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Pyridinic and graphitic nitrogen-rich graphene for high-performance supercapacitors and metal-free bifunctional electrocatalysts for ORR and OER†

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A facile synthesis method to produce nitrogen-doped graphene containing a high atomic percentage of pyridinic N and graphitic N *via* the thermal annealing of graphene oxide and uric acid is reported. The method yielded N-doped graphene nanosheets that had high surface areas and that contained 9.22 at% nitrogen, which was present in four bonding configurations with a high ratio of pyridinic N and graphitic N. The synthesized N-doped graphene showed excellent capacitance properties, which led to the fabrication of a stacked electrode supercapacitor cell of 2-electrode configuration with a specific capacitance of 230 F g⁻¹ at a current density of 1 A g⁻¹ and with a remarkably high energy density of 62.6 W h kg⁻¹ in aqueous electrolyte. Additionally, this material showed superior properties as an electrocatalyst for both the oxygen-reduction reaction (ORR) and oxygen evolution reaction (OER), producing high current density primarily *via* the four-electron pathway for ORR at a current density of -4.3 mA cm⁻² and OER activity of a generating current density of 10 mA cm⁻² at 1.74 V vs. RHE in alkaline media. Hence, this dual-natured N-doped graphene could provide a platform for developing nanostructured graphene-based advanced energy storage and conversion devices.

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1. Introduction

Carbon-based materials for clean and renewable energy are becoming promising candidates to help meet the global demand for sustainable energy development. In particular, nanostructured materials for efficient energy storage and conversion devices have been reported for a wide range of applications including as solar cells,¹ fuel cells, lithium-ion batteries and electrochemical double-layer capacitors.²⁻⁴ Among these energy storage devices, supercapacitors have gained attention due to their orders-of-magnitude-higher power density, cycle efficiency and charge/discharge rates than batteries.^{5,6} Besides energy storage, energy conversion is also crucial to meeting global energy demands. Fuel cells are among the promising methods of delivering energy in the future, but their widespread implementation is hindered by the expense of the cathode material (platinum), as well as by their deactivation by the by-products of fuel decomposition (the fuel crossover effect), making the development of alternative, precious-metal-free electrode materials essential.⁷⁻⁹ Graphene, composed of a one-atom-thick honeycomb lattice of sp²-hybridized carbons,

has shown great potential as an electrode material in energy storage and conversion due to its excellent electrical conductivity and high specific surface area.¹⁰⁻¹³

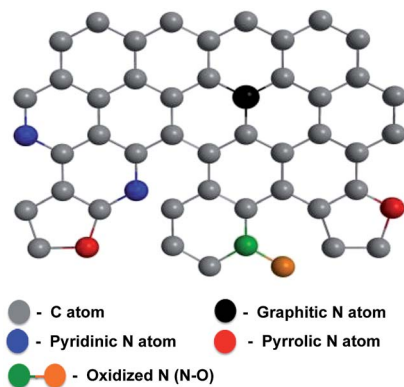
Chemical dopants can alter the electronic properties of graphene.¹⁴ The introduction of dopants, either by the adsorption of molecules on the graphene surface or by the substitution of dopant atoms into the carbon lattice, can produce a band gap in graphene, and this strategy has led to the development of high-performance energy storage and conversion devices.^{15,16} Various heteroatoms, including N, B, and S, have imparted excellent electrochemical properties,¹⁷⁻²¹ and N-doped graphene in particular has shown capacitance suitable for high-performance supercapacitor applications,²²⁻²⁶ excellent electrocatalytic activity for oxygen reduction in fuel cells^{27,28} and oxygen evolution reaction for rechargeable metal-air batteries.^{29,30} The properties of nitrogen-doped graphene (NG) are conferred by nitrogen atoms within the carbon lattice, where they can exist as pyridinic, pyrrolic, graphitic, and oxidized N (Scheme 1).³¹

The roles of the differently configured nitrogen atoms in enhancing capacitance or generating catalytically active sites are still controversial,^{32,33} but both quantum mechanical calculations³⁴⁻³⁶ and experimental^{24,37} investigations indicate that graphitic, pyridinic, and pyrrolic N can enhance the capacitance, whereas pyridinic N and graphitic N can be essential to catalyzing oxygen reduction reaction by enhancing the adsorption of O₂ on the adjacent carbon atoms and thus promotes the four-electron pathway.³⁸⁻⁴¹ Besides the capacitance and oxygen reduction

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Scheme 1 Schematic structure of NG.

reaction, oxygen evolution reaction has gained a considerable attention due to the promising future of rechargeable metal–air batteries.^{42–44} N-doped graphene,²⁹ NG/CNT composites,³⁰ N-doped coaxial carbon nanotube,⁴⁵ carbon nitrogen nanotubes⁴⁶ and nitrogen and phosphorous co-doped nanocarbons^{47,48} have shown to be promising alternatives to noble and transition metal based OER catalysts.

The enhanced supercapacitive and electrocatalytic properties are highly dependent on the type and amount of nitrogen present. To obtain both the capacitance and electrocatalytic properties, high atomic concentrations of pyridinic, pyrrolic and graphitic nitrogen are required. The types and amounts of the different nitrogen bonding configurations in an N-doped graphene sample depend on the nitrogen source as well as the synthesis method. Nitrogen has been incorporated into the graphitic lattice from sources like acetonitrile,⁴⁹ pyridine⁵⁰ and ammonia⁵¹ *via* chemical vapor deposition (CVD) or high-temperature treatment, and the resulting materials have been applied in lithium batteries,⁴⁹ field effect transistors,⁵⁰ and as catalysts for oxygen reduction.⁵¹ However, it is difficult to obtain materials with high levels of atomic nitrogen doping, and on a large scale, *via* CVD using gaseous or liquid sources of nitrogen, and the method also suffers from practical limitations due to possible contamination from the metal deposition catalyst as well as the toxicity of the precursors. On the other hand, the co-pyrolysis of graphene or a graphene precursor with a solid source of nitrogen, such as melamine⁵¹ or cyanamide,⁵² has produced NG with high levels of nitrogen (~9 at%). Co-pyrolysis methods are attractive because they can be scaled up without any catalyst.⁵³ However, the NG synthesized from melamine had a low surface area,⁵¹ whereas cyanamide, on the other hand, required a surfactant to bind to the graphene surface.⁵² Moreover, the NG synthesized from cyanamide did not contain pyrrolic nitrogen,⁵² which plays an important role in enhancing capacitance.³⁴ In addition to melamine and cyanamide, urea has been employed as a solid source of nitrogen in the synthesis of N-doped graphene, however both the capacitance and electrocatalytic properties in the material were not investigated.^{53–56} In order to show both high capacitance and electrocatalytic activity, an NG sample must contain sufficient percentages of pyridinic, pyrrolic, and graphitic nitrogen atoms, and have large surface area and high pore volume.

Herein, we report the synthesis of NG *via* the thermal annealing of graphene oxide (GO) with uric acid at 800 °C under an argon atmosphere. Uric acid has a high nitrogen content (~33 wt%), and its crystals can adhere to surfaces bearing various organic functionalities, and *via* multiple types of forces (electrostatic interactions, hydrogen bonding and/or van der Waals interactions), without the aid of surfactants.⁵⁷ These properties allowed the co-pyrolysis of graphene oxide and uric acid to produce a nitrogen-doped reduced graphene oxide (rGO) with nitrogen atoms in all of the possible configurations in the graphitic lattice, a high total nitrogen content with a high ratio of pyridinic N and graphitic N and a high surface area after pyrolysis. As a result, the N-doped graphene had both high capacitance and bifunctional electrocatalytic activity towards oxygen reduction and evolution.

2. Experimental

Synthesis of nitrogen-doped graphene

GO was prepared from naturally expanded graphite (Ashbury Carbons, Grade-3772) using a reported procedure for the generation of ultra-large GO sheets.⁵⁸ The GO sheets (approx. 20 mg) were dispersed in water by ultrasonication for 30 min at a frequency of 50 Hz, then mixed with uric acid in a ratio of 1 : 1, 1 : 5 or 1 : 10 by mass. The mixtures were then stirred continuously and heated at 80 °C for 10–12 h to remove the water. The resulting solids were transferred to a tubular furnace and heated in an argon atmosphere at 5 °C min⁻¹ to 800 °C, then annealed for 1 h. The synthesized nitrogen-doped graphenes produced with 1 : 1, 1 : 5 and 1 : 10 ratios of graphene oxide and uric acid were named NG1, NG5 and NG10, respectively. Pure GO was also thermally annealed to yield reduced graphene oxide (rGO) under the same conditions, and used as reference.

Characterization

The microstructures and morphology of the synthesized samples were investigated by Field Emission Scanning Electron Microscopy (FESEM, Zeiss ULTRA plus) and High Resolution Transmission Electron Microscopy (HRTEM, JEOL 2200FS) with an acceleration voltage of 200 kV. Thermogravimetric Analysis (TGA, TA Q500) was performed under Ar and at a heating rate of 5 °C min⁻¹. X-ray diffractometry (XRD) was performed on a Shimadzu S6000 equipped with Cu K α radiation ($\lambda = 0.154$ nm) with a resolution of 0.02°. Raman spectra were collected with Renishaw inVia Raman Spectrometer with laser wavelength of 514 nm. Nitrogen sorption isotherms and Brunauer–Emmett–Teller (BET) surface areas were measured over $P/P_0 = 0.05$ –0.3 at 77 K on an Autosorb IQ (Quantachrome, USA). X-ray photoelectron spectroscopy (XPS) analyses were performed by ESCALAB250Xi (Thermo Scientific, UK) with a monochromated Al K α (energy 1486.68 eV) operating at 150 W (13 kV \times 12 mA) under a vacuum of 2×10^{-9} mbar. The analysis spot was 500 μ m in diameter. The binding energies were referenced to the adventitious hydrocarbon C 1s signal at 284.9 eV.

Electrochemical measurement

For electrochemical characterization and measurement, the nitrogen-doped graphenes were deposited from a 1 mg mL⁻¹



mixture, prepared by sonicating in degassed isopropanol, on a glassy-carbon electrode. The glassy-carbon working electrode was polished mechanically with a slurry of 0.05 μm alumina particles to obtain a mirror-like surface, washed with Milli-Q water and acetone, and allowed to dry. It was then tested in a three-electrode cell, using platinum wire as the counter electrode and a Ag/AgCl, KCl (3 M) electrode as the reference electrode. Measurements were taken using a Bio-logic SP300 potentiostat. The binding agents were 10 wt% PTFE for capacitance measurements and 5 wt% Nafion for electrocatalytic measurements. Capacitance was determined using a T-cell device in a two-electrode cell configuration, as described previously.^{59,60} The device was fabricated using a Swagelok system with two stainless steel current collector. The PVDF separator was placed between the current collectors where the material to be studied was sandwiched by pressing together at high pressure. 0.5 M H_2SO_4 was used as the electrolyte and the cell was airtight. 2.0 mg active materials were deposited on the diameter of an 8 mm stainless steel electrode surface using a 10-wt% PTFE binder. The gravimetric capacitance from galvanostatic charge/discharge was calculated using eqn (1).⁶

$$C = \frac{4I}{m \frac{dV}{dt}} \quad (1)$$

where C is specific capacitance (F g^{-1}), I is the charge/discharge current density (A g^{-1}), m is the total mass of materials deposited on the electrodes and dV/dt was calculated from the slope of the discharge curve. The energy density (E) and power density (P) were calculated using eqn (2) and (3), respectively.^{59,60}

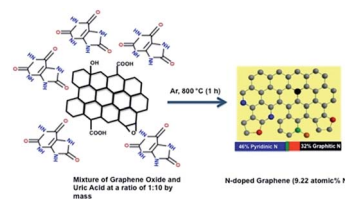
$$E = \frac{1}{2} \times C(\Delta V)^2 \quad (2)$$

$$P = Elt \quad (3)$$

where C is the specific capacitance of the device, ΔV is the potential range and t is the time to discharge. The electrocatalytic properties were evaluated by rotating disk electrode (RDE) experiments in an O_2 -saturated 0.1 M KOH solution. First, the oxygen-reduction capability was assessed by comparing cyclic voltammetric curves measured in argon- and oxygen-saturated electrolytes. Linear sweep voltammetry curves were recorded over the potential range of 0 to 1 V *versus* reversible hydrogen electrode (RHE) at a scan rate of 10 mV s^{-1} at ambient temperature and at rotation speeds of 100–1600 rpm. The effects of different mass loading on the glassy carbon electrode in electrocatalysis were evaluated by depositing 4, 8, 12, 16 and 20 μg of NG10. OER activity was measured by linear sweep voltammetry in the potential window ranging from 1.0 to 2.0 V *versus* reversible hydrogen electrode (RHE) scale by adding a value of 0.964 V. The commercial Pt/C (20 wt%) and Ru/C (20 wt%) into Nafion solution with same loading amount were taken for comparison.

3. Results and discussion

In the synthesis of NG (Scheme 2), graphene oxide and uric acid were mixed together in water (in ratios of 1 : 1, 1 : 5 or 1 : 10 by



Scheme 2 Schematic diagram of the formation of NG *via* the thermal annealing of graphene oxide with uric acid.

mass in mg scale for NG1, NG5 and NG10, respectively), and then stirred continuously at 80 °C to remove the water. We expect that the uric acid molecules adsorbed on the graphene oxide surface during this time. The solid mixtures were then heated in a tubular furnace at a rate of 5 °C min^{-1} to 800 °C in an argon environment. The formation of N-doped graphenes from the pyrolysis of GO with melamine⁵¹ and cyanamide⁵³ had been proposed to proceed in two steps *via* formation of carbon nitride first, which forms from the pyrolysis of cyanamide.⁶² For the present case, from the thermogravimetric-analysis curves examined under argon in a thermogravimetric analyser (TGA) (Fig. S1†), it is observed that the rapid weight losses started from 400 °C due to the decomposition of uric acid and the complete decomposition of the GO/uric acid mixture occurred around 750 °C. The thermogravimetric analysis of uric acid has also been performed to measure the possible amount of hard carbon transformed from the uric acid after carbonization. The TGA curve of uric acid showed the almost complete decomposition from 650 °C and only 0.38% of the initial mass of uric acid remained after the complete carbonization. This behaviour supports the mechanism of interaction of heterocyclic –NH groups of uric acid with the functional groups of GO, contributing to the incorporation of N into the graphene lattice rather than *via* formation of carbon nitride.⁵³ Due to organic in nature and high surface attachment ability,⁵⁷ uric acid can be attached on the graphene surface better than the inorganic solid precursor urea and thus higher atomic percentage of nitrogen with four possible bonding configurations has been achieved.

The similar mechanism has been observed with organo-nitrogen compounds having high nitrogen contents^{26,71} as well as the high nitrogen containing liquid organic isomers of phenylenediamine.⁷² The advantages of uric acid are, the less expensive commercial availability compared to organonitrogen compounds like aminoterephthalic acid²⁶ and methylimidazole⁷¹ as well as safe handling during synthesis due to its solid powder form rather than handling toxic liquids. The morphology, structure, and graphitic layers of the obtained NGs were characterized by scanning electron microscopy (SEM) and transmission electron microscopy (TEM). The SEM images of NG10 showed a laminar morphology (Fig. 1a), and its low-resolution TEM image showed transparent stacked graphene sheets. The high-resolution transmission electron micrograph (HRTEM) and selected area electron diffraction (SAED) pattern revealed the highly crystalline nature of NG10 (Fig. 1b). The well-defined diffraction spots and rings in the SAED pattern confirmed that the sample was comprised of the hexagonal carbon lattice typical of graphene-



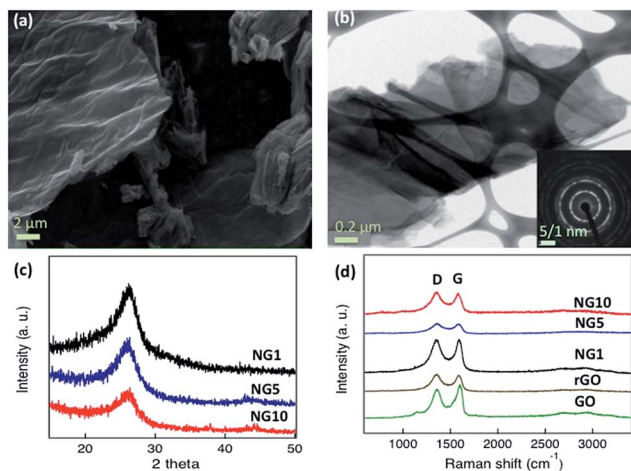


Fig. 1 (a) SEM image of NG10, (b) TEM image and selected area electron diffraction (SAED) pattern (inset) of NG10, (c) XRD pattern of NG1 (black), NG5 (blue) and NG10 (red), (d) Raman spectra of NG10 (red), NG5 (blue), NG1 (black), reduced graphene oxide (brown) and graphene oxide (green).

like materials.⁶³ The XRD patterns of NG1, NG5 and NG10 (Fig. 1c) showed broad peaks at around 26.0°, which corresponded to an interlayer spacing of 0.342 nm.

GO, rGO, and the N-doped materials NG1, NG5 and NG10, were characterized by Raman spectroscopy at an excitation wavelength of 514 nm under ambient conditions. As shown in Fig. 1d and Table 1, the G peaks of the GO, rGO, NG1, NG5 and NG10 occurred between 1582 and 1598 cm⁻¹. The G peaks shifted to lower energy from GO to NG10, indicating that a conjugated sp² structure was restored during pyrolysis.⁵³ The I_D/I_G ratios of GO, rGO, NG1, NG5 and NG10 (Table 1) were used to evaluate the defects or in-plane crystallite size after pyrolysis, and the crystallite sizes, L_a , were calculated according to the Tuinstra-Koenig (TK) relation, eqn (4).^{38,64}

$$L_a = (2.4 \times 10^{-10})\lambda^4(I_D/I_G)^{-1} \quad (4)$$

where λ is the Raman excitation wavelength (514 nm). The crystallite sizes of rGO, NG1, NG5 and NG10 were smaller than that of GO (Table 1), indicating that the crystallite sizes decreased upon thermal annealing.

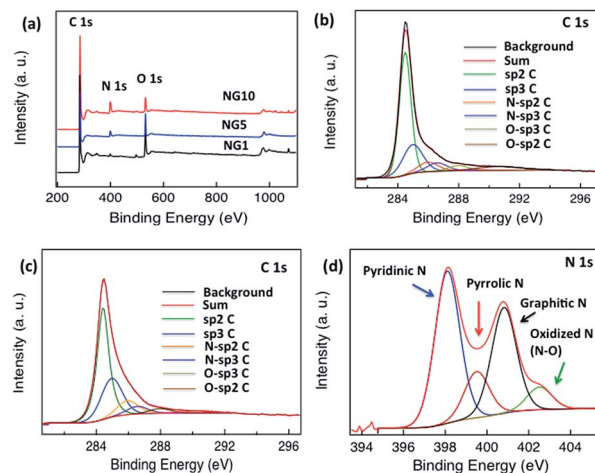


Fig. 2 (a) XPS spectra of NG10 (red), NG5 (blue) and NG1 (black), (b) high-resolution C 1s spectrum of NG10, (c) high-resolution C 1s spectrum of NG1, (d) N 1s spectrum of NG10.

To characterize the elemental composition and nitrogen content in the NGs, X-ray photoelectron spectra (XPS) of the materials were measured. A comparison of the spectra of GO and rGO showed a decrease in oxygen content in rGO; neither spectrum showed obvious N 1s peaks (Fig. S3†). As shown in Fig. 2a, the survey spectra of NG samples revealed the presence of C, O and N. In NG1, NG5 and NG10, the high-resolution C 1s peak was centred at 284.4 eV and tailed at higher binding energies, indicating the connection of carbon atoms with N and O atoms. Though NG10 (Fig. 2b), NG5 (Fig. S4†) and NG1 (Fig. 2c) showed significant amounts of graphite-like C (284.40 eV),⁶⁵ quantitative analysis of the C 1s region of the spectrum was impossible due to the likely presence of adventitious carbon in the samples. The deconvolutions of the N 1s peaks (Fig. 2d, S5 and S6,† Table 1) showed the presence of four nitrogen configurations. All the NG1, NG5 and NG10 contained primarily (>40%) pyridinic nitrogen and minor amounts (<10%) of N-oxides, but they differed in that NG1 contained much more pyrrolic N, whereas NG5 and NG10 was richer in graphitic N. The possible reason of high content of pyridinic N and graphitic N in NG5 and NG10 is due to interaction of higher amount of available nitrogen atoms with the active sites of created vacant

Table 1 Characterization data for GO, rGO, NG1, and NG10

| | $\nu(G)^a$ (cm ⁻¹) | I_D/I_G^a (-) | L_a^b (nm) | S_{BET}^c (m ² g ⁻¹) | V_{total}^d (cm ³ g ⁻¹) | N sites ^e (% of total N) | | | | |
|------|--------------------------------|-----------------|--------------|---|--|-------------------------------------|----------|-----------|-------|----------------------------|
| | | | | | | Pyridinic | Pyrrolic | Graphitic | Oxide | Total N ^e (at%) |
| GO | 1598 | 0.87 | 19 | — | — | — | — | — | — | — |
| rGO | 1594 | 1.03 | 16.3 | 75.68 | 0.07 | — | — | — | — | — |
| NG1 | 1592 | 0.99 | 16.8 | 96.06 | 0.20 | 41.4 | 34.9 | 19.1 | 4.61 | 1.52 |
| NG5 | 1582 | 1.03 | 16.2 | 164.04 | 0.68 | 46.3 | 13.6 | 28.4 | 9.33 | 6.75 |
| NG10 | 1582 | 1.03 | 16.2 | 197.80 | 0.80 | 46.6 | 14.1 | 32.4 | 6.83 | 9.22 |

^a Position of the G peak ($\nu(G)$) and ratio of D and G peak intensities (I_D/I_G) determined from Raman spectroscopy using $\lambda_{excitation} = 514$ nm. ^b Crystallite size (L_a) calculated from Raman data according to the Tuinstra-Koenig relation.⁶⁴ ^c BET surface area calculated over $P/P_0 = 0.05-0.30$ from N₂ adsorption at 77 K. ^d Total pore volume calculated from $P/P_0 = 0.995$ from N₂ adsorption at 77 K. ^e Calculated from X-ray photoelectron spectroscopy. Total nitrogen was calculated from detection-sensitivity-adjusted peak areas. ^f Not determined.



carbon bonds from the functional groups during the doping process at high temperature. Due to the presence of higher ratio of functional groups on the basal planes and six members benzene rings, the high amount of nitrogen doping occurred in pyridinic and graphitic sites.^{31,34,73}

Moreover, NG10 and NG5 had a much greater total nitrogen content (9.22 at% for NG10 and 6.75 at% for NG5, Table 1) and higher Brunauer–Emmett–Teller (BET) surface area ($197.8 \text{ m}^2 \text{ g}^{-1}$ measured over $P/P_0 = 0.05\text{--}0.30$ for NG10 and $164.04 \text{ m}^2 \text{ g}^{-1}$ measured over $P/P_0 = 0.05\text{--}0.30$ for NG5, Fig. S7, S8† and Table 1) than did NG1 (1.52 at% N, Table 1; $S_{\text{BET}} = 96.06 \text{ m}^2 \text{ g}^{-1}$ measured over $P/P_0 = 0.05\text{--}0.30$, Fig. S9† and Table 1); the NG materials had greater surface area than rGO ($75.68 \text{ m}^2 \text{ g}^{-1}$ measured over $P/P_0 = 0.05\text{--}0.30$, Fig. S10† and Table 1). NG10 also displayed high pore volume of $0.8 \text{ cm}^3 \text{ g}^{-1}$ (as defined at $P/P_0 = 0.995$), whereas the pore volumes of NG5, NG1 and rGO were 0.68, 0.2 and $0.07 \text{ cm}^3 \text{ g}^{-1}$ (as defined at $P/P_0 = 0.995$), respectively. Based on the above results, it is evident that the inclusion of N from uric acid contributed to raise the surface areas and pore volumes of the pyrolyzed N-doped graphene materials.

Cyclic voltammetry (CV) was used to investigate the electrochemical behavior of rGO, NG1, NG5 and NG10 in 0.5 M aqueous H_2SO_4 electrolyte deposited on a glassy carbon electrode at potential limits of 1.2 to $-0.2 \text{ V vs. Ag/AgCl}$ electrode in a three-electrode system at a scan rate of 100 mV s^{-1} (Fig. 3a), with 10 wt% PTFE in isopropanol being used as a binder. A comparison of the CVs for rGO, NG1, NG5 and NG10 revealed that the capacitive response became larger with a highly rectangular nature with increasing nitrogen content, and the higher doping percentage resulted in an enhancement of the double layer capacitance, indicating the more effective contribution of electrochemical double layer capacitance with pseudocapacitance for the sample with higher nitrogen content, NG10. Thus the electrochemical activity and reversibility of the capacitance of NG10 were recorded at different scan rates (Fig. S11†), and the capacitance increased at greater scan rates.

A T-cell device consisting of stainless steel electrodes with a Teflon cover was used to investigate the capacitive properties of NG10 for practical application (Fig. S12†).⁶⁰ The active material was deposited on the electrodes with a PTFE binder and dried in an oven at $110 \text{ }^\circ\text{C}$ for 1 h. A PVDF membrane separator was inserted between the two electrodes, which were then pressed and sandwiched under high pressure. The H_2SO_4 electrolyte was poured from the top to completely fill the device before it was sealed air-tight. CVs recorded in the T-cell at different scan rates (50, 100 and 200 mV s^{-1}) showed near-ideal rectangular voltammetric responses (Fig. 3b). The specific capacitance (C_{sp}) of the NG10 electrode material was calculated from the galvanostatic discharge curves (Fig. 3c) to be 230 F g^{-1} in 0.5 M H_2SO_4 at a current density of 1 A g^{-1} . At a high current density of 10 A g^{-1} , the specific capacitance value remained as high as 137 F g^{-1} . The RC constants at 1 and 10 A g^{-1} were 40.25 and 2.39 s, respectively. The energy density and power density of the device were calculated based on the discharge rates and plotted on a Ragone plot (Fig. S13†). The highest energy density, $62.61 \text{ W h kg}^{-1}$ was achieved at the power density of 2.79 kW kg^{-1} , and an energy density of $37.29 \text{ W h kg}^{-1}$ was retained even at a high power density of

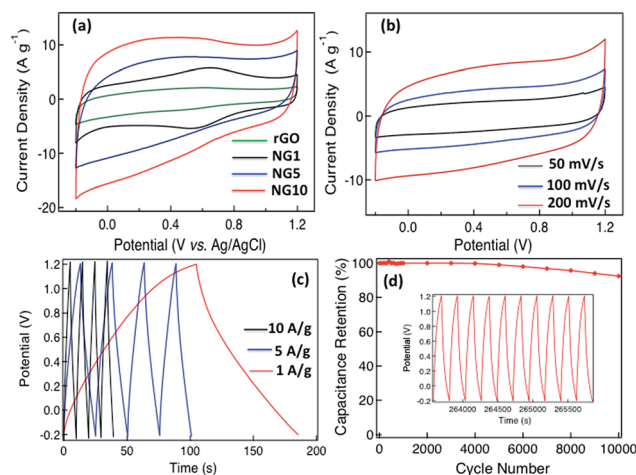


Fig. 3 (a) Cyclic voltammograms of rGO, NG1, NG5 and NG10 using a three-electrode cell at a scan rate of 100 mV s^{-1} in 0.5 M H_2SO_4 solution, (b) cyclic voltammograms of NG10 at different scan rates (50, 100 and 200 mV s^{-1}) in a two-electrode T-cell system, (c) galvanostatic charge/discharge curves for NG10 in a two-electrode T-cell system tested at current density of $1\text{--}10 \text{ A g}^{-1}$, (d) cycling stability of NG10 measured at 5 A g^{-1} over the potential range of 1.2 to -0.2 V . The inset shows the charge–discharge curves over the last ten cycles.

28.08 kW kg^{-1} ; this was superior than the values reported for chemical-doped graphene-based materials. (ESI Table S1†).

Besides the gravimetric capacitance, volumetric capacitance is also very important for fabricating devices for practical applications. The average density of the active material on the electrodes was calculated based on the thickness measured by SEM to be 0.5 g cm^{-3} , meaning that the volumetric capacitance was 115 F cm^{-3} , which is higher than that of nanostructured graphene-based supercapacitor materials such as activated graphene,⁶ graphene foam,⁶⁶ graphene hydrogel⁶⁷ and laser-scribed graphene.⁶⁸ The superior performance of NG10 could be attributed to its high contents of pyridinic, pyrrolic and graphitic N, as well as to its high surface area with large pore volume, which provided excellent electrical conductivity and pseudocapacitance with favourable electrolyte penetration. The stability of the device was also investigated over 10 000 cycles between 1.2 and -0.2 V at a high current density of 5 A g^{-1} (Fig. 3d). The material maintained almost 100% of its initial specific capacitance upto 4000 cycles and then gradually started decreasing and kept 92.4% of its initial specific capacitance after 10 000 cycles. The charge/discharge curves for the last ten cycles demonstrate the stability of the device (inset of Fig. 3d).

NG10 was also tested as an electrocatalyst for the oxygen-reduction reaction (ORR) under alkaline conditions. Thus, NG10 was deposited on a glassy carbon electrode and its cyclic voltammetry was measured in both argon- and oxygen-saturated 0.1 M $\text{KOH}(\text{aq})$ at a scan rate of 50 mV s^{-1} . As shown in Fig. 4a, a clean capacitive CV background was observed within a potential range of 1.0 to -0.05 V in the argon-saturated electrolyte. In contrast, in the oxygen-saturated electrolyte, a well-defined cathodic current was detected, and peaked at $\approx +0.6 \text{ V}$. This indicated that NG10 was an active catalyst for O_2 reduction. The electrocatalytic activities of NG10



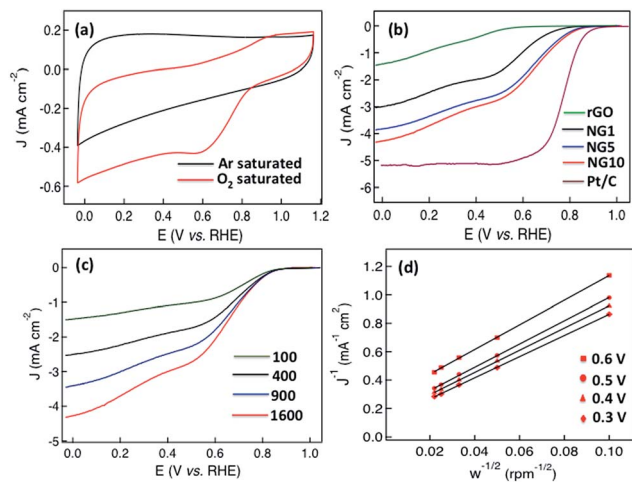


Fig. 4 (a) Cyclic voltammograms of NG10 at a scan rate of 50 mV s^{-1} in Ar-saturated (black) and O_2 -saturated (red) 0.1 M KOH(aq) solution, (b) linear sweep voltammograms (LSV) of rGO, NG1, NG5, NG10 and Pt/C at 1600 rpm and a scan rate of 10 mV s^{-1} in O_2 -saturated 0.1 M KOH(aq) , (c) RDE measurements of NG10 in O_2 -saturated 0.1 M KOH(aq) at scan rates of 10 mV s^{-1} , and (d) Koutecky–Levich plots of NG10 at electrode potentials of 0.6 , 0.5 , 0.4 and 0.3 V .

and the related materials rGO, NG1 and NG5 were further investigated using rotating disk electrodes (RDE). Fig. 4b shows the linear sweep voltammograms of rGO, NG1, NG5 and NG10, all measured at a scan rate of 10 mV s^{-1} and rotation frequency of 1600 rpm in oxygen-saturated 0.1 M KOH(aq) . To reveal the best performance in electrocatalysis, the effect of mass loading of the active material on the glassy carbon electrode has been investigated (Fig. S14†). Different masses of NG10 (4 , 8 , 12 , 16 and $20 \mu\text{g}$) have been deposited on a glassy carbon electrode having an area of 0.197 cm^2 . From the linear sweep voltammograms curves at 1600 rpm with the presence of oxygen, it was observed that the current density was increasing with the mass loading from 4 to $12 \mu\text{g}$ due to the increment of density of the active material (NG10). However, the current density was started to decrease even with the increasing mass loading from 16 to $20 \mu\text{g}$. The possible reason for this behaviour is due to the restacking of graphene layers after covering the whole surface area which hampering the permeation of electrolyte between the graphene layers causing less current density.⁶¹ The electrochemical impedance spectroscopy (EIS) study (Fig. S15†) also showed the highest conductivity as well as very small inherent resistance from 4 to $12 \mu\text{g}$ mass loading and a formation of semi Nyquist plot for $20 \mu\text{g}$ mass loading indicated higher charge transfer resistance compared to the rest of the plots in higher frequencies. Moreover, all the spectra in low frequencies exhibit a Warburg angle higher than 45° indicating the electrodes were highly controlled by ion diffusion/transport process.

The NG10 produced the most positive onset potential of $+0.75 \text{ V}$ and highest current density of -4.3 mA cm^{-2} , and further RDE measurements (Fig. 4c) were thus used to calculate, via the Koutecky–Levich equation (eqn (5)),²⁸ the number of electrons transferred to an oxygen molecule by the NG10 electrode.

$$\frac{1}{J} = \frac{\nu^{1/6}}{0.2nFC_0D_0^{2/3}\omega^{1/2}} + \frac{1}{J_K} \quad (5)$$

In the Koutecky–Levich equation, J , J_L , J_K are the measured current density, the diffusion-limiting current density and the kinetic-limiting current density, respectively; ω is the angular velocity of the disc, n is the overall number of electrons transferred in the oxygen reduction, F is the Faraday constant (96485 C mol^{-1}), C_0 is the bulk concentration of oxygen ($1.2 \times 10^{-6} \text{ mol cm}^{-3}$), D_0 is the diffusion coefficient of oxygen in 0.1 M KOH(aq) ($1.9 \times 10^{-5} \text{ cm}^2 \text{ s}^{-1}$), ν is the kinetic viscosity ($0.01 \text{ cm}^2 \text{ s}^{-1}$) and 0.2 is a constant that is valid when the rotation speed is expressed in rpm. The Koutecky–Levich plot of J^{-1} vs. $\omega^{-1/2}$ at a potential of 0.6 V on the NG10 electrode (Fig. 4d) exhibited good linearity ($R^2 = 0.998$). The average electron transfer number was 3.82 at different potentials of 0.6 , 0.5 , 0.4 and 0.3 V , which indicated that a four-electron process, with water as the product, was the preferred pathway. In addition, efficient electrocatalytic activity was observed from the cyclic voltammetry in the presence of 3 M methanol (Fig. S17†), indicating the stability of NG10 to the fuel crossover effect.⁵⁴

To further investigate the potential application of NG10 enriched with pyridinic N and graphitic N, OER activity was assessed by sweeping the RDE potential from 1.0 to 2.0 V . From the LSV curves of NG10, NG5, NG1, Ru/C and Pt/C plotted in Fig. 5a, the onset potentials were measured 1.52 , 1.56 , 1.72 , 1.46 , and 1.68 V respectively. The rGO did not show any significant response towards OER activity. The potentials for a current density of 10 mA cm^{-2} for NG10, NG5, Ru/C and Pt/C were evaluated at 1.74 , 1.76 , 1.64 and 1.88 V respectively as the required potential to oxidize water at the current density of 10 mA cm^{-2} is used to judge the OER activity. Therefore, NG10 exhibited the lowest onset potential and large current density among the studied catalysts of NG5 and Pt/C as well as close to Ru/C (one of the best electrocatalysts for OER at present). The effect of mass loading was also investigated for OER activity and the similar behaviour was obtained showing $12 \mu\text{g}$ mass loading as the optimum amount for achieving the best electrocatalytic performance (Fig. S16†). The Tafel slope for the NG10 and NG5 were about 104 and 122 mV per decade respectively, much smaller than the Pt/C (208 mV per decade), which indicated the better OER activity of NG10 (Fig. 5b). The Tafel slope of NG10 was further analyzed in two regions (low overpotential and high overpotential with the values of 106 and 99 mV per decade respectively) according to the following reports.^{74,75} The relatively higher low potential value indicated the first electron transfer reaction rate for OER.⁷⁶ The OER performance of NG10 was found to be superior to the commercial Pt/C as well as similar to transition metal oxides supported by carbon materials.^{45,69,70} Even though the exhibition of OER activities in metal-free nitrogen-doped few layers graphene is highly rare,^{30,47} the possible reason of the attribution of OER activity in NG10 even in two-dimensional nanostructure is due to the presence of high atomic percentages of N including high level of pyridinic N and graphitic N as well as the large surface area with high pore volume. The high percentage of pyridinic N and graphitic N on



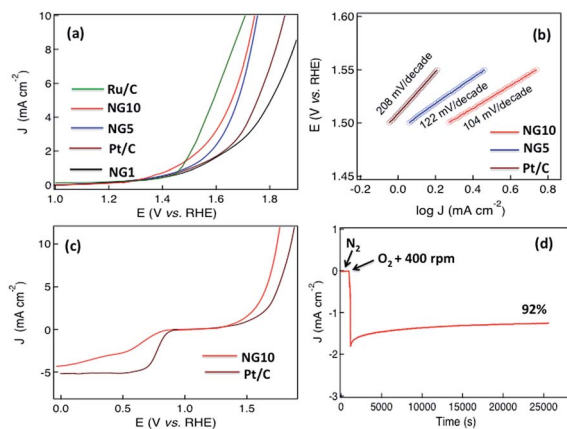


Fig. 5 (a) Linear sweep voltammograms (LSV) of Ru/C, NG10, NG5, Pt/C and NG1 at a scan rate of 10 mV s^{-1} in 0.1 M KOH(aq) ; (b) shows the Tafel plots for Pt/C, NG5 and NG10 and (c) LSV curves of NG10 and Pt/C on an RDE (1600 rpm) in O_2 -saturated 0.1 M KOH(aq) at scan rates of 10 mV s^{-1} showing the electrocatalytic activities towards both ORR and OER, (d) chronoamperometric response of ORR activity in oxygen saturated electrolyte for stability test.

the edge of the surface creates the active sites for efficient bi-catalysis reaction towards oxygen reduction and evolution. The comparison of bifunctional electrocatalysis with Pt/C measuring by LSV plots at 1600 rpm in oxygen-saturated 0.1 M KOH , (Fig. 5c) the comparable ORR with four electrons transfer pathway and better OER activity of NG10 were observed. In ORR the onset potential and current density of NG10 were lower than the commercial Pt/C as it is the state-of-the-art ORR electrode, however in OER the good catalytic activity of NG10 was reflected for its lower onset potential and higher current density than the Pt/C. Even though the OER activity of Ru/C is better than NG10; however, the ORR activity of Ru/C is poor compared to Pt/C and nitrogen-doped carbons, 48 thus not suitable as a bifunctional electrocatalyst. The electrocatalytic stability of NG10 was also investigated by continuous chronoamperometric measurement at a constant potential of $+0.6 \text{ V}$ (Fig. 5d). The nitrogen and oxygen gas was added in a flow rate of $50 \text{ cm}^3 \text{ min}^{-1}$ in 100 ml of electrolyte in a closed cell that saturated the electrolyte very quickly and helps to occur electrocatalysis instantly. The sharp increment of current density was observed after adding oxygen to the solution as well as keeping the rotation speed at 400 rpm . A nearly constant current density curve after stabilization was observed for a long cycle time and even at $25\,000 \text{ s}$ around 92% of its initial current density had been retained. Thus this active material can be a potential platform for developing bifunctional electrocatalyst for metal-air batteries and rechargeable fuel cells as well as further modification of structure and/or incorporation of nanomaterials can enhance the electrocatalytic activity superior than the commercial noble metal based catalysts.

4. Conclusions

In conclusion, we report a facile synthesis of N-doped graphenes with high surface areas as well as nitrogen atoms in all of

the reported nitrogen configurations, using uric acid as a solid source of nitrogen. The synthesized material NG10, produced from the co-pyrolysis of graphene oxide and uric acid in a 1 : 10 mass ratio, showed excellent capacitive properties in acidic media and efficient catalytic activity for oxygen reduction and evolution reaction in an alkaline solution, and could therefore be applied as a promising material in devices for energy storage and conversion. The superb stability and efficient bifunctional electrocatalytic activity could make this material an ideal candidate not only in the field of supercapacitors, fuel cell and metal-air battery electrodes, but also in the area of sensors and transistors. The incorporation of nanoparticles, polymers and other carbon-based nanostructures could further enhance the electrochemical properties for a broader range of applications.

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