



Cite this: RSC Adv., 2017, 7, 23279

Solvates and polymorphs of rebamipide: preparation, characterization, and physicochemical analysis†

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This study aimed to investigate the solid forms of the peptic ulcer agent rebamipide (RBM). Four RBM solvates, including dimethyl sulfoxide solvate (RBM/DMSO), dichloromethane solvate (RBM/CH₂Cl₂), hydrate (RBM/H₂O), and hemiethanol hemihydrate, were prepared *via* solvent crystallization. Single crystals of RBM/DMSO and RBM/CH₂Cl₂ were successfully obtained at 4 °C. These solid forms were both monoclinic with the space group *P2*₁/*c*. Hirshfeld surface and 2D fingerprint plot analyses indicated that RBM/DMSO and RBM/CH₂Cl₂ showed similarities in hydrogen bonds but exhibited differences in subtle intermolecular interactions. RBM/CH₂Cl₂ transformed into a novel solid form (form 3) after solvent removal and transformed into another solid form (form 4) when heated at 220 °C. RBM/H₂O was converted to form 5 after dehydration. Forms 3–5 were RBM polymorphs. The stability results indicated that RBM/H₂O, form 3, and form 4 showed good stability at 40 °C and 75% relative humidity. Furthermore, the three solid forms displayed higher solubility than commercial RBM (form 1) in both water and phosphate buffer solution (PBS). Compared with form 1, RBM/H₂O and form 4 showed a comparable and faster dissolution rate in water and PBS, respectively.

Received 10th March 2017

Accepted 20th April 2017

DOI: 10.1039/c7ra02895d

rsc.li/rsc-advances

1. Introduction

Polymorphism is the ability of a compound to crystallize in different forms, such as polymorphs, solvates, hydrates, and co-crystals, and is a common phenomenon with important applications in academic and industrial research.^{1–3} Various solid forms of active pharmaceutical ingredients (APIs) possess significantly different physicochemical properties, including solubility, dissolution rate, stability, and bioavailability.^{4–7} Therefore, medical researchers must select the optimal solid form of APIs to maximize therapeutic benefits. In addition, different polymorphs may be granted individual patent protection, and the commercial interests of a company could be threatened by the discovery of novel polymorphs.⁸ Thus, investigating the polymorphism of APIs is crucial and has practical utility in pharmaceutical research and development.^{9,10}

Solvates, the crystal forms of a compound that differ in the species or stoichiometry of the included solvent molecules, are widely applied in the pharmaceutical industry.¹⁰ Accidental solvate formation is undesirable during drug processing because of potentially uncontrollable changes in the physicochemical properties of the drug product.¹¹ Conversely, solvates

with pharmaceutically relevant solvents may improve the physicochemical properties of drugs.¹² The developmental capacity of APIs have been improved with solvates, such as darunavir ethanolate,¹³ indinavir sulfate ethanolate,¹⁴ acetone solvate of cabazitaxel,¹⁵ canagliflozin hydrate,¹⁶ and dimethyl sulfoxide (DMSO) solvate of trametinib.¹⁷ Moreover, novel modifications that may be unattainable *via* ordinary crystallization techniques can be formed by the desolvation (dehydration) of solvates (hydrates).^{18,19} Thus, investigating the solvates of APIs and corresponding solid forms after desolvation is of great importance.

Rebamipide (RBM) is used to treat gastric ulcers, acute gastritis, exacerbated chronic gastritis, and dry eye.^{20–22} Two polymorphs (forms 1 and 2),²³ one monohydrate,²⁴ two co-crystals,²⁵ one solvate (methanol solvate), and several salts²⁶ of RBM have been reported. In our previous study, we compared the amorphous forms of RBM that were obtained *via* milling and spray drying.²⁷ The existence of RBM polymorphs and solvates, however, has not been systematically investigated.

Thus, this study aimed to conduct a detailed investigation on the solid forms of RBM, including solvates and polymorphs. Four RBM solvates, namely, dimethyl sulfoxide solvate (RBM/DMSO), dichloromethane solvate (RBM/CH₂Cl₂), hydrate (RBM/H₂O), and hemiethanol hemihydrate (RBM/EtOH/H₂O), were prepared through solvent crystallization. After the single crystals of RBM/DMSO and RBM/CH₂Cl₂ were successfully obtained, their structures were determined *via* single-crystal X-ray diffraction (SXRD). Hirshfeld surface (HS) and 2D fingerprint

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† Electronic supplementary information (ESI) available. CCDC 1500324 and 1500325. For ESI and crystallographic data in CIF or other electronic format see DOI: 10.1039/c7ra02895d



plot (FP) were applied to analyze the intermolecular interactions of the two solvates. Although RBM/H₂O has been previously reported in a Chinese patent,²⁴ more detailed investigations were performed in the present study. In addition, the phase transitions of the solvates upon heating were investigated, and three novel polymorphs of RBM (forms 3–5) were obtained in this process. The stability of the solvates and polymorphs was explored. Based on the results, we further investigated the solubility and dissolution rate of RBM/H₂O, form 3, and form 4, the stable solid forms of RBM.

2. Experimental

2.1. Materials

RBM (99% pure) was purchased from YiTai Technology Co., Ltd. (Shanghai, China) and was consistent with form 1 as reported by Jeon and Sohn (Fig. S1†).²³ Ultrapure water (18 MΩ resistivity from a Millipore system) was used throughout the experiment. DMSO, CH₂Cl₂, methanol (MeOH), ethanol (EtOH), and other solvents were of analytical grade and were purchased from Kelong Co., Ltd. (Chengdu, China).

2.2. Preparation of RBM solvates and polymorphs

2.2.1. RBM/DMSO. Colorless single crystals of RBM/DMSO suitable for SXRD were obtained by dissolving 1 g of form 1 in 15 mL of DMSO in a sealed beaker flask. The solution was left to crystallize at 4 °C.

2.2.2. RBM/CH₂Cl₂. RBM/CH₂Cl₂ was obtained by dissolving 200 mg of form 1 in 150 mL of MeOH/CH₂Cl₂ (2 : 1 v/v) solvent mixture *via* ultrasonication. The resulting solution was then slowly evaporated at 4 °C. Finally, colorless single crystals suitable for SXRD were obtained.

2.2.3. RBM/H₂O. RBM/H₂O was also obtained in the binary system of MeOH and CH₂Cl₂. Form 1 (200 mg) was completely dissolved in 150 mL of MeOH/CH₂Cl₂ (3 : 1 v/v) with heating and stirring at 55 °C. Then, the solution was cooled to room temperature and slowly evaporated under ambient conditions. White globular solids were obtained after approximately 10 days. However, we failed to obtain single crystals suitable for SXRD.

2.2.4. RBM/EtOH/H₂O. Form 1 (200 mg) was dissolved in EtOH (200 mL) by heating and stirring. The resulting solution was filtered through Whatman filter paper and then evaporated under ambient conditions. Colorless crystals of RBM/EtOH/H₂O were obtained after approximately one month. Single crystals suitable for SXRD were still not obtained even after a series of crystallization experiments.

2.2.5. Polymorphs. Three polymorphs were obtained *via* the phase transitions of the solvates upon heating. Form 3 was generated by heating RBM/CH₂Cl₂ at 150 °C. Form 4 was obtained when RBM/CH₂Cl₂ was further heated at 220 °C. Form 5 was obtained by dehydrating RBM/H₂O at 150 °C.

2.3. Analytical methods

2.3.1. SXRD. Single-crystal data were collected using the Oxford Diffraction Xcalibur Nova system with Mo K α radiation ($\lambda = 0.71073$ Å). A suitable crystal was selected and held on an

Xcalibur Eos diffractometer. Cell refinement and data reduction were applied using Olex2,²⁸ and the structure was solved and refined using SHELX-97.²⁹ Hydrogen bond geometries and packing diagrams were determined with mercury.³⁰

2.3.2. HS analysis and 2D FP. HS analysis provides unique information on intermolecular interactions in crystal structures based on electron distribution, which is calculated as the sum of spherical atom electron densities.^{31,32} The normalized contact distance (d_{norm}) is given by the following equation:

$$d_{\text{norm}} = \frac{d_i - r_i^{\text{vdW}}}{r_i^{\text{vdW}}} + \frac{d_e - r_e^{\text{vdW}}}{r_e^{\text{vdW}}}, \quad (1)$$

where d_i and d_e represent the distances from the nearest atom interior and exterior to the surface, respectively, and r^{vdW} is the van der Waals radii of the atoms.^{33,34} The 2D FP presents a rapid quantitative summary in the percentage of each contact, which is constructed based on the d_e and d_i distances.³² All calculations were performed using CrystalExplorer.³⁵ HS diagrams were generated using standard (high) surface resolution. Solvent molecules were hidden.

2.3.3. Morphology. RBM/DMSO, RBM/CH₂Cl₂, and RBM/EtOH/H₂O crystals were placed in a transparent glass holder and recorded by an optical microscope (CX21, Olympus, Tokyo, Japan) at 40 \times magnification. Photomicrographs were captured by a digital camera (DSC-HX-50, Sony, Tokyo, Japan). Given its small crystal size, RBM/H₂O was observed using scanning electron microscopy (SEM). The SEM micrograph was obtained using a scanning electron microscope (JSM-7500F, JEOL, Tokyo, Japan) at 15.0 kV. An electrically conductive sample was obtained after coating a sample with a thin layer of gold in vacuum prior to examination.

2.3.4. X-ray powder diffraction (XRPD). XRPD data were collected at room temperature using a diffractometer (X'Pert PRO, PANalytical, Holland) with a PIXcel 1D detector and Cu K α radiation ($\lambda = 1.5406$ Å, generator settings: 40 kV and 40 mA). Samples were loaded on a rectangular glass sample holder. Diffraction data were collected in the 2 θ range of 4–50° with a step size of 0.01313° and a counting time of 30 ms per step.

2.3.5. Differential scanning calorimetry (DSC) and thermogravimetric analysis (TGA). Thermal transitions were determined with a differential scanning calorimeter (Q2000, TA, New Castle, DE). Samples (3–5 mg) were heated from 30 °C to 320 °C at a rate of 10 °C min⁻¹ under N₂ purge (20 mL min⁻¹) in sealed aluminum pans. The instrument was calibrated for temperature and heat flow using indium as the standard.

TGA was performed using an analyzer (TG209F1 Iris, NETZSCH, Germany). Samples (5–8 mg) were determined in aluminum crucibles by heating from 30 °C to 500 °C at a rate of 10 °C min⁻¹ under N₂ purging (60 mL min⁻¹).

2.4. Accelerated stability tests

The stability of RBM solid forms was investigated at 40 °C and 75% relative humidity (RH) International Conference on Harmonization conditions for three months.³⁶ In this study, 75% RH was achieved with saturated sodium chloride solution in a desiccator.³⁷



2.5. Solubility and dissolution studies

Solubility studies were conducted by adding excess amounts of RBM solid forms to 10 mL of triple distilled water and phosphate buffer solution (PBS, pH 6.8) in Erlenmeyer flasks.

The suspensions were stirred for 72 h at constant temperature (37 °C) and then were centrifuged at 10 000 rpm for 20 min using a high-speed refrigerated centrifuge (TGL-16M, Xiangyi, China). The supernatant was filtered through a 0.45 µm Millipore filter, appropriately diluted, and analyzed using a UV-vis spectrophotometer (TU-1901, Beijing, China) at 327 nm. Solid residues obtained by filtration were analyzed through XRPD. All solubility measurements were performed in triplicate.

Dissolution studies were conducted using the paddle method on a ZRC-8D dissolution tester (Chuangxing, Tianjin, China) at 100 rpm and 37 ± 0.5 °C. Samples (30 mg) were added to 900 mL of dissolution media (triple distilled water and pH 6.8 PBS). Aliquots (5 mL) of the sample solutions were withdrawn at 2, 5, 7, 10, 15, 20, 30, 45, 60, 90, and 120 min, and then an equivalent amount of fresh media were added. The sample solutions were

filtered through a 0.45 µm filter and analyzed with a UV-vis spectrophotometer (327 nm). Each sample was performed in triplicate.

3. Results and discussion

3.1. Characterization of RBM solvates

3.1.1. Crystal structure analyses. Unambiguous structural characterization is necessary for drug development and to elucidate structure–property relationships.²⁶ SXRD was used to determine the structures of RBM/DMSO and RBM/CH₂Cl₂. The crystallographic data and refinement details of these forms are listed in Table 1. The hydrogen bonds present in the crystal structures are listed in Table 2.

RBM/DMSO was solved in space group *P*2₁/*c*. The stoichiometric ratio of the RBM and DMSO molecules was 1 : 2. RBM molecules were connected to each other through R₂²(8) dimer motifs (O3–H3···O1 and N1–H1···O2) and weak hydrogen bonds of C–H···Cl (3.734 Å, 145°) [Fig. 1(a)].³⁸ One DMSO molecule was connected to RBM molecules through N2–H2···O6 hydrogen bonds and three weak auxiliary C–H···O6 interactions, while another DMSO molecule contacted with RBM molecules through C20–H20B···Cl1 weak interactions [Fig. 1(b)]. RBM molecules were arranged in layers that stretched in a “W” shape along the *ab* plane. DMSO molecules were located between the layers of RBM molecules [Fig. 1(c)]. In addition, π···π stacking was observed in the double-layer part of the “W”.

RBM/CH₂Cl₂ was solved in space group *P*2₁/*c*. The stoichiometric ratio of the RBM and CH₂Cl₂ molecules was 1 : 1. RBM molecules were also connected to each other through R₂²(8) dimer motifs (O4–H4···O1 and N1–H1···O3) [Fig. 2(a)].³⁸ Each dimer was connected to neighboring molecules through N2–H2···O1 hydrogen bonds and weak C2–H2A···O4 hydrogen bonds (3.355 Å, 159°), thus resulting in a sheet-like structure. The sheets were connected to each other through weak C–H···Cl (3.629 Å, 138°) hydrogen bonds, which were stronger than those in RBM/DMSO (Fig. S2†). CH₂Cl₂ molecules interacted with RBM molecules through the weak hydrogen bonds of C20–H20B···O2 (3.146 Å, 176°) [Fig. 2(b)]. In 2D packing, the molecules were arranged in a corrugated sheet when viewed down the *c*-axis [Fig. 2(c)]. π···π stacking was observed in the circular part of RBM/CH₂Cl₂ [Fig. 2(d)].

3.1.2. HS analysis and 2D FP. The HSs for RBM/DMSO and RBM/CH₂Cl₂ were mapped over *d*_{norm}, shape index, and curviness. The large, deep red, circular depressions in the maps of both RBM/DMSO (Fig. S3†) and RBM/CH₂Cl₂ (Fig. S4†)

Table 1 Crystallographic parameters of RBM forms

Parameters	RBM/DMSO	RBM/CH ₂ Cl ₂
Empirical formula	C ₂₃ H ₂₇ ClN ₂ O ₆ S ₂	C ₂₀ H ₁₇ Cl ₃ N ₂ O ₄
Formula weight	527.03	455.71
Temperature (K)	148.00(10)	142.95(10)
Crystal system	Monoclinic	Monoclinic
Space group	<i>P</i> 2 ₁ / <i>c</i>	<i>P</i> 2 ₁ / <i>c</i>
<i>a</i> (Å)	8.9468(9)	14.3285(16)
<i>b</i> (Å)	28.971(3)	9.7542(4)
<i>c</i> (Å)	9.8022(11)	21.475(2)
α (°)	90.00	90.00
β (°)	97.029(10)	136.08(2)
γ (°)	90.00	90.00
<i>V</i> (Å ³)	2521.6(5)	2081.9(3)
<i>Z</i>	4	4
<i>ρ</i> _{calc} (g cm ⁻³)	1.388	1.454
<i>μ</i> (mm ⁻¹)	0.358	0.470
Crystal size (mm ³)	1104.0	936.0
<i>F</i> (000)	0.4 × 0.3 × 0.25	0.4 × 0.35 × 0.3
2θ for data collection (°)	5.94 to 52.74	5.86 to 52.74
Index ranges	-11 ≤ <i>h</i> ≤ 11, -36 ≤ <i>k</i> ≤ 36, -12 ≤ <i>l</i> ≤ 12	-17 ≤ <i>h</i> ≤ 12, -12 ≤ <i>k</i> ≤ 11, -18 ≤ <i>l</i> ≤ 26
Reflns collected	9808	8491
Unique reflns	9808	8491
Goodness-of-fit on <i>F</i> ²	1.049	0.912
<i>R</i> ₁ (<i>I</i> ≥ 2(<i>I</i>))	0.0772	0.0590
<i>wR</i> ₂ (all)	0.2210	0.1755

Table 2 Hydrogen bond metrics in crystal structure

Crystal form	Interactions	H···A (Å)	D···A (Å)	<DHA (°)	Symmetry code
RBM/DMSO	O3–H3···O1	1.685	2.505	164.92	<i>x, y, z</i> + 1
	N1–H1···O2	2.005	2.869	166.83	<i>x, y, z</i> – 1
	N2–H2···O6	2.097	2.890	149.64	
RBM/CH ₂ Cl ₂	O4–H4···O1	1.602	2.505	172.14	<i>x, y</i> – 1, <i>z</i>
	N1–H1···O3	1.940	2.821	173.33	<i>x, y</i> + 1, <i>z</i>
	N2–H2···O1	2.352	3.073	139.20	– <i>x</i> + 1, <i>y</i> – 1/2, – <i>z</i> + 3/2



corresponded to strong H-bonding interactions ($\text{H}\cdots\text{O}/\text{O}\cdots\text{H}$), whereas the small spots were due to $\text{C}\cdots\text{H}$, $\text{H}\cdots\text{H}$ and $\text{Cl}\cdots\text{H}$ contacts. The bright red hollows that represent concave regions in the shape index map were due to the $\pi\cdots\pi$ interactions of the aromatic rings; moreover, $\pi\cdots\pi$ stacking was evident in the visible large flat area on the curvedness surface.³³ These results were consistent with the findings obtained in the crystal structure analyses. 2D FPs are useful in comparing the different crystal structures for a given molecule because of their uniqueness for a particular crystal form.³² The FPs of RBM/DMSO and RBM/CH₂Cl₂ displayed significant differences (Fig. 3). The two sharp spikes near $d_{\text{e}} + d_{\text{i}} \approx 1.5 \text{ \AA}$ in the FPs of RBM/DMSO and RBM/CH₂Cl₂ were due to the strong $\text{O}\cdots\text{H}$ interactions that resulted from N-H \cdots O hydrogen bonds. The contributions of these interactions to the total HS were both 28.0% in the two solvates because of similar hydrogen bonds. The moderate spikes with chicken wing-like features near $d_{\text{e}} + d_{\text{i}} \approx 2.8 \text{ \AA}$ in both solvates were due to $\text{Cl}\cdots\text{H}$ interactions. The contributions of these interactions to HS was higher in RBM/CH₂Cl₂ (21.5%) than in RBM/DMSO (9.1%) because of the stronger C-H \cdots Cl interactions and the CH₂Cl₂ molecules in the crystal structure of RBM/CH₂Cl₂. Other common and dominant interactions were C \cdots H contacts with HS contributions of 19.2% in RBM/DMSO and 18.5% in RBM/CH₂Cl₂. H \cdots H interactions contributed more in RBM/DMSO (32.0%) than in RBM/CH₂Cl₂ (21.2%). The triangular region at the center of the FP of both solvates was ascribed to C \cdots C contacts (2.7%) (Fig. S5†). These contacts are a typical and characteristic indicator of significant $\pi\cdots\pi$ interactions and the π -system is represented by the C atoms of the aromatic rings.³³

3.1.3. Morphological analysis. RBM/DMSO exhibited a transparent cuboid block structure with a smooth surface [Fig. 4(a)]. RBM/CH₂Cl₂ presented a hexagonal plate structure

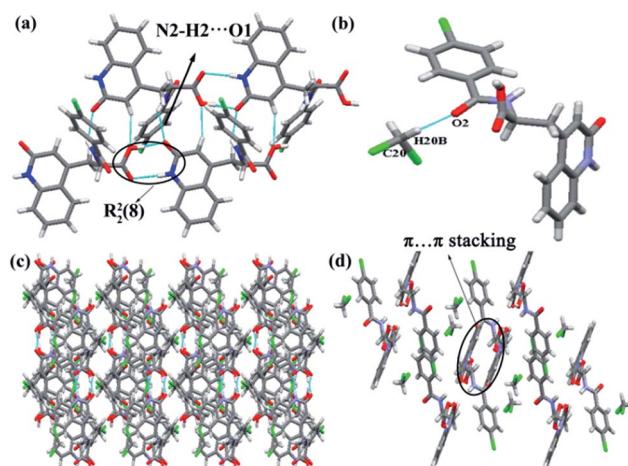


Fig. 2 (a) Hydrogen bond network of RBM/CH₂Cl₂ sustained by R₂(8) dimer motifs, N2-H2 \cdots O1 hydrogen bonds, and the weak hydrogen bonds of C2-H2A \cdots O4. (b) Interactions between CH₂Cl₂ and RBM molecules. (c) Corrugated sheet viewed down the c-axis. (d) $\pi\cdots\pi$ stacking in RBM/CH₂Cl₂.

that immediately lost transparency in air [Fig. 4(b)]. Meanwhile, RBM/EtOH/H₂O exhibited a transparent needle-like structure [Fig. 4(d)]. Although RBM/H₂O also displayed a needle-like structure, the needles were considerably smaller and were integrated as white spherules [Fig. 4(c)].

3.1.4. XRPD analysis. The experimental XRPD patterns of RBM/DMSO and RBM/CH₂Cl₂ matched well with the calculated patterns (Fig. S6†). The RBM solvates were readily distinguished from form 1 by the XRPD patterns and can be identified by the diagnostic peaks. The diagnostic peaks were at 6.10°, 18.22°, and 19.88° for RBM/DMSO; 9.32°, 16.53°, and 26.79° for RBM/CH₂Cl₂; 11.17°, 14.57°, and 25.58° for RBM/H₂O; and 7.98° and 18.91° for RBM/EtOH/H₂O [Fig. 5(a) and Tables S1–S4†]. The XRPD pattern of RBM/H₂O was consistent with that reported in a Chinese patent (Fig. S7†), which confirmed its monohydrate form.²⁴

3.1.5. DSC and TGA analyses. The DSC and TGA curves of the RBM solvates are illustrated in Fig. 5(b) and 6, respectively. An endothermic peak was observed in the DSC curve of RBM/DMSO at 100–150 °C, which was ascribed to the removal of DMSO. Mass reduction at the same temperature range was 27.77%, which was consistent with the theoretical DMSO content (29.65%) based on lattice DMSO with 1 : 2 stoichiometry. RBM/CH₂Cl₂ achieved a mass reduction of 12.55% during the first period, which was attributed to the removal of CH₂Cl₂ molecules. The experimental weight loss was lower than the theoretical weight loss (18.64%), thus indicating that molecules were partially lost before the TGA test because of the very low boiling point of CH₂Cl₂. In the DSC curve, RBM/CH₂Cl₂ exhibited a broad endothermic peak before 120 °C (removal of CH₂Cl₂) and further showed an exothermic phase transformation at 210–220 °C. The DSC curve of RBM/H₂O showed a broad endothermic peak before 120 °C and a small exothermic peak at 280–290 °C, which were ascribed to dehydration and transition, respectively. The weight loss of 4.28% during the

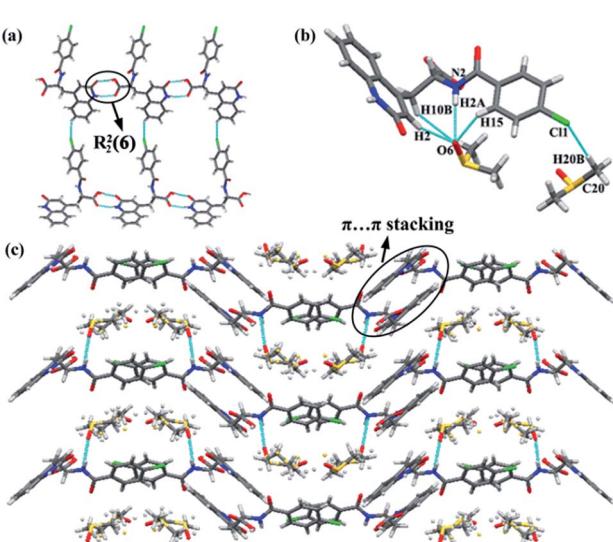
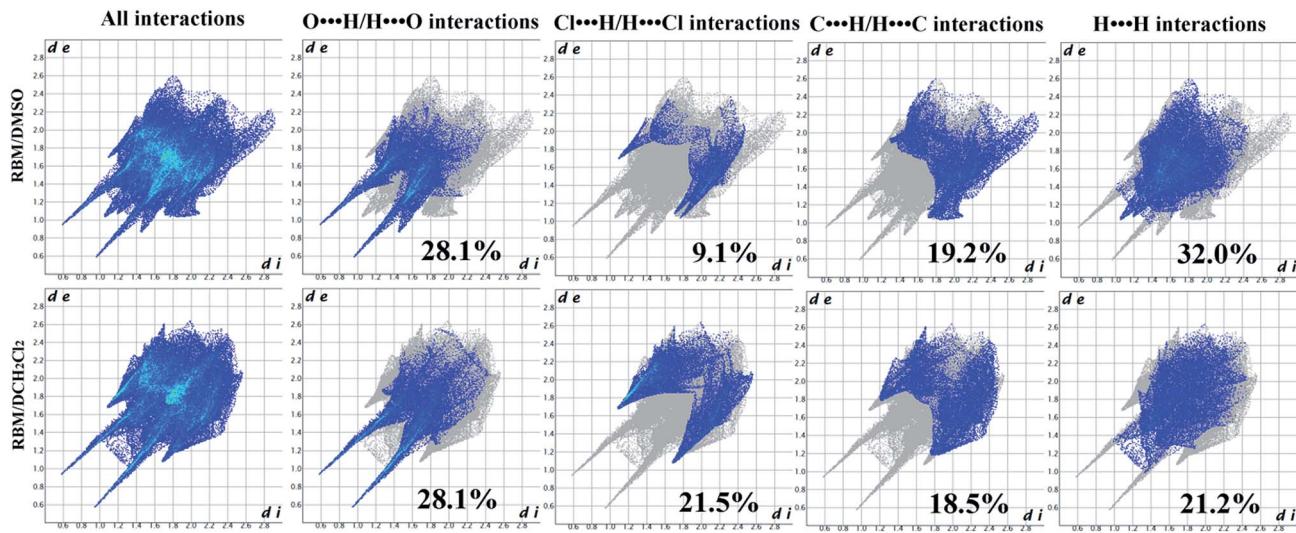
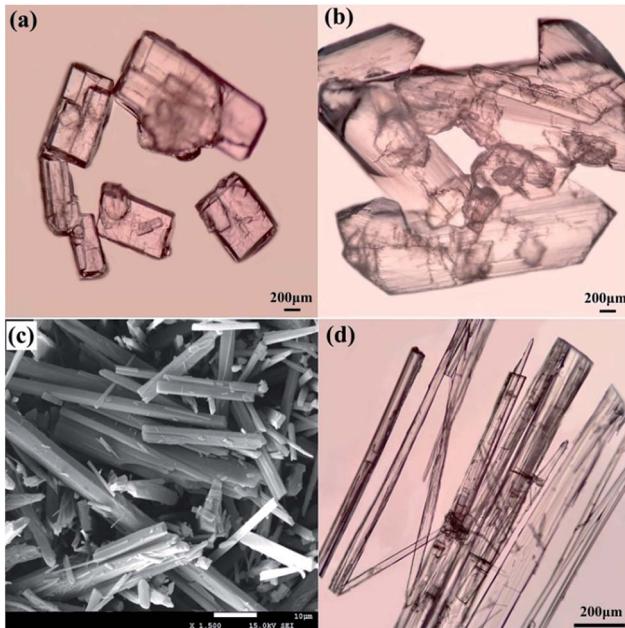


Fig. 1 (a) Hydrogen bond network of RBM/DMSO sustained by R₂(6) dimer motifs and weak C-H \cdots Cl hydrogen bonds. (b) Interactions between DMSO and RBM molecules. (c) View of the 2D packing of RBM/DMSO down the c-axis. DMSO molecules were located between the W-shaped layers.

Fig. 3 2D FPs of RBM/DMSO and RBM/CH₂Cl₂.Fig. 4 Morphological images of RBM solvates: (a) RBM/DMSO, (b) RBM/CH₂Cl₂, (c) RBM/H₂O, and (d) RBM/EtOH/H₂O.

first stage indicated that the stoichiometric ratio of RBM to water molecules was 1 : 1 [$\Delta m_{\text{cal.}} (-\text{H}_2\text{O}) = 4.63\%$]. RBM/EtOH/H₂O showed two mass steps of 5.83% and 2.80% at 50–100 °C and 140–160 °C, respectively. This phenomenon suggested the possible presence of 0.5EtOH [$\Delta m_{\text{cal.}} (-0.5\text{EtOH}) = 5.85\%$] and 0.5H₂O [$\Delta m_{\text{cal.}} (-0.5\text{H}_2\text{O}) = 2.37\%$] molecules in the crystal structure of RBM/EtOH/H₂O. The two broad endothermic peaks before 200 °C in the DSC curve further confirmed the hemiethanol hemihydrate structure of RBM/EtOH/H₂O. Furthermore, these two peaks were attributed to the removal of EtOH and H₂O molecules. The peak at approximately 300 °C for each DSC curve was broad or split, indicating that the solid forms

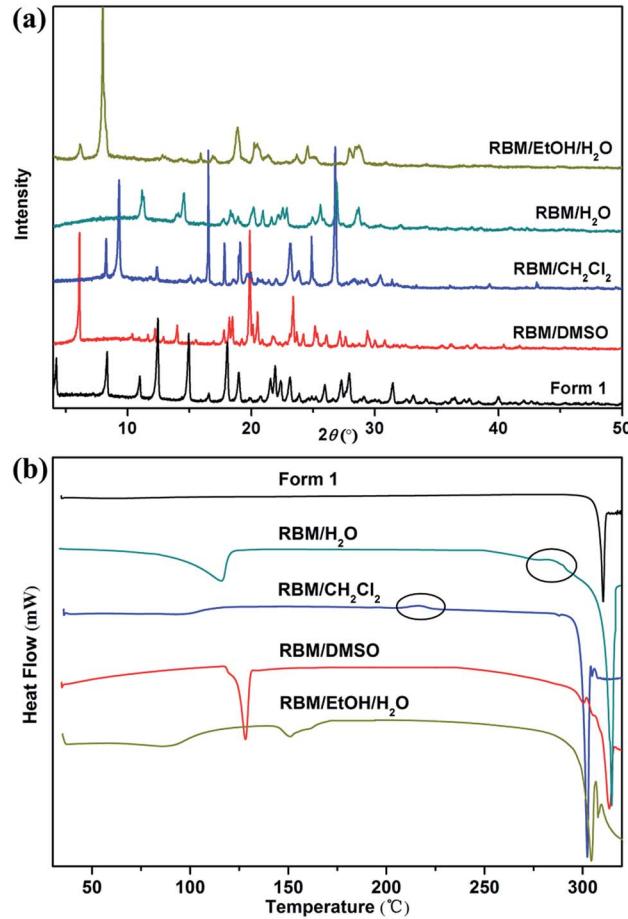


Fig. 5 (a) XRPD patterns and (b) DSC curves of the four solvates and form 1 of RBM.

obtained through desolvation or transformation melted with decomposition. This phenomenon was consistent with that indicated by the TGA results.

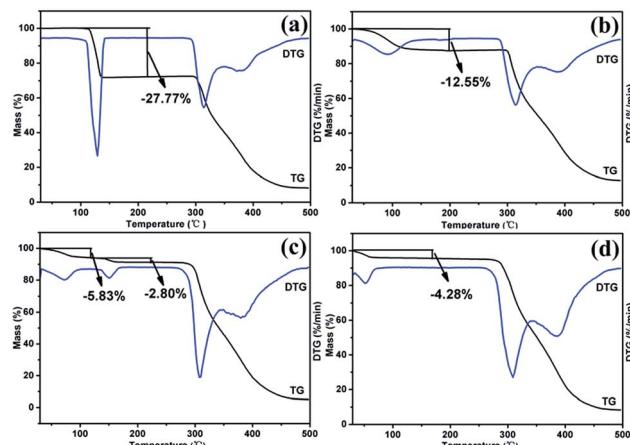


Fig. 6 TGA curves of the RBM solvates: (a) RBM/DMSO, (b) RBM/ CH_2Cl_2 , (c) RBM/ H_2O , and (d) RBM/EtOH/ H_2O .

3.2. Phase transitions upon heating

The modifications formed in solvates by solvent removal must be determined for screening novel solid forms. The phase transitions upon heating were investigated based on the thermal events displayed in the DSC curves.

RBM/DMSO and RBM/EtOH/ H_2O transformed to form 1 after desolvation (Fig. S8†). The XRPD patterns detected at 150 °C and 220 °C for RBM/ CH_2Cl_2 were considerably different from those for RBM/ CH_2Cl_2 [Fig. 7(a)]. No weight loss was observed in the TGA curves before melting and decomposition, and no desolvation peak was detected in the DSC curves [Fig. 7(a) and (c)].

Moreover, the two XRPD patterns were different from that of forms 1 and 2 (Fig. S9†).²³ Thus, two novel polymorphs (forms 3 and 4) were generated while heating RBM/ CH_2Cl_2 . Together with the DSC result, form 3 was generated by the removal of CH_2Cl_2 (150 °C) and was transformed to form 4 when further heated at 220 °C [Fig. 7(d)]. A novel polymorph (form 5) was obtained after dehydration of RBM/ H_2O , as indicated by the XRPD, TGA, and DSC results at 150 °C. The exothermic peak at 280–290 °C demonstrated that form 5 transformed into form 4 when heated at high temperature. In conclusion, three novel RBM polymorphs were obtained by the desolvation and transformation of the solvates. This phenomenon is significant for RBM development because the novel polymorphs may possess desirable physical and chemical properties.

3.3. Accelerated stability tests

Fig. 8 displays the XRPD patterns of the RBM solid forms, which were monitored at 40 °C and 75% RH. RBM/DMSO and RBM/EtOH/ H_2O were both converted to form 1 within three days [Fig. 8(a) and (d)]. RBM/ CH_2Cl_2 transformed to form 3, which could remain stable for three months at 40 °C and 75% RH [Fig. 8(b)]. Both RBM/ H_2O and form 4 showed relatively good stability because their XRPD patterns were nearly unchanged for three months [Fig. 8(c) and (e)]. Form 5 was unstable and could easily absorb water for reconversion to RBM/ H_2O [Fig. 8(f)].

3.4. Solubility and dissolution studies

Considering stability and security, the apparent solubility and dissolution rate of RBM/ H_2O , form 3, and form 4 were

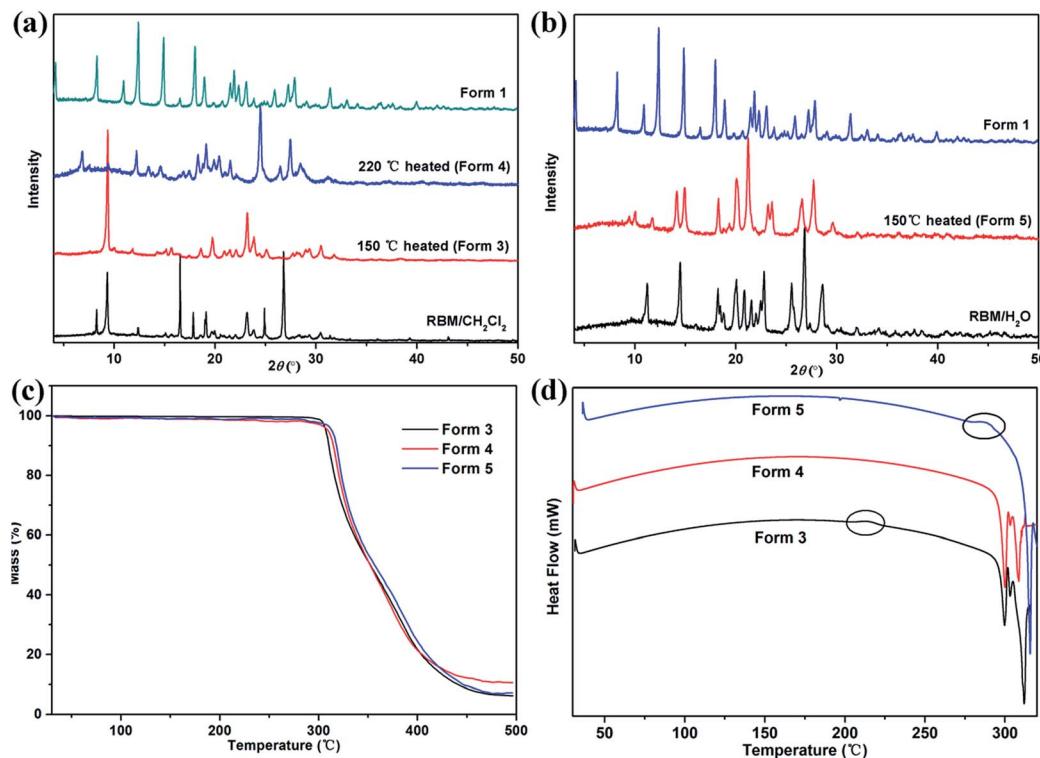


Fig. 7 XRPD patterns of (a) RBM/ CH_2Cl_2 heated at 150 °C and 220 °C, (b) RBM/ H_2O heated at 150 °C, (c) TGA, and (d) DSC curves of forms 3–5.



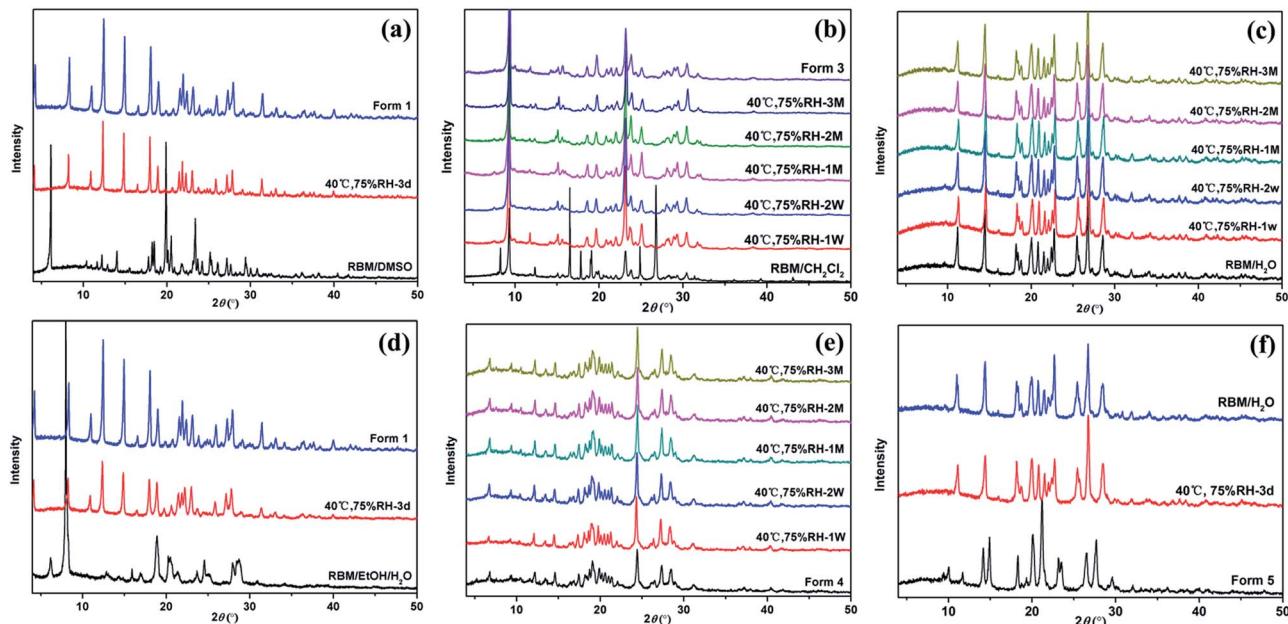


Fig. 8 Stability of (a) RBM/DMSO, (b) RBM/CH₂Cl₂ and form 3, (c) RBM/H₂O, (d) RBM/EtOH/H₂O, (e) form 4, and (f) form 5 monitored with XRPD at 40 °C and 75% RH. (d = day, W = week, M = month).

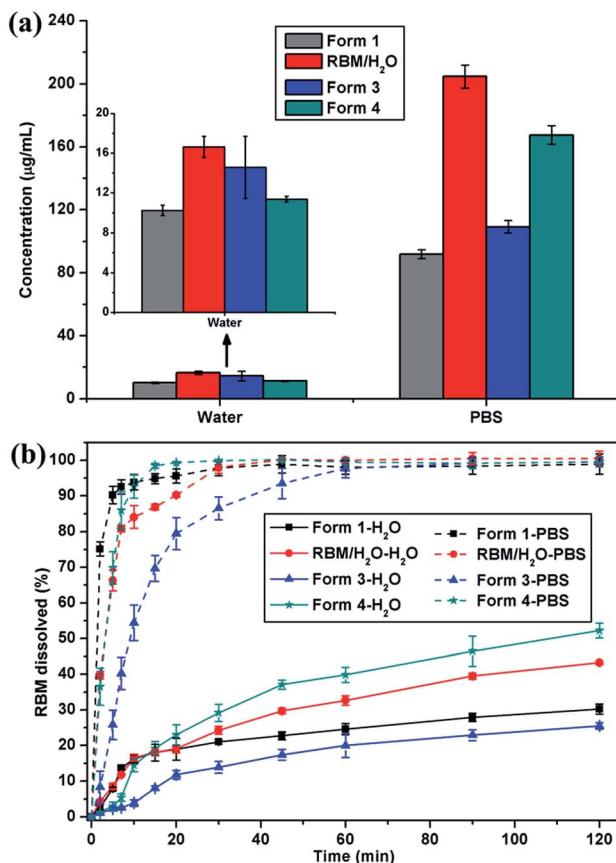


Fig. 9 (a) Solubility histograms and (b) dissolution profiles of form 1, RBM/H₂O, form 3, and form 4 in water and PBS (mean \pm standard deviation, $n = 3$).

investigated and compared with those of form 1. RBM/H₂O, form 3, and form 4 demonstrated higher solubility than form 1 in water and PBS [Fig. 9(a)]. Especially RBM/H₂O, the solubility of which was 1.62 and 2.23 times higher than form 1 in water and PBS, respectively. The solid residues obtained by centrifugation and filtration after solubility determination were analyzed by XRPD (Fig. S10†). Although forms 3 and 4 were converted to form 1, the apparent solubility of the two forms remained higher than that of the original form 1 in water and PBS. No transformation occurred for RBM/H₂O. Moreover, its solid residues showed increased crystallinity after solubility examination. Thus, RBM/H₂O is a potential substitute for form 1 in clinical treatment given its dominant solubility and good stability.

RBM/H₂O and form 4 showed a faster dissolution rate than form 1 after 20 min of dissolution in water; moreover, the cumulative drug release of the forms was considerably higher than that of form 1 after 120 min [Fig. 9(b)]. When dissolved in PBS, the three RBM solid forms exhibited a slower dissolution than form 1. However, their cumulative dissolution reached approximately 100% within 60 min. RBM/H₂O and form 4, in particular, almost completely dissolved after 30 min. These results indicated that RBM/H₂O and form 4 had a comparative or even better solubility and dissolution characteristics than form 1. Thus, RBM/H₂O and form 4 may be developed as novel solid forms of RBM for clinical applications.

4. Conclusions

Four solvates and three novel polymorphs of RBM were obtained. The crystal structures of RBM/DMSO and RBM/CH₂Cl₂ were determined *via* SXRD. Both structures were sustained by R₂(8) dimer motifs which formed through O–H···O and N–H···O

hydrogen bonds. The main difference between the two forms was that additional hydrogen bonds (N2–H2···O1) which were absent from RBM/DMSO, were present in RBM/CH₂Cl₂. The 2D FP result indicated that the subtle intermolecular interactions of RBM/DMSO and RBM/CH₂Cl₂ were significantly different. The stability results at 40 °C and 75% RH indicated that RBM/DMSO and RBM/EtOH/H₂O transformed to form 1 and that form 5 was converted to RBM/H₂O. RBM/H₂O, form 3, and form 4 showed good stability. In addition, RBM/H₂O, form 3, and form 4 displayed a higher solubility than form 1 in water and PBS. Furthermore, the dissolution rates of RBM/H₂O and form 4 in water were faster than that of form 1. Thus, RBM/H₂O and form 4 exhibited potential to be developed as novel solid forms of RBM for clinical applications, which is meaningful for the drug development of RBM.

Acknowledgements

This work was supported by the Applied Basic Research Project of Sichuan Province (Grant No. 2014JY0042), the National Development and Reform Commission and Education of China (Grant No. 2014BW011), and the Large-scale Science Instrument Shareable Platform Construction of Sichuan Province (Grant No. 2015JCPT0005-15010102).

References

- 1 C. Näther, I. Jess, P. G. Jones, C. Taouss and N. Teschmit, *Cryst. Growth Des.*, 2013, **13**, 1676–1684.
- 2 R. Bobrovs, L. Seton and N. Dempster, *CrystEngComm*, 2015, **17**, 5237–5251.
- 3 J. P. Brog, C. L. Chanez, A. Crochet and K. M. Fromm, *RSC Adv.*, 2013, **3**, 16905.
- 4 N. Gong, K. Hu, G. Jin, G. Du and Y. Lu, *RSC Adv.*, 2016, **6**, 38709–38715.
- 5 S. H. Thorat, M. V. Patwadkar, R. G. Gonnade and R. Vaidhyanathan, *CrystEngComm*, 2014, **16**, 8638–8641.
- 6 K. Xu, S. Zheng, Y. Zhai, L. Guo, P. Tang, J. Yan, D. Wu and H. Li, *Int. J. Pharm.*, 2015, **486**, 185–194.
- 7 D. E. Braun, P. G. Karamertzanis, J. B. Arlin, A. J. Florence, V. Kahlenberg, D. A. Tocher, U. J. Griesser and S. L. Price, *Cryst. Growth Des.*, 2011, **11**, 210–220.
- 8 S. Datta and D. J. Grant, *Nat. Rev. Drug Discovery*, 2004, **3**, 42–57.
- 9 D. Maddileti, B. Swapna and A. Nangia, *Cryst. Growth Des.*, 2015, **15**, 1745–1756.
- 10 C. Näther, I. Jess, L. Seyfarth, K. Bärwinkel and J. Senker, *Cryst. Growth Des.*, 2015, **15**, 366–373.
- 11 A. O. Surov, K. A. Solanko, A. D. Bond, A. Bauer-Brandl and G. L. Perlovich, *CrystEngComm*, 2015, **17**, 4089–4097.
- 12 U. J. Griesser, in *The importance of solvates, Polymorphism in the Pharmaceutical Industry*, 2006, pp. 211–233.
- 13 B. N. Patel, N. Bhanubhai and C. N. Suhagia, *Int. J. Pharm. Pharm. Sci.*, 2012, **4**, 270–273.
- 14 N. Variankaval, A. S. Cote and M. F. Doherty, *AIChE J.*, 2008, **54**, 1682–1688.
- 15 W. Xu, N. Gong, S. Yang, N. Zhang, L. He, G. Du and Y. Lu, *J. Pharm. Sci.*, 2015, **104**, 1256–1262.
- 16 H. X. Ding, C. A. Leverett, R. E. Kyne, K. K. C. Liu, S. J. Fink, A. C. Flick and C. J. O'Donnell, *Bioorg. Med. Chem.*, 2015, **23**, 1895–1922.
- 17 X. Wang, N. Gong, S. Yang, G. Du and Y. Lu, *J. Pharm. Sci.*, 2014, **103**, 2696–2703.
- 18 Z. Yin, Z. Li, W. Z. Yang, J. He and J. P. Cheng, *Struct. Chem.*, 2005, **16**, 641–647.
- 19 P. W. Cains, in *Polymorphism in pharmaceutical solids: classical methods of preparation of polymorphs and alternative solid forms*, 2009, pp. 76–138.
- 20 R. Pradhan, T. H. Tran, J. Y. Choi, I. S. Choi, H. G. Choi, C. S. Yong and J. O. Kim, *Arch. Pharmacal Res.*, 2014, **38**, 522–533.
- 21 N. T. Tung, C. W. Park, T. O. Oh, J. Y. Kim, J. M. Ha, Y. S. Rhee and E. S. Park, *J. Pharm. Pharmacol.*, 2011, **63**, 1539–1547.
- 22 S. Koh, Y. Inoue, T. Sugimoto, N. Maeda and K. Nishida, *Cornea*, 2013, **32**, 1219–1223.
- 23 S. H. Jeon and Y. T. Sohn, *Arch. Pharmacal Res.*, 2016, **39**, 508–515.
- 24 X. Zeng and C. Jia, Preparation of a rebamipide monohydrate crystal form for use in drugs for treating diseases related to gastric mucosa injury, China Patent, CN 104418802, March 18, 2015.
- 25 J. Holland, D. Gooding and A. Chorlton, Novel rebamipide complexes and cocrystals, U.S. Patent, 20140039005, February 6, 2014.
- 26 Y. Chi, C. Liu, T. Ren, X. Wang, Q. Yang, Z. Yang, Y. Yang, S. Yang, J. Gu and C. Hu, *Cryst. Growth Des.*, 2016, **16**, 3180–3189.
- 27 X. Xiong, K. Xu, S. Li, P. Tang, Y. Xiao and H. Li, *Drug Dev. Ind. Pharm.*, 2017, **43**, 283–292.
- 28 O. V. Dolomanov, L. J. Bourhis and R. J. Gildea, *J. Appl. Crystallogr.*, 2009, **42**, 339–341.
- 29 G. M. Sheldrick, *Acta Crystallogr., Sect. C: Struct. Chem.*, 2015, **71**, 3–8.
- 30 C. F. Macrae, P. R. Edgington, P. McCabe, E. Pidcock, G. P. Shields, R. Taylor, M. Towler and J. van de Streek, *J. Appl. Crystallogr.*, 2006, **39**, 453–457.
- 31 M. A. Spackman and P. G. Byrom, *Chem. Phys. Lett.*, 1997, **267**, 215–220.
- 32 M. A. Spackman and D. Jayatilaka, *CrystEngComm*, 2009, **11**, 19–32.
- 33 J. Bojarska, A. Fruziński and W. Maniukiewicz, *J. Mol. Struct.*, 2016, **1116**, 22–29.
- 34 M. A. Spackman and J. J. McKinnon, *CrystEngComm*, 2002, **4**, 378–392.
- 35 S. K. Wolff, D. J. Grimwood, J. J. McKinnon, M. J. Turner, D. Jayatilaka and M. A. Spackman, *Crystal explorer*, The University of Western Australia, 2012.
- 36 W. H. Organization, in *World Health Organization Technical Report Series: Annex 2: Stability testing of active pharmaceutical ingredients and finished pharmaceutical products*, 2009, pp. 87–130.
- 37 T. Lu and C. Chen, *Measurement*, 2007, **40**, 591–599.
- 38 J. Bernstein, R. E. Davis, L. Shimoni and N. L. Chang, *Angew. Chem., Int. Ed. Engl.*, 1995, **34**, 1555–1573.

