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# Extended visible photosensitivity of carboxyethyltin functionalized polyoxometalates with common organic dyes enabling enhanced photoelectric performance†

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A kind of broad spectral responsive photoelectrode has been assembled from open-chain carboxyethyltin functionalized Dawson sandwich-type polyoxometalates (POMs)  $[(\text{Sn}(\text{CH}_2)_2\text{COO})_2\text{M}_2(\text{P}_2\text{W}_{15}\text{O}_{56})_2]^{16-}$  (SnR-M-P<sub>2</sub>W<sub>15</sub>, M = Mn, Co, Cu and Zn) and organic dyes such as rhodamine B (RhB) and methylene blue (MB) through the layer-by-layer (LBL) self-assembly method. SnR-M-P<sub>2</sub>W<sub>15</sub> shows strong adsorption towards RhB and MB. The adsorption capacity of SnR-Co-P<sub>2</sub>W<sub>15</sub> towards RhB and MB reached 217.36 and 288.76 mg g<sup>-1</sup> in 10 min, respectively. Compared to Keggin sandwich-type open-chain carboxyethyltin-POM and closed-ring carboxyethyltin-POM derivatives such as  $[\text{Zn}_3(\text{Sn}(\text{CH}_2)_2\text{COO})_2(\text{PW}_9\text{O}_{34})_2]^{10-}$  (SnR-Zn-PW<sub>9</sub>) and  $[\text{K}_3(\text{Sn}(\text{CH}_2)_2\text{COO})_2(\text{PW}_9\text{O}_{34})_2]^{11-}$  (SnR-PW<sub>9</sub>), and the parent POMs  $[\text{M}_4(\text{H}_2\text{O})_2(\text{P}_2\text{W}_{15}\text{O}_{56})_2]^{16-}$  (M-P<sub>2</sub>W<sub>15</sub>, M = Mn, Co, Cu and Zn), SnR-M-P<sub>2</sub>W<sub>15</sub>-based photoelectrodes showed significantly enhanced photocurrent owing to their greater visible light harvesting. Also, the photoelectric performances of LBL multilayer films can be adjusted by controlling the number of deposited layers, adsorbed dye types and incorporated transition metal ions in the POMs.

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## Introduction

The development of high-efficiency solar cells represents a promising and desirable approach for the sustainable utilization of clean energy.<sup>1–7</sup> Typically, for achieving high photo-voltaic performances, the photoelectrodes must perform efficient light harvesting with absorption bands in the visible or near-IR region, preferably covering a broad range of wave-lengths, and have sufficient electrochemical driving force for solar conversion reactions, slow carrier recombination, stability in the electrolyte, and low cost.<sup>8–10</sup> Recently, polyoxometalates (POMs), which represent a diverse class of low cost metal-oxide clusters with tunable redox properties and high stability, have been found to be versatile inorganic building blocks for constructing photoelectrochemical and dye-sensitized solar cells.<sup>11–20</sup> Nevertheless, traditional POMs absorb mainly UV light, which limits their utilization of solar energy.<sup>21,22</sup> Thus, further boosting their photoelectrochemical performances in

the visible light region still remains a challenge for POM-based photovoltaic systems.

Carboxyethyltin functionalized POMs are a typical series of POM-based derivatives, which have attracted considerable interest recently due to their photoelectrochemical applications. Especially for double-armed type POMs that were co-functionalized by open-chain carboxyethyltin and transitional metals, in which the transitional metal can bring extra redox centers and enhance visible light photo-response, while the exposed –COO group plays a key role for building POM-based composite systems by further functionalization.<sup>23–26</sup> Our group has synthesized a series of carboxyethyltin-POM derivatives and further studied their photoelectric and electro-catalytic performances by loading them on semi-conductors<sup>24,25</sup> and carbon nanotubes.<sup>26</sup> But it should be noted that their utilization of solar energy is still limited and further functionalization for enhanced visible activity is needed. Liu and Zhang *et al.* reported the visible-light-driven POM-based photoelectrochemical cells by organic dye-functionalization.<sup>15</sup> As is known, rhodamine B (RhB) and methylene blue (MB) are of two dye contaminants in wastewater, which show intense absorbance in the visible region.<sup>27,28</sup> Hence, the composition of carboxyethyltin-POM derivatives with these dyes are expected to form new types of photoelectric materials, and also to be potential for decolorizing treatment of dye wastewater.

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In this work, open-chain carboxyethyltin functionalized Dawson sandwich-type tungstophosphates ( $\text{SnR-M-P}_2\text{W}_{15}$ ,  $\text{M} = \text{Mn, Co, Cu}$  and  $\text{Zn}$ ) were discovered to show strong adsorption capacity towards RhB or MB, which can form a kind of broad spectral responsive composite material  $\text{SnR-M-P}_2\text{W}_{15}/\text{dye}$  ( $\text{dye} = \text{RhB}$  or  $\text{MB}$ ). Furthermore, because the layer-by-layer (LBL) assembly is a simple and effective way to prepare films with controllable thicknesses, architectures and functionalities on the molecular level,<sup>29–32</sup> the LBL growth of  $\text{SnR-M-P}_2\text{W}_{15}$  and RhB or MB composite photoelectrodes have been carried out successfully. The self-assembly procedure of the films was illustrated in Scheme 1. Upon visible light irradiation ( $\lambda > 420$  nm), the photoelectrode ( $\text{SnR-M-P}_2\text{W}_{15}/\text{RhB}$ )<sub>n</sub> or ( $\text{SnR-M-P}_2\text{W}_{15}/\text{MB}$ )<sub>n</sub> ( $n = 1-6$ ) shows obvious anodic photocurrent owing to both the extended visible light harvesting by organic dyes and semiconductor-like characteristics of  $\text{SnR-M-P}_2\text{W}_{15}$ . Additionally, a series of study results demonstrate that the photoelectric performances of the LBL films are different depending on the carboxyethyltin structure, POM structure and incorporated transitional metal types. This work provides valuable insights for POM-based photovoltaic systems on further utilization of solar energy and environment protection.

## Experimental

### Chemicals and reagents

$\text{Na}_{12}[\alpha\text{-P}_2\text{W}_{15}\text{O}_{56}] \cdot 24\text{H}_2\text{O}$  ( $\text{P}_2\text{W}_{15}$ ) and  $\text{Na}_{16}[\text{M}_4(\text{H}_2\text{O})_2(\text{P}_2\text{W}_{15}\text{O}_{56})_2] \cdot n\text{H}_2\text{O}$  ( $\text{M-P}_2\text{W}_{15}$ ) were prepared according to the reported procedures.<sup>33–36</sup>  $\text{Cl}_3\text{SnCH}_2\text{CH}_2\text{COOCH}_3$  ( $\text{SnR}$ ) and  $\text{Cl}_3\text{SnCH}_2\text{CH}_2\text{COOH}$  ( $\text{SnR-COOH}$ ) were synthesized by the literature methods.<sup>37</sup>  $\{\text{C}(\text{NH}_2)_3\}_{12}\text{H}_4[\{\text{Sn}(\text{C}_3\text{H}_4\text{O}_2)_2\text{M}_2(\text{P}_2\text{W}_{15}\text{O}_{56})_2\}] \cdot n\text{H}_2\text{O}$  ( $\text{SnR-M-P}_2\text{W}_{15}$ ,  $\text{M} = \text{Mn}$ ,  $n = 22$ ;  $\text{M} = \text{Co}$ ,  $n = 14$ ) have been previously prepared in our group.<sup>25,38</sup> Herein, we added transitional metal ion  $\text{Cu}^{2+}$  and  $\text{Zn}^{2+}$  into the array of POMs, and complete details of this case will be provided at a later time. Together, these four representative POMs with the general abbreviative formula of  $\text{SnR-M-P}_2\text{W}_{15}$  ( $\text{M} = \text{Mn, Co, Cu, Zn}$ ) represent the Dawson sandwich-type open-chain carboxyethyltin functionalized POMs. We also prepared the Keggin sandwich-type open-chain carboxyethyltin functionalized POM

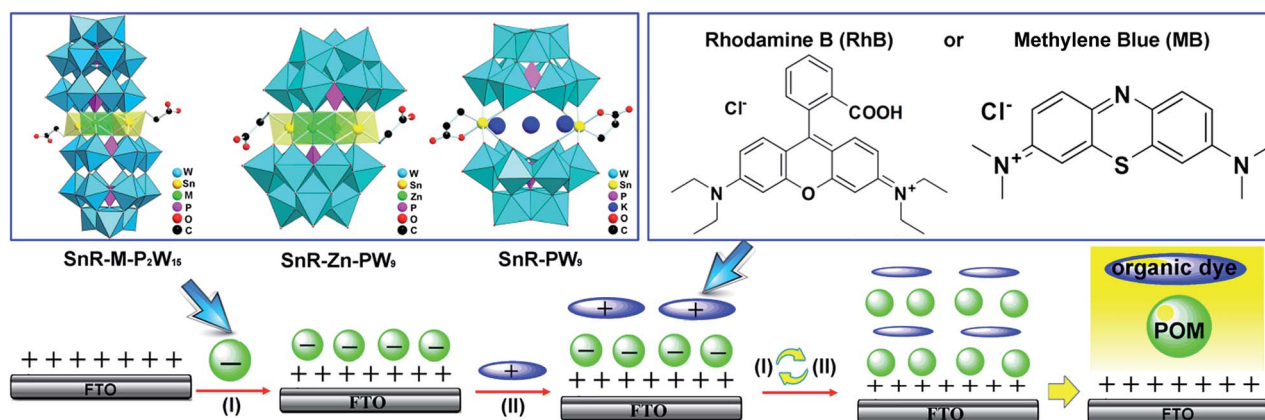
$\text{Na}_8\text{K}_2[\text{Zn}_2\{\text{Sn}(\text{CH}_2)_2\text{COO}\}_2[(\text{PW}_9\text{O}_{34})_2] \cdot 13\text{H}_2\text{O}$  ( $\text{SnR-Zn-PW}_9$ )<sup>39</sup> and Keggin sandwich-type closed-ring carboxyethyltin functionalized POM  $\text{Na}_3\text{K}_8[\text{K}_3\{\text{Sn}(\text{CH}_2)_2\text{COO}\}_2(\text{PW}_9\text{O}_{34})_2] \cdot 18\text{H}_2\text{O}$  ( $\text{SnR-PW}_9$ )<sup>40</sup> for comparative experiments. All of them were characterized by IR spectroscopy. Poly(ethylenimine) (PEI; MW = 750 000) and poly(styrenesulfonate) (PSS; MW = 70 000) were purchased from Aldrich and used without further treatment. All other reagents and chemicals were commercially purchased and used without further purification.

### Synthesis of $\{\text{C}(\text{NH}_2)_3\}_{12}\text{H}_4[\{\text{Sn}(\text{CH}_2)_2\text{COO}\}_2\text{Cu}_2(\text{P}_2\text{W}_{15}\text{O}_{56})_2] \cdot 14\text{H}_2\text{O}$ ( $\text{SnR-Cu-P}_2\text{W}_{15}$ )

$\text{SnR-Cu-P}_2\text{W}_{15}$  was successfully synthesized by referring to our previous method<sup>25,38</sup> and making small variations.  $\text{SnR}$  (0.12 g, 0.38 mmol) was dissolved in 30 mL of 0.5 mol L<sup>-1</sup> NaCl aqueous solution under vigorous stirring. Then,  $\text{Cu-P}_2\text{W}_{15}$  (0.20 g, 0.02 mmol) was added in small portions to obtain a clear, light-green solution. The resulting mixture was further stirred for 3 h at 80–90 °C. After cooling to room temperature, a small amount of insoluble precipitate was filtered off, and the filtrate was dealt with KCl (0.25 g) and  $\text{C}(\text{NH}_2)_3\text{Cl}$  aqueous solution (1.0 mol L<sup>-1</sup>, 1.8 mL). Slow evaporation of the mixture solution at 40–60 °C resulted in light-green block-like crystals after about two weeks (yield 31% based on W). Anal. calcd for  $\text{C}_{18}\text{H}_{112}\text{O}_{130}\text{N}_{36}\text{Cu}_2\text{Sn}_2\text{P}_4\text{W}_{30}$  (%): C 2.42, H 1.27, N 5.65, P 1.39, Cu 1.43, Sn 2.66, W 61.85. Found (%): C 2.45, H 1.31, N 5.69, P 1.46, Cu 1.48, Sn 2.70, W 61.91. FTIR (solid KBr pellet  $\nu/\text{cm}^{-1}$ ): 3434(s), 3266(w), 3185(w), 2918(w), 2855(w), 1668(s), 1395(w), 1097(s), 954(m), 903 (w), 761(m), 624(w), 524(w), 425(w).

### Synthesis of $\{\text{C}(\text{NH}_2)_3\}_{11}\text{H}_5[\{\text{Sn}(\text{CH}_2)_2\text{COO}\}_2\text{Zn}_2(\text{P}_2\text{W}_{15}\text{O}_{56})_2] \cdot 13\text{H}_2\text{O}$ ( $\text{SnR-Zn-P}_2\text{W}_{15}$ )

Detailed synthesis of  $\text{SnR-Zn-P}_2\text{W}_{15}$  is similar to that of  $\text{SnR-Cu-P}_2\text{W}_{15}$ , except using  $\text{Zn-P}_2\text{W}_{15}$  (0.20 g, 0.02 mmol) to replace  $\text{Cu-P}_2\text{W}_{15}$ , and colorless block-like crystalline  $\text{SnR-Zn-P}_2\text{W}_{15}$  was isolated after two weeks (yield 34% based on W). Anal. calcd for  $\text{C}_{17}\text{H}_{105}\text{O}_{129}\text{N}_{33}\text{Zn}_2\text{Sn}_2\text{P}_4\text{W}_{30}$  (%): C 2.31, H 1.20, N 5.23, P 1.40, Zn 1.48, Sn 2.68, W 62.36. Found (%): C 2.38, H 1.24, N 5.27, P 1.46, Zn 1.52, Sn 2.72, W 62.44. FTIR (solid KBr pellet  $\nu/\text{cm}^{-1}$ ):



Scheme 1 Schematic depiction of the buildup of the LBL multilayer films, and the structural representations of POMs and RhB and MB.



3436(s), 3259(w), 3187(w), 2916(w), 2853(w), 1666(s), 1400(w), 1096(s), 952(m), 912(w), 760(m), 624(w), 526(w), 421(w).

### Dye adsorption experiments

SnR-M-P<sub>2</sub>W<sub>15</sub> (M = Mn, Co, Cu and Zn) (34 mg) was dissolved in water (5 mL), followed by adding tetrabutylammonium bromide (TBA) to balance the charges in POMs. The pH of solution was adjusted to 4.0 by using a certain amount of NaOH or HCl solution, to which RhB (2.0 g L<sup>-1</sup>, 5 mL) solution was mixed together and stirred vigorously. Instantly, vast pink precipitate emerges, which was centrifugally separated for further analysis. The adsorption of SnR-M-P<sub>2</sub>W<sub>15</sub> towards MB is similar to that of SnR-M-P<sub>2</sub>W<sub>15</sub> towards RhB, except using MB solution (2.0 g L<sup>-1</sup>) to replace RhB solution, and cyanous precipitate momentarily appeared, which was centrifugally separated for further analysis.

### LBL growth of (SnR-M-P<sub>2</sub>W<sub>15</sub>/dye)<sub>n</sub> films (dye = RhB or MB)

Various substrates including quartz plate, silicon wafers, CaF<sub>2</sub> and FTO glasses were ultrasonically cleaned with detergent solution, deionized water, acetone and ethanol successively. Then these substrates were processed with piranha solution for 30 min (ref. 41) except for CaF<sub>2</sub>. After being rinsed with deionized water and dried under a stream of nitrogen, the cleaned substrate was immersed in PEI solution (1.0 mg mL<sup>-1</sup>) for 30 min, followed by rinsing with deionized water and drying under nitrogen atmosphere. The PEI-coated substrates were then alternately immersed into SnR-M-P<sub>2</sub>W<sub>15</sub> solution (10 mL, 3.69 mmol L<sup>-1</sup>) for 30 min and RhB solution (10 mL, 4.18 mmol L<sup>-1</sup>) for 30 min, and also rinsed with deionized water and dried under nitrogen after each dipping step. Subsequently, this process was repeated for certain times to obtain (SnR-M-P<sub>2</sub>W<sub>15</sub>/RhB)<sub>n</sub> films. Besides, (SnR-M-P<sub>2</sub>W<sub>15</sub>/MB)<sub>n</sub> films were fabricated with the same way except using MB instead of RhB.

### LBL growth of other composite films

The assembly of other composite films were similar to that of (SnR-M-P<sub>2</sub>W<sub>15</sub>/RhB)<sub>n</sub> films, except using SnR-Zn-PW<sub>9</sub>, SnR-PW<sub>9</sub>, M-P<sub>2</sub>W<sub>15</sub>, P<sub>2</sub>W<sub>15</sub> and PSS instead of SnR-M-P<sub>2</sub>W<sub>15</sub>, and more-over using PEI instead of RhB. They were marked as (SnR-Zn-PW<sub>9</sub>/RhB)<sub>n</sub>, (SnR-PW<sub>9</sub>/RhB)<sub>n</sub>, (M-P<sub>2</sub>W<sub>15</sub>/RhB)<sub>n</sub>, (P<sub>2</sub>W<sub>15</sub>/RhB)<sub>n</sub>, (PSS/RhB)<sub>n</sub> and (PEI/SnR-M-P<sub>2</sub>W<sub>15</sub>)<sub>n</sub>, respectively.

### Characterization methods

Single crystal X-ray diffraction data were collected on a Bruker Smart APEX II X-diffractometer equipped with graphite-monochromated Mo K $\alpha$  radiation ( $\lambda = 0.71073 \text{ \AA}$ ). C, H and N elemental analyses were performed on a VarioElcube elemental analyzer, and P, Sn, Cu, Zn and W were analyzed on a Prodigy XP emission spectrometer. IR spectra were recorded using KBr pellets on a Bruker AXS TENSOR-27 FTIR spectrometer in the range of 4000–400 cm<sup>-1</sup>. UV/Vis absorption spectra were recorded on a Lambda 35 UV/Vis spectrophotometer. TG analyses were performed on a Pyris Diamond TG-DTA thermal analyzer at the heating rate of 10 °C min<sup>-1</sup> in air atmosphere. X-

ray powder diffraction data were collected on a Bruker AXS D8 Advance diffractometer using Cu K $\alpha$  radiation ( $\lambda = 1.5418 \text{ \AA}$ ) in the  $2\theta$  range of 5–50° with a step size of 0.02°. NMR spectra were recorded at room temperature on a Bruker AVANCE 500 spectrometer with an inner tube containing D<sub>2</sub>O for instrumental lock. Phosphorus and tin chemical shifts were referenced to 85% H<sub>3</sub>PO<sub>4</sub> and (CH<sub>3</sub>)<sub>4</sub>Sn, respectively. The solid diffuse reflectivity spectra were collected on a Cary series UV/Vis spectrophotometer in reflectance mode, which was measured from 200 to 800 nm using BaSO<sub>4</sub> as a standard with 100% reflectance. The photoluminescence properties were determined on a RILI F-7000 fluorescence spectrophotometer in the solid state at room temperature. The X-ray photoelectron spectroscopy (XPS) spectra were recorded on a Thermo Electron ESCALAB-250 spectrometer with a monochromic X-ray source (Al K $\alpha$  line, 1486.6 eV). Atomic force microscopy (AFM) analysis were carried out using a Veeco Multimode Nanoscope IVA in tapping mode using silicon cantilevers (Bruker, RTESP) with radius of curvature less than 10 nm. The photoelectrochemical experiments were performed on a CHI604B electrochemical workstation. A three-electrode system was employed in a quartz cell with an Ag/AgCl electrode as the reference electrode, a platinum wire as the counter electrode and the hybrid films assembled on FTO electrode with an effective area of 1.0 × 1.0 cm<sup>2</sup> as the working electrode. A 300 W Xe lamp using the light filter of 420–780 nm was used as a light source, and the average intensity of irradiance reaching the composite films was measured to be ca. 100 mW cm<sup>-2</sup> with a Model ST-900M photometer. All photoelectrochemical measurements were done in 0.1 mol L<sup>-1</sup> Na<sub>2</sub>SO<sub>4</sub> electrolyte which was exposed to air.

### X-ray crystallography

The structures were solved by direct methods and refined on  $F^2$  by full-matrix least-squares fitting using SHELXTL-2014. Crystal data and structure refinement parameters of SnR-Cu-P<sub>2</sub>W<sub>15</sub> and SnR-Zn-P<sub>2</sub>W<sub>15</sub> are listed in Table 1. Hydrogen atoms were added on C/N atoms in calculated positions. In SnR-Cu-P<sub>2</sub>W<sub>15</sub> and SnR-Zn-P<sub>2</sub>W<sub>15</sub>, only some of lattice water molecules were accurately assigned from the residual electron peaks, whereas the rest were directly included in the molecular formula on the basis of the SQUEEZE calculation results, the elemental analysis and TG analysis. Selected bond lengths and angles are listed in Tables S1 and S2.† Hydrogen bonds are listed in Tables S3 and S4.†

## Results and discussion

### Crystal structure analysis

Single-crystal X-ray diffraction analysis reveals that SnR-Cu-P<sub>2</sub>W<sub>15</sub> and SnR-Zn-P<sub>2</sub>W<sub>15</sub> display the well-known Dawson sandwich-type structures, and the structures are isostructural with those of our previously reported compounds.<sup>25,33</sup> Two symmetrical units consisting of [P<sub>2</sub>W<sub>15</sub>O<sub>56</sub>]<sup>12-</sup> were subunit by two Cu<sup>2+</sup> or Zn<sup>2+</sup> ions and two carboxethyltin groups (Fig. S1†). In SnR-Cu-P<sub>2</sub>W<sub>15</sub> and SnR-Zn-P<sub>2</sub>W<sub>15</sub>, each [P<sub>2</sub>W<sub>15</sub>O<sub>56</sub>]<sup>12-</sup> unit supplies seven oxygen atoms (one central



**Table 1** Crystal and refinement data for SnR-Cu-P<sub>2</sub>W<sub>15</sub> and SnR-Zn-P<sub>2</sub>W<sub>15</sub>

Compound	SnR-Cu-P <sub>2</sub> W <sub>15</sub>	SnR-Zn-P <sub>2</sub> W <sub>15</sub>
Formula	C <sub>18</sub> H <sub>112</sub> O <sub>130</sub> N <sub>36</sub> <sup>-</sup> Cu <sub>2</sub> Sn <sub>2</sub> P <sub>4</sub> W <sub>30</sub>	C <sub>17</sub> H <sub>105</sub> O <sub>129</sub> N <sub>33</sub> <sup>-</sup> Zn <sub>2</sub> Sn <sub>2</sub> P <sub>4</sub> W <sub>30</sub>
Formula weight	8917.27	8843.83
<i>T</i> /K	296(2)	296(2)
Wavelength/Å	0.71073	0.71073
Crystal system	Monoclinic	Monoclinic
Space group	<i>P</i> 2(1)/ <i>n</i>	<i>P</i> 2(1)/ <i>n</i>
<i>a</i> /Å	13.3227(7)	13.3099(7)
<i>b</i> /Å	32.2012(16)	32.1189(18)
<i>c</i> /Å	19.0960(10)	18.9918(10)
$\beta$ /°	94.0540(10)	94.0490(10)
<i>V</i> /Å <sup>3</sup> , <i>Z</i>	8171.8(7), 2	8098.7(8), 2
<i>D</i> <sub>c</sub> /g cm <sup>-3</sup> , <i>F</i> <sub>000</sub>	3.624, 7900	3.627, 7820
GOF	1.015	1.038
Reflections	42 009	40 460
<i>R</i> <sub>int</sub>	0.0771	0.0606
$\theta$ range (°)	1.805–25.00	1.660–25.00
<i>R</i> <sub>1</sub> ( <i>I</i> > 2σ( <i>I</i> )) <sup>a</sup>	0.0501	0.0455
w <i>R</i> <sub>2</sub> (all data) <sup>a</sup>	0.1264	0.1109

$$^a R_1 = \sum ||F_o| - |F_c|| / \sum |F_o|. wR_2 = \sum [w(F_o^2 - F_c^2)^2] / \sum [w(F_o^2)^2]^{1/2}.$$

and six terminal oxygen atoms) that are able to coordinate with two edge-sharing CuO<sub>6</sub> or ZnO<sub>6</sub> octahedra as well as two open-chain [Sn(CH<sub>2</sub>)<sub>2</sub>COO]<sup>2+</sup> moieties. Sn1 is hexa-coordinated with five oxygen atoms (O3/O5, O27/O26, O35/O29, O39/43 and O52/O55) and one carbon atom (C1) derived from a carboxyethyl group (Fig. S2(a) and (b)†). The Sn–O bond distances are 2.056(13)–2.290(14) and 2.045(11)–2.300(11) Å, and the Sn–C bond distances are 2.12(3) and 2.12(2) Å for SnR-Cu-P<sub>2</sub>W<sub>15</sub> and SnR-Zn-P<sub>2</sub>W<sub>15</sub>, respectively. The bond lengths of Cu–O and Zn–O are in the range of 2.005(15)–2.241(15) and 2.028(11)–2.250(12) Å. The bond lengths of W–O are in the range of 1.683(13)–2.396(13) and 1.680(12)–2.411(11) Å (Tables S1 and S2†) for SnR-Cu-P<sub>2</sub>W<sub>15</sub> and SnR-Zn-P<sub>2</sub>W<sub>15</sub>, respectively. Moreover, the guanidinium [C(NH<sub>2</sub>)<sub>3</sub>]<sup>+</sup> cations and lattice water molecules locate around the polyoxoanions. Furthermore, in SnR-Zn-P<sub>2</sub>W<sub>15</sub>, two [C(NH<sub>2</sub>)<sub>3</sub>]<sup>+</sup> cations locate with 50% occupancy. For the packing arrangements of SnR-Cu-P<sub>2</sub>W<sub>15</sub> and SnR-Zn-P<sub>2</sub>W<sub>15</sub> (Fig. S3†), the sandwich-type polyoxoanions accumulated to form 3D supramolecular frameworks by hydrogen bonds (Tables S3 and S4,† N–H⋯O/OW 1.91–2.64 and 1.93–2.63 Å for SnR-Cu-P<sub>2</sub>W<sub>15</sub> and SnR-Zn-P<sub>2</sub>W<sub>15</sub>, respectively) between [C(NH<sub>2</sub>)<sub>3</sub>]<sup>+</sup> cations and polyoxoanions or water molecules, along with the electrostatic forces existing in [C(NH<sub>2</sub>)<sub>3</sub>]<sup>+</sup> cations and polyoxoanions. CCDC-1522095 and CCDC-1522096 contain the supplementary crystallographic data for this paper.

IR spectra (Fig. S4†) further confirmed the structures of SnR-Cu-P<sub>2</sub>W<sub>15</sub> and SnR-Zn-P<sub>2</sub>W<sub>15</sub>. The <sup>119</sup>Sn and <sup>31</sup>P NMR spectra (Fig. S5–S7†) proved the stability of SnR-Cu-P<sub>2</sub>W<sub>15</sub> and SnR-Zn-P<sub>2</sub>W<sub>15</sub> in aqueous solution. UV/Vis spectra of SnR-Cu-P<sub>2</sub>W<sub>15</sub> and SnR-Zn-P<sub>2</sub>W<sub>15</sub> exhibited a strong absorbance band at 273 nm in the near-UV region (Fig. S8†). From the TG analysis (Fig. S9†), the structures of carboxyethyltin–POMs still retain before 490 °C. And the purities of SnR-Cu-P<sub>2</sub>W<sub>15</sub> and SnR-Zn-P<sub>2</sub>W<sub>15</sub>

were confirmed by the consistent peak position of experimental and simulated results in XRPD (Fig. S10†). From the cyclic voltammograms in Fig. S11 and S12,† the electrochemical responses of the W centers were obviously observed.<sup>38</sup>

### Adsorption capability of SnR-M-P<sub>2</sub>W<sub>15</sub> towards organic dyes

To evaluate and compare the adsorption capability of SnR-M-P<sub>2</sub>W<sub>15</sub> and M-P<sub>2</sub>W<sub>15</sub>, the UV/Vis absorption spectra of RhB or MB solution were detected before and after the adsorption process (Fig. S13 and S14†). The adsorption capacity (*q<sub>t</sub>*, mg g<sup>-1</sup>) was calculated according to the following equations: *q<sub>t</sub>* = (*c<sub>o</sub>* – *c<sub>t</sub>*)/*V**m*, where *c<sub>o</sub>* and *c<sub>t</sub>* (mg L<sup>-1</sup>) represent the initial and final concentration of RhB or MB solution at certain interval time *t* (min); *V* (mL) was the final volume of RhB (MB) after mixing RhB (MB) and SnR-M-P<sub>2</sub>W<sub>15</sub> solution; *m* (mg) was the mass of SnR-M-P<sub>2</sub>W<sub>15</sub>. From Fig. S13,† it is clearly shown that SnR-M-P<sub>2</sub>W<sub>15</sub> displays a fast adsorption rate towards the cationic dyes RhB and MB. Also, Table 2 lists the *q<sub>t</sub>* values of various compounds, from which it is observed that the adsorption capacity of SnR-M-P<sub>2</sub>W<sub>15</sub> towards MB was larger than that of RhB; furthermore, the parent POMs (M-P<sub>2</sub>W<sub>15</sub>) showed the lower adsorption capacity than their corresponding carboxyethyltin derivatives, and the *q<sub>t</sub>* of SnR-Co-P<sub>2</sub>W<sub>15</sub> towards MB reached to 288.76 mg g<sup>-1</sup> in 10 min. So an interesting finding is that the functionalization of open-chain carboxyethyltin improved the adsorption ability of sandwich-type POMs.

After adsorption, the adsorbent was separated by centrifugation, resulting in the formation of pink precipitate (SnR-M-P<sub>2</sub>W<sub>15</sub>/RhB) or cyanous precipitate (SnR-M-P<sub>2</sub>W<sub>15</sub>/MB). SnR-M-P<sub>2</sub>W<sub>15</sub>/RhB was further characterized by FTIR (Fig. S15†) and solid UV/Vis absorption spectra (Fig. S16†). From the results of FTIR spectra, both SnR-M-P<sub>2</sub>W<sub>15</sub> and RhB maintain their stable structures after adsorption. Fig. S16† reveals that SnR-M-P<sub>2</sub>W<sub>15</sub>/RhB show two strong broad absorption bands centering at about 286–287 and 562–564 nm in the ultraviolet and visible regions, corresponding to the ligand to metal (O → W) charge transfer (LMCT) transition of POMs and π–π\* transition of RhB, respectively.<sup>25,42</sup> As compared to SnR-M-P<sub>2</sub>W<sub>15</sub>, the LMCT transition peak for SnR-M-P<sub>2</sub>W<sub>15</sub>/RhB blue shifted, which might be attributed to the decreased size of SnR-M-P<sub>2</sub>W<sub>15</sub> after adsorption.

**Table 2** Adsorption capacity of various compounds towards RhB and MB

Compounds	Adsorption capacity (RhB) (mg g <sup>-1</sup> )	Adsorption capacity (MB) (mg g <sup>-1</sup> )
SnR-Mn-P <sub>2</sub> W <sub>15</sub>	190.98	283.22
SnR-Co-P <sub>2</sub> W <sub>15</sub>	217.36	288.76
SnR-Cu-P <sub>2</sub> W <sub>15</sub>	179.02	263.47
SnR-Zn-P <sub>2</sub> W <sub>15</sub>	161.15	252.46
Mn-P <sub>2</sub> W <sub>15</sub>	158.07	237.28
Co-P <sub>2</sub> W <sub>15</sub>	70.42	174.93
Cu-P <sub>2</sub> W <sub>15</sub>	23.39	42.27
Zn-P <sub>2</sub> W <sub>15</sub>	35.29	44.54
P <sub>2</sub> W <sub>15</sub>	127.44	170.07



### Characterization of LBL assembly films

UV/Vis spectra were adopted to monitor the deposition process of LBL assembly films. Given that the polyoxoanions  $\text{SnR-M-P}_2\text{W}_{15}$  ( $\text{M} = \text{Mn, Co, Cu, Zn}$ ) are isostructural,  $(\text{SnR-Co-P}_2\text{W}_{15}/\text{RhB})_n$  or  $(\text{SnR-Co-P}_2\text{W}_{15}/\text{MB})_n$  films were taken as an example for describing the structural and morphological characterization of composite films. As shown in Fig. 1(a), the absorbance of  $(\text{SnR-Co-P}_2\text{W}_{15}/\text{RhB})_n$  ( $n = 1-6$ ) film on quartz substrate increased progressively as the bilayer number, indicating that the composite film was successfully fabricated. In addition, the inset of Fig. 1(a) shows the plot of absorbance at  $\lambda = 573$  nm versus the bilayer number  $n$  results in a nearly straight line, which proved the composite film is well-proportioned and both in the near-UV and visible regions, in which the LMCT absorption band of  $\text{SnR-Co-P}_2\text{W}_{15}$  in the UV region is overlapped by the absorption peak of RhB, and the visible peaks are assigned to the  $\pi-\pi^*$  transition of RhB, respectively. Meanwhile, compared with the absorption peak of RhB solution (554 nm), the absorption peak of RhB in  $(\text{SnR-Co-P}_2\text{W}_{15}/\text{RhB})_n$  film (573 nm) red shifted, indicating RhB molecules form J-aggregates.<sup>43</sup> Similarly, from Fig. 1(b) it was concluded that  $(\text{SnR-Co-P}_2\text{W}_{15}/\text{MB})_n$  film was successfully assembled and the formed composite film is well-proportioned and homogeneous.

The surface morphology and homogeneity of the multilayer films were observed by atomic force microscope (AFM) images (Fig. 2(a) and (b)). As can be seen, the surface of  $(\text{SnR-Co-P}_2\text{W}_{15}/\text{RhB})_4$  film ( $R_q = 3.19$  nm,  $R_a = 2.35$  nm) is more rough than that of  $(\text{PSS}/\text{RhB})_4$  film ( $R_q = 2.68$  nm,  $R_a = 1.66$  nm), which was probably because of the rigid structure and aggregation of POM anions. In addition, the AFM image  $(\text{Co-P}_2\text{W}_{15}/\text{RhB})_4$  film also shows rough surface and prominent aggregated particles (Fig. S17<sup>†</sup>). Fig. 2(c) portrays the FTIR spectra of  $(\text{SnR-M-P}_2\text{W}_{15}/\text{RhB})_4$  films ( $\text{M} = \text{Mn, Co, Cu}$  and  $\text{Zn}$ ) and  $(\text{PSS}/\text{RhB})_4$  film on  $\text{CaF}_2$  substrate. In comparison with  $(\text{PSS}/\text{RhB})_4$  film,  $(\text{SnR-M-P}_2\text{W}_{15}/\text{RhB})_4$  films exhibit the characteristic bands of  $\text{SnR-M-P}_2\text{W}_{15}$  observed at 1091, 954–949, 913–907 and 788–791  $\text{cm}^{-1}$ ,<sup>25,38</sup> respectively, suggesting that the structures of  $\text{SnR-M-P}_2\text{W}_{15}$  retained in the films. FTIR spectra of  $(\text{SnR-M-P}_2\text{W}_{15}/\text{MB})_4$

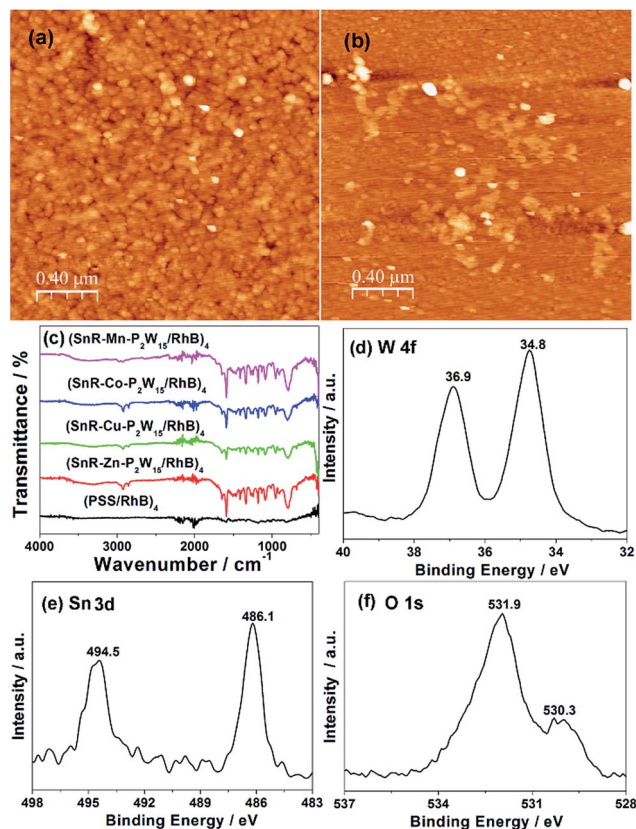


Fig. 2 The AFM images of  $(\text{SnR-Co-P}_2\text{W}_{15}/\text{RhB})_4$  film (a) and  $(\text{PSS}/\text{RhB})_4$  film (b) on the silicon wafer in the tapping mode, FTIR spectra of  $(\text{SnR-M-P}_2\text{W}_{15}/\text{RhB})_4$  ( $\text{M} = \text{Mn, Co, Cu, Zn}$ ) and  $(\text{PSS}/\text{RhB})_4$  films on  $\text{CaF}_2$  substrate (c), XPS spectra of  $(\text{SnR-Co-P}_2\text{W}_{15}/\text{RhB})_7$  film on Si substrate: W 4f (d), Sn 3d (e) and O 1s (f).

$\text{P}_2\text{W}_{15}$  observed at 1091, 954–949, 913–907 and 788–791  $\text{cm}^{-1}$ ,<sup>25,38</sup> respectively, suggesting that the structures of  $\text{SnR-M-P}_2\text{W}_{15}$  retained in the films. FTIR spectra of  $(\text{SnR-M-P}_2\text{W}_{15}/\text{MB})_4$

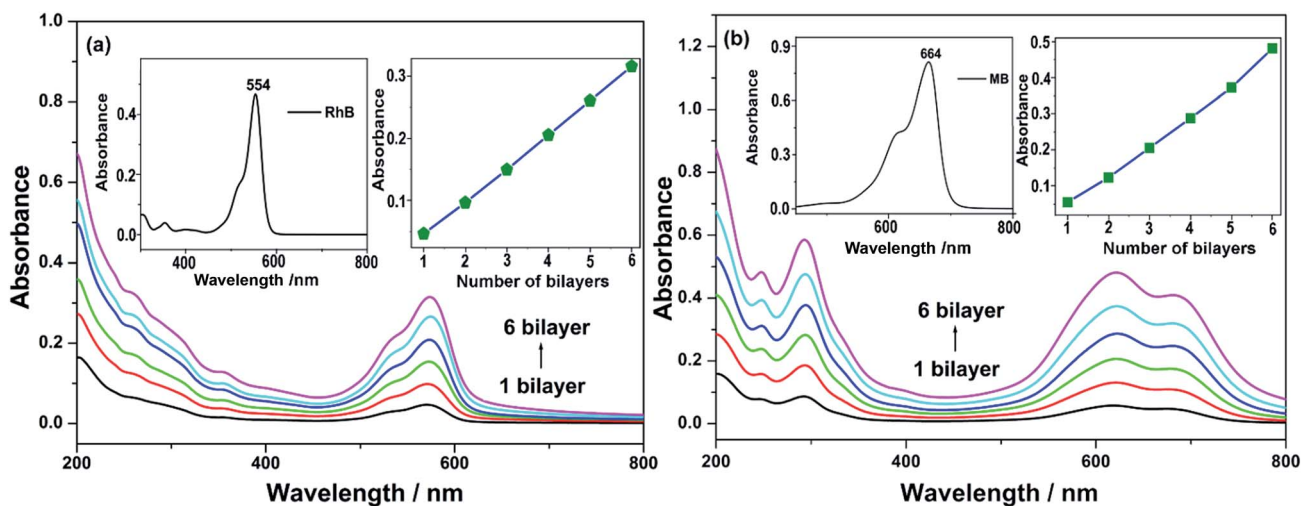


Fig. 1 UV/Vis absorption spectra of  $(\text{SnR-Co-P}_2\text{W}_{15}/\text{RhB})_n$  ( $n = 1-6$ ) multilayer films, and the insets of (a) show UV/Vis absorption spectrum of RhB aqueous solution and the absorbance of multilayer film at  $\lambda_{\text{max}} = 573$  nm as a function of  $n$  (a); UV/Vis absorption spectra of  $(\text{SnR-Co-P}_2\text{W}_{15}/\text{MB})_n$  ( $n = 1-6$ ) multilayer films, and the insets of (b) show UV/Vis absorption spectrum of MB aqueous solution and the absorbance of multilayer film at  $\lambda_{\text{max}} = 619$  nm as a function of  $n$  (b), respectively.



and (PSS/MB)<sub>4</sub> films were measured and shown in Fig. S18,<sup>†</sup> from which we can see that the peaks at 1096, 949–946, 901–892 and 765–761 cm<sup>-1</sup> in (SnR-M-P<sub>2</sub>W<sub>15</sub>/MB)<sub>4</sub> (M = Mn, Co, Cu and Zn) films can be attributed to the characteristic vibrations of the POM skeleton. The results indicate that the structures of SnR-M-P<sub>2</sub>W<sub>15</sub> retained in the films. XPS measurement was used to determine the existence of POMs and their oxidation states in the multilayer film. From Fig. 2(d)–(f), the binding energies of W 4f of SnR-Cu-P<sub>2</sub>W<sub>15</sub> were located at 36.9 eV and 34.8 eV, respectively, which are assigned to the W<sup>VI</sup> oxidation state.<sup>44</sup> The two peaks at 494.5 and 486.1 eV were consistent with the XPS spectrum of Sn 3d.<sup>45,46</sup> The above results suggested that SnR-Co-P<sub>2</sub>W<sub>15</sub> in the film exists in its oxidized form. The binding energy located at 530.3 and 531.9 eV in Fig. 2(f) were attributed to O 1s, in which the peak at 530.3 eV could be ascribed to the O<sup>2-</sup> in the framework of POMs, whereas the peak at 531.9 eV was assigned to O<sup>2-</sup> in carbon–oxygen bond.<sup>47,48</sup> Fig. 3 compares the optical absorption characteristics of different series of LBL assembly films. From Fig. 3(a) we can see that (SnR-Co-P<sub>2</sub>W<sub>15</sub>/RhB)<sub>4</sub> film exhibits the strongest UV/Vis

absorption intensity among its parent POMs and PSS-based films, which was ascribed to the superior adsorption ability of SnR-Co-P<sub>2</sub>W<sub>15</sub>.

Meantime, it is worth noting that (PEI/SnR-Co-P<sub>2</sub>W<sub>15</sub>)<sub>4</sub> shows almost no absorbance in the visible region, which was resulted from the weak visible activity and low content of SnR-Co-P<sub>2</sub>W<sub>15</sub>. To further understand the influence of structures of carboxyethyltin and POMs on the adsorption ability, Fig. 3(b) was obtained, which illustrates the following results: (1) Dawson-type open-chain carboxyethyltin–POM derivative SnR-Zn-P<sub>2</sub>W<sub>15</sub> displays significantly enhanced adsorption ability than that of Keggin-type analogue SnR-Zn-PW<sub>9</sub>, owing to the more negative charges of Dawson POM anions; (2) SnR-Zn-PW<sub>9</sub>, with the open-chain carboxyethyltin shows larger adsorbing capacity than that of SnR-PW<sub>9</sub>, with the closed-ring carboxyethyltin structure, indicating that the open-chain structure benefits for the adsorption; (3) SnR-COOH itself shows certain adsorption towards RhB in despite of their identical charge types in aqueous solution, which implies that the non-electrostatic force can form between the open-chain carboxyethyltin and RhB. Fig. 3(c) reflects the

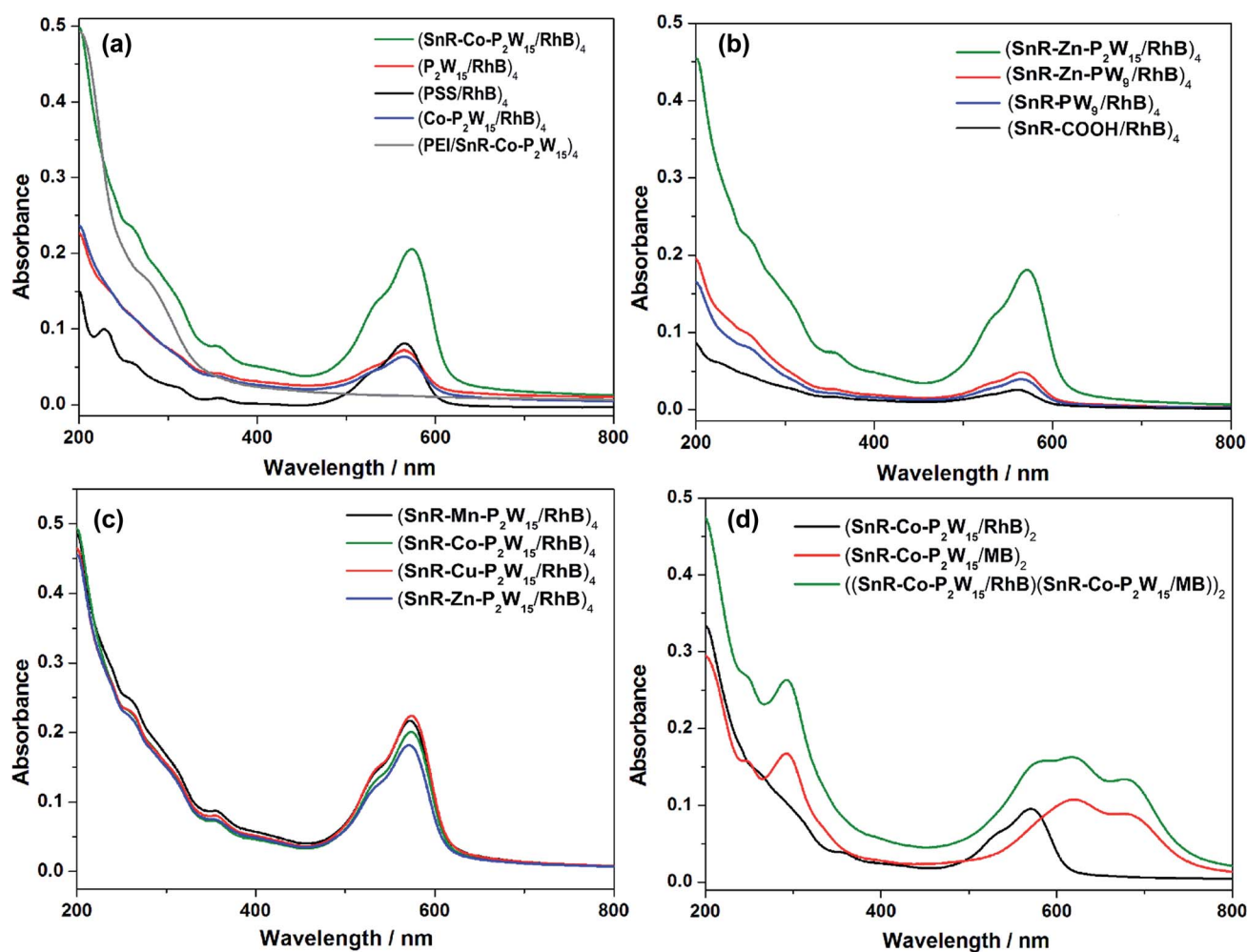


Fig. 3 UV/Vis absorption spectra of (SnR-Co-P<sub>2</sub>W<sub>15</sub>/RhB)<sub>4</sub>, (Co-P<sub>2</sub>W<sub>15</sub>/RhB)<sub>4</sub>, (P<sub>2</sub>W<sub>15</sub>/RhB)<sub>4</sub>, (PSS/RhB)<sub>4</sub> and (PEI/SnR-Co-P<sub>2</sub>W<sub>15</sub>)<sub>4</sub> films (a); (SnR-PW<sub>9</sub>/RhB)<sub>4</sub>, (SnR-COOH/RhB)<sub>4</sub>, (SnR-Zn-PW<sub>9</sub>/RhB)<sub>4</sub> and (SnR-Zn-P<sub>2</sub>W<sub>15</sub>/RhB)<sub>4</sub> films (b); (SnR-M-P<sub>2</sub>W<sub>15</sub>/RhB)<sub>4</sub> (M = Mn, Co, Cu and Zn) films (c); (SnR-Co-P<sub>2</sub>W<sub>15</sub>/RhB)<sub>2</sub>, (SnR-Co-P<sub>2</sub>W<sub>15</sub>/MB)<sub>2</sub> and ((SnR-Co-P<sub>2</sub>W<sub>15</sub>/RhB)(SnR-Co-P<sub>2</sub>W<sub>15</sub>/MB))<sub>2</sub> films (d), respectively.



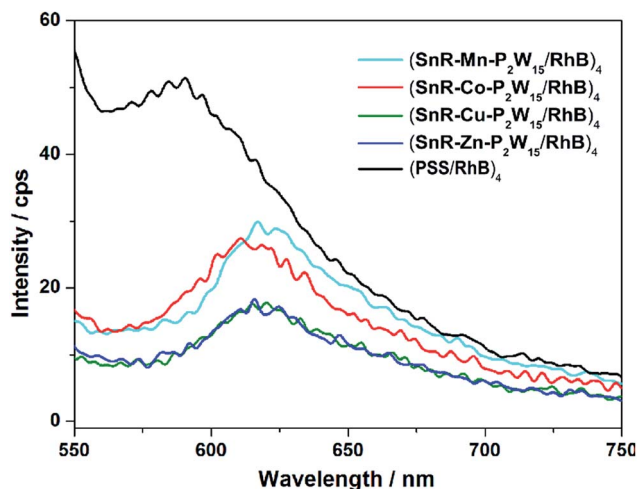


Fig. 4 Fluorescence spectra of  $(\text{PSS/RhB})_4$  and  $(\text{SnR-M-P}_2\text{W}_{15}/\text{RhB})_4$  ( $M = \text{Mn, Co, Cu, Zn}$ ) multilayer films ( $\lambda_{\text{ex}} = 515 \text{ nm}$ ) on the quartz substrate at room temperature.

absorbance differences of  $(\text{SnR-M-PW}_{15}/\text{RhB})_4$  ( $M = \text{Mn, Co, Cu}$  and  $\text{Zn}$ ) films are not obvious. Furthermore, Fig. 3(d) reveals the optical responsive behaviors of LBL films can be broadened to visible and near-IR regions by incorporating RhB and MB simultaneously into the films.

The fluorescence characteristics of  $(\text{PSS/RhB})_4$  and  $(\text{SnR-M-P}_2\text{W}_{15}/\text{RhB})_4$  ( $M = \text{Mn, Co, Cu}$  and  $\text{Zn}$ ) films have been investigated at room temperature. As shown in Fig. 4, the fluorescence spectra of LBL composite film infers that RhB can stably exist in the multilayer films. Additionally, compared with  $(\text{PSS/RhB})_4$  film, the emission peak  $(\text{SnR-M-P}_2\text{W}_{15}/\text{RhB})_4$  ( $M = \text{Mn, Co, Cu}$  and  $\text{Zn}$ ) films red shifted due to the aggregation of RhB. Meanwhile, a part of fluorescence of RhB was quenched after compositing with  $\text{SnR-M-P}_2\text{W}_{15}$ , indicating that charge transfer exists between  $\text{SnR-M-P}_2\text{W}_{15}$  and RhB.

### Photoelectrochemical properties

To investigate the visible photoelectric performance of the LBL films, the photocurrent response experiments were carried out by using a three-electrode system under irradiation ( $\lambda > 420 \text{ nm}$ ). Fig. S19 and S20† show the photocurrent responses

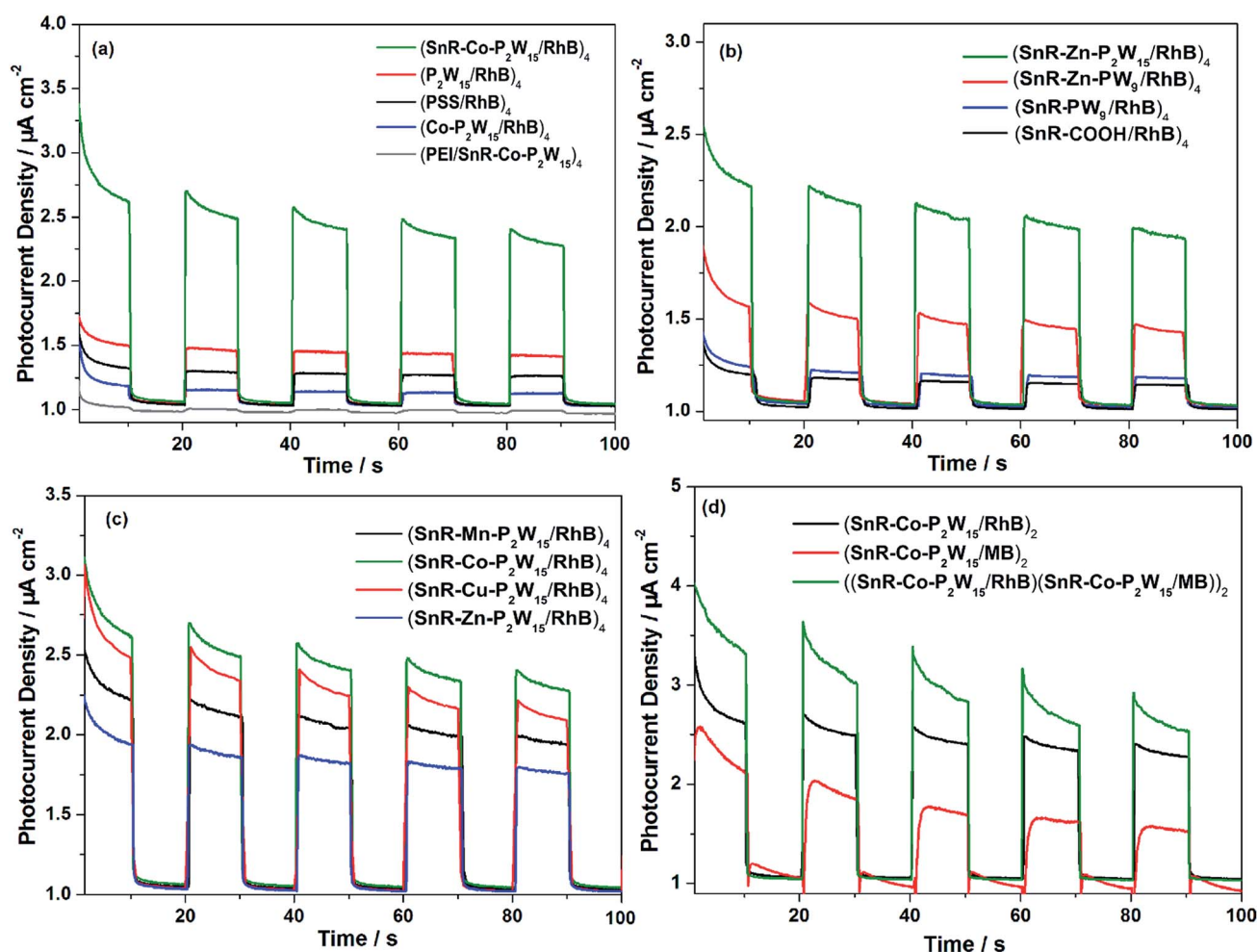


Fig. 5 Photocurrent responses of various series of LBL multilayer films under visible light irradiation ( $\lambda > 420 \text{ nm}$ ) in  $0.1 \text{ mol L}^{-1} \text{ Na}_2\text{SO}_4$  aqueous solution with the applied bias voltage of  $0.5 \text{ V}$  vs.  $\text{Ag}/\text{AgCl}$ .



observed for films at various pH values and with different numbers of bilayers ( $n = 1-6$ ), respectively. From Fig. S19<sup>†</sup> it was concluded that the photocurrent reaches to a plateau when the number of bilayers increase to 4, which was possibly caused by the slower electron transfer rate in the thicker films.

Fig. 5(a) shows the photocurrent responses of (SnR-Co-P<sub>2</sub>W<sub>15</sub>/RhB)<sub>4</sub> and the control films, from which we can see that all samples displayed a fast anodic photocurrent response when the illumination is switched on. Meanwhile, notably (SnR-Co-P<sub>2</sub>W<sub>15</sub>/RhB)<sub>4</sub> film performs the highest photocurrent, whereas the RhB-free film (PEI/SnR-Co-P<sub>2</sub>W<sub>15</sub>)<sub>4</sub> exhibits very little photocurrent due to its low visible light activity and low carrier separation efficiency. As a result, the visible photoelectric activity of SnR-Co-P<sub>2</sub>W<sub>15</sub> can be improved by the adsorption of RhB. Furthermore, (SnR-Co-P<sub>2</sub>W<sub>15</sub>/RhB)<sub>4</sub> electrode displays higher photocurrent than the photocurrent observed for (PSS/RhB)<sub>4</sub>, (P<sub>2</sub>W<sub>15</sub>/RhB)<sub>4</sub> and (Co-P<sub>2</sub>W<sub>15</sub>/RhB)<sub>4</sub> electrodes, which may be ascribed to the larger adsorbing capacity of SnR-Co-P<sub>2</sub>W<sub>15</sub> towards RhB and more efficient visible light harvesting. To further understand this phenomenon, the photocurrent responses for SnR-M-P<sub>2</sub>W<sub>15</sub> (M = Mn, Cu and Zn) electrodes and their corresponding control electrodes were measured in Fig. S21,<sup>†</sup> which shows that SnR-M-P<sub>2</sub>W<sub>15</sub> (M = Mn, Cu and Zn)-based films produced the highest photocurrent responses as compared to their parent POMs and PSS-based films. Besides, Fig. 5(b) compared the photocurrent responses of POMs with different carboxyethyltin structure and POM structure. In accordance with the UV/Vis spectral responses in Fig. 3(b), SnR-Zn-P<sub>2</sub>W<sub>15</sub> based film displays the highest photocurrent response as compared to SnR-PW<sub>9</sub>, SnR-Zn-PW<sub>9</sub> and SnR-COOH. Meantime, there is a slight difference between the photocurrents generated from four isostructural double armed sandwich-type carboxyethyltin-decorated POMs. From Fig. 5(c) it was observed that SnR-Co-P<sub>2</sub>W<sub>15</sub>-based photoelectrode presents the optimal photoelectric performance than other SnR-M-P<sub>2</sub>W<sub>15</sub>-based (M = Mn, Cu and Zn) electrodes. In addition, the other organic dye MB was applied to composite with SnR-Co-P<sub>2</sub>W<sub>15</sub>. Accordingly, the photocurrent response of (SnR-Co-P<sub>2</sub>W<sub>15</sub>/MB)<sub>4</sub> film was examined and shown in Fig. S22,<sup>†</sup> which also proved (SnR-Co-P<sub>2</sub>W<sub>15</sub>/MB)<sub>4</sub> film exhibited better photoelectric properties than the reference films including (P<sub>2</sub>W<sub>15</sub>/MB)<sub>4</sub>, (Co-P<sub>2</sub>W<sub>15</sub>/MB)<sub>4</sub> and (PSS/MB)<sub>4</sub> films. In order to further broaden the absorption spectra of POMs, RhB and MB were simultaneously assembled into SnR-Co-P<sub>2</sub>W<sub>15</sub>-based LBL films, and its photocurrent response was superior to the photocurrents of (SnR-Co-P<sub>2</sub>W<sub>15</sub>/MB)<sub>4</sub> and (SnR-Co-P<sub>2</sub>W<sub>15</sub>/RhB)<sub>4</sub> films in Fig. 5(d). Furthermore, some organic dye-POM LBL films have been reported, such as hemicyanine-POM films. They show obvious cathodic photocurrent response, which was generated based on charge transfer excitation of hemicyanine in the film.<sup>49-51</sup> Meantime, SnR-M-P<sub>2</sub>W<sub>15</sub>-based (M = Mn, Co, Cu and Zn) electrodes exhibit anodic photocurrent, which indicates that different types of photocurrent can be produced by assembling suitable POMs and organic dyes. In addition, the stability of the films was studied as shown in Fig. S23 and S24.<sup>†</sup> The absorbance of (SnR-Co-P<sub>2</sub>W<sub>15</sub>/RhB)<sub>4</sub> and (SnR-Co-P<sub>2</sub>W<sub>15</sub>/MB)<sub>4</sub> films reduced by 30.5% and 3.02%, respectively, after

dipping in aqueous solution for 2 h, which concluded that (SnR-Co-P<sub>2</sub>W<sub>15</sub>/MB)<sub>4</sub> film is more stable than RhB-based film.

## Conclusions

In summary, we present data demonstrating the photoelectric performance of the LBL multilayer photoelectrode (SnR-M-P<sub>2</sub>W<sub>15</sub>/dye)<sub>n</sub> (dye = RhB or MB,  $n = 1-6$ ) was obviously enhanced owing to their efficient light harvesting in visible region and strong adsorption capacity of SnR-M-P<sub>2</sub>W<sub>15</sub>. And the results demonstrate that Dawson structure and functionalization of open-chain carboxyethyltin play significant roles on the superior photoelectric properties. Furthermore, the transitional metal incorporated in SnR-M-P<sub>2</sub>W<sub>15</sub> also matters and SnR-Co-P<sub>2</sub>W<sub>15</sub> was proved to lead to the optimal effect. Meanwhile, the optical absorption characteristics of SnR-M-P<sub>2</sub>W<sub>15</sub> can be further expanded to visible and near-IR regions through simultaneously incorporating both RhB and MB in SnR-M-P<sub>2</sub>W<sub>15</sub>-based electrodes, which induces the ascending photoelectric performances. Our work opens up possibilities on the recycle of dye in wastewater for further use on solar energy conversion. Further work on compositing Dawson-type open-chain carboxyethyltin functionalized POMs with other photoactive materials are under going on.

## Acknowledgements

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