

PAPER

View Article Online
View Journal | View IssueCite this: *RSC Adv.*, 2017, 7, 23041

Received 7th March 2017

Accepted 19th April 2017

DOI: 10.1039/c7ra02788e

rsc.li/rsc-advances

Reusable ionic liquid-catalyzed oxidative esterification of carboxylic acids with benzylic hydrocarbons *via* benzylic Csp³–H bond activation under metal-free conditions†

Fen Mou, , Yadong Sun, , Weiwei Jin, , Yonghong Zhang, , Bin Wang, ,
Zhiqing Liu, , Lei Guo, , Jianbin Huang and Chenjiang Liu *

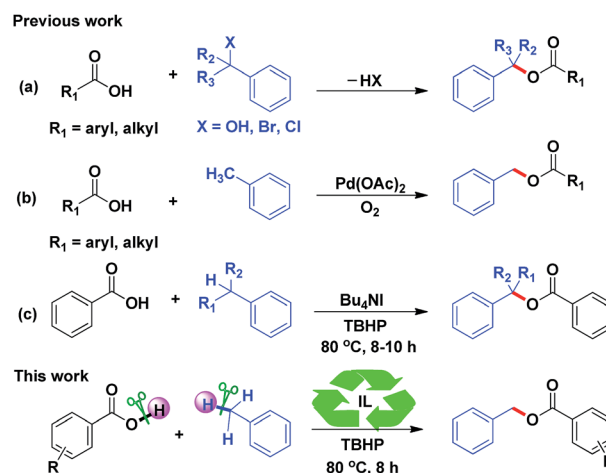
A metal-free protocol for the direct oxidative esterification of the Csp³–H bond in benzylic hydrocarbons with carboxylic acids using heterocyclic ionic liquid as catalyst has been reported. The catalyst 1-butylpyridinium iodide could be easily recycled and reused for at least four cycles without obvious loss of catalytic activity.

Introduction

It is well known that benzyl esters are prevalent functional groups in medicinal agents and natural products.¹ The traditional methods of preparing benzyl esters by utilizing benzyl alcohols/bromides/chlorides with carboxylic acids as coupling partners have been reported over the decades (Scheme 1a).² Furthermore, Khan's group proposed benzyl esters can also be obtained by the reaction of aromatic carboxylic acids with substituted toluenes in the presence of NaBrO₃/NaHSO₃ in early 2003.³ However, some approaches have disadvantages of long reaction time or sensitivity to the environment. With increasing interest in transition metal-catalyzed C–H bonds activation to construct C–O bonds, several types of transition metals, including Cu, Pd, Pt and Rh have been recently revealed admirable catalytic activity in esterification.⁴ Among them, the approach of Pd-catalyzed benzylation of toluene with carboxylic acids under 1 atm of oxygen has been disclosed by Zhang's group (Scheme 1b).⁴ It did not need to be pre-functionalized and provided a more facile, atom-economic method to prepare benzyl esters. Despite their obvious potential application, these transition metal-catalyzed C–H bond activation methods suffered from a few drawbacks such as expense and toxicity of transition metals, which limit their practicability. Hence, the development of a facile and practical approach for constructing benzyl esters under metal-free conditions is extremely desirable.

In 2012, Yu and co-workers established Bu₄Ni-catalyzed esterification of benzylic Csp³–H bonds (Scheme 1c).⁵ Subsequently, Singh's group applied carboxylic acids with alkyl benzenes to synthesis of benzyl esters catalyzed by NBS with TBHP as an oxidant under alkaline conditions.⁶ Moreover, Patel's group presented a strategy for the efficient preparation of benzylic esters through Bu₄Ni-catalyzed cross dehydrogenative coupling of alkylbenzenes as the only precursor.⁷ The main advantage of these methodologies is their environmentally friendliness.

In recent years, ionic liquids (ILs) are widely used as an environmentally eco-friendly and recyclable reaction medium or catalyst in green synthesis,⁸ which have attracted great attention of chemists all around the world. Compared with other organic agents, ILs have unique advantages such as low toxicity, easy recyclability, excellent chemical and thermal



Scheme 1 A variety of methods for the synthesis of benzyl esters.

The Key Laboratory of Oil and Gas Fine Chemicals, Ministry of Education, Xinjiang Uygur Autonomous Region, Urumqi Key Laboratory of Green Catalysis and Synthesis Technology, School of Chemistry and Chemical Engineering, Physics and Chemistry Detecting Center, Xinjiang University, Urumqi 830046, P. R. China. E-mail: pxylcj@126.com

† Electronic supplementary information (ESI) available. See DOI: 10.1039/c7ra02788e

stability, negligible vapor pressure, as well as good solubility with organic and inorganic compounds.⁹ Latterly, transition metal-catalyzed C–H activation reactions with ionic liquids as solvents have already been developed.¹⁰ Up to now, ionic liquids promoted C–C bond and catalyzed C–N bond formation reactions have been published by our group.¹¹ However, as far as we know, classical heterocyclic ILs-catalyzed benzylic Csp³–H activation for C–O bond formation to prepare benzyl esters have been rarely reported. With this background in mind, we herein report a heterocyclic IL-catalyzed the direct oxidative esterification between Csp³–H bond of benzylic hydrocarbons and carboxylic acids for the synthesis of benzyl esters under metal-free conditions.

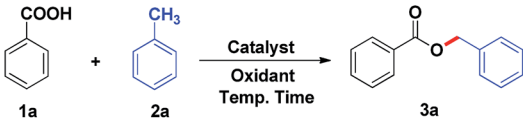
Results and discussion

We commenced our study using benzoic acid (**1a**) and toluene (**2a**) as a model system (Table 1). Firstly, the reaction of **1a** and **2a** was carried out at 80 °C for 8 h to give the desired benzyl ester product **3a** in 70% yield in the presence of 1.4 equiv. of TBHP and 20 mol% [BPy]I (Table 1, entry 1). Encouraged by this result, further optimization of conditions for the oxidative esterification was explored. Unfortunately, other ionic liquids such as 1-butylpyridinium bromide ([BPy]Br) and 1-butylpyridinium chloride ([BPy]Cl) were proved to be poor activity for this esterification (Table 1, entries 2–3). Meanwhile, it was found that essentially no reaction occurred in the catalyst-free system (Table 1, entry 4). The yield of product **3a** was not improved obviously by increasing the amount of TBHP to 2 equiv. (Table 1, entry 5). Subsequently, when the dosage of [BPy]I was increased from 20 mol% to 25 mol%, there was no significant

improvement in yield (Table 1, entry 6). In contrast, decreasing the amount of [BPy]I resulted in dramatically reducing the yield of **3a** (Table 1, entries 7–8). After a brief screening of oxidants such as 3-chloroperbenzoic acid, hydrogen peroxide, di-*tert*-butyl peroxide and iodobenzene diacetate, TBHP remained the optimal choice (Table 1, entries 1, 9–12). Finally, the effects of reaction temperature and reaction time were also screened (Table 1, entries 13–17), the results suggested that the optimized reaction temperature and time were 80 °C and 8 h, respectively. Therefore, the reaction was carried out with [BPy]I (20 mol%) as the catalyst and TBHP (1.4 mmol, 70% aqueous solution) as the oxidant at 80 °C for 8 h.

Under the optimized reaction conditions, the scope of carboxylic acids was explored. As shown in Table 2, a wide variety of benzoic acid derivatives with electron-donating substitutes and electron-withdrawing substitutes underwent this esterification smoothly and afforded the corresponding benzyl esters (**3b–3t**) in good to excellent yields. It is important to note that the methyl or nitro group substituted at *ortho*-, *meta*- or *para*-position of benzene ring were all compatible (**3b–3d**, **3q–3s**). In particular, the multisubstituted substrates such as 2,3,4,5,6-pentafluorobenzoic acid, 3,5-dinitrobenzoic acid could be transformed well with 73% and 93% yield, respectively (**3k**, **3t**). All of the above results indicated that the electronic and steric effects had no significant influence on the yield of products. To our delight, cinnamic acid and terephthalic acid were also amenable substrates (**3u**, **3v**). In order to better expand substrates scope, aromatic heterocyclic carboxylic acid such as 2-phenyl-2H-1,2,3-triazole-4-carboxylic acid and 2-furoic acid

Table 1 Optimizing the reaction conditions^a

					
Entry	Catalyst (mol%)	Oxidant ^b (1.4 mmol)	Temp. (°C)	Time (h)	Yield ^c (%)
1	[BPy]I (20)	TBHP	80	8	70
2	[BPy]Br (20)	TBHP	80	8	Trace
3	[BPy]Cl (20)	TBHP	80	8	N.R. ^d
4	—	TBHP	80	8	N.R.
5	[BPy]I (20)	TBHP	80	8	72 ^e
6	[BPy]I (25)	TBHP	80	8	71
7	[BPy]I (15)	TBHP	80	8	57
8	[BPy]I (10)	TBHP	80	8	Trace
9	[BPy]I (20)	<i>m</i> -CPBA	80	8	N.R.
10	[BPy]I (20)	H ₂ O ₂	80	8	N.R.
11	[BPy]I (20)	DTBP	80	8	N.R.
12	[BPy]I (20)	PhI(OAc) ₂	80	8	N.R.
13	[BPy]I (20)	TBHP	100	8	72
14	[BPy]I (20)	TBHP	60	8	26
15	[BPy]I (20)	TBHP	80	10	70
16	[BPy]I (20)	TBHP	80	6	60

^a Reaction conditions: **1a** (1 mmol), **2a** (20 mmol), oxidant (1.4 mmol). ^b TBHP: *tert*-butyl hydroperoxide 70% in water; *m*-CPBA: *m*-chloroperoxybenzoic acid; DTBP: di-*tert*-butyl peroxide; H₂O₂: 30% in water. ^c Isolated yield. ^d Not reaction. ^e TBHP: 2.0 mmol.



Table 2 Esterification of toluene with various carboxylic acids^a

 3a: 70%	 3b ^b : 94%	 3c: 74%
 3d ^b : 91%	 3e ^b : 75%	 3f: 94%
 3g ^b : 88%	 3h ^b : 84%	 3i ^b : 88%
 3j ^b : 88%	 3k: 73%	 3l: 92%
 3m: 92%	 3n: 84%	 3o: 73%
 3p: 80%	 3q: 86%	 3r: 76%
 3s: 87%	 3t: 93%	 3u: 65%
 3v: 48%	 3w: 85%	 3x: 55%

^a Reaction conditions: **1** (1 mmol), **2a** (20 mmol), [BPy]I (20 mol%), TBHP (1.4 mmol) at 80 °C for 8 h unless otherwise noted. ^b TBHP: 2.0 mmol.

were explored and the corresponding products (**3w**, **3x**) were obtained in 85% and 55% yields, respectively.

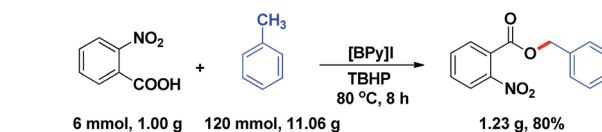
Subsequently, we investigated the scope of benzylic hydrocarbons. As shown in Table 3, a wide range of toluene derivatives with electron-donating groups and electron-withdrawing groups were well tolerated and obtained the desired products (**4a–4g**) in moderate to good yields. It is worth noting that *p*-xylene and mesitylene only gave monoesterification products (**4a**, **4b**) in 78% and 82% yields, respectively. When the substrate with *p*-OCH₃ was used, the target product (**4c**) was afforded in yield of 71%. In particular, the chloro group substituted at *ortho*-, *meta*- or *para*-position of benzene ring were all amenable substrates (**4e–4g**). Additionally, both ethylbenzene and cumene were suitable reactants, affording the corresponding products (**4h**, **4i**) in moderate yields.

To further demonstrate the potential application of the esterification method in industry, an IL-catalyzed gram-scale

Table 3 Esterification of benzoic acid with benzylic hydrocarbons^a

 4a: 78%	 4b ^b : 82%	 4c: 71%
 4d: 62%	 4e: 38%	 4f: 45%
 4g ^b : 64%	 4h ^b : 65%	 4i: 50%

^a Reaction conditions: **1a** (1 mmol), **2** (20 mmol), [BPy]I (20 mol%), TBHP (1.4 mmol) at 80 °C for 8 h unless otherwise noted. ^b TBHP: 2.0 mmol.



Scheme 2 Gram-scale oxidative esterification.

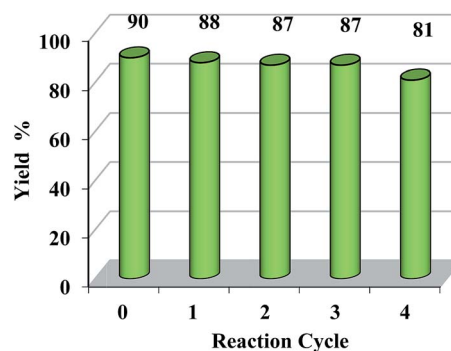
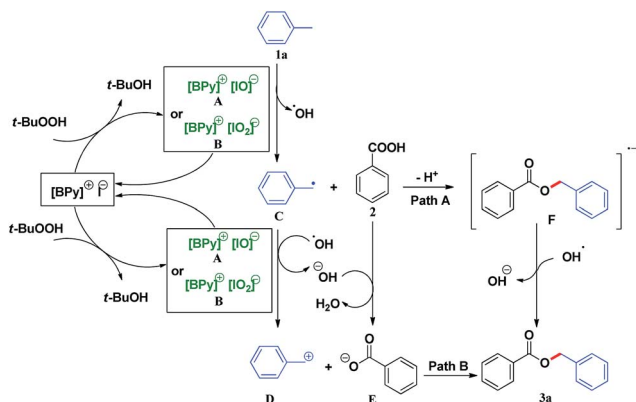


Fig. 1 Recycling reactions.

esterification between 2-nitrobenzoic acid and toluene was investigated (Scheme 2). The result showed that the reaction was easily performed under the optimized reaction conditions to give the desired benzyl ester in 80% isolated yield.

In order to validate the reaction mechanism, control experiments were carried out (see the ESI†). In the presence of 2 equiv. of TEMPO (2,2,6,6-tetramethylpiperidine-*N*-oxyl) as a free radical scavenger, the esterification was almost inhibited under the optimal conditions. When the amount of TEMPO was increased to 3 equiv., benzyl benzoate as the target could not be obtained at all. The results showed that the esterification may involve a radical pathway.





Scheme 3 The proposed mechanism.

Based on the results of control experiments and previous studies,^{5,11b,12} a possible radical mechanism is proposed (Scheme 3). Initially, [BPy]I is oxidized by TBHP to generate $\{[BPy]^+[IO]^- \}$ A or $\{[BPy]^+[IO_2]^- \}$ B species. Subsequently, the benzyl radical C is obtained from the homolytic cleavage of a benzylic C–H bond in the presence of A or B, followed by combination with benzoic acid to give the benzyl ester radical anion F. Finally, the desired product benzyl ester is produced by losing an electron from F with the assistance of hydroxyl radical (Scheme 3, path A). Furthermore, the benzyl radical is easily oxidized by active iodine species A or B to give the benzyl cation D and the hydroxide ion simultaneously. Subsequently, benzoic acid is deprotonated by the hydroxide ion to form the benzoate anion E. At last, the electrostatic attraction between the benzoate anion with specie E to yield corresponding product 3a (Scheme 3, path B).

It's worth noting that ionic liquid could be easily recycled. 2-Chlorobenzoic acid and methylbenzene were chosen as model substrates to study the reusability of the catalyst [BPy]I under standard reaction conditions. After completion of the reaction, ethyl acetate and water were added, and then the ionic liquid [BPy]I was recovered from the aqueous phase and reused without the significant loss of its activity (Fig. 1).

Conclusions

In summary, we have developed classic heterocyclic ionic liquid-catalyzed oxidative esterification of carboxylic acid derivatives with benzylic hydrocarbons. Most importantly, an inexpensive and recyclable ionic liquid [BPy]I is reused for at least four cycles with similar catalytic activity. Ionic liquid catalyzed C–H bond activation reactions are ongoing in our laboratory.

Acknowledgements

This work is supported by Grants from Key Laboratory of Xinjiang Uyghur Autonomous Region (2015KL014) and the National Natural Science Foundation of China (21572195 and 21262035).

Notes and references

- (a) V. Vindigni, R. Cortivo, L. Iacobellis, G. Abatangelo and B. Zavan, *Int. J. Mol. Sci.*, 2009, **10**, 2972; (b) A. Aliboni, A. D'Andrea and P. Massanisso, *J. Agric. Food Chem.*, 2011, **59**, 282; (c) C. S. Pappas, A. Malovikova, Z. Hromadkova, P. A. Tarantilis, A. Ebringerova and M. G. Polissiou, *Carbohydr. Polym.*, 2004, **56**, 465; (d) A. Isidro-Llobet, M. Álvarez and F. Albericio, *Chem. Rev.*, 2009, **109**, 2455.
- Selected traditional esterification methods: (a) J. C. Lee, Y. S. Oh, S. H. Cho and J. D. Lee, *Org. Prep. Proced. Int.*, 1996, **28**, 480; (b) C.-T. Chen and Y. S. Munot, *J. Org. Chem.*, 2005, **70**, 8625; (c) F. Cardellini, L. Brinchi, R. Germani and M. Tiecco, *Synth. Commun.*, 2014, **44**, 3248; (d) R. Kore and R. Srivastava, *Catal. Commun.*, 2011, **12**, 1420; (e) N. Iranpoor, H. Firouzabadi and D. Khalili, *Org. Biomol. Chem.*, 2010, **8**, 4436; (f) X. Wu, X. Han, L. Zhou and A. Li, *Indian J. Chem., Sect. A: Inorg., Bio-inorg., Phys., Theor. Anal. Chem.*, 2012, **51**, 791; (g) X. Li, W. Eli and G. Li, *Catal. Commun.*, 2008, **9**, 2264; (h) K. Kankanala, V. R. Reddy, K. Mukkanti and S. Pal, *J. Fluorine Chem.*, 2009, **130**, 505.
- K. M. Khan, G. M. Maharvi, S. Hayat, Z. Ullah, M. I. Choudhary and A.-u. Rahman, *Tetrahedron*, 2003, **59**, 5549.
- Selected esterification protocols through C–H activation: (a) X. Chen, X.-S. Hao, C. E. Goodhue and J.-Q. Yu, *J. Am. Chem. Soc.*, 2006, **128**, 6790; (b) A. R. Dick, K. L. Hull and M. S. Sanford, *J. Am. Chem. Soc.*, 2004, **126**, 2300; (c) M. S. Chen, N. Prabakaran, N. A. Labenz and M. C. White, *J. Am. Chem. Soc.*, 2005, **127**, 6970; (d) J. H. Delcamp and M. C. White, *J. Am. Chem. Soc.*, 2006, **128**, 15076; (e) D. J. Covell and M. C. White, *Angew. Chem., Int. Ed.*, 2008, **47**, 6448; (f) T. Yoneyama and R. H. Crabtree, *J. Mol. Catal. A: Chem.*, 1996, **108**, 35; (g) L. V. Desai, K. L. Hull and M. S. Sanford, *J. Am. Chem. Soc.*, 2004, **126**, 9542; (h) R. Giri, J. Liang, J.-G. Lei, J.-J. Li, D.-H. Wang, X. Chen, I. C. Naggar, C. Guo, B. M. Foxman and J.-Q. Yu, *Angew. Chem., Int. Ed.*, 2005, **44**, 7420; (i) D. Li, M. Yu, J. Zhang, Z. Liu and Y. Zhang, *Org. Lett.*, 2015, **17**, 5300; (j) H. Liu, G. Shi, S. Pan, Y. Jiang and Y. Zhang, *Org. Lett.*, 2013, **15**, 4098; (k) J. M. Lee and S. Chang, *Tetrahedron Lett.*, 2006, **47**, 1375; (l) Z. Ye, W. Wang, F. Luo, S. Zhang and J. Cheng, *Org. Lett.*, 2009, **11**, 3974.
- J. Feng, S. Liang, S.-Y. Chen, J. Zhang, S.-S. Fu and X.-Q. Yu, *Adv. Synth. Catal.*, 2012, **354**, 1287.
- T. Guntreddi, R. Vanjari and K. N. Singh, *Tetrahedron*, 2014, **70**, 3887.
- G. Majji, S. Guin, A. Gogoi, S. K. Rout and B. K. Patel, *Chem. Commun.*, 2013, **49**, 3031.
- (a) P. Wasserscheid and W. Keim, *Angew. Chem., Int. Ed.*, 2000, **39**, 3772; (b) J. Dupont, R. F. Souza and P. A. Z. Suarez, *Chem. Rev.*, 2002, **102**, 3667; (c) V. I. Pârvulescu and C. Hardacre, *Chem. Rev.*, 2007, **107**, 2615; (d) B. Xin and J. Hao, *Chem. Soc. Rev.*, 2014, **43**, 7171; (e) J. R. Harjani, R. D. Singer, M. T. Garcia and P. J. Scammells, *Green Chem.*, 2009, **11**, 83.



- 9 (a) M. Petkovic, K. R. Seddon, L. P. N. Rebelo and C. S. Pereira, *Chem. Soc. Rev.*, 2011, **40**, 1383; (b) C. D. Hubbard, P. Illner and R. Eldik, *Chem. Soc. Rev.*, 2011, **40**, 272; (c) J. P. Hallett and T. Welton, *Chem. Rev.*, 2011, **111**, 3508; (d) M. E. Zakrzewska, E. Bogel-Lukasik and R. Bogel-Lukasik, *Chem. Rev.*, 2011, **111**, 397; (e) Q. Zhang and J. M. Shreeve, *Chem. Rev.*, 2014, **114**, 10527; (f) M. V. Fedorov and A. A. Kornyshev, *Chem. Rev.*, 2014, **114**, 2978.
- 10 (a) J. Li, W. Yang, S. Yang, L. Huang, W. Wu, Y. Sun and H. Jiang, *Angew. Chem., Int. Ed.*, 2014, **53**, 7219; (b) J. Li, S. Yang, H. Jiang, W. Wu and J. Zhao, *J. Org. Chem.*, 2013, **78**, 12477; (c) O. Baslé, N. Borduas, P. Dubois, J. M. Chapuzet, T. H. Chan, J. Lessard and C. J. Li, *Chem.–Eur. J.*, 2010, **16**, 8162; (d) S. B. Park and H. Alper, *Chem. Commun.*, 2005, 1315; (e) P. Ehlers, A. Petrosyan, J. Baumgard, S. Jopp, N. Steinfeld, T. V. Ghochikyan, A. S. Saghyian, C. Fischer and P. Langer, *ChemCatChem*, 2013, **5**, 2504; (f) A. Biffis, L. Gazzola, C. Tubaro and M. Basato, *ChemSusChem*, 2010, **3**, 834.
- 11 (a) H. Li, C. Liu, Y. Zhang, Y. Sun, B. Wang and W. Liu, *Org. Lett.*, 2015, **17**, 932; (b) W. Liu, C. Liu, Y. Zhang, Y. Sun, A. Abdukader, B. Wang, H. Li, X. Ma and Z. Zhang, *Org. Biomol. Chem.*, 2015, **13**, 7154; (c) D. Cao, Y. Zhang, C. Liu, B. Wang, Y. Sun, A. Abdukader, H. Hu and Q. Liu, *Org. Lett.*, 2016, **18**, 2000.
- 12 (a) Q. Xue, J. Xie, H. Li, Y. Cheng and C. Zhu, *Chem. Commun.*, 2013, **49**, 3700; (b) L. Wang, K. Zhu, Q. Chen and M. He, *J. Org. Chem.*, 2014, **79**, 11780.

