



Cite this: *RSC Adv.*, 2017, 7, 18068

Theoretical exploration on the electronic and magnetic properties of $(\text{FeCp})_n$ – ($n = 1, 2$) ligand-functionalized graphene†

Xiuyun Zhang,^a Zujian Bao,^a Wenxian Xu,^a Lili Liu^b and Yongjun Liu^{*a}

Using first principles calculations, we systematically studied the electronic and magnetic properties of half sandwich organometallic ligand-functionalized single layer graphene, $(\text{FeCp})_n@SLGs$ ($n = 1, 2$), with different adsorption sites and coverage concentrations. Our results show that all the $(\text{FeCp})_n@SLGs$ are quite stable having high binding energies. Except the nonmagnetic $\text{FeCp}@G_{44}\text{-S}_1$, all the complexes are found to be robust ferromagnets. In particular, the magnetic moments of $(\text{FeCp})_2@G_{33}\text{-D}_1$, $(\text{FeCp})_2@G_{33}\text{-D}_2$, $(\text{FeCp})_2@G_{33}\text{-D}_3$ and $(\text{FeCp})_2@G_{44}\text{-D}_1$ per unit cell are as large as $2.0 \mu_B$, $1.64 \mu_B$, $2.0 \mu_B$ and $1.30 \mu_B$, respectively. Among the studied systems, spin polarized band gaps are opened in the Dirac points of $\text{FeCp}@G_{33}\text{-S}$, $(\text{FeCp})_2@G_{33}\text{-D}_1$, $(\text{FeCp})_2@G_{44}\text{-S}_2$, $(\text{FeCp})_2@G_{44}\text{-S}_3$, and $(\text{FeCp})_2@G_{44}\text{-D}_5$, in which $(\text{FeCp})_2@G_{33}\text{-D}_1$ is transformed into an intrinsic semiconductor. Moreover, displacing Fe by Co or Ni element is found to induce an increase in magnetic moment or induce a transition from metal to half metal.

Received 8th February 2017

Accepted 13th March 2017

DOI: 10.1039/c7ra01617d

rsc.li/rsc-advances

1. Introduction

Graphene, which is a two-dimensional single-atom-thick material, exhibits unusual electronic properties^{1–8} and has attracted great attention due to its potential application as a building block of nanoelectronic devices. However, the realization of graphene's commercial applications has to overcome two primary obstacles: (i) the synthesis of high-quality graphene on a large scale, and (ii) producing a band gap in graphene to replace silicon in future, which is the biggest challenge. Great efforts have been made to overcome the proposed challenges and various strategies have been proposed for the synthesis of graphene, such as mechanical peeling of graphite,² ultrasonic exfoliation,⁹ chemical reduction of graphite oxide (GO)¹⁰ and chemical vapor decomposition (CVD) growth on transition metal surfaces.^{11–14} Particularly, the CVD method has been proven to be an effective way to meet the requirement of overcoming challenge (i). For challenge (ii), a number of solutions have been proposed by theoretical and experimental studies. One straightforward method is to combine various graphene/X heterostructures,^{15,16} e.g. a 0.26 eV band gap is obtained by

placing graphene on a SiC surface.¹⁵ Another is substitutional doping with exotic elements (e.g. B and N)^{17–19} e.g. the band gap of graphene is found to increase linearly with the doped BN concentration.¹⁸ In addition, the quantum confinement effect by cutting graphene into nanoribbons^{19,20} has also been demonstrated as a way to open the gap of graphene. Alternatively, chemical/physical adsorption has been proven to be a simple and effective solution to tune the electronic properties of graphene,^{21–30} in which the approach of loading half organo-transition metal (TM) ligands (OTMLs) on graphene substrates has been theoretically explored.^{25–29} The adsorption of OTMLs on graphene has two significant advantages: (i) compared with two-end closed organometallic sandwich clusters,^{31–33} the bonding between half TM_nOLs and graphene is stronger due to coordinate d- π bonding. For example, the chemical attachment of a C_{60} fullerene on graphene is not energetically feasible,²⁴ whereas connecting graphene and C_{60} using a bridging Cr atom can largely enhance the binding energy;²⁵ (ii) with single TM atom adsorbed graphene, e.g. $\text{TM}@Gs$,^{34–39} the existence of organic ligands has the advantage of shielding the clustering and oxidation of the metal atoms. Indeed, OTMLs-functionalized graphene introduces versatile electronic properties, for example, the covalent bonding of TMBz to graphene opens up a 0.81 eV band gap²⁶ and induces a high magnetization energy.²⁷ In our previous studies, a density related band gap was observed for TMBz-functionalized graphene.^{28,29}

Compared with hexagonal hydrocarbon organic ligands (e.g. benzene), the bonding of TM atoms with the cyclopentadienyl ($\text{Cp} = \text{C}_5\text{H}_5$) ligand was found to display different bonding mechanisms due to the fact that it has one less valence

^aCollege of Physical Science and Technology, Yangzhou University, Yangzhou 225002, China. E-mail: xyzhang@yzu.edu.cn; yjliu@yzu.edu.cn

^bDepartment of Applied Physics and Key Laboratory of Soft Chemistry and Functional Materials (Ministry of Education), Nanjing University of Science and Technology, Nanjing, Jiangsu 210094, P. R. China

† Electronic supplementary information (ESI) available: Optimized structures, density of states and band structures of $\text{FeCp}@G_{44}\text{-Ss}$, $(\text{CoCp})_n@Gs$ and $(\text{NiCp})_n@Gs$, the CDDs plots of $(\text{FeCp})_n@Gs$, $(\text{CoCp})_n@Gs$ and $(\text{NiCp})_n@Gs$. See DOI: 10.1039/c7ra01617d



electron.³¹ Herein, we present a systematic theoretical study on the structural, electronic and magnetic properties of FeCp ligands-adsorbed graphene. For comparison, 3×3 and 4×4 graphene supercells with single-side and double-side adsorption are considered. Our results show that most systems are ferromagnetic in their ground states with the exception of FeCp@G₄₄-S₁. All the (FeCp)_n@G₃₃s and a number of (FeCp)_n@G₄₄s have a spin dependent band gap in the Dirac point. Particularly, (FeCp)₂@G₃₃-D₁ is transformed into an intrinsic semiconductor. Moreover, displacing Fe by Co or Ni elements is found to induce an increase in magnetic moments or induce a metallic properties transition.

2. Computational methodology and models

All calculations are performed within the framework of spin polarized DFT as implemented in the Vienna *Ab initio* Simulation Package (VASP).^{40,41} The exchange–correlation potentials are treated by the generalized gradient approximation (GGA) parametrized by Perdew, Burke, and Ernzerhof (PBE).⁴² The interaction between valence electrons and ion cores is described by the projected augmented wave (PAW) method.^{43,44} In order to consider the local coulomb interaction in the 3d orbitals of the TM atom, we incorporated an effective on-site coulomb interaction for the 3d states of TM atoms using the GGA+*U* ($U = 4$ eV and $J = 0.9$ eV) method. Furthermore, the DFT-D₂, method taking into account the van der Waals (vdW) interaction, was adopted.⁴⁵ Both 3×3 (G₃₃) and 4×4 (G₄₄) slabs with the lattice parameters of $a = b = 7.41$ Å and 9.88 Å, respectively, were used for the adsorption blocks. Accordingly, the Monkhorst–Pack grids of $13 \times 13 \times 1$ and $9 \times 9 \times 1$ were used for the (FeCp)_n@G₃₃s and (FeCp)_n@G₄₄s ($n = 1, 2$) geometry optimization, respectively. Moreover, much denser *k*-point grids of $25 \times 25 \times 1$ to $23 \times 23 \times 1$ were used for electronic properties exploration. To eliminate the interaction between the entities of neighbouring unit cells, the lattice parameters along the *z* axis were set as large as 15 Å and 18 Å for the single-side/two-side adsorbed systems. Moreover, a cutoff of 400 eV was used for the plane wave expansion of the wave function to converge the relevant quantities. Structure relaxations were carried out until all the atomic forces on each ion were less than 0.01 eV Å⁻¹.

3. Results and discussion

3.1 Structural, electronic and magnetic properties of (FeCp)_n@G₃₃s, $n = 1, 2$

3.1.1 Geometries and stabilities. A total of four configurations were prepared for (FeCp)_n@G₃₃s: (i) FeCp@G₃₃-S, in which one FeCp ligand is adsorbed on the hollow site of graphene (see Fig. 1a and e); (ii) (FeCp)₂@G₃₃-D₁, in which two FeCp ligands are adsorbed on opposite sides and same hollow site of graphene (see Fig. 1b and f); (iii) (FeCp)₂@G₃₃-D₂, isomer of configuration (ii) with two FeCp ligands adsorbed on the adjacent hollow sites of graphene (see Fig. 1c and g); and (iv) (FeCp)₂@G₃₃-D₃, isomer of configurations (ii) and (iii) by

binding the two FeCp ligands on the hollow sites spaced by one line of hexagons (see Fig. 1d and h). It is found that most of the (FeCp)_n@G₃₃s have normal sandwich configurations. An exception is found for (FeCp)₂@G₃₃-D₁, in which one FeCp ligand slightly deviates from the hollow site of graphene due to the coulomb repulsion (see Fig. 1b). The C–C and C–H bond lengths of the Cp ligands are in the range of 1.43–1.44 Å and 1.09–1.10 Å, respectively, whereas the C–C bond lengths of graphene are around 1.42–1.44 Å (see Table S1 in the (ESI†)). Moreover, the C–C bond lengths of graphene close to the Fe atom are a bit longer than those that are far away due to the local $\pi(\text{G})\text{--d}(\text{Fe})$ interaction. For example, the graphene C–C bond lengths of the carbon hexagon facing the Fe atom are in the range of 1.43–1.44 Å, whereas those located around are about 1.42 Å. Moreover, the Fe atoms are at a distance from the mass center of the Cp ligands ($D_{\text{Fe--Cp}}$) and graphene ($D_{\text{Fe--G}}$) in the range of 1.69–1.76 Å and 1.60–1.69 Å, respectively (see Table 1), among which, the $D_{\text{Fe--Cp}}$ s distances are a bit longer than that of $D_{\text{Fe--G}}$ s. Distinct geometric characters are derived from the different bonding characters, in which the charges donated to graphene are much smaller than that of Cp ligand (see Table 1 and Fig. 2e–j), leading to a more ionic bonding character of Fe–Cp and more covalent bonding characters of Fe–graphene. Different from the FeBz@G isomer,²⁸ the $D_{\text{Fe--Cp}}$ of FeCp@G₃₃-S is a bit larger, ~ 1.76 Å, whereas a different case is verified for that of $D_{\text{Fe--G}}$, ~ 1.64 Å. Compared with that of FeCp@G₃₃-S, the $D_{\text{Fe--G}}$ of (FeCp)₂@G₃₃-D₁ is much larger, ~ 1.69 Å, whereas for the (FeCp)₂@G₃₃-D₂ and (FeCp)₂@G₃₃-D₃ isomers, the $D_{\text{Fe--G}}$ s are reduced sharply to 1.60 Å (see Table 1).

To verify the structural stabilities of (FeCp)_n@SLGs, we calculated the binding energies, which are defined as follows:

$$E_b = -(E_{(\text{FeCp})_n@\text{SLG}} - nE_{\text{Cp}} - nE_{\text{Fe}} - E_{\text{G}})/n$$

where $E[\cdot]$ represents the total energy of (FeCp)_n@SLGs, the Cp ligands, Fe atom and pristine graphene, and n is the number of FeCp ligands. Our results show that the binding energies of all

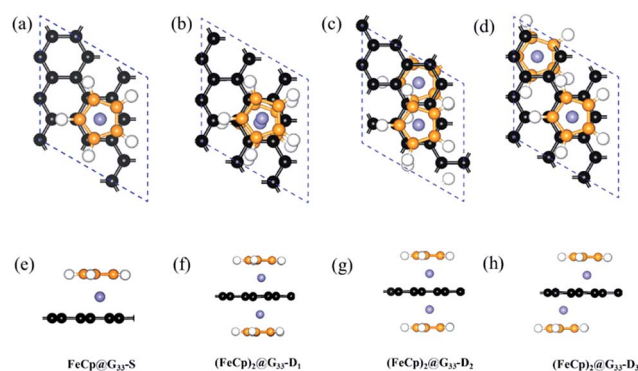


Fig. 1 Top and side views of the optimized structures of FeCp@G₃₃-S (a and e), (FeCp)₂@G₃₃-D₁ (b and f), (FeCp)₂@G₃₃-D₂ (c and g), and (FeCp)₂@G₃₃-D₃ (d and h). Black and orange balls represent the carbon atoms of graphene and Cp ligands; and white and purple balls represent H and Fe atoms, respectively. The complete descriptions of (FeCp)₂@G₃₃-D₂ (c) and (FeCp)₂@G₃₃-D₃ (d) are depicted in the 2×2 supercells in Fig. S1(a, b, e and f) in the ESI.†



Table 1 Systems (Sys), the distance of the Fe atom to Cp ligands ($D_{\text{Fe-Cp}}$) and graphene ($D_{\text{Fe-G}}$), charges transferred from Fe atom to graphene (Δe), binding energies (E_b), magnetic moments (MM) and the band gaps at the Dirac point (Δ_g) of $\text{FeCp}@G_{333}$ and $\text{FeCp}@G_{444}$ s

Sys	$D_{\text{Fe-Cp}}$ (Å)	$D_{\text{Fe-G}}$ (Å)	Δe (e)	E_b (kJ mol ⁻¹)	MM (μ_B)	Δ_g (eV)
$\text{FeCp}@G_{333}\text{-S}$	1.76	1.64	0.44	271.6	0.81	0.78
$(\text{FeCp})_2@G_{333}\text{-D}_1$	1.71	1.69, 1.69	0.72	256	2.00	0.40
$(\text{FeCp})_2@G_{333}\text{-D}_2$	1.69	1.60	0.83	277.6	1.64	0
$(\text{FeCp})_2@G_{333}\text{-D}_3$	1.73	1.60	0.80	278.7	2.00	0.61
$\text{FeCp}@G_{444}\text{-S}_1$	1.68	1.60	0.58	285.9	0	0
$(\text{FeCp})_2@G_{444}\text{-S}_2$	1.68	1.62	0.94	256.6	0.74	0.25
$(\text{FeCp})_2@G_{444}\text{-S}_3$	1.77, 1.64	1.61, 1.42	0.99	270.8	0.94	0.12
$(\text{FeCp})_2@G_{444}\text{-D}_1$	1.68	1.65	0.87	257.4	1.30	0
$(\text{FeCp})_2@G_{444}\text{-D}_2$	1.67	1.60	0.77	277.8	0.97	0
$(\text{FeCp})_2@G_{444}\text{-D}_3$	1.58	1.65, 1.69	1.06	281	0.50	0
$(\text{FeCp})_2@G_{444}\text{-D}_4$	1.65	1.58	1.02	280.4	0.49	0
$(\text{FeCp})_2@G_{444}\text{-D}_5$	1.65	1.58	1.03	277.6	0.61	0.25

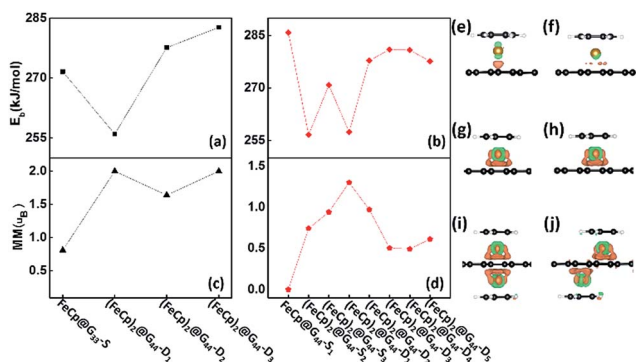


Fig. 2 Binding energies (a and b) and total magnetic moments (c and d) of $(\text{FeCp})_n@G_{333}$ and $(\text{FeCp})_n@G_{444}$ s, $n = 1, 2$. CDD plots of $\text{FeBz}@G_{333}\text{-S}$ (e), $\text{FeBz}@G_{444}\text{-S}$ (f), $\text{FeCp}@G_{333}\text{-S}$ (g), $\text{FeCp}@G_{444}\text{-S}$ (h), $(\text{FeCp})_2@G_{333}\text{-D}_1$ (i) and $(\text{FeCp})_2@G_{333}\text{-D}_2$ (j).

the studied $\text{FeCp}@G_{333}$ s are positive and range from 256 kJ mol⁻¹ to 285.9 kJ mol⁻¹, which is much larger than that of the $\text{FeBz}@G$ isomer (211 kJ mol⁻¹),²⁸ where such large binding energies show that the chemical attachment dominates the FeCp ligands and graphene interaction. As shown in the charge density difference plots (CDDs) (Fig. 2e–j), the electron densities around the graphene in $\text{FeCp}@SLGs$ (see Fig. 2g and h) is much higher than that of $\text{FeBz}@G_{333}$ and $\text{FeBz}@G_{444}$ (see Fig. 2e and f), which indicates a more covalent bonding character of $\text{FeCp}@SLGs$. The binding energies of $\text{FeCp}@G_{333}\text{-S}$, $(\text{FeCp})_2@G_{333}\text{-D}_2$ and $(\text{FeCp})_2@G_{333}\text{-D}_3$ are 271.6 kJ mol⁻¹, 277.6 kJ mol⁻¹ and 278.7 kJ mol⁻¹, respectively. In contrast, the binding energy of $(\text{FeCp})_2@G_{333}\text{-D}_1$ is a bit smaller, ~256 kJ mol⁻¹ (see Fig. 2a and Table 1). Such distinct binding characters can be simply understood from the electron transfer between Fe atoms and graphene. On one hand, with an increase in the distance between two FeCp ligands, the total charges transferred from the Fe atom to graphene also increase (see Table 1), resulting in an enhancement in the covalent bonding between the Fe atom and graphene. On the other hand, such energetic characters agree well with the Fe–G distances discussed above, that is, the larger FeCp ligand distances in $(\text{FeCp})_2@G_{333}\text{-D}_2$ and $(\text{FeCp})_2@G_{333}\text{-D}_3$ cause their coulomb repulsion to decrease,

which is the cause of their shorter Fe–graphene distances and larger stabilities.

3.1.2 Electronic and magnetic properties. Different from pristine graphene, the incorporation of FeCp ligands in $(\text{FeCp})_n@G_{333}$ s causes it to become ferromagnetic (FM) (see Fig. 2c). The magnetic moments of $\text{FeCp}@G_{333}\text{-S}$, $(\text{FeCp})_2@G_{333}\text{-D}_1$, $(\text{FeCp})_2@G_{333}\text{-D}_2$ and $(\text{FeCp})_2@G_{333}\text{-D}_3$ are 0.81 μ_B , 2.00 μ_B , 1.64 μ_B and 2.00 μ_B , respectively, in which the local magnetic moments of the Fe atoms are 0.73 μ_B , 1.07/1.09 μ_B , 0.82/0.82 μ_B , and 1.04/1.02 μ_B , respectively. Moreover, the antiferromagnetic (AFM) states are also extended, in which the AFM state of $(\text{FeCp})_2@G_{333}\text{-D}_1$ is 0.10 eV higher in energy than the FM state. The calculated density of states projected onto the atomic orbitals is plotted in Fig. 3a–d. In $(\text{FeCp})_n@G_{333}$ s, the states around the Fermi energy level are mainly contributed by the d electrons of the Fe atom, which indicates that the magnetic moments of the studied systems are mainly contributed by the unpaired d electrons of the Fe atom. Corresponding to $\text{FeBz}@SLG$ (2.0 μ_B),^{26,28} the magnetic moment of $\text{FeCp}@G_{333}\text{-S}$ is largely reduced. As shown in Fig. S3 in the ESI,[†] the d_{xz} electron on the major channel is shifted above the Fermi level, which leads to decrease in the magnetic moment by about 1 μ_B . Except for $(\text{FeCp})_2@G_{333}\text{-D}_2$, the spin polarized band gaps are opened in the Dirac points of graphene for the other three systems (see Fig. 3e–h), *i.e.*, the band gaps are about 0.78 eV, 0.40 eV and 0.61 eV for $\text{FeCp}@G_{333}\text{-S}$, $(\text{FeCp})_2@G_{333}\text{-D}_1$ and $(\text{FeCp})_2@G_{333}\text{-D}_3$, respectively. Particularly, $(\text{FeCp})_2@G_{333}\text{-D}_1$ is completely transformed into an intrinsic semiconductor, whereas $(\text{FeCp})_2@G_{333}\text{-S}$ and $(\text{FeCp})_2@G_{333}\text{-D}_3$ can be regarded as n-type semiconductors and can be produced to be intrinsic semiconductors by shifting their Fermi levels into the band gap.

3.2 Structural, electronic and magnetic properties of $(\text{FeCp})_n@G_{444}$ s, $n = 1, 2$

3.2.1 Geometries and stabilities. Similar to $(\text{FeCp})_n@G_{333}$ s, eight different configurations are considered for the $(\text{FeCp})_n@G_{444}$ s isomers: (i) $\text{FeCp}@G_{444}\text{-S}_1$, in which the Fe atom is fixed on the hollow site of graphene (see Fig. S2a and d in the ESI[†]); (ii) $(\text{FeCp})_2@G_{444}\text{-S}_2$, in which two FeCp ligands are adsorbed on adjacent hollow sites in the graphene slab,



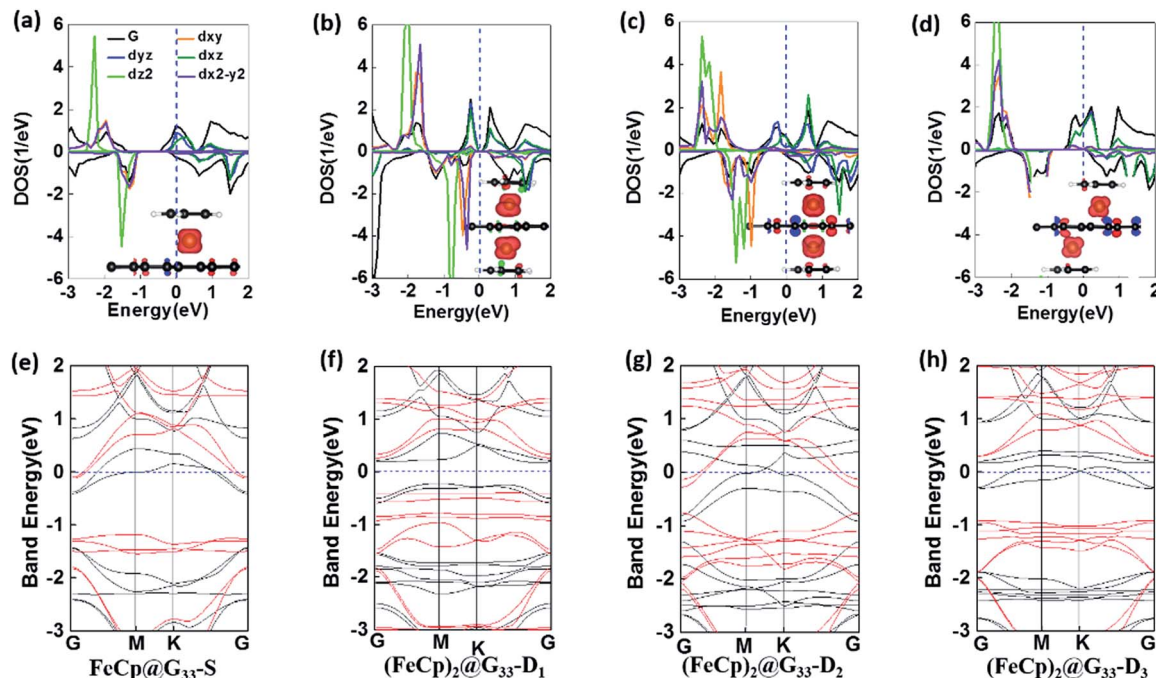


Fig. 3 (a–d) Projected density of states (DOS) and (e–h) electronic band structures of $(\text{FeCp})_n@G_{33}\text{-S}$, $n = 1, 2$. The black and red lines correspond to the spin-up and spin-down channels, respectively. Insets are the spin densities plots.

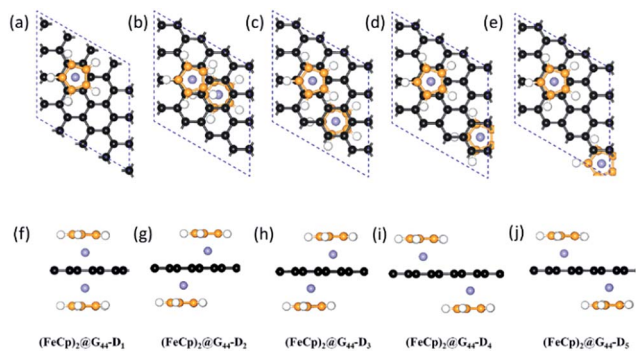


Fig. 4 (a–j) Top and side views of the optimized structures of $(\text{FeCp})_2@G_{44}\text{-D}_i$, $i = 1-5$. Black and orange balls represent the carbon atoms of graphene and Cp ligands; and white and purple balls represent H and Fe atoms, respectively. The complete descriptions of $(\text{FeCp})_2@G_{44}\text{-D}_4$ (c) and $(\text{FeCp})_2@G_{44}\text{-D}_5$ (d) are depicted in the 2×2 supercells in Fig. S1(c, d, g and f) in the ESI†

forming two FeCp-lines along one zigzag edge of graphene spaced by three rows of hexagons (see Fig. S2b and e in the ESI†); (iii) $(\text{FeCp})_2@G_{44}\text{-S}_3$, isomer of configuration (ii) with the FeCp ligands forming a large hexagonal loop (see Fig. S2c and f in the ESI†); (iv) $(\text{FeCp})_2@G_{44}\text{-D}_1$, isomer of $(\text{FeCp})_2@G_{33}\text{-D}_1$ (see Fig. 4a and f) with the G_{33} slab displaced by a G_{44} slab; (v) $(\text{FeCp})_2@G_{44}\text{-D}_2$, isomer of $(\text{FeCp})_2@G_{33}\text{-D}_2$ (see Fig. 4b and g); (vi) $(\text{FeCp})_2@G_{44}\text{-D}_3$, isomer of $(\text{FeCp})_2@G_{33}\text{-D}_3$ (see Fig. 4c and h); (vii) $(\text{FeCp})_2@G_{33}\text{-D}_3$, isomer of $(\text{FeCp})_2@G_{33}\text{-D}_3$ (see Fig. 4d and i); and (viii) $(\text{FeCp})_2@G_{44}\text{-D}_5$, isomer of configuration (iv)–(vii) having two-side FeCp ligands spaced by three carbon hexagons (see Fig. 4e and j).

Significant geometry distortions are found for the FeCp ligands in $(\text{FeCp})_2@G_{44}\text{-S}_2$ and $(\text{FeCp})_2@G_{44}\text{-S}_3$ (see Fig. S2b, c, e and f†), that is, the adjacent Cp ligands are largely tilted with the inclination angles of $\sim 10^\circ$ due to the van der Waals interaction. Differently, normal sandwich configurations are favoured for all the other systems. The C–C bond lengths of the Cp ligands and graphene are about 1.42–1.44 Å and 1.41–1.45 Å (see Table S1 in the ESI†), respectively. $D_{\text{Fe-CpS}}$ and $D_{\text{Fe-GS}}$ are in the range of 1.68–1.71 Å and 1.60–1.65 Å, respectively. Different from $\text{FeCp}@G_{44}\text{-S}_1$, the $D_{\text{Fe-GS}}$ of $\text{FeCp}@G_{44}\text{-S}_{2,3}$ and $(\text{FeCp})_2@G_{44}\text{-Ds}$ are a bit longer due to the enhanced atom electrostatic interaction of the FeCp ligands (see Table 1). All the $(\text{FeCp})_n@G_{44}\text{S}$ are chemically stable with high binding energies of 256.6–285.9 kJ mol^{-1} . Among them, $(\text{FeCp})_2@G_{44}\text{-S}_1$ has the largest binding energy of $\sim 285.9 \text{ kJ mol}^{-1}$, followed by $(\text{FeCp})_2@G_{44}\text{-D}_3$ and $(\text{FeCp})_2@G_{44}\text{-D}_4$ with values of $\sim 281 \text{ kJ mol}^{-1}$. In contrast, $(\text{FeCp})_2@G_{44}\text{-S}_2$ has the smallest binding energy of 256.6 kJ mol^{-1} . On one hand, the stability of these FeCp@SLGs isomers is slightly dependent on the coverage concentration. For example, the binding energies of $(\text{FeCp})_2@G_{33}\text{-D}_1$ and $(\text{FeCp})_2@G_{44}\text{-D}_1$ are 256 kJ mol^{-1} and 257.4 kJ mol^{-1} , respectively, whereas that of $(\text{FeCp})_2@G_{33}\text{-D}_2$ and $(\text{FeCp})_2@G_{44}\text{-D}_2$ is 277.6 kJ mol^{-1} and 277.8 kJ mol^{-1} , respectively. On the other hand, the stabilities of these $(\text{FeCp})_n@G_{44}\text{S}$ are largely dependent on the relative positions of FeCp ligands. For example, the binding energies of $(\text{FeCp})_2@G_{44}\text{-S}_2$, $(\text{FeCp})_2@G_{44}\text{-S}_3$ and $(\text{FeCp})_2@G_{44}\text{-D}_1$ per unit cell are much smaller than that of $\text{FeCp}@G_{44}\text{-D}_i$ ($i = 3, 4$ and 5) by about 13–24 kJ mol^{-1} .

3.2.2 Electronic and magnetic properties. Except $(\text{FeCp})_2@G_{44}\text{-S}_1$, which is nonferromagnetic, all the other



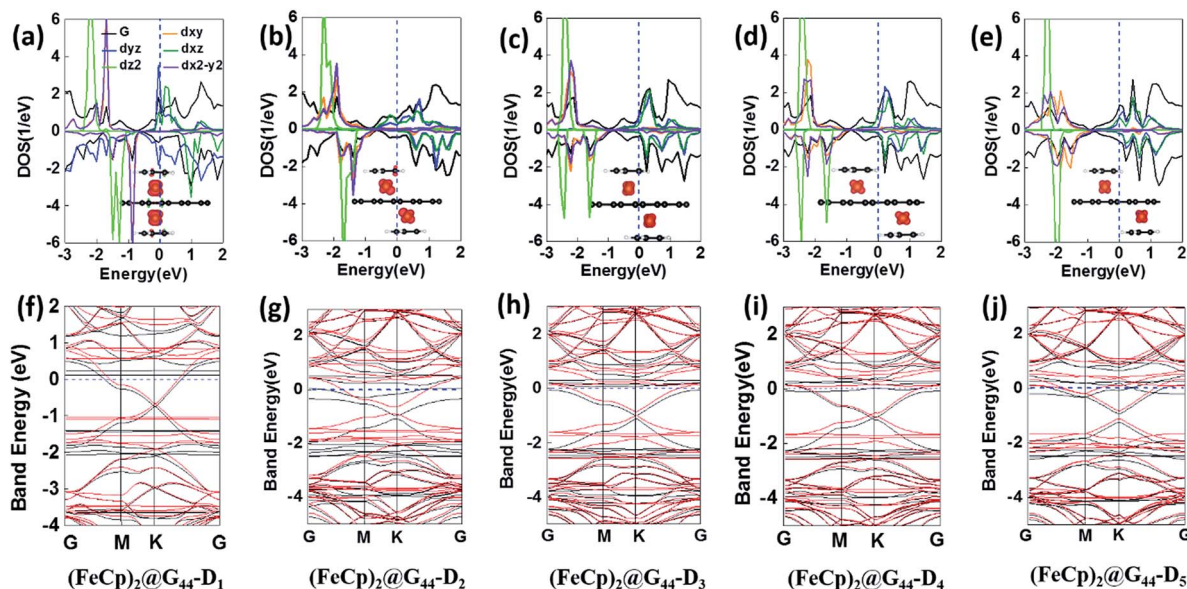


Fig. 5 (a–e) Density of states (DOS) and (f–j) band structure plots of $(\text{FeCp})_2@G_{44}s$. Insets are the spin densities plots.

$(\text{FeCp})_n@G_{44}s$ systems have ferromagnetic ground states, in which the AFM state of $(\text{FeCp})_2@G_{44}-D_1$ is less stable since it is about 0.11 eV higher in energy. Similar to $(\text{FeCp})_2@G_{33}-D_1$, $(\text{FeCp})_2@G_{44}-D_1$ has the largest magnetic moment of $\sim 1.30 \mu_B$, whereas that of $(\text{FeCp})_2@G_{44}-S_2$, $(\text{FeCp})_2@G_{44}-S_3$ and $(\text{FeCp})_2@G_{44}-D_2$ are $0.74 \mu_B$, $0.94 \mu_B$ and $0.97 \mu_B$, respectively. As for the others, their magnetic moments are relative smaller, which are about $0.50 \mu_B$, $0.49 \mu_B$ and $0.61 \mu_B$ for $(\text{FeCp})_2@G_{44}-D_i$ ($i = 3, 4$ and 5) respectively. Similar to $(\text{FeCp})_n@G_{33}s$, the ferromagnetic properties of these FeCp@SLGs are mainly derived from the unpaired d electrons of the Fe atoms (see Fig. 5f–j, S4†). Moreover, spin polarized band structures are found for a few $(\text{FeCp})_n@G_{44}s$ systems, in which a band gap of 0.25 eV, 0.12 eV and 0.36 eV is opened in the Dirac point for $(\text{FeCp})_2@G_{44}-S_2$, $(\text{FeCp})_2@G_{44}-S_3$ and $(\text{FeCp})_2@G_{44}-D_5$ due to the breaking of the intrinsic electronic symmetry of pristine graphene. The electronic properties of these $(\text{FeCp})_n@G_{44}s$ are largely sensitive to the adsorption densities of the FeCp ligands. For example, the n-type semiconducting $\text{FeCp}@G_{33}-S/(\text{FeCp})_2@G_{33}-D_1$ and the semiconducting $(\text{FeCp})_2@G_{33}-D_1$ are completely transformed into metals in their $(\text{FeCp})@G_{44}s$ isomers (see Table 2).

3.3 Structural, electronic and magnetic properties of other $(\text{TMCP})_n@Gs$, $n = 1, 2$, TM = Co, Ni

For comparison, we studied the electronic and magnetic properties of a number of $(\text{TMCP})_n@SLGs$ ($n = 1, 2$), TM = Co, Ni (see Fig. S5 and S6 in the ESI†). Different from $(\text{FeCp})_n@SLGs$, the Ni atoms prefer to sit close to the top site of graphene (see Fig. S6 in the ESI†) due to their stable valence electron configuration ($3d^{10}4s^2$). Regarding $(\text{FeCp})_n@SLGs$, the binding energies of $(\text{CoCp})_n@SLGs$ and $(\text{NiCp})_n@SLGs$ are relative smaller, which are around 164.4 – $178.3 \text{ kJ mol}^{-1}$ and 248.7 – $259.1 \text{ kJ mol}^{-1}$, respectively (see Table 2). This is due to the fact that the charges transferred from Co/Ni to graphene are much smaller than that of the Fe atoms (see Table 2 and Fig. S7 in the ESI†). On the other hand, different from $(\text{CoCp})_n@SLGs$ and $\text{FeCp}@SLGs$, the charge distribution on the Ni atoms and graphene in $\text{NiCp}@Gs$ is uneven, leading to their distinct interactions between the single-side adsorption and double-side adsorption.

Furthermore, the electronic and magnetic properties of these TMCp ligands-functionalized graphene are also dependent on the TM element. Different from $(\text{FeCp})_n@s$, the magnetic moments of $(\text{CoCp})_n@Gs$ and $(\text{NiCp})_n@Gs$ are almost not affected by the adsorption position of the TMCp ligand. For

Table 2 Systems (Sys), the distance of the TM atom to Cp ligands ($D_{\text{TM}-\text{Cp}}$) and graphene ($D_{\text{TM}-\text{G}}$), charges transferred from the Fe atom to graphene (Δe), binding energies (E_b), magnetic moments (MM) and band gaps (Δ_g) at the Dirac point of $(\text{CoCp})_n@Gs$ and $(\text{NiCp})_n@Gs$

Sys	$D_{\text{Fe}-\text{Cp}}$ (Å)	$D_{\text{Fe}-\text{G}}$ (Å)	Δe (e)	E_b (kJ mol^{-1})	MM (μ_B)	Δ_g (eV)
$\text{CoCp}@G_{44}-S_1$	1.80	1.70	0.28	178.3	1.94	0
$(\text{CoCp})_2@G_{44}-S_2$	1.79, 1.80	1.86, 1.86	0.50	164.4	3.92	2.0
$(\text{CoCp})_2@G_{44}-D_1$	1.71, 1.75	1.68, 1.68	0.55	172.8	3.81	1.83
$(\text{NiCp})_2@G_{33}-D_1$	1.75, 1.75	1.97, 2.03	0.36	259.1	2.0	2
$\text{NiCp}@G_{44}-S_1$	1.76	1.93	0.18	249.7	1.0	0
$(\text{NiCp})_2@G_{44}-S_2$	1.76, 1.77	2.06, 2.01	0.38	248.8	2.0	0.74
$(\text{NiCp})_2@G_{44}-D_5$	1.76, 1.76	1.93, 1.90	0.40	248.7	2.0	0.94



example, the magnetic moments per unit cell are $1.94 \mu_B$, $3.92 \mu_B$ and $3.81 \mu_B$ for $\text{CoCp}@G_{44}\text{-S}_1$, $(\text{CoCp})_2@G_{44}\text{-S}_2$ and $(\text{CoCp})_2@G_{44}\text{-D}_1$, respectively (see Table 2). As for the $(\text{NiCp})_n@G_s$ isomers, their magnetic moments are $1.0 \mu_B$ for $\text{NiCp}@G_{44}\text{-S}_1$ and $2.0 \mu_B$ for $(\text{NiCp})_2@G_{44}\text{-S}_2$, $(\text{NiCp})_2@G_{44}\text{-D}_1$ and $(\text{NiCp})_2@G_{44}\text{-D}_5$. The AFM states of $(\text{NiCp})_2@G_{33}\text{-D}_1$ and $(\text{CoCp})_2@G_{44}\text{-D}_1$ are less stable than that of the FM state by about 0.13 eV and 0.12 eV higher in energy. On the other hand, $\text{CoCp}@G_{44}\text{-S}_1$, $(\text{CoCp})_2@G_{44}\text{-D}_1$ (see Fig. S8d and f in the ESI†) and $(\text{NiCp})_2@G_{33}\text{-D}_1$ (see Fig. S9g in the ESI†) are metals. Differently, $\text{NiCp}@G_{44}\text{-S}_1$ is an intrinsic semiconductor with a band gap of 0.17 eV (see Fig. S9e in the ESI†). The others, $(\text{CoCp})_2@G_{44}\text{-S}_2$ (see Fig. S8e in the ESI†), $(\text{NiCp})_2@G_{44}\text{-S}_2$ (see Fig. S9f in the ESI†) and $(\text{NiCp})_2@G_{44}\text{-D}_5$ (see Fig. S9h in the ESI†) are half metals with half metal band gaps of 0.24 eV, 0.13 eV and 0.29 eV, respectively, which can be used for spintronic device applications.

4. Conclusions

The electronic and magnetic properties of half organic–metal sandwich ligands-functionalized monolayer graphene, $(\text{FeCp})_n@$ -SLGs ($n = 1, 2$), are systematically studied. All the $(\text{FeCp})_n@$ -SLGs are thermodynamically stable with the covalent bonding of TMCP ligands to graphene. Versatile electronic and magnetic properties are found for $(\text{FeCp})_n@$ -SLGs, depending on the adsorption sites and coverage concentrations. Particularly, $(\text{FeCp})_2@G_{33}\text{-D}_1$ is transformed into an intrinsic semiconductor with a band gap of 0.40 eV due to the breaking of the intrinsic electronic symmetry of pristine graphene. By displacing Fe by Co or Ni elements, distinct changes to the magnetic and electronic properties are found.

Acknowledgements

This study is supported by the NSFC (11574262, 11647316, 21503006). The authors thank the computational resources at the YZU.

References

- W. Han, R. K. Kawakami, M. Gmitra and J. Fabian, *Nat. Nanotechnol.*, 2014, **9**, 794–807.
- K. S. Novoselov, A. K. Geim, S. V. Morozov, D. Jiang, Y. Zhang, S. V. Dubonos, I. V. Grigorieva and A. A. Firsov, *Science*, 2004, **306**, 666–669.
- Y. Xu, Z. Li and W. Duan, *Small*, 2014, **10**, 2182–2199.
- Y. Zhang, Y. W. Tan, H. L. Stormer and P. Kim, *Nature*, 2005, **438**, 201–204.
- F. Bonaccorso, L. Colombo, G. Yu, M. Stoller, V. Tozzini, A. C. Ferrari, R. S. Ruoff and V. Pellegrini, *Science*, 2015, **347**, 1246501.
- M. J. Allen, V. C. Tung and R. B. Kaner, *Chem. Rev.*, 2010, **110**, 132–145.
- M. F. Craciuna, S. Russob, M. Yamamoto and S. Taruchac, *Nano Today*, 2011, **6**, 42–60.
- Y. Zhu, S. Murali, W. Cai, X. Li, J. W. Suk, J. R. Potts and R. S. Ruoff, *Adv. Mater.*, 2010, **22**, 3906–3924.
- A. C. Ferrari, J. C. Meyer, V. Scardaci, C. Casiraghi, M. Lazzeri, F. S. Mauri, D. Piscanec, K. Jiang, S. Novoselov, S. Roth and A. K. Geim, *Phys. Rev. Lett.*, 2006, **97**, 187401.
- H. L. Wang, J. T. Robinson, X. L. Li and H. J. Dai, *J. Am. Chem. Soc.*, 2009, **131**, 9910–9911.
- A. Reina, X. Jia, J. Ho, D. Nezich, H. Son, V. Bulovic, M. S. Dresselhaus and J. Kong, *Nano Lett.*, 2009, **9**, 30–35.
- A. Ismach, C. Druzgalski, S. Penwell, A. Schwartzberg, M. Zheng, A. Javey, J. Bokor and Y. Zhang, *Nano Lett.*, 2010, **10**, 1542–1548.
- L. Gao, J. R. Guest and N. P. Guisinger, *Nano Lett.*, 2010, **10**, 3512–3516.
- X. S. Li, W. W. Cai, J. H. An, S. Kim, J. Nah, D. X. Yang, R. D. Piner, A. Velamakanni, I. Jung and E. Tutuc, *Science*, 2009, **324**, 1312–1314.
- S. Y. Zhou, G. H. Gweon, A. V. Fedorov, P. N. First, W. A. D. Heer, D. H. Lee, F. Guinea, A. H. Castro Neto and A. Lanzara, *Nat. Mater.*, 2007, **6**, 770–775.
- L. Britnell, R. V. Gorbachev, R. Jalil, B. D. Belle, F. Schedin, A. Mishchenko, T. Georgiou, M. I. Katsnelson, L. Eaves, S. V. Morozov, N. M. R. Peres, J. Leist, A. K. Geim, K. S. Novoselov and L. A. Ponomarenko, *Science*, 2012, **335**, 947–950.
- V. V. Chaban and O. V. Prezhdo, *Nanoscale*, 2016, **8**, 15521–15528.
- X. F. Fan, Z. X. Shen, A. Q. Liu and J. L. Kuo, *Nanoscale*, 2012, **4**, 2157–2165.
- M. Y. Han, B. Ozyilmaz, Y. Zhang and P. Kim, *Phys. Rev. Lett.*, 2007, **98**, 206805.
- Y. Chen, D. G. de Oteyza, Z. Pedramrazi, C. Chen, F. R. Fischer and M. F. Crommie, *ACS Nano*, 2013, **7**, 6123–6128.
- P. A. Denis and F. Iribarne, *J. Phys. Chem. C*, 2011, **115**, 195–203.
- V. V. Chaban and O. V. Prezhdo, *J. Phys. Chem. Lett.*, 2015, **6**, 4397–4403.
- P. A. Denis, *J. Phys. Chem. C*, 2009, **113**, 5612–5619.
- B. Trzaskowski, L. Adamowicz, W. Beck, K. Muralidharan and P. A. Deymier, *J. Phys. Chem. C*, 2013, **117**, 19664–19671.
- H. M. Le, H. Hajime, Y. Kwwazoe and D. Nguyen-Manh, *Phys. Chem. Chem. Phys.*, 2013, **15**, 19395–19404.
- P. Plachinda, D. R. Evans and R. Solanki, *J. Chem. Phys.*, 2011, **135**, 044103.
- S. M. Avdoshenko, L. N. Loffe, G. Cuniberti, L. Dunsch and A. A. Popov, *ACS Nano*, 2011, **5**, 9939–9949.
- X. Y. Zhang, Y. J. Bian, W. K. Sun and Y. J. Liu, *RSC Adv.*, 2016, **6**, 97953–97960.
- X. Y. Yao, X. Y. Zhang, X. Y. Ye and J. L. Wang, *Phys. Chem. Chem. Phys.*, 2016, **18**, 22390–22398.
- H. Valencia, A. Gil and G. Frapper, *J. Phys. Chem. C*, 2015, **119**, 5506–5522.
- L. Shen, S. W. Yang, M. F. Ng, V. Ligatchev, L. P. Zhou and Y. P. Feng, *J. Am. Chem. Soc.*, 2008, **130**, 13956–13960.
- Y. C. Li, X. C. Chen, G. Zhou, W. H. Duan, Y. Kim, M. Kim and J. Ihm, *Phys. Rev. B: Condens. Matter Mater. Phys.*, 2011, **83**, 195443.



- 33 H. E. Kang and A. Pramanik, *J. Chem. Phys.*, 2011, **135**, 124708.
- 34 X. J. Liu, C. Z. Wang, Y. X. Yao, W. C. Lu, M. Hupalo, A. C. Tringides and K. M. Ho, *Phys. Rev. B: Condens. Matter Mater. Phys.*, 2011, **83**, 235411.
- 35 H. Johll and H. C. kang, *Phys. Rev. B: Condens. Matter Mater. Phys.*, 2009, **79**, 245416.
- 36 J. Hu, H. Alicea, R. Q. Wu and M. Franz, *Phys. Rev. Lett.*, 2011, **109**, 266801.
- 37 H. Johll and H. C. Kang, *Phys. Rev. B: Condens. Matter Mater. Phys.*, 2009, **79**, 245416.
- 38 I. Zanella, S. B. Fagan, R. Mota and A. Fazzio, *J. Phys. Chem. C*, 2008, **112**, 9163–9167.
- 39 X. L. Liu, C. Z. Wang, Y. X. Yao, W. C. Lu, M. Hupalo, M. C. Tringides and K. M. Ho, *Phys. Rev. B: Condens. Matter Mater. Phys.*, 2011, **83**, 235411.
- 40 G. Kresse and J. Hafner, *Phys. Rev. B: Condens. Matter Mater. Phys.*, 1993, **48**, 13115.
- 41 G. Kresse and J. Furthmuller, *Comput. Mater. Sci.*, 1996, **6**, 15.
- 42 J. P. Perdew, K. Burke and M. Ernzerhof, *Phys. Rev. Lett.*, 1998, **80**, 891.
- 43 P. E. Blochl, *Phys. Rev. B: Condens. Matter Mater. Phys.*, 1994, **50**, 17953.
- 44 G. Kresse and D. Joubert, *Phys. Rev. B: Condens. Matter Mater. Phys.*, 1999, **59**, 1758.
- 45 S. Grimme, *J. Comput. Chem.*, 2006, **27**, 1787.

