 21 B. Jang, M. Park, O. B. Chae, S. J. Park, Y. Kim, S. M. Oh, Y. Z. Piao and T. J. Hyeon, *J. Am. Chem. Soc.*, 2012, **134**, 15010–15015.
- 22 R. Bhandavat, L. David and G. J. Singh, *J. Phys. Chem. Lett.*, 2012, **3**, 1523–1530.
- 23 Y. R. Zhu, Z. B. Wu, M. J. Jing, W. X. Song, H. S. Hou, X. M. Yang, Q. Y. Chen and X. B. Ji, *Electrochim. Acta*, 2014, **149**, 144–151.
- 24 C. Z. Yuan, J. Y. Li, L. R. Hou, J. D. Lin, X. G. Zhang and S. L. Xiong, *J. Mater. Chem. A*, 2013, **1**, 11145–11151.
- 25 G. J. Hu, C. Xu, Z. H. Sun, S. G. Wang, H. M. Cheng, F. Li and W. C. Ren, *Adv. Mater.*, 2016, **28**, 1603–1609.
- 26 X. F. Wang, B. Liu, Q. F. Wang, W. F. Song, X. J. Hou, D. Chen, Y. B. Cheng and G. Z. Shen, *Adv. Mater.*, 2013, **25**, 1479–1486.
- 27 J. Cabana, L. Monconduit, D. Larcher and M. R. Palacin, *Adv. Mater.*, 2010, **22**, E170–E192.
- 28 C. D. Wang, T. Ding, Y. Sun, X. L. Zhou, Y. Liu and Q. Yang, *Nanoscale*, 2015, **7**, 19241–19249.
- 29 Y. Lu, X. L. Wang, Y. J. Mai, J. Y. Xiang, H. Zhang, L. Li, C. D. Gu, J. P. Tu and S. X. Mao, *J. Phys. Chem. C*, 2012, **116**, 22217–22225.
- 30 D. Yang, J. X. Zhu, X. H. Rui, H. T. Tan, R. Cai, H. E. Hoster, D. Y. W. Yu, H. H. Hng and Q. Y. Yan, *ACS Appl. Mater. Interfaces*, 2013, **5**, 1093–1099.
- 31 D. C. S. Souza, V. Pralong, A. J. Jacobson and L. F. Nazar, *Science*, 2002, **296**, 2012–2015.
- 32 Y. Lu, J. P. Tu, Q. Q. Xiong, J. Y. Xiang, Y. J. Mai, J. Zhang, Y. Q. Qiao, X. L. Wang, C. D. Gu and S. X. Mao, *Adv. Funct. Mater.*, 2012, **22**, 3927–3935.
- 33 S. L. Liu, C. L. Ma, L. B. Ma and H. Z. Zhang, *Chem. Phys. Lett.*, 2015, **638**, 52–55.
- 34 Z. L. Schaefer, M. L. Gross, M. A. Hickner and R. E. Schaak, *Angew. Chem., Int. Ed.*, 2010, **49**, 7045–7048.
- 35 L. L. Wang, J. W. Liang, Y. C. Zhu, T. Mei, X. Zhang, Q. Yang and Y. T. Qian, *Nanoscale*, 2013, **5**, 3627–3631.
- 36 J. Chen, J. Z. Wang, A. I. Minett, Y. Liu, C. Lynam, H. K. Liu and G. G. Wallace, *Energy Environ. Sci.*, 2009, **2**, 393–396.
- 37 Z. S. Wu, A. Winter, L. Chen, Y. Sun, A. Turchanin, X. L. Feng and K. Müllen, *Adv. Mater.*, 2012, **24**, 5130–5135.
- 38 B. J. Li, H. Q. Cao, J. Shao, G. Q. Li, M. Z. Qu and G. Yin, *Inorg. Chem.*, 2011, **50**, 1628–1632.
- 39 S. J. Guo and S. H. Sun, *J. Am. Chem. Soc.*, 2012, **134**, 2492–2495.
- 40 J. J. Yoo, K. Balakrishnan, J. S. Huang, V. Meunier, B. G. Sumpter, A. Srivastava, M. Conway, A. L. M. Reddy, J. Yu, R. Vajtai and P. M. Ajayan, *Nano Lett.*, 2011, **11**, 1423–1427.
- 41 Z. Weng, Y. Su, D. W. Wang, F. Li, J. H. Du and H. M. Cheng, *Adv. Energy Mater.*, 2011, **1**, 917–922.
- 42 Z. Wang, X. Q. Cao, D. N. Liu, S. Hao, G. Du, A. M. Asiri and X. P. Sun, *Chem. Commun.*, 2016, **52**, 14438–14441.
- 43 Y. Z. Su, Q. Z. Xu, G. F. Chen, H. Cheng, N. Li and Z. Q. Liu, *Electrochim. Acta*, 2015, **174**, 1216–1224.
- 44 Y. J. Bai, H. J. Zhang, L. Liu, H. T. Xu and Y. Wang, *Chem.–Eur. J.*, 2016, **22**, 1021–1029.
- 45 A. P. Grosvenor, S. D. Wik, R. G. Cavell and A. Mar, *Inorg. Chem.*, 2005, **44**, 8988–8998.
- 46 Q. Liu, J. Q. Tian, W. Cui, P. Jiang, N. Y. Cheng, A. M. Asiri and X. P. Sun, *Angew. Chem., Int. Ed.*, 2014, **53**, 6710–6714.
- 47 Y. Lu, J. P. Tu, Q. Q. Xiong, H. Zhang, C. D. Gu, X. L. Wang and S. X. Mao, *CrystEngComm*, 2012, **14**, 8633–8641.
- 48 Y. M. Hu, M. C. Liu, Q. Q. Yang, L. B. Kong and L. Kang, *J. Energy Chem.*, 2017, **26**, 49–55.

