

REVIEW

[View Article Online](#)
[View Journal](#) | [View Issue](#)
CrossMark
click for updatesCite this: *RSC Adv.*, 2017, **7**, 11987

Received 2nd February 2017
 Accepted 14th February 2017

DOI: 10.1039/c7ra01363a
rsc.li/rsc-advances

Tailored performance of layered transition metal dichalcogenides *via* integration with low dimensional nanostructures

De-Sheng Liu,^a Jiang Wu,^b Yanan Wang,^c Haining Ji,^{*a} Lei Gao,^a Xin Tong,^c Muhammad Usman,^c Peng Yu,^{*c} and Zhiming Wang^{*c}

Transition metal dichalcogenides (TMDs) with a unique sandwich structure have attracted tremendous attention in recent years due to their distinctive electrical and optical properties. The performance of TMDs can be tailored by integration with low dimensional nanostructures which possess quantum size effects and large specific surface areas. In this review, we give a brief description of integrated structures, their synthesis methods, and highlight the applications in hydrogen production, lithium batteries, sensors and detectors.

1. Introduction

Since the advent of graphene in 2004, two-dimensional (2D) materials with large specific surface areas and excellent electronic and optical properties have attracted significant attention in various fields such as energy storage, sensing, electronic and optoelectronic devices.^{1–7} Among them, transition metal dichalcogenides (TMDs, *e.g.* MoS₂, MoSe₂, WS₂, WSe₂) are gradually becoming hotspots of research due to their novel structure and remarkable properties.^{8,9} These materials present a MX₂ stoichiometry, where M is a transition metal and X is a chalcogenide.¹⁰ They possess a unique 2D X–M–X structure, in which the monolayer consists of transition metal atoms and chalcogen atoms interacting with strong covalent bonds, while adjacent TMDs layers are formed by weak van der Waals forces.^{11,12} TMDs with semiconducting properties have a tuneable bandgap around 1–2 eV, and the bandgap becomes direct when the thickness of crystals decreases to a single layer, indicating that TMDs are promising materials in optoelectronic applications.^{13–15}

TMDs-based heterostructures with outstanding properties and novel phenomena have been widely investigated.^{16,17} These heterostructures not only inherit the properties of individual components, but also present innovative properties, such as Moiré pattern-induced wave function localization in MoS₂/MoSe₂ heterostructures and larger on/off ratio in graphene/MoS₂ heterostructures-based transistor.^{18–20} With the development of

these heterostructures, new properties can be designed and tailored beyond the capability of conventional TMDs.²¹ These TMDs-based heterostructures can be integrated *via* various routes, including 0D quantum dots (QDs) and 1D nanotubes (NTs),²² or 0D QDs and 2D nanosheets (NSs),²³ or 1D nanorods (NRs) and 2D NSs as well.²⁴ Here, we focus on heterostructures composed of 2D TMD NSs and low dimensional nanostructures (LDNs).

LDNs such as nanodots (NDs), nanoparticles (NPs), nanowires (NWs) and nanobelts (NBs) with quantum size effects and large specific surface areas usually exhibit novel properties differ from their bulks.^{25–27} For instance, Nb₂O₅ shows an increased band gap from 3.4 eV to 4.2 eV when its particle size decreases to nanoscale.²⁸ Owing to tuneable optical properties and the effect of surface chemistry, some materials in low dimension show enhanced photoluminescence (PL), high sensitivity to detect gas or biological molecules, absorptivity of dye molecules, and improved charge transport ability.^{29–32} Therefore, a method of fabricating TMD-based heterostructures *via* integrating with LDNs is a potential approach to enhance performance in corresponding devices.^{21,33,34}

Although many comprehensive reviews regarding TMDs have been reported,^{16,17,35–38} there is still no systematic review with respect to LDNs/TMDs integration. In this review, we began with a brief description of integrated structures. We then outlined the synthesis methods of TMDs NSs and provided two main routes: self-assembly and hydrothermal/solvothermal to prepare LDNs/TMDs nanostructures. Subsequently, we highlight their applications and device performance. In the last part, a short summary is given.

2. Integrated structures

2.1 Low dimensional nanostructures

When the material size shrinks into nanoscale, the number of atoms or molecules forming the material becomes an

^aState Key Laboratory of Electronic Thin Film and Integrated Devices, University of Electronic Science and Technology of China, Chengdu 610054, P. R. China. E-mail: hainingji@uestc.edu.cn

^bDepartment of Electronic and Electrical Engineering, University College London, Torrington Place, London WC1E 7JE, UK

^cInstitute of Fundamental and Frontier Science, University of Electronic Science and Technology of China, Chengdu 610054, P. R. China. E-mail: azarisyu@gmail.com; Zhimingwang@uestc.edu.cn



influential role to the physicochemical properties and reactivity.³⁹ In order to simplify the size and dimensional effects, it is preferable to classify the materials according to the number of dimension with expansion freedom (not restricted in nano-size). For example, NDs and NPs, of which all dimensions are in the range of a few nanometres, belong to 0D category; and 3D materials usually refer to bulk materials. For the same material, the allotropes usually exhibit dimension-dependent distinctive properties. One of the typical instances is the allotropes of carbon materials. For a given carbon concentration, 1D multi-wall carbon nanotubes (MWNTs) has higher photomechanical stress response and energy conversion efficiency than 2D single layer graphene and 3D graphite under illumination of near-infrared (NIR) light.⁴⁰ In another case, silicon materials with different dimension show different optical properties. While the bulk silicon does not emit visible light due to the indirect bandgap, the low dimensional Si nanostructures with the quantum-confinement effect are able to emit visible light.^{41,42} Besides, Si NWs structures have a small reflection in a wide spectrum range compared to 2D thin films, which improves the optical absorption in the high-frequency regime.⁴³ In some metal materials, such as Au and Ag, when the dimension of metal crystals decreases to nanoscale, these nanostructures arrays demonstrate surface plasmon resonance phenomenon under light illumination.^{44,45} The collective motion of conduction band electrons related to fixed positive ions give rise to a plasmon oscillation in a metal under incident light.⁴⁶ This unique optical property provides some inspirations for applications such as biosensors and solar cells.⁴⁷

In LDNs materials, a small change of size will alter its optical properties due to quantum confinement effect.^{48–50} For instance, Xue *et al.* demonstrated that the TiO₂ NPs with a size varied from 6.8 nm to 14.2 nm have a large difference on Raman intensity, which can be attributed to the charge-transfer resonance between the TiO₂ and the adsorbed molecules.⁵¹ Similarly, owing to the size effect, an absorption peak redshifts of Au NPs is observed when the particle diameter increased from 20 nm to 100 nm.⁵² Therefore, the metal NPs with localized surface plasmon resonances can be employed to develop refractive index sensors through combining with microfiber.⁵³ In addition, some LDNs materials also exhibit distinct chemical, electrical or biological properties due to the size effect.^{54–56} For example, CdSe QDs employed for biological application, the smaller sized QDs have a greater improvement in enzymatic activity compared with the larger QDs.⁵⁷ In general, these materials in low dimension with unique characteristics are applied in a range of fields such as catalysis, sensing, electronic and optoelectronic.^{58,59}

2.2 Electronic and optical properties of 2D TMDs

TMDs materials are comprised of a transition metal element of groups IV – X and a chalcogen element such as sulfur (S) and selenium (Se).⁶⁰ Bulk TMDs exhibit a wide variety of polymorphs as compared to other 2D counterparts, and they cover a wide spectrum of properties such as semiconductivity, half-metallic magnetism and superconductivity.⁶¹ When the d-orbitals of

transition metal are partially occupied, TMDs display metallic conductivity properties, whereas these materials are semiconductors when the orbitals are fully occupied.¹⁰ In bulk phase of TMDs, there are three possible stacking structures: 1T, 2H and 3R with tetragonal symmetry, hexagonal closed packing and rhombohedral symmetry, respectively.⁹ However, monolayer TMDs are demonstrated to possess two polytypes, metallic 1T phase and semiconducting 2H phase, and the stacking structures are shown in Fig. 1.⁶² It has been demonstrated that the 2H stacking structure can be converted into 1T stacking structure by Li intercalation or plasmonic hot electron induced structure phase transition.^{5,63} The metallic 1T MoS₂ employed as electrode materials exhibits outstanding capacitive behaviour in both aqueous and organic electrolytes due to their intrinsic hydrophilicity and high electrical conductivity.⁶⁴

Semiconducting TMDs such as MoS₂ and WS₂ with a tunable band gap promote their applications in photoelectric field.^{9,65,66} With decreasing layer number, the band gap present an increasing trend, and the indirect band gap of bulk TMDs gradually shift to be direct when reduced to single layer.⁶⁷ Fig. 2 shows a band structure of bulk and monolayer MoS₂ calculated from first principles.⁶⁸ At the Γ -point, bulk MoS₂ has an indirect band gap of 1.2 eV, which can be transferred to a direct band gap of 1.9 eV in monolayer form. This phenomenon can be explained by quantum confinement effect.^{14,69} Owing to the direct band gap, a strong PL is observed in monolayer MoS₂, and the PL quantum yield (QY) has a 10⁴ fold enhancement compared with their bulk crystal.⁷⁰ Moreover, the PL of monolayer MoS₂ can be effectively improved *via* an air-stable, solution-based chemical treatment and the consequent QY is more than 95%.⁷¹ Therefore, single layer semiconducting TMDs with direct band gap indicates that they are promising candidates in optoelectronic applications.⁷²

2.3 LDNs/TMDs heterostructures

The integration of materials in different dimensionalities is able to combine the inherent advantages and mitigate the genetic disadvantages in terms of fundamental properties.⁷³ Noble metals, metal oxides, graphite derivatives crystals with sizes in nanoscale such as Au NPs and carbon NTs (CNTs) can be integrated with 2D TMDs NSs. The as-prepared heterostructures present an obviously enhanced performance in

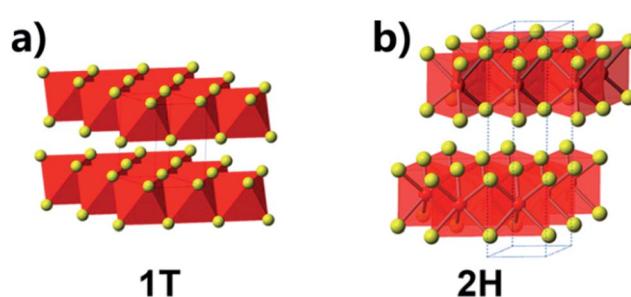


Fig. 1 Crystal structures of (a) 1T and (b) 2H crystal structures of the TMDs. Adapted with permission from ref. 6. Copyright 2013 American Chemical Society.



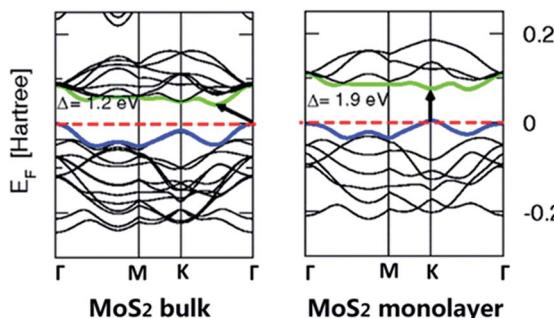


Fig. 2 Band structures of bulk MoS₂ and its monolayer. The arrows indicate the smallest value of the band gap (direct or indirect) for a given system. The top of valence band (blue) and bottom of conduction band (green) are highlighted. Adapted with permission from ref. 68. Copyright 2011 by the American Physical Society.

electrical and optical application.^{74–77} For instance, Zhou *et al.* prepared TiO₂ NBs@MoS₂ NSs heterostructure *via* a hydrothermal reaction, and the heterostructure exhibits highly photocatalytic hydrogen production.⁷⁸ Herein, the synergistic effect of TiO₂ and MoS₂ resulting in the heterostructure provides efficient charge separation to suppress the electron–hole recombination, and further display an improved photocatalytic activity. In another case, Zhang *et al.* reported a hydrothermal method to fabricate a ZnO NRs/MoS₂ NSs heterostructure without a catalyst.⁷⁹ Due to the light antennae effect of ZnO NRs, enhanced Raman and PL emission are observed in this type of heterostructure.

According to the dimension of nanostructures, these heterostructures can be divided into 0D/2D and 1D/2D, as shown in Fig. 3. Generally, 0D nanostructures usually tiled on TMDs NSs, and 1D nanostructure usually served as an axis for TMDs NSs, which were loaded on their surface in form of 1D/2D.^{21,80–82} Every individual structure in these heterostructures plays an important role and contributes to tailored performance of the whole system. Layered TMDs have several crystal phases with different lattice structures and electronic character such as 1T and 2H, although integrated with the same LDNs, these heterostructures exhibit distinct performance.^{83,84} Bai *et al.* demonstrated that TiO₂–MoS₂ (1T) hybrid structure exhibited excellent performance as a co-catalyst to product hydrogen, which is in sharp contrast to TiO₂–MoS₂ (2H).⁸⁵ This can be attributed to the active sites on the 1T-MoS₂ basal plane, which gives rise to a high charge diffusion rate and reduces the probability of charge recombination when integrated with TiO₂. On the contrary, 2H-MoS₂ does not provide reaction sites and may block the active sites on the TiO₂ nanocrystals surface. In addition, the layer numbers can affect the electrical and optical properties of TMDs. For example, monolayer MoS₂ has a large enhancement of the luminescence compared with the bulk material.¹³ Therefore, layer number and crystal phase of TMDs are indispensable considerations when integrated with LDNs. Besides, the morphology and distribution of LDNs have impact on the system performance.^{86,87} For instance, Wang *et al.* prepared various morphologies of spherical, triangle, clover-like and flower-like AuNPs (Fig. 4) by changing the ratio of MoS₂ and

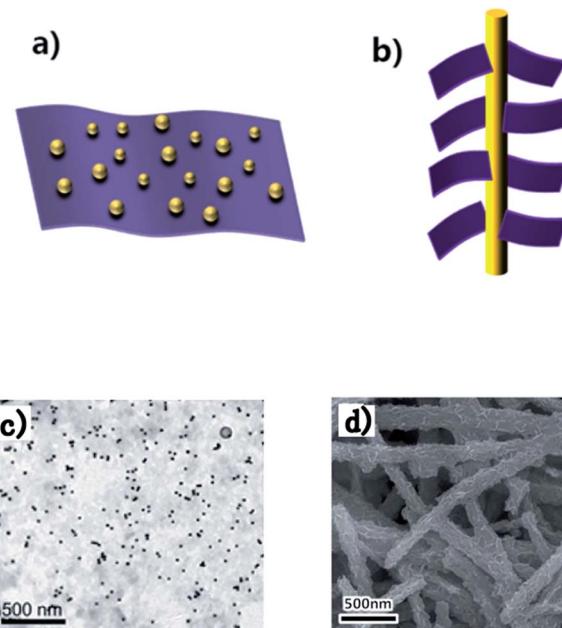


Fig. 3 Schematic of integrated heterostructure (a) 0D/2D and (b) 1D/2D. SEM images of (c) 0D Au NPs/2D MoS₂ NSs and (d) 1D TiO₂ NBs/2D MoSe₂ NSs. Adapted from ref. 75 and 89 with permission of The Royal Society of Chemistry.

thionine (Thi), and demonstrated the clover-like Au NP/Thi-MoS₂ nanocomposites-based electrochemical immunosensor exhibit an excellent selectivity, which can detect carcinoembryonic antigen (CEA) as low as 0.52 pg mL^{–1}.⁸⁸

3. Synthesis methods

Preparation of high quality TMDs NSs is an important process in their heterostructures fabrication. Both physical and chemical methods can be considered to obtain layered TMDs, such as exfoliation and chemical vapor deposition (CVD).^{21,90–94} For the synthesis of integrated nanostructures, different approaches have already been developed, as concluded in Table 1, which can be classified into two main methods: self-assembly and hydrothermal/solvothermal.

3.1 Layer TMDs

To date, exfoliation of layered solid and CVD growth method are mainstream strategies to obtain layered TMDs. Typically, mechanical exfoliation use micromechanical force to separate bulk materials and obtain single or few layers NSs, but the yield is relatively low and the size is quite small.⁶⁰ In comparison, chemical exfoliation can conquer these problems and obtain large-scale TMDs NSs. Zheng *et al.* used metal naphthalene as intercalating agents to exfoliate metal chalcogenides and obtained high quality sheets with large size up to 400 μm².⁹⁵ The size distribution of the flakes is much better than the sheets obtained by using organolithium salts, and the prepared TMDs sheets could be directly jet-printed on substrate. Shear exfoliation can also prepare TMDs NSs by separating bulk materials.



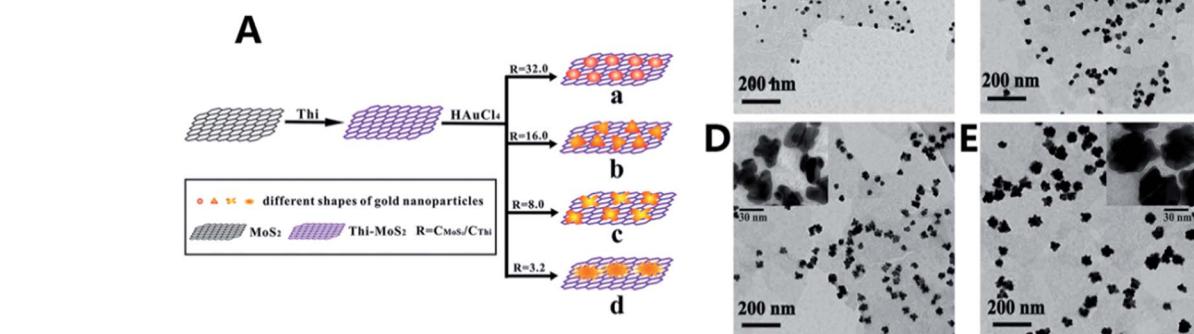


Fig. 4 (A) Schematic of preparation process of shape-controlled Au NPs supported on Thi- MoS_2 NSs. TEM images of AuNP-Thi- MoS_2 nanocomposites with (B) spherical, (C) triangle, (D) clover-like and (E) flower-like AuNPs. Adapted from ref. 88 with permission of The Royal Society of Chemistry.

Table 1 Summary of synthesis methods for integrated nanostructure

Dimensionality	Synthesis methods	Integrated nanostructures	Applications	Ref.
0D/2D	Self-assembly	Ag NPs- WS_2 Au NPs- MoS_2	Supercapacitor Sensor	102 103
	Hydrothermal	TiO_2 NPs- MoS_2 TiO_2 -RGO- MoS_2 ZnO - MoS_2	Hydrogen production Hydrogen production Hydrogen production	85 104 105
	Solvothermal	SnO_2 - MoS_2 GQDs- MoS_2	Sensor LIBs	106 107
	Electrodeposition	Au NPs- MoS_2	Sensor	108
	Thermal evaporation	Au NPs- MoS_2	Sensor or detector	82
	Drop casting	Pd NPs- MoS_2	Sensor	109
	Successive ion layer adsorption and reaction method	CdS NPs- WS_2	Water splitting	110
	Self-assembly	Au NRs- MoS_2	Hydrogen production	111
	Hydrothermal	C/TiO ₂ NBS- MoSe_2	LIBs	89
	Solvothermal	CoS ₂ NWs- MoS_2 CNTs- MoS_2 SnO_2 NTs- MoS_2	Hydrogen production Hydrogen production Hydrogen production	112 77 113
1D/2D				

Varrla *et al.* demonstrated the processing parameters for specific NSs concentration, length, and thickness and obtained large-scale shear-exfoliation of MoS_2 NSs (in the range ~ 40 –220 nm for length and ~ 2 –12 layers for thickness) in aqueous surfactant solution by using a kitchen blender.⁹⁶ CVD is a typical method to prepare high-quality TMDs layers with controllable thickness and scalable size.⁹⁹ Lee *et al.* has utilized CVD approach to synthesis large-area MoS_2 NSs with six-fold symmetry hexagonal lattice and high crystallinity on SiO_2/Si substrates.⁹⁷ In addition to preparing layered TMDs with certain thickness, CVD growth method provide a facile way to directly grow layered heterostructures, which eliminates interfacial contamination interfused during synthesis processes.⁹⁸ Moreover, wet chemical approach is another available method to prepared layered TMDs, which can be directly obtained in solution with a high yield.⁹⁹ However, improved purity and controllability over nanostructures are primary challenges

because of infancy.^{100,101} In general, chemical exfoliation is desirable method for mass production of TMDs NSs, because other methods usually have more or less limitations to some extent, such as low efficiency, long and complex process, requirement for expensive systems and severe synthetic conditions.¹¹⁴

3.2 LDNs/TMDs nanostructures

3.2.1 Self-assembly. Self-assembly approach offers a route for assembling various functional components and preserve lattice integrity.¹¹⁵ These discrete components spontaneously integrate to be a complete heterostructure *via* directly and/or indirectly specific interactions.¹¹⁶ The main synthesis process is immersing as-prepared NSs into nanostructure precursor solution and using available auxiliary technique to assist LDNs to be decorated on TMDs NSs. For instance, Shi and coworker



reported a simple method to decorate Au NPs on chemical exfoliation MoS₂ NSs and CVD-prepared MoS₂ NSs.⁸⁷ The CVD-MoS₂ NSs on sapphire substrates are directly immersed in AuCl₃ for 20 seconds, while the chemical exfoliation NSs need to disperse on water surface and then added several drops of AuCl₃. Au ions in the solution are reduced, leading to the anchoring of Au NPs on MoS₂ surface due to active nucleation sites of NSs.

Self-assembly is also available for metal oxide/TMDs heterostructure synthesis. Bai *et al.* prepared a TiO₂-MoS₂ (1T) hybrid structure *via* sonication.⁸⁸ The typical procedure is mixing the dispersion of MoS₂ NSs with aqueous dispersion of TiO₂ nanocrystals and then sonicated for 3 min. In another case, Sun *et al.* prepared a p-type semiconductor of Cu₂O decorated with MoS₂ NSs by using hexadecyl trimethylammonium bromide (CTAB) as the surfactant to control the morphology.¹¹⁷ As-prepared nanocomposite exhibits efficient solar hydrogen production under visible light.

Size distribution, morphology and density of metal NPs can be tuned by changing the defects density in TMDs layers.^{87,118} Lu *et al.* reported that a focused laser beam can achieve active surface domains in MoS₂ film, and the size of as-prepared Au NPs are controllable, which were preferentially decorated onto the modified regions.¹⁰³ The rough surface of laser-pruned MoS₂ could trap the precursor and provide initial nucleation sites for Au NP growth.

3.2.2 Hydrothermal/solvothermal. Hydrothermal and solvothermal methods are widely used to prepare the composite structures. The reaction is usually conducted in a heated solution beyond ambient temperature and pressure.¹¹⁹ Zhang *et al.* demonstrated a facile one-step hydrothermal synthesis to prepare Au NPs/MoS₂.¹²⁰ In an aqueous solution, hydrazine act as strong reductant reduced HAuCl₄ to Au NPs attached onto MoS₂ NSs with the S-Au bond. Su *et al.* reported a microwave-assisted hydrothermal method to prepare Au NPs/MoS₂ in a microwave reactor with a heated condition.¹²¹ In a typical synthesis process, MoS₂ NSs solution were first mixed with sodium citrate tribasic dehydrate, Tween and HAuCl₄·3H₂O, and then heated to 60 °C for 5 min in a microwave reactor. Finally, the target product of Au NPs/MoS₂ can be obtained with purification *via* a centrifugation step.

CNTs/TMDs heterostructure is widely studied in recent years due to their remarkable catalytic activity.^{77,122,123} Using *in situ* solid and hydrothermal reactions, Zhang *et al.* prepared Mo₂C-embedded N-doped CNTs/MoS₂ nanostructure.¹²⁴ In MoO₂/Mo₂C-NCNTs aqueous solution, O²⁻ ions are replaced with S²⁻ ions which resulted in the transformation of MoO₂ to MoS₂. Metal oxide/TMDs nanocomposites can be prepared through this method as well. Li *et al.* demonstrated that with the assistance of glucose, TiO₂ NWs can act as a backbone for the nucleation and growth of MoS₂ NSs.¹²⁵ As a binder, glucose suppressed the growth and stack of layered MoS₂, facilitating the growth of MoS₂ NSs with smaller sizes on the TiO₂ NWs surface.^{125,126} Liu *et al.* prepared a SnO₂@MoSe₂ nanostructure through facile electrospinning technique and solvothermal method.¹²⁷ SnO₂ NFs are first obtained *via* a simple electrospinning method, then mixed with Se solution and dispersed in

Na₂Mo₄ solution, which is subsequently heated to 180 °C and maintained for 12 h. The black precipitates of SnO₂@MoSe₂ after natural cooling exhibit excellent electrochemical activity as catalyst.

Wu *et al.* reported a green technique to prepare the WS₂ and MoS₂ QDs/NSs composites by sonication and solvothermal treatment.¹¹⁴ Typically, bulk MoS₂/WS₂ powders were sonicated in *N*-dimethylformamide for 3 h, and then stirred the top 2/3 of the dispersion for 6 h at 140 °C. After settled the suspensions for several hours, the light yellow supernatant of MoS₂/WS₂ QDs and the centrifugate of composite of MoS₂/WS₂ QDs and NSs composites were separated. In addition, Li *et al.* prepared a new type of 1D/2D heterostructures: GeS NWs are grown as seeds followed by GeS NSs epitaxially growing at the pre-grown NWs.⁷³ The pre-grown NWs exposing to air caused a mild degree of oxidation at the NW surface, which could prompt the nucleation of the NSs. This strategy opens a novel avenue to synthesis NWs/NSs heterostructures.

4. Applications

4.1 Hydrogen production

Hydrogen is regarded as a potential fuel in the future to address the environmental crises and energy shortage issues.^{128,129} Utilizing electricity to split water is a green and effective approach to produce hydrogen. In general, it can be divided into two strategies: electrocatalytic and photoelectrocatalytic.^{130,131} Noble metals such as Pt, Pd, Ru and Au are widely known as effective electrochemical catalysts due to their excellent catalytic activity.¹³² Pt and its alloys have usually been the first choice of catalysts, however, scarcity and high cost severely hinder its practical application.^{133,134} TMDs are regarded as ideal candidates to replace Pt due to high chemical stability and excellent electrocatalytic properties.¹³⁵ For instance, nanoscale MoS₂ exhibits a good reduction capacity for protons due to the quantum confinement effect, and the marginal edges served as active sites can promote the decomposition of water.⁷⁶ Xie *et al.* demonstrated defect-rich MoS₂ NSs could present excellent hydrogen evolution reaction (HER) activity with a low onset overpotential of 120 mV and small Tafel slope of 50 mV dec⁻¹.¹³⁶ However, low conductivity of TMDs largely limited the charge transfer rate and affected the total efficiency.¹³⁷ An efficient way to address this issue is integrating TMDs nanostructures with conductive substrates,¹²⁴ such as graphene, carbon NTs or carbon NFs. Guo *et al.* prepared a N-doped CNFs/defect and S-rich MoS₂ heterostructure for high efficiency HER with a small overpotential of 135 mV at 10 mA cm⁻² and a large cathodic current density of 65.6 mA cm⁻² at 200 mV (Fig. 5).¹²² Nitrogen doping promoted the conductivity of composites, and the embedded mode of MoS₂ NSs and CNFs greatly improved their stability and catalytic activity. CoS₂ NW/MoS₂ NSs supported on carbon cloth also exhibits excellent HER performance and prominent electrochemical durability.¹¹² However, most of the conductive substrates are inert for the HER and will decrease the density of active sites of TMDs NSs, therefore, balancing the active sites and conductivity of electrochemical catalysts remains a knotty problem.¹²⁴ On the other hand,



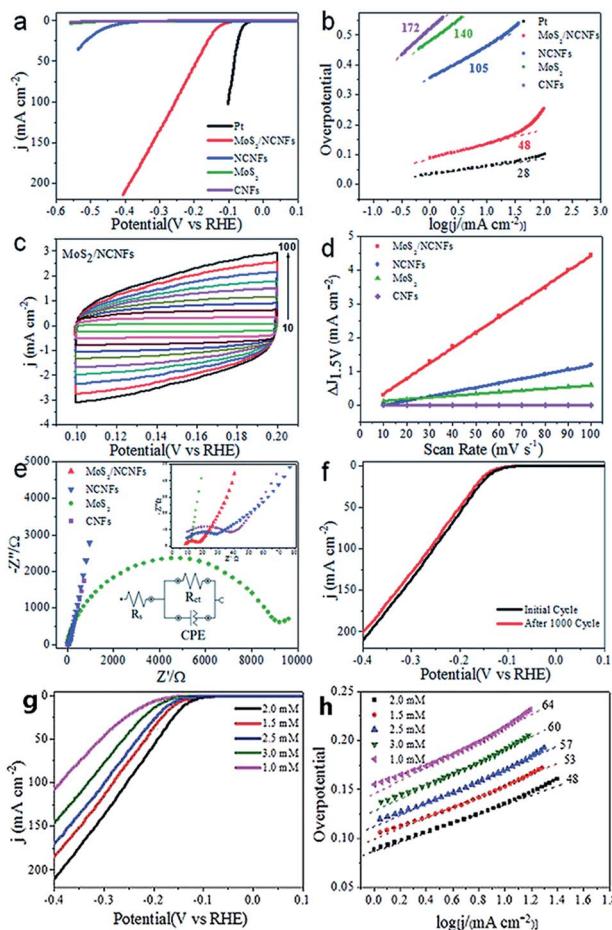


Fig. 5 Electrocatalytic performance of MoS₂/NCNFs. (a) Polarization curves, (b) Tafel plots, (c) electrochemical cyclic voltammogram of MoS₂/NCNFs at different potential scanning rates, (d) the C_{dl} of different materials obtained at 0.15 V vs. RHE, (e) Nyquist plots and (f) stability of the MoS₂/NCNFs with the initial polarization curve (black curve) and after 1000 cycles (red curve). (g) Polarization curves (g) and (h) the corresponding Tafel plots of MoS₂/NCNFs prepared at different concentration of Mo precursor. Adapted from ref. 122 with permission of The Royal Society of Chemistry.

integrated with noble metal nanostructures, TMDs can achieve a high electrocatalytic performance toward HER.¹³⁸ For instance, Au NPs coated on TMDs NSs can largely improve the charge transport along interplanar directions and inhibited restacking, which improved the accessibility of protons to the catalytically active sites.¹¹⁸ Hou *et al.* reported the growth of Pt NPs/MoS₂ NSs on the surface of carbon fiber and demonstrated an overpotential of -5 mV and a corresponding Tafel slope of 53.6 mV dec⁻¹.¹³⁹ The large specific surface area of TMDs not only effectively collect and transport charges, but also provides plenty of nucleation sites for the growth of metal NPs.¹³⁸

In 1972, Fujishima *et al.* demonstrated a green approach to split water utilizing TiO₂ as a photochemical catalyst, which offers a promising way for clean and low-cost production of hydrogen *via* solar energy.^{140,141} However, the rapid recombination of the electron–hole pairs and low activation ability of the surface reduced the amount of charge, resulting in a low photocatalytic efficiency.⁸⁵ Introducing TMDs as co-catalysts which

serve as the reaction sites can promote the charge separation and transport.¹⁰⁵ Indeed, MoS₂ NSs coated on a TiO₂ surface significantly facilitated the separation of the photo-induced carriers, which resulted in a higher photocurrent and improved broad spectrum photocatalytic properties.^{142,143} Moreover, Zhou *et al.* demonstrated that nitrogen self-doped MoS₂ NSs integrated with MnO₂ NBs exhibited excellent performance with an onset potential of -156 mV (vs. RHE), a large current density, a small Tafel slope of 47.5 mV dec⁻¹ and outstanding electrochemical durability.¹⁴⁴ Nitrogen doping provides a localized state near the Fermi energy in MoS₂ band gap and thus leads to a high electronic conductivity of the heterostructure. Besides, electron-rich sites are beneficial to attract H⁺ to the catalyst surface and gains an enhanced HER activity. Plasmon-excited hot electrons injection mechanism is also an efficient and promising route for the improvement of catalyst activity.¹¹¹ However, achieving a faster electrons injection from plasmonic metal to semiconductor than the recombination process is still a significant challenge.

4.2 Lithium battery

Rechargeable Li batteries with high energy density, good cycle life, and blameless power performance has become one of the most important energy storage devices in recent years.¹⁴⁵ To date, silicon, metal alloys, metal oxides and TMDs have been widely explored as the replacement of graphite anodes in Li-ion batteries (LIB).^{146–148} The high surface-to-volume ratio and short diffusion path of TMD NSs render it as a brilliant electrode material for LIBs. For instance, Hwang *et al.* reported graphene-like MoS₂ nanoplates showing a reversible capacity of 912 mA h g⁻¹ at a rate of 1C and excellent cycling stability when fabricated in LIBs.¹⁴⁹ To further improve the cycling stability and rate capability performance of TMDs, integrating LDNs such as Au NPs and TiO₂ NPs is an effective approach.^{120,150} Guo *et al.* demonstrated that GQDs/MoS₂ exhibits excellent electrochemical lithium storage properties with a high reversible capacity (1099 mA h g⁻¹ at 100 mA g⁻¹), good cyclic stability, as well as excellent rate performance (660 mA h g⁻¹ at 5000 mA g⁻¹) as shown in Fig. 6.¹⁰⁷ GQDs acted as binder enhanced cyclic stability and facilitate the charge transfer within the electrode. Furthermore, it can enlarge the interlayer distance between MoS₂ layers which is beneficial for Li⁺ intercalation. Zhou *et al.* demonstrated that CNFs@MoS₂ coaxial NFs showed an excellent Li storage properties, such as high specific capacity, good cycling stability, and superior high-rate capability.¹⁵¹ The synergistic effect of the two active components leads to a larger capacity and better cyclability.

TiO₂ with superior safety and rate capability exhibits excellent performance when employed as anodes materials.¹⁵² Li *et al.* demonstrated that TiO₂@MoS₂ heterostructure displayed a high initial discharge capacity of 862 mA h g⁻¹, and it maintained at 544 mA h g⁻¹ even after 100 cycles.¹²⁵ The synergistic effects of TiO₂ NW and ultrathin MoS₂ shorten the Li⁺ diffusion paths for rapid charge transfer, and improved cycling stability of the system at high rates. Xue *et al.* demonstrated a Fe₃O₄ NPs/MoS₂ nanostructures as an anode for LIBs with a superior cyclic



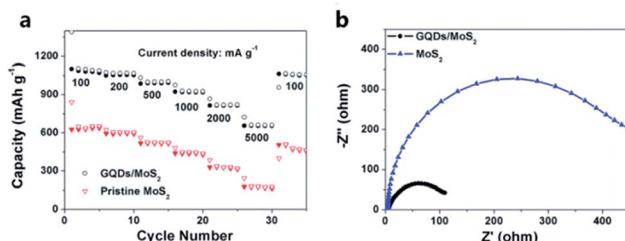


Fig. 6 (a) Rate capabilities and (b) electrochemical impedance spectroscopy spectra of GQDs/MoS₂ and pristine MoS₂ electrodes. Adapted from ref. 107 with permission of The Royal Society of Chemistry.

and rate performances.¹⁵³ Fe₃O₄ NPs serving as spacers and effectively stabilize the integrated structure, facilitating the MoS₂ surfaces accessible to electrolyte during charge/discharge processes, thereby hindering the fast fading of reversible capacities. The heterostructure contribute a 1033 and 224 mA h g⁻¹ at current densities of 2000 and 10 000 mA g⁻¹, respectively. The MoS₂ NSs provided a high reversible capacity and Fe₃O₄ NPs offered a high rate performance. In addition, MoS₂ has been reported as cathode material for lithium–oxygen (Li–O₂) batteries, which exhibits a 85% round-trip efficiency and good reversibility.¹⁵⁹ Table 2 lists some typical examples comprised of different components for lithium storage performances.

TMDs anodes have been studied for many years, yet there are still many challenges. The large irreversible capacity was observed in the first cycle, indicating that a large portion of Li⁺ ions could not be utilized again in the following cycles, and the relatively high voltage behavior of the TMDs limited its applications.¹⁶⁰ Nevertheless, TMDs is a promising anodes material owing to superior capacity and rate capability.⁸

4.3 Sensor and detector

TMDs materials with a tuneable band gap applied in sensors and detectors usually exhibited high selectivity and rapid detection.^{161–164} Using first-principles calculations, Yue *et al.* found that small gas molecules can be physically adsorbed on monolayer MoS₂, and determined the most stable adsorption configuration and adsorption energy.¹⁶⁵ The band structure of monolayer MoS₂ will not significantly be altered though small gas molecules are absorbed. Many reports have demonstrated

TMDs-based sensors, when integrated with noble metal or metal oxide semiconductors, exhibiting an enhanced performance compared with individual component.^{166–168} For instance, He *et al.* demonstrated that Pt NPs deposited on MoS₂ NSs exhibits high sensitivity for NO₂ detection.¹⁶⁹ Upon the adsorption of gas, the Schottky barrier between Pt NPs and MoS₂ channel is changed, which may help explain the increased sensitivity of MoS₂ thin film.^{170,171} In addition, Kuru *et al.* demonstrated Pd NPs/MoS₂ NS could show a highly sensitivity for H₂ detection at room temperature and the sensor exhibits a response of around 10 toward 50 000 ppm H₂ with a response and recovery time of 40 and 83 s, respectively.¹⁰⁹ TMDs-based biosensor can effectively and sensitively detect biomolecule.^{172,173} Using molecular dynamics and density functional theory simulation, Aluru *et al.* demonstrated that monolayer MoS₂ is an promising material for DNA sequencing technology.¹⁷⁴ Thiol modified monolayer MoS₂ integrated with Au NPs can largely enhance the DNA adsorption about 1 order of magnitude.¹⁷⁵ MoS₂ NSs exhibited different PL properties toward single-strained DNA *versus* double strained DNA when the concentration of DNA solution increased, which can be used for rapid and selective detection of DNA nucleoside chains. In another case, Su *et al.* reported a Au NPs/MoS₂ biosensor that could detect glucose in a concentration range from 10 to 300 μM with a detection limit of 2.8 μM.¹⁷⁶ Furthermore, the sensor also displays good reproducibility and long-term stability, indicating its potential to be applied for determination of glucose concentration in human serum. In addition, Shao *et al.* reported an electrochemical dopamine (DA) sensor based on Au NPs@MoS₂ nanocomposite, and the detection concentration of DA ranges from 0.1 to 200 mM, with a detection limit of 80 nM (S/N = 3),¹²¹ and the sensor can also detect DA in ascorbic acid existed environment.

Monolayer MoS₂ with a direct bandgap of 1.9 eV is extremely attractive for photodetector. Kis *et al.* reported a monolayer MoS₂ photodetector with a photoresponsivity of 880 A W⁻¹ at a wavelength of 561 nm and a photoresponse in the 400–680 nm range.¹⁷⁷ However, the slow responsivity is an important limiting factor for application. To address this situation, Kufer *et al.* demonstrated PbS QDs decorated MoS₂ for photodetector could achieve a high responsivity of 6 × 10⁵ A W⁻¹, and the responsivity is higher than individual PbS QDs and MoS₂-based detectors.¹⁷⁸ Owing to the effect of reverse electric induced by increasing separated charge carriers, the responsivity presents a decrease trend with increasing optical intensity. In addition,

Table 2 The lithium storage performances of TMDs based anode materials

Anode materials	Current rate	Initial capacity (mA h g ⁻¹)	Capacity retention (mA h g ⁻¹)/(cycles)	Ref.
MoS ₂	1C	1062	907/50	149
Graphene-MoS ₂	100 mA g ⁻¹	1571	1187/100	154
CNFs–MoS ₂	100 mA g ⁻¹	1489	1264/50	151
GQDs-MoS ₂	100 mA g ⁻¹	1394	1031/80	107
TiO ₂ –MoS ₂	0.1C	588	350/100	155
Carbon-TiO ₂ –MoS ₂	100 mA g ⁻¹	941.6	805/100	156
Polyaniline–MoS ₂	100 mA g ⁻¹	1063.9	~953/50	157
SnO ₂ –MoS ₂	0.2C	1306	~900/80	158



metal plasmon LDNs with near-field oscillation and scattering effect can effectively improve the photocurrent response and light absorption of TMDs, which can be applied to exploit plasmon resonance-based devices.¹⁷⁹ Although great progress has been made in this area, the mechanism and principles of electron transfer of TMDs is remain unclear.¹⁸⁰

5. Conclusion

In summary, the integration of LDNs and TMD exhibits a tailored performance due to synergistic effects. LDN materials with dissimilar properties that differ from their bulk counterparts effectively mitigate the genetic disadvantages of TMDs, while TMD NSs with controllable size and thickness provided suitable surface areas for LDNs to integrate. For LDN/TMD heterostructures, many synthesis methods such as self-assembly and hydrothermal methods have been developed for preparation *via* one step or multi-step approaches. The shape and distribution of LDNs can be controlled by changing the component ratio of precursor solution or modifying the surface of TMDs and have a large impact on the system performance. Through incorporating different types of LDNs, TMD heterostructures with unique properties have been applied in various areas. As a promising candidate to replace Pt catalyst for hydrogen production, TMDs integrated with LDNs solve the limitation of low conductivity, and shows a high chemical stability and excellent electrocatalytic activities. As electrode materials for LIBs, the heterostructure contributes a high reversible capacity and an excellent rate capability. Owing to the high surface-to-volume ratio and tunable band gap of TMDs, the sensors combine with LDNs exhibit a high selectivity to detect biomolecules and gas molecules, and have great impacts in the field of photodetection.

Although extensive investigations have focused on LDNs/TMDs heterostructures, several limitations still exist. Firstly, LDNs reported in most papers are noble metals, metal oxides, metal dichalcogenides and graphite derivatives, the fields may expand to organic materials in future. Secondly, a method to prepare TMD NSs with excellent morphology and desired properties is still necessary. Heteroatom doping can offer TMDs with various new or improved optical and structural properties,¹⁸¹ which may provide some inspiration for further development of LDNs/TMDs heterostructures. Finally, how the spatial organization and distribution affecting the system performances have not been fully investigated yet.¹⁸² Many reports are restricted to concern the component of the heterostructure, therefore more studies are still required to optimize the fabricated structures. Nonetheless, the study of TMDs is just in the infant stage and a greater potential is waiting for exploration.

Acknowledgements

The authors gratefully acknowledge funding from National Natural Science Foundation of China 51272038, National Natural Science Foundation of China 51302030, National Natural Science Foundation of China 61474015 and National Program on Key Basic Research Project (973 Program) 2013CB933301.

References

- 1 F. Bonaccorso, L. Colombo, G. Yu, M. Stoller, V. Tozzini, A. C. Ferrari, R. S. Ruoff and V. Pellegrini, *Science*, 2015, **347**, 1246501.
- 2 X. Wang, G. Sun, P. Routh, D. H. Kim, W. Huang and P. Chen, *Chem. Soc. Rev.*, 2014, **43**, 7067–7098.
- 3 F. Xia, H. Wang, D. Xiao, M. Dubey and A. Ramasubramaniam, *Nat. Photonics*, 2014, **8**, 899–907.
- 4 D. Jariwala, V. K. Sangwan, L. J. Lauhon, T. J. Marks and M. C. Hersam, *ACS Nano*, 2014, **8**, 1102–1120.
- 5 M. Xu, T. Liang, M. Shi and H. Chen, *Chem. Rev.*, 2013, **113**, 3766–3798.
- 6 S. Z. Butler, S. M. Hollen, L. Cao, Y. Cui, J. A. Gupta, H. R. Gutiérrez, T. F. Heinz, S. S. Hong, J. Huang, A. F. Ismach, E. Johnston-Halperin, M. Kuno, V. V. Plashnitsa, R. D. Robinson, R. S. Ruoff, S. Salahuddin, J. Shan, L. Shi, M. G. Spencer, M. Terrones, W. Windl and J. E. Goldberger, *ACS Nano*, 2013, **7**, 2898–2926.
- 7 L. Tao, D. Wang, S. Jiang, Y. Liu, Q. Xie, H. Tian, N. Deng, X. Wang, Y. Yang and T.-L. Ren, *J. Semicond.*, 2016, **37**, 041001.
- 8 H. Li, Y. Shi, M.-H. Chiu and L.-J. Li, *Nano Energy*, 2015, **18**, 293–305.
- 9 V. Sorkin, H. Pan, H. Shi, S. Y. Quek and Y. W. Zhang, *Crit. Rev. Solid State Mater. Sci.*, 2014, **39**, 319–367.
- 10 M. Chhowalla, H. S. Shin, G. Eda, L. J. Li, K. P. Loh and H. Zhang, *Nat. Chem.*, 2013, **5**, 263–275.
- 11 L. Yuwen, F. Xu, B. Xue, Z. Luo, Q. Zhang, B. Bao, S. Su, L. Weng, W. Huang and L. Wang, *Nanoscale*, 2014, **6**, 5762–5769.
- 12 J. A. Wilson and A. D. Yoffe, *Adv. Phys.*, 1969, **18**, 193–335.
- 13 K. F. Mak, C. Lee, J. Hone, J. Shan and T. F. Heinz, *Phys. Rev. Lett.*, 2010, **105**, 136805.
- 14 Q. H. Wang, K. Kalantar-Zadeh, A. Kis, J. N. Coleman and M. S. Strano, *Nat. Nanotechnol.*, 2012, **7**, 699–712.
- 15 X. Tong, E. Ashalley, F. Lin, H. Li and Z. M. Wang, *Nano-Micro Lett.*, 2015, **7**, 203–218.
- 16 A. K. Geim and I. V. Grigorieva, *Nature*, 2013, **499**, 419–425.
- 17 H. Wang, F. Liu, W. Fu, Z. Fang, W. Zhou and Z. Liu, *Nanoscale*, 2014, **6**, 12250–12272.
- 18 J. Kang, J. Li, S. S. Li, J. B. Xia and L. W. Wang, *Nano Lett.*, 2013, **13**, 5485–5490.
- 19 N. Myoung, K. Seo, S. J. Lee and G. Ihm, *ACS Nano*, 2013, **7**, 7021–7027.
- 20 H. Zheng, Y. Li, H. Liu, X. Yin and Y. Li, *Chem. Soc. Rev.*, 2011, **40**, 4506–4524.
- 21 B. Luo, G. Liu and L. Wang, *Nanoscale*, 2016, **8**, 6904–6920.
- 22 W. T. Sun, Y. Yu, H. Y. Pan, X. F. Gao, Q. Chen and L. M. Peng, *J. Am. Chem. Soc.*, 2008, **130**, 1124–1125.
- 23 A. Cao, Z. Liu, S. Chu, M. Wu, Z. Ye, Z. Cai, Y. Chang, S. Wang, Q. Gong and Y. Liu, *Adv. Mater.*, 2010, **22**, 103–106.
- 24 H. Kim and K. Yong, *Phys. Chem. Chem. Phys.*, 2013, **15**, 2109–2116.
- 25 H. Zhao, G. Sirigu, A. Parisini, A. Camellini, G. Nicotra, F. Rosei, V. Morandi, M. Zavelani-Rossi and A. Vomiero, *Nanoscale*, 2016, **8**, 4217–4226.

- 26 E. A. Anyebe, M. K. Rajpalke, T. D. Veal, C. J. Jin, Z. M. Wang and Q. D. Zhuang, *Nano Res.*, 2014, **8**, 1309–1319.
- 27 G. Vasseur, Y. Fagot-Revurat, M. Sicot, B. Kierren, L. Moreau, D. Malterre, L. Cardenas, G. Galeotti, J. Lipton-Duffin, F. Rosei, M. Di Giovannantonio, G. Contini, P. Le Fevre, F. Bertran, L. Liang, V. Meunier and D. F. Perekhchka, *Nat. Commun.*, 2016, **7**, 10235.
- 28 R. Brayner and F. Bozon-Verduraz, *Phys. Chem. Chem. Phys.*, 2003, **5**, 1457–1466.
- 29 Y. Xia, P. Yang, Y. Sun, Y. Wu, B. T. Mayers, B. D. Gates, Y. Yin, F. Kim and H. Yan, *Adv. Mater.*, 2003, **15**, 353–389.
- 30 C. Frigerio, D. S. Ribeiro, S. S. Rodrigues, V. L. Abreu, J. A. Barbosa, J. A. Prior, K. L. Marques and J. L. Santos, *Anal. Chim. Acta*, 2012, **735**, 9–22.
- 31 R.-H. Tao, J.-M. Wu, H.-X. Xue, X.-M. Song, X. Pan, X.-Q. Fang, X.-D. Fang and S.-Y. Dai, *J. Power Sources*, 2010, **195**, 2989–2995.
- 32 W. Wen, J. M. Wu, Y. Z. Jiang, S. L. Yu, J. Q. Bai, M. H. Cao and J. Cui, *Sci. Rep.*, 2015, **5**, 11804.
- 33 Z. Yin, B. Chen, M. Bosman, X. Cao, J. Chen, B. Zheng and H. Zhang, *Small*, 2014, **10**, 3537–3543.
- 34 B. Li, J. M. Wu, T. T. Guo, M. Z. Tang and W. Wen, *Nanoscale*, 2014, **6**, 3046–3050.
- 35 D. Deng, K. S. Novoselov, Q. Fu, N. Zheng, Z. Tian and X. Bao, *Nat. Nanotechnol.*, 2016, **11**, 218–230.
- 36 X. Chia, A. Y. Eng, A. Ambrosi, S. M. Tan and M. Pumera, *Chem. Rev.*, 2015, **115**, 11941–11966.
- 37 F. Reale, K. Sharda and C. Mattevi, *Applied Materials Today*, 2016, **3**, 11–22.
- 38 Z. He and W. Que, *Applied Materials Today*, 2016, **3**, 23–56.
- 39 W. Guo, J. Yin, H. Qiu, Y. Guo, H. Wu and M. Xue, *Friction*, 2014, **2**, 209–225.
- 40 J. Loomis and B. Panchapakesan, *Nanotechnology*, 2012, **23**, 215501.
- 41 K. Peng, Y. Xu, Y. Wu, Y. Yan, S. T. Lee and J. Zhu, *Small*, 2005, **1**, 1062–1067.
- 42 P. Yu, J. Wu, S. Liu, J. Xiong, C. Jagadish and Z. M. Wang, *Nano Today*, 2016, **11**, 704–737.
- 43 L. Hu and G. Chen, *Nano Lett.*, 2007, **7**, 3249–3252.
- 44 X. Huang, I. H. El-Sayed, W. Qian and M. A. El-Sayed, *J. Am. Chem. Soc.*, 2006, **128**, 2115–2120.
- 45 S. A. Maier and H. A. Atwater, *J. Appl. Phys.*, 2005, **98**, 011101.
- 46 M. R. Jones, K. D. Osberg, R. J. Macfarlane, M. R. Langille and C. A. Mirkin, *Chem. Rev.*, 2011, **111**, 3736–3827.
- 47 J. Wu, P. Yu, A. S. Susha, K. A. Sablon, H. Chen, Z. Zhou, H. Li, H. Ji, X. Niu, A. O. Govorov, A. L. Rogach and Z. M. Wang, *Nano Energy*, 2015, **13**, 827–835.
- 48 S. Eustis and M. A. El-Sayed, *Chem. Soc. Rev.*, 2006, **35**, 209–217.
- 49 B. O. Dabbousi, J. Rodriguez-Viejo, F. V. Mikulec, J. R. Heine, H. Mattoussi, R. Ober, K. F. Jensen and M. G. Bawendi, *J. Phys. Chem. B*, 1997, **101**, 9463–9475.
- 50 V.-V. Truong, P. Courteau and J. Singh, *J. Appl. Phys.*, 1987, **62**, 4863.
- 51 X. Xue, W. Ji, Z. Mao, H. Mao, Y. Wang, X. Wang, W. Ruan, B. Zhao and J. R. Lombardi, *J. Phys. Chem. C*, 2012, **116**, 8792–8797.
- 52 Y. Cui, P. P. Shum, G. Y. Wang, H. Chang, X. Q. Dinh, M. Jiang and G. Humbert, *Sensors*, 2011, 371–374.
- 53 A. Convertino, M. Cuscunà, F. Martelli, M. G. Manera and R. Rella, *J. Phys. Chem. C*, 2014, **118**, 685–690.
- 54 P. Claus, A. Bruckner, C. Mohr and H. Hofmeister, *J. Am. Chem. Soc.*, 2000, **122**, 11430–11439.
- 55 D. M. Sagar, R. R. Cooney, S. L. Sewall, E. A. Dias, M. M. Barsan, I. S. Butler and P. Kambhampati, *Phys. Rev. B: Condens. Matter Mater. Phys.*, 2008, 77.
- 56 D. Gao, H. Zhou, J. Wang, S. Miao, F. Yang, G. Wang, J. Wang and X. Bao, *J. Am. Chem. Soc.*, 2015, **137**, 4288–4291.
- 57 K. Das, K. Rawat, R. Patel and H. B. Bohidar, *RSC Adv.*, 2016, **6**, 46744–46754.
- 58 P. Yu, J. Wu, E. Ashalley, A. Govorov and Z. Wang, *J. Phys. D: Appl. Phys.*, 2016, **49**, 365101.
- 59 B. Hamawandi, M. Noroozi, G. Jayakumar, A. Ergül, K. Zahmatkesh, M. S. Toprak and H. H. Radamson, *J. Semicond.*, 2016, **37**, 102001.
- 60 S. A. Han, R. Bhatia and S.-W. Kim, *Nano Convergence*, 2015, 2.
- 61 C. Xia and J. Li, *J. Semicond.*, 2016, **37**, 051001.
- 62 W. Zhao, R. M. Ribeiro and G. Eda, *Acc. Chem. Res.*, 2015, **48**, 91–99.
- 63 Y. Kang, S. Najmaei, Z. Liu, Y. Bao, Y. Wang, X. Zhu, N. J. Halas, P. Nordlander, P. M. Ajayan, J. Lou and Z. Fang, *Adv. Mater.*, 2014, **26**, 6467–6471.
- 64 M. Acerce, D. Voiry and M. Chhowalla, *Nat. Nanotechnol.*, 2015, **10**, 313–318.
- 65 K. F. Mak, K. He, J. Shan and T. F. Heinz, *Nat. Nanotechnol.*, 2012, **7**, 494–498.
- 66 Z. Yin, H. Li, H. Li, L. Jiang, Y. Shi, Y. Sun, G. Lu, Q. Zhang, X. Chen and H. Zhang, *ACS Nano*, 2013, **6**, 74–80.
- 67 S. Bhattacharyya and A. K. Singh, *Phys. Rev. B: Condens. Matter Mater. Phys.*, 2012, 86.
- 68 A. Kuc, N. Zibouche and T. Heine, *Phys. Rev. B: Condens. Matter Mater. Phys.*, 2011, **83**, 245213.
- 69 A. Splendiani, L. Sun, Y. Zhang, T. Li, J. Kim, C. Y. Chim, G. Galli and F. Wang, *Nano Lett.*, 2010, **10**, 1271–1275.
- 70 K. F. Mak, C. Lee, J. Hone, J. Shan and T. F. Heinz, *Phys. Rev. Lett.*, 2010, **105**, 136805.
- 71 M. Amani, D. Lien, D. Kiriya, J. Xiao, A. Azcatl, J. Noh, S. R. Madhvapathy, R. Addou, K. C. Santosh and M. Dubey, *Science*, 2015, **350**, 1065–1068.
- 72 P. Tonndorf, R. Schmidt, P. Bottger, X. Zhang, J. Borner, A. Liebig, M. Albrecht, C. Kloc, O. Gordan, D. R. Zahn, S. Michaelis de Vasconcellos and R. Bratschitsch, *Opt. Express*, 2013, **21**, 4908–4916.
- 73 C. Li, Y. Yu, M. Chi and L. Cao, *Nano Lett.*, 2013, **13**, 948–953.
- 74 B. Mukherjee and E. Simsek, *Plasmonics*, 2015, **11**, 285–289.
- 75 X. Yang, W. Liu, M. Xiong, Y. Zhang, T. Liang, J. Yang, M. Xu, J. Ye and H. Chen, *J. Mater. Chem. A*, 2014, **2**, 14798.
- 76 B. Zhu, B. Lin, Y. Zhou, P. Sun, Q. Yao, Y. Chen and B. Gao, *J. Mater. Chem. A*, 2014, **2**, 3819.
- 77 Y. Yan, X. Ge, Z. Liu, J. Y. Wang, J. M. Lee and X. Wang, *Nanoscale*, 2013, **5**, 7768–7771.
- 78 W. Zhou, Z. Yin, Y. Du, X. Huang, Z. Zeng, Z. Fan, H. Liu, J. Wang and H. Zhang, *Small*, 2013, **9**, 140–147.

- 79 K. Zhang, Y. Zhang, T. Zhang, W. Dong, T. Wei, Y. Sun, X. Chen, G. Shen and N. Dai, *Nano Res.*, 2014, **8**, 743–750.
- 80 X. Dai, K. Du, Z. Li, H. Sun, Y. Yang, W. Zhang and X. Zhang, *Int. J. Hydrogen Energy*, 2015, **40**, 8877–8888.
- 81 M. Shen, Z. Yan, L. Yang, P. Du, J. Zhang and B. Xiang, *Chem. Commun.*, 2014, **50**, 15447–15449.
- 82 U. Bhanu, M. R. Islam, L. Tetard and S. I. Khondaker, *Sci. Rep.*, 2014, **4**, 5575.
- 83 K. K. Amara, Y. Chen, Y.-C. Lin, R. Kumar, E. Okunishi, K. Suenaga, S. Y. Quek and G. Eda, *Chem. Mater.*, 2016, **28**, 2308–2314.
- 84 B. Ouyang, Z. Mi and J. Song, *J. Phys. Chem. C*, 2016, **120**, 8927–8935.
- 85 S. Bai, L. Wang, X. Chen, J. Du and Y. Xiong, *Nano Res.*, 2014, **8**, 175–183.
- 86 W. Cao, V. Pankratov, M. Huttula, X. Shi, S. Saukko, Z. Huang and M. Zhang, *Mater. Chem. Phys.*, 2015, **158**, 89–95.
- 87 Y. Shi, J. K. Huang, L. Jin, Y. T. Hsu, S. F. Yu, L. J. Li and H. Y. Yang, *Sci. Rep.*, 2013, **3**, 1839.
- 88 S. Su, M. Zou, H. Zhao, C. Yuan, Y. Xu, C. Zhang, L. Wang, C. Fan and L. Wang, *Nanoscale*, 2015, **7**, 19129–19135.
- 89 H. Wu, Y. Wu, X. Chen, Y. Ma, M. Xu, W. Wei, J. Pan and X. Xiong, *RSC Adv.*, 2016, **6**, 23161–23168.
- 90 X. Wang, H. Feng, Y. Wu and L. Jiao, *J. Am. Chem. Soc.*, 2013, **135**, 5304–5307.
- 91 H. Li, X. Qi, J. Wu, Z. Zeng, J. Wei and H. Zhang, *ACS Nano*, 2013, **7**, 2842–2849.
- 92 S. M. Tan, Z. Sofer, J. Luxa and M. Pumera, *ACS Catal.*, 2016, **6**, 4594–4607.
- 93 A. Ambrosi, Z. Sofer and M. Pumera, *Small*, 2015, **11**, 605–612.
- 94 A. Ambrosi, Z. Sofer and M. Pumera, *Chem. Commun.*, 2015, **51**, 8450–8453.
- 95 J. Zheng, H. Zhang, S. Dong, Y. Liu, C. T. Nai, H. S. Shin, H. Y. Jeong, B. Liu and K. P. Loh, *Nat. Commun.*, 2014, **5**, 2995.
- 96 E. Varrla, C. Backes, K. R. Paton, A. Harvey, Z. Gholamvand, J. McCauley and J. N. Coleman, *Chem. Mater.*, 2015, **27**, 1129–1139.
- 97 Y. H. Lee, X. Q. Zhang, W. Zhang, M. T. Chang, C. T. Lin, K. D. Chang, Y. C. Yu, J. T. Wang, C. S. Chang, L. J. Li and T. W. Lin, *Adv. Mater.*, 2012, **24**, 2320–2325.
- 98 X. Li and H. Zhu, *Journal of Materomics*, 2015, **1**, 33–44.
- 99 R. Lv, J. A. Robinson, R. E. Schaak, D. Sun, Y. Sun, T. E. Mallouk and M. Terrones, *Acc. Chem. Res.*, 2015, **48**, 56–64.
- 100 J. T. Jang, S. Jeong, J. W. Seo, M. C. Kim, E. Sim, Y. Oh, S. Nam, B. Park and J. Cheon, *J. Am. Chem. Soc.*, 2011, **133**, 7636–7639.
- 101 S. Jeong, D. Yoo, J. T. Jang, M. Kim and J. Cheon, *J. Am. Chem. Soc.*, 2012, **134**, 18233–18236.
- 102 Y. Dai, X. H. Yan, X. Wu, D. W. Sha, M. Chen, H. Zou, J. Ren and X. N. Cheng, *Mater. Lett.*, 2016, **173**, 203–206.
- 103 J. Lu, J. H. Lu, H. Liu, B. Liu, L. Gong, E. S. Tok, K. P. Loh and C. H. Sow, *Small*, 2015, **11**, 1792–1800.
- 104 D. B. Nimbalkar, H.-H. Lo, P. V. R. K. Ramacharyulu and S.-C. Ke, *RSC Adv.*, 2016, **6**, 31661–31667.
- 105 Y. J. Yuan, F. Wang, B. Hu, H. W. Lu, Z. T. Yu and Z. G. Zou, *Dalton Trans.*, 2015, **44**, 10997–11003.
- 106 H. Yan, P. Song, S. Zhang, Z. Yang and Q. Wang, *RSC Adv.*, 2015, **5**, 79593–79599.
- 107 J. Guo, H. Zhu, Y. Sun, L. Tang and X. Zhang, *J. Mater. Chem. A*, 2016, **4**, 4783–4789.
- 108 H. Sun, J. Chao, X. Zuo, S. Su, X. Liu, L. Yuwen, C. Fan and L. Wang, *RSC Adv.*, 2014, **4**, 27625.
- 109 C. Kuru, C. Choi, A. Kargar, D. Choi, Y. J. Kim, C. H. Liu, S. Yavuz and S. Jin, *Adv. Sci.*, 2015, **2**, 1500004.
- 110 M. Zirak, M. Zhao, O. Moradlou, M. Samadi, N. Sarikhani, Q. Wang, H. L. Zhang and A. Z. Moshfegh, *Sol. Energy Mater. Sol. Cells*, 2015, **141**, 260–269.
- 111 Y. Shi, J. Wang, C. Wang, T. T. Zhai, W. J. Bao, J. J. Xu, X. H. Xia and H. Y. Chen, *J. Am. Chem. Soc.*, 2015, **137**, 7365–7370.
- 112 J. Huang, D. Hou, Y. Zhou, W. Zhou, G. Li, Z. Tang, L. Li and S. Chen, *J. Mater. Chem. A*, 2015, **3**, 22886–22891.
- 113 Y. Huang, Y. E. Miao, L. Zhang, W. W. Tjiu, J. Pan and T. Liu, *Nanoscale*, 2014, **6**, 10673–10679.
- 114 S. Xu, D. Li and P. Wu, *Adv. Funct. Mater.*, 2015, **25**, 1127–1136.
- 115 G. M. Whitesides and B. Grzybowski, *Science*, 2002, **295**, 2418.
- 116 M. Grzelczak, J. Vermant, E. M. Furst and L. M. Liz-Marzan, *ACS Nano*, 2010, **4**, 3591–3605.
- 117 Y.-F. Zhao, Z.-Y. Yang, Y.-X. Zhang, L. Jing, X. Guo, Z. Ke, P. Hu, G. Wang, Y.-M. Yan and K.-N. Sun, *J. Phys. Chem. C*, 2014, **118**, 14238–14245.
- 118 J. Kim, S. Byun, A. J. Smith, J. Yu and J. Huang, *J. Phys. Chem. Lett.*, 2013, **4**, 1227–1232.
- 119 M. Yoshimura and K. Byrappa, *J. Mater. Sci.*, 2007, **43**, 2085–2103.
- 120 P. Zhang, X. Lu, Y. Huang, J. Deng, L. Zhang, F. Ding, Z. Su, G. Wei and O. G. Schmidt, *J. Mater. Chem. A*, 2015, **3**, 14562–14566.
- 121 S. Su, H. Sun, F. Xu, L. Yuwen and L. Wang, *Electroanalysis*, 2013, **25**, 2523–2529.
- 122 Y. Guo, X. Zhang, X. Zhang and T. You, *J. Mater. Chem. A*, 2015, **3**, 15927–15934.
- 123 V. O. Koroteev, L. G. Bulusheva, I. P. Asanov, E. V. Shlyakhova, D. V. Vyalikh and A. V. Okotrub, *J. Phys. Chem. C*, 2011, **115**, 21199–21204.
- 124 K. Zhang, Y. Zhao, S. Zhang, H. Yu, Y. Chen, P. Gao and C. Zhu, *J. Mater. Chem. A*, 2014, **2**, 18715–18719.
- 125 X. Li, W. Li, M. Li, P. Cui, D. Chen, T. Gengenbach, L. Chu, H. Liu and G. Song, *J. Mater. Chem. A*, 2015, **3**, 2762–2769.
- 126 S. Ding, J. S. Chen and X. W. Lou, *Chemistry*, 2011, **17**, 13142–13145.
- 127 Y. Huang, Y.-E. Miao, J. Fu, S. Mo, C. Wei and T. Liu, *J. Mater. Chem. A*, 2015, **3**, 16263–16271.
- 128 J. D. Holladay, J. Hu, D. L. King and Y. Wang, *Catal. Today*, 2009, **139**, 244–260.
- 129 B. Johnston, M. C. Mayo and A. Khare, *Technovation*, 2005, **25**, 569–585.
- 130 A. B. Laursen, S. Kegnæs, S. Dahl and I. Chorkendorff, *Energy Environ. Sci.*, 2012, **5**, 5577.

- 131 Y. Qu, H. Medina, S.-W. Wang, Y.-C. Wang, C.-W. Chen, T.-Y. Su, A. Manikandan, K. Wang, Y.-C. Shih, J.-W. Chang, H.-C. Kuo, C.-Y. Lee, S.-Y. Lu, G. Shen, Z. M. Wang and Y.-L. Chueh, *Adv. Mater.*, 2016, **28**, 9831–9838.
- 132 E. J. Popczun, J. R. McKone, C. G. Read, A. J. Biacchi, A. M. Wiltrot, N. S. Lewis and R. E. Schaak, *J. Am. Chem. Soc.*, 2013, **135**, 9267–9270.
- 133 W. Sheng, H. A. Gasteiger and Y. Shaohorn, *J. Electrochem. Soc.*, 2010, 157.
- 134 Y. Zheng, Y. Jiao, M. Jaroniec and S. Z. Qiao, *Angew. Chem.*, 2015, **54**, 52–65.
- 135 T. F. Jaramillo, K. P. Jorgensen, J. Bonde, J. H. Nielsen, S. Horch and I. Chorkendorff, *Science*, 2007, **317**, 100–102.
- 136 J. Xie, H. Zhang, S. Li, R. Wang, X. Sun, M. Zhou, J. Zhou, X. W. Lou and Y. Xie, *Adv. Mater.*, 2013, **25**, 5807–5813.
- 137 D. McAteer, Z. Gholamvand, N. McEvoy, A. Harvey, E. O’Malley, G. S. Duesberg and J. N. Coleman, *ACS Nano*, 2016, **10**, 672–683.
- 138 X. Huang, Z. Zeng, S. Bao, M. Wang, X. Qi, Z. Fan and H. Zhang, *Molecular therapy : the journal of the American Society of Gene Therapy*, 2013, **4**, 1444.
- 139 D. Hou, W. Zhou, X. Liu, K. Zhou, J. Xie, G. Li and S. Chen, *Electrochim. Acta*, 2015, **166**, 26–31.
- 140 A. Fujishima and K. Honda, *Nature*, 1972, **238**, 37–38.
- 141 M. Ni, M. K. H. Leung, D. Y. C. Leung and K. Sumathy, *Renewable Sustainable Energy Rev.*, 2007, **11**, 401–425.
- 142 Q. Liu, Z. Pu, A. M. Asiri, A. H. Qusti, A. O. Al-Youbi and X. Sun, *J. Nanopart. Res.*, 2013, 15.
- 143 H. Li, Y. Wang, G. Chen, Y. Sang, H. Jiang, J. He, X. Li and H. Liu, *Nanoscale*, 2016, **8**, 6101–6109.
- 144 W. Zhou, D. Hou, Y. Sang, S. Yao, J. Zhou, G. Li, L. Li, H. Liu and S. Chen, *J. Mater. Chem. A*, 2014, **2**, 11358.
- 145 B. Dunn, H. Kamath and J.-M. Tarascon, *Science*, 2011, **334**, 928.
- 146 Y. Jing, Z. Zhou, C. R. Cabrera and Z. Chen, *J. Mater. Chem. A*, 2014, **2**, 12104.
- 147 W. Wen, J.-M. Wu and M.-H. Cao, *Nano Energy*, 2013, **2**, 1383–1390.
- 148 W. Wen, J. M. Wu and M. H. Cao, *Nanoscale*, 2014, **6**, 12476–12481.
- 149 H. Hwang, H. Kim and J. Cho, *Nano Lett.*, 2011, **11**, 4826–4830.
- 150 M. Mao, L. Mei, D. Guo, L. Wu, D. Zhang, Q. Li and T. Wang, *Nanoscale*, 2014, **6**, 12350–12353.
- 151 F. Zhou, S. Xin, H. W. Liang, L. T. Song and S. H. Yu, *Angew. Chem.*, 2014, **53**, 11552–11556.
- 152 Y. Ren, Z. Liu, F. Pourpoint, A. R. Armstrong, C. P. Grey and P. G. Bruce, *Angew. Chem.*, 2012, **51**, 2164–2167.
- 153 Y. Chen, B. Song, X. Tang, L. Lu and J. Xue, *Small*, 2014, **10**, 1536–1543.
- 154 K. Chang and W. Chen, *ACS Nano*, 2011, **5**, 4720–4728.
- 155 J.-Y. Liao, B. D. Luna and A. Manthiram, *J. Mater. Chem. A*, 2016, **4**, 801–806.
- 156 B. Chen, E. Liu, F. He, C. Shi, C. He, J. Li and N. Zhao, *Nano Energy*, 2016, **26**, 541–549.
- 157 L. Yang, S. Wang, J. Mao, J. Deng, Q. Gao, Y. Tang and O. G. Schmidt, *Adv. Mater.*, 2013, **25**, 1180–1184.
- 158 D.-A. Zhang, Q. Wang, Q. Wang, J. Sun, L.-L. Xing and X.-Y. Xue, *Electrochim. Acta*, 2015, **173**, 476–482.
- 159 M. Asadi, B. Kumar, C. Liu, P. Phillips, P. Yasaee, A. Behrangiinia, P. Zapol, R. F. Klie, L. A. Curtiss and A. Salehi-Khojin, *ACS Nano*, 2016, **10**, 2167–2175.
- 160 X. Xu, W. Liu, Y. Kim and J. Cho, *Nano Today*, 2014, **9**, 604–630.
- 161 J. Lee, P. Dak, Y. Lee, H. Park, W. Choi, M. A. Alam and S. Kim, *Sci. Rep.*, 2014, **4**, 7352.
- 162 B. Cho, M. G. Hahm, M. Choi, J. Yoon, A. R. Kim, Y. J. Lee, S. G. Park, J. D. Kwon, C. S. Kim, M. Song, Y. Jeong, K. S. Nam, S. Lee, T. J. Yoo, C. G. Kang, B. H. Lee, H. C. Ko, P. M. Ajayan and D. H. Kim, *Sci. Rep.*, 2015, **5**, 8052.
- 163 M. Park, Y. J. Park, X. Chen, Y. K. Park, M. S. Kim and J. H. Ahn, *Adv. Mater.*, 2016, **28**, 2556–2562.
- 164 J. Ye, X. Li, J. Zhao, X. Mei and Q. Li, *Nanoscale Res. Lett.*, 2015, **10**, 454.
- 165 Q. Yue, Z. Shao, S. Chang and J. Li, *Nanoscale Res. Lett.*, 2013, **8**, 425.
- 166 Z. Li, S. Jiang, S. Xu, C. Zhang, H. Qiu, C. Li, Y. Sheng, Y. Huo, C. Yang and B. Man, *Sens. Actuators, B*, 2016, **230**, 645–652.
- 167 J. Zhao, Z. Zhang, S. Yang, H. Zheng and Y. Li, *J. Alloys Compd.*, 2013, **559**, 87–91.
- 168 H. Yan, P. Song, S. Zhang, Z. Yang and Q. Wang, *J. Alloys Compd.*, 2016, **662**, 118–125.
- 169 Q. He, Z. Zeng, Z. Yin, H. Li, S. Wu, X. Huang and H. Zhang, *Small*, 2012, **8**, 2994–2999.
- 170 N. Singh, R. K. Gupta and P. S. Lee, *ACS Appl. Mater. Interfaces*, 2011, **3**, 2246–2252.
- 171 T. Wei, P. Yeh, S. Lu and Z. L. Wang, *J. Am. Chem. Soc.*, 2009, **131**, 17690–17695.
- 172 S. Wu, Z. Zeng, Q. He, Z. Wang, S. J. Wang, Y. Du, Z. Yin, X. Sun, W. Chen and H. Zhang, *Small*, 2012, **8**, 2264–2270.
- 173 L. Wang, Y. Wang, J. I. Wong, T. Palacios, J. Kong and H. Y. Yang, *Small*, 2014, **10**, 1101–1105.
- 174 A. B. Farimani, K. Min and N. R. Aluru, *ACS Nano*, 2014, **8**, 7914–7922.
- 175 K. Jin, L. Xie, Y. Tian and D. Liu, *J. Phys. Chem. C*, 2016, **120**, 11204–11209.
- 176 S. Su, H. Sun, F. Xu, L. Yuwen, C. Fan and L. Wang, *Microchim. Acta*, 2014, **181**, 1497–1503.
- 177 O. Lopez-Sanchez, D. Lembke, M. Kayci, A. Radenovic and A. Kis, *Nat. Nanotechnol.*, 2013, **8**, 497–501.
- 178 D. Kufer, I. Nikitskiy, T. Lasanta, G. Navickaite, F. H. Koppens and G. Konstantatos, *Adv. Mater.*, 2015, **27**, 176–180.
- 179 J. Miao, W. Hu, Y. Jing, W. Luo, L. Liao, A. Pan, S. Wu, J. Cheng, X. Chen and W. Lu, *Small*, 2015, **11**, 2392–2398.
- 180 W. Zhang, P. Zhang, Z. Su and G. Wei, *Nanoscale*, 2015, **7**, 18364–18378.
- 181 S. Mouri, Y. Miyauchi and K. Matsuda, *Nano Lett.*, 2013, **13**, 5944–5948.
- 182 C. Tan and H. Zhang, *Chem. Soc. Rev.*, 2015, **44**, 2713–2731.

