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## *In situ* strategy to prepare PDPB/SnO<sub>2</sub> p–n heterojunction with a high photocatalytic activity

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As a novel material for water depollution, conjugated polydiphenylbutadiyne (PDPB) nanofibers have received attention due to their visible light responsive photocatalytic activity. In this study, we successfully prepared a PDPB/SnO<sub>2</sub> p–n heterojunction by an *in situ* growth route. The XPS characterization indicates the generation of Sn–C bonds between SnO<sub>2</sub> and PDPB, which would decrease the bandgap of SnO<sub>2</sub> and promote the transference efficiency of electrons and holes between these two components. The PDPB could act as the sensitizer to enlarge the solar light absorption region of the heterojunction. The SnO<sub>2</sub> provides a stable mesoporous structure and enhanced hydrophilic properties. The solar-driven photodegradation rate of Rhodamine B over the PDPB/SnO<sub>2</sub> (ratio is 2 : 1) is 4 times higher than that of the pure SnO<sub>2</sub> and 2 times higher than that of the pure PDPB nanofibers. Our strategy gives a facile strategy for the preparation of an organic–inorganic hybrid heterojunction with high solar light activity.

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### 1. Introduction

In recent years, SnO<sub>2</sub> mesoporous single crystals (MSCs) have gained much attention due to their excellent properties such as perfect single-crystallinity, large specific surface areas, high electron mobility and exposed active sites.<sup>1–4</sup> These advantages have led to the successful research for their various applications in photocatalysis,<sup>5,6</sup> sensors<sup>7–9</sup> and solar cells.<sup>10,11</sup> However, SnO<sub>2</sub> MSCs have an inevitable drawback ascribed to their wide band gap, which limits solar light from being absorbed and then leads to the low photocatalytic activity. Thus, it is urgent to find a strategy to extend the light harvesting of SnO<sub>2</sub> MSCs. Much effort has been focused on enhancing the photoabsorption and the charge separation rate of SnO<sub>2</sub> by methods such as element doping,<sup>12,13</sup> surface modification,<sup>14,15</sup> and heterojunction design.<sup>16–18</sup> Note that coupling of two materials as a semiconductor heterojunction could substantially improve the photocatalytic activity by broadening the absorption range of light and increasing the separation efficiency of the photo-generated electrons and holes.<sup>19–23</sup> Zhang *et al.*<sup>24</sup> used a one-step fabrication to synthesize an Ag<sub>3</sub>PO<sub>4</sub>/SnO<sub>2</sub> composite, which showed improved photocatalytic activity and structural stability for photodecomposition of organic compounds. Zheng *et al.*<sup>25</sup> synthesized a network structured SnO<sub>2</sub>/ZnO nanocatalyst by a two-step solvothermal method. The photocatalytic activity of

SnO<sub>2</sub>/ZnO for the degradation of methyl orange was much higher than that of the pure SnO<sub>2</sub> and ZnO. Zhang *et al.*<sup>26</sup> used a combination method of sol–gel process and electrospinning technique to synthesize the one-dimensional ZnO–SnO<sub>2</sub> nanofibers, which exhibited excellent photocatalytic activity, superior to the electrospun pure ZnO and SnO<sub>2</sub> nanoparticles. According to these published studies, the heterojunction structure of two semiconductors will open an efficient pathway for the improvement of photocatalytic activity and stability.

Recently, conjugated polymer polydiphenylbutadiyne (PDPB) nanofibers<sup>27</sup> attracted wide attention due to their high visible light response, stable cycling property and low cost. They have a narrow band gap of 1.81 eV, which is attributed to the octamer structure. The electrons could be generated by the conjugated polymer chains when illuminated by energetic photons, but PDPB nanofibers have a highly hydrophobic feature that limits their photocatalytic performance in water. In order to overcome these drawbacks, we synthesized a p–n heterojunction structure of SnO<sub>2</sub> and PDPB. In this heterojunction structure system, PDPB nanofibers could enlarge the spectral absorption range, whereas SnO<sub>2</sub> MSCs could provide a stable mesoporous structure and a hydrophilic environment.

Herein, we employed a facile and green *in situ* strategy to synthesize a PDPB/SnO<sub>2</sub> heterojunction. This heterojunction has advantages of both single crystals and polymers, such as high photosensitivity, stable mesoporous structure and good hydrophilicity. The PDPB nanofibers were adsorbed and grown on the surface of the SnO<sub>2</sub> MSCs, and the Sn–C bonds were generated between them to realize the carbon doping on SnO<sub>2</sub>. When the heterojunction surface was illuminated, PDPB

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nanofibers absorbed the light and generated electrons. Moreover, carbon-doped SnO<sub>2</sub> (C-SnO<sub>2</sub>) MSCs offered charge transfer channels, which endowed the PDPB/SnO<sub>2</sub> composites with a high photocatalytic activity. The photodegradation rate of Rhodamine B (RhB) over the PDPB/SnO<sub>2</sub> composite (ratio is 2 : 1) is 4 times that of the single SnO<sub>2</sub> MSCs and 2 times that of single PDPB nanofibers.

## 2. Experimental

### 2.1 Materials

Tetraethyl orthosilicate (TEOS), aqueous solution of ammonia (NH<sub>3</sub>·H<sub>2</sub>O, 25 wt%), ethanol, distilled water, stannic chloride (SnCl<sub>4</sub>·5H<sub>2</sub>O), polyethylene pyrrole (PVP), hydrochloric acid (HCl), sodium hydroxide (NaOH), sodium chloride, sodium dodecyl sulphate, 1,4-diphenylbutadiene (DPB), benzoin methyl ether (BME), cyclohexane, and pentanol-1 were used. All chemicals were of analytical grade and used without further purification.

### 2.2 Preparation of photocatalysts

**Preparation of SiO<sub>2</sub> spheres.** The silica spheres were synthesized by the Stöber method. First, 167.2 mL ethanol was mixed with 28.8 mL water in a beaker. The mixture was stirred for 1 min at room temperature. Simultaneously, 18.0 mL tetraethoxysilane (TEOS) was mixed with 182.0 mL absolute ethanol at room temperature. The TEOS solution was stirred for 1 min and poured into the abovementioned mixture. After that, 4.0 mL ammonium hydroxide solution was added into the mixture. The precursor liquid was stirred for 24 h to obtain a suspension (silica colloids). Then, it was centrifuged at 12 000 rpm for 1 h and dried in a vacuum oven at 60 °C for 12 h.

**Preparation of SnO<sub>2</sub> MSCs.** The SnO<sub>2</sub> MSCs were prepared by a modifying the method developed by Zheng *et al.*<sup>28</sup> Herein, 12 mL H<sub>2</sub>O, 12 mL ethanol and 0.1 g PVP were added in the autoclave (50 mL volume). After being completely dissolved, 1.346 g SnCl<sub>4</sub>·5H<sub>2</sub>O was added into the autoclave. Then, the mixture was stirred at room temperature for 10 min, and 1.8 mL hydrochloric acid (HCl, 37%) was added and stirred for another 10 min. Then, 0.8 g SiO<sub>2</sub> was added, and the sealed vessel was heated at 200 °C for 12 h. Finally, the silica template was removed by etching in 2 M NaOH at 80 °C for 1 h. The remaining products were collected by centrifugation and washed several times with H<sub>2</sub>O and ethanol.

**Preparation of PDPB/SnO<sub>2</sub> MSC heterojunction.** The *in situ* route to PDPB/SnO<sub>2</sub> MSCs was prepared by modifying the method developed by Ghosh *et al.*<sup>27</sup> Herein, 0.035 g NaCl was dissolved in 2 mL H<sub>2</sub>O, and 1 g of a surfactant (sodium dodecyl sulphate) was dissolved in the mixture solution. After vigorous agitation for 5 h, the surfactant had completely dissolved to form a transparent and viscous micellar solution. After that, we weighed 0.1 g DPB and 0.01 g BME into a beaker and added 5 mL cyclohexane, followed by ultrasound treatment of the solution for complete dissolution. This solution was added to the abovementioned micellar solution under stirring, which led to a white unstable emulsion. After reaction for 0.5 h, a co-

surfactant, pentanol-1 (420 μL) was added to the mixture and stirred for 3 h. Then, different masses of SnO<sub>2</sub> MSCs were added into the mixture. The glass tube was illuminated using a 300 W xenon lamp at a distance of 5 cm for 12 h. Then, the products were extracted into a 15 mL water and 15 mL ethanol solution and dried in a vacuum oven at 60 °C for 12 h.

### 2.3 Characterizations

The crystal structure of the sample was detected by X-ray diffraction (XRD) measurements (Rigaku Ultima IV (Cu Kα radiation, λ = 1.5406 Å) in the range of 10–80° (2θ)). The morphologies of the PDPB/SnO<sub>2</sub> heterojunctions were characterized by transmission electron microscopy (TEM, JEM2100) and field emission scanning electron microscopy (FESEM, HITACHI, S4800). BET specific surface area measurements were studied by N<sub>2</sub> adsorption at 77 K using an ASAP2020 instrument. The X-ray photoelectron spectroscopy (XPS) studies were carried out using a Perkin-Elmer PHI 5000C ESCA system with Al Kα radiation. The binding energy shift was referenced to the C1s level at 284.6 eV as an internal standard. Infrared spectra (IR) were recorded with KBr disks containing the powder sample with an FTIR spectrometer (Nicolet Magna 550). Thermogravimetric and differential thermal analyses were conducted on a Pyris Diamond TG/DTA (PerkinElmer) apparatus at a heating rate of 20 K min<sup>-1</sup> from 40 °C to 800 °C in air flow. UV-vis diffuse reflectance spectra (DRS) were obtained with a SHIMADZU UV-2600 spectroscope equipped with an integrating sphere assembly and using BaSO<sub>4</sub> as the reflectance sample. Photocurrent measurements were carried out on an electrochemical analyzer (CHI 660 D electrochemical station, CHI Instruments Inc.) at room temperature. Transient photocurrent responses of different samples were carried out in 0.5 M Na<sub>2</sub>SO<sub>4</sub> aqueous solution under various irradiation conditions (300 W Xe lamp).

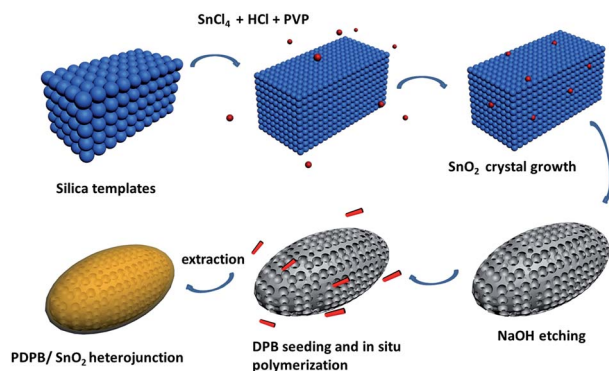
### 2.4 Measurements of photocatalytic activity

The photocatalytic activities of the PDPB/SnO<sub>2</sub> MSCs have been evaluated for the photodecomposition of RhB in water. The photocatalyst (0.01 g) was added into a 100 mL quartz photo-reactor containing 50 mL of RhB solution (10 mg L<sup>-1</sup>). The suspension containing PDPB/SnO<sub>2</sub> with pollutants was stirred in the dark for 1 h to establish adsorption-desorption equilibrium before irradiation. A 300 W Xe lamp with an AM1.5 air mass filter was used as a simulated solar light source. Then, the solutions were irradiated for 3 h and quantified every 30 min. The degradation efficiency of RhB was detected by ultraviolet-visible spectrophotometry.

## 3. Results and discussion

The overall synthesis of the PDPB/SnO<sub>2</sub> MSC heterojunction consisted of two steps (Scheme 1). First, we used a hard template method to synthesize the SnO<sub>2</sub> MSCs by a solvothermal process. SnO<sub>2</sub> was generated in the gap of the silica template, and the silica was etched by NaOH to form the mesoporous structure of SnO<sub>2</sub>. Second, the SnO<sub>2</sub> MSCs were added into the soft template hexagonal mesophases, which contained





Scheme 1 Synthetic steps of SnO<sub>2</sub> MSCs and PDPB/SnO<sub>2</sub> MSC heterojunction.

DPB and BME. After that, the DPB would grow and adsorb on the surface of SnO<sub>2</sub> MSCs under light irradiation. We synthesized the PDPB/SnO<sub>2</sub> heterojunction with different weight ratios (PDPB : SnO<sub>2</sub> = 4 : 1, 2 : 1, 1.33 : 1, 1 : 1).

### 3.1 Heterojunction structure

The crystal structures of the PDPB/SnO<sub>2</sub> MSCs were examined by X-ray diffraction (XRD). Fig. 1 shows the XRD patterns of pure SnO<sub>2</sub> MSCs, pure PDPB nanofibers and PDPB/SnO<sub>2</sub> composites. The SnO<sub>2</sub> MSCs showed a rutile phase of tetragonal SnO<sub>2</sub>. The sharp peaks are observed at  $2\theta = 26.58^\circ, 33.9^\circ, 37.92^\circ, 51.76^\circ, 54.7^\circ, 57.9^\circ, 61.86^\circ, 64.78^\circ$  and  $65.92^\circ$ , which could be indexed as the (110), (101), (200), (211), (220), (002), (310), (312) and (301) diffraction planes, respectively. The characteristic peaks of PDPB nanofibers also appeared in XRD patterns. There is no evident peak ascribed to the PDPB over the PDPB/SnO<sub>2</sub> composite, which is due to the low loading content of PDPB on the surface of SnO<sub>2</sub>.

### 3.2 Morphology and surface area

The morphologies of the SnO<sub>2</sub> MSCs and PDPB/SnO<sub>2</sub> composites were analyzed by TEM and FESEM, and the results are shown in Fig. 2. From Fig. 2a and c, we can find the obvious, well-defined, and uniform mesoporous structure of the shuttle shape SnO<sub>2</sub> from the TEM and FESEM images. The length and

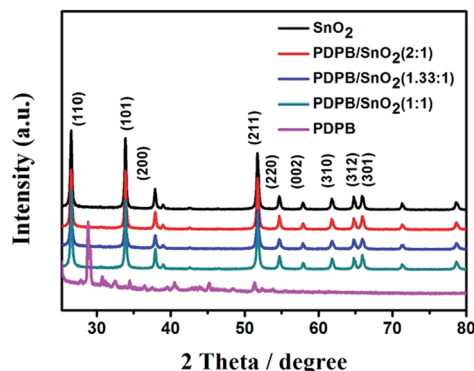


Fig. 1 XRD patterns of pure PDPB, pure SnO<sub>2</sub> and PDPB/SnO<sub>2</sub> composites.

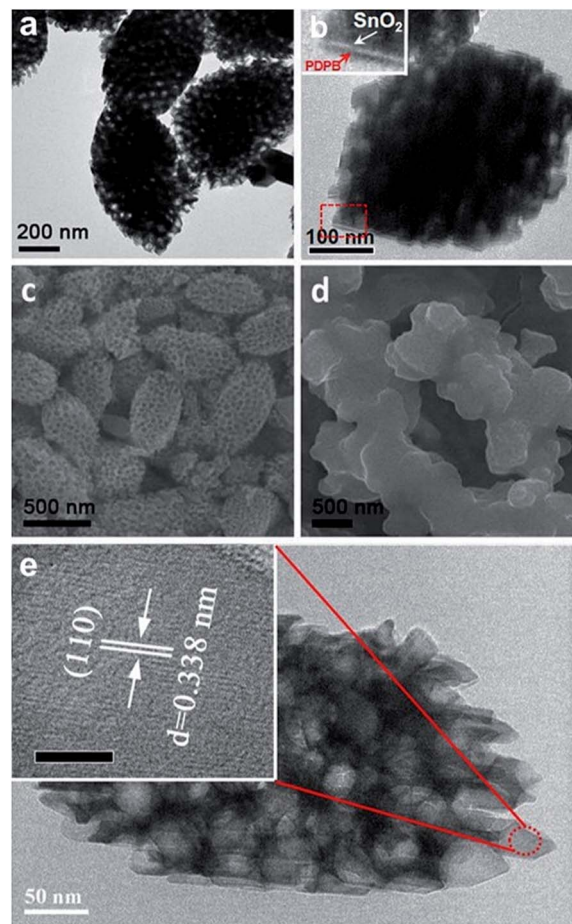


Fig. 2 Morphology characterization results of as-prepared samples. TEM images of pure SnO<sub>2</sub> MSCs (a) and PDPB/SnO<sub>2</sub> composites (2 : 1) (inset is the amplification of the red circle) (b); FESEM images of pure SnO<sub>2</sub> MSCs (c) and PDPB/SnO<sub>2</sub> composites (2 : 1) (d); HRTEM image of SnO<sub>2</sub> MSCs (e).

width of the SnO<sub>2</sub> MSCs are  $\sim 400$  nm and  $\sim 200$  nm, respectively, and the pore size is uniform at  $\sim 50$  nm. The HRTEM images of SnO<sub>2</sub> MSCs (Fig. 2e) show that the uniform lattice distance is 0.338 nm, which indicates the (110) plane of SnO<sub>2</sub>. Compared to pure SnO<sub>2</sub> MSCs, Fig. 2b and d demonstrate that the pore structure of PDPB/SnO<sub>2</sub> becomes hazy, which is attributed to the *in situ* growth of PDPB on the surface of SnO<sub>2</sub> MCMs. This speculation is confirmed by the BET results. The BET curve is shown in Fig. 3; the hysteresis loop and the pore size distribution curve could prove that the PDPB does not enter into the pore structure of SnO<sub>2</sub> MSCs. The BET surface area of pure SnO<sub>2</sub> MSCs and PDPB/SnO<sub>2</sub> is  $23.2 \text{ m}^2 \text{ g}^{-1}$  and  $31.2 \text{ m}^2 \text{ g}^{-1}$ , respectively, indicating that the PDPB was covered on the surface rather than in the pores of SnO<sub>2</sub>. The surface area of PDPB/SnO<sub>2</sub> is a little larger compared to the pure SnO<sub>2</sub>, which also implies that the loading of PDPB could not change the original mesoporous structure of SnO<sub>2</sub>.

### 3.3 XPS characterization

The surface composition and chemical states of SnO<sub>2</sub> MSCs and PDPB/SnO<sub>2</sub> composites were investigated by XPS spectra. Fig. 4



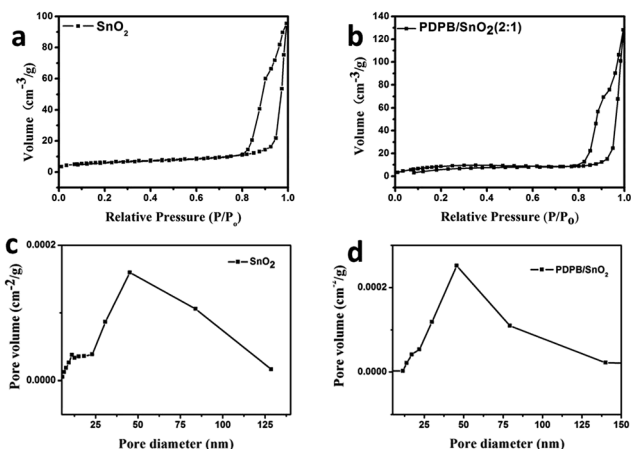


Fig. 3 Nitrogen adsorption/desorption isotherm of SnO<sub>2</sub> MSCs (a) and PDPB/SnO<sub>2</sub> composites (b); pore size distributions of SnO<sub>2</sub> MSCs (c) and PDPB/SnO<sub>2</sub> composites (d).

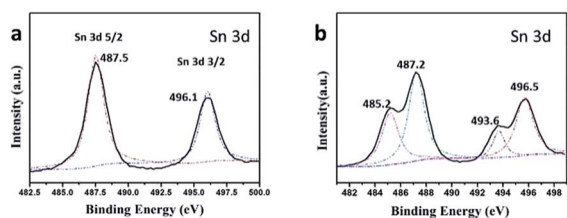


Fig. 4 Sn 3d XPS spectra of pure SnO<sub>2</sub> (a) and PDPB/SnO<sub>2</sub> composites (2 : 1) (b).

shows the Sn 3d XPS spectrum of SnO<sub>2</sub> MSCs and PDPB/SnO<sub>2</sub> composites. The peaks at 487.5 eV and 496.1 eV belong to the Sn 3d<sub>5/2</sub> and Sn 3d<sub>3/2</sub>, respectively, which are the characteristic peaks of Sn(IV). Interestingly, compared to pure SnO<sub>2</sub>, the Sn 3d spectrum of PDPB/SnO<sub>2</sub> has two extra peaks at 485.2 eV and 493.6 eV. These two peaks indicate the changes of chemical states of Sn element, which may be caused by the combination between SnO<sub>2</sub> and PDPB. A similar phenomenon was also reported in previous literature.<sup>29</sup> We inferred that this result is due to the combination between SnO<sub>2</sub> and PDPB and it may generate a new impurity level with the Sn–C bond. C replaced O from SnO<sub>2</sub> and acted as the electron donor to enlarge the electron cloud density of Sn element, which led to the appearance of new XPS peaks at a lower binding energy.

### 3.4 IR and TGA analysis

In order to confirm the existence of PDPB and find the interaction between SnO<sub>2</sub> and PDPB, IR spectra were obtained, as shown in Fig. 5. The pure PDPB shows the characteristic peaks at 2850 cm<sup>-1</sup>, 2919 cm<sup>-1</sup> and 3048 cm<sup>-1</sup>, which can be attributed to the C–H vibration of the polymer molecule chains. There are no significant peaks of pure SnO<sub>2</sub> in this region. Compared to the pure SnO<sub>2</sub> and PDPB, the characteristic peaks at 2850 cm<sup>-1</sup> and 2919 cm<sup>-1</sup> are consistent with the pure PDPB, but the 3048 cm<sup>-1</sup> peak is too weak to locate and could be induced by the chemical bonding between SnO<sub>2</sub> and PDPB. The XPS results

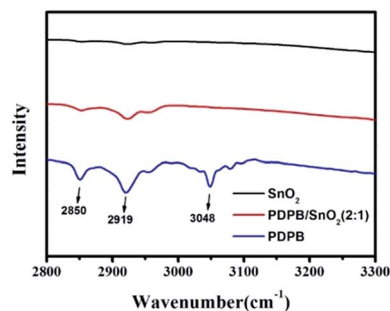


Fig. 5 IR spectra of pure SnO<sub>2</sub>, pure PDPB and PDPB/SnO<sub>2</sub> composites (2 : 1).

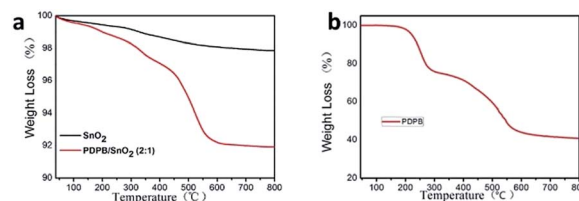


Fig. 6 (a) TGA spectra of SnO<sub>2</sub> and PDPB/SnO<sub>2</sub> composites (2 : 1). (b) TGA spectra of PDPB nanofibers.

indicate the generation of Sn–C bonds between polymers and semiconductors, which could be formed by the dehydration between C–H bonds and Sn–OH bonds.

The loading content of PDPB in the heterojunction was examined by the thermogravimetric analysis (TGA). Fig. 6a demonstrates that the TGA curve of pure SnO<sub>2</sub> is like a horizontal line, indicating the stable property of SnO<sub>2</sub>. The curve of the PDPB/SnO<sub>2</sub> heterojunction shows a weight loss at 160 °C. The loss process has two steps: (1) from 160 °C to 250 °C, it indicates the thermal decomposition of polymer chains; (2) from 250 °C to 450 °C, it indicates the decomposition of carbons. From the TGA results, it is observed that PDPB mass in the heterojunction is about 8% (2 : 1 ratio), and the interaction between PDPB and SnO<sub>2</sub> causes the heterojunction (160 °C) to have a weight-loss temperature lower than that of the pure PDPB (200 °C) (Fig. 6b).

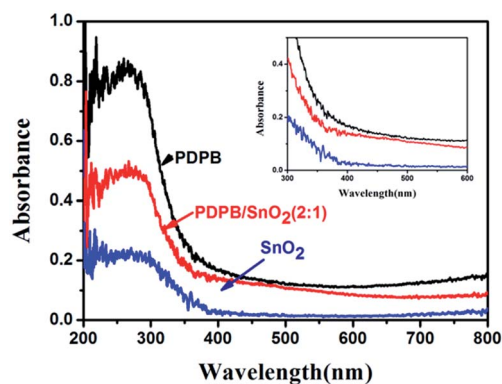


Fig. 7 UV-vis diffuse reflectance spectra of pure PDPB, pure SnO<sub>2</sub> MSCs and PDPB/SnO<sub>2</sub> composites (2 : 1).



### 3.5 UV-vis diffuse reflectance spectra

The absorbance of the pure SnO<sub>2</sub> MSCs, pure PDPB and PDPB/SnO<sub>2</sub> composites were measured by UV-vis diffuse reflectance spectra (Fig. 7). The pure PDPB shows a high absorption from 200–700 nm wavelength, which shows the excellent light harvest. The PDPB/SnO<sub>2</sub> composites present a lower photo response than pure PDPB due to the low loading amounts of PDPB in the composites. It is clear that the PDPB/SnO<sub>2</sub> composite has a higher absorption value than the pure SnO<sub>2</sub> MSCs from 200–700 nm wavelengths. Particularly in the ultraviolet region, the absorption value over heterojunction has a remarkable increase than the pure SnO<sub>2</sub> MSCs, which indicates that PDPB could enhance the light absorption of SnO<sub>2</sub>. From the UV-vis results, we can demonstrate that PDPB could be used as the sensitizer to enlarge the solar light response range of SnO<sub>2</sub>, which is expected to improve the photocatalytic activity of SnO<sub>2</sub> under solar light.

### 3.6 Photochemical property measurement

In order to compare the photochemical properties of pure SnO<sub>2</sub> and PDPB/SnO<sub>2</sub> composites, we used the transient photocurrent response measurements to characterize the separation efficiency of photogenerated electrons and holes. The photocurrents of these samples were recorded in the dark or under solar light illumination using the xenon lamp assembled with an AM1.5 optical filter as the light source, as shown in Fig. 8. The results show that the photocurrent of PDPB/SnO<sub>2</sub> is much higher than that of pure SnO<sub>2</sub> MSCs. For PDPB/SnO<sub>2</sub> composites, the photocurrent response was 4 μA cm<sup>-2</sup>, which is about 4 times that of the SnO<sub>2</sub> MSCs (below 1 μA cm<sup>-2</sup>). As we know, carbon-doped SnO<sub>2</sub> is the n-type semiconductor and PDPB is the p-type polymer semiconductor.<sup>30</sup> Owing to the p–n heterojunction structure of PDPB and SnO<sub>2</sub> MSCs, the composites display a higher transfer rate of photogenerated electron–hole pairs, which confirms the lower recombination rate for the PDPB/SnO<sub>2</sub> composites. When solar light is irradiated on the PDPB/SnO<sub>2</sub> heterojunction, the photo-generated electrons can be easily excited from the PDPB valence band (VB) to the conduction band (CB) and subsequently injected to the CB of SnO<sub>2</sub> MSCs. Moreover, the hole still remains in the valence band of PDPB. Then, the electrons and holes further react with O<sub>2</sub> or H<sub>2</sub>O to generate the superoxide radical and hydroxyl radicals, and these radicals can oxidize the RhB and degrade it,

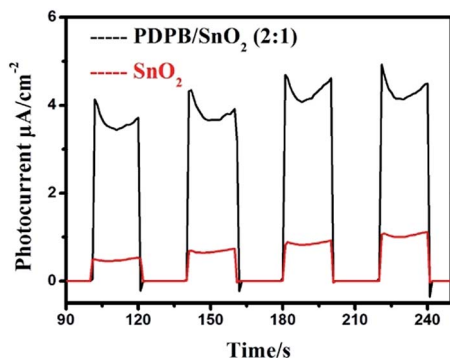


Fig. 8 Photocurrent spectra of pure SnO<sub>2</sub> MSCs (red) and PDPB/SnO<sub>2</sub> composites (black) (2 : 1).

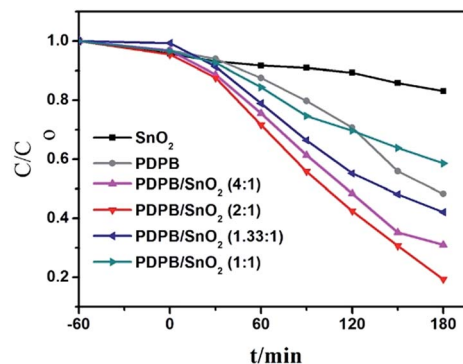


Fig. 9 Photodegradation of RhB of pure PDPB, pure SnO<sub>2</sub> and PDPB/SnO<sub>2</sub> composites under the solar light irradiation (Xe lamp with AM1.5 filter).

which is in accordance with the highly photocatalytic activity of PDPB/SnO<sub>2</sub>.

### 3.7 Photocatalytic activity measurement

The photocatalytic activity of the as-prepared samples was evaluated by measuring the decomposition rate of the model pollutants of Rhodamine B (RhB) in water under solar light (Xe lamp with AM1.5 filter). The photocatalytic activities at different ratios of PDPB/SnO<sub>2</sub> were compared with the activities of pure PDPB and pure SnO<sub>2</sub> MSCs. Fig. 9 shows the degradation rate of all samples, and it is clearly shown that all PDPB/SnO<sub>2</sub> composites presented an evidently enhanced photocatalytic activity than pure SnO<sub>2</sub> MSCs. The samples with 4 : 1, 2 : 1, and 1.33 : 1 ratios also showed a photocatalytic activity higher than the pure PDPB. The SnO<sub>2</sub> MSCs with a wide band gap showed a low solar light activity to degrade RhB. The p–n heterojunction showed a much higher photocatalytic activity that is caused due to the sensitization of PDPB and the narrowed band gap of SnO<sub>2</sub> resulting due to carbon doping. The PDPB/SnO<sub>2</sub> showed a higher activity than pure PDPB owing to the synergistic effect of the p–n heterojunction and PDPB sensitization. That is, the photo-generated electrons and holes have an efficient separation at the interface between PDPB and the SnO<sub>2</sub> MSCs. The 2 : 1 ratio sample had the highest photocatalytic activity, whose degradation rate of RhB is 4 times and 2 times that of pure SnO<sub>2</sub> MSCs and pure PDPB, respectively. The sample with a 1 : 1 ratio showed a relatively lower activity, which indicates that the extra PDPB covered on the surface of SnO<sub>2</sub> would decrease the light absorption of C–SnO<sub>2</sub>. Therefore, the ratio of PDPB and SnO<sub>2</sub> MSCs has an optimum value. In addition, the PDPB amount in the PDPB/SnO<sub>2</sub> (1 : 1) composite is relatively low (lower than 8%), and hence the main body of the PDPB/SnO<sub>2</sub> composite is SnO<sub>2</sub>. Thus, the composite with 1 : 1 ratio shows a lower activity than PDPB. When the amount of PDPB is too high, it generates more recombination centers of electrons and holes. This indicates that the ratio of PDPB and SnO<sub>2</sub> has an optimal value.

### 3.8 Mechanism

The proposed mechanism of the PDPB/SnO<sub>2</sub> heterojunction is illustrated in Fig. 10. In this heterojunction structure, n-type SnO<sub>2</sub> MSCs provided a stable mesoporous structure,



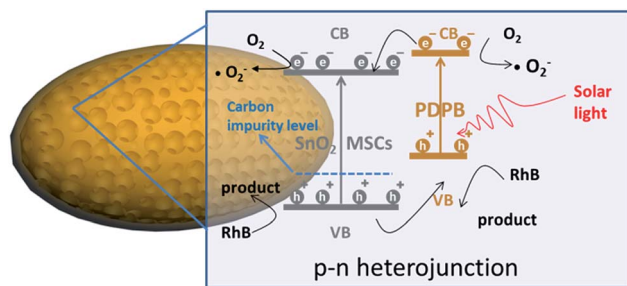


Fig. 10 Mechanism of the PDPB/SnO<sub>2</sub> composite degradation of the RhB under solar light.

a hydrophilic environment, a high electron mobility, and exposed facets with active sites, which could expedite the electron transfer rate. The p-type PDPB on the surface of the heterojunction acted as a sensitizer to improve the solar light absorption.<sup>31</sup> The generation of Sn–C bonds between SnO<sub>2</sub> and PDPB introduced a carbon impurity level above the VB of SnO<sub>2</sub>, which led to the solar light response of SnO<sub>2</sub>. The construction of a p–n heterojunction provided an ohmic interface to enhance the separation of photogenerated electrons and holes.

## 4. Conclusion

In summary, we used an *in situ* strategy to synthesize a PDPB/SnO<sub>2</sub> p–n heterojunction that exhibited a high photocatalytic activity for the degradation of RhB under solar light irradiation owing to its excellent light harvesting and effective separation of electrons and holes. The chemical bonds of Sn–C generated on the interface of heterojunctions could decrease the bandgap of SnO<sub>2</sub> and enhance the interaction between the two components. As a result, the as-prepared catalysts had a high solar-driven photocatalytic performance. Our study gives a facile strategy for the preparation of an organic–inorganic hybrid heterojunction with high solar light activity.

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